

# N-ethyl-4-nitroaniline

<b>Other names:</b>	N-ethyl-4-nitrobenzenamine
<b>Inchi:</b>	InChI=1S/C8H10N2O2/c1-2-9-7-3-5-8(6-4-7)10(11)12/h3-6,9H,2H2,1H3
<b>InchiKey:</b>	XBNNLAWQCMDISJ-UHFFFAOYSA-N
<b>Formula:</b>	C8H10N2O2
<b>SMILES:</b>	CCNc1ccc([N+](=O)[O-])cc1
<b>Mol. weight [g/mol]:</b>	166.18

## Physical Properties

Property code	Value	Unit	Source
gf	244.20	kJ/mol	Joback Method
hf	59.32	kJ/mol	Joback Method
hfus	22.70	kJ/mol	DSC measurement and prediction of phase diagrams for binary mixtures of energetic materials' stabilizers
hvap	59.37	kJ/mol	Joback Method
log10ws	-2.55		Crippen Method
logp	2.027		Crippen Method
mcvol	127.220	ml/mol	McGowan Method
pc	3695.46	kPa	Joback Method
tb	616.11	K	Joback Method
tc	858.52	K	Joback Method
tf	368.21	K	DSC measurement and prediction of phase diagrams for binary mixtures of energetic materials' stabilizers
vc	0.492	m3/kmol	Joback Method

## Temperature Dependent Properties

Property code	Value	Unit	Temperature [K]	Source
cpg	309.89	J/molxK	616.11	Joback Method
cpg	322.05	J/molxK	656.51	Joback Method
cpg	333.31	J/molxK	696.91	Joback Method
cpg	343.71	J/molxK	737.31	Joback Method

cpg	353.30	J/mol×K	777.71	Joback Method
cpg	362.13	J/mol×K	818.12	Joback Method
cpg	370.24	J/mol×K	858.52	Joback Method

## Sources

DSC measurement and prediction of phase diagrams for binary mixtures of energetic materials' stabilizers:	<a href="https://www.doi.org/10.1016/j.tca.2013.04.021">https://www.doi.org/10.1016/j.tca.2013.04.021</a>
Joback Method:	<a href="https://en.wikipedia.org/wiki/Joback_method">https://en.wikipedia.org/wiki/Joback_method</a>
McGowan Method:	<a href="http://link.springer.com/article/10.1007/BF02311772">http://link.springer.com/article/10.1007/BF02311772</a>
Crippen Method:	<a href="http://pubs.acs.org/doi/abs/10.1021/ci9903071">http://pubs.acs.org/doi/abs/10.1021/ci9903071</a>
Crippen Method:	<a href="https://www.chemeo.com/doc/models/crippen_log10ws">https://www.chemeo.com/doc/models/crippen_log10ws</a>

## Legend

<b>cpg:</b>	Ideal gas heat capacity
<b>gf:</b>	Standard Gibbs free energy of formation
<b>hf:</b>	Enthalpy of formation at standard conditions
<b>hfus:</b>	Enthalpy of fusion at standard conditions
<b>hvap:</b>	Enthalpy of vaporization at standard conditions
<b>log10ws:</b>	Log10 of Water solubility in mol/l
<b>logp:</b>	Octanol/Water partition coefficient
<b>mcvol:</b>	McGowan's characteristic volume
<b>pc:</b>	Critical Pressure
<b>tb:</b>	Normal Boiling Point Temperature
<b>tc:</b>	Critical Temperature
<b>tf:</b>	Normal melting (fusion) point
<b>vc:</b>	Critical Volume

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