

1-Propanol

Other names:	1-Hydroxypropane 1-Propyl alcohol Alcohol, propyl Alcool propilico Alcool propylique Ethylcarbinol N-PROPYL ALCOHOL NSC 30300 Optal Osmosol extra PROPYL ALCOHOL Propan-1-ol Propanol Propanol-1 Propanole Propanolen Propanoli Propylan-propyl alcohol Propylic alcohol Propylowy alkohol UN 1274 n-C ₃ H ₇ OH n-Propan-1-ol n-Propanol n-Propyl alkohol
Inchi:	InChI=1S/C ₃ H ₈ O/c1-2-3-4/h4H,2-3H ₂ ,1H ₃
InchiKey:	BDERNNFJNOPAEC-UHFFFAOYSA-N
Formula:	C ₃ H ₈ O
SMILES:	CCCO
Mol. weight [g/mol]:	60.09
CAS:	71-23-8

Physical Properties

Property code	Value	Unit	Source
af	0.6230		KDB
affp	786.50	kJ/mol	NIST Webbook

aigt	644.26	K	KDB
basg	756.10	kJ/mol	NIST Webbook
chl	-2019.40 ± 0.30	kJ/mol	NIST Webbook
chl	-2017.70 ± 1.00	kJ/mol	NIST Webbook
chl	-2021.40 ± 0.75	kJ/mol	NIST Webbook
chl	-2032.59	kJ/mol	NIST Webbook
chl	-2021.31 ± 0.25	kJ/mol	NIST Webbook
cpl	143.41	J/molxK	THERMODYNAMICS OF MIXTURES CONTAINING AMINES. XIV. CpEm OF BENZYLAMINE WITH HEPTANE AT 293.15 K OR WITH METHANOL, 1-PROPANOL OR 1-PENTANOL AT (293.15-308.15) K
dm	1.70	debye	KDB
dvisc	0.0019544	Paxs	Densities and Viscosities of Binary Liquid Mixtures of Trichloroethylene and Tetrachloroethylene with Some Polar and Nonpolar Solvents
fil	2.10	% in Air	KDB
flu	13.50	% in Air	KDB
fpc	300.37	K	KDB
fpo	298.15	K	KDB
gf	-161.90	kJ/mol	KDB
gyrad	2.7360		KDB
hf	-257.20	kJ/mol	NIST Webbook
hf	-255.20 ± 0.30	kJ/mol	NIST Webbook
hf	-256.60	kJ/mol	KDB
hf	-257.30 ± 0.40	kJ/mol	NIST Webbook
hf	-255.60 ± 1.30	kJ/mol	NIST Webbook
hf	-254.70 ± 4.40	kJ/mol	NIST Webbook
hf	-258.80 ± 1.10	kJ/mol	NIST Webbook
hf	-255.10	kJ/mol	NIST Webbook
hfl	-306.30 ± 1.00	kJ/mol	NIST Webbook
hfl	-302.54 ± 0.25	kJ/mol	NIST Webbook
hfl	-303.00 ± 1.30	kJ/mol	NIST Webbook
hfl	-304.60 ± 0.40	kJ/mol	NIST Webbook
hfl	-302.50 ± 4.20	kJ/mol	NIST Webbook
hfus	5.40	kJ/mol	Heat Capacities and Derived Thermodynamic Functions of 1-Propanol between 10 K and 350 K and of 1-Pentanol between 85 K and 370 K
hvap	38.95	kJ/mol	Joback Method
ie	10.22 ± 0.06	eV	NIST Webbook

ie	10.22 ± 0.07	eV	NIST Webbook
ie	10.00	eV	NIST Webbook
ie	10.15 ± 0.03	eV	NIST Webbook
ie	10.16 ± 0.03	eV	NIST Webbook
ie	10.32 ± 0.02	eV	NIST Webbook
ie	10.25	eV	NIST Webbook
ie	10.22 ± 0.04	eV	NIST Webbook
ie	10.48	eV	NIST Webbook
ie	10.20	eV	NIST Webbook
ie	10.52 ± 0.03	eV	NIST Webbook
ie	10.51	eV	NIST Webbook
ie	10.49	eV	NIST Webbook
ie	10.51	eV	NIST Webbook
log10ws	0.62		Estimated Solubility Method
log10ws	0.62		Aqueous Solubility Prediction Method
logp	0.389		Crippen Method
mcvol	59.000	ml/mol	McGowan Method
nfpaf	%!d(float64=3)		KDB
nfpah	%!d(float64=1)		KDB
pc	5170.00	kPa	NIST Webbook
pc	5170.00	kPa	NIST Webbook
pc	5082.00	kPa	NIST Webbook
pc	5218.00	kPa	NIST Webbook
pc	5082.00	kPa	NIST Webbook
pc	5168.00 ± 20.00	kPa	NIST Webbook
pc	5168.00 ± 20.00	kPa	NIST Webbook
pc	5397.00	kPa	NIST Webbook
pc	5155.00	kPa	NIST Webbook
pc	5182.00 ± 10.00	kPa	NIST Webbook
pc	5169.00	kPa	KDB
pc	5170.00 ± 20.00	kPa	NIST Webbook
pc	5170.00	kPa	NIST Webbook
rinpol	524.00		NIST Webbook
rinpol	551.40		NIST Webbook
rinpol	544.60		NIST Webbook
rinpol	542.30		NIST Webbook
rinpol	540.50		NIST Webbook
rinpol	541.50		NIST Webbook
rinpol	556.50		NIST Webbook
rinpol	553.50		NIST Webbook
rinpol	551.40		NIST Webbook
rinpol	565.00		NIST Webbook
rinpol	575.00		NIST Webbook

rinpol	546.00	NIST Webbook
rinpol	555.00	NIST Webbook
rinpol	530.00	NIST Webbook
rinpol	539.90	NIST Webbook
rinpol	530.00	NIST Webbook
rinpol	544.00	NIST Webbook
rinpol	525.00	NIST Webbook
rinpol	562.00	NIST Webbook
rinpol	520.00	NIST Webbook
rinpol	556.50	NIST Webbook
rinpol	523.00	NIST Webbook
rinpol	574.00	NIST Webbook
rinpol	561.00	NIST Webbook
rinpol	555.00	NIST Webbook
rinpol	595.00	NIST Webbook
rinpol	551.00	NIST Webbook
rinpol	568.00	NIST Webbook
rinpol	560.00	NIST Webbook
rinpol	568.00	NIST Webbook
rinpol	568.00	NIST Webbook
rinpol	541.50	NIST Webbook
rinpol	521.00	NIST Webbook
rinpol	551.00	NIST Webbook
rinpol	555.00	NIST Webbook
rinpol	557.00	NIST Webbook
rinpol	556.00	NIST Webbook
rinpol	537.00	NIST Webbook
rinpol	540.20	NIST Webbook
rinpol	540.20	NIST Webbook
rinpol	546.00	NIST Webbook
rinpol	552.00	NIST Webbook
rinpol	532.00	NIST Webbook
rinpol	552.00	NIST Webbook
rinpol	552.00	NIST Webbook
rinpol	536.00	NIST Webbook
rinpol	533.00	NIST Webbook
rinpol	530.00	NIST Webbook
rinpol	537.00	NIST Webbook
rinpol	544.00	NIST Webbook
rinpol	535.00	NIST Webbook
rinpol	507.00	NIST Webbook
rinpol	557.00	NIST Webbook
rinpol	561.00	NIST Webbook
rinpol	549.00	NIST Webbook

rinpol	553.00	NIST Webbook
rinpol	536.00	NIST Webbook
rinpol	559.00	NIST Webbook
rinpol	559.00	NIST Webbook
rinpol	559.00	NIST Webbook
rinpol	548.00	NIST Webbook
rinpol	557.90	NIST Webbook
rinpol	568.00	NIST Webbook
rinpol	568.00	NIST Webbook
rinpol	568.00	NIST Webbook
rinpol	554.00	NIST Webbook
rinpol	553.00	NIST Webbook
rinpol	560.00	NIST Webbook
rinpol	539.00	NIST Webbook
rinpol	538.64	NIST Webbook
rinpol	521.00	NIST Webbook
rinpol	535.00	NIST Webbook
rinpol	548.00	NIST Webbook
rinpol	561.00	NIST Webbook
rinpol	555.00	NIST Webbook
rinpol	568.00	NIST Webbook
rinpol	547.00	NIST Webbook
rinpol	544.00	NIST Webbook
rinpol	548.00	NIST Webbook
rinpol	544.00	NIST Webbook
rinpol	557.00	NIST Webbook
rinpol	558.00	NIST Webbook
rinpol	551.00	NIST Webbook
rinpol	544.00	NIST Webbook
rinpol	548.00	NIST Webbook
rinpol	548.00	NIST Webbook
rinpol	535.00	NIST Webbook
rinpol	570.00	NIST Webbook
rinpol	560.00	NIST Webbook
rinpol	539.00	NIST Webbook
rinpol	560.00	NIST Webbook
rinpol	562.00	NIST Webbook
rinpol	543.00	NIST Webbook
rinpol	546.00	NIST Webbook
rinpol	546.00	NIST Webbook
rinpol	539.00	NIST Webbook
rinpol	571.00	NIST Webbook
rinpol	534.00	NIST Webbook
rinpol	555.00	NIST Webbook

rinpol	535.00	NIST Webbook
rinpol	581.00	NIST Webbook
rinpol	571.00	NIST Webbook
rinpol	572.00	NIST Webbook
rinpol	540.50	NIST Webbook
rinpol	543.70	NIST Webbook
rinpol	540.50	NIST Webbook
rinpol	544.60	NIST Webbook
rinpol	553.50	NIST Webbook
rinpol	555.00	NIST Webbook
rinpol	525.00	NIST Webbook
rinpol	561.00	NIST Webbook
rinpol	557.00	NIST Webbook
rinpol	559.00	NIST Webbook
rinpol	568.00	NIST Webbook
rinpol	560.00	NIST Webbook
rinpol	548.00	NIST Webbook
rinpol	542.30	NIST Webbook
rinpol	544.60	NIST Webbook
rinpol	521.00	NIST Webbook
rinpol	536.90	NIST Webbook
rinpol	543.70	NIST Webbook
rinpol	549.70	NIST Webbook
rinpol	558.20	NIST Webbook
rinpol	559.30	NIST Webbook
rinpol	561.40	NIST Webbook
rinpol	561.20	NIST Webbook
rinpol	553.60	NIST Webbook
rinpol	561.20	NIST Webbook
rinpol	547.00	NIST Webbook
ripol	1040.00	NIST Webbook
ripol	1030.00	NIST Webbook
ripol	1051.00	NIST Webbook
ripol	1046.00	NIST Webbook
ripol	1036.00	NIST Webbook
ripol	1038.00	NIST Webbook
ripol	1047.00	NIST Webbook
ripol	1002.00	NIST Webbook
ripol	1037.00	NIST Webbook
ripol	1025.00	NIST Webbook
ripol	1049.00	NIST Webbook
ripol	1014.00	NIST Webbook
ripol	1033.00	NIST Webbook
ripol	1026.00	NIST Webbook

ripol	1045.00	NIST Webbook
ripol	1064.00	NIST Webbook
ripol	1033.00	NIST Webbook
ripol	1072.00	NIST Webbook
ripol	1052.00	NIST Webbook
ripol	1031.00	NIST Webbook
ripol	1046.00	NIST Webbook
ripol	1046.90	NIST Webbook
ripol	1002.00	NIST Webbook
ripol	1031.00	NIST Webbook
ripol	1027.00	NIST Webbook
ripol	1025.00	NIST Webbook
ripol	1042.00	NIST Webbook
ripol	1047.00	NIST Webbook
ripol	1042.00	NIST Webbook
ripol	1042.00	NIST Webbook
ripol	1036.00	NIST Webbook
ripol	1042.00	NIST Webbook
ripol	1038.00	NIST Webbook
ripol	1036.00	NIST Webbook
ripol	1044.00	NIST Webbook
ripol	1035.00	NIST Webbook
ripol	1032.00	NIST Webbook
ripol	1036.00	NIST Webbook
ripol	1036.00	NIST Webbook
ripol	1036.00	NIST Webbook
ripol	1046.00	NIST Webbook
ripol	1052.00	NIST Webbook
ripol	1051.00	NIST Webbook
ripol	1036.00	NIST Webbook
ripol	1036.00	NIST Webbook
ripol	1086.00	NIST Webbook
ripol	1036.00	NIST Webbook
ripol	1030.00	NIST Webbook
ripol	992.00	NIST Webbook
ripol	1045.00	NIST Webbook
ripol	1038.00	NIST Webbook
ripol	1036.00	NIST Webbook
ripol	1040.00	NIST Webbook
ripol	1042.00	NIST Webbook
ripol	996.00	NIST Webbook
ripol	990.00	NIST Webbook
ripol	996.00	NIST Webbook
ripol	990.00	NIST Webbook

ripol	1030.00	NIST Webbook
ripol	1002.00	NIST Webbook
ripol	1022.00	NIST Webbook
ripol	1042.00	NIST Webbook
ripol	1002.00	NIST Webbook
ripol	1049.00	NIST Webbook
ripol	1046.00	NIST Webbook
ripol	1069.00	NIST Webbook
ripol	1002.00	NIST Webbook
ripol	1032.00	NIST Webbook
ripol	1038.00	NIST Webbook
ripol	1006.00	NIST Webbook
ripol	1045.00	NIST Webbook
ripol	1000.00	NIST Webbook
ripol	1038.00	NIST Webbook
ripol	1040.00	NIST Webbook
ripol	1025.00	NIST Webbook
ripol	1029.00	NIST Webbook
ripol	1037.00	NIST Webbook
ripol	1043.00	NIST Webbook
ripol	1030.00	NIST Webbook
ripol	1030.00	NIST Webbook
ripol	1002.00	NIST Webbook
ripol	1035.00	NIST Webbook
ripol	1025.00	NIST Webbook
ripol	1034.00	NIST Webbook
ripol	1045.00	NIST Webbook
ripol	1037.00	NIST Webbook
ripol	1040.00	NIST Webbook
ripol	1023.00	NIST Webbook
ripol	1028.00	NIST Webbook
ripol	1033.00	NIST Webbook
ripol	1031.00	NIST Webbook
ripol	1034.00	NIST Webbook
ripol	1032.00	NIST Webbook
ripol	1041.00	NIST Webbook
ripol	1029.00	NIST Webbook
ripol	1037.00	NIST Webbook
ripol	1052.00	NIST Webbook
ripol	1052.00	NIST Webbook
ripol	1051.00	NIST Webbook
ripol	1052.00	NIST Webbook
ripol	1052.00	NIST Webbook
ripol	1031.00	NIST Webbook

ripol	1025.00	NIST Webbook
ripol	1025.00	NIST Webbook
ripol	1049.00	NIST Webbook
ripol	1064.00	NIST Webbook
ripol	1038.00	NIST Webbook
ripol	1038.00	NIST Webbook
ripol	1041.00	NIST Webbook
ripol	1073.00	NIST Webbook
ripol	1045.00	NIST Webbook
ripol	1040.00	NIST Webbook
ripol	1052.00	NIST Webbook
ripol	1041.00	NIST Webbook
ripol	1045.00	NIST Webbook
ripol	1026.00	NIST Webbook
ripol	1038.00	NIST Webbook
ripol	1015.00	NIST Webbook
ripol	1025.00	NIST Webbook
ripol	1049.00	NIST Webbook
ripol	1049.00	NIST Webbook
ripol	1030.00	NIST Webbook
ripol	1061.00	NIST Webbook
ripol	1060.00	NIST Webbook
ripol	1037.00	NIST Webbook
ripol	1025.00	NIST Webbook
ripol	1002.00	NIST Webbook
ripol	1005.00	NIST Webbook
ripol	1043.00	NIST Webbook
ripol	1030.00	NIST Webbook
ripol	1037.00	NIST Webbook
ripol	993.00	NIST Webbook
ripol	1014.00	NIST Webbook
ripol	1021.00	NIST Webbook
ripol	1056.00	NIST Webbook
ripol	1026.00	NIST Webbook
ripol	1036.00	NIST Webbook
ripol	1046.00	NIST Webbook
ripol	1030.00	NIST Webbook
ripol	1031.00	NIST Webbook
ripol	993.00	NIST Webbook
ripol	1035.00	NIST Webbook
ripol	1040.00	NIST Webbook
ripol	997.00	NIST Webbook
ripol	1000.00	NIST Webbook
ripol	988.00	NIST Webbook

ripol	1072.00		NIST Webbook
ripol	1054.90		NIST Webbook
ripol	1051.70		NIST Webbook
ripol	1046.90		NIST Webbook
ripol	1037.60		NIST Webbook
ripol	1024.00		NIST Webbook
ripol	1031.00		NIST Webbook
ripol	1037.00		NIST Webbook
ripol	1056.00		NIST Webbook
sg	322.49	J/molxK	NIST Webbook
sl	192.80	J/molxK	NIST Webbook
sl	214.20	J/molxK	NIST Webbook
ss	112.70	J/molxK	NIST Webbook
tb	370.40 ± 0.50	K	NIST Webbook
tb	370.37 ± 0.10	K	NIST Webbook
tb	369.60 ± 1.00	K	NIST Webbook
tb	370.30 ± 0.50	K	NIST Webbook
tb	370.60 ± 0.30	K	NIST Webbook
tb	370.41 ± 0.20	K	NIST Webbook
tb	370.40 ± 0.30	K	NIST Webbook
tb	370.40 ± 0.30	K	NIST Webbook
tb	370.50 ± 0.50	K	NIST Webbook
tb	370.00 ± 1.00	K	NIST Webbook
tb	368.70 ± 0.50	K	NIST Webbook
tb	370.34 ± 0.20	K	NIST Webbook
tb	369.80 ± 0.50	K	NIST Webbook
tb	370.40 ± 0.50	K	NIST Webbook
tb	370.40 ± 0.10	K	NIST Webbook
tb	369.20 ± 0.10	K	NIST Webbook
tb	369.30 ± 0.50	K	NIST Webbook
tb	369.90 ± 2.00	K	NIST Webbook
tb	369.75 ± 0.50	K	NIST Webbook
tb	371.00 ± 1.00	K	NIST Webbook
tb	371.00 ± 0.50	K	NIST Webbook
tb	370.60 ± 0.50	K	NIST Webbook
tb	370.56 ± 0.30	K	NIST Webbook
tb	363.00 ± 4.00	K	NIST Webbook
tb	370.35 ± 0.05	K	NIST Webbook
tb	370.33 ± 0.05	K	NIST Webbook
tb	370.34 ± 0.05	K	NIST Webbook
tb	371.25 ± 0.20	K	NIST Webbook
tb	370.30 ± 0.20	K	NIST Webbook
tb	370.13 ± 0.20	K	NIST Webbook
tb	370.90 ± 0.20	K	NIST Webbook

tb	370.60 ± 0.50	K	NIST Webbook
tb	370.34 ± 0.10	K	NIST Webbook
tb	370.20 ± 0.50	K	NIST Webbook
tb	370.50 ± 0.50	K	NIST Webbook
tb	371.00 ± 0.30	K	NIST Webbook
tb	370.40 ± 0.30	K	NIST Webbook
tb	370.70 ± 1.00	K	NIST Webbook
tb	370.40 ± 0.50	K	NIST Webbook
tb	370.31 ± 0.30	K	NIST Webbook
tb	370.70 ± 1.00	K	NIST Webbook
tb	370.40 ± 0.50	K	NIST Webbook
tb	370.40 ± 0.50	K	NIST Webbook
tb	370.40 ± 0.50	K	NIST Webbook
tb	370.70 ± 0.50	K	NIST Webbook
tb	370.50 ± 0.50	K	NIST Webbook
tb	370.40 ± 0.30	K	NIST Webbook
tb	370.50 ± 0.50	K	NIST Webbook
tb	370.60 ± 0.50	K	NIST Webbook
tb	370.40 ± 0.50	K	NIST Webbook
tb	370.34 ± 0.07	K	NIST Webbook
tb	370.40 ± 0.10	K	NIST Webbook
tb	370.30 ± 0.05	K	NIST Webbook
tb	370.30 ± 0.30	K	NIST Webbook
tb	370.95 ± 0.10	K	NIST Webbook
tb	370.35 ± 0.03	K	NIST Webbook
tb	369.60 ± 1.00	K	NIST Webbook
tb	370.30 ± 0.20	K	NIST Webbook
tb	371.00 ± 0.30	K	NIST Webbook
tb	369.80 ± 0.40	K	NIST Webbook
tb	369.60 ± 1.00	K	NIST Webbook
tb	355.30 ± 0.50	K	NIST Webbook
tb	370.50 ± 0.50	K	NIST Webbook
tb	370.40 ± 0.50	K	NIST Webbook
tb	370.70 ± 0.60	K	NIST Webbook
tb	371.20 ± 0.50	K	NIST Webbook
tb	370.40 ± 0.30	K	NIST Webbook
tb	370.40 ± 0.30	K	NIST Webbook
tb	370.40 ± 0.40	K	NIST Webbook
tb	370.40 ± 0.50	K	NIST Webbook
tb	370.10 ± 0.50	K	NIST Webbook
tb	355.40 ± 0.30	K	NIST Webbook
tb	369.60 ± 0.50	K	NIST Webbook
tb	370.00 ± 1.00	K	NIST Webbook
tb	367.60 ± 1.00	K	NIST Webbook

tb	370.34 ± 0.20	K	NIST Webbook
tb	370.60 ± 0.50	K	NIST Webbook
tb	370.40 ± 0.50	K	NIST Webbook
tb	370.00 ± 1.00	K	NIST Webbook
tb	368.60 ± 1.00	K	NIST Webbook
tb	370.40 ± 0.30	K	NIST Webbook
tb	368.70 ± 1.00	K	NIST Webbook
tb	370.30 ± 0.20	K	NIST Webbook
tb	370.20 ± 0.30	K	NIST Webbook
tb	370.40 ± 0.20	K	NIST Webbook
tb	370.59 ± 0.33	K	NIST Webbook
tb	370.00 ± 0.50	K	NIST Webbook
tb	370.40 ± 0.30	K	NIST Webbook
tb	370.00 ± 1.00	K	NIST Webbook
tb	370.60 ± 0.30	K	NIST Webbook
tb	366.70 ± 3.00	K	NIST Webbook
tb	370.70 ± 1.00	K	NIST Webbook
tb	370.40 ± 0.30	K	NIST Webbook
tb	370.34 ± 0.30	K	NIST Webbook
tb	370.60 ± 0.30	K	NIST Webbook
tb	370.38 ± 0.10	K	NIST Webbook
tb	370.35 ± 0.30	K	NIST Webbook
tb	370.00 ± 1.00	K	NIST Webbook
tb	370.00 ± 1.50	K	NIST Webbook
tb	370.23 ± 0.20	K	NIST Webbook
tb	370.40 ± 0.20	K	NIST Webbook
tb	370.30 ± 0.05	K	NIST Webbook
tb	370.44 ± 0.15	K	NIST Webbook
tb	370.25 ± 0.20	K	NIST Webbook
tb	370.35 ± 0.30	K	NIST Webbook
tb	370.35	K	NIST Webbook
tb	370.35 ± 0.30	K	NIST Webbook
tb	370.23 ± 0.12	K	NIST Webbook
tb	370.35 ± 0.20	K	NIST Webbook
tb	370.30 ± 0.50	K	NIST Webbook
tb	370.30 ± 0.50	K	NIST Webbook
tb	370.29 ± 0.30	K	NIST Webbook
tb	370.40 ± 0.30	K	NIST Webbook
tb	370.30 ± 0.50	K	NIST Webbook
tb	370.32 ± 0.06	K	NIST Webbook
tb	370.30 ± 0.50	K	NIST Webbook
tb	370.60 ± 0.30	K	NIST Webbook
tb	370.15 ± 0.50	K	NIST Webbook
tb	370.40 ± 0.30	K	NIST Webbook

tb	370.40 ± 0.30	K	NIST Webbook
tb	370.40 ± 0.40	K	NIST Webbook
tb	370.30	K	NIST Webbook
tb	370.50 ± 0.20	K	NIST Webbook
tb	370.60	K	NIST Webbook
tb	370.10 ± 0.25	K	NIST Webbook
tb	370.30 ± 0.30	K	NIST Webbook
tb	370.40 ± 0.20	K	NIST Webbook
tb	370.30 ± 0.20	K	NIST Webbook
tb	370.28 ± 0.08	K	NIST Webbook
tb	370.26 ± 0.20	K	NIST Webbook
tb	370.26 ± 0.20	K	NIST Webbook
tb	370.20 ± 0.40	K	NIST Webbook
tb	370.26 ± 0.20	K	NIST Webbook
tb	370.26 ± 0.20	K	NIST Webbook
tb	370.30 ± 0.20	K	NIST Webbook
tb	370.21	K	Vapor Liquid Equilibrium Behaviors of 5-Methyl-2-(1-methylethyl)phenol in Alcohol
tb	370.21	K	Vapor-Liquid Equilibrium Behaviors of 3-Ethoxy-4-hydroxybenzaldehyde in Alcohol
tb	370.40	K	Evaluation of the 2-Methoxyethanol as Entrainer in Ethanol Water and 1-Propanol Water Mixtures
tb	370.05	K	P, rho, T Measurements and Isobaric Vapor-Liquid-Equilibria of the 1,3,3-Trimethyl-2-oxabicyclo[2,2,2]octane + Propan-1-ol Mixture: Cubic and Statistical Associating Fluid Theory-Based Equation of State Analysis
tb	370.24	K	Vapor-Liquid Equilibria for the Ternary System Acetonitrile + 1-Propanol + Dimethyl Sulfoxide and the Corresponding Binary Systems at 101.3 kPa
tb	370.21	K	Vapor-Liquid Equilibrium Behaviors of Coumarin and Vanillin in Ethanol, 1-Propanol, and 2-Propanol
tb	370.19	K	Vapor-Liquid Equilibrium Behavior of Tolan in Alcohol

tb	370.31	K	Measurement and Correlation of Excess Molar Enthalpies for Ethylene Glycol + Alkanol Systems at the Temperatures (298.15, 308.15, and 323.15) K
tb	370.30	K	Isobaric Phase Equilibria of Diethyl Carbonate with Five Alcohols at 101.3 kPa
tb	370.33	K	Isobaric Vapor-Liquid Equilibrium of the Acetonitrile + 1-Propanol + Ionic Liquids at an Atmospheric Pressure
tb	370.35	K	Measurement and Modelization of VLE for Butyl Acetate with Methanol, Ethanol, 1-Propanol, and 1-Butanol. Experimental Data at 0.15 MPa
tb	369.90	K	Separation of the mixture (isopropyl alcohol + diisopropyl ether + n-propanol): Entrainer selection, interaction exploration and vapour-liquid equilibrium measurements
tb	370.35	K	Isobaric vapor-liquid equilibria for extractive distillation of 1-propanol + water mixture using thiocyanate-based ionic liquids
tb	370.45	K	(Vapour + liquid) equilibria in the ternary system (acetonitrile + n-propanol + ethylene glycol) and corresponding binary systems at 101.3 kPa
tb	369.75	K	Isobaric (vapour + liquid + liquid) equilibrium data for (di-n-propyl ether + n-propyl alcohol + water) and (diisopropyl ether + isopropyl alcohol + water) systems at 100 kPa
tb	370.20	K	Heterogeneous azeotropic distillation for the separation of n-propanol + water mixture using n-propyl acetate as entrainer
tb	370.35	K	Isobaric (vapor-liquid) equilibria for binary systems of methanol + 1-(methoxymethoxy)-propane and 1-propanol + 1-(methoxymethoxy)-propane at 101.33 kPa

tb	370.31	K	Isobaric vapor liquid equilibrium for the binary systems of 1-propanol + 1-(methoxymethoxy)-butane and 1-butanol + 1-(methoxymethoxy)-butane at 101.3 kPa
tb	370.44	K	Isobaric VLE at 0.6 MPa for binary systems isobutyl acetate + ethanol, + 1-propanol or + 2-propanol
tb	370.40	K	Excess molar volumes of ternary mixtures of 1,3-dichlorobenzene and methyl ethyl ketone with 1-alkanols at 303.15K
tb	370.30	K	KDB
tb	369.90 ± 0.50	K	NIST Webbook
tb	369.80 ± 0.30	K	NIST Webbook
tb	370.20 ± 0.50	K	NIST Webbook
tb	370.40 ± 0.50	K	NIST Webbook
tb	370.50 ± 0.30	K	NIST Webbook
tc	537.00	K	Measurement and correlation of critical properties for binary mixtures and ternary mixtures containing gasoline additives
tc	536.80	K	KDB
tc	537.00	K	Measurement of critical temperatures and critical pressures for binary mixtures of methyl tert-butyl ether (MTBE) + alcohol and MTBE + alkane
tf	147.00	K	KDB
tf	147.00 ± 3.00	K	NIST Webbook
tf	146.70 ± 0.50	K	NIST Webbook
tf	146.95	K	NIST Webbook
tf	146.70	K	Phase equilibria of didecyldimethylammonium nitrate ionic liquid with water and organic solvents
tf	146.62	K	Aqueous Solubility Prediction Method
tt	148.75	K	KDB
tt	147.00 ± 0.30	K	NIST Webbook
tt	148.75 ± 0.02	K	NIST Webbook
tt	148.75 ± 0.02	K	NIST Webbook
vc	0.218	m ³ /kmol	NIST Webbook
vc	0.218	m ³ /kmol	KDB
vc	0.216 ± 0.001	m ³ /kmol	NIST Webbook
zc	0.2524720		KDB

