1-Propanol

Other names:	1-Hydroxypropane
	1-Propyl alcohol
	Alcohol, propyl
	Alcool propilico
	Alcool propylique
	Ethylcarbinol
	N-PROPYL ALCOHOL
	NSC 30300
	Optal
	Osmosol extra
	PROPYL ALCOHOL
	Propan-1-ol
	Propanol
	Propanol-1
	Propanole
	Propanolen
	Propanoli
	Propylan-propyl alcohol
	Propylic alcohol
	Propylowy alkohol
	UN 1274
	n-C3H7OH
	n-Propan-1-ol
	n-Propanol
	n-Propyl alkohol
Inchi:	InChI=1S/C3H8O/c1-2-3-4/h4H,2-3H2,1H3
InchiKey:	BDERNNFJNOPAEC-UHFFFAOYSA-N
Formula:	C3H8O
SMILES:	CCCO
Mol. weight [g/mol]:	60.09
CAS:	71-23-8

Physical Properties

Property code	Value	Unit	Source
af	0.6230		KDB
affp	786.50	kJ/mol	NIST Webbook

aigt	644.26	K	KDB
basg	756.10	kJ/mol	NIST Webbook
chl	-2021.31 ± 0.25	kJ/mol	NIST Webbook
chl	-2021.40 ± 0.75	kJ/mol	NIST Webbook
chl	-2019.40 ± 0.30	kJ/mol	NIST Webbook
chl	-2032.59	kJ/mol	NIST Webbook
chl	-2017.70 ± 1.00	kJ/mol	NIST Webbook
cpl	143.41	J/mol×K	THERMODYNAMICS OF MIXTURES CONTAINING AMINES. XIV. CpEm OF BENZYLAMINE WITH HEPTANE AT 293.15 K OR WITH METHANOL, 1-PROPANOL OR 1-PENTANOL AT (293.15-308.15) K
dm	1.70	debye	KDB
dvisc	0.0019544	Pa×s	Densities and Viscosities of Binary Liquid Mixtures of Trichloroethylene and Tetrachloroethylene with Some Polar and Nonpolar Solvents
fll	2.10	% in Air	KDB
flu	13.50	% in Air	KDB
fpc	300.37	K	KDB
fpo	298.15	K	KDB
gf	-161.90	kJ/mol	KDB
gyrad	2.7360		KDB
hf	-255.10	kJ/mol	NIST Webbook
hf	-257.30 ± 0.40	kJ/mol	NIST Webbook
hf	-257.20	kJ/mol	NIST Webbook
hf	-258.80 ± 1.10	kJ/mol	NIST Webbook
hf	-254.70 ± 4.40	kJ/mol	NIST Webbook
hf	-256.60	kJ/mol	KDB
hf	-255.60 ± 1.30	kJ/mol	NIST Webbook
hf	-255.20 ± 0.30	kJ/mol	NIST Webbook
hfl	-304.60 ± 0.40	kJ/mol	NIST Webbook
hfl	-302.50 ± 4.20	kJ/mol	NIST Webbook
hfl	-303.00 ± 1.30	kJ/mol	NIST Webbook
hfl	-302.54 ± 0.25	kJ/mol	NIST Webbook
hfl	-306.30 ± 1.00	kJ/mol	NIST Webbook
hfus	5.40	kJ/mol	Heat Capacities and Derived Thermodynamic Functions of 1-Propanol between 10 K and 350 K and of 1-Pentanol between 85 K and 370 K
hvap	38.95	kJ/mol	Joback Method
ie	10.52 ± 0.03	eV	NIST Webbook

ie	10.48	eV	NIST Webbook
ie	10.20	eV	NIST Webbook
ie	10.49	eV	NIST Webbook
ie	10.25	eV	NIST Webbook
ie	10.22 ± 0.06	eV	NIST Webbook
ie	10.32 ± 0.02	eV	NIST Webbook
ie	10.16 ± 0.03	eV	NIST Webbook
ie	10.51	eV	NIST Webbook
ie	10.15 ± 0.03	eV	NIST Webbook
ie	10.00	eV	NIST Webbook
ie	10.22 ± 0.07	eV	NIST Webbook
ie	10.51	eV	NIST Webbook
ie	10.22 ± 0.04	eV	NIST Webbook
log10ws	0.62		Estimated Solubility Method
log10ws	0.62		Aqueous Solubility Prediction Method
logp	0.389		Crippen Method
mcvol	59.000	ml/mol	McGowan Method
nfpaf	%!d(float64=3)		KDB
nfpah	%!d(float64=1)		KDB
pc	5169.00	kPa	KDB
pc	5170.00 ± 20.00	kPa	NIST Webbook
pc	5170.00	kPa	NIST Webbook
pc	5182.00 ± 10.00	kPa	NIST Webbook
pc	5168.00 ± 20.00	kPa	NIST Webbook
pc	5170.00	kPa	NIST Webbook
pc	5170.00	kPa	NIST Webbook
pc	5082.00	kPa	NIST Webbook
pc	5218.00	kPa	NIST Webbook
pc	5082.00	kPa	NIST Webbook
pc	5397.00	kPa	NIST Webbook
pc	5155.00	kPa	NIST Webbook
pc	5168.00 ± 20.00	kPa	NIST Webbook
rinpol	551.00		NIST Webbook
rinpol	557.00		NIST Webbook
rinpol	561.00		NIST Webbook
rinpol	549.00		NIST Webbook
rinpol	553.00		NIST Webbook
rinpol	536.00		NIST Webbook
rinpol	559.00		NIST Webbook
rinpol	559.00		NIST Webbook
rinpol	559.00		NIST Webbook
rinpol	548.00		NIST Webbook
rinpol	557.90		NIST Webbook

rinpol	568.00	NIST Webbook
rinpol	568.00	NIST Webbook
rinpol	568.00	NIST Webbook
rinpol	554.00	NIST Webbook
rinpol	553.00	NIST Webbook
rinpol	560.00	NIST Webbook
rinpol	539.00	NIST Webbook
rinpol	538.64	NIST Webbook
rinpol	521.00	NIST Webbook
rinpol	535.00	NIST Webbook
rinpol	548.00	NIST Webbook
rinpol	561.00	NIST Webbook
rinpol	555.00	NIST Webbook
rinpol	568.00	NIST Webbook
rinpol	547.00	NIST Webbook
rinpol	544.00	NIST Webbook
rinpol	548.00	NIST Webbook
rinpol	544.00	NIST Webbook
rinpol	557.00	NIST Webbook
rinpol	558.00	NIST Webbook
rinpol	535.00	NIST Webbook
rinpol	544.00	NIST Webbook
rinpol	548.00	NIST Webbook
rinpol	548.00	NIST Webbook
rinpol	535.00	NIST Webbook
rinpol	570.00	NIST Webbook
rinpol	560.00	NIST Webbook
rinpol	539.00	NIST Webbook
rinpol	560.00	NIST Webbook
rinpol	562.00	NIST Webbook
rinpol	543.00	NIST Webbook
rinpol	546.00	NIST Webbook
rinpol	546.00	NIST Webbook
rinpol	539.00	NIST Webbook
rinpol	571.00	NIST Webbook
rinpol	534.00	NIST Webbook
rinpol	555.00	NIST Webbook
rinpol	535.00	NIST Webbook
rinpol	581.00	NIST Webbook
rinpol	571.00	NIST Webbook
rinpol	572.00	NIST Webbook
rinpol	561.20	NIST Webbook
rinpol	543.70	NIST Webbook
rinpol	540.50	NIST Webbook

rinpol	544.60	NIST Webbook
rinpol	553.50	NIST Webbook
rinpol	555.00	NIST Webbook
rinpol	525.00	NIST Webbook
rinpol	561.00	NIST Webbook
rinpol	557.00	NIST Webbook
rinpol	559.00	NIST Webbook
rinpol	568.00	NIST Webbook
rinpol	560.00	NIST Webbook
rinpol	548.00	NIST Webbook
rinpol	544.00	NIST Webbook
rinpol	507.00	NIST Webbook
rinpol	537.00	NIST Webbook
rinpol	530.00	NIST Webbook
rinpol	533.00	NIST Webbook
rinpol	536.00	NIST Webbook
rinpol	552.00	NIST Webbook
rinpol	552.00	NIST Webbook
rinpol	532.00	NIST Webbook
rinpol	552.00	NIST Webbook
rinpol	546.00	NIST Webbook
rinpol	540.20	NIST Webbook
rinpol	540.20	NIST Webbook
rinpol	537.00	NIST Webbook
rinpol	556.00	NIST Webbook
rinpol	557.00	NIST Webbook
rinpol	555.00	NIST Webbook
rinpol	551.00	NIST Webbook
rinpol	521.00	NIST Webbook
rinpol	524.00	NIST Webbook
rinpol	568.00	NIST Webbook
rinpol	568.00	NIST Webbook
rinpol	568.00	NIST Webbook
rinpol	551.00	NIST Webbook
rinpol	595.00	NIST Webbook
rinpol	555.00	NIST Webbook
rinpol	561.00	NIST Webbook
rinpol	574.00	NIST Webbook
rinpol	523.00	NIST Webbook
rinpol	547.00	NIST Webbook
rinpol	520.00	NIST Webbook
rinpol	562.00	NIST Webbook
rinpol	525.00	NIST Webbook
rinpol	544.00	NIST Webbook

rinpol	530.00	NIST Webbook
rinpol	539.90	NIST Webbook
rinpol	530.00	NIST Webbook
rinpol	555.00	NIST Webbook
rinpol	546.00	NIST Webbook
rinpol	575.00	NIST Webbook
rinpol	551.40	NIST Webbook
rinpol	553.50	NIST Webbook
rinpol	556.50	NIST Webbook
rinpol	541.50	NIST Webbook
rinpol	540.50	NIST Webbook
rinpol	542.30	NIST Webbook
rinpol	544.60	NIST Webbook
rinpol	551.40	NIST Webbook
rinpol	553.60	NIST Webbook
rinpol	556.50	NIST Webbook
rinpol	541.50	NIST Webbook
rinpol	540.50	NIST Webbook
rinpol	542.30	NIST Webbook
rinpol	544.60	NIST Webbook
rinpol	521.00	NIST Webbook
rinpol	536.90	NIST Webbook
rinpol	543.70	NIST Webbook
rinpol	549.70	NIST Webbook
rinpol	558.20	NIST Webbook
rinpol	559.30	NIST Webbook
rinpol	561.40	NIST Webbook
rinpol	561.20	NIST Webbook
rinpol	565.00	NIST Webbook
rinpol	560.00	NIST Webbook
ripol	1026.00	NIST Webbook
ripol	1056.00	NIST Webbook
ripol	1021.00	NIST Webbook
ripol	1014.00	NIST Webbook
ripol	993.00	NIST Webbook
ripol	1036.00	NIST Webbook
ripol	1030.00	NIST Webbook
ripol	1043.00	NIST Webbook
ripol	1005.00	NIST Webbook
ripol	1002.00	NIST Webbook
ripol	1025.00	NIST Webbook
ripol	1037.00	NIST Webbook
ripol	1060.00	NIST Webbook
ripol	1061.00	NIST Webbook

ripol	1030.00	NIST Webbook
ripol	1049.00	NIST Webbook
ripol	1049.00	NIST Webbook
ripol	1025.00	NIST Webbook
ripol	1015.00	NIST Webbook
ripol	1038.00	NIST Webbook
ripol	1026.00	NIST Webbook
ripol	1045.00	NIST Webbook
ripol	1041.00	NIST Webbook
ripol	1052.00	NIST Webbook
ripol	1040.00	NIST Webbook
ripol	1045.00	NIST Webbook
ripol	1073.00	NIST Webbook
ripol	1041.00	NIST Webbook
ripol	1038.00	NIST Webbook
ripol	1038.00	NIST Webbook
ripol	1064.00	NIST Webbook
ripol	1049.00	NIST Webbook
ripol	1025.00	NIST Webbook
ripol	1025.00	NIST Webbook
ripol	1031.00	NIST Webbook
ripol	1052.00	NIST Webbook
ripol	1052.00	NIST Webbook
ripol	1051.00	NIST Webbook
ripol	1052.00	NIST Webbook
ripol	1052.00	NIST Webbook
ripol	1037.00	NIST Webbook
ripol	1029.00	NIST Webbook
ripol	1041.00	NIST Webbook
ripol	1032.00	NIST Webbook
ripol	1034.00	NIST Webbook
ripol	1031.00	NIST Webbook
ripol	1033.00	NIST Webbook
ripol	1028.00	NIST Webbook
ripol	1023.00	NIST Webbook
ripol	1040.00	NIST Webbook
ripol	1037.00	NIST Webbook
ripol	1045.00	NIST Webbook
ripol	1034.00	NIST Webbook
ripol	1025.00	NIST Webbook
ripol	1035.00	NIST Webbook
ripol	1002.00	NIST Webbook
ripol	1030.00	NIST Webbook
ripol	1030.00	NIST Webbook

ripol	1043.00	NIST Webbook
ripol	1037.00	NIST Webbook
ripol	1029.00	NIST Webbook
ripol	1025.00	NIST Webbook
ripol	1040.00	NIST Webbook
ripol	1038.00	NIST Webbook
ripol	1045.00	NIST Webbook
ripol	1006.00	NIST Webbook
ripol	1038.00	NIST Webbook
ripol	1032.00	NIST Webbook
ripol	1002.00	NIST Webbook
ripol	1069.00	NIST Webbook
ripol	1046.00	NIST Webbook
ripol	1046.00	NIST Webbook
ripol	1002.00	NIST Webbook
ripol	1042.00	NIST Webbook
ripol	1022.00	NIST Webbook
ripol	1002.00	NIST Webbook
ripol	1030.00	NIST Webbook
ripol	990.00	NIST Webbook
ripol	996.00	NIST Webbook
ripol	990.00	NIST Webbook
ripol	996.00	NIST Webbook
ripol	1042.00	NIST Webbook
ripol	1040.00	NIST Webbook
ripol	1036.00	NIST Webbook
ripol	1038.00	NIST Webbook
ripol	1045.00	NIST Webbook
ripol	992.00	NIST Webbook
ripol	1030.00	NIST Webbook
ripol	1036.00	NIST Webbook
ripol	1086.00	NIST Webbook
ripol	1036.00	NIST Webbook
ripol	1036.00	NIST Webbook
ripol	1051.00	NIST Webbook
ripol	1052.00	NIST Webbook
ripol	1046.00	NIST Webbook
ripol	1036.00	NIST Webbook
ripol	1036.00	NIST Webbook
ripol	1036.00	NIST Webbook
ripol	1032.00	NIST Webbook
ripol	1035.00	NIST Webbook
ripol	1044.00	NIST Webbook
ripol	1036.00	NIST Webbook

ripol	1038.00	NIST Webbook
ripol	1042.00	NIST Webbook
ripol	1036.00	NIST Webbook
ripol	1042.00	NIST Webbook
ripol	1042.00	NIST Webbook
ripol	1047.00	NIST Webbook
ripol	1042.00	NIST Webbook
ripol	1025.00	NIST Webbook
ripol	1030.00	NIST Webbook
ripol	1031.00	NIST Webbook
ripol	1002.00	NIST Webbook
ripol	1072.00	NIST Webbook
ripol	1033.00	NIST Webbook
ripol	1033.00	NIST Webbook
ripol	1014.00	NIST Webbook
ripol	1040.00	NIST Webbook
ripol	1030.00	NIST Webbook
ripol	1051.00	NIST Webbook
ripol	1036.00	NIST Webbook
ripol	1038.00	NIST Webbook
ripol	1047.00	NIST Webbook
ripol	1002.00	NIST Webbook
ripol	1037.00	NIST Webbook
ripol	1025.00	NIST Webbook
ripol	1049.00	NIST Webbook
ripol	1026.00	NIST Webbook
ripol	1045.00	NIST Webbook
ripol	1064.00	NIST Webbook
ripol	1052.00	NIST Webbook
ripol	1031.00	NIST Webbook
ripol	1046.00	NIST Webbook
ripol	1046.90	NIST Webbook
ripol	1000.00	NIST Webbook
ripol	1056.00	NIST Webbook
ripol	1031.00	NIST Webbook
ripol	1049.00	NIST Webbook
ripol	993.00	NIST Webbook
ripol	1035.00	NIST Webbook
ripol	1040.00	NIST Webbook
ripol	997.00	NIST Webbook
ripol	1000.00	NIST Webbook
ripol	988.00	NIST Webbook
ripol	1072.00	NIST Webbook
ripol	1054.90	NIST Webbook

