

# Urea, N-(2-methylcyclohexyl)-N'-phenyl-

**Other names:**

1-(2-Methylcyclohexyl)-3-phenylurea  
Du Pont 1318  
Du pont herbicide 1,318  
H 1318  
N-Phenyl-N'-(2-methylcyclohexyl)urea  
NSC 131951  
Siduron  
Trey  
Tupersan  
Urea, 1-(2-methylcyclohexyl)-3-phenyl-

**Inchi:** InChI=1S/C14H20N2O/c1-11-7-5-6-10-13(11)16-14(17)15-12-8-3-2-4-9-12/h2-4,8-9,11,1**InchiKey:** JXVIIQLNUPXOII-UHFFFAOYSA-N**Formula:** C14H20N2O**SMILES:** CC1CCCCC1NC(=O)Nc1ccccc1**Mol. weight [g/mol]:** 232.32**CAS:** 1982-49-6

## Physical Properties

Property code	Value	Unit	Source
gf	246.01	kJ/mol	Joback Method
hf	-67.42	kJ/mol	Joback Method
hfus	30.76	kJ/mol	Joback Method
hvap	68.77	kJ/mol	Joback Method
log10ws	-4.11		Aqueous Solubility Prediction Method
log10ws	-4.11		Estimated Solubility Method
logp	3.387		Crippen Method
mvol	195.030	ml/mol	McGowan Method
pc	2576.72	kPa	Joback Method
tb	715.49	K	Joback Method
tc	949.61	K	Joback Method
tf	432.35	K	Joback Method
vc	0.720	m <sup>3</sup> /kmol	Joback Method

# Temperature Dependent Properties

Property code	Value	Unit	Temperature [K]	Source
cpg	573.69	J/molxK	715.49	Joback Method
cpg	592.18	J/molxK	754.51	Joback Method
cpg	609.22	J/molxK	793.53	Joback Method
cpg	624.86	J/molxK	832.55	Joback Method
cpg	639.16	J/molxK	871.57	Joback Method
cpg	652.20	J/molxK	910.59	Joback Method
cpg	664.03	J/molxK	949.61	Joback Method

## Sources

<b>Crippen Method:</b>	<a href="http://pubs.acs.org/doi/abs/10.1021/ci9903071">http://pubs.acs.org/doi/abs/10.1021/ci9903071</a>
<b>Joback Method:</b>	<a href="https://en.wikipedia.org/wiki/Joback_method">https://en.wikipedia.org/wiki/Joback_method</a>
<b>Aqueous Solubility Prediction Method:</b>	<a href="http://onschallenge.wikispaces.com/file/view/AqueousDataset002.xlsx/351826032/AqueousDa">http://onschallenge.wikispaces.com/file/view/AqueousDataset002.xlsx/351826032/AqueousDa</a>
<b>Estimated Solubility Method:</b>	<a href="http://pubs.acs.org/doi/suppl/10.1021/ci034243x/suppl_file/ci034243xsi20040112_053635.txt">http://pubs.acs.org/doi/suppl/10.1021/ci034243x/suppl_file/ci034243xsi20040112_053635.txt</a>
<b>McGowan Method:</b>	<a href="http://link.springer.com/article/10.1007/BF02311772">http://link.springer.com/article/10.1007/BF02311772</a>
<b>NIST Webbook:</b>	<a href="http://webbook.nist.gov/cgi/cbook.cgi?ID=C1982496&amp;Units=SI">http://webbook.nist.gov/cgi/cbook.cgi?ID=C1982496&amp;Units=SI</a>

## Legend

<b>cpg:</b>	Ideal gas heat capacity
<b>gf:</b>	Standard Gibbs free energy of formation
<b>hf:</b>	Enthalpy of formation at standard conditions
<b>hfus:</b>	Enthalpy of fusion at standard conditions
<b>hvac:</b>	Enthalpy of vaporization at standard conditions
<b>log10ws:</b>	Log10 of Water solubility in mol/l
<b>logp:</b>	Octanol/Water partition coefficient
<b>mccvol:</b>	McGowan's characteristic volume
<b>pc:</b>	Critical Pressure
<b>tb:</b>	Normal Boiling Point Temperature
<b>tc:</b>	Critical Temperature
<b>tf:</b>	Normal melting (fusion) point
<b>vc:</b>	Critical Volume

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