Temperature Dependent Properties

Property code	Value	Unit	Temperature [K]	Source
cpg	115.56 ± 0.96	J/mol×K	420.75	NIST Webbook
cpg	106.44 ± 0.21	J/mol×K	391.20	NIST Webbook
cpg	114.35 ± 0.23	J/mol×K	431.20	NIST Webbook
cpg	110.42 ± 0.22	J/mol×K	411.20	NIST Webbook
cpg	113.59 ± 0.96	J/mol×K	409.95	NIST Webbook
cpg	111.21 ± 0.96	J/mol×K	396.95	NIST Webbook
cpg	109.42 ± 0.96	J/mol×K	387.15	NIST Webbook
cpg	115.97 ± 0.96	J/mol×K	422.95	NIST Webbook
cpg	108.67 ± 0.96	J/mol×K	383.05	NIST Webbook
cpg	107.28 ± 0.96	J/mol×K	375.45	NIST Webbook
cpg	102.26 ± 0.20	J/mol×K	371.20	NIST Webbook
cpg	118.71 ± 0.96	J/mol×K	437.95	NIST Webbook
cpg	118.62 ± 0.24	J/mol×K	451.20	NIST Webbook
cpg	122.94 ± 0.96	J/mol×K	461.05	NIST Webbook
cpg	125.55 ± 0.96	J/mol×K	475.35	NIST Webbook
cpg	130.97 ± 0.96	J/mol×K	504.95	NIST Webbook
cpg	132.23 ± 0.96	J/mol×K	511.85	NIST Webbook
cpg	135.98 ± 0.96	J/mol×K	532.35	NIST Webbook
cpg	141.05 ± 0.96	J/mol×K	560.05	NIST Webbook
cpg	144.49 ± 0.96	J/mol×K	578.85	NIST Webbook
cpg	148.95 ± 0.96	J/mol×K	603.25	NIST Webbook
cpl	167.80	J/mol×K	333.15	Excess Molar Enthalpies, Molar Heat Capacities, Densities, Viscosities, and Refractive Indices of Dimethyl Sulfoxide + 1-Propanol at (288.15, 298.15, and 308.15) K and at Normal Pressure
cpl	141.80	J/mol×K	293.15	NIST Webbook
cpl	146.88	J/mol×K	298.15	NIST Webbook
cpl	138.40	J/mol×K	288.15	NIST Webbook
cpl	144.44	J/mol×K	298.15	NIST Webbook
cpl	143.96	J/mol×K	298.15	NIST Webbook
cpl	144.60	J/mol×K	298.00	NIST Webbook

cpl	146.34	J/molxK	298.22	NIST Webbook
cpl	164.50	J/molxK	328.15	Excess Molar Enthalpies, Molar Heat Capacities, Densities, Viscosities, and Refractive Indices of Dimethyl Sulfoxide + 1-Propanol at (288.15, 298.15, and 308.15) K and at Normal Pressure
cpl	160.10	J/molxK	323.15	Excess Molar Enthalpies, Molar Heat Capacities, Densities, Viscosities, and Refractive Indices of Dimethyl Sulfoxide + 1-Propanol at (288.15, 298.15, and 308.15) K and at Normal Pressure
cpl	157.10	J/molxK	318.15	Excess Molar Enthalpies, Molar Heat Capacities, Densities, Viscosities, and Refractive Indices of Dimethyl Sulfoxide + 1-Propanol at (288.15, 298.15, and 308.15) K and at Normal Pressure
cpl	153.70	J/molxK	313.15	Excess Molar Enthalpies, Molar Heat Capacities, Densities, Viscosities, and Refractive Indices of Dimethyl Sulfoxide + 1-Propanol at (288.15, 298.15, and 308.15) K and at Normal Pressure

cpl	148.20	J/mol×K	303.15	Excess Molar Enthalpies, Molar Heat Capacities, Densities, Viscosities, and Refractive Indices of Dimethyl Sulfoxide + 1-Propanol at (288.15, 298.15, and 308.15) K and at Normal Pressure
cpl	144.90	J/mol×K	298.15	Excess Molar Enthalpies, Molar Heat Capacities, Densities, Viscosities, and Refractive Indices of Dimethyl Sulfoxide + 1-Propanol at (288.15, 298.15, and 308.15) K and at Normal Pressure
cpl	142.30	J/mol×K	293.15	Excess Molar Enthalpies, Molar Heat Capacities, Densities, Viscosities, and Refractive Indices of Dimethyl Sulfoxide + 1-Propanol at (288.15, 298.15, and 308.15) K and at Normal Pressure
cpl	139.00	J/mol×K	288.15	Excess Molar Enthalpies, Molar Heat Capacities, Densities, Viscosities, and Refractive Indices of Dimethyl Sulfoxide + 1-Propanol at (288.15, 298.15, and 308.15) K and at Normal Pressure

cpl	151.00	J/molxK	308.15	Excess Molar Enthalpies, Molar Heat Capacities, Densities, Viscosities, and Refractive Indices of Dimethyl Sulfoxide + 1-Propanol at (288.15, 298.15, and 308.15) K and at Normal Pressure
cpl	147.90	J/molxK	303.40	NIST Webbook
cpl	143.77	J/molxK	298.15	NIST Webbook
cpl	143.78	J/molxK	298.15	NIST Webbook
cpl	143.87	J/molxK	298.15	NIST Webbook
cpl	144.06	J/molxK	298.15	NIST Webbook
cpl	158.60	J/molxK	313.20	NIST Webbook
cpl	143.80	J/molxK	298.15	NIST Webbook
cpl	146.10	J/molxK	298.00	NIST Webbook
cpl	131.30	J/molxK	274.60	NIST Webbook
cpl	155.60	J/molxK	320.00	NIST Webbook
cpl	140.21	J/molxK	303.00	NIST Webbook
cpl	145.60	J/molxK	298.10	NIST Webbook
cpl	136.00	J/molxK	270.00	NIST Webbook
cpl	192.90	J/molxK	298.10	NIST Webbook
cpl	133.50	J/molxK	275.40	NIST Webbook
cpl	133.50	J/molxK	275.00	NIST Webbook
cpl	144.80	J/molxK	298.00	NIST Webbook
cpl	164.80	J/molxK	301.20	NIST Webbook
cpl	149.00	J/molxK	298.15	NIST Webbook
cps	106.30	J/molxK	150.00	NIST Webbook
dvisc	0.0013516	Paxs	313.15	Densities and Viscosities of (1-Propanol + 1,2-Dichloroethane), (1-Propanol + Benzaldehyde), (Benzaldehyde + 1,2-Dichloroethane), and (1-Propanol + 1,2-Dichloroethane + Benzaldehyde) Mixtures from T = 288.15 K to 313.15 K

dvisc	0.0016145	Paxs	303.15	Thermodynamic Study of Binary Mixture of $x_1[\text{C6mim}][\text{BF}_4] + x_2\text{1-propanol}$: Measurements and Molecular Modeling
dvisc	0.0013907	Paxs	313.15	Densities and Viscosities of Binary Mixture of the Ionic Liquid Bis(2-hydroxyethyl)ammonium Propionate with Methanol, Ethanol, and 1-Propanol at T = (293.15, 303.15, 313.15, and 323.15) K and at P = 0.1 MPa
dvisc	0.0010977	Paxs	323.15	Densities and Viscosities of Binary Mixture of the Ionic Liquid Bis(2-hydroxyethyl)ammonium Propionate with Methanol, Ethanol, and 1-Propanol at T = (293.15, 303.15, 313.15, and 323.15) K and at P = 0.1 MPa
dvisc	0.0025070	Paxs	288.15	Densities and Viscosities of Four Binary Diethyl Carbonate + 1-Alcohol Systems from (288.15 to 313.15) K
dvisc	0.0022020	Paxs	293.15	Densities and Viscosities of Four Binary Diethyl Carbonate + 1-Alcohol Systems from (288.15 to 313.15) K
dvisc	0.0019730	Paxs	298.15	Densities and Viscosities of Four Binary Diethyl Carbonate + 1-Alcohol Systems from (288.15 to 313.15) K

dvisc	0.0017330	Paxs	303.15	Densities and Viscosities of Four Binary Diethyl Carbonate + 1-Alcohol Systems from (288.15 to 313.15) K
dvisc	0.0013810	Paxs	313.15	Densities and Viscosities of Four Binary Diethyl Carbonate + 1-Alcohol Systems from (288.15 to 313.15) K
dvisc	0.0019414	Paxs	298.15	Densities and Viscosities for Binary Mixtures of the Ionic Liquid N-Ethyl Piperazinium Propionate with n-Alcohols at Several Temperatures
dvisc	0.0017282	Paxs	303.15	Densities and Viscosities for Binary Mixtures of the Ionic Liquid N-Ethyl Piperazinium Propionate with n-Alcohols at Several Temperatures
dvisc	0.0015448	Paxs	308.15	Densities and Viscosities for Binary Mixtures of the Ionic Liquid N-Ethyl Piperazinium Propionate with n-Alcohols at Several Temperatures
dvisc	0.0013847	Paxs	313.15	Densities and Viscosities for Binary Mixtures of the Ionic Liquid N-Ethyl Piperazinium Propionate with n-Alcohols at Several Temperatures

dvisc	0.0021920	Paxs	293.15	Densities and Viscosities of Naphthalen-1-ol, Naphthalen-2-ol, and 1-Aminonaphthalene in the Solvents of Ethanol, Propan-1-ol, and Butan-1-ol
dvisc	0.0017260	Paxs	303.15	Densities and Viscosities of Naphthalen-1-ol, Naphthalen-2-ol, and 1-Aminonaphthalene in the Solvents of Ethanol, Propan-1-ol, and Butan-1-ol
dvisc	0.0013730	Paxs	313.15	Densities and Viscosities of Naphthalen-1-ol, Naphthalen-2-ol, and 1-Aminonaphthalene in the Solvents of Ethanol, Propan-1-ol, and Butan-1-ol
dvisc	0.0011020	Paxs	323.15	Densities and Viscosities of Naphthalen-1-ol, Naphthalen-2-ol, and 1-Aminonaphthalene in the Solvents of Ethanol, Propan-1-ol, and Butan-1-ol
dvisc	0.0008980	Paxs	333.15	Densities and Viscosities of Naphthalen-1-ol, Naphthalen-2-ol, and 1-Aminonaphthalene in the Solvents of Ethanol, Propan-1-ol, and Butan-1-ol
dvisc	0.0019550	Paxs	298.15	Densities and Viscosities of Binary Mixtures of 2,2,4-Trimethylpentane + 1-Propanol, + 1-Pentanol, + 1-Hexanol, and + 1-Heptanol from (298.15 to 323.15) K

dvisc	0.0017370	Paxs	303.15	Densities and Viscosities of Binary Mixtures of 2,2,4-Trimethylpentane + 1-Propanol, + 1-Pentanol, + 1-Hexanol, and + 1-Heptanol from (298.15 to 323.15) K
dvisc	0.0015480	Paxs	308.15	Densities and Viscosities of Binary Mixtures of 2,2,4-Trimethylpentane + 1-Propanol, + 1-Pentanol, + 1-Hexanol, and + 1-Heptanol from (298.15 to 323.15) K
dvisc	0.0013840	Paxs	313.15	Densities and Viscosities of Binary Mixtures of 2,2,4-Trimethylpentane + 1-Propanol, + 1-Pentanol, + 1-Hexanol, and + 1-Heptanol from (298.15 to 323.15) K
dvisc	0.0012410	Paxs	318.15	Densities and Viscosities of Binary Mixtures of 2,2,4-Trimethylpentane + 1-Propanol, + 1-Pentanol, + 1-Hexanol, and + 1-Heptanol from (298.15 to 323.15) K
dvisc	0.0011170	Paxs	323.15	Densities and Viscosities of Binary Mixtures of 2,2,4-Trimethylpentane + 1-Propanol, + 1-Pentanol, + 1-Hexanol, and + 1-Heptanol from (298.15 to 323.15) K
dvisc	0.0022050	Paxs	293.15	Viscosities and Densities of Binary Mixtures of (N-Acetylmorpholine + Alkanols) from (293.15 to 323.15) K

dvisc	0.0017140	Paxs	303.15	Viscosities and Densities of Binary Mixtures of (N-Acetylmorpholine + Alkanols) from (293.15 to 323.15) K
dvisc	0.0011050	Paxs	323.15	Viscosities and Densities of Binary Mixtures of (N-Acetylmorpholine + Alkanols) from (293.15 to 323.15) K
dvisc	0.0022040	Paxs	293.15	Thermodynamic Study of Binary Mixture of x1[C6mim][BF4] + x21-propanol: Measurements and Molecular Modeling
dvisc	0.0013830	Paxs	313.15	Properties of pure 1-methylimidazolium acetate ionic liquid and its binary mixtures with alcohols
dvisc	0.0015480	Paxs	308.15	Properties of pure 1-methylimidazolium acetate ionic liquid and its binary mixtures with alcohols
dvisc	0.0017370	Paxs	303.15	Properties of pure 1-methylimidazolium acetate ionic liquid and its binary mixtures with alcohols
dvisc	0.0019760	Paxs	298.15	Properties of pure 1-methylimidazolium acetate ionic liquid and its binary mixtures with alcohols
dvisc	0.0022060	Paxs	293.15	Properties of pure 1-methylimidazolium acetate ionic liquid and its binary mixtures with alcohols

dvisc	0.0010819	Paxs	333.15	Density and viscosity of 1-butyl-3-methylimidazolium nitrate with ethanol, 1-propanol, or 1-butanol at several temperatures
dvisc	0.0011523	Paxs	328.15	Density and viscosity of 1-butyl-3-methylimidazolium nitrate with ethanol, 1-propanol, or 1-butanol at several temperatures
dvisc	0.0012560	Paxs	323.15	Density and viscosity of 1-butyl-3-methylimidazolium nitrate with ethanol, 1-propanol, or 1-butanol at several temperatures
dvisc	0.0013833	Paxs	318.15	Density and viscosity of 1-butyl-3-methylimidazolium nitrate with ethanol, 1-propanol, or 1-butanol at several temperatures
dvisc	0.0015237	Paxs	313.15	Density and viscosity of 1-butyl-3-methylimidazolium nitrate with ethanol, 1-propanol, or 1-butanol at several temperatures
dvisc	0.0016709	Paxs	308.15	Density and viscosity of 1-butyl-3-methylimidazolium nitrate with ethanol, 1-propanol, or 1-butanol at several temperatures
dvisc	0.0018742	Paxs	303.15	Density and viscosity of 1-butyl-3-methylimidazolium nitrate with ethanol, 1-propanol, or 1-butanol at several temperatures

dvisc	0.0021178	Paxs	298.15	Density and viscosity of 1-butyl-3-methylimidazolium nitrate with ethanol, 1-propanol, or 1-butanol at several temperatures
dvisc	0.0024104	Paxs	293.15	Density and viscosity of 1-butyl-3-methylimidazolium nitrate with ethanol, 1-propanol, or 1-butanol at several temperatures
dvisc	0.0027676	Paxs	288.15	Density and viscosity of 1-butyl-3-methylimidazolium nitrate with ethanol, 1-propanol, or 1-butanol at several temperatures
dvisc	0.0034744	Paxs	283.15	Density and viscosity of 1-butyl-3-methylimidazolium nitrate with ethanol, 1-propanol, or 1-butanol at several temperatures
dvisc	0.0011004	Paxs	323.15	Densities and viscosities of binary mixtures of {dimethylsulfoxide + aliphatic lower alkanols (C1 C3)} at temperatures from T = 303.15 K to T = 323.15 K
dvisc	0.0012340	Paxs	318.15	Densities and viscosities of binary mixtures of {dimethylsulfoxide + aliphatic lower alkanols (C1 C3)} at temperatures from T = 303.15 K to T = 323.15 K
dvisc	0.0011883	Paxs	313.15	Thermodynamic Study of Binary Mixture of x ₁ [C6mim][BF ₄] + x ₂ 1-propanol: Measurements and Molecular Modeling

dvisc	0.0013672	Paxs	313.15	Densities and viscosities of binary mixtures of {dimethylsulfoxide + aliphatic lower alkanols (C1 C3)} at temperatures from T = 303.15 K to T = 323.15 K
dvisc	0.0015344	Paxs	308.15	Densities and viscosities of binary mixtures of {dimethylsulfoxide + aliphatic lower alkanols (C1 C3)} at temperatures from T = 303.15 K to T = 323.15 K
dvisc	0.0017325	Paxs	303.15	Densities and viscosities of binary mixtures of {dimethylsulfoxide + aliphatic lower alkanols (C1 C3)} at temperatures from T = 303.15 K to T = 323.15 K
dvisc	0.0022720	Paxs	293.15	Densities and Kinematic Viscosities of Ten Binary 1-Alkanol Liquid Systems at Temperatures of (293.15 and 298.15) K
dvisc	0.0019150	Paxs	298.15	Dynamic viscosities of binary mixtures of cycloalkanes with primary alcohols at T = (293.15, 298.15, and 303.15) K: New UNIFAC-VISCO interaction parameters
dvisc	0.0021590	Paxs	293.15	Dynamic viscosities of binary mixtures of cycloalkanes with primary alcohols at T = (293.15, 298.15, and 303.15) K: New UNIFAC-VISCO interaction parameters

dvisc	0.0015330	Paxs	308.15	Densities, viscosities, and ultrasonic velocity studies of binary mixtures of trichloromethane with methanol, ethanol, propan-1-ol, and butan-1-ol at T=(298.15 and 308.15) K
dvisc	0.0019500	Paxs	298.15	Densities, viscosities, and ultrasonic velocity studies of binary mixtures of trichloromethane with methanol, ethanol, propan-1-ol, and butan-1-ol at T=(298.15 and 308.15) K
dvisc	0.0017130	Paxs	303.15	Viscosity, density, and speed of sound of methylcyclopentane with primary and secondary alcohols at T = (293.15, 298.15, and 303.15) K
dvisc	0.0019150	Paxs	298.15	Viscosity, density, and speed of sound of methylcyclopentane with primary and secondary alcohols at T = (293.15, 298.15, and 303.15) K
dvisc	0.0021590	Paxs	293.15	Viscosity, density, and speed of sound of methylcyclopentane with primary and secondary alcohols at T = (293.15, 298.15, and 303.15) K
dvisc	0.0013400	Paxs	313.00	Ultrasonic velocity, viscosity and excess properties of binary mixture of tetrahydrofuran with 1-propanol and 2-propanol

dvisc	0.0016100	Paxs	303.00	Ultrasonic velocity, viscosity and excess properties of binary mixture of tetrahydrofuran with 1-propanol and 2-propanol
dvisc	0.0021900	Paxs	293.00	Ultrasonic velocity, viscosity and excess properties of binary mixture of tetrahydrofuran with 1-propanol and 2-propanol
dvisc	0.0017300	Paxs	303.15	Viscosities and densities for binary mixtures of N-methylpiperazine with methanol, ethanol, n-propanol, iso-propanol, n-butanol and iso-butanol at 293.15, 298.15 and 303.15K
dvisc	0.0019450	Paxs	298.15	Viscosities and densities for binary mixtures of N-methylpiperazine with methanol, ethanol, n-propanol, iso-propanol, n-butanol and iso-butanol at 293.15, 298.15 and 303.15K
dvisc	0.0021940	Paxs	293.15	Viscosities and densities for binary mixtures of N-methylpiperazine with methanol, ethanol, n-propanol, iso-propanol, n-butanol and iso-butanol at 293.15, 298.15 and 303.15K

dvisc	0.0007420	Paxs	343.15	Density and Viscosity of Binary Mixtures of Diethyl Carbonate with Alcohols at (293.15 to 363.15) K and Predictive Results by UNIFAC-VISCO Group Contribution Method
dvisc	0.0019660	Paxs	298.15	Densities and Kinematic Viscosities of Ten Binary 1-Alkanol Liquid Systems at Temperatures of (293.15 and 298.15) K
dvisc	0.0019400	Paxs	298.15	Temperature and Composition Dependence of the Density and Viscosity of Binary Mixtures of {1-Butyl-3-methylimidazolium Thiocyanate + 1-Alcohols}
dvisc	0.0015400	Paxs	308.15	Temperature and Composition Dependence of the Density and Viscosity of Binary Mixtures of {1-Butyl-3-methylimidazolium Thiocyanate + 1-Alcohols}
dvisc	0.0012400	Paxs	318.15	Temperature and Composition Dependence of the Density and Viscosity of Binary Mixtures of {1-Butyl-3-methylimidazolium Thiocyanate + 1-Alcohols}
dvisc	0.0010200	Paxs	328.15	Temperature and Composition Dependence of the Density and Viscosity of Binary Mixtures of {1-Butyl-3-methylimidazolium Thiocyanate + 1-Alcohols}

dvisc	0.0008420	Paxs	338.15	Temperature and Composition Dependence of the Density and Viscosity of Binary Mixtures of {1-Butyl-3-methylimidazolium Thiocyanate + 1-Alcohols}
dvisc	0.0007100	Paxs	348.15	Temperature and Composition Dependence of the Density and Viscosity of Binary Mixtures of {1-Butyl-3-methylimidazolium Thiocyanate + 1-Alcohols}
dvisc	0.0019550	Paxs	298.15	Densities and Viscosities of Binary Mixtures of Cyclopropanecarboxylic Acid with Methanol, Ethanol, Propan-1-ol, and Butan-1-ol at Different Temperatures
dvisc	0.0015380	Paxs	308.15	Densities and Viscosities of Binary Mixtures of Cyclopropanecarboxylic Acid with Methanol, Ethanol, Propan-1-ol, and Butan-1-ol at Different Temperatures
dvisc	0.0012240	Paxs	318.15	Densities and Viscosities of Binary Mixtures of Cyclopropanecarboxylic Acid with Methanol, Ethanol, Propan-1-ol, and Butan-1-ol at Different Temperatures

dvisc	0.0009860	Paxs	328.15	Densities and Viscosities of Binary Mixtures of Cyclopropanecarboxylic Acid with Methanol, Ethanol, Propan-1-ol, and Butan-1-ol at Different Temperatures
dvisc	0.0008010	Paxs	338.15	Densities and Viscosities of Binary Mixtures of Cyclopropanecarboxylic Acid with Methanol, Ethanol, Propan-1-ol, and Butan-1-ol at Different Temperatures
dvisc	0.0024202	Paxs	288.15	Densities and Viscosities of (1-Propanol + 1,2-Dichloroethane), (1-Propanol + Benzaldehyde), (Benzaldehyde + 1,2-Dichloroethane), and (1-Propanol + 1,2-Dichloroethane + Benzaldehyde) Mixtures from T = 288.15 K to 313.15 K
dvisc	0.0021372	Paxs	293.15	Densities and Viscosities of (1-Propanol + 1,2-Dichloroethane), (1-Propanol + Benzaldehyde), (Benzaldehyde + 1,2-Dichloroethane), and (1-Propanol + 1,2-Dichloroethane + Benzaldehyde) Mixtures from T = 288.15 K to 313.15 K

dvisc	0.0018939	Paxs	298.15	Densities and Viscosities of (1-Propanol + 1,2-Dichloroethane), (1-Propanol + Benzaldehyde), (Benzaldehyde + 1,2-Dichloroethane), and (1-Propanol + 1,2-Dichloroethane + Benzaldehyde) Mixtures from T = 288.15 K to 313.15 K
dvisc	0.0016841	Paxs	303.15	Densities and Viscosities of (1-Propanol + 1,2-Dichloroethane), (1-Propanol + Benzaldehyde), (Benzaldehyde + 1,2-Dichloroethane), and (1-Propanol + 1,2-Dichloroethane + Benzaldehyde) Mixtures from T = 288.15 K to 313.15 K
dvisc	0.0021412	Paxs	293.15	Densities and Viscosities of Binary Mixture of the Ionic Liquid Bis(2-hydroxyethyl)ammonium Propionate with Methanol, Ethanol, and 1-Propanol at T = (293.15, 303.15, 313.15, and 323.15) K and at P = 0.1 MPa
dvisc	0.0005150	Paxs	363.15	Density and Viscosity of Binary Mixtures of Diethyl Carbonate with Alcohols at (293.15 to 363.15) K and Predictive Results by UNIFAC-VISCO Group Contribution Method

dvisc	0.0006150	Paxs	353.15	Density and Viscosity of Binary Mixtures of Diethyl Carbonate with Alcohols at (293.15 to 363.15) K and Predictive Results by UNIFAC-VISCO Group Contribution Method
dvisc	0.0015037	Paxs	308.15	Densities and Viscosities of (1-Propanol + 1,2-Dichloroethane), (1-Propanol + Benzaldehyde), (Benzaldehyde + 1,2-Dichloroethane), and (1-Propanol + 1,2-Dichloroethane + Benzaldehyde) Mixtures from T = 288.15 K to 313.15 K
dvisc	0.0009070	Paxs	333.15	Density and Viscosity of Binary Mixtures of Diethyl Carbonate with Alcohols at (293.15 to 363.15) K and Predictive Results by UNIFAC-VISCO Group Contribution Method
dvisc	0.0011150	Paxs	323.15	Density and Viscosity of Binary Mixtures of Diethyl Carbonate with Alcohols at (293.15 to 363.15) K and Predictive Results by UNIFAC-VISCO Group Contribution Method

dvisc	0.0013810	Paxs	313.15	Density and Viscosity of Binary Mixtures of Diethyl Carbonate with Alcohols at (293.15 to 363.15) K and Predictive Results by UNIFAC-VISCO Group Contribution Method
dvisc	0.0017450	Paxs	303.15	Density and Viscosity of Binary Mixtures of Diethyl Carbonate with Alcohols at (293.15 to 363.15) K and Predictive Results by UNIFAC-VISCO Group Contribution Method
dvisc	0.0019810	Paxs	298.15	Density and Viscosity of Binary Mixtures of Diethyl Carbonate with Alcohols at (293.15 to 363.15) K and Predictive Results by UNIFAC-VISCO Group Contribution Method
dvisc	0.0022380	Paxs	293.15	Density and Viscosity of Binary Mixtures of Diethyl Carbonate with Alcohols at (293.15 to 363.15) K and Predictive Results by UNIFAC-VISCO Group Contribution Method

dvisc	0.0013630	Paxs	313.15	Densities, Viscosities, and Ultrasonic Velocity Studies of Binary Mixtures of Chloroform with Propan-1-ol and Butan-1-ol at (303.15 and 313.15) K
dvisc	0.0017190	Paxs	303.15	Densities, Viscosities, and Ultrasonic Velocity Studies of Binary Mixtures of Chloroform with Propan-1-ol and Butan-1-ol at (303.15 and 313.15) K
dvisc	0.0013780	Paxs	313.15	Density and Viscosity of Binary Mixtures of Ethyl-2-methylbutyrate and Ethyl Hexanoate with Methanol, Ethanol, and 1-Propanol at (293.15, 303.15, and 313.15) K
dvisc	0.0017130	Paxs	303.15	Density and Viscosity of Binary Mixtures of Ethyl-2-methylbutyrate and Ethyl Hexanoate with Methanol, Ethanol, and 1-Propanol at (293.15, 303.15, and 313.15) K
dvisc	0.0021880	Paxs	293.15	Density and Viscosity of Binary Mixtures of Ethyl-2-methylbutyrate and Ethyl Hexanoate with Methanol, Ethanol, and 1-Propanol at (293.15, 303.15, and 313.15) K

dvisc	0.0017880	Paxs	303.15	Densities and Viscosities of Binary Mixture of the Ionic Liquid Bis(2-hydroxyethyl)ammonium Propionate with Methanol, Ethanol, and 1-Propanol at T = (293.15, 303.15, 313.15, and 323.15) K and at P = 0.1 MPa
dvisc	0.0015650	Paxs	308.15	Viscosities, Densities, and Ultrasonic Velocities of Binary Mixtures of Ethylbenzene with Ethanol, 1-Propanol, and 1-Butanol at (298.15 and 308.15)K
dvisc	0.0020420	Paxs	298.15	Viscosities, Densities, and Ultrasonic Velocities of Binary Mixtures of Ethylbenzene with Ethanol, 1-Propanol, and 1-Butanol at (298.15 and 308.15)K
dvisc	0.0015370	Paxs	308.15	Excess Molar Volumes and Viscosity Deviations of Binary Mixtures of 2,4,6-Trimethyl-1,3,5-trioxane + Ethanol, 1-Propanol, and 1-Butanol at (298.15, 303.15, and 308.15) K
dvisc	0.0017250	Paxs	303.15	Excess Molar Volumes and Viscosity Deviations of Binary Mixtures of 2,4,6-Trimethyl-1,3,5-trioxane + Ethanol, 1-Propanol, and 1-Butanol at (298.15, 303.15, and 308.15) K

dvisc	0.0019430	Paxs	298.15	Excess Molar Volumes and Viscosity Deviations of Binary Mixtures of 2,4,6-Trimethyl-1,3,5-trioxane + Ethanol, 1-Propanol, and 1-Butanol at (298.15, 303.15, and 308.15) K
dvisc	0.0005985	Paxs	333.15	Thermodynamic Study of Binary Mixture of x1[C6mim][BF4] + x21-propanol: Measurements and Molecular Modeling
dvisc	0.0008796	Paxs	323.15	Thermodynamic Study of Binary Mixture of x1[C6mim][BF4] + x21-propanol: Measurements and Molecular Modeling
dvisc	0.0013820	Paxs	313.15	Viscosities and Densities of Binary Mixtures of (N-Acetylmorpholine + Alkanols) from (293.15 to 323.15) K
dvisc	0.0017130	Paxs	303.15	Dynamic viscosities of binary mixtures of cycloalkanes with primary alcohols at T = (293.15, 298.15, and 303.15) K: New UNIFAC-VISCO interaction parameters
hfust	5.37	kJ/mol	148.75	NIST Webbook
hfust	5.19	kJ/mol	147.00	NIST Webbook
hfust	5.37	kJ/mol	148.80	NIST Webbook
hfust	5.37	kJ/mol	148.80	NIST Webbook
hfust	5.40	kJ/mol	148.70	NIST Webbook
hvapt	41.44	kJ/mol	370.30	NIST Webbook
hvapt	41.20	kJ/mol	371.00	NIST Webbook
hvapt	35.20	kJ/mol	423.00	NIST Webbook
hvapt	21.00	kJ/mol	498.00	NIST Webbook
hvapt	11.40	kJ/mol	528.00	NIST Webbook
hvapt	47.00	kJ/mol	336.50	NIST Webbook

hvapt	42.90	kJ/mol	368.50	NIST Webbook
hvapt	48.00	kJ/mol	214.00	NIST Webbook
hvapt	43.50	kJ/mol	366.00	NIST Webbook
hvapt	42.30	kJ/mol	388.00	NIST Webbook
hvapt	40.10	kJ/mol	441.50	NIST Webbook
hvapt	36.50	kJ/mol	492.50	NIST Webbook
hvapt	46.40 ± 0.10	kJ/mol	313.00	NIST Webbook
hvapt	45.70 ± 0.10	kJ/mol	323.00	NIST Webbook
hvapt	44.90 ± 0.10	kJ/mol	333.00	NIST Webbook
hvapt	41.76	kJ/mol	370.40	KDB
hvapt	44.00 ± 0.10	kJ/mol	343.00	NIST Webbook
hvapt	43.20 ± 0.10	kJ/mol	353.00	NIST Webbook
hvapt	42.40 ± 0.10	kJ/mol	363.00	NIST Webbook
hvapt	49.30	kJ/mol	324.00	NIST Webbook
hvapt	46.90	kJ/mol	331.00	NIST Webbook
hvapt	46.70	kJ/mol	318.00	NIST Webbook
hvapt	40.70	kJ/mol	471.00	NIST Webbook
hvapt	44.30	kJ/mol	358.00	NIST Webbook
hvapt	44.10	kJ/mol	364.00	NIST Webbook
hvapt	43.90 ± 0.10	kJ/mol	343.00	NIST Webbook
hvapt	42.30 ± 0.10	kJ/mol	360.00	NIST Webbook
hvapt	41.20 ± 0.10	kJ/mol	370.00	NIST Webbook
hvapt	40.30 ± 0.10	kJ/mol	378.00	NIST Webbook
hvapt	39.70 ± 0.10	kJ/mol	384.00	NIST Webbook
hvapt	45.50	kJ/mol	344.00	NIST Webbook
hvapt	43.20	kJ/mol	354.00	NIST Webbook
hvapt	44.99 ± 0.42	kJ/mol	333.13	NIST Webbook
hvapt	29.40	kJ/mol	453.00	NIST Webbook
hvapt	44.70	kJ/mol	355.00	NIST Webbook
pvap	15.00	kPa	327.25	P, rho, T Measurements and Isobaric Vapor-Liquid-Equilibria of the 1,3,3-Trimethyl-2-oxabicyclo[2,2,2]octane + Propan-1-ol Mixture: Cubic and Statistical Associating Fluid Theory-Based Equation of State Analysis

pvap	19.93	kPa	333.15	Isothermal (vapor + liquid) equilibria for the binary mixtures of (propylene oxide + ethanol) and (propylene oxide + 1-propanol) at several temperatures
pvap	76.56	kPa	363.20	(Vapour + liquid) equilibria, (VLE) excess molar enthalpies and infinite dilution activity coefficients of selected binary systems involving n-hexyl pyridinium bis(trifluoromethylsulphonyl)imide ionic liquid: Experimental and predictions using modified UNIFAC (Dortmund)
pvap	100.00	kPa	369.80	Isobaric vapor-liquid equilibria for the 1-propanol + water + 1-ethyl-3-methylimidazolium dicyanamide system at 100 kPa
pvap	101.30	kPa	370.35	Isobaric vapor-liquid equilibria for extractive distillation of 1-propanol + water mixture using thiocyanate-based ionic liquids
pvap	100.51	kPa	370.10	Vapor-Liquid Equilibrium of Binary Mixtures Containing Isopropyl Acetate and Alkanols at 101.32 kPa
pvap	125.86	kPa	376.30	Vapor-Liquid Equilibrium of Binary Mixtures Containing Isopropyl Acetate and Alkanols at 101.32 kPa

pvap	151.03	kPa	381.50	Vapor-Liquid Equilibrium of Binary Mixtures Containing Isopropyl Acetate and Alkanols at 101.32 kPa
pvap	176.20	kPa	385.89	Vapor-Liquid Equilibrium of Binary Mixtures Containing Isopropyl Acetate and Alkanols at 101.32 kPa
pvap	201.37	kPa	389.89	Vapor-Liquid Equilibrium of Binary Mixtures Containing Isopropyl Acetate and Alkanols at 101.32 kPa
pvap	226.54	kPa	393.59	Vapor-Liquid Equilibrium of Binary Mixtures Containing Isopropyl Acetate and Alkanols at 101.32 kPa
pvap	251.71	kPa	396.19	Vapor-Liquid Equilibrium of Binary Mixtures Containing Isopropyl Acetate and Alkanols at 101.32 kPa
pvap	101.30	kPa	370.33	Isobaric Vapor-Liquid Equilibrium of the Acetonitrile + 1-Propanol + Ionic Liquids at an Atmospheric Pressure
pvap	2.88	kPa	298.15	Vapor-Pressure Measurements of Liquid Solutions at Different Temperatures: Apparatus for Use over an Extended Temperature Range and Some New Data

pvap	9.01	kPa	318.15	Isothermal (vapor + liquid) equilibria for the binary mixtures of (propylene oxide + ethanol) and (propylene oxide + 1-propanol) at several temperatures
pvap	5.37	kPa	308.15	Vapor-Pressure Measurements of Liquid Solutions at Different Temperatures: Apparatus for Use over an Extended Temperature Range and Some New Data
pvap	7.20	kPa	313.15	Vapor-Pressure Measurements of Liquid Solutions at Different Temperatures: Apparatus for Use over an Extended Temperature Range and Some New Data
pvap	9.57	kPa	318.15	Vapor-Pressure Measurements of Liquid Solutions at Different Temperatures: Apparatus for Use over an Extended Temperature Range and Some New Data
pvap	12.42	kPa	323.15	Vapor-Pressure Measurements of Liquid Solutions at Different Temperatures: Apparatus for Use over an Extended Temperature Range and Some New Data

pvap	16.14	kPa	328.15	Vapor-Pressure Measurements of Liquid Solutions at Different Temperatures: Apparatus for Use over an Extended Temperature Range and Some New Data
pvap	20.72	kPa	333.15	Vapor-Pressure Measurements of Liquid Solutions at Different Temperatures: Apparatus for Use over an Extended Temperature Range and Some New Data
pvap	26.41	kPa	338.15	Vapor-Pressure Measurements of Liquid Solutions at Different Temperatures: Apparatus for Use over an Extended Temperature Range and Some New Data
pvap	33.23	kPa	343.15	Vapor-Pressure Measurements of Liquid Solutions at Different Temperatures: Apparatus for Use over an Extended Temperature Range and Some New Data
pvap	41.63	kPa	348.15	Vapor-Pressure Measurements of Liquid Solutions at Different Temperatures: Apparatus for Use over an Extended Temperature Range and Some New Data

pvap	51.67	kPa	353.15	Vapor-Pressure Measurements of Liquid Solutions at Different Temperatures: Apparatus for Use over an Extended Temperature Range and Some New Data
pvap	63.69	kPa	358.15	Vapor-Pressure Measurements of Liquid Solutions at Different Temperatures: Apparatus for Use over an Extended Temperature Range and Some New Data
pvap	3.63	kPa	303.15	Isothermal (vapor + liquid) equilibria for the binary mixtures of (propylene oxide + ethanol) and (propylene oxide + 1-propanol) at several temperatures
pvap	5.88	kPa	310.17	Vapor-Liquid Equilibrium for Benzene + 2-Methylpentane and Allyl Alcohol + 1-Propanol
pvap	6.98	kPa	313.15	Vapor-Liquid Equilibrium for Benzene + 2-Methylpentane and Allyl Alcohol + 1-Propanol
pvap	7.69	kPa	314.84	Vapor-Liquid Equilibrium for Benzene + 2-Methylpentane and Allyl Alcohol + 1-Propanol
pvap	10.18	kPa	319.89	Vapor-Liquid Equilibrium for Benzene + 2-Methylpentane and Allyl Alcohol + 1-Propanol
pvap	13.14	kPa	324.71	Vapor-Liquid Equilibrium for Benzene + 2-Methylpentane and Allyl Alcohol + 1-Propanol

pvap	15.75	kPa	328.14	Vapor-Liquid Equilibrium for Benzene + 2-Methylpentane and Allyl Alcohol + 1-Propanol
pvap	19.31	kPa	332.18	Vapor-Liquid Equilibrium for Benzene + 2-Methylpentane and Allyl Alcohol + 1-Propanol
pvap	20.27	kPa	333.15	Vapor-Liquid Equilibrium for Benzene + 2-Methylpentane and Allyl Alcohol + 1-Propanol
pvap	22.56	kPa	335.34	Vapor-Liquid Equilibrium for Benzene + 2-Methylpentane and Allyl Alcohol + 1-Propanol
pvap	26.91	kPa	339.01	Vapor-Liquid Equilibrium for Benzene + 2-Methylpentane and Allyl Alcohol + 1-Propanol
pvap	31.12	kPa	342.11	Vapor-Liquid Equilibrium for Benzene + 2-Methylpentane and Allyl Alcohol + 1-Propanol
pvap	35.82	kPa	345.18	Vapor-Liquid Equilibrium for Benzene + 2-Methylpentane and Allyl Alcohol + 1-Propanol
pvap	40.12	kPa	347.71	Vapor-Liquid Equilibrium for Benzene + 2-Methylpentane and Allyl Alcohol + 1-Propanol
pvap	44.26	kPa	349.93	Vapor-Liquid Equilibrium for Benzene + 2-Methylpentane and Allyl Alcohol + 1-Propanol
pvap	50.95	kPa	353.15	Vapor-Liquid Equilibrium for Benzene + 2-Methylpentane and Allyl Alcohol + 1-Propanol

