Ammonium sulfate

Other names: ammonium sulphate

InChl=1S/H8N2O4S/c1-5-7(3,4)6-2/h1-2H4

InchiKey: KTMPTTWZBSFVPI-UHFFFAOYSA-N

Formula: H8N2O4S

SMILES: NOS(=O)(=O)ON

Mol. weight [g/mol]: 132.14 CAS: 7783-20-2

Physical Properties

| Property code | Value | Unit | Source |
|---------------|----------|---------|--|
| gf | -596.52 | kJ/mol | Joback Method |
| hf | -693.54 | kJ/mol | Joback Method |
| hfus | 2.85 | kJ/mol | Phase transitions of some sulfur-containing ammonium salts |
| hvap | 60.33 | kJ/mol | Joback Method |
| log10ws | 0.34 | | Crippen Method |
| logp | -1.988 | | Crippen Method |
| mcvol | 70.650 | ml/mol | McGowan Method |
| рс | 10348.87 | kPa | Joback Method |
| tb | 437.08 | K | Joback Method |
| tc | 635.89 | K | Joback Method |
| tf | 339.30 | K | Joback Method |
| VC | 0.256 | m3/kmol | Joback Method |

Temperature Dependent Properties

| Property code | Value | Unit | Temperature [K] | Source |
|---------------|--------|---------|-----------------|---------------|
| cpg | 144.70 | J/mol×K | 437.08 | Joback Method |
| cpg | 149.29 | J/mol×K | 470.22 | Joback Method |
| cpg | 153.87 | J/mol×K | 503.35 | Joback Method |
| cpg | 158.40 | J/mol×K | 536.49 | Joback Method |
| cpg | 162.86 | J/mol×K | 569.62 | Joback Method |
| cpg | 167.20 | J/mol×K | 602.76 | Joback Method |
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Sources

organic salts: The study of phase behavior of The study of phase behavior of aqueous two-phase system containing to the study of phase system containing to the system containing to the system system containing to the system Solutions Containing Ammonium
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N-butylpyridiniumtetrafluoroborate and https://www.doi.org/10.1021/je301276s
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Phase equilibria in aqueous two-phase https://www.doi.org/10.1016/j.tca.2010.01.020 Phase equilibria in aqueous two-phase systems containing [Bmim]BF4 and Maasoname surfactors of the Sphibility of tall the form clattice of the Sphibility of tall the systems containing [Bmim]BF4 and Maasoname surfactors of the Sphibility of tall the systems of the sy http://webbook.nist.gov/cgi/cbook.cgi?ID=C7783202&Units=SI https://www.doi.org/10.1016/j.fluid.2014.09.029 https://www.doi.org/10.1021/acs.jced.9b00226 https://www.doi.org/10.1021/acs.jced.6b00844 https://www.doi.org/10.1016/j.jct.2013.10.015 https://www.doi.org/10.1021/je8008649 https://www.doi.org/10.1021/je400364b https://www.doi.org/10.1016/j.jct.2008.09.019 https://www.doi.org/10.1021/je401034k https://www.doi.org/10.1021/acs.jced.8b00740 https://www.doi.org/10.1021/acs.jced.7b01113 https://www.doi.org/10.1016/j.fluid.2017.05.002 https://en.wikipedia.org/wiki/Joback_method https://www.doi.org/10.1021/acs.jced.7b00433 https://www.doi.org/10.1016/j.jct.2019.03.003 https://www.doi.org/10.1021/je700284r http://pubs.acs.org/doi/abs/10.1021/ci990307l https://www.doi.org/10.1021/je9010129 https://www.doi.org/10.1016/j.jct.2011.04.024 https://www.doi.org/10.1021/je900504e https://www.doi.org/10.1016/j.tca.2014.08.035 https://www.chemeo.com/doc/models/crippen_log10ws https://www.doi.org/10.1021/acs.jced.7b01015 http://link.springer.com/article/10.1007/BF02311772 https://www.doi.org/10.1021/je201390r

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https://www.doi.org/10.1021/je400453b

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Potassines Sulfate, and Aluminum
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https://www.doi.org/10.1016/j.tca.2013.07.002 Measurement and correlation of phase diagram data for acetone and sulfate Experimental phase ministers of the Experimental phase of th https://www.doi.org/10.1016/j.fluid.2016.06.019 https://www.doi.org/10.1016/j.jct.2013.08.018

Legend

Ideal gas heat capacity cpg:

gf: Standard Gibbs free energy of formation hf: Enthalpy of formation at standard conditions hfus: Enthalpy of fusion at standard conditions

hvap: Enthalpy of vaporization at standard conditions

log10ws: Log10 of Water solubility in mol/l Octanol/Water partition coefficient logp: mcvol: McGowan's characteristic volume

pc: Critical Pressure

Normal Boiling Point Temperature tb:

Critical Temperature tc:

Normal melting (fusion) point tf:

Critical Volume vc:

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