ripol	1051.70		NIST Webbook
ripol	1046.90		NIST Webbook
ripol	1037.60		NIST Webbook
ripol	1024.00		NIST Webbook
ripol	1031.00		NIST Webbook
ripol	1037.00		NIST Webbook
ripol	1046.00		NIST Webbook
ripol	1027.00		NIST Webbook
ripol	1037.00		NIST Webbook
sg	322.49	J/mol×K	NIST Webbook
sl	214.20	J/mol×K	NIST Webbook
sl	192.80	J/mol×K	NIST Webbook
ss	112.70	J/mol×K	NIST Webbook
tb	370.35	K	Isobaric (vapor-liquid) equilibria for binary systems of methanol + 1-(methoxymethoxy)-propane and 1-propanol + 1-(methoxymethoxy)-propane at 101.33 kPa
tb	369.90	K	Separation of the mixture (isopropyl alcohol + diisopropyl ether + n-propanol): Entrainer selection, interaction exploration and vapour-liquid equilibrium measurements
tb	370.35	K	Measurement and Modelization of VLE for Butyl Acetate with Methanol, Ethanol, 1-Propanol, and 1-Butanol. Experimental Data at 0.15 MPa
tb	370.33	K	Isobaric Vapor-Liquid Equilibrium of the Acetonitrile + 1-Propanol + Ionic Liquids at an Atmospheric Pressure
tb	370.30	K	Isobaric Phase Equilibria of Diethyl Carbonate with Five Alcohols at 101.3 kPa
tb	370.31	K	Measurement and Correlation of Excess Molar Enthalpies for Ethylene Glycol + Alkanol Systems at the Temperatures (298.15, 308.15, and 323.15) K
tb	370.19	K	Vapor-Liquid Equilibrium Behavior of Tolan in Alcohol

tb	370.21	K	Vapor-Liquid Equilibrium Behaviors of Coumarin and Vanillin in Ethanol, 1-Propanol, and 2-Propanol
tb	370.24	K	Vapor-Liquid Equilibria for the Ternary System Acetonitrile + 1-Propanol + Dimethyl Sulfoxide and the Corresponding Binary Systems at 101.3 kPa
tb	370.05	K	P, rho, T Measurements and Isobaric Vapor-Liquid-Equilibria of the 1,3,3-Trimethyl-2-oxabicycle[2,2,2]octane + Propan-1-ol Mixture: Cubic and Statistical Associating Fluid Theory-Based Equation of State Analysis
tb	370.40	K	Evaluation of the 2-Methoxyethanol as Entrainer in Ethanol Water and 1-Propanol Water Mixtures
tb	370.21	K	Vapor-Liquid Equilibrium Behaviors of 3-Ethoxy-4-hydroxybenzaldehyde in Alcohol
tb	370.21	K	Vapor Liquid Equilibrium Behaviors of 5-Methyl-2-(1-methylethyl)phenol in Alcohol
tb	370.30 ± 0.20	K	NIST Webbook
tb	370.26 ± 0.20	K	NIST Webbook
tb	370.26 ± 0.20	K	NIST Webbook
tb	370.20 ± 0.40	K	NIST Webbook
tb	370.26 ± 0.20	K	NIST Webbook
tb	370.26 ± 0.20	K	NIST Webbook
tb	370.28 ± 0.08	K	NIST Webbook
tb	370.30 ± 0.20	K	NIST Webbook
tb	370.40 ± 0.20	K	NIST Webbook
tb	370.30 ± 0.30	K	NIST Webbook
tb	370.10 ± 0.25	K	NIST Webbook
tb	370.60	K	NIST Webbook
tb	370.50 ± 0.20	K	NIST Webbook
tb	370.30	K	NIST Webbook
tb	370.40 ± 0.40	K	NIST Webbook
tb	370.40 ± 0.30	K	NIST Webbook
tb	370.40 ± 0.30	K	NIST Webbook
tb	370.15 ± 0.50	K	NIST Webbook
tb	370.60 ± 0.30	K	NIST Webbook
tb	370.30 ± 0.50	K	NIST Webbook

tb	370.32 ± 0.06	K	NIST Webbook
tb	370.30 ± 0.50	K	NIST Webbook
tb	370.40 ± 0.30	K	NIST Webbook
tb	370.29 ± 0.30	K	NIST Webbook
tb	370.30 ± 0.50	K	NIST Webbook
tb	370.30 ± 0.50	K	NIST Webbook
tb	370.35 ± 0.20	K	NIST Webbook
tb	370.23 ± 0.12	K	NIST Webbook
tb	370.50 ± 0.30	K	NIST Webbook
tb	370.35 ± 0.30	K	NIST Webbook
tb	370.35	K	NIST Webbook
tb	370.35 ± 0.30	K	NIST Webbook
tb	370.25 ± 0.20	K	NIST Webbook
tb	370.44 ± 0.15	K	NIST Webbook
tb	370.30 ± 0.05	K	NIST Webbook
tb	370.40 ± 0.20	K	NIST Webbook
tb	370.23 ± 0.20	K	NIST Webbook
tb	370.00 ± 1.50	K	NIST Webbook
tb	370.00 ± 1.00	K	NIST Webbook
tb	370.35 ± 0.30	K	NIST Webbook
tb	370.38 ± 0.10	K	NIST Webbook
tb	370.60 ± 0.30	K	NIST Webbook
tb	370.34 ± 0.30	K	NIST Webbook
tb	370.40 ± 0.30	K	NIST Webbook
tb	370.70 ± 1.00	K	NIST Webbook
tb	366.70 ± 3.00	K	NIST Webbook
tb	370.60 ± 0.30	K	NIST Webbook
tb	370.00 ± 1.00	K	NIST Webbook
tb	370.40 ± 0.50	K	NIST Webbook
tb	370.00 ± 0.50	K	NIST Webbook
tb	370.59 ± 0.33	K	NIST Webbook
tb	370.40 ± 0.20	K	NIST Webbook
tb	370.20 ± 0.30	K	NIST Webbook
tb	370.30 ± 0.20	K	NIST Webbook
tb	368.70 ± 1.00	K	NIST Webbook
tb	370.40 ± 0.30	K	NIST Webbook
tb	368.60 ± 1.00	K	NIST Webbook
tb	370.00 ± 1.00	K	NIST Webbook
tb	370.40 ± 0.50	K	NIST Webbook
tb	370.60 ± 0.50	K	NIST Webbook
tb	370.34 ± 0.20	K	NIST Webbook
tb	367.60 ± 1.00	K	NIST Webbook
tb	370.00 ± 1.00	K	NIST Webbook
tb	369.60 ± 0.50	K	NIST Webbook

tb	355.40 ± 0.30	K	NIST Webbook
tb	370.10 ± 0.50	K	NIST Webbook
tb	370.40 ± 0.50	K	NIST Webbook
tb	370.40 ± 0.40	K	NIST Webbook
tb	370.40 ± 0.30	K	NIST Webbook
tb	370.40 ± 0.30	K	NIST Webbook
tb	371.20 ± 0.50	K	NIST Webbook
tb	370.70 ± 0.60	K	NIST Webbook
tb	370.40 ± 0.50	K	NIST Webbook
tb	370.50 ± 0.50	K	NIST Webbook
tb	370.35	K	Isobaric vapor-liquid equilibria for extractive distillation of 1-propanol + water mixture using thiocyanate-based ionic liquids
tb	369.60 ± 1.00	K	NIST Webbook
tb	369.80 ± 0.40	K	NIST Webbook
tb	371.00 ± 0.30	K	NIST Webbook
tb	370.30 ± 0.20	K	NIST Webbook
tb	369.60 ± 1.00	K	NIST Webbook
tb	370.35 ± 0.03	K	NIST Webbook
tb	370.95 ± 0.10	K	NIST Webbook
tb	370.30 ± 0.30	K	NIST Webbook
tb	370.30 ± 0.05	K	NIST Webbook
tb	370.40 ± 0.10	K	NIST Webbook
tb	370.34 ± 0.07	K	NIST Webbook
tb	370.40 ± 0.50	K	NIST Webbook
tb	370.60 ± 0.50	K	NIST Webbook
tb	370.50 ± 0.50	K	NIST Webbook
tb	370.40 ± 0.30	K	NIST Webbook
tb	370.50 ± 0.50	K	NIST Webbook
tb	370.70 ± 0.50	K	NIST Webbook
tb	370.40 ± 0.50	K	NIST Webbook
tb	370.40 ± 0.50	K	NIST Webbook
tb	370.40 ± 0.50	K	NIST Webbook
tb	370.70 ± 1.00	K	NIST Webbook
tb	370.31 ± 0.30	K	NIST Webbook
tb	370.40 ± 0.50	K	NIST Webbook
tb	370.70 ± 1.00	K	NIST Webbook
tb	370.31	K	Isobaric vapor liquid equilibrium for the binary systems of 1-propanol + 1-(methoxymethoxy)-butane and 1-butanol + 1-(methoxymethoxy)-butane at 101.3 kPa
tb	371.00 ± 0.30	K	NIST Webbook

tb	370.50 ± 0.50	K	NIST Webbook
tb	370.20 ± 0.50	K	NIST Webbook
tb	370.34 ± 0.10	K	NIST Webbook
tb	370.60 ± 0.50	K	NIST Webbook
tb	370.90 ± 0.20	K	NIST Webbook
tb	370.13 ± 0.20	K	NIST Webbook
tb	370.30 ± 0.20	K	NIST Webbook
tb	371.25 ± 0.20	K	NIST Webbook
tb	370.34 ± 0.05	K	NIST Webbook
tb	370.33 ± 0.05	K	NIST Webbook
tb	370.35 ± 0.05	K	NIST Webbook
tb	370.40 ± 0.50	K	NIST Webbook
tb	370.20 ± 0.50	K	NIST Webbook
tb	369.80 ± 0.30	K	NIST Webbook
tb	369.90 ± 0.50	K	NIST Webbook
tb	370.40 ± 0.30	K	NIST Webbook
tb	370.37 ± 0.10	K	NIST Webbook
tb	369.60 ± 1.00	K	NIST Webbook
tb	370.30 ± 0.50	K	NIST Webbook
tb	370.60 ± 0.30	K	NIST Webbook
tb	370.41 ± 0.20	K	NIST Webbook
tb	370.40 ± 0.30	K	NIST Webbook
tb	370.40 ± 0.30	K	NIST Webbook
tb	370.50 ± 0.50	K	NIST Webbook
tb	370.00 ± 1.00	K	NIST Webbook
tb	368.70 ± 0.50	K	NIST Webbook
tb	370.34 ± 0.20	K	NIST Webbook
tb	369.80 ± 0.50	K	NIST Webbook
tb	370.40 ± 0.50	K	NIST Webbook
tb	370.40 ± 0.10	K	NIST Webbook
tb	369.20 ± 0.10	K	NIST Webbook
tb	369.30 ± 0.50	K	NIST Webbook
tb	369.90 ± 2.00	K	NIST Webbook
tb	369.75 ± 0.50	K	NIST Webbook
tb	371.00 ± 1.00	K	NIST Webbook
tb	371.00 ± 0.50	K	NIST Webbook
tb	370.60 ± 0.50	K	NIST Webbook
tb	370.56 ± 0.30	K	NIST Webbook
tb	363.00 ± 4.00	K	NIST Webbook
tb	370.45	K	(Vapour + liquid) equilibria in the ternary system (acetonitrile + n-propanol + ethylene glycol) and corresponding binary systems at 101.3 kPa

tb	369.75	K	Isobaric (vapour + liquid + liquid) equilibrium data for (di-n-propyl ether + n-propyl alcohol + water) and (diisopropyl ether + isopropyl alcohol + water) systems at 100 kPa
tb	370.20	K	Heterogeneous azeotropic distillation for the separation of n-propanol + water mixture using n-propyl acetate as entrainer
tb	370.44	K	Isobaric VLE at 0.6 MPa for binary systems isobutyl acetate + ethanol, + 1-propanol or + 2-propanol
tb	370.40	K	Excess molar volumes of ternary mixtures of 1,3-dichlorobenzene and methyl ethyl ketone with 1-alkanols at 303.15K
tb	355.30 ± 0.50	K	NIST Webbook
tb	370.30	K	KDB
tb	370.40 ± 0.30	K	NIST Webbook
tc	536.80	K	KDB
tc	537.00	K	Measurement and correlation of critical properties for binary mixtures and ternary mixtures containing gasoline additives
tc	537.00	K	Measurement of critical temperatures and critical pressures for binary mixtures of methyl tert-butyl ether (MTBE) + alcohol and MTBE + alkane
tf	146.95	K	NIST Webbook
tf	146.70	K	Phase equilibria of didecyldimethylammonium nitrate ionic liquid with water and organic solvents
tf	146.62	K	Aqueous Solubility Prediction Method
tf	146.70 ± 0.50	K	NIST Webbook
tf	147.00 ± 3.00	K	NIST Webbook
tf	147.00	K	KDB
tt	148.75	K	KDB
tt	148.75 ± 0.02	K	NIST Webbook
tt	148.75 ± 0.02	K	NIST Webbook
tt	147.00 ± 0.30	K	NIST Webbook
vc	0.218	m3/kmol	KDB
vc	0.216 ± 0.001	m3/kmol	NIST Webbook

vc	0.218	m3/kmol	NIST Webbook
zc	0.2524720		KDB
zra	0.25		KDB

Temperature Dependent Properties

Property code	Value	Unit	Temperature [K]	Source
cpg	107.28 ± 0.96	J/mol×K	375.45	NIST Webbook
cpg	108.67 ± 0.96	J/mol×K	383.05	NIST Webbook
cpg	109.42 ± 0.96	J/mol×K	387.15	NIST Webbook
cpg	106.44 ± 0.21	J/mol×K	391.20	NIST Webbook
cpg	111.21 ± 0.96	J/mol×K	396.95	NIST Webbook
cpg	113.59 ± 0.96	J/mol×K	409.95	NIST Webbook
cpg	110.42 ± 0.22	J/mol×K	411.20	NIST Webbook
cpg	115.56 ± 0.96	J/mol×K	420.75	NIST Webbook
cpg	115.97 ± 0.96	J/mol×K	422.95	NIST Webbook
cpg	114.35 ± 0.23	J/mol×K	431.20	NIST Webbook
cpg	118.71 ± 0.96	J/mol×K	437.95	NIST Webbook
cpg	118.62 ± 0.24	J/mol×K	451.20	NIST Webbook
cpg	122.94 ± 0.96	J/mol×K	461.05	NIST Webbook
cpg	125.55 ± 0.96	J/mol×K	475.35	NIST Webbook
cpg	130.97 ± 0.96	J/mol×K	504.95	NIST Webbook
cpg	144.49 ± 0.96	J/mol×K	578.85	NIST Webbook
cpg	141.05 ± 0.96	J/mol×K	560.05	NIST Webbook
cpg	135.98 ± 0.96	J/mol×K	532.35	NIST Webbook
cpg	132.23 ± 0.96	J/mol×K	511.85	NIST Webbook
cpg	148.95 ± 0.96	J/mol×K	603.25	NIST Webbook
cpg	102.26 ± 0.20	J/mol×K	371.20	NIST Webbook
cpl	144.06	J/mol×K	298.15	NIST Webbook
cpl	164.80	J/mol×K	301.20	NIST Webbook
cpl	145.60	J/mol×K	298.10	NIST Webbook
cpl	140.21	J/mol×K	303.00	NIST Webbook
cpl	155.60	J/mol×K	320.00	NIST Webbook
cpl	146.10	J/mol×K	298.00	NIST Webbook
cpl	143.80	J/mol×K	298.15	NIST Webbook
cpl	158.60	J/mol×K	313.20	NIST Webbook
cpl	192.90	J/mol×K	298.10	NIST Webbook
cpl	143.87	J/mol×K	298.15	NIST Webbook
cpl	143.78	J/mol×K	298.15	NIST Webbook
cpl	149.00	J/mol×K	298.15	NIST Webbook
cpl	143.77	J/mol×K	298.15	NIST Webbook

cpl	147.90	J/mol×K	303.40	NIST Webbook
cpl	146.34	J/mol×K	298.22	NIST Webbook
cpl	146.88	J/mol×K	298.15	NIST Webbook
cpl	138.40	J/mol×K	288.15	NIST Webbook
cpl	144.44	J/mol×K	298.15	NIST Webbook
cpl	143.96	J/mol×K	298.15	NIST Webbook
cpl	144.60	J/mol×K	298.00	NIST Webbook
cpl	167.80	J/mol×K	333.15	Excess Molar Enthalpies, Molar Heat Capacities, Densities, Viscosities, and Refractive Indices of Dimethyl Sulfoxide + 1-Propanol at (288.15, 298.15, and 308.15) K and at Normal Pressure
cpl	164.50	J/mol×K	328.15	Excess Molar Enthalpies, Molar Heat Capacities, Densities, Viscosities, and Refractive Indices of Dimethyl Sulfoxide + 1-Propanol at (288.15, 298.15, and 308.15) K and at Normal Pressure
cpl	160.10	J/mol×K	323.15	Excess Molar Enthalpies, Molar Heat Capacities, Densities, Viscosities, and Refractive Indices of Dimethyl Sulfoxide + 1-Propanol at (288.15, 298.15, and 308.15) K and at Normal Pressure

cpl	157.10	J/mol×K	318.15	Excess Molar Enthalpies, Molar Heat Capacities, Densities, Viscosities, and Refractive Indices of Dimethyl Sulfoxide + 1-Propanol at (288.15, 298.15, and 308.15) K and at Normal Pressure	
cpl	153.70	J/mol×K	313.15	Excess Molar Enthalpies, Molar Heat Capacities, Densities, Viscosities, and Refractive Indices of Dimethyl Sulfoxide + 1-Propanol at (288.15, 298.15, and 308.15) K and at Normal Pressure	
cpl	151.00	J/mol×K	308.15	Excess Molar Enthalpies, Molar Heat Capacities, Densities, Viscosities, and Refractive Indices of Dimethyl Sulfoxide + 1-Propanol at (288.15, 298.15, and 308.15) K and at Normal Pressure	
cpl	148.20	J/mol×K	303.15	Excess Molar Enthalpies, Molar Heat Capacities, Densities, Viscosities, and Refractive Indices of Dimethyl Sulfoxide + 1-Propanol at (288.15, 298.15, and 308.15) K and at Normal Pressure	