pvap	50.95	kPa	353.17	Vapor-Liquid Equilibrium for Benzene + 2-Methylpentane and Allyl Alcohol + 1-Propanol
pvap	59.13	kPa	356.70	Vapor-Liquid Equilibrium for Benzene + 2-Methylpentane and Allyl Alcohol + 1-Propanol
pvap	9.47	kPa	318.58	Vapor-Liquid Equilibrium for 1-Propanol + 1-Butene, + cis-2-Butene, + 2-Methyl-propene, + trans-2-Butene, + n-Butane, and + 2-Methyl-propane
pvap	12.07	kPa	323.07	Vapor-Liquid Equilibrium for 1-Propanol + 1-Butene, + cis-2-Butene, + 2-Methyl-propene, + trans-2-Butene, + n-Butane, and + 2-Methyl-propane
pvap	12.27	kPa	323.26	Vapor-Liquid Equilibrium for 1-Propanol + 1-Butene, + cis-2-Butene, + 2-Methyl-propene, + trans-2-Butene, + n-Butane, and + 2-Methyl-propane
pvap	17.57	kPa	330.16	Vapor-Liquid Equilibrium for 1-Propanol + 1-Butene, + cis-2-Butene, + 2-Methyl-propene, + trans-2-Butene, + n-Butane, and + 2-Methyl-propane
pvap	77.81	kPa	363.15	Vapor-Pressure Measurements of Liquid Solutions at Different Temperatures: Apparatus for Use over an Extended Temperature Range and Some New Data

pvap	19.17	kPa	332.05	Vapor-Liquid Equilibrium for 1-Propanol + 1-Butene, + cis-2-Butene, + 2-Methyl-propene, + trans-2-Butene, + n-Butane, and + 2-Methyl-propane
pvap	81.30	kPa	364.51	Isothermal Vapor Liquid Equilibrium for 2-Methylpropene + Methanol, + 1-Propanol, + 2-Propanol, + 2-Butanol, and + 2-Methyl-2-propanol Binary Systems at 364.5 K
pvap	1558.00	kPa	468.15	Vapor-Liquid Equilibria for the n-Pentane + 1-Propanol and n-Pentane + 2-Methyl-1-propanol Systems near the Critical Region
pvap	2081.00	kPa	483.15	Vapor-Liquid Equilibria for the n-Pentane + 1-Propanol and n-Pentane + 2-Methyl-1-propanol Systems near the Critical Region
pvap	2733.00	kPa	498.15	Vapor-Liquid Equilibria for the n-Pentane + 1-Propanol and n-Pentane + 2-Methyl-1-propanol Systems near the Critical Region
pvap	3546.00	kPa	513.15	Vapor-Liquid Equilibria for the n-Pentane + 1-Propanol and n-Pentane + 2-Methyl-1-propanol Systems near the Critical Region
pvap	406.60	kPa	413.20	High-temperature vapour liquid equilibrium for the (water + alcohol) systems and modelling with SAFT-VR: 2. Water-1-propanol

pvap	303.10	kPa	403.20	High-temperature vapour liquid equilibrium for the (water + alcohol) systems and modelling with SAFT-VR: 2. Water-1-propanol
pvap	223.30	kPa	393.20	High-temperature vapour liquid equilibrium for the (water + alcohol) systems and modelling with SAFT-VR: 2. Water-1-propanol
pvap	160.40	kPa	383.20	High-temperature vapour liquid equilibrium for the (water + alcohol) systems and modelling with SAFT-VR: 2. Water-1-propanol
pvap	114.90	kPa	373.65	Isobaric (vapour + liquid) equilibria for the (1-propanol + 1-butanol) binary mixture at (53.3 and 91.3) kPa
pvap	101.80	kPa	370.40	Isobaric (vapour + liquid) equilibria for the (1-propanol + 1-butanol) binary mixture at (53.3 and 91.3) kPa
pvap	86.80	kPa	366.25	Isobaric (vapour + liquid) equilibria for the (1-propanol + 1-butanol) binary mixture at (53.3 and 91.3) kPa
pvap	74.70	kPa	362.45	Isobaric (vapour + liquid) equilibria for the (1-propanol + 1-butanol) binary mixture at (53.3 and 91.3) kPa
pvap	60.70	kPa	357.35	Isobaric (vapour + liquid) equilibria for the (1-propanol + 1-butanol) binary mixture at (53.3 and 91.3) kPa

pvap	51.10	kPa	353.25	Isobaric (vapour + liquid) equilibria for the (1-propanol + 1-butanol) binary mixture at (53.3 and 91.3) kPa
pvap	46.00	kPa	350.80	Isobaric (vapour + liquid) equilibria for the (1-propanol + 1-butanol) binary mixture at (53.3 and 91.3) kPa
pvap	37.40	kPa	346.15	Isobaric (vapour + liquid) equilibria for the (1-propanol + 1-butanol) binary mixture at (53.3 and 91.3) kPa
pvap	30.10	kPa	341.35	Isobaric (vapour + liquid) equilibria for the (1-propanol + 1-butanol) binary mixture at (53.3 and 91.3) kPa
pvap	25.60	kPa	337.95	Isobaric (vapour + liquid) equilibria for the (1-propanol + 1-butanol) binary mixture at (53.3 and 91.3) kPa
pvap	12.32	kPa	323.15	Vapor-Liquid Equilibria and HE for Binary Systems of Dimethyl Ether (DME) with C1-C4 Alkan-1-ols at 323.15 K and Liquid-Liquid Equilibria for Ternary System of DME + Methanol + Water at 313.15 K
pvap	19.07	kPa	331.92	Vapor-Liquid Equilibrium for 1-Propanol + 1-Butene, + cis-2-Butene, + 2-Methyl-propene, + trans-2-Butene, + n-Butane, and + 2-Methyl-propane

pvap	101.00	kPa	370.20	Heterogeneous azeotropic distillation for the separation of n-propanol + water mixture using n-propyl acetate as entrainer
pvap	49.80	kPa	352.68	P-x data for binary systems using a novel static total pressure apparatus
pvap	31.45	kPa	342.83	P-x data for binary systems using a novel static total pressure apparatus
pvap	6.99	kPa	313.17	P-x data for binary systems using a novel static total pressure apparatus
pvap	102.70	kPa	370.44	Phase equilibria on binary systems containing diethyl sulfide
pvap	91.60	kPa	367.51	Phase equilibria on binary systems containing diethyl sulfide
pvap	80.70	kPa	364.30	Phase equilibria on binary systems containing diethyl sulfide
pvap	71.40	kPa	361.25	Phase equilibria on binary systems containing diethyl sulfide
pvap	62.40	kPa	357.95	Phase equilibria on binary systems containing diethyl sulfide
pvap	52.70	kPa	353.96	Phase equilibria on binary systems containing diethyl sulfide
pvap	42.90	kPa	349.12	Phase equilibria on binary systems containing diethyl sulfide

pvap	31.20	kPa	342.07	Phase equilibria on binary systems containing diethyl sulfide
pvap	21.90	kPa	334.53	Phase equilibria on binary systems containing diethyl sulfide
pvap	6.99	kPa	313.15	Thermodynamic properties of mixtures containing ionic liquids Vapor pressures and activity coefficients of n-alcohols and benzene in binary mixtures with 1-methyl-3-butyl-imidazolium bis(trifluoromethyl-sulfonyl) imide
pvap	5.20	kPa	308.15	Thermodynamic properties of mixtures containing ionic liquids Vapor pressures and activity coefficients of n-alcohols and benzene in binary mixtures with 1-methyl-3-butyl-imidazolium bis(trifluoromethyl-sulfonyl) imide
pvap	3.83	kPa	303.15	Thermodynamic properties of mixtures containing ionic liquids Vapor pressures and activity coefficients of n-alcohols and benzene in binary mixtures with 1-methyl-3-butyl-imidazolium bis(trifluoromethyl-sulfonyl) imide

pvap	2.79	kPa	298.15	Thermodynamic properties of mixtures containing ionic liquids Vapor pressures and activity coefficients of n-alcohols and benzene in binary mixtures with 1-methyl-3-butyl-imidazolium bis(trifluoromethyl-sulfonyl) imide
pvap	9.31	kPa	318.15	Thermodynamics of mixtures with strongly negative deviations from Raoult s law Part 9. Vapor liquid equilibria for the system 1-propanol + di-n-propylamine at six temperatures between 293.15 and 318.15K
pvap	7.03	kPa	313.15	Thermodynamics of mixtures with strongly negative deviations from Raoult s law Part 9. Vapor liquid equilibria for the system 1-propanol + di-n-propylamine at six temperatures between 293.15 and 318.15K
pvap	10.00	kPa	319.65	P, rho, T Measurements and Isobaric Vapor-Liquid-Equilibria of the 1,3,3-Trimethyl-2-oxabicyclo[2,2,2]octane + Propan-1-ol Mixture: Cubic and Statistical Associating Fluid Theory-Based Equation of State Analysis

pvap	5.25	kPa	308.15	Thermodynamics of mixtures with strongly negative deviations from Raoult's law Part 9. Vapor liquid equilibria for the system 1-propanol + di-n-propylamine at six temperatures between 293.15 and 318.15K
pvap	20.00	kPa	332.95	P, rho, T Measurements and Isobaric Vapor-Liquid-Equilibria of the 1,3,3-Trimethyl-2-oxabicyclo[2,2,2]octane + Propan-1-ol Mixture: Cubic and Statistical Associating Fluid Theory-Based Equation of State Analysis
pvap	3.88	kPa	303.15	Thermodynamics of mixtures with strongly negative deviations from Raoult's law Part 9. Vapor liquid equilibria for the system 1-propanol + di-n-propylamine at six temperatures between 293.15 and 318.15K
pvap	2.81	kPa	298.15	Thermodynamics of mixtures with strongly negative deviations from Raoult's law Part 9. Vapor liquid equilibria for the system 1-propanol + di-n-propylamine at six temperatures between 293.15 and 318.15K
pvap	98.66	kPa	369.55	Isobaric Vapor-Liquid Equilibria for Tetrahydropyran and Alcohol Systems

pvap	93.32	kPa	368.08	Isobaric Vapor-Liquid Equilibria for Tetrahydropyran and Alcohol Systems
pvap	79.99	kPa	364.11	Isobaric Vapor-Liquid Equilibria for Tetrahydropyran and Alcohol Systems
pvap	66.66	kPa	359.56	Isobaric Vapor-Liquid Equilibria for Tetrahydropyran and Alcohol Systems
pvap	53.33	kPa	354.18	Isobaric Vapor-Liquid Equilibria for Tetrahydropyran and Alcohol Systems
pvap	40.00	kPa	347.55	Isobaric Vapor-Liquid Equilibria for Tetrahydropyran and Alcohol Systems
pvap	81.20	kPa	364.53	Vapor Liquid Equilibrium for the Systems trans-2-Butene + Methanol, + 1-Propanol, + 2-Propanol, + 2-Butanol, and + 2-Methyl-2-propanol at 364.5 K
pvap	101.30	kPa	370.40	Evaluation of the 2-Methoxyethanol as Entrainer in Ethanol Water and 1-Propanol Water Mixtures
pvap	600.00	kPa	427.20	Experimental Determination of Vapor Liquid Equilibria. Binary Systems of Methyl Acetate, Ethyl Acetate, and Propyl Acetate with 1-Propanol at 0.6 MPa

pvap	12.18	kPa	323.15	Vapor Pressures and Activity Coefficients of (1-Propanol + 1,8-Cineole) at 10 Temperatures between 278.15 K and 323.15 K
pvap	9.29	kPa	318.15	Vapor Pressures and Activity Coefficients of (1-Propanol + 1,8-Cineole) at 10 Temperatures between 278.15 K and 323.15 K
pvap	7.01	kPa	313.15	Vapor Pressures and Activity Coefficients of (1-Propanol + 1,8-Cineole) at 10 Temperatures between 278.15 K and 323.15 K
pvap	5.23	kPa	308.15	Vapor Pressures and Activity Coefficients of (1-Propanol + 1,8-Cineole) at 10 Temperatures between 278.15 K and 323.15 K
pvap	3.87	kPa	303.15	Vapor Pressures and Activity Coefficients of (1-Propanol + 1,8-Cineole) at 10 Temperatures between 278.15 K and 323.15 K
pvap	2.84	kPa	298.15	Vapor Pressures and Activity Coefficients of (1-Propanol + 1,8-Cineole) at 10 Temperatures between 278.15 K and 323.15 K
pvap	546.10	kPa	423.20	High-temperature vapour liquid equilibrium for the (water + alcohol) systems and modelling with SAFT-VR: 2. Water-1-propanol

pvap	2.04	kPa	293.15	Vapor Pressures and Activity Coefficients of (1-Propanol + 1,8-Cineole) at 10 Temperatures between 278.15 K and 323.15 K
pvap	1.45	kPa	288.15	Vapor Pressures and Activity Coefficients of (1-Propanol + 1,8-Cineole) at 10 Temperatures between 278.15 K and 323.15 K
pvap	1.01	kPa	283.15	Vapor Pressures and Activity Coefficients of (1-Propanol + 1,8-Cineole) at 10 Temperatures between 278.15 K and 323.15 K
pvap	2.03	kPa	293.15	Thermodynamics of mixtures with strongly negative deviations from Raoult s law Part 9. Vapor liquid equilibria for the system 1-propanol + di-n-propylamine at six temperatures between 293.15 and 318.15K
pvap	100.00	kPa	370.00	Isobaric Vapor Liquid Equilibria for the 1-Propanol + Ethylene Glycol Monopropyl Ether and 1-Butanol + Ethylene Glycol Monopropyl Ether Systems
pvap	80.00	kPa	364.20	Isobaric Vapor Liquid Equilibria for the 1-Propanol + Ethylene Glycol Monopropyl Ether and 1-Butanol + Ethylene Glycol Monopropyl Ether Systems

pvap	60.00	kPa	357.10	Isobaric Vapor Liquid Equilibria for the 1-Propanol + Ethylene Glycol Monopropyl Ether and 1-Butanol + Ethylene Glycol Monopropyl Ether Systems
pvap	101.33	kPa	370.05	P, rho, T Measurements and Isobaric Vapor-Liquid-Equilibria of the 1,3,3-Trimethyl-2-oxabicyclo[2,2,2]octane + Propan-1-ol Mixture: Cubic and Statistical Associating Fluid Theory-Based Equation of State Analysis
pvap	90.00	kPa	367.15	P, rho, T Measurements and Isobaric Vapor-Liquid-Equilibria of the 1,3,3-Trimethyl-2-oxabicyclo[2,2,2]octane + Propan-1-ol Mixture: Cubic and Statistical Associating Fluid Theory-Based Equation of State Analysis
pvap	85.00	kPa	365.65	P, rho, T Measurements and Isobaric Vapor-Liquid-Equilibria of the 1,3,3-Trimethyl-2-oxabicyclo[2,2,2]octane + Propan-1-ol Mixture: Cubic and Statistical Associating Fluid Theory-Based Equation of State Analysis
pvap	80.00	kPa	364.15	P, rho, T Measurements and Isobaric Vapor-Liquid-Equilibria of the 1,3,3-Trimethyl-2-oxabicyclo[2,2,2]octane + Propan-1-ol Mixture: Cubic and Statistical Associating Fluid Theory-Based Equation of State Analysis

pvap	75.00	kPa	362.55	P, rho, T Measurements and Isobaric Vapor-Liquid-Equilibria of the 1,3,3-Trimethyl-2-oxabicyclo[2,2,2]octane + Propan-1-ol Mixture: Cubic and Statistical Associating Fluid Theory-Based Equation of State Analysis
pvap	70.00	kPa	360.85	P, rho, T Measurements and Isobaric Vapor-Liquid-Equilibria of the 1,3,3-Trimethyl-2-oxabicyclo[2,2,2]octane + Propan-1-ol Mixture: Cubic and Statistical Associating Fluid Theory-Based Equation of State Analysis
pvap	66.66	kPa	359.65	P, rho, T Measurements and Isobaric Vapor-Liquid-Equilibria of the 1,3,3-Trimethyl-2-oxabicyclo[2,2,2]octane + Propan-1-ol Mixture: Cubic and Statistical Associating Fluid Theory-Based Equation of State Analysis
pvap	65.00	kPa	359.05	P, rho, T Measurements and Isobaric Vapor-Liquid-Equilibria of the 1,3,3-Trimethyl-2-oxabicyclo[2,2,2]octane + Propan-1-ol Mixture: Cubic and Statistical Associating Fluid Theory-Based Equation of State Analysis
pvap	60.00	kPa	357.05	P, rho, T Measurements and Isobaric Vapor-Liquid-Equilibria of the 1,3,3-Trimethyl-2-oxabicyclo[2,2,2]octane + Propan-1-ol Mixture: Cubic and Statistical Associating Fluid Theory-Based Equation of State Analysis

pvap	55.00	kPa	355.05	P, rho, T Measurements and Isobaric Vapor-Liquid-Equilibria of the 1,3,3-Trimethyl-2-oxabicyclo[2,2,2]octane + Propan-1-ol Mixture: Cubic and Statistical Associating Fluid Theory-Based Equation of State Analysis
pvap	50.00	kPa	352.75	P, rho, T Measurements and Isobaric Vapor-Liquid-Equilibria of the 1,3,3-Trimethyl-2-oxabicyclo[2,2,2]octane + Propan-1-ol Mixture: Cubic and Statistical Associating Fluid Theory-Based Equation of State Analysis
pvap	45.00	kPa	350.35	P, rho, T Measurements and Isobaric Vapor-Liquid-Equilibria of the 1,3,3-Trimethyl-2-oxabicyclo[2,2,2]octane + Propan-1-ol Mixture: Cubic and Statistical Associating Fluid Theory-Based Equation of State Analysis
pvap	40.00	kPa	347.65	P, rho, T Measurements and Isobaric Vapor-Liquid-Equilibria of the 1,3,3-Trimethyl-2-oxabicyclo[2,2,2]octane + Propan-1-ol Mixture: Cubic and Statistical Associating Fluid Theory-Based Equation of State Analysis

pvap	35.00	kPa	344.65	P, rho, T Measurements and Isobaric Vapor-Liquid-Equilibria of the 1,3,3-Trimethyl-2-oxabicyclo[2,2,2]octane + Propan-1-ol Mixture: Cubic and Statistical Associating Fluid Theory-Based Equation of State Analysis
pvap	33.33	kPa	343.55	P, rho, T Measurements and Isobaric Vapor-Liquid-Equilibria of the 1,3,3-Trimethyl-2-oxabicyclo[2,2,2]octane + Propan-1-ol Mixture: Cubic and Statistical Associating Fluid Theory-Based Equation of State Analysis
pvap	30.00	kPa	341.25	P, rho, T Measurements and Isobaric Vapor-Liquid-Equilibria of the 1,3,3-Trimethyl-2-oxabicyclo[2,2,2]octane + Propan-1-ol Mixture: Cubic and Statistical Associating Fluid Theory-Based Equation of State Analysis
pvap	25.00	kPa	337.45	P, rho, T Measurements and Isobaric Vapor-Liquid-Equilibria of the 1,3,3-Trimethyl-2-oxabicyclo[2,2,2]octane + Propan-1-ol Mixture: Cubic and Statistical Associating Fluid Theory-Based Equation of State Analysis
pvap	49.36	kPa	352.44	Vapor-Liquid Equilibrium for Benzene + 2-Methylpentane and Allyl Alcohol + 1-Propanol

pvap	0.70	kPa	278.15	Vapor Pressures and Activity Coefficients of (1-Propanol + 1,8-Cineole) at 10 Temperatures between 278.15 K and 323.15 K
pvap	4.00	kPa	303.15	Vapor-Pressure Measurements of Liquid Solutions at Different Temperatures: Apparatus for Use over an Extended Temperature Range and Some New Data
rfi	1.37660		313.15	Thermophysical properties of {(+)-linalool + propan-1-ol}: A first stage towards the development of a green process
rfi	1.37929		308.15	Experimental Determination and Modeling of Densities and Refractive Indices of the Binary Systems Alcohol + Dicyclohexylamine at T = (288.15 to 323.15) K
rfi	1.38308		298.15	Isobaric Vapor-Liquid Equilibria at 101.32 kPa and Densities, Speeds of Sound, and Refractive Indices at 298.15 K for MTBE or DIPE or TAME + 1-Propanol Binary Systems
rfi	1.38310		298.15	Synthesis and Physical Properties of 1-Ethylpyridinium Ethylsulfate and its Binary Mixtures with Ethanol and 1-Propanol at Several Temperatures

rfi	1.38370	298.15	Experimental Determination, Correlation, and Prediction of Physical Properties of the Ternary Mixtures Ethanol and 1-Propanol + Water + 1-Ethyl-3-methylpyridinium Ethylsulfate at 298.15 K
rfi	1.38530	293.15	Measurement and Correlation of the Solubilities of m-Phthalic Acid in Monobasic Alcohols
rfi	1.38309	298.15	Density and Viscosity Experimental Data of the Ternary Mixtures 1-Propanol or 2-Propanol + Water + 1-Ethyl-3-methylimidazolium Ethylsulfate. Correlation and Prediction of Physical Properties of the Ternary Systems
rfi	1.38305	298.15	Experimental Liquid-Liquid Equilibria of 1-Alkyl-3-methylimidazolium Hexafluorophosphate with 1-Alcohols
rfi	1.38310	298.15	Physical Properties of Binary Mixtures of the Ionic Liquid 1-Ethyl-3-methylimidazolium Ethyl Sulfate with Several Alcohols at T = (298.15, 313.15, and 328.15) K and Atmospheric Pressure
rfi	1.38540	295.15	Isobaric Vapor Liquid Equilibrium Data for Binary Mixtures of 1-Phenylethanone with a Few Alcohols at 95.2 kPa

rfi	1.38510	293.15	Isobaric Vapor Liquid Equilibrium Data of 2-Methyl-propan-2-ol (1) + Heptan-1-ol (2), Methanol (1) + Heptan-1-ol (2), Ethanol (1) + Heptan-1-ol (2), and Propan-1-ol (1) + Heptan-1-ol (2) at 96.5 kPa
rfi	1.38510	293.20	Vapor Liquid Equilibrium Data for Binary Systems of 1H-Pyrrole with Butan-1-ol, Propan-1-ol, or Pentan-1-ol
rfi	1.38510	293.20	Vapor Liquid Equilibrium Data for Binary Systems of 1-Methyl-4-(1-methylethenyl)-cyclohexene + {Ethanol, Propan-1-ol, Propan-2-ol, Butan-1-ol, Pentan-1-ol, or Hexan-1-ol} at 40 kPa
rfi	1.38200	293.15	Isothermal Vapor Liquid Equilibrium Data for the Propan-1-ol + Dodecane System at (323.0, 343.4, 353.2, 363.1, and 369.2) K
rfi	1.38340	298.15	Density, Viscosity, and Refractive Index Properties for the Binary Mixtures of n-Butylammonium Acetate Ionic Liquid + Alkanols at Several Temperatures
rfi	1.38330	298.15	Isothermal Vapor Liquid Equilibrium Data of Propan-1-ol + 2,2,4-Trimethylpentane and Butan-1-ol + 2,2,4-Trimethylpentane at 318.15 K

rfi	1.38305	298.15	Liquid Liquid Phase Equilibria of 1-Propanol + Water + n-Alcohol Ternary Systems at 298.15 K and Atmospheric Pressure
rfi	1.38510	293.15	Vapor Liquid Equilibrium Data for Binary Systems of n-Dodecane + {Propan-1-ol, Butan-1-ol, 2-Methylpropan-1-ol} at 40 kPa
rfi	1.38300	298.15	Excess Molar Enthalpies of 2-Methyl-2-butanol (1) + 1-Alkanols (C1-C5) (2) at 298.15 K
rfi	1.38560	298.15	Isobaric Vapor-Liquid Equilibria for Binary and Ternary Mixtures of Propanal, Propanol, and Propanoic Acid
rfi	1.38552	293.15	Excess Volume of the 1-Propanol + 1-Alkene Systems in Terms of an Equation of State with Association
rfi	1.38330	298.15	Excess Molar Enthalpies of Benzyl Alcohol + Alkanols (C1-C6) and Their Correlations at 298.15 K and Ambient Pressure
rfi	1.38360	298.15	Phase Equilibria Involved in Extractive Distillation of Dipropyl Ether + 1-Propyl Alcohol Using N,N-Dimethylformamide as Entrainer

rfi	1.38360	298.15	Isobaric Vapor-Liquid Equilibria for Binary and Ternary Mixtures of Dipropyl Ether, 1-Propyl Alcohol, and Butyl Propionate
rfi	1.38330	298.15	Effect of Pressure on the Static Relative Permittivities of Alkan-1-ols at 298.15 K
rfi	1.38302	298.15	Physical Properties of Binary Mixtures of the Ionic Liquid 1-Methyl-3-octylimidazolium Chloride with Methanol, Ethanol, and 1-Propanol at T = (298.15, 313.15, and 328.15) K and at P) 0.1 MPa
rfi	1.38330	298.15	Excess Molar Enthalpies of 1,2-Propanediol + Alkan-1-ols (C1-C6) and Their Correlations at 298.15 K and Ambient Pressure (81.5 kPa)
rfi	1.38304	298.15	Density, Refractive Index, Speed of Sound at 298.15 K, and Vapor-Liquid Equilibria at 101.3 kPa for Binary Mixtures of Propanol + 2-Methyl-1-butanol and Propanol + 3-Methyl-1-butanol
rfi	1.38300	298.15	Density, Surface Tension, and Refractive Index of Octane + 1-Alkanol Mixtures at T) 298.15 K.

rfi	1.38550	293.15	Limiting Activity Coefficients by Comparative Tensimetry: 1-Propanol and 1-Butanol in Heptane and in Octane
rfi	1.37890	308.15	Excess Molar Volumes, Viscosities, and Refractive Indexes for Binary Mixtures of 1-Chlorobutane with Four Alcohols at T = (288.15, 298.15 and 308.15) K
rfi	1.38290	298.15	Excess Molar Volumes, Viscosities, and Refractive Indexes for Binary Mixtures of 1-Chlorobutane with Four Alcohols at T = (288.15, 298.15 and 308.15) K
rfi	1.38690	288.15	Excess Molar Volumes, Viscosities, and Refractive Indexes for Binary Mixtures of 1-Chlorobutane with Four Alcohols at T = (288.15, 298.15 and 308.15) K
rfi	1.38070	298.15	Bubble Temperatures of the Binary Mixtures of Dimethylcarbonate with Some Alcohols at 95.8 kPa
rfi	1.38540	293.15	Solubility Data for Roflumilast and Maraviroc in Various Solvents between T = (278.2-323.2) K

rfi	1.38520	293.15	Solid-Liquid Equilibrium Measurements for Posaconazole and Voriconazole in Several Solvents between T = 278.2 and 323.2 K Using Differential Thermal Analysis/Thermal Gravimetric Analysis
rfi	1.38530	293.15	Isothermal Vapor-Liquid Equilibrium Measurements for Alcohol + Water/n-Hexane Azeotropic Systems Using Both Dynamic and Automated Static-Synthetic Methods
rfi	1.38480	293.15	Isothermal Vapor-Liquid Equilibrium Measurements for the Pentan-2-one + Propan-1-ol/Butan-1-ol System within 342-363 K
rfi	1.38520	293.15	P-rho-T Data and Modeling for Propan-1-ol + n-Octane or n-Nonane or n-Decane from 313.15 K to 363.15 K and 1 MPa to 20 MPa
rfi	1.38340	298.00	Determination of Physicochemical Parameters of Sodium Dodecyl Sulfate in Aqueous Micellar Solutions Containing Short-Chain Alcohols
rfi	1.38512	293.15	Vapor Liquid Equilibrium for Methyl Isobutyl Ketone (MIBK) + (1-Propanol or 2-Propanol) Binary Mixtures

rfi	1.38350	298.15	Molar excess enthalpies and molar excess volumes of formamide + 1-propanol or 2-propanol and thermodynamic modeling by Prigogine-Flory-Patterson theory and Treszczanowicz-Benson association model
rfi	1.38306	298.15	Effect of anion fluorination in 1-ethyl-3-methylimidazolium as solvent for the liquid extraction of ethanol from ethyl tert-butyl ether
rfi	1.38370	298.15	Activity coefficients of the binary mixtures of o-cresol or p-cresol with C1-C4 aliphatic alcohols near ambient pressure
rfi	1.38360	298.15	Isobaric vapor-liquid equilibria for the binary systems 1-propyl alcohol + dipropyl ether and 1-butyl alcohol + dibutyl ether at 20 and 101.3 kPa
rfi	1.38360	298.15	Vapor liquid equilibria in the ternary system dipropyl ether + 1-propanol + 1-pentanol and the binary systems dipropyl ether + 1-pentanol, 1-propanol + 1-pentanol at 101.3 kPa
rfi	1.38390	298.20	Vapor liquid equilibria for the ternary mixture of carbon dioxide + 1-propanol + propyl acetate at elevated pressures

rfi	1.38500	293.15	Phase equilibria of water + 1-propanol + solvent (n-amyl acetate, cyclohexanol, and cyclohexyl acetate) at T = 298.2K
rfi	1.38360	298.15	Phase equilibria involved in extractive distillation of dipropyl ether + 1-propyl alcohol using 2-ethoxyethanol as entrainer
rfi	1.38333	298.15	Vapor liquid equilibrium, densities, and interfacial tensions for the system ethyl 1,1-dimethylethyl ether (ETBE) + propan-1-ol
rfi	1.38360	298.15	Liquid liquid equilibria of the systems dipropyl ether + n-propanol + water and dipropyl ether + n-propanol + ethylene glycol at different temperatures
rfi	1.38305	298.15	Isobaric vapour liquid equilibria for binary systems of 2-butanone with ethanol, 1-propanol, and 2-propanol at 20 and 101.3 kPa
rfi	1.38312	298.15	Isothermal vapor liquid equilibrium at 323.15K and excess molar volumes and refractive indices at 298.15K for the ternary system propyl vinyl ether + 1-propanol + benzene and its binary sub-systems

rfi	1.38350	298.20	Measurement and prediction of tie-line data for mixtures of (water + 1-propanol + diisopropyl ether): LLE diagrams as a function of temperature
rfi	1.38320	298.15	Excess molar volumes, excess molar enthalpies and refractive index deviations for binary mixtures of propan-1-ol, butan-1-ol and pentan-1-ol with 2,2,4-trimethylpentane at 298.15 K
rfi	1.38290	298.15	Experimental and predicted volumetric and refractive index properties of ternary mixtures of iodoethane with toluene and alcohols at temperature 298.15 K and pressure 101 kPa
rfi	1.38340	298.15	Excess molar volumes and partial molar volumes for (propionitrile + an alkanol) at T = 298.15 K and p = 0.1 MPa
rfi	1.38400	298.15	Physico-chemical and excess properties of the binary mixtures of methylcyclohexane + ethanol, + propan-1-ol, + propan-2-ol, + butan-1-ol, + 2-methyl-1-propanol, or 3-methyl-1-butanol at T = (298.15, 303.15, and 308.15) K

rfi	1.38190	303.15	Physico-chemical and excess properties of the binary mixtures of methylcyclohexane + ethanol, + propan-1-ol, + propan-2-ol, + butan-1-ol, + 2-methyl-1-propanol, or 3-methyl-1-butanol at T = (298.15, 303.15, and 308.15) K
rfi	1.38000	308.15	Physico-chemical and excess properties of the binary mixtures of methylcyclohexane + ethanol, + propan-1-ol, + propan-2-ol, + butan-1-ol, + 2-methyl-1-propanol, or 3-methyl-1-butanol at T = (298.15, 303.15, and 308.15) K
rfi	1.38260	298.15	Thermodynamic properties of (tetradecane + benzene, + toluene, + chlorobenzene, + bromobenzene, + anisole) binary mixtures at T = (298.15, 303.15, and 308.15) K
rfi	1.38080	303.15	Thermodynamic properties of (tetradecane + benzene, + toluene, + chlorobenzene, + bromobenzene, + anisole) binary mixtures at T = (298.15, 303.15, and 308.15) K
rfi	1.37820	308.15	Thermodynamic properties of (tetradecane + benzene, + toluene, + chlorobenzene, + bromobenzene, + anisole) binary mixtures at T = (298.15, 303.15, and 308.15) K

rfi	1.38370	298.15	(Vapor + liquid) equilibria of the binary mixtures of m-cresol with C1 C4 aliphatic alcohols at 95.5 kPa
rfi	1.38260	298.15	Thermodynamic interactions in binary mixtures of anisole with ethanol, propan-1-ol, propan-2-ol, butan-1-ol, pentan-1-ol, and 3-methylbutan-1-ol at T = (298.15, 303.15, and 308.15) K
rfi	1.38080	303.15	Thermodynamic interactions in binary mixtures of anisole with ethanol, propan-1-ol, propan-2-ol, butan-1-ol, pentan-1-ol, and 3-methylbutan-1-ol at T = (298.15, 303.15, and 308.15) K
rfi	1.37820	308.15	Thermodynamic interactions in binary mixtures of anisole with ethanol, propan-1-ol, propan-2-ol, butan-1-ol, pentan-1-ol, and 3-methylbutan-1-ol at T = (298.15, 303.15, and 308.15) K
rfi	1.38307	298.15	Thermophysical properties of the binary mixtures (1,8-cineole + 1-alkanol) at T = (298.15 and 313.15) K and at atmospheric pressure
rfi	1.37674	313.15	Thermophysical properties of the binary mixtures (1,8-cineole + 1-alkanol) at T = (298.15 and 313.15) K and at atmospheric pressure

rfi	1.38906	283.15	Thermophysical properties of {(+)-linalool + propan-1-ol}: A first stage towards the development of a green process
rfi	1.38288	298.15	Thermophysical properties of {(+)-linalool + propan-1-ol}: A first stage towards the development of a green process
rfi	1.37252	323.15	Experimental Determination and Modeling of Densities and Refractive Indices of the Binary Systems Alcohol + Dicyclohexylamine at T = (288.15 to 323.15) K
rfi	1.37015	328.15	Thermophysical properties of {(+)-linalool + propan-1-ol}: A first stage towards the development of a green process
rfi	1.38220	298.15	Application of the ERAS model to volumetric properties of binary mixtures of banana oil with primary and secondary alcohols (C1- C4) at different temperatures
rfi	1.38307	298.15	Effect of the temperature on the physical properties of pure 1-propyl-3-methylimidazolium bis(trifluoromethylsulfonyl)imide and characterization of its binary mixtures with alcohols

