

# Benzoic acid, 3,5-dinitro-, phenylmethyl ester

<b>Other names:</b>	BenzyI 3,5-dinitrobenzoate Benzoic acid, 3,5-dinitro, benzyl ester
<b>Inchi:</b>	InChI=1S/C14H10N2O6/c17-14(22-9-10-4-2-1-3-5-10)11-6-12(15(18)19)8-13(7-11)16(20)
<b>InchiKey:</b>	ULVFRRGXEDYWEM-UHFFFAOYSA-N
<b>Formula:</b>	C14H10N2O6
<b>SMILES:</b>	O=C(OCc1ccccc1)c1cc([N+](=O)[O-])cc([N+](=O)[O-])c1
<b>Mol. weight [g/mol]:</b>	302.24
<b>CAS:</b>	10478-07-6

## Physical Properties

Property code	Value	Unit	Source
gf	109.74	kJ/mol	Joback Method
hf	-148.49	kJ/mol	Joback Method
hfus	44.83	kJ/mol	Joback Method
hvap	94.97	kJ/mol	Joback Method
log10ws	-5.13		Crippen Method
logp	2.860		Crippen Method
mcvol	202.880	ml/mol	McGowan Method
pc	2890.51	kPa	Joback Method
rinpol	2325.00		NIST Webbook
rinpol	2368.00		NIST Webbook
rinpol	2358.00		NIST Webbook
rinpol	2342.00		NIST Webbook
rinpol	2358.00		NIST Webbook
rinpol	2325.00		NIST Webbook
rinpol	2325.00		NIST Webbook
rinpol	2377.00		NIST Webbook
rinpol	2357.00		NIST Webbook
tb	963.01	K	Joback Method
tc	1238.81	K	Joback Method
tf	684.80	K	Joback Method
vc	0.791	m3/kmol	Joback Method

# Temperature Dependent Properties

Property code	Value	Unit	Temperature [K]	Source
cpg	591.56	J/mol×K	963.01	Joback Method
cpg	600.00	J/mol×K	1008.98	Joback Method
cpg	607.20	J/mol×K	1054.94	Joback Method
cpg	613.25	J/mol×K	1100.91	Joback Method
cpg	618.21	J/mol×K	1146.88	Joback Method
cpg	622.17	J/mol×K	1192.85	Joback Method
cpg	625.19	J/mol×K	1238.81	Joback Method

## Sources

<b>NIST Webbook:</b>	<a href="http://webbook.nist.gov/cgi/cbook.cgi?ID=C10478076&amp;Units=SI">http://webbook.nist.gov/cgi/cbook.cgi?ID=C10478076&amp;Units=SI</a>
<b>Crippen Method:</b>	<a href="http://pubs.acs.org/doi/abs/10.1021/ci990307l">http://pubs.acs.org/doi/abs/10.1021/ci990307l</a>
<b>Crippen Method:</b>	<a href="https://www.chemeo.com/doc/models/crippen_log10ws">https://www.chemeo.com/doc/models/crippen_log10ws</a>
<b>Joback Method:</b>	<a href="https://en.wikipedia.org/wiki/Joback_method">https://en.wikipedia.org/wiki/Joback_method</a>
<b>McGowan Method:</b>	<a href="http://link.springer.com/article/10.1007/BF02311772">http://link.springer.com/article/10.1007/BF02311772</a>

## Legend

<b>cpg:</b>	Ideal gas heat capacity
<b>gf:</b>	Standard Gibbs free energy of formation
<b>hf:</b>	Enthalpy of formation at standard conditions
<b>hfus:</b>	Enthalpy of fusion at standard conditions
<b>hvap:</b>	Enthalpy of vaporization at standard conditions
<b>log10ws:</b>	Log10 of Water solubility in mol/l
<b>logp:</b>	Octanol/Water partition coefficient
<b>mcvol:</b>	McGowan's characteristic volume
<b>pc:</b>	Critical Pressure
<b>rinpol:</b>	Non-polar retention indices
<b>tb:</b>	Normal Boiling Point Temperature
<b>tc:</b>	Critical Temperature
<b>tf:</b>	Normal melting (fusion) point
<b>vc:</b>	Critical Volume

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