cpl	144.90	J/mol×K	298.15	Excess Molar Enthalpies, Molar Heat Capacities, Densities, Viscosities, and Refractive Indices of Dimethyl Sulfoxide + 1-Propanol at (288.15, 298.15, and 308.15) K and at Normal Pressure
cpl	142.30	J/mol×K	293.15	Excess Molar Enthalpies, Molar Heat Capacities, Densities, Viscosities, and Refractive Indices of Dimethyl Sulfoxide + 1-Propanol at (288.15, 298.15, and 308.15) K and at Normal Pressure
cpl	139.00	J/mol×K	288.15	Excess Molar Enthalpies, Molar Heat Capacities, Densities, Viscosities, and Refractive Indices of Dimethyl Sulfoxide + 1-Propanol at (288.15, 298.15, and 308.15) K and at Normal Pressure
cpl	133.50	J/mol×K	275.40	NIST Webbook
cpl	133.50	J/mol×K	275.00	NIST Webbook
cpl	131.30	J/mol×K	274.60	NIST Webbook
cpl	136.00	J/mol×K	270.00	NIST Webbook
cpl	144.80	J/mol×K	298.00	NIST Webbook
cpl	141.80	J/mol×K	293.15	NIST Webbook
cps	106.30	J/mol×K	150.00	NIST Webbook
dvisc	0.0013907	Pa×s	313.15	Densities and Viscosities of Binary Mixture of the Ionic Liquid Bis(2-hydroxyethyl)ammonium Propionate with Methanol, Ethanol, and 1-Propanol at T = (293.15, 303.15, 313.15, and 323.15) K and at P = 0.1 MPa

dvisc	0.0021412	Pa×s	293.15	Densities and Viscosities of Binary Mixture of the Ionic Liquid Bis(2-hydroxyethyl)ammonium Propionate with Methanol, Ethanol, and 1-Propanol at T = (293.15, 303.15, 313.15, and 323.15) K and at P = 0.1 MPa
dvisc	0.0017880	Pa×s	303.15	Densities and Viscosities of Binary Mixture of the Ionic Liquid Bis(2-hydroxyethyl)ammonium Propionate with Methanol, Ethanol, and 1-Propanol at T = (293.15, 303.15, 313.15, and 323.15) K and at P = 0.1 MPa
dvisc	0.0005150	Pa×s	363.15	Density and Viscosity of Binary Mixtures of Diethyl Carbonate with Alcohols at (293.15 to 363.15) K and Predictive Results by UNIFAC-VISCO Group Contribution Method
dvisc	0.0010977	Pa×s	323.15	Densities and Viscosities of Binary Mixture of the Ionic Liquid Bis(2-hydroxyethyl)ammonium Propionate with Methanol, Ethanol, and 1-Propanol at T = (293.15, 303.15, 313.15, and 323.15) K and at P = 0.1 MPa
dvisc	0.0025070	Pa×s	288.15	Densities and Viscosities of Four Binary Diethyl Carbonate + 1-Alcohol Systems from (288.15 to 313.15) K

dvisc	0.0022020	Pa×s	293.15	Densities and Viscosities of Four Binary Diethyl Carbonate + 1-Alcohol Systems from (288.15 to 313.15) K
dvisc	0.0019730	Pa×s	298.15	Densities and Viscosities of Four Binary Diethyl Carbonate + 1-Alcohol Systems from (288.15 to 313.15) K
dvisc	0.0017330	Pa×s	303.15	Densities and Viscosities of Four Binary Diethyl Carbonate + 1-Alcohol Systems from (288.15 to 313.15) K
dvisc	0.0013810	Pa×s	313.15	Densities and Viscosities of Four Binary Diethyl Carbonate + 1-Alcohol Systems from (288.15 to 313.15) K
dvisc	0.0019414	Pa×s	298.15	Densities and Viscosities for Binary Mixtures of the Ionic Liquid N-Ethyl Piperazinium Propionate with n-Alcohols at Several Temperatures
dvisc	0.0017282	Pa×s	303.15	Densities and Viscosities for Binary Mixtures of the Ionic Liquid N-Ethyl Piperazinium Propionate with n-Alcohols at Several Temperatures

dvisc	0.0015448	Pa×s	308.15	Densities and Viscosities for Binary Mixtures of the Ionic Liquid N-Ethyl Piperazinium Propionate with n-Alcohols at Several Temperatures
dvisc	0.0013847	Pa×s	313.15	Densities and Viscosities for Binary Mixtures of the Ionic Liquid N-Ethyl Piperazinium Propionate with n-Alcohols at Several Temperatures
dvisc	0.0021920	Pa×s	293.15	Densities and Viscosities of Naphthalen-1-ol, Naphthalen-2-ol, and 1-Aminonaphthalene in the Solvents of Ethanol, Propan-1-ol, and Butan-1-ol
dvisc	0.0017260	Pa×s	303.15	Densities and Viscosities of Naphthalen-1-ol, Naphthalen-2-ol, and 1-Aminonaphthalene in the Solvents of Ethanol, Propan-1-ol, and Butan-1-ol
dvisc	0.0013730	Pa×s	313.15	Densities and Viscosities of Naphthalen-1-ol, Naphthalen-2-ol, and 1-Aminonaphthalene in the Solvents of Ethanol, Propan-1-ol, and Butan-1-ol
dvisc	0.0011020	Pa×s	323.15	Densities and Viscosities of Naphthalen-1-ol, Naphthalen-2-ol, and 1-Aminonaphthalene in the Solvents of Ethanol, Propan-1-ol, and Butan-1-ol

dvisc	0.0006150	Pa×s	353.15	Density and Viscosity of Binary Mixtures of Diethyl Carbonate with Alcohols at (293.15 to 363.15) K and Predictive Results by UNIFAC-VISCO Group Contribution Method
dvisc	0.0019550	Pa×s	298.15	Densities and Viscosities of Binary Mixtures of 2,2,4-Trimethylpentane + 1-Propanol, + 1-Pentanol, + 1-Hexanol, and + 1-Heptanol from (298.15 to 323.15) K
dvisc	0.0017370	Pa×s	303.15	Densities and Viscosities of Binary Mixtures of 2,2,4-Trimethylpentane + 1-Propanol, + 1-Pentanol, + 1-Hexanol, and + 1-Heptanol from (298.15 to 323.15) K
dvisc	0.0015480	Pa×s	308.15	Densities and Viscosities of Binary Mixtures of 2,2,4-Trimethylpentane + 1-Propanol, + 1-Pentanol, + 1-Hexanol, and + 1-Heptanol from (298.15 to 323.15) K
dvisc	0.0013840	Pa×s	313.15	Densities and Viscosities of Binary Mixtures of 2,2,4-Trimethylpentane + 1-Propanol, + 1-Pentanol, + 1-Hexanol, and + 1-Heptanol from (298.15 to 323.15) K

dvisc	0.0007420	Pa×s	343.15	Density and Viscosity of Binary Mixtures of Diethyl Carbonate with Alcohols at (293.15 to 363.15) K and Predictive Results by UNIFAC-VISCO Group Contribution Method
dvisc	0.0011170	Pa×s	323.15	Densities and Viscosities of Binary Mixtures of 2,2,4-Trimethylpentane + 1-Propanol, + 1-Pentanol, + 1-Hexanol, and + 1-Heptanol from (298.15 to 323.15) K
dvisc	0.0022050	Pa×s	293.15	Viscosities and Densities of Binary Mixtures of (N-Acetylmorpholine + Alkanols) from (293.15 to 323.15) K
dvisc	0.0017140	Pa×s	303.15	Viscosities and Densities of Binary Mixtures of (N-Acetylmorpholine + Alkanols) from (293.15 to 323.15) K
dvisc	0.0013820	Pa×s	313.15	Viscosities and Densities of Binary Mixtures of (N-Acetylmorpholine + Alkanols) from (293.15 to 323.15) K
dvisc	0.0011050	Pa×s	323.15	Viscosities and Densities of Binary Mixtures of (N-Acetylmorpholine + Alkanols) from (293.15 to 323.15) K

dvisc	0.0022720	Pa×s	293.15	Densities and Kinematic Viscosities of Ten Binary 1-Alkanol Liquid Systems at Temperatures of (293.15 and 298.15) K
dvisc	0.0019660	Pa×s	298.15	Densities and Kinematic Viscosities of Ten Binary 1-Alkanol Liquid Systems at Temperatures of (293.15 and 298.15) K
dvisc	0.0019400	Pa×s	298.15	Temperature and Composition Dependence of the Density and Viscosity of Binary Mixtures of {1-Butyl-3-methylimidazolium Thiocyanate + 1-Alcohols}
dvisc	0.0015400	Pa×s	308.15	Temperature and Composition Dependence of the Density and Viscosity of Binary Mixtures of {1-Butyl-3-methylimidazolium Thiocyanate + 1-Alcohols}
dvisc	0.0012400	Pa×s	318.15	Temperature and Composition Dependence of the Density and Viscosity of Binary Mixtures of {1-Butyl-3-methylimidazolium Thiocyanate + 1-Alcohols}
dvisc	0.0010200	Pa×s	328.15	Temperature and Composition Dependence of the Density and Viscosity of Binary Mixtures of {1-Butyl-3-methylimidazolium Thiocyanate + 1-Alcohols}

dvisc	0.0008420	Pa×s	338.15	Temperature and Composition Dependence of the Density and Viscosity of Binary Mixtures of {1-Butyl-3-methylimidazolium Thiocyanate + 1-Alcohols}
dvisc	0.0007100	Pa×s	348.15	Temperature and Composition Dependence of the Density and Viscosity of Binary Mixtures of {1-Butyl-3-methylimidazolium Thiocyanate + 1-Alcohols}
dvisc	0.0019550	Pa×s	298.15	Densities and Viscosities of Binary Mixtures of Cyclopropanecarboxylic Acid with Methanol, Ethanol, Propan-1-ol, and Butan-1-ol at Different Temperatures
dvisc	0.0015380	Pa×s	308.15	Densities and Viscosities of Binary Mixtures of Cyclopropanecarboxylic Acid with Methanol, Ethanol, Propan-1-ol, and Butan-1-ol at Different Temperatures
dvisc	0.0012240	Pa×s	318.15	Densities and Viscosities of Binary Mixtures of Cyclopropanecarboxylic Acid with Methanol, Ethanol, Propan-1-ol, and Butan-1-ol at Different Temperatures

dvisc	0.0009860	Pa×s	328.15	Densities and Viscosities of Binary Mixtures of Cyclopropanecarboxylic Acid with Methanol, Ethanol, Propan-1-ol, and Butan-1-ol at Different Temperatures
dvisc	0.0008010	Pa×s	338.15	Densities and Viscosities of Binary Mixtures of Cyclopropanecarboxylic Acid with Methanol, Ethanol, Propan-1-ol, and Butan-1-ol at Different Temperatures
dvisc	0.0024202	Pa×s	288.15	Densities and Viscosities of (1-Propanol + 1,2-Dichloroethane), (1-Propanol + Benzaldehyde), (Benzaldehyde + 1,2-Dichloroethane), and (1-Propanol + 1,2-Dichloroethane + Benzaldehyde) Mixtures from T = 288.15 K to 313.15 K
dvisc	0.0012410	Pa×s	318.15	Densities and Viscosities of Binary Mixtures of 2,2,4-Trimethylpentane + 1-Propanol, + 1-Pentanol, + 1-Hexanol, and + 1-Heptanol from (298.15 to 323.15) K
dvisc	0.0018939	Pa×s	298.15	Densities and Viscosities of (1-Propanol + 1,2-Dichloroethane), (1-Propanol + Benzaldehyde), (Benzaldehyde + 1,2-Dichloroethane), and (1-Propanol + 1,2-Dichloroethane + Benzaldehyde) Mixtures from T = 288.15 K to 313.15 K

dvisc	0.0016841	Pa×s	303.15	Densities and Viscosities of (1-Propanol + 1,2-Dichloroethane), (1-Propanol + Benzaldehyde), (Benzaldehyde + 1,2-Dichloroethane), and (1-Propanol + 1,2-Dichloroethane + Benzaldehyde) Mixtures from T = 288.15 K to 313.15 K
dvisc	0.0015037	Pa×s	308.15	Densities and Viscosities of (1-Propanol + 1,2-Dichloroethane), (1-Propanol + Benzaldehyde), (Benzaldehyde + 1,2-Dichloroethane), and (1-Propanol + 1,2-Dichloroethane + Benzaldehyde) Mixtures from T = 288.15 K to 313.15 K
dvisc	0.0013516	Pa×s	313.15	Densities and Viscosities of (1-Propanol + 1,2-Dichloroethane), (1-Propanol + Benzaldehyde), (Benzaldehyde + 1,2-Dichloroethane), and (1-Propanol + 1,2-Dichloroethane + Benzaldehyde) Mixtures from T = 288.15 K to 313.15 K
dvisc	0.0009070	Pa×s	333.15	Density and Viscosity of Binary Mixtures of Diethyl Carbonate with Alcohols at (293.15 to 363.15) K and Predictive Results by UNIFAC-VISCO Group Contribution Method

dvisc	0.0013810	Pa×s	313.15	Density and Viscosity of Binary Mixtures of Diethyl Carbonate with Alcohols at (293.15 to 363.15) K and Predictive Results by UNIFAC-VISCO Group Contribution Method
dvisc	0.0017450	Pa×s	303.15	Density and Viscosity of Binary Mixtures of Diethyl Carbonate with Alcohols at (293.15 to 363.15) K and Predictive Results by UNIFAC-VISCO Group Contribution Method
dvisc	0.0019810	Pa×s	298.15	Density and Viscosity of Binary Mixtures of Diethyl Carbonate with Alcohols at (293.15 to 363.15) K and Predictive Results by UNIFAC-VISCO Group Contribution Method
dvisc	0.0022380	Pa×s	293.15	Density and Viscosity of Binary Mixtures of Diethyl Carbonate with Alcohols at (293.15 to 363.15) K and Predictive Results by UNIFAC-VISCO Group Contribution Method

dvisc	0.0013630	Pa×s	313.15	Densities, Viscosities, and Ultrasonic Velocity Studies of Binary Mixtures of Chloroform with Propan-1-ol and Butan-1-ol at (303.15 and 313.15) K
dvisc	0.0017190	Pa×s	303.15	Densities, Viscosities, and Ultrasonic Velocity Studies of Binary Mixtures of Chloroform with Propan-1-ol and Butan-1-ol at (303.15 and 313.15) K
dvisc	0.0021940	Pa×s	293.15	Viscosities and densities for binary mixtures of N-methylpiperazine with methanol, ethanol, n-propanol, iso-propanol, n-butanol and iso-butanol at 293.15, 298.15 and 303.15K
dvisc	0.0013780	Pa×s	313.15	Density and Viscosity of Binary Mixtures of Ethyl-2-methylbutyrate and Ethyl Hexanoate with Methanol, Ethanol, and 1-Propanol at (293.15, 303.15, and 313.15) K
dvisc	0.0017130	Pa×s	303.15	Density and Viscosity of Binary Mixtures of Ethyl-2-methylbutyrate and Ethyl Hexanoate with Methanol, Ethanol, and 1-Propanol at (293.15, 303.15, and 313.15) K

dvisc	0.0021880	Pa×s	293.15	Density and Viscosity of Binary Mixtures of Ethyl-2-methylbutyrate and Ethyl Hexanoate with Methanol, Ethanol, and 1-Propanol at (293.15, 303.15, and 313.15) K
dvisc	0.0015650	Pa×s	308.15	Viscosities, Densities, and Ultrasonic Velocities of Binary Mixtures of Ethylbenzene with Ethanol, 1-Propanol, and 1-Butanol at (298.15 and 308.15)K
dvisc	0.0020420	Pa×s	298.15	Viscosities, Densities, and Ultrasonic Velocities of Binary Mixtures of Ethylbenzene with Ethanol, 1-Propanol, and 1-Butanol at (298.15 and 308.15)K
dvisc	0.0015370	Pa×s	308.15	Excess Molar Volumes and Viscosity Deviations of Binary Mixtures of 2,4,6-Trimethyl-1,3,5-trioxane + Ethanol, 1-Propanol, and 1-Butanol at (298.15, 303.15, and 308.15) K
dvisc	0.0017250	Pa×s	303.15	Excess Molar Volumes and Viscosity Deviations of Binary Mixtures of 2,4,6-Trimethyl-1,3,5-trioxane + Ethanol, 1-Propanol, and 1-Butanol at (298.15, 303.15, and 308.15) K

dvisc	0.0019430	Pa×s	298.15	Excess Molar Volumes and Viscosity Deviations of Binary Mixtures of 2,4,6-Trimethyl-1,3,5-trioxane + Ethanol, 1-Propanol, and 1-Butanol at (298.15, 303.15, and 308.15) K	
dvisc	0.0005985	Pa×s	333.15	Thermodynamic Study of Binary Mixture of x1[C6mim][BF4] + x21-propanol: Measurements and Molecular Modeling	
dvisc	0.0008796	Pa×s	323.15	Thermodynamic Study of Binary Mixture of x1[C6mim][BF4] + x21-propanol: Measurements and Molecular Modeling	
dvisc	0.0011883	Pa×s	313.15	Thermodynamic Study of Binary Mixture of x1[C6mim][BF4] + x21-propanol: Measurements and Molecular Modeling	
dvisc	0.0016145	Pa×s	303.15	Thermodynamic Study of Binary Mixture of x1[C6mim][BF4] + x21-propanol: Measurements and Molecular Modeling	
dvisc	0.0021372	Pa×s	293.15	Densities and Viscosities of (1-Propanol + 1,2-Dichloroethane), (1-Propanol + Benzaldehyde), (Benzaldehyde + 1,2-Dichloroethane), and (1-Propanol + 1,2-Dichloroethane + Benzaldehyde) Mixtures from T = 288.15 K to 313.15 K	