rfi	1.38210	303.15	Experimental study on the calorimetric data of cyclohexanol with alkanols (C1-C4) and correlation with Wilson, NRTL and UNIQUAC models at 300 K
rfi	1.38780	288.15	Experimental determination and modeling of excess molar volumes, viscosities and refractive indices of the binary systems (pyridine + 1-propanol, + 1,2-propanediol, + 1,3-propanediol, and + glycerol). New UNIFAC-VISCO parameters determination
rfi	1.38576	293.15	Experimental determination and modeling of excess molar volumes, viscosities and refractive indices of the binary systems (pyridine + 1-propanol, + 1,2-propanediol, + 1,3-propanediol, and + glycerol). New UNIFAC-VISCO parameters determination
rfi	1.38369	298.15	Experimental determination and modeling of excess molar volumes, viscosities and refractive indices of the binary systems (pyridine + 1-propanol, + 1,2-propanediol, + 1,3-propanediol, and + glycerol). New UNIFAC-VISCO parameters determination

rfi 1.38162 303.15 Experimental determination and modeling of excess molar volumes, viscosities and refractive indices of the binary systems (pyridine + 1-propanol, + 1,2-propanediol, + 1,3-propanediol, and + glycerol).
New UNIFAC-VISCO parameters determination

rfi 1.37918 308.15 Experimental determination and modeling of excess molar volumes, viscosities and refractive indices of the binary systems (pyridine + 1-propanol, + 1,2-propanediol, + 1,3-propanediol, and + glycerol).
New UNIFAC-VISCO parameters determination

rfi 1.37695 313.15 Experimental determination and modeling of excess molar volumes, viscosities and refractive indices of the binary systems (pyridine + 1-propanol, + 1,2-propanediol, + 1,3-propanediol, and + glycerol).
New UNIFAC-VISCO parameters determination

rfi	1.37469	318.15	Experimental determination and modeling of excess molar volumes, viscosities and refractive indices of the binary systems (pyridine + 1-propanol, + 1,2-propanediol, + 1,3-propanediol, and + glycerol). New UNIFAC-VISCO parameters determination
rfi	1.37231	323.15	Experimental determination and modeling of excess molar volumes, viscosities and refractive indices of the binary systems (pyridine + 1-propanol, + 1,2-propanediol, + 1,3-propanediol, and + glycerol). New UNIFAC-VISCO parameters determination
rfi	1.38334	298.15	Densities, speeds of sound, and refractive indices for binary mixtures of 1-butyl-3-methylimidazolium methyl sulphate ionic liquid with alcohols at T = (298.15, 303.15, 308.15, and 313.15) K
rfi	1.38131	303.15	Densities, speeds of sound, and refractive indices for binary mixtures of 1-butyl-3-methylimidazolium methyl sulphate ionic liquid with alcohols at T = (298.15, 303.15, 308.15, and 313.15) K

rfi	1.37926	308.15	Densities, speeds of sound, and refractive indices for binary mixtures of 1-butyl-3-methylimidazolium methyl sulphate ionic liquid with alcohols at T = (298.15, 303.15, 308.15, and 313.15) K
rfi	1.37722	313.15	Densities, speeds of sound, and refractive indices for binary mixtures of 1-butyl-3-methylimidazolium methyl sulphate ionic liquid with alcohols at T = (298.15, 303.15, 308.15, and 313.15) K
rfi	1.38330	298.15	Properties of pure n-butylammonium nitrate ionic liquid and its binary mixtures of with alcohols at T = (293.15 to 313.15) K
rfi	1.38432	293.15	Separation of thiophene from heptane with ionic liquids
rfi	1.38286	298.15	Effect of the temperature on the physical properties of the pure ionic liquid 1-ethyl-3-methylimidazolium methylsulfate and characterization of its binary mixtures with alcohols
rfi	1.38785	288.15	Experimental Determination and Modeling of Densities and Refractive Indices of the Binary Systems Alcohol + Dicyclohexylamine at T = (288.15 to 323.15) K

rfi	1.38578		293.15	Experimental Determination and Modeling of Densities and Refractive Indices of the Binary Systems Alcohol + Dicyclohexylamine at T = (288.15 to 323.15) K
rfi	1.38364		298.15	Experimental Determination and Modeling of Densities and Refractive Indices of the Binary Systems Alcohol + Dicyclohexylamine at T = (288.15 to 323.15) K
rfi	1.38153		303.15	Experimental Determination and Modeling of Densities and Refractive Indices of the Binary Systems Alcohol + Dicyclohexylamine at T = (288.15 to 323.15) K
rfi	1.37483		318.15	Experimental Determination and Modeling of Densities and Refractive Indices of the Binary Systems Alcohol + Dicyclohexylamine at T = (288.15 to 323.15) K
rfi	1.37702		313.15	Experimental Determination and Modeling of Densities and Refractive Indices of the Binary Systems Alcohol + Dicyclohexylamine at T = (288.15 to 323.15) K
rhol	799.56	kg/m ³	298.15	Thermophysical Characterization of the Mixtures of the Ionic Liquid 1-Ethyl-3-Methylimidazolium Acetate with 1-Propanol or 2-Propanol

rho1	787.72	kg/m3	313.15	Densities and Viscosities of Binary Mixtures Containing 1,3-Dimethylimidazolium Dimethylphosphate and Alcohols
rho1	783.56	kg/m3	318.15	Densities and Viscosities of Binary Mixtures Containing 1,3-Dimethylimidazolium Dimethylphosphate and Alcohols
rho1	779.36	kg/m3	323.15	Densities and Viscosities of Binary Mixtures Containing 1,3-Dimethylimidazolium Dimethylphosphate and Alcohols
rho1	807.75	kg/m3	288.15	Volumetric and Viscometric Study of Binary Systems of Ethyl Butyrate with Alcohols
rho1	803.77	kg/m3	293.15	Volumetric and Viscometric Study of Binary Systems of Ethyl Butyrate with Alcohols
rho1	799.76	kg/m3	298.15	Volumetric and Viscometric Study of Binary Systems of Ethyl Butyrate with Alcohols
rho1	795.72	kg/m3	303.15	Volumetric and Viscometric Study of Binary Systems of Ethyl Butyrate with Alcohols
rho1	791.64	kg/m3	308.15	Volumetric and Viscometric Study of Binary Systems of Ethyl Butyrate with Alcohols
rho1	787.52	kg/m3	313.15	Volumetric and Viscometric Study of Binary Systems of Ethyl Butyrate with Alcohols

rho	783.36	kg/m ³	318.15	Volumetric and Viscometric Study of Binary Systems of Ethyl Butyrate with Alcohols
rho	779.14	kg/m ³	323.15	Volumetric and Viscometric Study of Binary Systems of Ethyl Butyrate with Alcohols
rho	803.48	kg/m ³	293.15	Speed of sound, density and related thermodynamic excess properties of binary mixtures of butan-2-one with C1-C4 nalkanols and chloroform
rho	791.83	kg/m ³	308.15	Densities and Viscosities of Binary Mixtures Containing 1,3-Dimethylimidazolium Dimethylphosphate and Alcohols
rho	795.44	kg/m ³	303.15	Speed of sound, density and related thermodynamic excess properties of binary mixtures of butan-2-one with C1-C4 nalkanols and chloroform
rho	791.37	kg/m ³	308.15	Speed of sound, density and related thermodynamic excess properties of binary mixtures of butan-2-one with C1-C4 nalkanols and chloroform
rho	787.26	kg/m ³	313.15	Speed of sound, density and related thermodynamic excess properties of binary mixtures of butan-2-one with C1-C4 nalkanols and chloroform

rho1	783.10	kg/m3	318.15	Speed of sound, density and related thermodynamic excess properties of binary mixtures of butan-2-one with C1-C4 nalkanols and chloroform
rho1	778.90	kg/m3	323.15	Speed of sound, density and related thermodynamic excess properties of binary mixtures of butan-2-one with C1-C4 nalkanols and chloroform
rho1	804.97	kg/m3	293.15	Densities of Ionic Liquids, 1-Butyl-3-methylimidazolium Hexafluorophosphate and 1-Butyl-3-methylimidazolium Tetrafluoroborate, with Benzene, Acetonitrile, and 1-Propanol at T = (293.15 to 343.15) K
rho1	801.01	kg/m3	298.15	Densities of Ionic Liquids, 1-Butyl-3-methylimidazolium Hexafluorophosphate and 1-Butyl-3-methylimidazolium Tetrafluoroborate, with Benzene, Acetonitrile, and 1-Propanol at T = (293.15 to 343.15) K
rho1	796.89	kg/m3	303.15	Densities of Ionic Liquids, 1-Butyl-3-methylimidazolium Hexafluorophosphate and 1-Butyl-3-methylimidazolium Tetrafluoroborate, with Benzene, Acetonitrile, and 1-Propanol at T = (293.15 to 343.15) K

rho	792.83	kg/m ³	308.15	Densities of Ionic Liquids, 1-Butyl-3-methylimidazolium Hexafluorophosphate and 1-Butyl-3-methylimidazolium Tetrafluoroborate, with Benzene, Acetonitrile, and 1-Propanol at T = (293.15 to 343.15) K
rho	789.07	kg/m ³	313.15	Densities of Ionic Liquids, 1-Butyl-3-methylimidazolium Hexafluorophosphate and 1-Butyl-3-methylimidazolium Tetrafluoroborate, with Benzene, Acetonitrile, and 1-Propanol at T = (293.15 to 343.15) K
rho	780.74	kg/m ³	323.15	Densities of Ionic Liquids, 1-Butyl-3-methylimidazolium Hexafluorophosphate and 1-Butyl-3-methylimidazolium Tetrafluoroborate, with Benzene, Acetonitrile, and 1-Propanol at T = (293.15 to 343.15) K
rho	771.75	kg/m ³	333.15	Densities of Ionic Liquids, 1-Butyl-3-methylimidazolium Hexafluorophosphate and 1-Butyl-3-methylimidazolium Tetrafluoroborate, with Benzene, Acetonitrile, and 1-Propanol at T = (293.15 to 343.15) K
rho	762.86	kg/m ³	343.15	Densities of Ionic Liquids, 1-Butyl-3-methylimidazolium Hexafluorophosphate and 1-Butyl-3-methylimidazolium Tetrafluoroborate, with Benzene, Acetonitrile, and 1-Propanol at T = (293.15 to 343.15) K

rho	799.70	kg/m ³	298.15	Thermodynamic Properties of 1-Butyl-3-methylpyridinium Tetrafluoroborate and Its Mixtures with Water and Alkanols
rho	783.16	kg/m ³	318.15	Thermodynamic Properties of 1-Butyl-3-methylpyridinium Tetrafluoroborate and Its Mixtures with Water and Alkanols
rho	796.55	kg/m ³	303.15	Isothermal VLE and VE at 303.15 K for the Binary and Ternary Mixtures of Di-isopropyl Ether (DIPE) + 1-Propanol + 2,2,4-Trimethylpentane
rho	799.52	kg/m ³	298.15	Densities and Excess Molar Volumes of N-Methylmorpholine + 1-Alkanol Systems at 298.15 K
rho	799.94	kg/m ³	298.15	Liquid Liquid Equilibrium for Ternary Systems of Propyl Vinyl Ether + C ₃ or C ₄ Alcohols + Water at 298.15 K and Excess Molar Enthalpies for Ternary and Constituent Binary Systems of Propyl Vinyl Ether + Ethanol + Isooctane at 303.15 K
rho	803.59	kg/m ³	293.15	Volumetric Properties of Binary and Ternary Liquid Mixtures of 1-Propanol (1) + 2-Propanol (2) + Water (3) at Different Temperatures and Ambient Pressure (81.5 kPa)

rho	795.53	kg/m ³	303.15	Volumetric Properties of Binary and Ternary Liquid Mixtures of 1-Propanol (1) + 2-Propanol (2) + Water (3) at Different Temperatures and Ambient Pressure (81.5 kPa)
rho	787.38	kg/m ³	313.15	Volumetric Properties of Binary and Ternary Liquid Mixtures of 1-Propanol (1) + 2-Propanol (2) + Water (3) at Different Temperatures and Ambient Pressure (81.5 kPa)
rho	778.92	kg/m ³	323.15	Volumetric Properties of Binary and Ternary Liquid Mixtures of 1-Propanol (1) + 2-Propanol (2) + Water (3) at Different Temperatures and Ambient Pressure (81.5 kPa)
rho	799.98	kg/m ³	298.15	Binary Liquid-Liquid Equilibrium (LLE) for Dibutyl Ether (DBE) + Water from (288.15 to 318.15) K and Ternary LLE for Systems of DBE + C1 less than C4 Alcohols + Water at 298.15 K
rho	799.89	kg/m ³	298.15	Binary Liquid-Liquid Equilibrium (LLE) for Methyl tert-Amyl Ether (TAME) + Water from (288.15 to 313.15) K and Ternary LLE for Systems of TAME + C1-C4 Alcohols + Water at 298.15 K

rho1	799.81	kg/m3	298.15	Densities and Excess Molar Volumes of Cyclopentane (1) + 1-Alkanol (2) Systems at (298.15 and 308.15) K
rho1	791.69	kg/m3	308.15	Densities and Excess Molar Volumes of Cyclopentane (1) + 1-Alkanol (2) Systems at (298.15 and 308.15) K
rho1	800.61	kg/m3	298.15	Binary and Ternary Vapor-Liquid Equilibrium at 323.15 K and Excess Molar Volumes at 298.15 K for the Mixtures of Propyl Vinyl Ether + 1-Propanol + Toluene
rho1	799.70	kg/m3	298.15	Solubility and Liquid-Liquid Equilibrium of Aqueous Systems of Iodoethane with Methanol, Ethanol, or 1-Propanol at Temperature 298.15 K and Pressure 101.2 kPa
rho1	779.30	kg/m3	323.15	Liquid-Liquid Equilibria, Density, Viscosity, and Surface and Interfacial Tension of the System Water + n-Butyl Acetate + 1-Propanol at 323.15 K and Atmospheric Pressure

rhoI	800.70	kg/m ³	298.15	Apparent Molal Volumes and Viscosity B-Coefficients of Acetyl Salicylic Acid (2-Acetoxy Benzoic Acid) Solutions in Higher Alcohols at Different Temperatures
rhoI	799.56	kg/m ³	298.15	Refractive Indices and Deviations in Refractive Indices for Binary Mixtures of 1-Ethyl-3-methylimidazolium Trifluoromethanesulfonate with Methanol, Ethanol, 1-Propanol, and 2-Propanol at Several Temperatures
rhoI	795.87	kg/m ³	303.15	Densities and Viscosities of Binary Mixtures Containing 1,3-Dimethylimidazolium Dimethylphosphate and Alcohols
rhoI	799.67	kg/m ³	298.15	Densities and Viscosities of Binary Mixtures Containing 1,3-Dimethylimidazolium Dimethylphosphate and Alcohols
rhoI	803.65	kg/m ³	293.15	Densities and Viscosities of Binary Mixtures Containing 1,3-Dimethylimidazolium Dimethylphosphate and Alcohols
rhoI	770.32	kg/m ³	333.15	Measurement and Correlation of the Excess Properties of Ternary Mixture of {x ₁ [Hmim][BF ₄] + x ₂ 1-Propanol + x ₃ 2-Propanol} at Different Temperatures

rho	778.93	kg/m ³	323.15	Measurement and Correlation of the Excess Properties of Ternary Mixture of {x ₁ [Hmim][BF ₄] + x ₂ 1-Propanol + x ₃ 2-Propanol} at Different Temperatures
rho	787.30	kg/m ³	313.15	Measurement and Correlation of the Excess Properties of Ternary Mixture of {x ₁ [Hmim][BF ₄] + x ₂ 1-Propanol + x ₃ 2-Propanol} at Different Temperatures
rho	795.47	kg/m ³	303.15	Measurement and Correlation of the Excess Properties of Ternary Mixture of {x ₁ [Hmim][BF ₄] + x ₂ 1-Propanol + x ₃ 2-Propanol} at Different Temperatures
rho	803.52	kg/m ³	293.15	Measurement and Correlation of the Excess Properties of Ternary Mixture of {x ₁ [Hmim][BF ₄] + x ₂ 1-Propanol + x ₃ 2-Propanol} at Different Temperatures
rho	770.42	kg/m ³	333.15	Densities and Viscosities of Diaminotoluene with Water, Ethanol, Propan-1-ol, and Butan-1-ol from (293.15 to 333.15) K
rho	779.01	kg/m ³	323.15	Densities and Viscosities of Diaminotoluene with Water, Ethanol, Propan-1-ol, and Butan-1-ol from (293.15 to 333.15) K

rhoI	787.37	kg/m3	313.15	Densities and Viscosities of Diaminotoluene with Water, Ethanol, Propan-1-ol, and Butan-1-ol from (293.15 to 333.15) K
rhoI	795.54	kg/m3	303.15	Densities and Viscosities of Diaminotoluene with Water, Ethanol, Propan-1-ol, and Butan-1-ol from (293.15 to 333.15) K
rhoI	803.86	kg/m3	293.15	Densities and Viscosities of Diaminotoluene with Water, Ethanol, Propan-1-ol, and Butan-1-ol from (293.15 to 333.15) K
rhoI	764.00	kg/m3	333.15	Composition and Temperature Dependence of Density, Surface Tension, and Viscosity of EMIM DEP/MMIM DMP + Water + 1-Propanol/2-Propanol Ternary Mixtures and Their Mathematical Representation Using the Jouyban Acree Model
rhoI	780.00	kg/m3	323.15	Composition and Temperature Dependence of Density, Surface Tension, and Viscosity of EMIM DEP/MMIM DMP + Water + 1-Propanol/2-Propanol Ternary Mixtures and Their Mathematical Representation Using the Jouyban Acree Model

rho1 788.00 kg/m3 313.15 Composition and
Temperature
Dependence of
Density, Surface
Tension, and
Viscosity of
EMIM
DEP/MMIM DMP
+ Water +
1-Propanol/2-Propanol
Ternary Mixtures
and Their
Mathematical
Representation
Using the
Jouyban Acree
Model

rho1 796.00 kg/m3 303.15 Composition and
Temperature
Dependence of
Density, Surface
Tension, and
Viscosity of
EMIM
DEP/MMIM DMP
+ Water +
1-Propanol/2-Propanol
Ternary Mixtures
and Their
Mathematical
Representation
Using the
Jouyban Acree
Model

rho1 800.10 kg/m3 298.15 Composition and
Temperature
Dependence of
Density, Surface
Tension, and
Viscosity of
EMIM
DEP/MMIM DMP
+ Water +
1-Propanol/2-Propanol
Ternary Mixtures
and Their
Mathematical
Representation
Using the
Jouyban Acree
Model

rho	804.00	kg/m ³	293.15	Composition and Temperature Dependence of Density, Surface Tension, and Viscosity of EMIM DEP/MMIM DMP + Water + 1-Propanol/2-Propanol Ternary Mixtures and Their Mathematical Representation Using the Jouyban Acree Model
rho	779.32	kg/m ³	323.15	Volumetric and Transport Properties of Binary Mixtures of n-Octane + Ethanol, + 1-Propanol, + 1-Butanol, and + 1-Pentanol from (293.15 to 323.15) K at Atmospheric Pressure
rho	787.64	kg/m ³	313.15	Volumetric and Transport Properties of Binary Mixtures of n-Octane + Ethanol, + 1-Propanol, + 1-Butanol, and + 1-Pentanol from (293.15 to 323.15) K at Atmospheric Pressure
rho	791.74	kg/m ³	308.15	Volumetric and Transport Properties of Binary Mixtures of n-Octane + Ethanol, + 1-Propanol, + 1-Butanol, and + 1-Pentanol from (293.15 to 323.15) K at Atmospheric Pressure

rho	795.81	kg/m ³	303.15	Volumetric and Transport Properties of Binary Mixtures of n-Octane + Ethanol, + 1-Propanol, + 1-Butanol, and + 1-Pentanol from (293.15 to 323.15) K at Atmospheric Pressure
rho	799.85	kg/m ³	298.15	Volumetric and Transport Properties of Binary Mixtures of n-Octane + Ethanol, + 1-Propanol, + 1-Butanol, and + 1-Pentanol from (293.15 to 323.15) K at Atmospheric Pressure
rho	803.84	kg/m ³	293.15	Volumetric and Transport Properties of Binary Mixtures of n-Octane + Ethanol, + 1-Propanol, + 1-Butanol, and + 1-Pentanol from (293.15 to 323.15) K at Atmospheric Pressure
rho	770.38	kg/m ³	333.15	Density and Viscosity Measurements of Binary Alkanol Mixtures from (293.15 to 333.15) K at Atmospheric Pressure
rho	778.98	kg/m ³	323.15	Density and Viscosity Measurements of Binary Alkanol Mixtures from (293.15 to 333.15) K at Atmospheric Pressure

rho	787.35	kg/m ³	313.15	Density and Viscosity Measurements of Binary Alkanol Mixtures from (293.15 to 333.15) K at Atmospheric Pressure
rho	795.53	kg/m ³	303.15	Density and Viscosity Measurements of Binary Alkanol Mixtures from (293.15 to 333.15) K at Atmospheric Pressure
rho	803.57	kg/m ³	293.15	Density and Viscosity Measurements of Binary Alkanol Mixtures from (293.15 to 333.15) K at Atmospheric Pressure
rho	799.67	kg/m ³	298.15	Thermophysical Properties of the Pure Ionic Liquid 1-Butyl-1-methylpyrrolidinium Dicyanamide and Its Binary Mixtures with Alcohols
rho	799.70	kg/m ³	298.15	Excess Volumes of Ternary Mixtures 2,2,4-Trimethylpentane + Diisopropyl Ether or Methyl tert-Butyl Ether + Methanol, Ethanol, or 1-Propanol at 298.15 K
rho	799.56	kg/m ³	298.15	ACSExcess Molar Enthalpies of Mixtures of (+-)-Linalool with Several Alkanols
rho	804.27	kg/m ³	298.15	Phase Equilibria for Reactive Distillation of Propyl Propanoate. Pure Component Property Data, Vapor-Liquid Equilibria, and Liquid-Liquid Equilibria

rhoI	797.60	kg/m ³	298.15	Volumetric Properties for (Ionic Liquid + Methanol or Ethanol or 1-Propanol + Nitromethane) at 298.15 K and Atmospheric Pressure
rhoI	775.97	kg/m ³	328.15	Excess Molar Volumes of 1,3-Diethyl Propanedioate with Methanol, Ethanol, Propan-1-ol, Propan-2-ol, Butan-2-ol, 2-Methyl-propan-1-ol, and Pentan-1-ol at T = (288.15, 298.15, 313.15, and 328.15) K
rhoI	788.69	kg/m ³	313.15	Excess Molar Volumes of 1,3-Diethyl Propanedioate with Methanol, Ethanol, Propan-1-ol, Propan-2-ol, Butan-2-ol, 2-Methyl-propan-1-ol, and Pentan-1-ol at T = (288.15, 298.15, 313.15, and 328.15) K
rhoI	800.97	kg/m ³	298.15	Excess Molar Volumes of 1,3-Diethyl Propanedioate with Methanol, Ethanol, Propan-1-ol, Propan-2-ol, Butan-2-ol, 2-Methyl-propan-1-ol, and Pentan-1-ol at T = (288.15, 298.15, 313.15, and 328.15) K
rhoI	783.55	kg/m ³	318.15	Volumetric and Transport Properties of Binary Mixtures of n-Octane + Ethanol, + 1-Propanol, + 1-Butanol, and + 1-Pentanol from (293.15 to 323.15) K at Atmospheric Pressure

rho1	809.01	kg/m3	288.15	Excess Molar Volumes of 1,3-Diethyl Propanedioate with Methanol, Ethanol, Propan-1-ol, Propan-2-ol, Butan-2-ol, 2-Methyl-propan-1-ol, and Pentan-1-ol at T = (288.15, 298.15, 313.15, and 328.15) K
rho1	770.80	kg/m3	333.15	Densities and Excess Molar Volumes for Binary Glycerol + 1-Propanol, + 2-Propanol, + 1,2-Propanediol, and + 1,3-Propanediol Mixtures at Different Temperatures
rho1	774.90	kg/m3	328.15	Densities and Excess Molar Volumes for Binary Glycerol + 1-Propanol, + 2-Propanol, + 1,2-Propanediol, and + 1,3-Propanediol Mixtures at Different Temperatures
rho1	779.00	kg/m3	323.15	Densities and Excess Molar Volumes for Binary Glycerol + 1-Propanol, + 2-Propanol, + 1,2-Propanediol, and + 1,3-Propanediol Mixtures at Different Temperatures
rho1	783.30	kg/m3	318.15	Densities and Excess Molar Volumes for Binary Glycerol + 1-Propanol, + 2-Propanol, + 1,2-Propanediol, and + 1,3-Propanediol Mixtures at Different Temperatures

rho	787.30	kg/m ³	313.15	Densities and Excess Molar Volumes for Binary Glycerol + 1-Propanol, + 2-Propanol, + 1,2-Propanediol, and + 1,3-Propanediol Mixtures at Different Temperatures
rho	791.50	kg/m ³	308.15	Densities and Excess Molar Volumes for Binary Glycerol + 1-Propanol, + 2-Propanol, + 1,2-Propanediol, and + 1,3-Propanediol Mixtures at Different Temperatures
rho	795.50	kg/m ³	303.15	Densities and Excess Molar Volumes for Binary Glycerol + 1-Propanol, + 2-Propanol, + 1,2-Propanediol, and + 1,3-Propanediol Mixtures at Different Temperatures
rho	799.60	kg/m ³	298.15	Densities and Excess Molar Volumes for Binary Glycerol + 1-Propanol, + 2-Propanol, + 1,2-Propanediol, and + 1,3-Propanediol Mixtures at Different Temperatures
rho	795.50	kg/m ³	303.00	Fluid-Phase Behavior of {1-Hexyl-3-methylimidazolium Bis(trifluoromethylsulfonyl) Imide, [C6mim][NTf2], + C2-C8 n-Alcohol} Mixtures: Liquid-Liquid Equilibrium and Excess Volumes

rho1	795.50	kg/m3	303.00	Fluid-Phase Behavior of {1-Hexyl-3-methylimidazolium Bis(trifluoromethylsulfonyl) Imide, [C6mim][NTf2], + C2-C8 n-Alcohol} Mixtures: Liquid-Liquid Equilibrium and Excess Volumes
rho1	799.50	kg/m3	298.00	Fluid-Phase Behavior of {1-Hexyl-3-methylimidazolium Bis(trifluoromethylsulfonyl) Imide, [C6mim][NTf2], + C2-C8 n-Alcohol} Mixtures: Liquid-Liquid Equilibrium and Excess Volumes
rho1	799.50	kg/m3	298.00	Fluid-Phase Behavior of {1-Hexyl-3-methylimidazolium Bis(trifluoromethylsulfonyl) Imide, [C6mim][NTf2], + C2-C8 n-Alcohol} Mixtures: Liquid-Liquid Equilibrium and Excess Volumes
rho1	803.50	kg/m3	293.00	Fluid-Phase Behavior of {1-Hexyl-3-methylimidazolium Bis(trifluoromethylsulfonyl) Imide, [C6mim][NTf2], + C2-C8 n-Alcohol} Mixtures: Liquid-Liquid Equilibrium and Excess Volumes
rho1	803.50	kg/m3	293.00	Fluid-Phase Behavior of {1-Hexyl-3-methylimidazolium Bis(trifluoromethylsulfonyl) Imide, [C6mim][NTf2], + C2-C8 n-Alcohol} Mixtures: Liquid-Liquid Equilibrium and Excess Volumes

rhoI	799.75	kg/m3	298.15	Measurement and Correlation of Vapor-Liquid Equilibria at T) 333.15 K and Excess Molar Volumes at T) 298.15 K for Ethanol + Dimethyl Carbonate (DMC), DMC + 1-Propanol, and DMC + 1-Butanol
rhoI	799.50	kg/m3	298.15	Dynamic Viscosities of Diethyl Carbonate with Linear and Secondary Alcohols at Several Temperatures
rhoI	803.63	kg/m3	293.15	Excess Molar Enthalpies for Binary Mixtures of Ethanol + Acetone, + Octane, + Cyclohexane and 1-Propanol + Acetone, + Octane, + Heptane at 323.15
rhoI	779.20	kg/m3	323.15	Liquid Densities and Speed of Sound for Ionic Liquid (2-HEAA and 2-HDEAA) + Alcohol (1-Propanol and 2-Propanol) Mixtures at T = (293.15-323.15 K) and Atmospheric Pressure
rhoI	787.50	kg/m3	313.15	Liquid Densities and Speed of Sound for Ionic Liquid (2-HEAA and 2-HDEAA) + Alcohol (1-Propanol and 2-Propanol) Mixtures at T = (293.15-323.15 K) and Atmospheric Pressure

rho1	795.70	kg/m3	303.15	Liquid Densities and Speed of Sound for Ionic Liquid (2-HEAA and 2-HDEAA) + Alcohol (1-Propanol and 2-Propanol) Mixtures at T = (293.15-323.15 K) and Atmospheric Pressure
rho1	803.60	kg/m3	293.15	Liquid Densities and Speed of Sound for Ionic Liquid (2-HEAA and 2-HDEAA) + Alcohol (1-Propanol and 2-Propanol) Mixtures at T = (293.15-323.15 K) and Atmospheric Pressure
rho1	778.90	kg/m3	323.15	Density and Viscosity of 2-Butanol + (1-Propanol, 2-Propanol, or 3-Amino-1-propanol) Mixtures at Temperatures of (293.15 to 323.15) K: Application of the ERAS Model
rho1	783.10	kg/m3	318.15	Density and Viscosity of 2-Butanol + (1-Propanol, 2-Propanol, or 3-Amino-1-propanol) Mixtures at Temperatures of (293.15 to 323.15) K: Application of the ERAS Model
rho1	787.20	kg/m3	313.15	Density and Viscosity of 2-Butanol + (1-Propanol, 2-Propanol, or 3-Amino-1-propanol) Mixtures at Temperatures of (293.15 to 323.15) K: Application of the ERAS Model

rho1	791.30	kg/m3	308.15	Density and Viscosity of 2-Butanol + (1-Propanol, 2-Propanol, or 3-Amino-1-propanol) Mixtures at Temperatures of (293.15 to 323.15) K: Application of the ERAS Model
rho1	795.40	kg/m3	303.15	Density and Viscosity of 2-Butanol + (1-Propanol, 2-Propanol, or 3-Amino-1-propanol) Mixtures at Temperatures of (293.15 to 323.15) K: Application of the ERAS Model
rho1	799.40	kg/m3	298.15	Density and Viscosity of 2-Butanol + (1-Propanol, 2-Propanol, or 3-Amino-1-propanol) Mixtures at Temperatures of (293.15 to 323.15) K: Application of the ERAS Model
rho1	803.50	kg/m3	293.15	Density and Viscosity of 2-Butanol + (1-Propanol, 2-Propanol, or 3-Amino-1-propanol) Mixtures at Temperatures of (293.15 to 323.15) K: Application of the ERAS Model
rho1	783.12	kg/m3	318.15	Vapor-Liquid Equilibrium, Volumetric, and Compressibility Properties of 1-Propanol + Poly(ethylene glycol) Dimethyl Ether 250 and 500 Binary Mixtures

rho	791.40	kg/m ³	308.15	Vapor-Liquid Equilibrium, Volumetric, and Compressibility Properties of 1-Propanol + Poly(ethylene glycol) Dimethyl Ether 250 and 500 Binary Mixtures
rho	799.60	kg/m ³	298.15	Vapor-Liquid Equilibrium, Volumetric, and Compressibility Properties of 1-Propanol + Poly(ethylene glycol) Dimethyl Ether 250 and 500 Binary Mixtures
rho	761.53	kg/m ³	343.15	Density, Viscosity, and Speed of Sound of Pure and Binary Mixtures of Ionic Liquids Based on Sulfonium and Imidazolium Cations and Bis(trifluoromethylsulfonyl)imide Anion with 1-Propanol
rho	766.03	kg/m ³	338.15	Density, Viscosity, and Speed of Sound of Pure and Binary Mixtures of Ionic Liquids Based on Sulfonium and Imidazolium Cations and Bis(trifluoromethylsulfonyl)imide Anion with 1-Propanol
rho	770.44	kg/m ³	333.15	Density, Viscosity, and Speed of Sound of Pure and Binary Mixtures of Ionic Liquids Based on Sulfonium and Imidazolium Cations and Bis(trifluoromethylsulfonyl)imide Anion with 1-Propanol

rho1	774.79	kg/m3	328.15	Density, Viscosity, and Speed of Sound of Pure and Binary Mixtures of Ionic Liquids Based on Sulfonium and Imidazolium Cations and Bis(trifluoromethylsulfonyl)imide Anion with 1-Propanol
rho1	779.05	kg/m3	323.15	Density, Viscosity, and Speed of Sound of Pure and Binary Mixtures of Ionic Liquids Based on Sulfonium and Imidazolium Cations and Bis(trifluoromethylsulfonyl)imide Anion with 1-Propanol
rho1	783.27	kg/m3	318.15	Density, Viscosity, and Speed of Sound of Pure and Binary Mixtures of Ionic Liquids Based on Sulfonium and Imidazolium Cations and Bis(trifluoromethylsulfonyl)imide Anion with 1-Propanol
rho1	787.43	kg/m3	313.15	Density, Viscosity, and Speed of Sound of Pure and Binary Mixtures of Ionic Liquids Based on Sulfonium and Imidazolium Cations and Bis(trifluoromethylsulfonyl)imide Anion with 1-Propanol
rho1	791.54	kg/m3	308.15	Density, Viscosity, and Speed of Sound of Pure and Binary Mixtures of Ionic Liquids Based on Sulfonium and Imidazolium Cations and Bis(trifluoromethylsulfonyl)imide Anion with 1-Propanol

