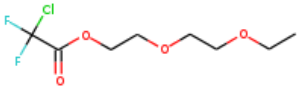
Diethylene glycol monoethyl ether, chlorodifluoroacetate

InChI: InChI=1S/C8H13ClF2O4/c1-2-13-3-4-14-5-6-15-7(12)8(9,10)11/h2-6H2,1H3
InChI Key: CECOVAGVDQHUQX-UHFFFAOYSA-N
Formula: C8H13ClF2O4
SMILES: CCOCCOCCOC(=O)C(F)(F)Cl
Molecular Weight: 246.64



Physical Properties

Property	Value	Unit	Source
$\Delta_f G°$	-826.15	kJ/mol	Joback Method
$\Delta_f H°_{gas}$	-1134.40	kJ/mol	Joback Method
$\Delta_{fus} H°$	24.58	kJ/mol	Joback Method
$\Delta_{vap} H°$	48.83	kJ/mol	Joback Method
$logP_{oct/wat}$	1.414		Crippen Method
P_c	2306.95	kPa	Joback Method
T_{boil}	536.31	K	Joback Method
T_c	709.18	K	Joback Method
T_{fus}	330.06	K	Joback Method
V_c	0.618	m^3/kg-mol	Joback Method

Temperature Dependent Properties

Property	Value	Unit	Temperature (K)	Source
$C_{p,gas}$	375.31	J/mol×K	536.31	Joback Method

Sources

Joback Method: https://en.wikipedia.org/wiki/Joback_method

NIST Webbook: http://webbook.nist.gov/cgi/inchi/InChI=1S/C8H13ClF2O4/c1-2-13-3-4-14-5-6-15-7(12)8(9,10)11/h2-6H2,1H3

Crippen Method: http://pubs.acs.org/doi/abs/10.1021/ci990307l

Legend

$C_{p,gas}$: Ideal gas heat capacity (J/mol×K).
$\Delta_f G°$: Standard Gibbs free energy of formation (kJ/mol).
$\Delta_f H°_{gas}$: Enthalpy of formation at standard conditions (kJ/mol).
$\Delta_{fus} H°$: Enthalpy of fusion at standard conditions (kJ/mol).
$\Delta_{vap} H°$: Enthalpy of vaporization at standard conditions (kJ/mol).
***log*$P_{oct/wat}$**: Octanol/Water partition coefficient .
P_c: Critical Pressure (kPa).
T_{boil}: Normal Boiling Point Temperature (K).
T_c: Critical Temperature (K).
T_{fus}: Normal melting (fusion) point (K).
V_c: Critical Volume (m^3/kg-mol).