

# Succinic acid, 2,2-dichloroethyl 2,3-dimethylphenyl ester

<b>Inchi:</b>	InChI=1S/C14H16Cl2O4/c1-9-4-3-5-11(10(9)2)20-14(18)7-6-13(17)19-8-12(15)16/h3-5,1
<b>InchiKey:</b>	IWBYHJNVSMXTJ-UHFFFAOYSA-N
<b>Formula:</b>	C14H16Cl2O4
<b>SMILES:</b>	<chem>Cc1cccc(OC(=O)CCC(=O)OCC(Cl)Cl)c1C</chem>
<b>Mol. weight [g/mol]:</b>	319.18

## Physical Properties

Property code	Value	Unit	Source
gf	-333.99	kJ/mol	Joback Method
hf	-645.06	kJ/mol	Joback Method
hfus	35.72	kJ/mol	Joback Method
hvap	77.05	kJ/mol	Joback Method
log10ws	-4.19		Crippen Method
logp	3.336		Crippen Method
mvol	223.720	ml/mol	McGowan Method
pc	2001.91	kPa	Joback Method
rinpol	2255.00		NIST Webbook
rinpol	2255.00		NIST Webbook
tb	783.36	K	Joback Method
tc	1000.10	K	Joback Method
tf	488.16	K	Joback Method
vc	0.852	m <sup>3</sup> /kmol	Joback Method

## Temperature Dependent Properties

Property code	Value	Unit	Temperature [K]	Source
cpg	589.53	J/mol×K	783.36	Joback Method
cpg	601.85	J/mol×K	819.48	Joback Method
cpg	613.20	J/mol×K	855.61	Joback Method
cpg	623.60	J/mol×K	891.73	Joback Method
cpg	633.05	J/mol×K	927.85	Joback Method
cpg	641.56	J/mol×K	963.98	Joback Method
cpg	649.13	J/mol×K	1000.10	Joback Method
dvisc	0.0006953	Paxs	488.16	Joback Method

dvisc	0.0004209	Paxs	537.36	Joback Method
dvisc	0.0002771	Paxs	586.56	Joback Method
dvisc	0.0001947	Paxs	635.76	Joback Method
dvisc	0.0001439	Paxs	684.96	Joback Method
dvisc	0.0001107	Paxs	734.16	Joback Method
dvisc	0.0000881	Paxs	783.36	Joback Method

## Sources

<b>NIST Webbook:</b>	<a href="http://webbook.nist.gov/cgi/cbook.cgi?ID=U390019&amp;Units=SI">http://webbook.nist.gov/cgi/cbook.cgi?ID=U390019&amp;Units=SI</a>
<b>Crippen Method:</b>	<a href="http://pubs.acs.org/doi/abs/10.1021/ci9903071">http://pubs.acs.org/doi/abs/10.1021/ci9903071</a>
<b>Crippen Method:</b>	<a href="https://www.chemeo.com/doc/models/crippen_log10ws">https://www.chemeo.com/doc/models/crippen_log10ws</a>
<b>Joback Method:</b>	<a href="https://en.wikipedia.org/wiki/Joback_method">https://en.wikipedia.org/wiki/Joback_method</a>
<b>McGowan Method:</b>	<a href="http://link.springer.com/article/10.1007/BF02311772">http://link.springer.com/article/10.1007/BF02311772</a>

## Legend

<b>cp<sub>g</sub>:</b>	Ideal gas heat capacity
<b>dvisc:</b>	Dynamic viscosity
<b>gf:</b>	Standard Gibbs free energy of formation
<b>hf:</b>	Enthalpy of formation at standard conditions
<b>hfus:</b>	Enthalpy of fusion at standard conditions
<b>hvap:</b>	Enthalpy of vaporization at standard conditions
<b>log<sub>10</sub>ws:</b>	Log <sub>10</sub> of Water solubility in mol/l
<b>logp:</b>	Octanol/Water partition coefficient
<b>m<sub>c</sub>vol:</b>	McGowan's characteristic volume
<b>pc:</b>	Critical Pressure
<b>rinpol:</b>	Non-polar retention indices
<b>tb:</b>	Normal Boiling Point Temperature
<b>tc:</b>	Critical Temperature
<b>tf:</b>	Normal melting (fusion) point
<b>vc:</b>	Critical Volume

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