rho	795.61	kg/m ³	303.15	Density, Viscosity, and Speed of Sound of Pure and Binary Mixtures of Ionic Liquids Based on Sulfonium and Imidazolium Cations and Bis(trifluoromethylsulfonyl)imide Anion with 1-Propanol
rho	799.65	kg/m ³	298.15	Density, Viscosity, and Speed of Sound of Pure and Binary Mixtures of Ionic Liquids Based on Sulfonium and Imidazolium Cations and Bis(trifluoromethylsulfonyl)imide Anion with 1-Propanol
rho	799.47	kg/m ³	298.15	Speed of sound, density and related thermodynamic excess properties of binary mixtures of butan-2-one with C1-C4 nalkanols and chloroform
rho	803.65	kg/m ³	293.15	Density, Viscosity, and Speed of Sound of Pure and Binary Mixtures of Ionic Liquids Based on Sulfonium and Imidazolium Cations and Bis(trifluoromethylsulfonyl)imide Anion with 1-Propanol
rho	807.63	kg/m ³	288.15	Density, Viscosity, and Speed of Sound of Pure and Binary Mixtures of Ionic Liquids Based on Sulfonium and Imidazolium Cations and Bis(trifluoromethylsulfonyl)imide Anion with 1-Propanol

rho1	811.60	kg/m3	283.15	Density, Viscosity, and Speed of Sound of Pure and Binary Mixtures of Ionic Liquids Based on Sulfonium and Imidazolium Cations and Bis(trifluoromethylsulfonyl)imide Anion with 1-Propanol
rho1	815.52	kg/m3	278.15	Density, Viscosity, and Speed of Sound of Pure and Binary Mixtures of Ionic Liquids Based on Sulfonium and Imidazolium Cations and Bis(trifluoromethylsulfonyl)imide Anion with 1-Propanol
rho1	783.64	kg/m3	318.15	Densities and Excess Molar Volumes for the Binary and Ternary Systems of (1,4-Dioxane, 1-Propanol or 2-Propanol, and 1,2-Dichloroethane) at T = (288.15 to 318.15) K. Experimental Measurements and Prigogine-Flory-Patterson Modeling
rho1	791.71	kg/m3	308.15	Densities and Excess Molar Volumes for the Binary and Ternary Systems of (1,4-Dioxane, 1-Propanol or 2-Propanol, and 1,2-Dichloroethane) at T = (288.15 to 318.15) K. Experimental Measurements and Prigogine-Flory-Patterson Modeling

rho1	799.78	kg/m3	298.15	Densities and Excess Molar Volumes for the Binary and Ternary Systems of (1,4-Dioxane, 1-Propanol or 2-Propanol, and 1,2-Dichloroethane) at T = (288.15 to 318.15) K. Experimental Measurements and Prigogine-Flory-Patterson Modeling
rho1	807.86	kg/m3	288.15	Densities and Excess Molar Volumes for the Binary and Ternary Systems of (1,4-Dioxane, 1-Propanol or 2-Propanol, and 1,2-Dichloroethane) at T = (288.15 to 318.15) K. Experimental Measurements and Prigogine-Flory-Patterson Modeling
rho1	799.50	kg/m3	298.15	Isobaric Vapor-Liquid Phase Equilibrium Measurements, Correlation, and Prediction for Separation of the Mixtures of Cyclohexanone and Alcohols
rho1	803.60	kg/m3	293.15	Three Binary Azeotropic Systems for 1-(Methoxymethoxy)-propane, 1-(Ethoxymethoxy)-propane, and Methoxy(methoxymethoxy)methane with Three Alcohols at 101.33 kPa: Experimental Data, Correlation, and Purification

rho	783.12	kg/m ³	318.15	Measurement and Correlation of Activity, Density, and Speed of Sound for Binary Mixtures of 1-Propanol + Poly(Propylene Glycol) 400, 725, and 1025
rho	791.40	kg/m ³	308.15	Measurement and Correlation of Activity, Density, and Speed of Sound for Binary Mixtures of 1-Propanol + Poly(Propylene Glycol) 400, 725, and 1025
rho	799.60	kg/m ³	298.15	Measurement and Correlation of Activity, Density, and Speed of Sound for Binary Mixtures of 1-Propanol + Poly(Propylene Glycol) 400, 725, and 1025
rho	807.83	kg/m ³	288.15	Measurement and Correlation of Activity, Density, and Speed of Sound for Binary Mixtures of 1-Propanol + Poly(Propylene Glycol) 400, 725, and 1025
rho	802.00	kg/m ³	298.15	Isobaric Vapor-Liquid Equilibrium for the Binary Systems of Sec-butyl Acetate and Ethanol, 1-Propanol, or 2-Propanol at 101.3 kPa
rho	800.94	kg/m ³	298.15	Volumetric Properties, Viscosity, and Refractive Indices of Different Naringenin Solutions at Several Temperatures

rhoI	803.67	kg/m ³	293.15	Volumetric Properties, Viscosity, and Refractive Indices of Different Naringenin Solutions at Several Temperatures
rhoI	795.54	kg/m ³	303.15	Isobaric Vapor Liquid Equilibrium for the Binary Systems Dimethyl Disulfide + C ₁ C ₄ n-Alkanol at 40.000 and 101.325 kPa
rhoI	792.76	kg/m ³	308.15	Physical Properties of the Pure Deep Eutectic Solvent, [ChCl]:[Lev] (1:2) DES, and Its Binary Mixtures with Alcohols
rhoI	799.69	kg/m ³	298.15	Physical Properties of the Pure Deep Eutectic Solvent, [ChCl]:[Lev] (1:2) DES, and Its Binary Mixtures with Alcohols
rhoI	742.77	kg/m ³	363.15	Densities and Viscosities for Binary Liquid Mixtures of n-Undecane + 1-Propanol, + 1-Butanol, + 1-Pentanol, and + 1-Hexanol from 283.15 to 363.15 K at 0.1 MPa
rhoI	747.65	kg/m ³	358.15	Densities and Viscosities for Binary Liquid Mixtures of n-Undecane + 1-Propanol, + 1-Butanol, + 1-Pentanol, and + 1-Hexanol from 283.15 to 363.15 K at 0.1 MPa

rho	752.44	kg/m ³	353.15	Densities and Viscosities for Binary Liquid Mixtures of n-Undecane + 1-Propanol, + 1-Butanol, + 1-Pentanol, and + 1-Hexanol from 283.15 to 363.15 K at 0.1 MPa
rho	757.09	kg/m ³	348.15	Densities and Viscosities for Binary Liquid Mixtures of n-Undecane + 1-Propanol, + 1-Butanol, + 1-Pentanol, and + 1-Hexanol from 283.15 to 363.15 K at 0.1 MPa
rho	761.67	kg/m ³	343.15	Densities and Viscosities for Binary Liquid Mixtures of n-Undecane + 1-Propanol, + 1-Butanol, + 1-Pentanol, and + 1-Hexanol from 283.15 to 363.15 K at 0.1 MPa
rho	766.15	kg/m ³	338.15	Densities and Viscosities for Binary Liquid Mixtures of n-Undecane + 1-Propanol, + 1-Butanol, + 1-Pentanol, and + 1-Hexanol from 283.15 to 363.15 K at 0.1 MPa
rho	770.54	kg/m ³	333.15	Densities and Viscosities for Binary Liquid Mixtures of n-Undecane + 1-Propanol, + 1-Butanol, + 1-Pentanol, and + 1-Hexanol from 283.15 to 363.15 K at 0.1 MPa

rho	774.86	kg/m ³	328.15	Densities and Viscosities for Binary Liquid Mixtures of n-Undecane + 1-Propanol, + 1-Butanol, + 1-Pentanol, and + 1-Hexanol from 283.15 to 363.15 K at 0.1 MPa
rho	779.11	kg/m ³	323.15	Densities and Viscosities for Binary Liquid Mixtures of n-Undecane + 1-Propanol, + 1-Butanol, + 1-Pentanol, and + 1-Hexanol from 283.15 to 363.15 K at 0.1 MPa
rho	783.30	kg/m ³	318.15	Densities and Viscosities for Binary Liquid Mixtures of n-Undecane + 1-Propanol, + 1-Butanol, + 1-Pentanol, and + 1-Hexanol from 283.15 to 363.15 K at 0.1 MPa
rho	787.44	kg/m ³	313.15	Densities and Viscosities for Binary Liquid Mixtures of n-Undecane + 1-Propanol, + 1-Butanol, + 1-Pentanol, and + 1-Hexanol from 283.15 to 363.15 K at 0.1 MPa
rho	791.53	kg/m ³	308.15	Densities and Viscosities for Binary Liquid Mixtures of n-Undecane + 1-Propanol, + 1-Butanol, + 1-Pentanol, and + 1-Hexanol from 283.15 to 363.15 K at 0.1 MPa

rho	795.58	kg/m ³	303.15	Densities and Viscosities for Binary Liquid Mixtures of n-Undecane + 1-Propanol, + 1-Butanol, + 1-Pentanol, and + 1-Hexanol from 283.15 to 363.15 K at 0.1 MPa
rho	799.59	kg/m ³	298.15	Densities and Viscosities for Binary Liquid Mixtures of n-Undecane + 1-Propanol, + 1-Butanol, + 1-Pentanol, and + 1-Hexanol from 283.15 to 363.15 K at 0.1 MPa
rho	803.58	kg/m ³	293.15	Densities and Viscosities for Binary Liquid Mixtures of n-Undecane + 1-Propanol, + 1-Butanol, + 1-Pentanol, and + 1-Hexanol from 283.15 to 363.15 K at 0.1 MPa
rho	807.54	kg/m ³	288.15	Densities and Viscosities for Binary Liquid Mixtures of n-Undecane + 1-Propanol, + 1-Butanol, + 1-Pentanol, and + 1-Hexanol from 283.15 to 363.15 K at 0.1 MPa
rho	811.48	kg/m ³	283.15	Densities and Viscosities for Binary Liquid Mixtures of n-Undecane + 1-Propanol, + 1-Butanol, + 1-Pentanol, and + 1-Hexanol from 283.15 to 363.15 K at 0.1 MPa
rho	756.91	kg/m ³	348.15	Thermophysical Characterization of the Mixtures of the Ionic Liquid 1-Ethyl-3-Methylimidazolium Acetate with 1-Propanol or 2-Propanol

rho	765.98	kg/m ³	338.15	Thermophysical Characterization of the Mixtures of the Ionic Liquid 1-Ethyl-3-Methylimidazolium Acetate with 1-Propanol or 2-Propanol
rho	774.72	kg/m ³	328.15	Thermophysical Characterization of the Mixtures of the Ionic Liquid 1-Ethyl-3-Methylimidazolium Acetate with 1-Propanol or 2-Propanol
rho	783.20	kg/m ³	318.15	Thermophysical Characterization of the Mixtures of the Ionic Liquid 1-Ethyl-3-Methylimidazolium Acetate with 1-Propanol or 2-Propanol
rho	791.46	kg/m ³	308.15	Thermophysical Characterization of the Mixtures of the Ionic Liquid 1-Ethyl-3-Methylimidazolium Acetate with 1-Propanol or 2-Propanol
rho	807.55	kg/m ³	288.15	Thermophysical Characterization of the Mixtures of the Ionic Liquid 1-Ethyl-3-Methylimidazolium Acetate with 1-Propanol or 2-Propanol
rho	799.70	kg/m ³	298.15	Thermodynamics of 1,3-dimethylurea in eight alcohols
rho	778.93	kg/m ³	323.15	Densities and Viscosities of Binary Mixtures of 2-Ethyl-1,1,3,3-tetramethylguanidinium Ionic Liquids with Ethanol and 1-Propanol
rho	783.14	kg/m ³	318.15	Densities and Viscosities of Binary Mixtures of 2-Ethyl-1,1,3,3-tetramethylguanidinium Ionic Liquids with Ethanol and 1-Propanol

rho	787.29	kg/m ³	313.15	Densities and Viscosities of Binary Mixtures of 2-Ethyl-1,1,3,3-tetramethylguanidinium Ionic Liquids with Ethanol and 1-Propanol
rho	791.39	kg/m ³	308.15	Densities and Viscosities of Binary Mixtures of 2-Ethyl-1,1,3,3-tetramethylguanidinium Ionic Liquids with Ethanol and 1-Propanol
rho	795.46	kg/m ³	303.15	Densities and Viscosities of Binary Mixtures of 2-Ethyl-1,1,3,3-tetramethylguanidinium Ionic Liquids with Ethanol and 1-Propanol
rho	799.49	kg/m ³	298.15	Densities and Viscosities of Binary Mixtures of 2-Ethyl-1,1,3,3-tetramethylguanidinium Ionic Liquids with Ethanol and 1-Propanol
rho	803.49	kg/m ³	293.15	Densities and Viscosities of Binary Mixtures of 2-Ethyl-1,1,3,3-tetramethylguanidinium Ionic Liquids with Ethanol and 1-Propanol
rho	765.96	kg/m ³	338.15	Density, Viscosity, and Sound Speed of Bis(trifluoromethylsulfonyl)imide-Based Ionic Liquids + 1-Propanol Mixtures
rho	783.18	kg/m ³	318.15	Density, Viscosity, and Sound Speed of Bis(trifluoromethylsulfonyl)imide-Based Ionic Liquids + 1-Propanol Mixtures

rho1	799.55	kg/m3	298.15	Density, Viscosity, and Sound Speed of Bis(trifluoromethylsulfonyl)imide-Based Ionic Liquids + 1-Propanol Mixtures
rho1	815.46	kg/m3	278.15	Density, Viscosity, and Sound Speed of Bis(trifluoromethylsulfonyl)imide-Based Ionic Liquids + 1-Propanol Mixtures
rho1	801.08	kg/m3	298.15	Modified Method for Measuring the Solubility of Pharmaceutical Compounds in Organic Solvents by Visual Camera
rho1	783.15	kg/m3	318.15	The excess molar volume and the molar surface Gibbs energy of the binary of the ether-functionalized ionic liquids [C22O1IM][TfO] with ethanol and isomeric propanols at T = (288.15-318.15) K
rho1	787.31	kg/m3	313.15	The excess molar volume and the molar surface Gibbs energy of the binary of the ether-functionalized ionic liquids [C22O1IM][TfO] with ethanol and isomeric propanols at T = (288.15-318.15) K
rho1	791.42	kg/m3	308.15	The excess molar volume and the molar surface Gibbs energy of the binary of the ether-functionalized ionic liquids [C22O1IM][TfO] with ethanol and isomeric propanols at T = (288.15-318.15) K

rho1	795.48	kg/m3	303.15	The excess molar volume and the molar surface Gibbs energy of the binary of the ether-functionalized ionic liquids [C22O1IM][TfO] with ethanol and isomeric propanols at T = (288.15-318.15) K
rho1	799.51	kg/m3	298.15	The excess molar volume and the molar surface Gibbs energy of the binary of the ether-functionalized ionic liquids [C22O1IM][TfO] with ethanol and isomeric propanols at T = (288.15-318.15) K
rho1	803.52	kg/m3	293.15	The excess molar volume and the molar surface Gibbs energy of the binary of the ether-functionalized ionic liquids [C22O1IM][TfO] with ethanol and isomeric propanols at T = (288.15-318.15) K
rho1	807.51	kg/m3	288.15	The excess molar volume and the molar surface Gibbs energy of the binary of the ether-functionalized ionic liquids [C22O1IM][TfO] with ethanol and isomeric propanols at T = (288.15-318.15) K

rho1	778.70	kg/m3	323.15	Excess volumes, Viscosities, and Excess Gibbs Energy of Activation for Viscous Flow, for binary and ternary mixtures 1- propanol + N-N dimethylformamid + chloroform at different temperatures
rho1	787.00	kg/m3	313.15	Excess volumes, Viscosities, and Excess Gibbs Energy of Activation for Viscous Flow, for binary and ternary mixtures 1- propanol + N-N dimethylformamid + chloroform at different temperatures
rho1	795.10	kg/m3	303.15	Excess volumes, Viscosities, and Excess Gibbs Energy of Activation for Viscous Flow, for binary and ternary mixtures 1- propanol + N-N dimethylformamid + chloroform at different temperatures
rho1	803.30	kg/m3	293.15	Excess volumes, Viscosities, and Excess Gibbs Energy of Activation for Viscous Flow, for binary and ternary mixtures 1- propanol + N-N dimethylformamid + chloroform at different temperatures

rho1	779.14	kg/m3	323.15	Experimental measurements and modelling of volumetric properties, refractive index and viscosity of selected binary systems with butyl lactate at 288.15 to 323.15 K and atmospheric pressure. New UNIFAC-VISCO interaction parameters.
rho1	783.36	kg/m3	318.15	Experimental measurements and modelling of volumetric properties, refractive index and viscosity of selected binary systems with butyl lactate at 288.15 to 323.15 K and atmospheric pressure. New UNIFAC-VISCO interaction parameters.
rho1	787.52	kg/m3	313.15	Experimental measurements and modelling of volumetric properties, refractive index and viscosity of selected binary systems with butyl lactate at 288.15 to 323.15 K and atmospheric pressure. New UNIFAC-VISCO interaction parameters.

rho1	791.64	kg/m3	308.15	Experimental measurements and modelling of volumetric properties, refractive index and viscosity of selected binary systems with butyl lactate at 288.15 to 323.15 K and atmospheric pressure. New UNIFAC-VISCO interaction parameters.
rho1	795.72	kg/m3	303.15	Experimental measurements and modelling of volumetric properties, refractive index and viscosity of selected binary systems with butyl lactate at 288.15 to 323.15 K and atmospheric pressure. New UNIFAC-VISCO interaction parameters.
rho1	799.75	kg/m3	298.15	Experimental measurements and modelling of volumetric properties, refractive index and viscosity of selected binary systems with butyl lactate at 288.15 to 323.15 K and atmospheric pressure. New UNIFAC-VISCO interaction parameters.

rho1	803.77	kg/m3	293.15	Experimental measurements and modelling of volumetric properties, refractive index and viscosity of selected binary systems with butyl lactate at 288.15 to 323.15 K and atmospheric pressure. New UNIFAC-VISCO interaction parameters.
rho1	807.75	kg/m3	288.15	Experimental measurements and modelling of volumetric properties, refractive index and viscosity of selected binary systems with butyl lactate at 288.15 to 323.15 K and atmospheric pressure. New UNIFAC-VISCO interaction parameters.
rho1	787.95	kg/m3	313.15	Excess molar volumes of Diisopropylamine + (C1-C5) Alkan-1-ols: application of the ERAS model and cubic EOS
rho1	795.05	kg/m3	303.15	Excess molar volumes of Diisopropylamine + (C1-C5) Alkan-1-ols: application of the ERAS model and cubic EOS
rho1	797.81	kg/m3	298.15	Excess molar volumes of Diisopropylamine + (C1-C5) Alkan-1-ols: application of the ERAS model and cubic EOS

rho1	800.03	kg/m3	293.15	Excess molar volumes of Diisopropylamine + (C1-C5) Alkan-1-ols: application of the ERAS model and cubic EOS
rho1	799.28	kg/m3	298.15	Topological investigations of the molecular species and molecular interactions that characterize pyrrolidin-2-one + lower alkanol mixtures
rho1	795.61	kg/m3	303.15	Volumetric and transport properties of ternary mixtures containing 1-propanol + ethyl ethanoate + cyclohexane or benzene at 303.15 K: Experimental data, correlation and prediction by ERAS model
rho1	799.30	kg/m3	298.15	Thermodynamic properties of binary mixtures of 2,2,2-Trifluoroethanol with Water or Alkanols at T=298.15 K
rho1	795.48	kg/m3	303.15	Excess molar enthalpies and heat capacities of dimethyl sulfoxide + seven normal alkanols at 303.15K and atmospheric pressure
rho1	799.10	kg/m3	298.15	Excess enthalpies of binary mixtures of some propylamines + some propanols at 298.15K
rho1	799.51	kg/m3	298.15	Excess molar enthalpies of methyl isobutyl ketone (MIBK) with alkan-1-ols (C1-C6) and their correlations at 298.15 K

rho1	779.09	kg/m3	323.15	Vapor-liquid equilibrium and excess properties of the binary mixtures formed by ethyl isobutyrate and n-alkanols
rho1	795.53	kg/m3	303.15	Vapor-liquid equilibrium and excess properties of the binary mixtures formed by ethyl isobutyrate and n-alkanols
rho1	799.97	kg/m3	298.15	Excess volumes and partial molar volumes of binary liquid mixtures of furfural or 2-methylfuran with alcohols at 298.15 K
rho1	778.96	kg/m3	323.15	Volumetric properties of monoethanolamine and alcohol binary mixtures at different temperatures and 0.1 MPa
rho1	783.17	kg/m3	318.15	Volumetric properties of monoethanolamine and alcohol binary mixtures at different temperatures and 0.1 MPa
rho1	787.42	kg/m3	313.15	Volumetric properties of monoethanolamine and alcohol binary mixtures at different temperatures and 0.1 MPa
rho1	791.52	kg/m3	308.15	Volumetric properties of monoethanolamine and alcohol binary mixtures at different temperatures and 0.1 MPa

rho1	795.58	kg/m3	303.15	Volumetric properties of monoethanolamine and alcohol binary mixtures at different temperatures and 0.1 MPa
rho1	799.69	kg/m3	298.15	Volumetric properties of monoethanolamine and alcohol binary mixtures at different temperatures and 0.1 MPa
rho1	803.77	kg/m3	293.15	Volumetric properties of monoethanolamine and alcohol binary mixtures at different temperatures and 0.1 MPa
rho1	774.70	kg/m3	328.15	Mass density, sound velocity, mixing enthalpy, ¹ H NMR, Ab initio calculations and intermolecular interactions in binary mixtures of N-methylimidazole + water, +methanol, +ethanol, +1-propanol, +2-propanol
rho1	783.18	kg/m3	318.15	Mass density, sound velocity, mixing enthalpy, ¹ H NMR, Ab initio calculations and intermolecular interactions in binary mixtures of N-methylimidazole + water, +methanol, +ethanol, +1-propanol, +2-propanol

rhoI	791.44	kg/m3	308.15	Mass density, sound velocity, mixing enthalpy, ¹ H NMR, Ab initio calculations and intermolecular interactions in binary mixtures of N-methylimidazole + water, +methanol, +ethanol, +1-propanol, +2-propanol
rhoI	799.53	kg/m3	298.15	Mass density, sound velocity, mixing enthalpy, ¹ H NMR, Ab initio calculations and intermolecular interactions in binary mixtures of N-methylimidazole + water, +methanol, +ethanol, +1-propanol, +2-propanol
rhoI	807.50	kg/m3	288.15	Mass density, sound velocity, mixing enthalpy, ¹ H NMR, Ab initio calculations and intermolecular interactions in binary mixtures of N-methylimidazole + water, +methanol, +ethanol, +1-propanol, +2-propanol
rhoI	800.00	kg/m3	298.15	Density, speed of sound and refractive index of mixtures containing 2-phenoxyethanol with propanol or butanol at various temperatures

rho1	780.00	kg/m3	323.15	Density, speed of sound and refractive index of mixtures containing 2-phenoxyethanol with propanol or butanol at various temperatures
rho1	788.00	kg/m3	313.15	Density, speed of sound and refractive index of mixtures containing 2-phenoxyethanol with propanol or butanol at various temperatures
rho1	796.00	kg/m3	303.15	Density, speed of sound and refractive index of mixtures containing 2-phenoxyethanol with propanol or butanol at various temperatures
rho1	803.00	kg/m3	293.15	Density, speed of sound and refractive index of mixtures containing 2-phenoxyethanol with propanol or butanol at various temperatures
rho1	787.88	kg/m3	313.15	Intermolecular interactions in binary mixtures of 2-diethylethanolamine with 1-propanol and 1-butanol at different temperatures
rho1	795.82	kg/m3	303.15	Intermolecular interactions in binary mixtures of 2-diethylethanolamine with 1-propanol and 1-butanol at different temperatures

rho1	803.76	kg/m3	293.15	Intermolecular interactions in binary mixtures of 2-diethylethanolamine with 1-propanol and 1-butanol at different temperatures
rho1	791.55	kg/m3	308.15	Thermodynamic evidence for nano-heterogeneity in solutions of the macrocycle C-butylresorcin[4]arene in non-aqueous solvents
rho1	795.62	kg/m3	303.15	Thermodynamic evidence for nano-heterogeneity in solutions of the macrocycle C-butylresorcin[4]arene in non-aqueous solvents
rho1	799.65	kg/m3	298.15	Thermodynamic evidence for nano-heterogeneity in solutions of the macrocycle C-butylresorcin[4]arene in non-aqueous solvents
rho1	803.66	kg/m3	293.15	Thermodynamic evidence for nano-heterogeneity in solutions of the macrocycle C-butylresorcin[4]arene in non-aqueous solvents
rho1	807.64	kg/m3	288.15	Thermodynamic evidence for nano-heterogeneity in solutions of the macrocycle C-butylresorcin[4]arene in non-aqueous solvents
rho1	811.60	kg/m3	283.15	Thermodynamic evidence for nano-heterogeneity in solutions of the macrocycle C-butylresorcin[4]arene in non-aqueous solvents

rho	815.55	kg/m ³	278.15	Thermodynamic evidence for nano-heterogeneity in solutions of the macrocycle C-butylresorcin[4]arene in non-aqueous solvents
rho	778.79	kg/m ³	323.15	Electrostriction of water and lower alcohols around ammonium nitrate - Volumetric approach
rho	783.11	kg/m ³	318.15	Electrostriction of water and lower alcohols around ammonium nitrate - Volumetric approach
rho	787.31	kg/m ³	313.15	Electrostriction of water and lower alcohols around ammonium nitrate - Volumetric approach
rho	791.43	kg/m ³	308.15	Electrostriction of water and lower alcohols around ammonium nitrate - Volumetric approach
rho	795.50	kg/m ³	303.15	Electrostriction of water and lower alcohols around ammonium nitrate - Volumetric approach
rho	799.52	kg/m ³	298.15	Electrostriction of water and lower alcohols around ammonium nitrate - Volumetric approach
rho	803.52	kg/m ³	293.15	Electrostriction of water and lower alcohols around ammonium nitrate - Volumetric approach

rho1	774.63	kg/m3	328.15	Thermophysical and volumetric study of mixtures {p-cymene + propan-1-ol} at several temperatures and atmospheric pressure. Modeling with COSMO-RS
rho1	783.10	kg/m3	318.15	Thermophysical and volumetric study of mixtures {p-cymene + propan-1-ol} at several temperatures and atmospheric pressure. Modeling with COSMO-RS
rho1	791.36	kg/m3	308.15	Thermophysical and volumetric study of mixtures {p-cymene + propan-1-ol} at several temperatures and atmospheric pressure. Modeling with COSMO-RS
rho1	799.46	kg/m3	298.15	Thermophysical and volumetric study of mixtures {p-cymene + propan-1-ol} at several temperatures and atmospheric pressure. Modeling with COSMO-RS
rho1	787.73	kg/m3	313.15	Temperature and composition dependence of the volumetric and acoustic properties of ionic liquid [emim][HSO4] with polar protic and aprotic co-solvents

rho	791.82	kg/m ³	308.15	Temperature and composition dependence of the volumetric and acoustic properties of ionic liquid [emim][HSO ₄] with polar protic and aprotic co-solvents
rho	795.92	kg/m ³	303.15	Temperature and composition dependence of the volumetric and acoustic properties of ionic liquid [emim][HSO ₄] with polar protic and aprotic co-solvents
rho	766.20	kg/m ³	338.15	Thermophysical properties of binary mixtures of 1-butyl-1-methylpyrrolidinium trifluoromethanesulfonate ionic liquid with alcohols at several temperatures
rho	774.90	kg/m ³	328.15	Thermophysical properties of binary mixtures of 1-butyl-1-methylpyrrolidinium trifluoromethanesulfonate ionic liquid with alcohols at several temperatures
rho	783.30	kg/m ³	318.15	Thermophysical properties of binary mixtures of 1-butyl-1-methylpyrrolidinium trifluoromethanesulfonate ionic liquid with alcohols at several temperatures
rho	791.60	kg/m ³	308.15	Thermophysical properties of binary mixtures of 1-butyl-1-methylpyrrolidinium trifluoromethanesulfonate ionic liquid with alcohols at several temperatures

rho1	799.70	kg/m3	298.15	Thermophysical properties of binary mixtures of 1-butyl-1-methylpyrrolidinium trifluoromethanesulfonate ionic liquid with alcohols at several temperatures
rho1	807.70	kg/m3	288.15	Thermophysical properties of binary mixtures of 1-butyl-1-methylpyrrolidinium trifluoromethanesulfonate ionic liquid with alcohols at several temperatures
rho1	815.60	kg/m3	278.15	Thermophysical properties of binary mixtures of 1-butyl-1-methylpyrrolidinium trifluoromethanesulfonate ionic liquid with alcohols at several temperatures
rho1	783.20	kg/m3	318.15	Excess molar volume and excess Gibbs energy of activation for viscous flow for the binary mixtures of N-ethylpyridinium dicyanamide [C2py][DCA] with alcohols
rho1	787.36	kg/m3	313.15	Excess molar volume and excess Gibbs energy of activation for viscous flow for the binary mixtures of N-ethylpyridinium dicyanamide [C2py][DCA] with alcohols

rho1	791.47	kg/m3	308.15	Excess molar volume and excess Gibbs energy of activation for viscous flow for the binary mixtures of N-ethylpyridinium dicyanamide [C2py][DCA] with alcohols
rho1	795.55	kg/m3	303.15	Excess molar volume and excess Gibbs energy of activation for viscous flow for the binary mixtures of N-ethylpyridinium dicyanamide [C2py][DCA] with alcohols
rho1	799.58	kg/m3	298.15	Excess molar volume and excess Gibbs energy of activation for viscous flow for the binary mixtures of N-ethylpyridinium dicyanamide [C2py][DCA] with alcohols
rho1	803.59	kg/m3	293.15	Excess molar volume and excess Gibbs energy of activation for viscous flow for the binary mixtures of N-ethylpyridinium dicyanamide [C2py][DCA] with alcohols
rho1	807.55	kg/m3	288.15	Excess molar volume and excess Gibbs energy of activation for viscous flow for the binary mixtures of N-ethylpyridinium dicyanamide [C2py][DCA] with alcohols

rho1	803.70	kg/m3	293.15	Measurement and correlation of (vapour-liquid) equilibrium for binary mixtures composed of 1-(ethoxymethoxy)-propane with ethanol and 1-propanol at 101.33 kPa
rho1	789.00	kg/m3	313.00	Comparative study of physical properties of binary mixtures of halogen free ionic liquids with alcohols
rho1	792.00	kg/m3	308.00	Comparative study of physical properties of binary mixtures of halogen free ionic liquids with alcohols
rho1	796.00	kg/m3	303.00	Comparative study of physical properties of binary mixtures of halogen free ionic liquids with alcohols
rho1	800.00	kg/m3	298.00	Comparative study of physical properties of binary mixtures of halogen free ionic liquids with alcohols
rho1	799.92	kg/m3	298.20	A green process for recovery of 1-propanol/2-propanol from their aqueous solutions: Experimental and MD simulation studies
rho1	799.50	kg/m3	298.15	Fully automatized apparatus for determining speed of sound for liquids in the temperature and pressure interval (283.15-343.15) K and (0.1-95) MPa

rhoI	787.41	kg/m3	313.15	Thermodynamic and spectroscopic properties of binary mixtures of n-butylammonium butanoate ionic liquid with alcohols at T = (293.15-313.15) K
rhoI	791.51	kg/m3	308.15	Thermodynamic and spectroscopic properties of binary mixtures of n-butylammonium butanoate ionic liquid with alcohols at T = (293.15-313.15) K
rhoI	795.58	kg/m3	303.15	Thermodynamic and spectroscopic properties of binary mixtures of n-butylammonium butanoate ionic liquid with alcohols at T = (293.15-313.15) K
rhoI	799.62	kg/m3	298.15	Thermodynamic and spectroscopic properties of binary mixtures of n-butylammonium butanoate ionic liquid with alcohols at T = (293.15-313.15) K
rhoI	803.62	kg/m3	293.15	Thermodynamic and spectroscopic properties of binary mixtures of n-butylammonium butanoate ionic liquid with alcohols at T = (293.15-313.15) K

rho	799.61	kg/m ³	298.15	Experimental study on the calorimetric data of 2-butoxyethanol with aliphatic alcohols (C1-C4) and correlation with the Wilson, NRTL and UNIQUAC models at T = 298 K
rho	803.85	kg/m ³	293.15	(Liquid + liquid) equilibria of four alcohol-water systems containing 1,8-cineole at T = 298.15 K
rho	799.70	kg/m ³	298.15	Solubility and solution thermodynamics of thymol in six pure organic solvents
rho	799.56	kg/m ³	298.15	Measurements and equation-of-state modelling of thermodynamic properties of binary mixtures of 1-butyl-1-methylpyrrolidinium tetracyanoborate ionic liquid with molecular compounds
rho	799.56	kg/m ³	298.15	Excess molar enthalpies of R-fenchone + propan-1-ol or +propan-2-ol. Modeling with COSMO-RS and UNIFAC
rho	787.42	kg/m ³	313.15	Properties of pure 1,1,3,3-tetramethylguanidine imidazole ionic liquid and its binary mixtures with alcohols at T = (293.15 to 313.15) K

rho1	791.56	kg/m3	308.15	Properties of pure 1,1,3,3-tetramethylguanidine imidazole ionic liquid and its binary mixtures with alcohols at T = (293.15 to 313.15) K
rho1	795.65	kg/m3	303.15	Properties of pure 1,1,3,3-tetramethylguanidine imidazole ionic liquid and its binary mixtures with alcohols at T = (293.15 to 313.15) K
rho1	799.71	kg/m3	298.15	Properties of pure 1,1,3,3-tetramethylguanidine imidazole ionic liquid and its binary mixtures with alcohols at T = (293.15 to 313.15) K
rho1	803.75	kg/m3	293.15	Properties of pure 1,1,3,3-tetramethylguanidine imidazole ionic liquid and its binary mixtures with alcohols at T = (293.15 to 313.15) K
rho1	803.85	kg/m3	293.15	(Liquid + liquid) equilibria for (water + 1-propanol or acetone + .beta.-citronellol) at different temperatures
rho1	799.56	kg/m3	298.15	A combined experimental and computational investigation of excess molar enthalpies of (nitrobenzene + alkanol) mixtures
rho1	787.88	kg/m3	313.15	Molecular interactions in binary mixtures of 1-butoxy-2-propanol with alcohols at different temperatures: A thermophysical and spectroscopic approach

