

# Triethylamine

<b>Other names:</b>	Benzenemethanamine, «alpha», «alpha»-diphenyl- Methylamine, 1,1,1-triphenyl- Triphenylmethanamine
<b>Inchi:</b>	InChI=1S/C19H17N/c20-19(16-10-4-1-5-11-16,17-12-6-2-7-13-17)18-14-8-3-9-15-18/h1-
<b>InchiKey:</b>	BZVJOYBTLHNRDW-UHFFFAOYSA-N
<b>Formula:</b>	C19H17N
<b>SMILES:</b>	NC(c1ccccc1)(c1ccccc1)c1ccccc1
<b>Mol. weight [g/mol]:</b>	259.35
<b>CAS:</b>	5824-40-8

## Physical Properties

Property code	Value	Unit	Source
gf	515.62	kJ/mol	Joback Method
hf	299.14	kJ/mol	Joback Method
hfus	24.87	kJ/mol	Joback Method
hvap	74.06	kJ/mol	Joback Method
log10ws	-4.89		Crippen Method
logp	3.937		Crippen Method
mcpvol	217.270	ml/mol	McGowan Method
pc	2568.89	kPa	Joback Method
tb	783.46	K	Joback Method
tc	1065.59	K	Joback Method
tf	468.83	K	Joback Method
vc	0.793	m <sup>3</sup> /kmol	Joback Method

## Temperature Dependent Properties

Property code	Value	Unit	Temperature [K]	Source
cpg	615.28	J/mol×K	783.46	Joback Method
cpg	632.36	J/mol×K	830.48	Joback Method
cpg	647.72	J/mol×K	877.50	Joback Method
cpg	661.61	J/mol×K	924.52	Joback Method
cpg	674.27	J/mol×K	971.54	Joback Method
cpg	685.91	J/mol×K	1018.56	Joback Method

cpg

696.79

J/mol×K

1065.59

Joback Method

## Pressure Dependent Properties

Property code	Value	Unit	Pressure [kPa]	Source
tbrp	496.20	K	1.90	NIST Webbook

## Sources

McGowan Method:	<a href="http://link.springer.com/article/10.1007/BF02311772">http://link.springer.com/article/10.1007/BF02311772</a>
NIST Webbook:	<a href="http://webbook.nist.gov/cgi/cbook.cgi?ID=C5824408&amp;Units=SI">http://webbook.nist.gov/cgi/cbook.cgi?ID=C5824408&amp;Units=SI</a>
Crippen Method:	<a href="http://pubs.acs.org/doi/abs/10.1021/ci9903071">http://pubs.acs.org/doi/abs/10.1021/ci9903071</a>
Crippen Method:	<a href="https://www.chemeo.com/doc/models/crippen_log10ws">https://www.chemeo.com/doc/models/crippen_log10ws</a>
Joback Method:	<a href="https://en.wikipedia.org/wiki/Joback_method">https://en.wikipedia.org/wiki/Joback_method</a>

## Legend

<b>cpg:</b>	Ideal gas heat capacity
<b>gf:</b>	Standard Gibbs free energy of formation
<b>hf:</b>	Enthalpy of formation at standard conditions
<b>hfus:</b>	Enthalpy of fusion at standard conditions
<b>hvap:</b>	Enthalpy of vaporization at standard conditions
<b>log10ws:</b>	Log10 of Water solubility in mol/l
<b>logp:</b>	Octanol/Water partition coefficient
<b>mcvol:</b>	McGowan's characteristic volume
<b>pc:</b>	Critical Pressure
<b>tb:</b>	Normal Boiling Point Temperature
<b>tbrp:</b>	Boiling point at reduced pressure
<b>tc:</b>	Critical Temperature
<b>tf:</b>	Normal melting (fusion) point
<b>vc:</b>	Critical Volume

Latest version available from:

<https://www.chemeo.com/cid/14-504-6/Triethylamine.pdf>

Generated by Cheméo on 2024-04-27 23:10:31.51294663 +0000 UTC m=+16548680.433523942.

Cheméo (<https://www.chemeo.com>) is the biggest free database of chemical and physical data for the process industry.