dvisc	0.0022040	Pa×s	293.15	Thermodynamic Study of Binary Mixture of x1[C6mim][BF4] + x21-propanol: Measurements and Molecular Modeling
dvisc	0.0013830	Pa×s	313.15	Properties of pure 1-methylimidazolium acetate ionic liquid and its binary mixtures with alcohols
dvisc	0.0015480	Pa×s	308.15	Properties of pure 1-methylimidazolium acetate ionic liquid and its binary mixtures with alcohols
dvisc	0.0017370	Pa×s	303.15	Properties of pure 1-methylimidazolium acetate ionic liquid and its binary mixtures with alcohols
dvisc	0.0019760	Pa×s	298.15	Properties of pure 1-methylimidazolium acetate ionic liquid and its binary mixtures with alcohols
dvisc	0.0022060	Pa×s	293.15	Properties of pure 1-methylimidazolium acetate ionic liquid and its binary mixtures with alcohols
dvisc	0.0010819	Pa×s	333.15	Density and viscosity of 1-butyl-3-methylimidazolium nitrate with ethanol, 1-propanol, or 1-butanol at several temperatures
dvisc	0.0011523	Pa×s	328.15	Density and viscosity of 1-butyl-3-methylimidazolium nitrate with ethanol, 1-propanol, or 1-butanol at several temperatures

dvisc	0.0012560	Pa×s	323.15	Density and viscosity of 1-butyl-3-methylimidazolium nitrate with ethanol, 1-propanol, or 1-butanol at several temperatures
dvisc	0.0013833	Pa×s	318.15	Density and viscosity of 1-butyl-3-methylimidazolium nitrate with ethanol, 1-propanol, or 1-butanol at several temperatures
dvisc	0.0015237	Pa×s	313.15	Density and viscosity of 1-butyl-3-methylimidazolium nitrate with ethanol, 1-propanol, or 1-butanol at several temperatures
dvisc	0.0016709	Pa×s	308.15	Density and viscosity of 1-butyl-3-methylimidazolium nitrate with ethanol, 1-propanol, or 1-butanol at several temperatures
dvisc	0.0018742	Pa×s	303.15	Density and viscosity of 1-butyl-3-methylimidazolium nitrate with ethanol, 1-propanol, or 1-butanol at several temperatures
dvisc	0.0021178	Pa×s	298.15	Density and viscosity of 1-butyl-3-methylimidazolium nitrate with ethanol, 1-propanol, or 1-butanol at several temperatures
dvisc	0.0024104	Pa×s	293.15	Density and viscosity of 1-butyl-3-methylimidazolium nitrate with ethanol, 1-propanol, or 1-butanol at several temperatures

dvisc	0.0034744	Pa×s	283.15	Density and viscosity of 1-butyl-3-methylimidazolium nitrate with ethanol, 1-propanol, or 1-butanol at several temperatures
dvisc	0.0011004	Pa×s	323.15	Densities and viscosities of binary mixtures of {dimethylsulfoxide + aliphatic lower alkanols (C1 C3)} at temperatures from T = 303.15 K to T = 323.15 K
dvisc	0.0012340	Pa×s	318.15	Densities and viscosities of binary mixtures of {dimethylsulfoxide + aliphatic lower alkanols (C1 C3)} at temperatures from T = 303.15 K to T = 323.15 K
dvisc	0.0013672	Pa×s	313.15	Densities and viscosities of binary mixtures of {dimethylsulfoxide + aliphatic lower alkanols (C1 C3)} at temperatures from T = 303.15 K to T = 323.15 K
dvisc	0.0015344	Pa×s	308.15	Densities and viscosities of binary mixtures of {dimethylsulfoxide + aliphatic lower alkanols (C1 C3)} at temperatures from T = 303.15 K to T = 323.15 K
dvisc	0.0017325	Pa×s	303.15	Densities and viscosities of binary mixtures of {dimethylsulfoxide + aliphatic lower alkanols (C1 C3)} at temperatures from T = 303.15 K to T = 323.15 K

dvisc	0.0017130	Pa×s	303.15	Dynamic viscosities of binary mixtures of cycloalkanes with primary alcohols at T = (293.15, 298.15, and 303.15) K: New UNIFAC-VISCO interaction parameters
dvisc	0.0019150	Pa×s	298.15	Dynamic viscosities of binary mixtures of cycloalkanes with primary alcohols at T = (293.15, 298.15, and 303.15) K: New UNIFAC-VISCO interaction parameters
dvisc	0.0021590	Pa×s	293.15	Dynamic viscosities of binary mixtures of cycloalkanes with primary alcohols at T = (293.15, 298.15, and 303.15) K: New UNIFAC-VISCO interaction parameters
dvisc	0.0015330	Pa×s	308.15	Densities, viscosities, and ultrasonic velocity studies of binary mixtures of trichloromethane with methanol, ethanol, propan-1-ol, and butan-1-ol at T=(298.15 and 308.15) K
dvisc	0.0019500	Pa×s	298.15	Densities, viscosities, and ultrasonic velocity studies of binary mixtures of trichloromethane with methanol, ethanol, propan-1-ol, and butan-1-ol at T=(298.15 and 308.15) K

dvisc	0.0017130	Pa×s	303.15	Viscosity, density, and speed of sound of methylcyclopentane with primary and secondary alcohols at T = (293.15, 298.15, and 303.15) K
dvisc	0.0019150	Pa×s	298.15	Viscosity, density, and speed of sound of methylcyclopentane with primary and secondary alcohols at T = (293.15, 298.15, and 303.15) K
dvisc	0.0021590	Pa×s	293.15	Viscosity, density, and speed of sound of methylcyclopentane with primary and secondary alcohols at T = (293.15, 298.15, and 303.15) K
dvisc	0.0013400	Pa×s	313.00	Ultrasonic velocity, viscosity and excess properties of binary mixture of tetrahydrofuran with 1-propanol and 2-propanol
dvisc	0.0016100	Pa×s	303.00	Ultrasonic velocity, viscosity and excess properties of binary mixture of tetrahydrofuran with 1-propanol and 2-propanol
dvisc	0.0021900	Pa×s	293.00	Ultrasonic velocity, viscosity and excess properties of binary mixture of tetrahydrofuran with 1-propanol and 2-propanol

dvisc	0.0017300	Pa×s	303.15	Viscosities and densities for binary mixtures of N-methylpiperazine with methanol, ethanol, n-propanol, iso-propanol, n-butanol and iso-butanol at 293.15, 298.15 and 303.15K
dvisc	0.0019450	Pa×s	298.15	Viscosities and densities for binary mixtures of N-methylpiperazine with methanol, ethanol, n-propanol, iso-propanol, n-butanol and iso-butanol at 293.15, 298.15 and 303.15K
dvisc	0.0027676	Pa×s	288.15	Density and viscosity of 1-butyl-3-methylimidazolium nitrate with ethanol, 1-propanol, or 1-butanol at several temperatures
dvisc	0.0011150	Pa×s	323.15	Density and Viscosity of Binary Mixtures of Diethyl Carbonate with Alcohols at (293.15 to 363.15) K and Predictive Results by UNIFAC-VISCO Group Contribution Method
dvisc	0.0008980	Pa×s	333.15	Densities and Viscosities of Naphthalen-1-ol, Naphthalen-2-ol, and 1-Aminonaphthalene in the Solvents of Ethanol, Propan-1-ol, and Butan-1-ol
hfust	5.37	kJ/mol	148.75	NIST Webbook
hfust	5.40	kJ/mol	148.70	NIST Webbook
hfust	5.37	kJ/mol	148.80	NIST Webbook
hfust	5.37	kJ/mol	148.80	NIST Webbook

hfust	5.19	kJ/mol	147.00	NIST Webbook
hvapt	44.00 ± 0.10	kJ/mol	343.00	NIST Webbook
hvapt	46.70	kJ/mol	318.00	NIST Webbook
hvapt	40.70	kJ/mol	471.00	NIST Webbook
hvapt	44.30	kJ/mol	358.00	NIST Webbook
hvapt	44.10	kJ/mol	364.00	NIST Webbook
hvapt	43.90 ± 0.10	kJ/mol	343.00	NIST Webbook
hvapt	42.30 ± 0.10	kJ/mol	360.00	NIST Webbook
hvapt	41.20 ± 0.10	kJ/mol	370.00	NIST Webbook
hvapt	44.90 ± 0.10	kJ/mol	333.00	NIST Webbook
hvapt	39.70 ± 0.10	kJ/mol	384.00	NIST Webbook
hvapt	45.50	kJ/mol	344.00	NIST Webbook
hvapt	43.20	kJ/mol	354.00	NIST Webbook
hvapt	44.99 ± 0.42	kJ/mol	333.13	NIST Webbook
hvapt	45.70 ± 0.10	kJ/mol	323.00	NIST Webbook
hvapt	46.40 ± 0.10	kJ/mol	313.00	NIST Webbook
hvapt	46.90	kJ/mol	331.00	NIST Webbook
hvapt	40.10	kJ/mol	441.50	NIST Webbook
hvapt	42.30	kJ/mol	388.00	NIST Webbook
hvapt	43.50	kJ/mol	366.00	NIST Webbook
hvapt	48.00	kJ/mol	214.00	NIST Webbook
hvapt	42.90	kJ/mol	368.50	NIST Webbook
hvapt	47.00	kJ/mol	336.50	NIST Webbook
hvapt	11.40	kJ/mol	528.00	NIST Webbook
hvapt	21.00	kJ/mol	498.00	NIST Webbook
hvapt	29.40	kJ/mol	453.00	NIST Webbook
hvapt	35.20	kJ/mol	423.00	NIST Webbook
hvapt	41.20	kJ/mol	371.00	NIST Webbook
hvapt	41.44	kJ/mol	370.30	NIST Webbook
hvapt	41.76	kJ/mol	370.40	KDB
hvapt	44.70	kJ/mol	355.00	NIST Webbook
hvapt	49.30	kJ/mol	324.00	NIST Webbook
hvapt	42.40 ± 0.10	kJ/mol	363.00	NIST Webbook
hvapt	43.20 ± 0.10	kJ/mol	353.00	NIST Webbook
hvapt	40.30 ± 0.10	kJ/mol	378.00	NIST Webbook
hvapt	36.50	kJ/mol	492.50	NIST Webbook
pvap	2733.00	kPa	498.15	Vapor-Liquid Equilibria for the n-Pentane + 1-Propanol and n-Pentane + 2-Methyl-1-propanol Systems near the Critical Region

pvap	10.00	kPa	319.65	P, rho, T Measurements and Isobaric Vapor-Liquid-Equilibria of the 1,3,3-Trimethyl-2-oxabicycle[2,2,2]octane + Propan-1-ol Mixture: Cubic and Statistical Associating Fluid Theory-Based Equation of State Analysis
pvap	15.00	kPa	327.25	P, rho, T Measurements and Isobaric Vapor-Liquid-Equilibria of the 1,3,3-Trimethyl-2-oxabicycle[2,2,2]octane + Propan-1-ol Mixture: Cubic and Statistical Associating Fluid Theory-Based Equation of State Analysis
pvap	20.00	kPa	332.95	P, rho, T Measurements and Isobaric Vapor-Liquid-Equilibria of the 1,3,3-Trimethyl-2-oxabicycle[2,2,2]octane + Propan-1-ol Mixture: Cubic and Statistical Associating Fluid Theory-Based Equation of State Analysis
pvap	25.00	kPa	337.45	P, rho, T Measurements and Isobaric Vapor-Liquid-Equilibria of the 1,3,3-Trimethyl-2-oxabicycle[2,2,2]octane + Propan-1-ol Mixture: Cubic and Statistical Associating Fluid Theory-Based Equation of State Analysis
pvap	30.00	kPa	341.25	P, rho, T Measurements and Isobaric Vapor-Liquid-Equilibria of the 1,3,3-Trimethyl-2-oxabicycle[2,2,2]octane + Propan-1-ol Mixture: Cubic and Statistical Associating Fluid Theory-Based Equation of State Analysis

pvap	33.33	kPa	343.55	P, rho, T Measurements and Isobaric Vapor-Liquid-Equilibria of the 1,3,3-Trimethyl-2-oxabicycle[2,2,2]octane + Propan-1-ol Mixture: Cubic and Statistical Associating Fluid Theory-Based Equation of State Analysis
pvap	35.00	kPa	344.65	P, rho, T Measurements and Isobaric Vapor-Liquid-Equilibria of the 1,3,3-Trimethyl-2-oxabicycle[2,2,2]octane + Propan-1-ol Mixture: Cubic and Statistical Associating Fluid Theory-Based Equation of State Analysis
pvap	40.00	kPa	347.65	P, rho, T Measurements and Isobaric Vapor-Liquid-Equilibria of the 1,3,3-Trimethyl-2-oxabicycle[2,2,2]octane + Propan-1-ol Mixture: Cubic and Statistical Associating Fluid Theory-Based Equation of State Analysis
pvap	45.00	kPa	350.35	P, rho, T Measurements and Isobaric Vapor-Liquid-Equilibria of the 1,3,3-Trimethyl-2-oxabicycle[2,2,2]octane + Propan-1-ol Mixture: Cubic and Statistical Associating Fluid Theory-Based Equation of State Analysis

pvap	50.00	kPa	352.75	P, rho, T Measurements and Isobaric Vapor-Liquid-Equilibria of the 1,3,3-Trimethyl-2-oxabicycle[2,2,2]octane + Propan-1-ol Mixture: Cubic and Statistical Associating Fluid Theory-Based Equation of State Analysis
pvap	55.00	kPa	355.05	P, rho, T Measurements and Isobaric Vapor-Liquid-Equilibria of the 1,3,3-Trimethyl-2-oxabicycle[2,2,2]octane + Propan-1-ol Mixture: Cubic and Statistical Associating Fluid Theory-Based Equation of State Analysis
pvap	60.00	kPa	357.05	P, rho, T Measurements and Isobaric Vapor-Liquid-Equilibria of the 1,3,3-Trimethyl-2-oxabicycle[2,2,2]octane + Propan-1-ol Mixture: Cubic and Statistical Associating Fluid Theory-Based Equation of State Analysis
pvap	65.00	kPa	359.05	P, rho, T Measurements and Isobaric Vapor-Liquid-Equilibria of the 1,3,3-Trimethyl-2-oxabicycle[2,2,2]octane + Propan-1-ol Mixture: Cubic and Statistical Associating Fluid Theory-Based Equation of State Analysis
pvap	66.66	kPa	359.65	P, rho, T Measurements and Isobaric Vapor-Liquid-Equilibria of the 1,3,3-Trimethyl-2-oxabicycle[2,2,2]octane + Propan-1-ol Mixture: Cubic and Statistical Associating Fluid Theory-Based Equation of State Analysis

pvap	3546.00	kPa	513.15	Vapor-Liquid Equilibria for the n-Pentane + 1-Propanol and n-Pentane + 2-Methyl-1-propanol Systems near the Critical Region
pvap	75.00	kPa	362.55	P, rho, T Measurements and Isobaric Vapor-Liquid-Equilibria of the 1,3,3-Trimethyl-2-oxabicycle[2,2,2]octane + Propan-1-ol Mixture: Cubic and Statistical Associating Fluid Theory-Based Equation of State Analysis
pvap	80.00	kPa	364.15	P, rho, T Measurements and Isobaric Vapor-Liquid-Equilibria of the 1,3,3-Trimethyl-2-oxabicycle[2,2,2]octane + Propan-1-ol Mixture: Cubic and Statistical Associating Fluid Theory-Based Equation of State Analysis
pvap	85.00	kPa	365.65	P, rho, T Measurements and Isobaric Vapor-Liquid-Equilibria of the 1,3,3-Trimethyl-2-oxabicycle[2,2,2]octane + Propan-1-ol Mixture: Cubic and Statistical Associating Fluid Theory-Based Equation of State Analysis
pvap	90.00	kPa	367.15	P, rho, T Measurements and Isobaric Vapor-Liquid-Equilibria of the 1,3,3-Trimethyl-2-oxabicycle[2,2,2]octane + Propan-1-ol Mixture: Cubic and Statistical Associating Fluid Theory-Based Equation of State Analysis