rho	792.00	kg/m ³	308.15	Molecular interactions in binary mixtures of 1-butoxy-2-propanol with alcohols at different temperatures: A thermophysical and spectroscopic approach
rho	796.08	kg/m ³	303.15	Molecular interactions in binary mixtures of 1-butoxy-2-propanol with alcohols at different temperatures: A thermophysical and spectroscopic approach
rho	800.13	kg/m ³	298.15	Molecular interactions in binary mixtures of 1-butoxy-2-propanol with alcohols at different temperatures: A thermophysical and spectroscopic approach
rho	804.14	kg/m ³	293.15	Molecular interactions in binary mixtures of 1-butoxy-2-propanol with alcohols at different temperatures: A thermophysical and spectroscopic approach
rho	799.52	kg/m ³	298.15	Extraction desulfurization process of fuels with ionic liquids
rho	799.52	kg/m ³	298.15	Effect of the alkyl side chain of the 1-alkylpiperidinium-based ionic liquids on desulfurization of fuels

rhoI	779.09	kg/m ³	323.15	Osmotic coefficients and apparent molar volumes of 1-hexyl-3-methylimidazolium trifluoromethanesulfonate ionic liquid in alcohols
rhoI	799.52	kg/m ³	298.15	Separation of sulfur compounds from alkanes with 1-alkylcyanopyridinium-based ionic liquids
rhoI	799.66	kg/m ³	298.15	Thermodynamics of (ketone + amine) mixtures. Part XI. Excess molar enthalpies at T = 298.15 K for the (1-propanol + N,N,N-triethylamine + 2-butanone) system
rhoI	798.00	kg/m ³	303.15	Study of molecular interactions in the mixtures of some primary alcohols with equimolar mixture of 1-propanol and alkylbenzoates at T = 303.15 K
rhoI	791.78	kg/m ³	308.15	Physical properties of the pure 1-methyl-1-propylpyrrolidinium bis(trifluoromethylsulfonyl)imide ionic liquid and its binary mixtures with alcohols
rhoI	795.94	kg/m ³	303.15	Physical properties of the pure 1-methyl-1-propylpyrrolidinium bis(trifluoromethylsulfonyl)imide ionic liquid and its binary mixtures with alcohols
rhoI	799.87	kg/m ³	298.15	Physical properties of the pure 1-methyl-1-propylpyrrolidinium bis(trifluoromethylsulfonyl)imide ionic liquid and its binary mixtures with alcohols

rho1	791.65	kg/m3	308.15	FT-IR studies on excess thermodynamic properties of binary liquid mixtures o-chlorotoluene with 1-propanol, 1-butanol, 1-pentanol, 1-hexanol and 1-heptanol at different temperatures
rho1	795.65	kg/m3	303.15	FT-IR studies on excess thermodynamic properties of binary liquid mixtures o-chlorotoluene with 1-propanol, 1-butanol, 1-pentanol, 1-hexanol and 1-heptanol at different temperatures
rho1	799.65	kg/m3	298.15	FT-IR studies on excess thermodynamic properties of binary liquid mixtures o-chlorotoluene with 1-propanol, 1-butanol, 1-pentanol, 1-hexanol and 1-heptanol at different temperatures
rho1	779.33	kg/m3	323.15	Osmotic and apparent molar properties of binary mixtures alcohol + 1-butyl-3-methylimidazolium trifluoromethanesulfonate ionic liquid
rho1	791.84	kg/m3	308.15	Osmotic and apparent molar properties of binary mixtures alcohol + 1-butyl-3-methylimidazolium trifluoromethanesulfonate ionic liquid

rhoI	803.92	kg/m ³	293.15	Osmotic and apparent molar properties of binary mixtures alcohol + 1-butyl-3-methylimidazolium trifluoromethanesulfonate ionic liquid
rhoI	783.16	kg/m ³	318.15	Volume effects for binary mixtures of propane-1,2-diol with methanol, propan-1-ol, hexan-1-ol, octan-1-ol, or nonan-1-ol at temperatures (293.15 to 318.15) K
rhoI	787.32	kg/m ³	313.15	Volume effects for binary mixtures of propane-1,2-diol with methanol, propan-1-ol, hexan-1-ol, octan-1-ol, or nonan-1-ol at temperatures (293.15 to 318.15) K
rhoI	791.43	kg/m ³	308.15	Volume effects for binary mixtures of propane-1,2-diol with methanol, propan-1-ol, hexan-1-ol, octan-1-ol, or nonan-1-ol at temperatures (293.15 to 318.15) K
rhoI	795.49	kg/m ³	303.15	Volume effects for binary mixtures of propane-1,2-diol with methanol, propan-1-ol, hexan-1-ol, octan-1-ol, or nonan-1-ol at temperatures (293.15 to 318.15) K

rho	799.53	kg/m ³	298.15	Volume effects for binary mixtures of propane-1,2-diol with methanol, propan-1-ol, hexan-1-ol, octan-1-ol, or nonan-1-ol at temperatures (293.15 to 318.15) K
rho	803.53	kg/m ³	293.15	Volume effects for binary mixtures of propane-1,2-diol with methanol, propan-1-ol, hexan-1-ol, octan-1-ol, or nonan-1-ol at temperatures (293.15 to 318.15) K
rho	799.45	kg/m ³	298.15	Apparent molar volumes and compressibilities of tetrabutyl-ammonium bromide in organic solvents
rho	788.60	kg/m ³	313.15	Excess molar volumes and isentropic compressibility of binary systems {trioctylmethylammonium bis(trifluoromethylsulfonyl)imide + methanol or ethanol or 1-propanol} at different temperatures
rho	796.40	kg/m ³	303.15	Excess molar volumes and isentropic compressibility of binary systems {trioctylmethylammonium bis(trifluoromethylsulfonyl)imide + methanol or ethanol or 1-propanol} at different temperatures

rho1	799.40	kg/m3	298.15	Excess molar volumes and isentropic compressibility of binary systems {trioctylmethylammonium bis(trifluoromethylsulfonyl)imide + methanol or ethanol or 1-propanol} at different temperatures
rho1	799.67	kg/m3	298.15	Volumetric, acoustic, and viscometric studies of molecular interactions in binary mixtures of dipropylene glycol dimethyl ether with 1-alkanols at 298.15 K
rho1	778.97	kg/m3	323.15	Densities and volumetric properties of (N-(2-hydroxyethyl)morpholine + ethanol, + 1-propanol, + 2-propanol, + 1-butanol, and + 2-butanol) at (293.15, 298.15, 303.15, 313.15, and 323.15) K
rho1	787.34	kg/m3	313.15	Densities and volumetric properties of (N-(2-hydroxyethyl)morpholine + ethanol, + 1-propanol, + 2-propanol, + 1-butanol, and + 2-butanol) at (293.15, 298.15, 303.15, 313.15, and 323.15) K
rho1	795.58	kg/m3	303.15	Densities and volumetric properties of (N-(2-hydroxyethyl)morpholine + ethanol, + 1-propanol, + 2-propanol, + 1-butanol, and + 2-butanol) at (293.15, 298.15, 303.15, 313.15, and 323.15) K

rho	799.71	kg/m ³	298.15	Densities and volumetric properties of (N-(2-hydroxyethyl)morpholine + ethanol, + 1-propanol, + 2-propanol, + 1-butanol, and + 2-butanol) at (293.15, 298.15, 303.15, 313.15, and 323.15) K
rho	803.56	kg/m ³	293.15	Densities and volumetric properties of (N-(2-hydroxyethyl)morpholine + ethanol, + 1-propanol, + 2-propanol, + 1-butanol, and + 2-butanol) at (293.15, 298.15, 303.15, 313.15, and 323.15) K
rho	787.36	kg/m ³	313.15	Excess molar volumes of binary mixtures of 1,3-dimethylimidazolidin-2-one with an alkan-1-ol at the temperatures 283.15 K, 298.15 K, and 313.15 K
rho	799.59	kg/m ³	298.15	Excess molar volumes of binary mixtures of 1,3-dimethylimidazolidin-2-one with an alkan-1-ol at the temperatures 283.15 K, 298.15 K, and 313.15 K
rho	811.54	kg/m ³	283.15	Excess molar volumes of binary mixtures of 1,3-dimethylimidazolidin-2-one with an alkan-1-ol at the temperatures 283.15 K, 298.15 K, and 313.15 K
rho	761.51	kg/m ³	343.15	Densities and derived thermodynamic properties of (2-methoxyethanol + 1-propanol, or 2-propanol, or 1,2-propandiol) at temperatures from T = (293.15 to 343.15) K

rhoI	770.42	kg/m3	333.15	Densities and derived thermodynamic properties of (2-methoxyethanol + 1-propanol, or 2-propanol, or 1,2-propandiol) at temperatures from T = (293.15 to 343.15) K
rhoI	779.02	kg/m3	323.15	Densities and derived thermodynamic properties of (2-methoxyethanol + 1-propanol, or 2-propanol, or 1,2-propandiol) at temperatures from T = (293.15 to 343.15) K
rhoI	787.39	kg/m3	313.15	Densities and derived thermodynamic properties of (2-methoxyethanol + 1-propanol, or 2-propanol, or 1,2-propandiol) at temperatures from T = (293.15 to 343.15) K
rhoI	795.57	kg/m3	303.15	Densities and derived thermodynamic properties of (2-methoxyethanol + 1-propanol, or 2-propanol, or 1,2-propandiol) at temperatures from T = (293.15 to 343.15) K
rhoI	803.61	kg/m3	293.15	Densities and derived thermodynamic properties of (2-methoxyethanol + 1-propanol, or 2-propanol, or 1,2-propandiol) at temperatures from T = (293.15 to 343.15) K

rhoI	783.22	kg/m3	318.15	Volumetric properties of binary mixtures of ionic liquid 1-butyl-3-methylimidazolium octylsulfate with water or propanol in the temperature range of 278.15 K to 328.15 K
rhoI	791.44	kg/m3	308.15	Volumetric properties of binary mixtures of ionic liquid 1-butyl-3-methylimidazolium octylsulfate with water or propanol in the temperature range of 278.15 K to 328.15 K
rhoI	799.54	kg/m3	298.15	Volumetric properties of binary mixtures of ionic liquid 1-butyl-3-methylimidazolium octylsulfate with water or propanol in the temperature range of 278.15 K to 328.15 K
rhoI	807.54	kg/m3	288.15	Volumetric properties of binary mixtures of ionic liquid 1-butyl-3-methylimidazolium octylsulfate with water or propanol in the temperature range of 278.15 K to 328.15 K
rhoI	815.98	kg/m3	278.15	Volumetric properties of binary mixtures of ionic liquid 1-butyl-3-methylimidazolium octylsulfate with water or propanol in the temperature range of 278.15 K to 328.15 K
rhoI	799.60	kg/m3	298.15	Bubble point temperatures of the binary mixtures of nitrobenzene with C1 C4 aliphatic alcohols at 94.95 kPa

rhoI	799.74	kg/m ³	298.15	(Vapor + liquid) equilibria for the binary mixtures (1-propanol + dibromomethane, or + bromochloromethane, or + 1,2-dichloroethane or +1-bromo-2-chloroethane) at T = 313.15 K.
rhoI	791.30	kg/m ³	308.15	Temperature dependence of the volumetric properties of some alkoxypropanols + n-alkanol mixtures
rhoI	799.40	kg/m ³	298.15	Temperature dependence of the volumetric properties of some alkoxypropanols + n-alkanol mixtures
rhoI	807.30	kg/m ³	288.15	Temperature dependence of the volumetric properties of some alkoxypropanols + n-alkanol mixtures
rhoI	800.06	kg/m ³	298.15	Solid-liquid equilibria for selected binary systems containing diphenyl carbonate
rhoI	791.82	kg/m ³	308.15	A systematic study on volumetric and transport properties of binary systems 1-propanol + n-hexadecane, 1-butanol + n-hexadecane and 1-propanol + ethyl oleate at different temperatures: Experimental and modeling

rho1	795.89	kg/m3	303.15	A systematic study on volumetric and transport properties of binary systems 1-propanol + n-hexadecane, 1-butanol + n-hexadecane and 1-propanol + ethyl oleate at different temperatures: Experimental and modeling
rho1	799.93	kg/m3	298.15	A systematic study on volumetric and transport properties of binary systems 1-propanol + n-hexadecane, 1-butanol + n-hexadecane and 1-propanol + ethyl oleate at different temperatures: Experimental and modeling
rho1	799.52	kg/m3	298.15	Separation of pyridine from heptane with tricyanomethanide-based ionic liquids
rho1	779.14	kg/m3	323.15	Measurement and modeling of volumetric properties and speeds of sound of several mixtures of alcohol liquids containing 1-propanol and 2-propanol at T=(298.15 - 323.15) K and ambient pressure
rho1	783.36	kg/m3	318.15	Measurement and modeling of volumetric properties and speeds of sound of several mixtures of alcohol liquids containing 1-propanol and 2-propanol at T=(298.15 - 323.15) K and ambient pressure

rhoI	787.52	kg/m3	313.15	Measurement and modeling of volumetric properties and speeds of sound of several mixtures of alcohol liquids containing 1-propanol and 2-propanol at T=(298.15 - 323.15) K and ambient pressure
rhoI	791.63	kg/m3	308.15	Measurement and modeling of volumetric properties and speeds of sound of several mixtures of alcohol liquids containing 1-propanol and 2-propanol at T=(298.15 - 323.15) K and ambient pressure
rhoI	795.70	kg/m3	303.15	Measurement and modeling of volumetric properties and speeds of sound of several mixtures of alcohol liquids containing 1-propanol and 2-propanol at T=(298.15 - 323.15) K and ambient pressure
rhoI	799.71	kg/m3	298.15	Measurement and modeling of volumetric properties and speeds of sound of several mixtures of alcohol liquids containing 1-propanol and 2-propanol at T=(298.15 - 323.15) K and ambient pressure

rho1	799.62	kg/m3	298.15	Liquid liquid equilibria of 4-methyl-2-pentanone + 1-propanol or 2-propanol + water ternary systems: Measurements and correlation at different temperatures
rho1	800.13	kg/m3	298.15	Thermodynamic and spectral investigations of binary liquid mixtures of 2-butoxy ethanol with alcohols at temperature range of 293.15-313.15 K
rho1	800.61	kg/m3	298.15	Isothermal vapor liquid equilibrium at 333.15K and excess molar volumes at 298.15K for the ternary system di-isopropyl ether + n-propyl alcohol + toluene and its binary subsystems
rho1	799.96	kg/m3	298.15	Liquid liquid equilibria for the binary system of di-isopropyl ether (DIPE) +water in between 288.15 and 323.15K and the ternary systems of DIPE +water + C1 C4 alcohols at 298.15K
rho1	791.76	kg/m3	308.15	Thermodynamic properties of mixtures containing alkoxypropanol and n-alkanol
rho1	799.67	kg/m3	298.15	Thermodynamic properties of mixtures containing alkoxypropanol and n-alkanol
rho1	807.90	kg/m3	288.15	Thermodynamic properties of mixtures containing alkoxypropanol and n-alkanol

rho1	787.47	kg/m3	313.15	Influence of chain length and degree of branching of alcohol + chlorobenzene mixtures on determination and modelling of VE by CEOS and CEOS/GE mixing rules
rho1	791.58	kg/m3	308.15	Influence of chain length and degree of branching of alcohol + chlorobenzene mixtures on determination and modelling of VE by CEOS and CEOS/GE mixing rules
rho1	795.65	kg/m3	303.15	Influence of chain length and degree of branching of alcohol + chlorobenzene mixtures on determination and modelling of VE by CEOS and CEOS/GE mixing rules
rho1	799.69	kg/m3	298.15	Influence of chain length and degree of branching of alcohol + chlorobenzene mixtures on determination and modelling of VE by CEOS and CEOS/GE mixing rules
rho1	803.70	kg/m3	293.15	Influence of chain length and degree of branching of alcohol + chlorobenzene mixtures on determination and modelling of VE by CEOS and CEOS/GE mixing rules

rho	807.69	kg/m ³	288.15	Influence of chain length and degree of branching of alcohol + chlorobenzene mixtures on determination and modelling of VE by CEOS and CEOS/GE mixing rules
rho	799.65	kg/m ³	298.15	Excess molar enthalpy of 1-alkanol + 1-octene mixtures at 298.15K Experimental results and theoretical description by means of the ERAS and TB models
rho	770.39	kg/m ³	333.15	Volumetric properties of the boldine + alcohol mixtures at atmospheric pressure from 283.15 to 333.15K A new method for the determination of the density of pure boldine
rho	774.73	kg/m ³	328.15	Volumetric properties of the boldine + alcohol mixtures at atmospheric pressure from 283.15 to 333.15K A new method for the determination of the density of pure boldine
rho	779.01	kg/m ³	323.15	Volumetric properties of the boldine + alcohol mixtures at atmospheric pressure from 283.15 to 333.15K A new method for the determination of the density of pure boldine

rho1	787.37	kg/m3	313.15	Volumetric properties of the boldine + alcohol mixtures at atmospheric pressure from 283.15 to 333.15K A new method for the determination of the density of pure boldine
rho1	791.49	kg/m3	308.15	Volumetric properties of the boldine + alcohol mixtures at atmospheric pressure from 283.15 to 333.15K A new method for the determination of the density of pure boldine
rho1	795.58	kg/m3	303.15	Volumetric properties of the boldine + alcohol mixtures at atmospheric pressure from 283.15 to 333.15K A new method for the determination of the density of pure boldine
rho1	799.62	kg/m3	298.15	Volumetric properties of the boldine + alcohol mixtures at atmospheric pressure from 283.15 to 333.15K A new method for the determination of the density of pure boldine
rho1	803.63	kg/m3	293.15	Volumetric properties of the boldine + alcohol mixtures at atmospheric pressure from 283.15 to 333.15K A new method for the determination of the density of pure boldine

rho1	807.62	kg/m3	288.15	Volumetric properties of the boldine + alcohol mixtures at atmospheric pressure from 283.15 to 333.15K A new method for the determination of the density of pure boldine
rho1	811.57	kg/m3	283.15	Volumetric properties of the boldine + alcohol mixtures at atmospheric pressure from 283.15 to 333.15K A new method for the determination of the density of pure boldine
rho1	795.80	kg/m3	303.15	Viscous synergy and antagonism and isentropic compressibility of ternary mixtures containing 1,3-dioxolane, water and monoalkanols at 303.15K
rho1	799.51	kg/m3	298.15	Fluid Phase Topology of Benzene + Cyclohexane + 1-Propanol at 101.3 kPa
rho1	799.60	kg/m3	298.15	Experimental Determination of Densities and Isobaric Vapor Liquid Equilibria of Methyl Acetate and Ethyl Acetate with Alcohols (C3 and C4) at 0.3 MPa
rho1	799.28	kg/m3	298.15	Topological Investigations of Excess Molar Volumes and Excess Isentropic Compressibilities of Ternary Mixtures Containing Pyrrolidin-2-one at 308.15 K

rho	783.86	kg/m ³	318.15	Densities and Volumetric Properties of Binary Mixtures of Aniline with 1-Propanol, 2-Propanol, 2-Methyl-1-Propanol, and 2-Methyl-2-Propanol at Temperatures from 293.15 to 318.15 K
rho	787.83	kg/m ³	313.15	Densities and Volumetric Properties of Binary Mixtures of Aniline with 1-Propanol, 2-Propanol, 2-Methyl-1-Propanol, and 2-Methyl-2-Propanol at Temperatures from 293.15 to 318.15 K
rho	791.80	kg/m ³	308.15	Densities and Volumetric Properties of Binary Mixtures of Aniline with 1-Propanol, 2-Propanol, 2-Methyl-1-Propanol, and 2-Methyl-2-Propanol at Temperatures from 293.15 to 318.15 K
rho	795.77	kg/m ³	303.15	Densities and Volumetric Properties of Binary Mixtures of Aniline with 1-Propanol, 2-Propanol, 2-Methyl-1-Propanol, and 2-Methyl-2-Propanol at Temperatures from 293.15 to 318.15 K
rho	799.74	kg/m ³	298.15	Densities and Volumetric Properties of Binary Mixtures of Aniline with 1-Propanol, 2-Propanol, 2-Methyl-1-Propanol, and 2-Methyl-2-Propanol at Temperatures from 293.15 to 318.15 K

rho1	803.71	kg/m3	293.15	Densities and Volumetric Properties of Binary Mixtures of Aniline with 1-Propanol, 2-Propanol, 2-Methyl-1-Propanol, and 2-Methyl-2-Propanol at Temperatures from 293.15 to 318.15 K
rho1	800.62	kg/m3	298.15	Speeds of Sound and Isentropic Compressibilities in Binary Mixtures of 2-Propanol with Several 1-Alkanols at 298.15K
rho1	804.00	kg/m3	293.00	KDB
rho1	783.22	kg/m3	318.15	Volumetric properties of the boldine + alcohol mixtures at atmospheric pressure from 283.15 to 333.15K A new method for the determination of the density of pure boldine
rho1	784.45	kg/m3	318.15	Physical Properties of the Pure Deep Eutectic Solvent, [ChCl]:[Lev] (1:2) DES, and Its Binary Mixtures with Alcohols
rho1	799.96	kg/m3	298.15	Temperature and composition dependence of the volumetric and acoustic properties of ionic liquid [emim][HSO4] with polar protic and aprotic co-solvents
sfust	36.11	J/molxK	148.75	NIST Webbook
sfust	35.30	J/molxK	147.00	NIST Webbook

speedsl	1257.90	m/s	283.15	Excess Enthalpy, Excess Volume, and Speed of Sound Deviation for the Mixtures ss-Pinene + Ethanol and ss-Pinene + 1-propanol at (283.15, 298.15 and, 313.15) K
speedsl	1137.83	m/s	318.15	Volumetric and Ultrasonic Studies of 1-Ethyl-3-methylimidazolium Trifluoromethanesulfonate Ionic Liquid with Methanol, Ethanol, 1-Propanol, and Water at Several Temperatures
speedsl	1171.41	m/s	308.15	Volumetric and Ultrasonic Studies of 1-Ethyl-3-methylimidazolium Trifluoromethanesulfonate Ionic Liquid with Methanol, Ethanol, 1-Propanol, and Water at Several Temperatures
speedsl	1205.37	m/s	298.15	Volumetric and Ultrasonic Studies of 1-Ethyl-3-methylimidazolium Trifluoromethanesulfonate Ionic Liquid with Methanol, Ethanol, 1-Propanol, and Water at Several Temperatures
speedsl	1239.86	m/s	288.15	Volumetric and Ultrasonic Studies of 1-Ethyl-3-methylimidazolium Trifluoromethanesulfonate Ionic Liquid with Methanol, Ethanol, 1-Propanol, and Water at Several Temperatures

speedsl	1275.27	m/s	278.15	Volumetric and Ultrasonic Studies of 1-Ethyl-3-methylimidazolium Trifluoromethanesulfonate Ionic Liquid with Methanol, Ethanol, 1-Propanol, and Water at Several Temperatures
speedsl	1104.40	m/s	328.15	Volumetric and Ultrasonic Studies of 1-Ethyl-3-methylimidazolium Trifluoromethanesulfonate Ionic Liquid with Methanol, Ethanol, 1-Propanol, and Water at Several Temperatures
speedsl	1155.10	m/s	313.15	Excess Enthalpy, Excess Volume, and Speed of Sound Deviation for the Mixtures ss-Pinene + Ethanol and ss-Pinene + 1-propanol at (283.15, 298.15 and, 313.15) K
speedsl	1206.00	m/s	298.15	Excess Enthalpy, Excess Volume, and Speed of Sound Deviation for the Mixtures ss-Pinene + Ethanol and ss-Pinene + 1-propanol at (283.15, 298.15 and, 313.15) K
speedsl	1070.97	m/s	338.15	Volumetric and Ultrasonic Studies of 1-Ethyl-3-methylimidazolium Trifluoromethanesulfonate Ionic Liquid with Methanol, Ethanol, 1-Propanol, and Water at Several Temperatures

speedsl	1121.80	m/s	323.15	Acoustic, volumetric and osmotic properties of binary mixtures containing the ionic liquid 1-butyl-3-methylimidazolim dicyanamide mixed with primary and secondary alcohols
speedsl	1172.30	m/s	308.15	Acoustic, volumetric and osmotic properties of binary mixtures containing the ionic liquid 1-butyl-3-methylimidazolim dicyanamide mixed with primary and secondary alcohols
speedsl	1223.40	m/s	293.15	Acoustic, volumetric and osmotic properties of binary mixtures containing the ionic liquid 1-butyl-3-methylimidazolim dicyanamide mixed with primary and secondary alcohols
speedsl	1205.80	m/s	298.15	Ultrasonic speeds and isentropic compressibilities of {difurylmethane + (C1 C6) n-alkanol} binary mixtures at T = 298.15 K
speedsl	1172.04	m/s	308.15	Densities, excess molar volumes, speeds of sound and isothermal compressibilities for {2-(2-hexyloxyethoxy)ethanol + n-alkanol} systems at temperatures between (288.15 and 308.15) K

speedsl	1189.86	m/s	303.15	Densities, excess molar volumes, speeds of sound and isothermal compressibilities for {2-(2-hexyloxyethoxy)ethanol + n-alkanol} systems at temperatures between (288.15 and 308.15) K
speedsl	1206.47	m/s	298.15	Densities, excess molar volumes, speeds of sound and isothermal compressibilities for {2-(2-hexyloxyethoxy)ethanol + n-alkanol} systems at temperatures between (288.15 and 308.15) K
speedsl	1223.17	m/s	293.15	Densities, excess molar volumes, speeds of sound and isothermal compressibilities for {2-(2-hexyloxyethoxy)ethanol + n-alkanol} systems at temperatures between (288.15 and 308.15) K
speedsl	1240.01	m/s	288.15	Densities, excess molar volumes, speeds of sound and isothermal compressibilities for {2-(2-hexyloxyethoxy)ethanol + n-alkanol} systems at temperatures between (288.15 and 308.15) K
speedsl	1207.00	m/s	298.15	Compressibility Studies of Binary Solutions Involving Water as a Solute in Nonaqueous Solvents at T) 298.15 K
srf	0.02	N/m	313.15	Surface Tension of Binary Mixtures of 2,2,4-Trimethylpentane +1-Alkanols from 298.15 to 323.15 K

srf	0.02	N/m	293.15	Density and Surface Tension of Binary Mixtures of Acetonitrile + 1-Alkanol at 293.15 K
srf	0.02	N/m	323.15	Surface Tension of Dilute Solutions of Linear Alcohols in Benzyl Alcohol
srf	0.02	N/m	318.15	Surface Tension of Dilute Solutions of Linear Alcohols in Benzyl Alcohol
srf	0.02	N/m	313.15	Surface Tension of Dilute Solutions of Linear Alcohols in Benzyl Alcohol
srf	0.02	N/m	308.15	Surface Tension of Dilute Solutions of Linear Alcohols in Benzyl Alcohol
srf	0.02	N/m	303.15	Surface Tension of Dilute Solutions of Linear Alcohols in Benzyl Alcohol
srf	0.02	N/m	298.15	Surface Tension of Dilute Solutions of Linear Alcohols in Benzyl Alcohol
srf	0.02	N/m	293.15	Surface Tension of Dilute Solutions of Linear Alcohols in Benzyl Alcohol
srf	0.02	N/m	323.15	Surface Tension of Binary Mixtures of 2,2,4-Trimethylpentane +1-Alkanols from 298.15 to 323.15 K
srf	0.02	N/m	318.15	Surface Tension of Binary Mixtures of 2,2,4-Trimethylpentane +1-Alkanols from 298.15 to 323.15 K

srf	0.02	N/m	298.00	Surface Tension and Density of Pure Ionic Liquids and Some Binary Mixtures with 1-Propanol and 1-Butanol
srf	0.02	N/m	308.15	Surface Tension of Binary Mixtures of 2,2,4-Trimethylpentane +1-Alkanols from 298.15 to 323.15 K
srf	0.02	N/m	303.15	Surface Tension of Binary Mixtures of 2,2,4-Trimethylpentane +1-Alkanols from 298.15 to 323.15 K
srf	0.02	N/m	298.15	Surface Tension of Binary Mixtures of 2,2,4-Trimethylpentane +1-Alkanols from 298.15 to 323.15 K
srf	0.02	N/m	333.20	Surface tension and interfacial compositions of binary glycerol/alcohol mixtures
srf	0.02	N/m	323.20	Surface tension and interfacial compositions of binary glycerol/alcohol mixtures
srf	0.02	N/m	313.20	Surface tension and interfacial compositions of binary glycerol/alcohol mixtures
srf	0.02	N/m	303.20	Surface tension and interfacial compositions of binary glycerol/alcohol mixtures
srf	0.02	N/m	293.20	Surface tension and interfacial compositions of binary glycerol/alcohol mixtures

srf	0.02	N/m	318.15	The molar surface Gibbs energy and its application to the binary mixtures of N-butylpyridinium dicyanamide [C4py][DCA] with alcohols
srf	0.02	N/m	313.15	The molar surface Gibbs energy and its application to the binary mixtures of N-butylpyridinium dicyanamide [C4py][DCA] with alcohols
srf	0.02	N/m	308.15	The molar surface Gibbs energy and its application to the binary mixtures of N-butylpyridinium dicyanamide [C4py][DCA] with alcohols
srf	0.02	N/m	303.15	The molar surface Gibbs energy and its application to the binary mixtures of N-butylpyridinium dicyanamide [C4py][DCA] with alcohols
srf	0.02	N/m	298.15	The molar surface Gibbs energy and its application to the binary mixtures of N-butylpyridinium dicyanamide [C4py][DCA] with alcohols
srf	0.02	N/m	293.15	The molar surface Gibbs energy and its application to the binary mixtures of N-butylpyridinium dicyanamide [C4py][DCA] with alcohols

srf	0.02	N/m	288.15	The molar surface Gibbs energy and its application to the binary mixtures of N-butylpyridinium dicyanamide [C4py][DCA] with alcohols
srf	0.02	N/m	328.15	Study of surface tension and surface properties of binary systems of DMSO with long chain alcohols at various temperatures
srf	0.02	N/m	318.15	Study of surface tension and surface properties of binary systems of DMSO with long chain alcohols at various temperatures
srf	0.02	N/m	308.15	Study of surface tension and surface properties of binary systems of DMSO with long chain alcohols at various temperatures
srf	0.02	N/m	298.15	Study of surface tension and surface properties of binary systems of DMSO with long chain alcohols at various temperatures
srf	0.02	N/m	288.15	Study of surface tension and surface properties of binary systems of DMSO with long chain alcohols at various temperatures
srf	0.02	N/m	318.15	Surface thermodynamics of binary mixtures of aliphatic alcohols in heavy water

srf	0.02	N/m	308.15	Surface thermodynamics of binary mixtures of aliphatic alcohols in heavy water
srf	0.02	N/m	298.15	Surface thermodynamics of binary mixtures of aliphatic alcohols in heavy water
srf	0.02	N/m	288.15	Surface thermodynamics of binary mixtures of aliphatic alcohols in heavy water
srf	0.02	N/m	293.20	KDB
srf	0.02	N/m	308.15	Application of the Extended Langmuir Model for the Determination of Lyophobicity of 1-Propanol in Acetonitrile
srf	0.02	N/m	303.15	Application of the Extended Langmuir Model for the Determination of Lyophobicity of 1-Propanol in Acetonitrile
srf	0.02	N/m	298.15	Application of the Extended Langmuir Model for the Determination of Lyophobicity of 1-Propanol in Acetonitrile
srf	0.02	N/m	293.15	Application of the Extended Langmuir Model for the Determination of Lyophobicity of 1-Propanol in Acetonitrile
srf	0.02	N/m	288.15	Application of the Extended Langmuir Model for the Determination of Lyophobicity of 1-Propanol in Acetonitrile

srf	0.02	N/m	283.15	Application of the Extended Langmuir Model for the Determination of Lyophobicity of 1-Propanol in Acetonitrile
srf	0.02	N/m	278.15	Application of the Extended Langmuir Model for the Determination of Lyophobicity of 1-Propanol in Acetonitrile
srf	0.02	N/m	313.15	Application of the Extended Langmuir Model for the Determination of Lyophobicity of 1-Propanol in Acetonitrile

Pressure Dependent Properties

Property code	Value	Unit	Pressure [kPa]	Source
tbp	360.86	K	69.79	A novel dynamic recirculating apparatus for vapour-liquid equilibrium measurements at moderate pressures and temperatures
tbp	362.60	K	74.92	A novel dynamic recirculating apparatus for vapour-liquid equilibrium measurements at moderate pressures and temperatures
tbp	364.11	K	79.60	A novel dynamic recirculating apparatus for vapour-liquid equilibrium measurements at moderate pressures and temperatures

tbp	365.70	K	84.80	A novel dynamic recirculating apparatus for vapour-liquid equilibrium measurements at moderate pressures and temperatures
tbp	367.14	K	89.78	A novel dynamic recirculating apparatus for vapour-liquid equilibrium measurements at moderate pressures and temperatures
tbp	368.52	K	94.77	A novel dynamic recirculating apparatus for vapour-liquid equilibrium measurements at moderate pressures and temperatures
tbp	369.82	K	99.70	A novel dynamic recirculating apparatus for vapour-liquid equilibrium measurements at moderate pressures and temperatures
tbp	371.30	K	105.14	A novel dynamic recirculating apparatus for vapour-liquid equilibrium measurements at moderate pressures and temperatures
tbp	372.42	K	110.10	A novel dynamic recirculating apparatus for vapour-liquid equilibrium measurements at moderate pressures and temperatures
tbp	374.88	K	120.75	A novel dynamic recirculating apparatus for vapour-liquid equilibrium measurements at moderate pressures and temperatures

tbp	373.15	K	113.19	A novel dynamic recirculating apparatus for vapour-liquid equilibrium measurements at moderate pressures and temperatures
tbp	376.91	K	130.10	A novel dynamic recirculating apparatus for vapour-liquid equilibrium measurements at moderate pressures and temperatures
tbp	384.57	K	171.00	A novel dynamic recirculating apparatus for vapour-liquid equilibrium measurements at moderate pressures and temperatures
tbp	389.14	K	199.90	A novel dynamic recirculating apparatus for vapour-liquid equilibrium measurements at moderate pressures and temperatures
tbp	393.15	K	228.50	A novel dynamic recirculating apparatus for vapour-liquid equilibrium measurements at moderate pressures and temperatures
tbp	393.35	K	229.90	A novel dynamic recirculating apparatus for vapour-liquid equilibrium measurements at moderate pressures and temperatures
tbp	397.13	K	260.00	A novel dynamic recirculating apparatus for vapour-liquid equilibrium measurements at moderate pressures and temperatures

tbp	400.57	K	290.00	A novel dynamic recirculating apparatus for vapour-liquid equilibrium measurements at moderate pressures and temperatures
tbp	402.71	K	310.00	A novel dynamic recirculating apparatus for vapour-liquid equilibrium measurements at moderate pressures and temperatures
tbp	404.74	K	330.00	A novel dynamic recirculating apparatus for vapour-liquid equilibrium measurements at moderate pressures and temperatures
tbp	406.69	K	350.00	A novel dynamic recirculating apparatus for vapour-liquid equilibrium measurements at moderate pressures and temperatures
tbp	409.43	K	380.00	A novel dynamic recirculating apparatus for vapour-liquid equilibrium measurements at moderate pressures and temperatures
tbp	411.18	K	400.00	A novel dynamic recirculating apparatus for vapour-liquid equilibrium measurements at moderate pressures and temperatures
tbp	412.84	K	420.00	A novel dynamic recirculating apparatus for vapour-liquid equilibrium measurements at moderate pressures and temperatures

tbp	415.23	K	450.00	A novel dynamic recirculating apparatus for vapour-liquid equilibrium measurements at moderate pressures and temperatures
tbp	417.52	K	480.00	A novel dynamic recirculating apparatus for vapour-liquid equilibrium measurements at moderate pressures and temperatures
tbp	418.96	K	500.00	A novel dynamic recirculating apparatus for vapour-liquid equilibrium measurements at moderate pressures and temperatures
tbp	420.37	K	520.00	A novel dynamic recirculating apparatus for vapour-liquid equilibrium measurements at moderate pressures and temperatures
tbp	421.06	K	530.00	A novel dynamic recirculating apparatus for vapour-liquid equilibrium measurements at moderate pressures and temperatures
tbp	422.41	K	550.00	A novel dynamic recirculating apparatus for vapour-liquid equilibrium measurements at moderate pressures and temperatures
tbp	423.16	K	561.50	A novel dynamic recirculating apparatus for vapour-liquid equilibrium measurements at moderate pressures and temperatures

Correlations

Information	Value
Property code	pvap
Equation	$\ln(P_{vp}) = A + B/(T + C)$
Coeff. A	1.58913e+01
Coeff. B	-3.49716e+03
Coeff. C	-6.00740e+01
Temperature range (K), min.	284.20
Temperature range (K), max.	536.78

Information	Value
Property code	pvap
Equation	$\ln(P_{vp}) = A + B/T + C \cdot \ln(T) + D \cdot T^2$
Coeff. A	7.17170e+01
Coeff. B	-8.00269e+03
Coeff. C	-7.70223e+00
Coeff. D	3.95045e-07
Temperature range (K), min.	146.95
Temperature range (K), max.	536.71

Datasets

Viscosity, Pa*s

Temperature, K - Liquid	Pressure, kPa - Liquid	Viscosity, Pa*s - Liquid
298.15	81.50	0.0019430

Reference <https://www.doi.org/10.1016/j.jct.2016.12.036>

Pressure, kPa	Temperature, K	Viscosity, Pa*s
100.00	293.15	0.0022040
100.00	313.15	0.0013830

100.00	333.15	0.0009130
100.00	353.15	0.0006300
20000.00	293.15	0.0025450
20000.00	313.15	0.0016030
20000.00	333.15	0.0010860
20000.00	353.15	0.0007340
40000.00	293.15	0.0028550
40000.00	313.15	0.0018710
40000.00	333.15	0.0012630
40000.00	353.15	0.0008450
60000.00	293.15	0.0031910
60000.00	313.15	0.0021460
60000.00	333.15	0.0014390
60000.00	353.15	0.0009620
80000.00	293.15	0.0036130
80000.00	313.15	0.0024010
80000.00	333.15	0.0016140
80000.00	353.15	0.0010920
100000.00	293.15	0.0041490
100000.00	313.15	0.0026180
100000.00	333.15	0.0017860
100000.00	353.15	0.0012420

Reference

<https://www.doi.org/10.1021/je9002477>

Refractive index (Na D-line)

Pressure, kPa - Liquid

Temperature, K - Liquid

Refractive index (Na D-line) - Liquid

100.00	298.15	1.3836
--------	--------	--------

Reference

<https://www.doi.org/10.1016/j.jct.2016.01.012>

Temperature, K

Pressure, kPa

Refractive index (Na D-line)

294.15	100.00	1.3846
--------	--------	--------

Reference

<https://www.doi.org/10.1021/acs.jced.5b01024>

Mass density, kg/m³

Pressure, kPa - Liquid	Temperature, K - Liquid	Mass density, kg/m ³ - Liquid
100.00	293.15	803.5
100.00	303.15	795.2
100.00	313.15	787.6
100.00	323.15	779.3
100.00	333.15	770.5
100.00	343.15	761.0
100.00	353.15	752.2
5000.00	293.15	807.3
5000.00	303.15	799.2
5000.00	313.15	791.9
5000.00	323.15	783.7
5000.00	333.15	775.4
5000.00	343.15	766.0
5000.00	353.15	757.6
10000.00	293.15	810.9
10000.00	303.15	802.9
10000.00	313.15	795.9
10000.00	323.15	788.0
10000.00	333.15	780.0
10000.00	343.15	770.7
10000.00	353.15	762.7
15000.00	293.15	814.3
15000.00	303.15	806.7
15000.00	313.15	799.7
15000.00	323.15	792.0
15000.00	333.15	784.3
15000.00	343.15	775.3
15000.00	353.15	767.6
20000.00	293.15	817.7
20000.00	303.15	810.1
20000.00	313.15	803.3
20000.00	323.15	795.8
20000.00	333.15	788.4
20000.00	343.15	779.6
20000.00	353.15	772.1
25000.00	293.15	821.0
25000.00	303.15	813.6
25000.00	313.15	806.8
25000.00	323.15	799.7

25000.00	333.15	792.0
25000.00	343.15	783.7
25000.00	353.15	776.3
30000.00	293.15	824.0
30000.00	303.15	816.7
30000.00	313.15	810.2
30000.00	323.15	803.2
30000.00	333.15	795.7
30000.00	343.15	787.6
30000.00	353.15	780.4
35000.00	293.15	827.0
35000.00	303.15	819.8
35000.00	313.15	813.5
35000.00	323.15	806.6
35000.00	333.15	799.2
35000.00	343.15	791.2
35000.00	353.15	784.1
40000.00	293.15	829.9
40000.00	303.15	822.9
40000.00	313.15	816.6
40000.00	323.15	809.9
40000.00	333.15	802.6
40000.00	343.15	794.7
40000.00	353.15	787.7
45000.00	293.15	832.7
45000.00	303.15	825.8
45000.00	313.15	819.7
45000.00	323.15	813.0
45000.00	333.15	805.9
45000.00	343.15	798.3
45000.00	353.15	791.8
50000.00	293.15	835.4
50000.00	303.15	828.8
50000.00	313.15	822.7
50000.00	323.15	816.1
50000.00	333.15	809.1
50000.00	343.15	801.6
50000.00	353.15	795.1
55000.00	293.15	838.1
55000.00	303.15	831.6
55000.00	313.15	825.5
55000.00	323.15	819.0
55000.00	333.15	812.3
55000.00	343.15	804.6

55000.00	353.15	798.5
60000.00	293.15	840.7
60000.00	303.15	834.2
60000.00	313.15	828.3
60000.00	323.15	821.8
60000.00	333.15	815.2
60000.00	343.15	807.7
60000.00	353.15	801.7
65000.00	293.15	843.2
65000.00	303.15	837.0
65000.00	313.15	831.0
65000.00	323.15	824.6
65000.00	333.15	818.1
65000.00	343.15	810.6
65000.00	353.15	805.0

Reference

<https://www.doi.org/10.1016/j.fluid.2006.05.032>

Temperature, K	Pressure, kPa	Mass density, kg/m ³
298.15	110.00	799.54
298.15	5270.00	803.6
298.15	9870.00	807.01
298.15	14700.00	810.46
298.15	20530.00	814.46
298.15	25030.00	817.44
298.15	30100.00	820.7
298.15	35010.00	823.76
298.15	39560.00	826.52
323.15	590.00	779.78
323.15	5940.00	784.71
323.15	10600.00	788.81
323.15	16000.00	793.2
323.15	20640.00	796.73
323.15	24880.00	799.83
323.15	30800.00	803.99
323.15	35220.00	807.01
323.15	39790.00	810.02
348.15	140.00	757.11
348.15	5170.00	762.68
348.15	10260.00	767.5
348.15	15790.00	772.52
348.15	20560.00	776.41
348.15	24810.00	780.32

348.15	29300.00	784.09
348.15	35560.00	789.01
348.15	39950.00	792.23
373.15	360.00	733.24
373.15	5480.00	739.29
373.15	10290.00	745.11
373.15	15700.00	750.9
373.15	19960.00	755.19
373.15	25170.00	760.17
373.15	30510.00	765.02
373.15	35920.00	769.7
373.15	39780.00	772.91
398.15	620.00	705.55
398.15	5670.00	713.23
398.15	10620.00	720.1
398.15	15580.00	726.46
398.15	20350.00	732.35
398.15	25090.00	737.77
398.15	31150.00	744.07
398.15	35410.00	748.41
398.15	39390.00	752.32
423.15	1570.00	674.63
423.15	5870.00	682.79
423.15	10160.00	690.93
423.15	15320.00	699.32
423.15	20500.00	707.2
423.15	26280.00	715.01
423.15	31600.00	721.36
423.15	35780.00	726.03
423.15	39700.00	730.17

Reference

<https://www.doi.org/10.1016/j.fluid.2008.03.009>

Temperature, K	Pressure, kPa	Mass density, kg/m ³
293.15	100.00	803.8
293.15	10000.00	811.3
293.15	20000.00	818.1
293.15	30000.00	824.4
293.15	40000.00	830.0
293.15	50000.00	835.6
293.15	60000.00	841.0
293.15	70000.00	845.9
293.15	80000.00	850.7

293.15	90000.00	855.2
293.15	100000.00	859.6
293.15	110000.00	863.7
293.15	120000.00	867.8
293.15	130000.00	871.6
293.15	140000.00	875.3
313.15	100.00	787.1
313.15	10000.00	795.5
313.15	20000.00	803.0
313.15	30000.00	810.0
313.15	40000.00	816.4
313.15	50000.00	822.3
313.15	60000.00	827.9
313.15	70000.00	833.2
313.15	80000.00	838.3
313.15	90000.00	843.0
313.15	100000.00	847.9
313.15	110000.00	852.3
313.15	120000.00	856.6
313.15	130000.00	860.7
313.15	140000.00	864.6
333.15	100.00	769.9
333.15	10000.00	780.0
333.15	20000.00	788.3
333.15	30000.00	795.7
333.15	40000.00	802.5
333.15	50000.00	809.0
333.15	60000.00	815.0
333.15	70000.00	820.6
333.15	80000.00	826.2
333.15	90000.00	831.2
333.15	100000.00	836.3
333.15	110000.00	840.9
333.15	120000.00	845.3
333.15	130000.00	849.6
333.15	140000.00	853.6
353.15	100.00	752.0
353.15	10000.00	762.5
353.15	20000.00	771.8
353.15	30000.00	780.2
353.15	40000.00	787.9
353.15	50000.00	795.0
353.15	60000.00	801.4
353.15	70000.00	807.7

353.15	80000.00	813.4
353.15	90000.00	818.7
353.15	100000.00	823.8
353.15	110000.00	828.8
353.15	120000.00	833.5
353.15	130000.00	837.9
353.15	140000.00	842.1

Reference

<https://www.doi.org/10.1016/j.fluid.2013.11.031>

Temperature, K	Pressure, kPa	Mass density, kg/m3
278.15	100.00	815.7
278.15	1000.00	816.4
278.15	5000.00	819.4
278.15	10000.00	822.8
278.15	15000.00	826.2
278.15	20000.00	829.4
278.15	25000.00	832.6
278.15	30000.00	835.7
278.15	35000.00	838.6
278.15	40000.00	841.5
278.15	45000.00	844.3
278.15	50000.00	847.1
278.15	55000.00	849.7
278.15	60000.00	852.1
288.15	100.00	807.8
288.15	1000.00	808.5
288.15	5000.00	811.6
288.15	10000.00	815.3
288.15	15000.00	818.9
288.15	20000.00	822.3
288.15	25000.00	825.5
288.15	30000.00	828.7
288.15	35000.00	831.8
288.15	40000.00	834.7
288.15	45000.00	837.6
288.15	50000.00	840.4
288.15	55000.00	843.1
288.15	60000.00	845.6
298.15	100.00	799.8
298.15	1000.00	800.5
298.15	5000.00	803.8
298.15	10000.00	807.6

298.15	15000.00	811.4
298.15	20000.00	814.9
298.15	25000.00	818.3
298.15	30000.00	821.6
298.15	35000.00	824.8
298.15	40000.00	827.9
298.15	45000.00	830.9
298.15	50000.00	833.8
298.15	55000.00	836.6
298.15	60000.00	839.2
308.15	100.00	791.6
308.15	1000.00	792.4
308.15	5000.00	795.8
308.15	10000.00	800.1
308.15	15000.00	803.8
308.15	20000.00	807.5
308.15	25000.00	811.1
308.15	30000.00	814.5
308.15	35000.00	817.8
308.15	40000.00	821.0
308.15	45000.00	824.1
308.15	50000.00	827.1
308.15	55000.00	830.0
308.15	60000.00	832.7
318.15	100.00	783.4
318.15	1000.00	784.2
318.15	5000.00	787.9
318.15	10000.00	792.2
318.15	15000.00	796.3
318.15	20000.00	800.2
318.15	25000.00	803.9
318.15	30000.00	807.5
318.15	35000.00	810.9
318.15	40000.00	814.2
318.15	45000.00	817.5
318.15	50000.00	820.5
318.15	55000.00	823.6
318.15	60000.00	825.8
328.15	100.00	775.0
328.15	1000.00	775.9
328.15	5000.00	779.7
328.15	10000.00	784.2
328.15	15000.00	788.5
328.15	20000.00	792.6

328.15	25000.00	796.5
328.15	30000.00	800.2
328.15	35000.00	803.8
328.15	40000.00	807.2
328.15	45000.00	810.6
328.15	50000.00	813.8
328.15	55000.00	816.9
328.15	60000.00	819.8
338.15	100.00	766.3
338.15	1000.00	767.2
338.15	5000.00	771.3
338.15	10000.00	776.1
338.15	15000.00	780.6
338.15	20000.00	784.9
338.15	25000.00	789.0
338.15	30000.00	792.9
338.15	35000.00	796.6
338.15	40000.00	800.2
338.15	45000.00	803.7
338.15	50000.00	807.0
338.15	55000.00	810.3
338.15	60000.00	813.4
348.15	100.00	757.2
348.15	1000.00	758.3
348.15	5000.00	762.6
348.15	10000.00	767.7
348.15	15000.00	772.5
348.15	20000.00	777.0
348.15	25000.00	781.3
348.15	30000.00	785.4
348.15	35000.00	789.3
348.15	40000.00	793.0
348.15	45000.00	796.6
348.15	50000.00	800.1
348.15	55000.00	803.5
348.15	60000.00	806.5
358.15	100.00	747.7
358.15	1000.00	748.8
358.15	5000.00	753.5
358.15	10000.00	758.8
358.15	15000.00	764.0
358.15	20000.00	768.7
358.15	25000.00	773.3
358.15	30000.00	777.5

358.15	35000.00	781.6
358.15	40000.00	785.5
358.15	45000.00	789.3
358.15	50000.00	792.9
358.15	55000.00	796.4
358.15	60000.00	799.6

Reference

<https://www.doi.org/10.1016/j.jct.2011.10.023>

Temperature, K	Pressure, kPa	Mass density, kg/m ³
303.20	1000.00	796.7
303.20	2000.00	797.5
303.20	3000.00	798.3
303.20	4000.00	799.2
303.20	5000.00	799.9
303.20	6000.00	800.7
303.20	7000.00	801.5
303.20	8000.00	802.3
303.20	9000.00	803.1
303.20	10000.00	803.9
323.20	1000.00	779.0
323.20	2000.00	779.9
323.20	3000.00	780.8
323.20	4000.00	781.7
323.20	5000.00	782.6
323.20	6000.00	783.5
323.20	7000.00	784.4
323.20	8000.00	785.3
323.20	9000.00	786.2
323.20	10000.00	787.0

Reference

<https://www.doi.org/10.1016/j.jct.2012.10.002>

Temperature, K	Pressure, kPa	Mass density, kg/m ³
293.15	100.00	803.85

Reference

<https://www.doi.org/10.1016/j.jct.2016.12.011>

Temperature, K	Pressure, kPa	Mass density, kg/m ³
298.15	100.00	799.66

Temperature, K	Pressure, kPa	Mass density, kg/m ³
288.15	100.00	807.62
288.15	5000.00	811.32
288.15	10000.00	814.7
288.15	15000.00	817.95
288.15	20000.00	821.27
288.15	25000.00	824.24
288.15	30000.00	827.24
288.15	35000.00	830.04
288.15	40000.00	832.73
293.15	100.00	803.57
293.15	5000.00	807.24
293.15	10000.00	810.8
293.15	15000.00	814.22
293.15	20000.00	817.46
293.15	25000.00	820.74
293.15	30000.00	823.7
293.15	35000.00	826.5
293.15	40000.00	829.09
298.15	100.00	799.61
298.15	5000.00	803.49
298.15	10000.00	807.15
298.15	15000.00	810.5
298.15	20000.00	813.78
298.15	25000.00	816.92
298.15	30000.00	820.04
298.15	35000.00	822.87
298.15	40000.00	825.81
303.15	100.00	795.62
303.15	5000.00	799.57
303.15	10000.00	803.13
303.15	15000.00	806.69
303.15	20000.00	810.14
303.15	25000.00	813.48
303.15	30000.00	816.64
303.15	35000.00	819.75
303.15	40000.00	822.6
308.15	100.00	791.31
308.15	5000.00	795.4
308.15	10000.00	799.21

308.15	15000.00	802.79
308.15	20000.00	806.37
308.15	25000.00	809.72
308.15	30000.00	813.03
308.15	35000.00	816.12
308.15	40000.00	819.06

Reference

<https://www.doi.org/10.1016/j.jct.2019.02.011>

Temperature, K

Pressure, kPa

Mass density, kg/m3

298.15	81.50	799.54
--------	-------	--------

Reference

<https://www.doi.org/10.1021/acs.jced.5b00162>

Temperature, K

Pressure, kPa

Mass density, kg/m3

293.15	100.00	803.17
293.15	5000.00	807.25
293.15	10000.00	811.02
293.15	15000.00	814.54
293.15	20000.00	817.87
293.15	25000.00	821.05
293.15	30000.00	824.15
293.15	35000.00	827.21
303.15	100.00	795.52
303.15	5000.00	799.83
303.15	10000.00	803.8
303.15	15000.00	807.49
303.15	20000.00	810.95
303.15	25000.00	814.23
303.15	30000.00	817.38
303.15	35000.00	820.48
313.15	100.00	787.3
313.15	5000.00	791.61
313.15	10000.00	795.67
313.15	15000.00	799.52
313.15	20000.00	803.19
313.15	25000.00	806.72
313.15	30000.00	810.15
313.15	35000.00	813.5

Reference

<https://www.doi.org/10.1021/acs.jced.8b00975>

Temperature, K	Pressure, kPa	Mass density, kg/m3
303.15	100.00	795.6
303.15	5000.00	799.5
303.15	10000.00	803.4
303.15	15000.00	807.0
303.15	20000.00	810.5
303.15	25000.00	813.9
303.15	30000.00	817.2
313.15	100.00	787.6
313.15	5000.00	791.6
313.15	10000.00	795.7
313.15	15000.00	799.5
313.15	20000.00	803.2
313.15	25000.00	806.7
313.15	30000.00	810.1
323.15	100.00	779.1
323.15	5000.00	783.4
323.15	10000.00	787.7
323.15	15000.00	791.7
323.15	20000.00	795.6
323.15	25000.00	799.3
323.15	30000.00	802.9
333.15	100.00	770.6
333.15	5000.00	775.1
333.15	10000.00	779.6
333.15	15000.00	783.9
333.15	20000.00	787.9
333.15	25000.00	791.8
333.15	30000.00	795.5

Reference

<https://www.doi.org/10.1021/je049685z>

Temperature, K	Pressure, kPa	Mass density, kg/m3
283.15	100.00	810.9
283.15	5000.00	814.7
283.15	10000.00	818.1
283.15	15000.00	821.7
283.15	20000.00	824.9
298.15	100.00	799.0
298.15	5000.00	803.0
298.15	10000.00	806.9
298.15	15000.00	810.6
298.15	20000.00	814.1

313.15	100.00	786.8
313.15	5000.00	791.2
313.15	10000.00	795.3
313.15	15000.00	799.3
313.15	20000.00	803.1
328.15	100.00	774.5
328.15	5000.00	779.1
328.15	10000.00	783.7
328.15	15000.00	787.8
328.15	20000.00	791.9

Reference

<https://www.doi.org/10.1021/je100581m>

Speed of sound, m/s

Temperature, K - Liquid	Pressure, kPa - Liquid	Speed of sound, m/s - Liquid
253.15	102.00	1368.034
253.15	5133.00	1393.016
253.15	10109.00	1416.633
253.15	15198.00	1439.865
253.15	20092.00	1461.442
253.15	25341.00	1483.73
253.15	30423.00	1504.606
273.15	101.00	1293.525
273.15	5181.00	1320.802
273.15	10163.00	1346.251
273.15	15287.00	1371.412
273.15	20261.00	1394.803
273.15	25076.00	1416.612
273.15	30164.00	1438.924
293.15	105.00	1222.539
293.15	5042.00	1251.176
293.15	10141.00	1279.332
293.15	15272.00	1306.262
293.15	20136.00	1330.72
293.15	25068.00	1354.567
293.15	30231.00	1378.484
313.15	108.00	1154.108
313.15	5139.00	1185.752
313.15	10136.00	1215.403
313.15	15128.00	1243.591

313.15	20151.00	1270.557
313.15	25136.00	1296.129
313.01	30237.00	1321.177
333.15	101.00	1086.768
333.15	5018.00	1120.458
333.15	10050.00	1152.816
333.15	15186.00	1183.885
333.15	20183.00	1212.543
333.15	25121.00	1239.487
333.15	30028.00	1265.042
353.15	104.00	1018.903
353.15	5034.00	1055.962
353.15	10009.00	1090.768
353.15	15132.00	1124.22
353.15	20099.00	1154.783
353.15	25054.00	1183.626
353.15	30493.00	1213.605

Reference

<https://www.doi.org/10.1016/j.jct.2015.10.006>

Temperature, K	Pressure, kPa	Speed of sound, m/s
249.99	100.00	1382.6
218.20	102.00	1512.0
299.97	138.00	1199.6
249.99	177.00	1383.0
350.05	270.00	1030.8
400.05	307.00	847.6
218.26	875.00	1515.2
400.05	947.00	854.9
350.07	997.00	1036.3
450.12	1081.00	636.4
249.99	1128.00	1387.7
299.96	1174.00	1206.0
350.07	2073.00	1044.4
218.21	2111.00	1520.8
400.05	2155.00	867.0
450.12	2181.00	653.9
250.01	2296.00	1393.4
299.96	2560.00	1214.3
500.31	4798.00	438.4
400.03	4879.00	893.3
218.26	5018.00	1533.1
299.98	5042.00	1228.8

450.11	5172.00	697.0
350.07	5244.00	1067.5
249.97	5362.00	1408.3
249.94	9322.00	1427.1
218.32	9465.00	1551.0
299.98	9515.00	1254.1
500.32	9558.00	542.3
350.06	9697.00	1098.3
400.07	9768.00	936.6
450.11	9782.00	754.2
218.35	19810.00	1591.6
299.96	19849.00	1308.7
500.32	19968.00	690.2
400.07	19974.00	1015.3
350.08	20259.00	1164.3
450.10	20385.00	860.0
249.90	20903.00	1478.7
400.06	27823.00	1068.3
218.34	29071.00	1626.6
450.11	30113.00	938.6
399.98	30784.00	1087.2
500.31	30851.00	799.1
299.96	39521.00	1400.6
249.93	39564.00	1553.9
500.28	39800.00	870.0
450.13	40189.00	1008.7
399.96	40232.00	1142.9
350.04	40537.00	1273.0
218.31	49499.00	1698.1
299.96	49960.00	1444.6
450.13	50091.00	1069.5
399.98	50284.00	1196.8
500.30	50346.00	946.6
350.03	52186.00	1327.6
249.94	59448.00	1625.8
399.97	59616.00	1242.9
500.31	60509.00	1009.7
350.04	60588.00	1364.3
218.32	61345.00	1736.4
500.31	77533.00	1102.4

Reference

<https://www.doi.org/10.1021/acs.jced.8b00938>

Pressure, kPa

Temperature, K

Speed of sound, m/s

100.00	293.19	1222.98
100.00	298.12	1205.93
100.00	303.12	1188.72
100.00	308.20	1171.37
100.00	313.17	1154.51
100.00	318.20	1137.48
15200.00	293.12	1308.47
15200.00	298.11	1292.86
15200.00	303.14	1277.31
15200.00	308.20	1263.59
15200.00	313.16	1248.7
30400.00	293.22	1379.93
30400.00	298.15	1365.6
30400.00	303.24	1350.89
30400.00	308.13	1336.85
30400.00	313.16	1322.58
30400.00	318.16	1308.73
45590.00	293.21	1444.18
45590.00	298.14	1430.65
45590.00	303.13	1417.04
45590.00	308.13	1403.65
45590.00	313.16	1390.41
45590.00	318.14	1377.4
60790.00	293.21	1502.84
60790.00	298.14	1490.03
60790.00	303.20	1477.02
60790.00	308.13	1464.46
60790.00	313.17	1451.85
60790.00	318.12	1439.64
75990.00	293.20	1557.22
75990.00	298.14	1544.87
75990.00	303.15	1532.53
75990.00	308.12	1520.4
75990.00	313.16	1508.49
75990.00	318.12	1496.85
91190.00	293.20	1607.57
91190.00	298.14	1595.85
91190.00	303.17	1584.03
91190.00	308.20	1572.36
91190.00	313.15	1561.08
91190.00	318.12	1549.89
101320.00	293.18	1639.47
101320.00	298.14	1628.0
101320.00	303.17	1616.49

101320.00	308.19	1605.09
101320.00	313.15	1594.14
101320.00	318.11	1583.33
111450.00	293.17	1669.92
111450.00	298.14	1658.78
111450.00	303.19	1647.55
111450.00	308.16	1636.69
111450.00	313.22	1625.76
111450.00	318.17	1615.27
121580.00	293.16	1699.12
121580.00	298.14	1688.24
121580.00	303.18	1677.36
121580.00	308.11	1666.98
121580.00	313.17	1656.22
121580.00	318.17	1645.95

Reference

<https://www.doi.org/10.1021/je030136n>

Molar volume, m³/mol

Temperature, K - Liquid	Pressure, kPa - Liquid	Molar volume, m ³ /mol - Liquid
298.15	100.00	0.0001
298.15	10000.00	0.0001
313.15	100.00	0.0001
313.15	10000.00	0.0001
328.15	100.00	0.0001
328.15	10000.00	0.0001

Reference

<https://www.doi.org/10.1021/je800334m>

Molar heat capacity at constant volume, J/K/mol

Temperature, K - Liquid	Pressure, kPa - Liquid	Molar heat capacity at constant volume, J/K/mol - Liquid
314.15	4353.00	130.947
315.15	5217.00	131.0071
316.15	6081.00	131.7283
317.15	6946.00	132.6297

318.15	7808.00	133.2307
319.15	8666.00	133.2908
320.15	9526.00	133.6513
321.15	10392.00	134.6128
322.15	11260.00	134.9133
323.15	12130.00	135.8147
324.15	13000.00	135.8748
325.15	13871.00	136.3556
326.15	14740.00	137.4974
327.15	15599.00	136.9566
328.15	16461.00	139.1801
329.15	17326.00	138.519
330.15	18193.00	138.6993
331.15	19066.00	139.9012
332.15	19945.00	139.6007
333.15	20825.00	140.1416
334.15	21705.00	140.8026
335.15	22585.00	141.4637
336.15	23467.00	141.644
337.15	24350.00	141.2834
338.15	25234.00	142.5454
339.15	26119.00	142.7858
340.15	27006.00	143.7473
341.15	27893.00	143.0862
342.15	28774.00	145.0093
343.15	3874.00	148.615
343.15	29657.00	143.9276
344.15	4675.00	146.9924
345.15	5478.00	149.3361
346.15	6283.00	148.1943
347.15	7088.00	149.5164
348.15	7896.00	148.8554
349.15	8704.00	149.9972
350.15	9514.00	150.7784
351.15	10326.00	149.9972
352.15	11138.00	151.3193
353.15	11953.00	151.7399
354.15	12769.00	152.4611
355.15	13587.00	151.6798
356.15	14406.00	152.5813
357.15	15227.00	152.9418
358.15	16050.00	153.4827
359.15	16870.00	153.4226
360.15	17670.00	155.3456

361.15	18489.00	154.1437
362.15	19314.00	155.5259
363.15	20144.00	154.9851
364.15	20977.00	154.9851
365.15	21811.00	155.4057
366.15	22647.00	156.6677
367.15	23484.00	156.0067
368.15	24324.00	156.0668
369.15	25165.00	156.9682
370.15	26008.00	156.3071
371.15	26853.00	157.0884
372.15	2876.00	161.3551
372.15	27699.00	156.5475
373.15	3622.00	161.9561
373.15	28548.00	157.5691
374.15	4369.00	162.1965
374.15	29398.00	157.8095
375.15	5118.00	162.4969
376.15	5868.00	163.158
377.15	6620.00	163.158
378.15	7374.00	162.8575
379.15	8129.00	163.2181
380.15	8887.00	163.6387
381.15	9646.00	163.7589
382.15	10407.00	162.8575
383.15	11170.00	164.6003
384.15	11935.00	162.7373
385.15	12702.00	164.4801
386.15	13470.00	163.819
387.15	14240.00	164.2397
388.15	15013.00	164.1796
389.15	15786.00	164.2998
390.15	16562.00	164.2397
391.15	17338.00	164.8406
392.15	18116.00	163.4585
393.15	18896.00	164.3599
394.15	19677.00	164.2998
395.15	20459.00	164.3599
396.15	21244.00	164.5402
397.15	22030.00	164.6604
398.15	22817.00	164.7204
399.15	23606.00	164.8406
400.15	24395.00	163.9993
401.15	25186.00	164.3599

Liquid Liquid Equilibria of Quaternary Systems Composed of 1,3-Propanediol, Nitroethane, Acetic Acid, and Salt: Comparative Tensimetry, 1-Propanol Liquid-Liquid Equilibrium Data in Systems n-Propyl Acetate or Isopropyl Acetate + 1-Halobutane and 1,5-K and Solubility Measurements of Phase Interactions Modeling mixtures of 1-Propanol with alcohols, 1-Propanol, 2-Propanol, Water): Containing Hydrogen Phosphate of dimethyl sulfoxide + Acetic Acid of the FAS model for and volumetric properties of binary mixtures of butane with methanol and Solubility Determination of Flavonoid Temperature with Application Properties Parameters of 3-O-methyl-tert-butyl ether (MTE) in Various Organic Solvents: Leukopiperidinium-based ionic liquids in binary mixtures with alcohols and Refractive Indexes for Binary Mixtures of Ethanol and 2-methyl-2-butanol in 2-methyl-2-butanol and 2-methyl-2-butanol and 2-methyl-2-butanol and 2-methyl-2-butanol): Solubility of Glycidol in Hydrogen Halide: Excess Enthalpy of Lower alcohols: alkanols at T=298.15K, Measurement and Correlation of Solubility and Dissolution Parameters of 2,2,6,6-Tetramethylpiperidine-1-oxyl (TEMPO) in Binary Mixtures of (propylene oxide + Ethanol) and (propylene oxide + Methanol) and Propylene Glycol in Water at 313.15 K and Atmospheric Pressure: Functionalized Ionic Liquids Using Inverse Gas Chromatography: Mixtures of Diethyl Carbonate with Ketones (ethyl acetate, 3-pentanone) and Benzene with the Ionic Liquid CO₂-Bis(2-trifluoromethyl)ethylammonium Hexafluorophosphate Binary Mixtures of n-Undecane + 1-Propanol, n-Butanol, 2-Propanol, binary liquid mixtures of ethyl acetate, 1-propanol, 2-propanol, and ethanol and ethanol liquid mixtures of 1-propanol and 2-propanol at 298.15 and 323.15 K: New ionic liquid [P(4,4,4,4)NTf₂] in bio-butanol extraction on investigation Solubility Determination, Modeling, and Thermodynamic Dissolution Properties of Binary mixtures of n-Undecane + 1-Propanol, n-Butanol, 2-Propanol, binary liquid mixtures of ethyl acetate, 1-propanol, 2-propanol, and ethanol and ethanol liquid mixtures of 1-propanol and 2-propanol at 298.15 and 323.15 K: Vapor-Liquid Equilibrium Behaviors of 5-Methyl-2-(1-methylethyl)phenol in Aqueous Solutions of Fluconazole Solubility in Various Solvents at Different temperatures in 1-ethyl-3-methylimidazolium as solvent for Solubility Measurement and Correlation of Fluconazole in Six Organic Solvents and Different Binary Solvent Systems: Acetic Acid in Benzene and Methylacetone and Methylacetone Solubility of m-Phthalic Acid in Methylacetone and Solubility for Imatinib Mesylate in Nine Solvents and Its Correlation for Binary Diethyl Carbonate + 1-Alcohol Systems from 298.15 to 323.15 K: Potassium in Organic Solvents and Experimental study of the Density and Refractive Index (at 298.15 K and 303.15 K and partial molar volumes) properties of aqueous 1-propanol mixtures at temperatures from 298K to 582K and pressures up to 40MPa:

<https://www.doi.org/10.1021/acs.jced.6b00472>
<https://www.doi.org/10.1021/je0302295>
<https://www.doi.org/10.1021/acs.jced.8b00776>
<https://www.doi.org/10.1016/j.jct.2018.12.019>
<https://www.doi.org/10.1021/acs.jced.9b00350>
<https://www.doi.org/10.1021/je060232z>
<https://www.doi.org/10.1016/j.tca.2006.10.025>
<https://www.doi.org/10.1016/j.jct.2010.12.009>
<https://www.doi.org/10.1021/acs.jced.7b00118>
<https://www.doi.org/10.1016/j.jct.2011.02.018>
<https://www.doi.org/10.1016/j.jct.2013.12.029>
<https://www.doi.org/10.1021/je030226s>
<https://www.doi.org/10.1016/j.jct.2017.05.044>
<https://www.doi.org/10.1021/je200067e>
<https://www.doi.org/10.1016/j.tca.2005.03.009>
<https://www.doi.org/10.1021/je500012b>
<https://www.doi.org/10.1021/acs.jced.7b00206>
<https://www.doi.org/10.1016/j.jct.2015.02.006>
<https://www.doi.org/10.1016/j.fluid.2015.07.053>
<https://www.doi.org/10.1021/je900838a>
<https://www.doi.org/10.1021/je0600808>
<https://www.doi.org/10.1021/je200822w>
<https://www.doi.org/10.1021/acs.jced.6b00121>
<https://www.doi.org/10.1016/j.jct.2013.03.022>
<https://www.doi.org/10.1021/je800158z>
<https://www.doi.org/10.1021/acs.jced.9b00593>
<https://www.doi.org/10.1016/j.fluid.2018.07.028>
<https://www.doi.org/10.1021/acs.jced.9b00360>
<https://www.doi.org/10.1016/j.fluid.2010.10.008>
<https://www.doi.org/10.1016/j.jct.2019.02.002>
<https://www.doi.org/10.1016/j.jct.2011.11.021>
<https://www.doi.org/10.1016/j.fluid.2005.07.023>
<https://www.doi.org/10.1021/je900599q>
<https://www.doi.org/10.1021/je7007582>
<https://www.doi.org/10.1021/je4010257>
<https://www.doi.org/10.1016/j.fluid.2006.01.008>
<https://www.doi.org/10.1021/acs.jced.7b00638>
<https://www.doi.org/10.1021/je800676s>
<https://www.doi.org/10.1021/je800520w>
<https://www.doi.org/10.1021/acs.jced.8b00551>
<https://www.doi.org/10.1021/je1013476>
<https://www.doi.org/10.1021/acs.jced.7b00978>
<https://www.doi.org/10.1016/j.jct.2013.11.036>

Acoustic, volumetric and osmotic properties of binary mixtures of ethanol and 1-propanol

Thermophysical Properties of the Pure Ionic Liquid 1-butyl-3-methylimidazolium hexafluorophosphate with Binary and Ternary Mixtures of Ethanol, 1-Propanol, or 2-Propanol at 100 kPa: Measurements and Correlations of the Solubility of Resorcinol Bis(cyclic 2,2-dimethyl-1,3-dioxane) in Binary Solvent Mixtures: Experimental Measurements and Thermodynamic Modeling of Phenyl-pyrrole-2,5-dione in Binary and Ternary Vapor-Liquid Equilibrium at 323.15 K and Excess Molar Volumes at 298.15 K for the Methyl-*o*-(methylsulfanyl)benzene Solvent System: Thermodynamic Solubility of Florfenicol in Different Solvents at Temperatures from 278 to 323 K: Excess molar volumes of binary mixtures of

Study of molecular interactions in the mixtures of some secondary alcohols with different alcohols of ethanol and glycerol: coefficients and vapor pressures of binary systems 1- and 2-propanol with some aprotic solvents and mixtures: Co-solubility and thermodynamic properties of binary and ternary liquid mixtures: Measurements of activity coefficients at infinite dilution for organic solutes and water in the compressibility behaviour of ethyl (propyl) glycol - nitrate mixtures: Solubility of different 1-ethylimidazolium Nitrate and its binary mixtures with alcohols at several temperatures: Measurements of activity coefficients at infinite dilution for organic solutes in water: Measurements of the thermodynamic properties using GLC: Lactone in a pure and mixed solvents (p-cymene + propan-1-ol) at several temperatures: The physical properties of the COSMO-RS: Mixing thermodynamic properties of bis(1-methyl-2-pyrrolidone) and measurements of the critical mixtures of binary and ternary mixtures of ethanol, 1-propanol, 2-propanol, and water: Equilibrium of the systems: Ethanol and 1-propanol with 2,2,2-trifluoroethane in solvent mixtures of methyl acetate: Coefficient at infinite dilution of organic solutes in different solvents: The solution enthalpies of some organic solutes and water in different solvents: Graphy: Vapor-liquid equilibrium of different tetrahydrofuran

Excess molar enthalpies of ternary mixtures of Ethanol + 1-Propanol + Tetrahydrofuran: Properties of binary mixtures under elevated pressure: Dynamic Properties of Mixtures Containing Ionic Liquids. 4. LLE of binary mixtures containing 1-propanol, tetrahydrofuran, and ethyl acetate: Equilibrium measurements of the ternary systems: Ethanol + 1-propanol + water: Thermodynamic and volumetric and refractive index properties of methyl acetate and ethanol with solvents and alcohols at two solvent volumes: Properties of the binary alcohol mixtures at atmospheric pressure from 283.15 to 333.15 K: A new method for the determination of the density of pure boldine:

<https://www.doi.org/10.1016/j.jct.2012.02.007>

<https://www.doi.org/10.1021/jc300384g>

<https://www.doi.org/10.1021/acs.jced.7b00445>

<https://www.doi.org/10.1021/jc900888e>

<https://www.doi.org/10.1021/acs.jced.7b00695>

<https://www.doi.org/10.1021/acs.jced.9b00097>

<https://www.doi.org/10.1021/acs.jced.9b00560>

<https://www.doi.org/10.1021/jc800972b>

<https://www.doi.org/10.1016/j.jct.2016.08.021>

<https://www.doi.org/10.1021/jc101230a>

<https://www.doi.org/10.1021/jc1008284>

<https://www.doi.org/10.1016/j.jct.2006.09.017>

<https://www.doi.org/10.1016/j.jct.2013.09.005>

<https://www.doi.org/10.1016/j.jct.2011.03.012>

<https://www.doi.org/10.1016/j.jct.2016.07.013>

<https://www.doi.org/10.1021/jc0602723>

<https://www.doi.org/10.1016/j.fluid.2010.11.026>

<https://www.doi.org/10.1016/j.jct.2013.07.004>

<https://www.doi.org/10.1016/j.jct.2015.06.024>

<https://www.doi.org/10.1021/jc901058s>

<http://pubs.acs.org/doi/abs/10.1021/ci9903071>

<https://www.doi.org/10.1016/j.jct.2013.02.004>

<https://www.doi.org/10.1021/acs.jced.8b00720>

<https://www.doi.org/10.1016/j.jct.2018.03.021>

<https://www.doi.org/10.1016/j.jct.2011.08.028>

<https://www.doi.org/10.1016/j.jct.2008.02.019>

<https://www.doi.org/10.1021/acs.jced.6b00361>

<https://www.doi.org/10.1021/jc201047m>

<https://www.doi.org/10.1021/jc100577v>

<https://www.doi.org/10.1021/acs.jced.7b00982>

<https://www.doi.org/10.1016/j.fluid.2018.01.019>

<https://www.doi.org/10.1016/j.jct.2011.02.012>

<https://www.doi.org/10.1016/j.fluid.2017.09.007>

<https://www.doi.org/10.1016/j.jct.2011.09.028>

<https://www.doi.org/10.1021/jc049852v>

<https://www.doi.org/10.1021/jc030136n>

<https://www.doi.org/10.1021/jc049564z>

<https://www.doi.org/10.1016/j.tca.2008.09.006>

<https://www.doi.org/10.1021/jc050052+>

<https://www.doi.org/10.1016/j.jct.2012.03.005>

<https://www.doi.org/10.1016/j.jct.2005.05.004>

<https://www.doi.org/10.1016/j.jct.2015.11.004>

<https://www.doi.org/10.1016/j.fluid.2007.04.030>

Vapor-Liquid Equilibrium for Benzene + 2-Methylpentane and Allyl Alcohol + Propanol and Viscosities of Diaminotoluene with Water, Ethanol, Propanol, and the Binary Mixtures of Propanol with Ethanol and Methanol from Tugstien's Opticscrnazole in nine Solvents and in Sulfolane from 15 to 25 °C and in Different Solvents from 278.15 to 300.15 K: Liquid Equilibria in the Ternary System Dipropyl Ether + 1-Propanol + Potassium Acetate in Binary Systems Dipropyl Ether in Selected Solvents: Propanol Measurement and Calculation of the Solubility of Measurements of Activity Coefficients at Infinite Dilution of Aliphatic and Aromatic Hydrocarbons Equilibrium of 1-Propanol in Water and in Binary Systems of Water with Ethanol, Methanol, Ethyl Acetate, and Propanol, Ethyl Acetate, and Propanol and Reaction Isothermal Titration Calorimeters: An Evaluation of their Main Characteristics and the Transfer Properties of Binary Systems from the Data of the Literature and of the Measurements of the Activity Coefficients of Propanol in Different Solvents: Thermodynamic Modeling: 3-Methylimidazolium [12] Neat and in Aqueous Solutions of Various Organic Solutes at Infinite Dilution of Organic Solutes in the Ionic Liquid Solubility Determination and Thermodynamic Modelling of Gas-Liquid Equilibria in the System (3-Methylimidazolium) + 3H-pyrazol-2-one in Various Organic Solvents from T = 278.15 to 300.15 K: Measurement of the Solubility of Organic Solutes in Organic Solvents + Argon in Potassium/Sodium Acetate Solvent Systems: Experimental Determination of Propanol + Water, and Experimental Work of the System of 1-Alkyl-3-methylimidazolium Acetate in Different Solvents: Alcohols: Correlation, and Solvent Effect in the Solvent Extraction of Propanol: Pyrene Dissolved in Alcohol + Acetonitrile: Measurements of the Vapor Pressure of 1-Propanol/2-Propanol from their Experimental and Theoretical Study of Interactions between Organic Compounds in the 1-Propanol + 2-Propanol Systems in Terms of the Parameters of the Thermodynamic Properties of Binary Alkanol + Boldine Mixtures: Compressed Liquid Thermodynamic Properties of Binary Mixtures of n-Alcohols with Acrylonitrile at Infinite Dilution: Experimental and Theoretical Study of the Solubility of 3-Methylimidazolium Acetate in Different Solvents: Thermodynamic Modeling of the Solubility of Organic Solutes in Binary Systems: Water + Methanol, Ethanol, Propanol, and Acetonitrile: Infinite Dilution of Propanol and Ketones in Binary Systems: Experimental and Theoretical Study of the Solubility of Propanol in Water + NH₄Cl + 1-Propanol Systems at 0.1 MPa: Activity Coefficients of Organic Solutes and Water in the Ionic Liquid 1-butyl-3-methylimidazolium Hexafluorophosphate in Binary Systems: Solubility of Organic Solutes in Binary Systems: Acetone + Acetonitrile Binary Solvents from the Thermodynamic Properties and Solubility in Pure Organic Solvents: Topological Investigations of the Molecular Species and Molecular Interactions that Characterize Pyrrolidin-2-one + Lower Alkanol Mixtures:

<https://www.doi.org/10.1021/je0499519>

<https://www.doi.org/10.1021/je400968v>

<https://www.doi.org/10.1016/j.jct.2016.11.032>

<https://www.doi.org/10.1021/je800689a>

<https://www.doi.org/10.1016/j.fluid.2006.07.003>

<https://www.doi.org/10.1016/j.fluid.2012.05.003>

<https://www.doi.org/10.1016/j.jct.2012.06.005>

<https://www.doi.org/10.1016/j.jct.2008.12.018>

<https://www.doi.org/10.1021/je2002539>

<https://www.doi.org/10.1021/je1006289>

<https://www.doi.org/10.1016/j.jct.2011.12.018>

<https://www.doi.org/10.1016/j.fluid.2018.05.028>

<https://www.doi.org/10.1016/j.jct.2013.01.005>

<https://www.doi.org/10.1021/acs.jced.8b01250>

<https://www.doi.org/10.1016/j.jct.2015.09.010>

<https://www.doi.org/10.1016/j.jct.2010.12.019>

<https://www.doi.org/10.1016/j.jct.2016.08.007>

<https://www.doi.org/10.1021/acs.jced.8b00600>

<https://www.doi.org/10.1021/acs.jced.5b00122>

<https://www.doi.org/10.1021/acs.jced.5b00200>

<https://www.doi.org/10.1021/je901006c>

<https://www.doi.org/10.1021/je700099w>

<https://www.doi.org/10.1021/acs.jced.8b01101>

<https://www.doi.org/10.1021/je020226c>

<https://www.doi.org/10.1016/j.jct.2016.10.001>

<https://www.doi.org/10.1016/j.fluid.2014.06.021>

<https://www.doi.org/10.1021/je100595s>

<https://www.doi.org/10.1016/j.jct.2012.03.003>

<https://www.doi.org/10.1016/j.jct.2013.08.013>

<https://www.doi.org/10.1016/j.jct.2012.09.033>

<https://www.doi.org/10.1021/je700266n>

<https://www.doi.org/10.1016/j.jct.2009.11.015>

<https://www.doi.org/10.1021/je800022j>

<https://www.doi.org/10.1016/j.fluid.2015.09.003>

<https://www.doi.org/10.1016/j.fluid.2016.07.024>

<https://www.doi.org/10.1016/j.fluid.2018.03.003>

<https://www.doi.org/10.1016/j.jct.2012.08.016>

<https://www.doi.org/10.1021/je100352r>

<https://www.doi.org/10.1016/j.jct.2019.07.017>

<https://www.doi.org/10.1021/acs.jced.8b00717>

<https://www.thermofisher.com/files/research/kdb/mol/mol819.mol>

<https://www.doi.org/10.1016/j.fluid.2014.04.028>

<https://www.doi.org/10.1016/j.tca.2009.03.014>

Thermodynamic solubility modelling, solvent effect and preferential solvation
Measurement and correlation of the
solubility of (n-propanol, ethanol,
isopropanol and dimethyl sulfoxide)
alcohols with alcohols at temperatures
from 25 to 50 °C using density
guides (nitrophenyl) Ethanone and
Binary Solubility and Solubility
Determination and Model Correlation of
Binary and Ternary Systems of
Properties of binary systems of DMSO
with long chain alcohols and
Modeling of Densities and Refractive
Indices of the binary systems of alcohol
1-Dimethylammonium perchlorate
Apply of the Pitzer coefficients and
thermodynamic chemical properties for
organic solvent liquid phase for the
binary systems of propanol +
1-Dimethylammonium perchlorate
activity
coefficients measurements for organic
solvents at different temperatures and
of density of 1,2-dimethylaziridine in
liquid phase measurements from the
293.15 to 313.15 K Equilibrium of Ternary
Liquid Mixtures under Pressure:
Thermodynamic Properties of Solutions
Measurement and correlation solubility
and mixing properties of dimethyl
Sulfoxide of alcohols and aromatic
Compounds in liquid phase Based
on Pitzer and Correlation of the
Excess Properties of Ternary Mixture of
Excess Properties of Ternary
Mixtures of Ethanol, Acetone, +
Dimethylammonium perchlorate
solubilities of organic solvents in
aqueous urea and correlation between
dependence of the volumetric and
solubility properties of organic liquid
in aqueous urea with respect to the
hydrogen bonding capacity of urea and
a binary liquid phase of the pure ionic
liquid and urea in the presence of
binary liquid mixtures (alcohols)
of acetonitrile, propylene carbonate or
methyl formate of ethanol in
solubility of organic solvent mixtures of
ethyl acetate, several organic solvents:
Densities and refractive indices of
binary mixtures of ionic liquids
and alcohols in the presence of
and densities of binary mixtures and
density dependence between organic
solvents and urea and correlation with
dependence of the volumetric and
solubility properties of organic liquid
in aqueous urea
Thermodynamic modeling of
volumetric properties and speeds of
sound and refractive indices of binary
liquid mixtures of propanol and urea
of organic solubility urea and
Binary liquid phase of the pure ionic
liquid and urea in the presence of
binary liquid mixtures (alcohols)
of acetonitrile, propylene carbonate or
methyl formate of ethanol in
solubility of organic solvent mixtures of
ethyl acetate, several organic solvents:
Densities and refractive indices of
binary mixtures of ionic liquids
and alcohols in the presence of
and densities of binary mixtures and
density dependence between organic
solvents and urea and correlation with
dependence of the volumetric and
solubility properties of organic liquid
in aqueous urea
Thermodynamic modeling of
volumetric properties and speeds of
sound and refractive indices of binary
liquid mixtures of propanol and urea
of organic solubility urea and
Binary liquid phase of the pure ionic
liquid and urea in the presence of
binary liquid mixtures (alcohols)
of acetonitrile, propylene carbonate or
methyl formate of ethanol in
solubility of organic solvent mixtures of
ethyl acetate, several organic solvents:
Densities and refractive indices of
binary mixtures of ionic liquids
and alcohols in the presence of
and densities of binary mixtures and
density dependence between organic
solvents and urea and correlation with
dependence of the volumetric and
solubility properties of organic liquid
in aqueous urea

<https://www.doi.org/10.1016/j.jct.2019.02.016>
<https://www.doi.org/10.1016/j.tca.2016.01.009>
<https://www.doi.org/10.1016/j.jct.2013.09.032>
<https://www.doi.org/10.1021/acs.jced.8b00192>
<https://www.doi.org/10.1021/acs.jced.8b01181>
<https://www.doi.org/10.1016/j.jct.2013.04.009>
<https://www.doi.org/10.1016/j.tca.2011.08.002>
<https://www.doi.org/10.1016/j.jct.2016.01.026>
<https://www.doi.org/10.1016/j.jct.2013.01.007>
<https://www.doi.org/10.1016/j.fluid.2014.12.029>
<https://www.doi.org/10.1016/j.jct.2016.08.008>
<https://www.doi.org/10.1016/j.jct.2016.11.014>
<https://www.doi.org/10.1021/acs.jced.7b00110>
<https://www.doi.org/10.1021/je049685z>
<https://www.doi.org/10.1016/j.fluid.2013.06.037>
<https://www.doi.org/10.1021/je800331n>
<https://www.doi.org/10.1021/je401101u>
<https://www.doi.org/10.1021/je020149l>
<https://www.doi.org/10.1021/je100870x>
<https://www.doi.org/10.1016/j.jct.2018.03.007>
<https://www.doi.org/10.1021/acs.jced.7b00948>
<https://www.doi.org/10.1016/j.jct.2014.01.030>
<https://www.doi.org/10.1016/j.jct.2013.06.002>
<https://www.doi.org/10.1016/j.jct.2016.10.043>
<https://www.doi.org/10.1021/acs.jced.9b00258>
<https://www.doi.org/10.1021/je050134y>
<https://www.doi.org/10.1021/je500696p>
<https://www.doi.org/10.1016/j.fluid.2014.11.020>
<https://www.doi.org/10.1016/j.jct.2011.11.007>
<https://www.doi.org/10.1016/j.fluid.2014.06.005>
<https://www.doi.org/10.1021/acs.jced.9b00844>
<https://www.doi.org/10.1021/acs.jced.8b00023>
<https://www.doi.org/10.1021/acs.jced.6b00037>
<https://www.doi.org/10.1021/je050474j>
<https://www.doi.org/10.1021/je8004485>
<https://www.doi.org/10.1021/acs.jced.8b00139>
<https://www.doi.org/10.1016/j.jct.2013.03.002>
<https://www.doi.org/10.1021/je700426k>
<https://www.doi.org/10.1021/acs.jced.8b00043>
<https://www.doi.org/10.1021/acs.jced.8b00785>
<https://www.doi.org/10.1016/j.jct.2013.12.031>
<https://www.doi.org/10.1021/je1000582>
<https://www.doi.org/10.1021/je049778g>

Measurement and correlation of solubilities of apigenin and apigenin 7-O-glucuronide in different alcohols: Research on Dissolution Capability of Several Antifloxacin Salts: Solubility of Phenacetinum in Methanol, Ethanol, 1-Propanol, Separation of Purine from caffeine with Tri-n-butylammonium based ionic liquid: Vapor-Liquid Phase Equilibrium Measurements, Solubility of n-Propylamine in Water, Methanol, Ethanol, 2-Propanol, Ethylacetate, n-Butylacetate, and Tetrahydrofuran: Experimental and Thermodynamic Data of 2-butoxyethanol with aliphatic alcohols (1-butanol and 1-propanol) and partial molar volumes of binary mixtures of alcohols and 2-butoxyethanol: Properties of (2-methoxyethanol + 1-propanol) of solid-liquid phase diagrams in ternary systems by group contribution correlation, and Solvent Effect of Solubility of 2-isopropylindazole in Pure and Organic Solvents and Liquid Mixtures of Methanol and Acetate) Thermodynamic Modeling of Binary Mixtures of the Properties of Binary Mixtures of 2-(2-methoxyethoxy)ethylamine with 1-Propanol (LLE) for Methyl tert-Amyl Ether (TAME) + Water at 298.15 and 313.15 K and Properties of Binary Mixtures of TAME + 1-Propanol in the Temperature Range 298.15 K: Solubility of n-Propylamine in Ethanol, Methanol, and Acetate: Thermodynamic Studies of Molecular Interactions in Binary Mixtures of n-Propylamine, Hexane, Acetone, and n-Propylamine at 298.15 K: Thermodynamic Properties of Mixtures Containing n-Propylamine (2.9 Activity Coefficients) in the Dilution of Ethylacetate, n-Propylacetate, and Methylacetate with n-Propylamine at 298.15 K: Thermodynamic Properties of Binary Mixtures of Alcohols and Water: Vapor-Liquid Phase Equilibrium and Vapor Pressure Data for n-Propylamine and Water at Different Temperatures: Properties for Use over a Broad Range of Temperatures and Pressures: Thermodynamic Properties and their Description with the Soave-Redlich-Kwong Equation of State: Ternary Phase Diagrams of the Phase Separation Ability of KCl in Organic Solvent Systems Composed of Organic Solvents: Determination of the Phase Separation Ability of n-Propylamine in Short-Chain Alcohols: Vapor-Liquid Equilibria for 1-propanol + water + copper(II) chloride at 300 K: Physical Characterization of the Mixtures of the Ionic Liquid Solubility of Benzoic Acid and Acetic in Pure Solvents Using Focused Beam Reflective Measurement: Methanol, Ethanol, 1-Propanol, 2-Propanol, 1-Butanol and Water in the Binary Mixtures of Thermodynamic Properties of Mixtures Containing Ionic Liquids with Activity Coefficients: Partition of Hydrocarbons, Alcohols, Esters, and Aldehydes in Solubility of ciclesonide in seven pure solvents: Determination and Modeling of the Diffusion Coefficient of the Solvent Molecules in Ethylacetate and n-Propylamine: Ethanol, n-Propanol and Isopropylalcohol in the Binary Systems of 3,20-dione Bis(1-methyl-2-methylsulfonyl)imide-Based Ionic Liquids with Polar and Non-polar alcohols at different temperatures:

<https://www.doi.org/10.1016/j.jct.2010.09.013>
<https://www.doi.org/10.1016/j.jct.2011.10.023>
<https://www.doi.org/10.1021/acs.jced.8b00288>
<https://www.doi.org/10.1021/je700209v>
<https://www.doi.org/10.1016/j.fluid.2015.03.027>
<https://www.doi.org/10.1021/acs.jced.8b00033>
<https://www.doi.org/10.1021/je800196k>
<https://www.doi.org/10.1016/j.jct.2016.08.003>
<https://www.doi.org/10.1016/j.jct.2019.02.027>
<https://www.doi.org/10.1016/j.jct.2006.06.001>
<https://www.doi.org/10.1016/j.tca.2006.02.017>
<https://www.doi.org/10.1021/acs.jced.8b01226>
<https://www.doi.org/10.1016/j.jct.2016.12.028>
<https://www.doi.org/10.1016/j.jct.2016.07.009>
<https://www.doi.org/10.1016/j.tca.2007.04.012>
<https://www.doi.org/10.1021/je8006698>
<https://www.doi.org/10.1016/j.jct.2017.02.022>
<https://www.doi.org/10.1016/j.fluid.2007.01.021>
<https://www.doi.org/10.1016/j.jct.2008.01.008>
<https://www.doi.org/10.1021/je800739y>
<https://www.doi.org/10.1021/je050440b>
<https://www.doi.org/10.1016/j.jct.2005.04.015>
<https://www.doi.org/10.1021/je0497303>
<https://www.doi.org/10.1016/j.fluid.2012.11.003>
<https://www.doi.org/10.1021/je034197x>
<https://www.doi.org/10.1016/j.jct.2019.01.025>
<https://www.doi.org/10.1016/j.fluid.2009.03.011>
<https://www.doi.org/10.1021/je300568h>
<https://www.doi.org/10.1016/j.jct.2019.04.016>
<https://www.doi.org/10.1016/j.fluid.2011.09.027>
<https://www.doi.org/10.1016/j.jct.2013.07.025>
<https://www.doi.org/10.1016/j.fluid.2004.11.014>
<https://www.doi.org/10.1021/acs.jced.5b01023>
<https://www.doi.org/10.1021/acs.jced.5b00746>
<https://www.doi.org/10.1021/je1012839>
<https://www.doi.org/10.1016/j.jct.2016.01.012>
<https://www.doi.org/10.1021/je0503554>
<https://www.therich.org/research/kdb/hcprop/showprop.php?cmpid=819>
<https://www.doi.org/10.1016/j.jct.2016.10.014>
<https://www.doi.org/10.1021/acs.jced.8b00555>
<https://www.doi.org/10.1016/j.jct.2016.10.046>
<https://www.doi.org/10.1021/acs.jced.5b00181>
<https://www.doi.org/10.1016/j.jct.2014.04.011>

Liquid liquid equilibria for the ternary systems water + 1-propanol + methyl
 excess molar enthalpies of binary and
 Alcohol + Alkanols (C4-C6) and Their
 Correlations at 298.15 K and Activity
 coefficients of potassium phosphate in
 Ethanol/water mixed solvent
 methyl acetate ternary range of
 25.0 to 50.0 mole-% methyl acetate
 (operation) with seawater in 5.0
 Ethanol/water mixtures as
 Extractant in Ethanol/Water and
 1-Propanol/Water mixtures for Methyl
 Isobutyl Ketone (MIBK) + (1-Propanol
 Solubility Measurements of Lauric,
 Palmitic, and Stearic Acids in Ethanol,
 Propanol, and 1-Propanol Using
 Differential Scanning Calorimetry
 by Ethanol/Water and 1-Propanol
 and triethylphenylenediamine in nine
 Measurments and Correlation of Ethanol
 Solubility of Gatifloxacin in 12 Pure
 Solvents from 298.15 to 313.15 K and
 Ethyl-3-methylimidazolium acetate
 Chromatographic and Diffusion Excess
 enthalpies of binary mixtures
 solubility of 2-propanol in different
 solvent systems at different
 temperatures using a new flow
 Solubility, liquid-liquid equilibrium and
 critical states for the system acetic acid
 infinite dilution activity coefficient and
 Solutes Dissolved in Two
 Non-aqueous Media by Hydrophilic Ionic
 Solvents: Mixtures of
 2-Ethoxyethanol and Propanol
 for the Binary and Ternary Systems of
 Propanol and Propanol
 the molecular parameters in 288.15
 description of the binary mixtures of
 methyl ethanol mixtures from (293.15 to
 densities and volumetric properties of
 (formamide + ethanol, or 1-propanol,
 solubility of oil in 2-propanol in nine
 organic solvents: Experimental
 data and correlations from the molecular
 Properties Databank);
 Determination of activity coefficients at
 infinite dilution of organic solutes in
 the a Equilibrium Properties of Binary
 and Ternary Mixtures Containing
 Alcohols and Glycols in Infinite
 dilution Solubility at T = 313.15 K:
 Ethyl-3-methylimidazolium acetate
 using gas-liquid chromatography:
 measurement and correlation
 solubility of carbendazim in water at T =
 liquid-liquid equilibria for Reactive
 Distillation of Propyl Propanoate. Pure
 Component Properties for (Ionic Liquid
 and Ethanol/Ethanol and Propanol +
 Ethanol and Ethanol). K and
 Anomalous Pressure:
 2,6-Dichloro-4-nitrophenol and 2 Pure
 binary systems: experimental and
 computational investigation of excess
 Measurments of Infinite Dilution
 Activity Coefficients of Alcohols,
 Ketones, and Aldehydes and Propanol
 investigations of binary liquid
 systems of 2-propanol with water in
 Methanol at temperature Propanol of
 water and 2-propanol in the water
 solubility of 2-(4-methoxyphenyl)acetic Acid
 in alcohols and binary systems pure
 alkanols:
 Liquid-Liquid Equilibrium of
 Hydrophilic Alcohols with Three
 Binary systems by ionic-based ionic
 liquid in the separation processes:
 Present status of the modified UNIFAC
 Model for the prediction of phase
 Study of the liquid-liquid equilibrium of
 systems with manganese sulphate and
 2-propanol + lithium sulphate aqueous
 two-phase systems at different
 temperatures: Experiment and
 correlation:

<https://www.doi.org/10.1016/j.fluid.2007.05.018>

<https://www.doi.org/10.1021/je100095p>

<https://www.doi.org/10.1016/j.jct.2008.01.018>

<https://www.doi.org/10.1016/j.tca.2005.11.041>

<https://www.doi.org/10.1016/j.jct.2015.05.014>

<https://www.doi.org/10.1021/je400791n>

<https://www.doi.org/10.1021/acs.jced.7b00026>

<https://www.doi.org/10.1021/acs.jced.8b01044>

<https://www.doi.org/10.1021/acs.jced.7b00098>

<https://www.doi.org/10.1016/j.jct.2017.01.006>

<https://www.doi.org/10.1021/acs.jced.8b00902>

<https://www.doi.org/10.1016/j.jct.2011.11.005>

<https://www.doi.org/10.1016/j.jct.2008.12.004>

<https://www.doi.org/10.1016/j.jct.2016.07.049>

<https://www.doi.org/10.1021/acs.jced.5b00882>

<https://www.doi.org/10.1016/j.fluid.2014.04.034>

<https://www.doi.org/10.1021/je500050p>

<https://www.doi.org/10.1021/je7005049>

<https://www.doi.org/10.1021/acs.jced.8b00067>

<https://www.doi.org/10.1016/j.jct.2014.12.023>

<https://www.doi.org/10.1021/je3011634>

<https://www.doi.org/10.1016/j.jct.2006.07.021>

<https://www.doi.org/10.1016/j.jct.2018.09.017>

<https://www.thermo.com/research/kdb/hcprop/showprop.php?cmpid=819>

<https://www.doi.org/10.1016/j.fluid.2005.04.021>

<https://www.doi.org/10.1021/je060253t>

<https://www.doi.org/10.1016/j.fluid.2012.05.006>

<https://www.doi.org/10.1016/j.jct.2009.12.001>

<https://www.doi.org/10.1016/j.tca.2017.12.007>

<https://www.doi.org/10.1021/je101302p>

<https://www.doi.org/10.1021/je100830b>

<https://www.doi.org/10.1021/acs.jced.9b00778>

<https://www.doi.org/10.1016/j.jct.2019.06.018>

<https://www.doi.org/10.1016/j.jct.2014.09.001>

<https://www.doi.org/10.1021/je800246v>

<https://www.doi.org/10.1016/j.fluid.2013.06.021>

<https://www.doi.org/10.1021/je100964a>

<https://www.doi.org/10.1021/acs.jced.6b00957>

<https://www.doi.org/10.1016/j.fluid.2008.11.005>

<https://www.doi.org/10.1021/acs.jced.5b00771>

<https://www.doi.org/10.1016/j.jct.2016.01.017>

<https://www.doi.org/10.1016/j.fluid.2014.03.006>

<https://www.doi.org/10.1016/j.fluid.2011.10.004>

Solid-liquid equilibria for selected binary systems containing diphenyl ether
Densities and Viscosities of Binary Mixtures Containing Diphenyl ether
Densities and viscosities of binary mixtures of diphenyl ether with aliphatic lower alkanols (C1-C3) at temperatures from T = 303.15 K to T = 323.15 K:

<https://www.doi.org/10.1016/j.fluid.2018.09.023>

<https://www.doi.org/10.1021/je401124t>

<https://www.doi.org/10.1016/j.jct.2006.10.016>

Legend

af:	Acentric Factor
affp:	Proton affinity
aigt:	Autoignition Temperature
basg:	Gas basicity
chl:	Standard liquid enthalpy of combustion
cpg:	Ideal gas heat capacity
cpl:	Liquid phase heat capacity
cps:	Solid phase heat capacity
dm:	Dipole Moment
dvisc:	Dynamic viscosity
fll:	Lower Flammability Limit
flu:	Upper Flammability Limit
fpc:	Flash Point (Closed Cup Method)
fpo:	Flash Point (Open Cup Method)
gf:	Standard Gibbs free energy of formation
gyrad:	Radius of Gyration
hf:	Enthalpy of formation at standard conditions
hfl:	Liquid phase enthalpy of formation at standard conditions
hfus:	Enthalpy of fusion at standard conditions
hfust:	Enthalpy of fusion at a given temperature
hvap:	Enthalpy of vaporization at standard conditions
hvapt:	Enthalpy of vaporization at a given temperature
ie:	Ionization energy
log10ws:	Log10 of Water solubility in mol/l
logp:	Octanol/Water partition coefficient
mcvol:	McGowan's characteristic volume
nfpaf:	NFPA Fire Rating
nfpah:	NFPA Health Rating
pc:	Critical Pressure
pvap:	Vapor pressure
rfi:	Refractive Index
rho:	Liquid Density
rinpol:	Non-polar retention indices
ripol:	Polar retention indices
sfust:	Entropy of fusion at a given temperature

sg:	Molar entropy at standard conditions
sl:	Liquid phase molar entropy at standard conditions
speedsl:	Speed of sound in fluid
srf:	Surface Tension
ss:	Solid phase molar entropy at standard conditions
tb:	Normal Boiling Point Temperature
tbp:	Boiling point at given pressure
tc:	Critical Temperature
tf:	Normal melting (fusion) point
tt:	Triple Point Temperature
vc:	Critical Volume
volm:	Molar Volume
zc:	Critical Compressibility
zra:	Rackett Parameter

Latest version available from:

<https://www.chemeo.com/cid/15-582-9/1-Propanol.pdf>

Generated by Cheméo on 2025-12-22 19:39:19.555048808 +0000 UTC m=+6180557.085089474.

Cheméo (<https://www.chemeo.com>) is the biggest free database of chemical and physical data for the process industry.