pvap	101.33	kPa	370.05	P, rho, T Measurements and Isobaric Vapor-Liquid-Equilibria of the 1,3,3-Trimethyl-2-oxabicycle[2,2,2]octane + Propan-1-ol Mixture: Cubic and Statistical Associating Fluid Theory-Based Equation of State Analysis
pvap	60.00	kPa	357.10	Isobaric Vapor Liquid Equilibria for the 1-Propanol + Ethylene Glycol Monopropyl Ether and 1-Butanol + Ethylene Glycol Monopropyl Ether Systems
pvap	80.00	kPa	364.20	Isobaric Vapor Liquid Equilibria for the 1-Propanol + Ethylene Glycol Monopropyl Ether and 1-Butanol + Ethylene Glycol Monopropyl Ether Systems
pvap	100.00	kPa	370.00	Isobaric Vapor Liquid Equilibria for the 1-Propanol + Ethylene Glycol Monopropyl Ether and 1-Butanol + Ethylene Glycol Monopropyl Ether Systems
pvap	0.70	kPa	278.15	Vapor Pressures and Activity Coefficients of (1-Propanol + 1,8-Cineole) at 10 Temperatures between 278.15 K and 323.15 K
pvap	1.01	kPa	283.15	Vapor Pressures and Activity Coefficients of (1-Propanol + 1,8-Cineole) at 10 Temperatures between 278.15 K and 323.15 K

pvap	1.45	kPa	288.15	Vapor Pressures and Activity Coefficients of (1-Propanol + 1,8-Cineole) at 10 Temperatures between 278.15 K and 323.15 K
pvap	2.04	kPa	293.15	Vapor Pressures and Activity Coefficients of (1-Propanol + 1,8-Cineole) at 10 Temperatures between 278.15 K and 323.15 K
pvap	2.84	kPa	298.15	Vapor Pressures and Activity Coefficients of (1-Propanol + 1,8-Cineole) at 10 Temperatures between 278.15 K and 323.15 K
pvap	3.87	kPa	303.15	Vapor Pressures and Activity Coefficients of (1-Propanol + 1,8-Cineole) at 10 Temperatures between 278.15 K and 323.15 K
pvap	5.23	kPa	308.15	Vapor Pressures and Activity Coefficients of (1-Propanol + 1,8-Cineole) at 10 Temperatures between 278.15 K and 323.15 K
pvap	7.01	kPa	313.15	Vapor Pressures and Activity Coefficients of (1-Propanol + 1,8-Cineole) at 10 Temperatures between 278.15 K and 323.15 K
pvap	9.29	kPa	318.15	Vapor Pressures and Activity Coefficients of (1-Propanol + 1,8-Cineole) at 10 Temperatures between 278.15 K and 323.15 K

pvap	12.18	kPa	323.15	Vapor Pressures and Activity Coefficients of (1-Propanol + 1,8-Cineole) at 10 Temperatures between 278.15 K and 323.15 K
pvap	600.00	kPa	427.20	Experimental Determination of Vapor Liquid Equilibria. Binary Systems of Methyl Acetate, Ethyl Acetate, and Propyl Acetate with 1-Propanol at 0.6 MPa
pvap	101.30	kPa	370.40	Evaluation of the 2-Methoxyethanol as Entrainer in Ethanol Water and 1-Propanol Water Mixtures
pvap	81.20	kPa	364.53	Vapor Liquid Equilibrium for the Systems trans-2-Butene + Methanol, + 1-Propanol, + 2-Propanol, + 2-Butanol, and + 2-Methyl-2-propanol at 364.5 K
pvap	40.00	kPa	347.55	Isobaric Vapor-Liquid Equilibria for Tetrahydropyran and Alcohol Systems
pvap	53.33	kPa	354.18	Isobaric Vapor-Liquid Equilibria for Tetrahydropyran and Alcohol Systems
pvap	66.66	kPa	359.56	Isobaric Vapor-Liquid Equilibria for Tetrahydropyran and Alcohol Systems
pvap	79.99	kPa	364.11	Isobaric Vapor-Liquid Equilibria for Tetrahydropyran and Alcohol Systems

pvap	93.32	kPa	368.08	Isobaric Vapor-Liquid Equilibria for Tetrahydropyran and Alcohol Systems	
pvap	98.66	kPa	369.55	Isobaric Vapor-Liquid Equilibria for Tetrahydropyran and Alcohol Systems	
pvap	7.03	kPa	313.15	Thermodynamics of mixtures with strongly negative deviations from Raoult s law Part 9. Vapor liquid equilibria for the system 1-propanol + di-n-propylamine at six temperatures between 293.15 and 318.15K	
pvap	2.03	kPa	293.15	Thermodynamics of mixtures with strongly negative deviations from Raoult s law Part 9. Vapor liquid equilibria for the system 1-propanol + di-n-propylamine at six temperatures between 293.15 and 318.15K	
pvap	2.81	kPa	298.15	Thermodynamics of mixtures with strongly negative deviations from Raoult s law Part 9. Vapor liquid equilibria for the system 1-propanol + di-n-propylamine at six temperatures between 293.15 and 318.15K	

pvap	3.88	kPa	303.15	Thermodynamics of mixtures with strongly negative deviations from Raoult s law Part 9. Vapor liquid equilibria for the system 1-propanol + di-n-propylamine at six temperatures between 293.15 and 318.15K	
pvap	5.25	kPa	308.15	Thermodynamics of mixtures with strongly negative deviations from Raoult s law Part 9. Vapor liquid equilibria for the system 1-propanol + di-n-propylamine at six temperatures between 293.15 and 318.15K	
pvap	9.31	kPa	318.15	Thermodynamics of mixtures with strongly negative deviations from Raoult s law Part 9. Vapor liquid equilibria for the system 1-propanol + di-n-propylamine at six temperatures between 293.15 and 318.15K	
pvap	2.79	kPa	298.15	Thermodynamic properties of mixtures containing ionic liquids Vapor pressures and activity coefficients of n-alcohols and benzene in binary mixtures with 1-methyl-3-butyl-imidazolium bis(trifluoromethyl-sulfonyl) imide	

pvap	3.83	kPa	303.15	Thermodynamic properties of mixtures containing ionic liquids Vapor pressures and activity coefficients of n-alcohols and benzene in binary mixtures with 1-methyl-3-butyl-imidazolium bis(trifluoromethyl-sulfonyl) imide
pvap	5.20	kPa	308.15	Thermodynamic properties of mixtures containing ionic liquids Vapor pressures and activity coefficients of n-alcohols and benzene in binary mixtures with 1-methyl-3-butyl-imidazolium bis(trifluoromethyl-sulfonyl) imide
pvap	6.99	kPa	313.15	Thermodynamic properties of mixtures containing ionic liquids Vapor pressures and activity coefficients of n-alcohols and benzene in binary mixtures with 1-methyl-3-butyl-imidazolium bis(trifluoromethyl-sulfonyl) imide
pvap	21.90	kPa	334.53	Phase equilibria on binary systems containing diethyl sulfide
pvap	31.20	kPa	342.07	Phase equilibria on binary systems containing diethyl sulfide
pvap	42.90	kPa	349.12	Phase equilibria on binary systems containing diethyl sulfide

pvap	52.70	kPa	353.96	Phase equilibria on binary systems containing diethyl sulfide
pvap	62.40	kPa	357.95	Phase equilibria on binary systems containing diethyl sulfide
pvap	71.40	kPa	361.25	Phase equilibria on binary systems containing diethyl sulfide
pvap	80.70	kPa	364.30	Phase equilibria on binary systems containing diethyl sulfide
pvap	91.60	kPa	367.51	Phase equilibria on binary systems containing diethyl sulfide
pvap	102.70	kPa	370.44	Phase equilibria on binary systems containing diethyl sulfide
pvap	6.99	kPa	313.17	P-x data for binary systems using a novel static total pressure apparatus
pvap	31.45	kPa	342.83	P-x data for binary systems using a novel static total pressure apparatus
pvap	49.80	kPa	352.68	P-x data for binary systems using a novel static total pressure apparatus
pvap	101.00	kPa	370.20	Heterogeneous azeotropic distillation for the separation of n-propanol + water mixture using n-propyl acetate as entrainer

pvap	25.60	kPa	337.95	Isobaric (vapour + liquid) equilibria for the (1-propanol + 1-butanol) binary mixture at (53.3 and 91.3) kPa
pvap	30.10	kPa	341.35	Isobaric (vapour + liquid) equilibria for the (1-propanol + 1-butanol) binary mixture at (53.3 and 91.3) kPa
pvap	37.40	kPa	346.15	Isobaric (vapour + liquid) equilibria for the (1-propanol + 1-butanol) binary mixture at (53.3 and 91.3) kPa
pvap	46.00	kPa	350.80	Isobaric (vapour + liquid) equilibria for the (1-propanol + 1-butanol) binary mixture at (53.3 and 91.3) kPa
pvap	51.10	kPa	353.25	Isobaric (vapour + liquid) equilibria for the (1-propanol + 1-butanol) binary mixture at (53.3 and 91.3) kPa
pvap	60.70	kPa	357.35	Isobaric (vapour + liquid) equilibria for the (1-propanol + 1-butanol) binary mixture at (53.3 and 91.3) kPa
pvap	74.70	kPa	362.45	Isobaric (vapour + liquid) equilibria for the (1-propanol + 1-butanol) binary mixture at (53.3 and 91.3) kPa
pvap	86.80	kPa	366.25	Isobaric (vapour + liquid) equilibria for the (1-propanol + 1-butanol) binary mixture at (53.3 and 91.3) kPa
pvap	101.80	kPa	370.40	Isobaric (vapour + liquid) equilibria for the (1-propanol + 1-butanol) binary mixture at (53.3 and 91.3) kPa

pvap	114.90	kPa	373.65	Isobaric (vapour + liquid) equilibria for the (1-propanol + 1-butanol) binary mixture at (53.3 and 91.3) kPa	
pvap	160.40	kPa	383.20	High-temperature vapour liquid equilibrium for the (water + alcohol) systems and modelling with SAFT-VR: 2. Water-1-propanol	
pvap	223.30	kPa	393.20	High-temperature vapour liquid equilibrium for the (water + alcohol) systems and modelling with SAFT-VR: 2. Water-1-propanol	
pvap	303.10	kPa	403.20	High-temperature vapour liquid equilibrium for the (water + alcohol) systems and modelling with SAFT-VR: 2. Water-1-propanol	
pvap	406.60	kPa	413.20	High-temperature vapour liquid equilibrium for the (water + alcohol) systems and modelling with SAFT-VR: 2. Water-1-propanol	
pvap	546.10	kPa	423.20	High-temperature vapour liquid equilibrium for the (water + alcohol) systems and modelling with SAFT-VR: 2. Water-1-propanol	
pvap	3.63	kPa	303.15	Isothermal (vapor + liquid) equilibria for the binary mixtures of (propylene oxide + ethanol) and (propylene oxide + 1-propanol) at several temperatures	

pvap	9.01	kPa	318.15	Isothermal (vapor + liquid) equilibria for the binary mixtures of (propylene oxide + ethanol) and (propylene oxide + 1-propanol) at several temperatures
pvap	19.93	kPa	333.15	Isothermal (vapor + liquid) equilibria for the binary mixtures of (propylene oxide + ethanol) and (propylene oxide + 1-propanol) at several temperatures
pvap	76.56	kPa	363.20	(Vapour + liquid) equilibria, (VLE) excess molar enthalpies and infinite dilution activity coefficients of selected binary systems involving n-hexyl pyridinium bis(trifluoromethylsulphonyl)imide ionic liquid: Experimental and predictions using modified UNIFAC (Dortmund)
pvap	100.00	kPa	369.80	Isobaric vapor-liquid equilibria for the 1-propanol + water + 1-ethyl-3-methylimidazolium dicyanamide system at 100 kPa
pvap	101.30	kPa	370.35	Isobaric vapor-liquid equilibria for extractive distillation of 1-propanol + water mixture using thiocyanate-based ionic liquids
pvap	100.51	kPa	370.10	Vapor-Liquid Equilibrium of Binary Mixtures Containing Isopropyl Acetate and Alkanols at 101.32 kPa

pvap	9.47	kPa	318.58	Vapor-Liquid Equilibrium for 1-Propanol + 1-Butene, + cis-2-Butene, + 2-Methyl-propene, + trans-2-Butene, + n-Butane, and + 2-Methyl-propane
pvap	125.86	kPa	376.30	Vapor-Liquid Equilibrium of Binary Mixtures Containing Isopropyl Acetate and Alkanols at 101.32 kPa
pvap	151.03	kPa	381.50	Vapor-Liquid Equilibrium of Binary Mixtures Containing Isopropyl Acetate and Alkanols at 101.32 kPa
pvap	176.20	kPa	385.89	Vapor-Liquid Equilibrium of Binary Mixtures Containing Isopropyl Acetate and Alkanols at 101.32 kPa
pvap	201.37	kPa	389.89	Vapor-Liquid Equilibrium of Binary Mixtures Containing Isopropyl Acetate and Alkanols at 101.32 kPa
pvap	226.54	kPa	393.59	Vapor-Liquid Equilibrium of Binary Mixtures Containing Isopropyl Acetate and Alkanols at 101.32 kPa
pvap	251.71	kPa	396.19	Vapor-Liquid Equilibrium of Binary Mixtures Containing Isopropyl Acetate and Alkanols at 101.32 kPa
pvap	101.30	kPa	370.33	Isobaric Vapor-Liquid Equilibrium of the Acetonitrile + 1-Propanol + Ionic Liquids at an Atmospheric Pressure

pvap	2.88	kPa	298.15	Vapor-Pressure Measurements of Liquid Solutions at Different Temperatures: Apparatus for Use over an Extended Temperature Range and Some New Data
pvap	4.00	kPa	303.15	Vapor-Pressure Measurements of Liquid Solutions at Different Temperatures: Apparatus for Use over an Extended Temperature Range and Some New Data
pvap	5.37	kPa	308.15	Vapor-Pressure Measurements of Liquid Solutions at Different Temperatures: Apparatus for Use over an Extended Temperature Range and Some New Data
pvap	7.20	kPa	313.15	Vapor-Pressure Measurements of Liquid Solutions at Different Temperatures: Apparatus for Use over an Extended Temperature Range and Some New Data
pvap	9.57	kPa	318.15	Vapor-Pressure Measurements of Liquid Solutions at Different Temperatures: Apparatus for Use over an Extended Temperature Range and Some New Data

pvap	12.42	kPa	323.15	Vapor-Pressure Measurements of Liquid Solutions at Different Temperatures: Apparatus for Use over an Extended Temperature Range and Some New Data	
pvap	16.14	kPa	328.15	Vapor-Pressure Measurements of Liquid Solutions at Different Temperatures: Apparatus for Use over an Extended Temperature Range and Some New Data	
pvap	20.72	kPa	333.15	Vapor-Pressure Measurements of Liquid Solutions at Different Temperatures: Apparatus for Use over an Extended Temperature Range and Some New Data	
pvap	26.41	kPa	338.15	Vapor-Pressure Measurements of Liquid Solutions at Different Temperatures: Apparatus for Use over an Extended Temperature Range and Some New Data	
pvap	33.23	kPa	343.15	Vapor-Pressure Measurements of Liquid Solutions at Different Temperatures: Apparatus for Use over an Extended Temperature Range and Some New Data	

pvap	41.63	kPa	348.15	Vapor-Pressure Measurements of Liquid Solutions at Different Temperatures: Apparatus for Use over an Extended Temperature Range and Some New Data	
pvap	51.67	kPa	353.15	Vapor-Pressure Measurements of Liquid Solutions at Different Temperatures: Apparatus for Use over an Extended Temperature Range and Some New Data	
pvap	63.69	kPa	358.15	Vapor-Pressure Measurements of Liquid Solutions at Different Temperatures: Apparatus for Use over an Extended Temperature Range and Some New Data	
pvap	77.81	kPa	363.15	Vapor-Pressure Measurements of Liquid Solutions at Different Temperatures: Apparatus for Use over an Extended Temperature Range and Some New Data	
pvap	5.88	kPa	310.17	Vapor-Liquid Equilibrium for Benzene + 2-Methylpentane and Allyl Alcohol + 1-Propanol	
pvap	6.98	kPa	313.15	Vapor-Liquid Equilibrium for Benzene + 2-Methylpentane and Allyl Alcohol + 1-Propanol	
pvap	7.69	kPa	314.84	Vapor-Liquid Equilibrium for Benzene + 2-Methylpentane and Allyl Alcohol + 1-Propanol	

pvap	10.18	kPa	319.89	Vapor-Liquid Equilibrium for Benzene + 2-Methylpentane and Allyl Alcohol + 1-Propanol
pvap	13.14	kPa	324.71	Vapor-Liquid Equilibrium for Benzene + 2-Methylpentane and Allyl Alcohol + 1-Propanol
pvap	15.75	kPa	328.14	Vapor-Liquid Equilibrium for Benzene + 2-Methylpentane and Allyl Alcohol + 1-Propanol
pvap	19.31	kPa	332.18	Vapor-Liquid Equilibrium for Benzene + 2-Methylpentane and Allyl Alcohol + 1-Propanol
pvap	12.32	kPa	323.15	Vapor-Liquid Equilibria and HE for Binary Systems of Dimethyl Ether (DME) with C1-C4 Alkan-1-ols at 323.15 K and Liquid-Liquid Equilibria for Ternary System of DME + Methanol + Water at 313.15 K
pvap	20.27	kPa	333.15	Vapor-Liquid Equilibrium for Benzene + 2-Methylpentane and Allyl Alcohol + 1-Propanol
pvap	22.56	kPa	335.34	Vapor-Liquid Equilibrium for Benzene + 2-Methylpentane and Allyl Alcohol + 1-Propanol
pvap	26.91	kPa	339.01	Vapor-Liquid Equilibrium for Benzene + 2-Methylpentane and Allyl Alcohol + 1-Propanol

pvap	31.12	kPa	342.11	Vapor-Liquid Equilibrium for Benzene + 2-Methylpentane and Allyl Alcohol + 1-Propanol
pvap	35.82	kPa	345.18	Vapor-Liquid Equilibrium for Benzene + 2-Methylpentane and Allyl Alcohol + 1-Propanol
pvap	40.12	kPa	347.71	Vapor-Liquid Equilibrium for Benzene + 2-Methylpentane and Allyl Alcohol + 1-Propanol
pvap	44.26	kPa	349.93	Vapor-Liquid Equilibrium for Benzene + 2-Methylpentane and Allyl Alcohol + 1-Propanol
pvap	49.36	kPa	352.44	Vapor-Liquid Equilibrium for Benzene + 2-Methylpentane and Allyl Alcohol + 1-Propanol
pvap	50.95	kPa	353.15	Vapor-Liquid Equilibrium for Benzene + 2-Methylpentane and Allyl Alcohol + 1-Propanol
pvap	50.95	kPa	353.17	Vapor-Liquid Equilibrium for Benzene + 2-Methylpentane and Allyl Alcohol + 1-Propanol
pvap	59.13	kPa	356.70	Vapor-Liquid Equilibrium for Benzene + 2-Methylpentane and Allyl Alcohol + 1-Propanol
pvap	12.07	kPa	323.07	Vapor-Liquid Equilibrium for 1-Propanol + 1-Butene, + cis-2-Butene, + 2-Methyl-propene, + trans-2-Butene, + n-Butane, and + 2-Methyl-propane

pvap	12.27	kPa	323.26	Vapor-Liquid Equilibrium for 1-Propanol + 1-Butene, + cis-2-Butene, + 2-Methyl-propene, + trans-2-Butene, + n-Butane, and + 2-Methyl-propane
pvap	17.57	kPa	330.16	Vapor-Liquid Equilibrium for 1-Propanol + 1-Butene, + cis-2-Butene, + 2-Methyl-propene, + trans-2-Butene, + n-Butane, and + 2-Methyl-propane
pvap	19.07	kPa	331.92	Vapor-Liquid Equilibrium for 1-Propanol + 1-Butene, + cis-2-Butene, + 2-Methyl-propene, + trans-2-Butene, + n-Butane, and + 2-Methyl-propane
pvap	19.17	kPa	332.05	Vapor-Liquid Equilibrium for 1-Propanol + 1-Butene, + cis-2-Butene, + 2-Methyl-propene, + trans-2-Butene, + n-Butane, and + 2-Methyl-propane
pvap	81.30	kPa	364.51	Isothermal Vapor Liquid Equilibrium for 2-Methylpropene + Methanol, + 1-Propanol, + 2-Propanol, + 2-Butanol, and + 2-Methyl-2-propanol Binary Systems at 364.5 K
pvap	1558.00	kPa	468.15	Vapor-Liquid Equilibria for the n-Pentane + 1-Propanol and n-Pentane + 2-Methyl-1-propanol Systems near the Critical Region

pvap	2081.00	kPa	483.15	Vapor-Liquid Equilibria for the n-Pentane + 1-Propanol and n-Pentane + 2-Methyl-1-propanol Systems near the Critical Region
pvap	70.00	kPa	360.85	P, rho, T Measurements and Isobaric Vapor-Liquid-Equilibria of the 1,3,3-Trimethyl-2-oxabicycle[2,2,2]octane + Propan-1-ol Mixture: Cubic and Statistical Associating Fluid Theory-Based Equation of State Analysis
rfi	1.38306		298.15	Effect of anion fluorination in 1-ethyl-3-methylimidazolium as solvent for the liquid extraction of ethanol from ethyl tert-butyl ether
rfi	1.38370		298.15	Activity coefficients of the binary mixtures of a-cresol or p-cresol with C l-C4 aliphatic alcohols near ambient pressure
rfi	1.38360		298.15	Isobaric vapor-liquid equilibria for the binary systems 1-propyl alcohol + dipropyl ether and 1-butyl alcohol + dibutyl ether at 20 and 101.3 kPa
rfi	1.38360		298.15	Vapor liquid equilibria in the ternary system dipropyl ether + 1-propanol + 1-pentanol and the binary systems dipropyl ether + 1-pentanol, 1-propanol + 1-pentanol at 101.3 kPa

rfi	1.38390	298.20	Vapor liquid equilibria for the ternary mixture of carbon dioxide + 1-propanol + propyl acetate at elevated pressures
rfi	1.38500	293.15	Phase equilibria of water + 1-propanol + solvent (n-amyl acetate, cyclohexanol, and cyclohexyl acetate) at T = 298.2K
rfi	1.38360	298.15	Phase equilibria involved in extractive distillation of dipropyl ether + 1-propyl alcohol using 2-ethoxyethanol as entrainer
rfi	1.38333	298.15	Vapor liquid equilibrium, densities, and interfacial tensions for the system ethyl 1,1-dimethylethyl ether (ETBE) + propan-1-ol
rfi	1.38080	303.15	Thermodynamic properties of (tetradecane + benzene, + toluene, + chlorobenzene, + bromobenzene, + anisole) binary mixtures at T = (298.15, 303.15, and 308.15) K
rfi	1.38360	298.15	Liquid liquid equilibria of the systems dipropyl ether + n-propanol +water and dipropyl ether + n-propanol + ethylene glycol at different temperatures

rfi	1.38305	298.15	Isobaric vapour liquid equilibria for binary systems of 2-butanone with ethanol, 1-propanol, and 2-propanol at 20 and 101.3 kPa
rfi	1.38312	298.15	Isothermal vapor liquid equilibrium at 323.15K and excess molar volumes and refractive indices at 298.15K for the ternary system propyl vinyl ether + 1-propanol + benzene and its binary sub-systems
rfi	1.38350	298.20	Measurement and prediction of tie-line data for mixtures of (water + 1-propanol + diisopropyl ether): LLE diagrams as a function of temperature
rfi	1.38320	298.15	Excess molar volumes, excess molar enthalpies and refractive index deviations for binary mixtures of propan-1-ol, butan-1-ol and pentan-1-ol with 2,2,4-trimethylpentane at 298.15 K
rfi	1.38290	298.15	Experimental and predicted volumetric and refractive index properties of ternary mixtures of iodoethane with toluene and alcohols at temperature 298.15 K and pressure 101 kPa

rfi	1.38340	298.15	Excess molar volumes and partial molar volumes for (propionitrile + an alkanol) at T = 298.15 K and p = 0.1 MPa
rfi	1.38400	298.15	Physico-chemical and excess properties of the binary mixtures of methylcyclohexane + ethanol, + propan-1-ol, + propan-2-ol, + butan-1-ol, + 2-methyl-1-propanol, or 3-methyl-1-butanol at T = (298.15, 303.15, and 308.15) K
rfi	1.38190	303.15	Physico-chemical and excess properties of the binary mixtures of methylcyclohexane + ethanol, + propan-1-ol, + propan-2-ol, + butan-1-ol, + 2-methyl-1-propanol, or 3-methyl-1-butanol at T = (298.15, 303.15, and 308.15) K
rfi	1.38000	308.15	Physico-chemical and excess properties of the binary mixtures of methylcyclohexane + ethanol, + propan-1-ol, + propan-2-ol, + butan-1-ol, + 2-methyl-1-propanol, or 3-methyl-1-butanol at T = (298.15, 303.15, and 308.15) K

rfi	1.38260	298.15	Thermodynamic properties of (tetradecane + benzene, + toluene, + chlorobenzene, + bromobenzene, + anisole) binary mixtures at T = (298.15, 303.15, and 308.15) K	
rfi	1.37820	308.15	Thermodynamic properties of (tetradecane + benzene, + toluene, + chlorobenzene, + bromobenzene, + anisole) binary mixtures at T = (298.15, 303.15, and 308.15) K	
rfi	1.38370	298.15	(Vapor + liquid) equilibria of the binary mixtures of m-cresol with C1 C4 aliphatic alcohols at 95.5 kPa	
rfi	1.38260	298.15	Thermodynamic interactions in binary mixtures of anisole with ethanol, propan-1-ol, propan-2-ol, butan-1-ol, pentan-1-ol, and 3-methylbutan-1-ol at T = (298.15, 303.15, and 308.15) K	
rfi	1.38080	303.15	Thermodynamic interactions in binary mixtures of anisole with ethanol, propan-1-ol, propan-2-ol, butan-1-ol, pentan-1-ol, and 3-methylbutan-1-ol at T = (298.15, 303.15, and 308.15) K	

rfi	1.37820	308.15	Thermodynamic interactions in binary mixtures of anisole with ethanol, propan-1-ol, propan-2-ol, butan-1-ol, pentan-1-ol, and 3-methylbutan-1-ol at T = (298.15, 303.15, and 308.15) K
rfi	1.38307	298.15	Thermophysical properties of the binary mixtures (1,8-cineole + 1-alkanol) at T = (298.15 and 313.15) K and at atmospheric pressure
rfi	1.37674	313.15	Thermophysical properties of the binary mixtures (1,8-cineole + 1-alkanol) at T = (298.15 and 313.15) K and at atmospheric pressure
rfi	1.38906	283.15	Thermophysical properties of {(+-)-linalool + propan-1-ol}: A first stage towards the development of a green process
rfi	1.38288	298.15	Thermophysical properties of {(+-)-linalool + propan-1-ol}: A first stage towards the development of a green process
rfi	1.37660	313.15	Thermophysical properties of {(+-)-linalool + propan-1-ol}: A first stage towards the development of a green process
rfi	1.37015	328.15	Thermophysical properties of {(+-)-linalool + propan-1-ol}: A first stage towards the development of a green process

rfi	1.38220	298.15	Application of the ERAS model to volumetric properties of binary mixtures of banana oil with primary and secondary alcohols (C1- C4) at different temperatures
rfi	1.38307	298.15	Effect of the temperature on the physical properties of pure 1-propyl-3-methylimidazolium bis(trifluoromethylsulfonyl)imide and characterization of its binary mixtures with alcohols
rfi	1.38210	303.15	Experimental study on the calorimetric data of cyclohexanol with alkanols (C1-C4) and correlation with Wilson, NRTL and UNIQUAC models at 300 K
rfi	1.38780	288.15	Experimental determination and modeling of excess molar volumes, viscosities and refractive indices of the binary systems (pyridine + 1-propanol, + 1,2-propanediol, + 1,3-propanediol, and + glycerol). New UNIFAC-VISCO parameters determination

rfi	1.38576		293.15	Experimental determination and modeling of excess molar volumes, viscosities and refractive indices of the binary systems (pyridine + 1-propanol, + 1,2-propanediol, + 1,3-propanediol, and + glycerol). New UNIFAC-VISCO parameters determination
rfi	1.38369		298.15	Experimental determination and modeling of excess molar volumes, viscosities and refractive indices of the binary systems (pyridine + 1-propanol, + 1,2-propanediol, + 1,3-propanediol, and + glycerol). New UNIFAC-VISCO parameters determination
rfi	1.38162		303.15	Experimental determination and modeling of excess molar volumes, viscosities and refractive indices of the binary systems (pyridine + 1-propanol, + 1,2-propanediol, + 1,3-propanediol, and + glycerol). New UNIFAC-VISCO parameters determination

rfi	1.37918	308.15	Experimental determination and modeling of excess molar volumes, viscosities and refractive indices of the binary systems (pyridine + 1-propanol, + 1,2-propanediol, + 1,3-propanediol, and + glycerol). New UNIFAC-VISCO parameters determination
rfi	1.37695	313.15	Experimental determination and modeling of excess molar volumes, viscosities and refractive indices of the binary systems (pyridine + 1-propanol, + 1,2-propanediol, + 1,3-propanediol, and + glycerol). New UNIFAC-VISCO parameters determination
rfi	1.37469	318.15	Experimental determination and modeling of excess molar volumes, viscosities and refractive indices of the binary systems (pyridine + 1-propanol, + 1,2-propanediol, + 1,3-propanediol, and + glycerol). New UNIFAC-VISCO parameters determination

rfi	1.37231	323.15	Experimental determination and modeling of excess molar volumes, viscosities and refractive indices of the binary systems (pyridine + 1-propanol, + 1,2-propanediol, + 1,3-propanediol, and + glycerol). New UNIFAC-VISCO parameters determination
rfi	1.38334	298.15	Densities, speeds of sound, and refractive indices for binary mixtures of 1-butyl-3-methylimidazolium methyl sulphate ionic liquid with alcohols at T = (298.15, 303.15, 308.15, and 313.15) K
rfi	1.38131	303.15	Densities, speeds of sound, and refractive indices for binary mixtures of 1-butyl-3-methylimidazolium methyl sulphate ionic liquid with alcohols at T = (298.15, 303.15, 308.15, and 313.15) K
rfi	1.37926	308.15	Densities, speeds of sound, and refractive indices for binary mixtures of 1-butyl-3-methylimidazolium methyl sulphate ionic liquid with alcohols at T = (298.15, 303.15, 308.15, and 313.15) K

rfi	1.37722	313.15	Densities, speeds of sound, and refractive indices for binary mixtures of 1-butyl-3-methylimidazolium methyl sulphate ionic liquid with alcohols at T = (298.15, 303.15, 308.15, and 313.15) K	
rfi	1.38330	298.15	Properties of pure n-butylammonium nitrate ionic liquid and its binary mixtures of with alcohols at T = (293.15 to 313.15) K	
rfi	1.38432	293.15	Separation of thiophene from heptane with ionic liquids	
rfi	1.38785	288.15	Experimental Determination and Modeling of Densities and Refractive Indices of the Binary Systems Alcohol + Dicyclohexylamine at T = (288.15 to 323.15) K	
rfi	1.38578	293.15	Experimental Determination and Modeling of Densities and Refractive Indices of the Binary Systems Alcohol + Dicyclohexylamine at T = (288.15 to 323.15) K	
rfi	1.38364	298.15	Experimental Determination and Modeling of Densities and Refractive Indices of the Binary Systems Alcohol + Dicyclohexylamine at T = (288.15 to 323.15) K	

rfi	1.38153	303.15	Experimental Determination and Modeling of Densities and Refractive Indices of the Binary Systems Alcohol + Dicyclohexylamine at T = (288.15 to 323.15) K
rfi	1.37929	308.15	Experimental Determination and Modeling of Densities and Refractive Indices of the Binary Systems Alcohol + Dicyclohexylamine at T = (288.15 to 323.15) K
rfi	1.37702	313.15	Experimental Determination and Modeling of Densities and Refractive Indices of the Binary Systems Alcohol + Dicyclohexylamine at T = (288.15 to 323.15) K
rfi	1.37483	318.15	Experimental Determination and Modeling of Densities and Refractive Indices of the Binary Systems Alcohol + Dicyclohexylamine at T = (288.15 to 323.15) K
rfi	1.37252	323.15	Experimental Determination and Modeling of Densities and Refractive Indices of the Binary Systems Alcohol + Dicyclohexylamine at T = (288.15 to 323.15) K

rfi	1.38350	298.15	Molar excess enthalpies and molar excess volumes of formamide + 1-propanol or 2-propanol and thermodynamic modeling by Prigogine-Flory-Patterson theory and Treszczanowicz-Benson association model
rfi	1.38512	293.15	Vapor Liquid Equilibrium for Methyl Isobutyl Ketone (MIBK) + (1-Propanol or 2-Propanol) Binary Mixtures
rfi	1.38340	298.00	Determination of Physicochemical Parameters of Sodium Dodecyl Sulfate in Aqueous Micellar Solutions Containing Short-Chain Alcohols
rfi	1.38520	293.15	P-rho-T Data and Modeling for Propan-1-ol + n-Octane or n-Nonane or n-Decane from 313.15 K to 363.15 K and 1 MPa to 20 MPa
rfi	1.38480	293.15	Isothermal Vapor-Liquid Equilibrium Measurements for the Pentan-2-one + Propan-1-ol/Butan-1-ol System within 342-363 K
rfi	1.38530	293.15	Isothermal Vapor-Liquid Equilibrium Measurements for Alcohol + Water/n-Hexane Azeotropic Systems Using Both Dynamic and Automated Static-Synthetic Methods

rfi	1.38520	293.15	Solid-Liquid Equilibrium Measurements for Posaconazole and Voriconazole in Several Solvents between T = 278.2 and 323.2 K Using Differential Thermal Analysis/Thermal Gravimetric Analysis
rfi	1.38540	293.15	Solubility Data for Roflumilast and Maraviroc in Various Solvents between T = (278.2-323.2) K
rfi	1.38070	298.15	Bubble Temperatures of the Binary Mixtures of Dimethylcarbonate with Some Alcohols at 95.8 kPa
rfi	1.38690	288.15	Excess Molar Volumes, Viscosities, and Refractive Indexes for Binary Mixtures of 1-Chlorobutane with Four Alcohols at T = (288.15, 298.15 and 308.15) K
rfi	1.38290	298.15	Excess Molar Volumes, Viscosities, and Refractive Indexes for Binary Mixtures of 1-Chlorobutane with Four Alcohols at T = (288.15, 298.15 and 308.15) K

rfi	1.37890	308.15	Excess Molar Volumes, Viscosities, and Refractive Indexes for Binary Mixtures of 1-Chlorobutane with Four Alcohols at T = (288.15, 298.15 and 308.15) K
rfi	1.38550	293.15	Limiting Activity Coefficients by Comparative Tensimetry: 1-Propanol and 1-Butanol in Heptane and in Octane
rfi	1.38300	298.15	Density, Surface Tension, and Refractive Index of Octane + 1-Alkanol Mixtures at T) 298.15 K.
rfi	1.38304	298.15	Density, Refractive Index, Speed of Sound at 298.15 K, and Vapor-Liquid Equilibria at 101.3 kPa for Binary Mixtures of Propanol + 2-Methyl-1-butanol and Propanol + 3-Methyl-1-butanol
rfi	1.38330	298.15	Excess Molar Enthalpies of 1,2-Propanediol + Alkan-1-ols (C1-C6) and Their Correlations at 298.15 K and Ambient Pressure (81.5 kPa)
rfi	1.38302	298.15	Physical Properties of Binary Mixtures of the Ionic Liquid 1-Methyl-3-octylimidazolium Chloride with Methanol, Ethanol, and 1-Propanol at T = (298.15, 313.15, and 328.15) K and at P) 0.1 MPa

rfi	1.38330	298.15	Effect of Pressure on the Static Relative Permittivities of Alkan-1-ols at 298.15 K
rfi	1.38360	298.15	Isobaric Vapor-Liquid Equilibria for Binary and Ternary Mixtures of Dipropyl Ether, 1-Propyl Alcohol, and Butyl Propionate
rfi	1.38360	298.15	Phase Equilibria Involved in Extractive Distillation of Dipropyl Ether + 1-Propyl Alcohol Using N,N-Dimethylformamide as Entrainer
rfi	1.38330	298.15	Excess Molar Enthalpies of Benzyl Alcohol + Alkanols (C1-C6) and Their Correlations at 298.15 K and Ambient Pressure
rfi	1.38552	293.15	Excess Volume of the 1-Propanol + 1-Alkene Systems in Terms of an Equation of State with Association
rfi	1.38560	298.15	Isobaric Vapor-Liquid Equilibria for Binary and Ternary Mixtures of Propanal, Propanol, and Propanoic Acid
rfi	1.38300	298.15	Excess Molar Enthalpies of 2-Methyl-2-butanol (1) + 1-Alkanols (C1-C5) (2) at 298.15 K
rfi	1.38305	298.15	Liquid Liquid Phase Equilibria of 1-Propanol + Water + n-Alcohol Ternary Systems at 298.15 K and Atmospheric Pressure

rfi	1.38330	298.15	Isothermal Vapor Liquid Equilibrium Data of Propan-1-ol + 2,2,4-Trimethylpentane and Butan-1-ol + 2,2,4-Trimethylpentane at 318.15 K
rfi	1.38340	298.15	Density, Viscosity, and Refractive Index Properties for the Binary Mixtures of n-Butylammonium Acetate Ionic Liquid + Alkanols at Several Temperatures
rfi	1.38200	293.15	Isothermal Vapor Liquid Equilibrium Data for the Propan-1-ol + Dodecane System at (323.0, 343.4, 353.2, 363.1, and 369.2) K
rfi	1.38510	293.20	Vapor Liquid Equilibrium Data for Binary Systems of 1-Methyl-4-(1-methylethenyl)-cyclohexene + {Ethanol, Propan-1-ol, Propan-2-ol, Butan-1-ol, Pentan-1-ol, or Hexan-1-ol} at 40 kPa
rfi	1.38510	293.20	Vapor Liquid Equilibrium Data for Binary Systems of 1H-Pyrrole with Butan-1-ol, Propan-1-ol, or Pentan-1-ol
rfi	1.38510	293.15	Isobaric Vapor Liquid Equilibrium Data of 2-Methyl-propan-2-ol (1) + Heptan-1-ol (2), Methanol (1) + Heptan-1-ol (2), Ethanol (1) + Heptan-1-ol (2), and Propan-1-ol (1) + Heptan-1-ol (2) at 96.5 kPa

rfi	1.38540	295.15	Isobaric Vapor Liquid Equilibrium Data for Binary Mixtures of 1-Phenylethanone with a Few Alcohols at 95.2 kPa
rfi	1.38510	293.15	Vapor Liquid Equilibrium Data for Binary Systems of n-Dodecane + {Propan-1-ol, Butan-1-ol, 2-Methylpropan-1-ol} at 40 kPa
rfi	1.38310	298.15	Physical Properties of Binary Mixtures of the Ionic Liquid 1-Ethyl-3-methylimidazolium Ethyl Sulfate with Several Alcohols at T = (298.15, 313.15, and 328.15) K and Atmospheric Pressure
rfi	1.38305	298.15	Experimental Liquid-Liquid Equilibria of 1-Alkyl-3-methylimidazolium Hexafluorophosphate with 1-Alcohols
rfi	1.38309	298.15	Density and Viscosity Experimental Data of the Ternary Mixtures 1-Propanol or 2-Propanol + Water + 1-Ethyl-3-methylimidazolium Ethylsulfate. Correlation and Prediction of Physical Properties of the Ternary Systems
rfi	1.38530	293.15	Measurement and Correlation of the Solubilities of m-Phthalic Acid in Monobasic Alcohols

rfi	1.38370		298.15	Experimental Determination, Correlation, and Prediction of Physical Properties of the Ternary Mixtures Ethanol and 1-Propanol + Water + 1-Ethyl-3-methylpyridinium Ethylsulfate at 298.15 K
rfi	1.38310		298.15	Synthesis and Physical Properties of 1-Ethylpyridinium Ethylsulfate and its Binary Mixtures with Ethanol and 1-Propanol at Several Temperatures
rfi	1.38286		298.15	Effect of the temperature on the physical properties of the pure ionic liquid 1-ethyl-3-methylimidazolium methylsulfate and characterization of its binary mixtures with alcohols
rfi	1.38308		298.15	Isobaric Vapor-Liquid Equilibria at 101.32 kPa and Densities, Speeds of Sound, and Refractive Indices at 298.15 K for MTBE or DIPE or TAME + 1-Propanol Binary Systems
rhol	783.10	kg/m3	318.15	Thermophysical and volumetric study of mixtures {p-cymene + propan-1-ol} at several temperatures and atmospheric pressure. Modeling with COSMO-RS

rhol	774.63	kg/m3	328.15	Thermophysical and volumetric study of mixtures {p-cymene + propan-1-ol} at several temperatures and atmospheric pressure. Modeling with COSMO-RS
rhol	803.52	kg/m3	293.15	Electrostriction of water and lower alcohols around ammonium nitrate - Volumetric approach
rhol	799.52	kg/m3	298.15	Electrostriction of water and lower alcohols around ammonium nitrate - Volumetric approach
rhol	795.50	kg/m3	303.15	Electrostriction of water and lower alcohols around ammonium nitrate - Volumetric approach
rhol	791.43	kg/m3	308.15	Electrostriction of water and lower alcohols around ammonium nitrate - Volumetric approach
rhol	787.31	kg/m3	313.15	Electrostriction of water and lower alcohols around ammonium nitrate - Volumetric approach
rhol	783.11	kg/m3	318.15	Electrostriction of water and lower alcohols around ammonium nitrate - Volumetric approach
rhol	778.79	kg/m3	323.15	Electrostriction of water and lower alcohols around ammonium nitrate - Volumetric approach

rhol	815.55	kg/m3	278.15	Thermodynamic evidence for nano-heterogeneity in solutions of the macrocycle C-butylresorcin[4]arene in non-aqueous solvents
rhol	811.60	kg/m3	283.15	Thermodynamic evidence for nano-heterogeneity in solutions of the macrocycle C-butylresorcin[4]arene in non-aqueous solvents
rhol	807.64	kg/m3	288.15	Thermodynamic evidence for nano-heterogeneity in solutions of the macrocycle C-butylresorcin[4]arene in non-aqueous solvents
rhol	803.66	kg/m3	293.15	Thermodynamic evidence for nano-heterogeneity in solutions of the macrocycle C-butylresorcin[4]arene in non-aqueous solvents
rhol	791.36	kg/m3	308.15	Thermophysical and volumetric study of mixtures {p-cymene + propan-1-ol} at several temperatures and atmospheric pressure. Modeling with COSMO-RS
rhol	795.62	kg/m3	303.15	Thermodynamic evidence for nano-heterogeneity in solutions of the macrocycle C-butylresorcin[4]arene in non-aqueous solvents
rhol	791.55	kg/m3	308.15	Thermodynamic evidence for nano-heterogeneity in solutions of the macrocycle C-butylresorcin[4]arene in non-aqueous solvents

rhol	803.76	kg/m3	293.15	Intermolecular interactions in binary mixtures of 2-diethylethanolamine with 1-propanol and 1-butanol at different temperatures
rhol	795.82	kg/m3	303.15	Intermolecular interactions in binary mixtures of 2-diethylethanolamine with 1-propanol and 1-butanol at different temperatures
rhol	787.88	kg/m3	313.15	Intermolecular interactions in binary mixtures of 2-diethylethanolamine with 1-propanol and 1-butanol at different temperatures
rhol	803.00	kg/m3	293.15	Density, speed of sound and refractive index of mixtures containing 2-phenoxyethanol with propanol or butanol at various temperatures
rhol	796.00	kg/m3	303.15	Density, speed of sound and refractive index of mixtures containing 2-phenoxyethanol with propanol or butanol at various temperatures
rhol	788.00	kg/m3	313.15	Density, speed of sound and refractive index of mixtures containing 2-phenoxyethanol with propanol or butanol at various temperatures

rhol	780.00	kg/m3	323.15	Density, speed of sound and refractive index of mixtures containing 2-phenoxyethanol with propanol or butanol at various temperatures
rhol	800.00	kg/m3	298.15	Density, speed of sound and refractive index of mixtures containing 2-phenoxyethanol with propanol or butanol at various temperatures
rhol	807.50	kg/m3	288.15	Mass density, sound velocity, mixing enthalpy, 1H NMR, Ab initio calculations and intermolecular interactions in binary mixtures of N-methylimidazole + water, +methanol, +ethanol, +1-propanol, +2-propanol
rhol	799.53	kg/m3	298.15	Mass density, sound velocity, mixing enthalpy, 1H NMR, Ab initio calculations and intermolecular interactions in binary mixtures of N-methylimidazole + water, +methanol, +ethanol, +1-propanol, +2-propanol

rhol	791.44	kg/m3	308.15	Mass density, sound velocity, mixing enthalpy, 1H NMR, Ab initio calculations and intermolecular interactions in binary mixtures of N-methylimidazole + water, +methanol, +ethanol, +1-propanol, +2-propanol
rhol	783.18	kg/m3	318.15	Mass density, sound velocity, mixing enthalpy, 1H NMR, Ab initio calculations and intermolecular interactions in binary mixtures of N-methylimidazole + water, +methanol, +ethanol, +1-propanol, +2-propanol
rhol	774.70	kg/m3	328.15	Mass density, sound velocity, mixing enthalpy, 1H NMR, Ab initio calculations and intermolecular interactions in binary mixtures of N-methylimidazole + water, +methanol, +ethanol, +1-propanol, +2-propanol
rhol	803.77	kg/m3	293.15	Volumetric properties of monoethanolamine and alcohol binary mixtures at different temperatures and 0.1 MPa
rhol	799.69	kg/m3	298.15	Volumetric properties of monoethanolamine and alcohol binary mixtures at different temperatures and 0.1 MPa

rhol	795.58	kg/m3	303.15	Volumetric properties of monoethanolamine and alcohol binary mixtures at different temperatures and 0.1 MPa
rhol	791.52	kg/m3	308.15	Volumetric properties of monoethanolamine and alcohol binary mixtures at different temperatures and 0.1 MPa
rhol	787.42	kg/m3	313.15	Volumetric properties of monoethanolamine and alcohol binary mixtures at different temperatures and 0.1 MPa
rhol	783.17	kg/m3	318.15	Volumetric properties of monoethanolamine and alcohol binary mixtures at different temperatures and 0.1 MPa
rhol	778.96	kg/m3	323.15	Volumetric properties of monoethanolamine and alcohol binary mixtures at different temperatures and 0.1 MPa
rhol	799.97	kg/m3	298.15	Excess volumes and partial molar volumes of binary liquid mixtures of furfural or 2-methylfuran with alcohols at 298.15 K
rhol	795.53	kg/m3	303.15	Vapor-liquid equilibrium and excess properties of the binary mixtures formed by ethyl isobutyrate and n-alkanols

rhol	779.09	kg/m3	323.15	Vapor-liquid equilibrium and excess properties of the binary mixtures formed by ethyl isobutyrate and n-alkanols
rhol	799.51	kg/m3	298.15	Excess molar enthalpies of methyl isobutyl ketone (MIBK) with alkan-1-ols (C1-C6) and their correlations at 298.15 K
rhol	799.10	kg/m3	298.15	Excess enthalpies of binary mixtures of some propylamines + some propanols at 298.15K
rhol	795.48	kg/m3	303.15	Excess molar enthalpies and heat capacities of dimethyl sulfoxide + seven normal alkanols at 303.15K and atmospheric pressure
rhol	799.30	kg/m3	298.15	Thermodynamic properties of binary mixtures of 2.2.2-Trifluoroethanol with Water or Alkanols at T=298.15 K
rhol	795.61	kg/m3	303.15	Volumetric and transport properties of ternary mixtures containing 1-propanol + ethyl ethanoate + cyclohexane or benzene at 303.15 K: Experimental data, correlation and prediction by ERAS model
rhol	799.28	kg/m3	298.15	Topological investigations of the molecular species and molecular interactions that characterize pyrrolidin-2-one + lower alkanol mixtures

rhol	800.03	kg/m3	293.15	Excess molar volumes of Diisopropylamine + (C1-C5) Alkan-1-ols: application of the ERAS model and cubic EOS	
rhol	797.81	kg/m3	298.15	Excess molar volumes of Diisopropylamine + (C1-C5) Alkan-1-ols: application of the ERAS model and cubic EOS	
rhol	795.05	kg/m3	303.15	Excess molar volumes of Diisopropylamine + (C1-C5) Alkan-1-ols: application of the ERAS model and cubic EOS	
rhol	787.95	kg/m3	313.15	Excess molar volumes of Diisopropylamine + (C1-C5) Alkan-1-ols: application of the ERAS model and cubic EOS	
rhol	807.75	kg/m3	288.15	Experimental measurements and modelling of volumetric properties, refractive index and viscosity of selected binary systems with butyl lactate at 288.15 to 323.15 K and atmospheric pressure. New UNIFAC-VISCO interaction parameters.	

rhol	803.77	kg/m3	293.15	Experimental measurements and modelling of volumetric properties, refractive index and viscosity of selected binary systems with butyl lactate at 288.15 to 323.15 K and atmospheric pressure. New UNIFAC-VISCO interaction parameters.	
rhol	799.75	kg/m3	298.15	Experimental measurements and modelling of volumetric properties, refractive index and viscosity of selected binary systems with butyl lactate at 288.15 to 323.15 K and atmospheric pressure. New UNIFAC-VISCO interaction parameters.	
rhol	795.72	kg/m3	303.15	Experimental measurements and modelling of volumetric properties, refractive index and viscosity of selected binary systems with butyl lactate at 288.15 to 323.15 K and atmospheric pressure. New UNIFAC-VISCO interaction parameters.	

rhol	791.64	kg/m3	308.15	Experimental measurements and modelling of volumetric properties, refractive index and viscosity of selected binary systems with butyl lactate at 288.15 to 323.15 K and atmospheric pressure. New UNIFAC-VISCO interaction parameters.
rhol	787.52	kg/m3	313.15	Experimental measurements and modelling of volumetric properties, refractive index and viscosity of selected binary systems with butyl lactate at 288.15 to 323.15 K and atmospheric pressure. New UNIFAC-VISCO interaction parameters.
rhol	783.36	kg/m3	318.15	Experimental measurements and modelling of volumetric properties, refractive index and viscosity of selected binary systems with butyl lactate at 288.15 to 323.15 K and atmospheric pressure. New UNIFAC-VISCO interaction parameters.

rhol	779.14	kg/m3	323.15	Experimental measurements and modelling of volumetric properties, refractive index and viscosity of selected binary systems with butyl lactate at 288.15 to 323.15 K and atmospheric pressure. New UNIFAC-VISCO interaction parameters.	
rhol	803.30	kg/m3	293.15	Excess volumes, Viscosities, and Excess Gibbs Energy of Activation for Viscous Flow, for binary and ternary mixtures 1- propanol + N-N dimethylformamid + chloroform at different temperatures	
rhol	795.10	kg/m3	303.15	Excess volumes, Viscosities, and Excess Gibbs Energy of Activation for Viscous Flow, for binary and ternary mixtures 1- propanol + N-N dimethylformamid + chloroform at different temperatures	
rhol	787.00	kg/m3	313.15	Excess volumes, Viscosities, and Excess Gibbs Energy of Activation for Viscous Flow, for binary and ternary mixtures 1- propanol + N-N dimethylformamid + chloroform at different temperatures	

rhol	778.70	kg/m3	323.15	Excess volumes, Viscosities, and Excess Gibbs Energy of Activation for Viscous Flow, for binary and ternary mixtures 1- propanol + N-N dimethylformamid + chloroform at different temperatures
rhol	807.51	kg/m3	288.15	The excess molar volume and the molar surface Gibbs energy of the binary of the ether-functionalized ionic liquids [C22O1IM][TfO] with ethanol and isomeric propanols at T = (288.15-318.15) K
rhol	803.52	kg/m3	293.15	The excess molar volume and the molar surface Gibbs energy of the binary of the ether-functionalized ionic liquids [C22O1IM][TfO] with ethanol and isomeric propanols at T = (288.15-318.15) K
rhol	799.51	kg/m3	298.15	The excess molar volume and the molar surface Gibbs energy of the binary of the ether-functionalized ionic liquids [C22O1IM][TfO] with ethanol and isomeric propanols at T = (288.15-318.15) K

rhol	795.48	kg/m3	303.15	The excess molar volume and the molar surface Gibbs energy of the binary of the ether-functionalized ionic liquids [C22O1IM][TfO] with ethanol and isomeric propanols at T = (288.15-318.15) K
rhol	791.42	kg/m3	308.15	The excess molar volume and the molar surface Gibbs energy of the binary of the ether-functionalized ionic liquids [C22O1IM][TfO] with ethanol and isomeric propanols at T = (288.15-318.15) K
rhol	787.31	kg/m3	313.15	The excess molar volume and the molar surface Gibbs energy of the binary of the ether-functionalized ionic liquids [C22O1IM][TfO] with ethanol and isomeric propanols at T = (288.15-318.15) K
rhol	783.15	kg/m3	318.15	The excess molar volume and the molar surface Gibbs energy of the binary of the ether-functionalized ionic liquids [C22O1IM][TfO] with ethanol and isomeric propanols at T = (288.15-318.15) K
rhol	801.08	kg/m3	298.15	Modified Method for Measuring the Solubility of Pharmaceutical Compounds in Organic Solvents by Visual Camera

rhol	815.46	kg/m3	278.15	Density, Viscosity, and Sound Speed of Bis(trifluoromethylsulfonyl)imide-Based Ionic Liquids + 1-Propanol Mixtures
rhol	799.55	kg/m3	298.15	Density, Viscosity, and Sound Speed of Bis(trifluoromethylsulfonyl)imide-Based Ionic Liquids + 1-Propanol Mixtures
rhol	783.18	kg/m3	318.15	Density, Viscosity, and Sound Speed of Bis(trifluoromethylsulfonyl)imide-Based Ionic Liquids + 1-Propanol Mixtures
rhol	765.96	kg/m3	338.15	Density, Viscosity, and Sound Speed of Bis(trifluoromethylsulfonyl)imide-Based Ionic Liquids + 1-Propanol Mixtures
rhol	803.49	kg/m3	293.15	Densities and Viscosities of Binary Mixtures of 2-Ethyl-1,1,3,3-tetramethylguanidinium Ionic Liquids with Ethanol and 1-Propanol
rhol	799.49	kg/m3	298.15	Densities and Viscosities of Binary Mixtures of 2-Ethyl-1,1,3,3-tetramethylguanidinium Ionic Liquids with Ethanol and 1-Propanol
rhol	795.46	kg/m3	303.15	Densities and Viscosities of Binary Mixtures of 2-Ethyl-1,1,3,3-tetramethylguanidinium Ionic Liquids with Ethanol and 1-Propanol

rhol	791.39	kg/m3	308.15	Densities and Viscosities of Binary Mixtures of 2-Ethyl-1,1,3,3-tetramethylguanidinium Ionic Liquids with Ethanol and 1-Propanol
rhol	787.29	kg/m3	313.15	Densities and Viscosities of Binary Mixtures of 2-Ethyl-1,1,3,3-tetramethylguanidinium Ionic Liquids with Ethanol and 1-Propanol
rhol	783.14	kg/m3	318.15	Densities and Viscosities of Binary Mixtures of 2-Ethyl-1,1,3,3-tetramethylguanidinium Ionic Liquids with Ethanol and 1-Propanol
rhol	778.93	kg/m3	323.15	Densities and Viscosities of Binary Mixtures of 2-Ethyl-1,1,3,3-tetramethylguanidinium Ionic Liquids with Ethanol and 1-Propanol
rhol	799.70	kg/m3	298.15	Thermodynamics of 1,3-dimethylurea in eight alcohols
rhol	807.55	kg/m3	288.15	Thermophysical Characterization of the Mixtures of the Ionic Liquid 1-Ethyl-3-Methylimidazolium Acetate with 1-Propanol or 2-Propanol
rhol	799.56	kg/m3	298.15	Thermophysical Characterization of the Mixtures of the Ionic Liquid 1-Ethyl-3-Methylimidazolium Acetate with 1-Propanol or 2-Propanol

rhol	791.46	kg/m3	308.15	Thermophysical Characterization of the Mixtures of the Ionic Liquid 1-Ethyl-3-Methylimidazolium Acetate with 1-Propanol or 2-Propanol
rhol	783.20	kg/m3	318.15	Thermophysical Characterization of the Mixtures of the Ionic Liquid 1-Ethyl-3-Methylimidazolium Acetate with 1-Propanol or 2-Propanol
rhol	774.72	kg/m3	328.15	Thermophysical Characterization of the Mixtures of the Ionic Liquid 1-Ethyl-3-Methylimidazolium Acetate with 1-Propanol or 2-Propanol
rhol	765.98	kg/m3	338.15	Thermophysical Characterization of the Mixtures of the Ionic Liquid 1-Ethyl-3-Methylimidazolium Acetate with 1-Propanol or 2-Propanol
rhol	756.91	kg/m3	348.15	Thermophysical Characterization of the Mixtures of the Ionic Liquid 1-Ethyl-3-Methylimidazolium Acetate with 1-Propanol or 2-Propanol
rhol	811.48	kg/m3	283.15	Densities and Viscosities for Binary Liquid Mixtures of n-Undecane + 1-Propanol, + 1-Butanol, + 1-Pentanol, and + 1-Hexanol from 283.15 to 363.15 K at 0.1 MPa
rhol	807.54	kg/m3	288.15	Densities and Viscosities for Binary Liquid Mixtures of n-Undecane + 1-Propanol, + 1-Butanol, + 1-Pentanol, and + 1-Hexanol from 283.15 to 363.15 K at 0.1 MPa

rhol	803.58	kg/m3	293.15	Densities and Viscosities for Binary Liquid Mixtures of n-Undecane + 1-Propanol, + 1-Butanol, + 1-Pentanol, and + 1-Hexanol from 283.15 to 363.15 K at 0.1 MPa
rhol	799.59	kg/m3	298.15	Densities and Viscosities for Binary Liquid Mixtures of n-Undecane + 1-Propanol, + 1-Butanol, + 1-Pentanol, and + 1-Hexanol from 283.15 to 363.15 K at 0.1 MPa
rhol	795.58	kg/m3	303.15	Densities and Viscosities for Binary Liquid Mixtures of n-Undecane + 1-Propanol, + 1-Butanol, + 1-Pentanol, and + 1-Hexanol from 283.15 to 363.15 K at 0.1 MPa
rhol	791.53	kg/m3	308.15	Densities and Viscosities for Binary Liquid Mixtures of n-Undecane + 1-Propanol, + 1-Butanol, + 1-Pentanol, and + 1-Hexanol from 283.15 to 363.15 K at 0.1 MPa
rhol	787.44	kg/m3	313.15	Densities and Viscosities for Binary Liquid Mixtures of n-Undecane + 1-Propanol, + 1-Butanol, + 1-Pentanol, and + 1-Hexanol from 283.15 to 363.15 K at 0.1 MPa

rhol	783.30	kg/m3	318.15	Densities and Viscosities for Binary Liquid Mixtures of n-Undecane + 1-Propanol, + 1-Butanol, + 1-Pentanol, and + 1-Hexanol from 283.15 to 363.15 K at 0.1 MPa
rhol	779.11	kg/m3	323.15	Densities and Viscosities for Binary Liquid Mixtures of n-Undecane + 1-Propanol, + 1-Butanol, + 1-Pentanol, and + 1-Hexanol from 283.15 to 363.15 K at 0.1 MPa
rhol	774.86	kg/m3	328.15	Densities and Viscosities for Binary Liquid Mixtures of n-Undecane + 1-Propanol, + 1-Butanol, + 1-Pentanol, and + 1-Hexanol from 283.15 to 363.15 K at 0.1 MPa
rhol	770.54	kg/m3	333.15	Densities and Viscosities for Binary Liquid Mixtures of n-Undecane + 1-Propanol, + 1-Butanol, + 1-Pentanol, and + 1-Hexanol from 283.15 to 363.15 K at 0.1 MPa
rhol	766.15	kg/m3	338.15	Densities and Viscosities for Binary Liquid Mixtures of n-Undecane + 1-Propanol, + 1-Butanol, + 1-Pentanol, and + 1-Hexanol from 283.15 to 363.15 K at 0.1 MPa

rhol	761.67	kg/m3	343.15	Densities and Viscosities for Binary Liquid Mixtures of n-Undecane + 1-Propanol, + 1-Butanol, + 1-Pentanol, and + 1-Hexanol from 283.15 to 363.15 K at 0.1 MPa	
rhol	757.09	kg/m3	348.15	Densities and Viscosities for Binary Liquid Mixtures of n-Undecane + 1-Propanol, + 1-Butanol, + 1-Pentanol, and + 1-Hexanol from 283.15 to 363.15 K at 0.1 MPa	
rhol	752.44	kg/m3	353.15	Densities and Viscosities for Binary Liquid Mixtures of n-Undecane + 1-Propanol, + 1-Butanol, + 1-Pentanol, and + 1-Hexanol from 283.15 to 363.15 K at 0.1 MPa	
rhol	747.65	kg/m3	358.15	Densities and Viscosities for Binary Liquid Mixtures of n-Undecane + 1-Propanol, + 1-Butanol, + 1-Pentanol, and + 1-Hexanol from 283.15 to 363.15 K at 0.1 MPa	
rhol	742.77	kg/m3	363.15	Densities and Viscosities for Binary Liquid Mixtures of n-Undecane + 1-Propanol, + 1-Butanol, + 1-Pentanol, and + 1-Hexanol from 283.15 to 363.15 K at 0.1 MPa	
rhol	799.69	kg/m3	298.15	Physical Properties of the Pure Deep Eutectic Solvent, [ChCl]:[Lev] (1:2) DES, and Its Binary Mixtures with Alcohols	

rhol	792.76	kg/m3	308.15	Physical Properties of the Pure Deep Eutectic Solvent, [ChCl]:[Lev] (1:2) DES, and Its Binary Mixtures with Alcohols	
rhol	784.45	kg/m3	318.15	Physical Properties of the Pure Deep Eutectic Solvent, [ChCl]:[Lev] (1:2) DES, and Its Binary Mixtures with Alcohols	
rhol	795.54	kg/m3	303.15	Isobaric Vapor Liquid Equilibrium for the Binary Systems Dimethyl Disulfide + C1 C4 n-Alkanol at 40.000 and 101.325 kPa	
rhol	803.67	kg/m3	293.15	Volumetric Properties, Viscosity, and Refractive Indices of Different Naringenin Solutions at Several Temperatures	
rhol	800.94	kg/m3	298.15	Volumetric Properties, Viscosity, and Refractive Indices of Different Naringenin Solutions at Several Temperatures	
rhol	802.00	kg/m3	298.15	Isobaric Vapor-Liquid Equilibrium for the Binary Systems of Sec-butyl Acetate and Ethanol, 1-Propanol, or 2-Propanol at 101.3 kPa	

rhol	807.83	kg/m3	288.15	Measurement and Correlation of Activity, Density, and Speed of Sound for Binary Mixtures of 1-Propanol + Poly(Propylene Glycol) 400, 725, and 1025
rhol	799.60	kg/m3	298.15	Measurement and Correlation of Activity, Density, and Speed of Sound for Binary Mixtures of 1-Propanol + Poly(Propylene Glycol) 400, 725, and 1025
rhol	791.40	kg/m3	308.15	Measurement and Correlation of Activity, Density, and Speed of Sound for Binary Mixtures of 1-Propanol + Poly(Propylene Glycol) 400, 725, and 1025
rhol	783.12	kg/m3	318.15	Measurement and Correlation of Activity, Density, and Speed of Sound for Binary Mixtures of 1-Propanol + Poly(Propylene Glycol) 400, 725, and 1025
rhol	803.60	kg/m3	293.15	Three Binary Azeotropic Systems for 1-(Methoxymethoxy)-propane, 1-(Ethoxymethoxy)-propane, and Methoxy(methoxymethoxy)methane with Three Alcohols at 101.33 kPa: Experimental Data, Correlation, and Purification

rhol	799.50	kg/m3	298.15	Isobaric Vapor-Liquid Phase Equilibrium Measurements, Correlation, and Prediction for Separation of the Mixtures of Cyclohexanone and Alcohols
rhol	807.86	kg/m3	288.15	Densities and Excess Molar Volumes for the Binary and Ternary Systems of (1,4-Dioxane, 1-Propanol or 2-Propanol, and 1,2-Dichloroethane) at T = (288.15 to 318.15) K. Experimental Measurements and Prigogine-Flory-Patterson Modeling
rhol	799.78	kg/m3	298.15	Densities and Excess Molar Volumes for the Binary and Ternary Systems of (1,4-Dioxane, 1-Propanol or 2-Propanol, and 1,2-Dichloroethane) at T = (288.15 to 318.15) K. Experimental Measurements and Prigogine-Flory-Patterson Modeling
rhol	791.71	kg/m3	308.15	Densities and Excess Molar Volumes for the Binary and Ternary Systems of (1,4-Dioxane, 1-Propanol or 2-Propanol, and 1,2-Dichloroethane) at T = (288.15 to 318.15) K. Experimental Measurements and Prigogine-Flory-Patterson Modeling

rhol	783.64	kg/m3	318.15	Densities and Excess Molar Volumes for the Binary and Ternary Systems of (1,4-Dioxane, 1-Propanol or 2-Propanol, and 1,2-Dichloroethane) at T = (288.15 to 318.15) K. Experimental Measurements and Prigogine-Flory-Patterson Modeling
rhol	815.52	kg/m3	278.15	Density, Viscosity, and Speed of Sound of Pure and Binary Mixtures of Ionic Liquids Based on Sulfonium and Imidazolium Cations and Bis(trifluoromethylsulfonyl)imide Anion with 1-Propanol
rhol	811.60	kg/m3	283.15	Density, Viscosity, and Speed of Sound of Pure and Binary Mixtures of Ionic Liquids Based on Sulfonium and Imidazolium Cations and Bis(trifluoromethylsulfonyl)imide Anion with 1-Propanol
rhol	807.63	kg/m3	288.15	Density, Viscosity, and Speed of Sound of Pure and Binary Mixtures of Ionic Liquids Based on Sulfonium and Imidazolium Cations and Bis(trifluoromethylsulfonyl)imide Anion with 1-Propanol

rhol	803.65	kg/m3	293.15	Density, Viscosity, and Speed of Sound of Pure and Binary Mixtures of Ionic Liquids Based on Sulfonium and Imidazolium Cations and Bis(trifluoromethylsulfonyl)imide Anion with 1-Propanol
rhol	799.65	kg/m3	298.15	Density, Viscosity, and Speed of Sound of Pure and Binary Mixtures of Ionic Liquids Based on Sulfonium and Imidazolium Cations and Bis(trifluoromethylsulfonyl)imide Anion with 1-Propanol
rhol	795.61	kg/m3	303.15	Density, Viscosity, and Speed of Sound of Pure and Binary Mixtures of Ionic Liquids Based on Sulfonium and Imidazolium Cations and Bis(trifluoromethylsulfonyl)imide Anion with 1-Propanol
rhol	791.54	kg/m3	308.15	Density, Viscosity, and Speed of Sound of Pure and Binary Mixtures of Ionic Liquids Based on Sulfonium and Imidazolium Cations and Bis(trifluoromethylsulfonyl)imide Anion with 1-Propanol
rhol	787.43	kg/m3	313.15	Density, Viscosity, and Speed of Sound of Pure and Binary Mixtures of Ionic Liquids Based on Sulfonium and Imidazolium Cations and Bis(trifluoromethylsulfonyl)imide Anion with 1-Propanol

rhol	783.27	kg/m3	318.15	Density, Viscosity, and Speed of Sound of Pure and Binary Mixtures of Ionic Liquids Based on Sulfonium and Imidazolium Cations and Bis(trifluoromethylsulfonyl)imide Anion with 1-Propanol
rhol	779.05	kg/m3	323.15	Density, Viscosity, and Speed of Sound of Pure and Binary Mixtures of Ionic Liquids Based on Sulfonium and Imidazolium Cations and Bis(trifluoromethylsulfonyl)imide Anion with 1-Propanol
rhol	774.79	kg/m3	328.15	Density, Viscosity, and Speed of Sound of Pure and Binary Mixtures of Ionic Liquids Based on Sulfonium and Imidazolium Cations and Bis(trifluoromethylsulfonyl)imide Anion with 1-Propanol
rhol	770.44	kg/m3	333.15	Density, Viscosity, and Speed of Sound of Pure and Binary Mixtures of Ionic Liquids Based on Sulfonium and Imidazolium Cations and Bis(trifluoromethylsulfonyl)imide Anion with 1-Propanol

rhol	766.03	kg/m3	338.15	Density, Viscosity, and Speed of Sound of Pure and Binary Mixtures of Ionic Liquids Based on Sulfonium and Imidazolium Cations and Bis(trifluoromethylsulfonyl)imide Anion with 1-Propanol
rhol	761.53	kg/m3	343.15	Density, Viscosity, and Speed of Sound of Pure and Binary Mixtures of Ionic Liquids Based on Sulfonium and Imidazolium Cations and Bis(trifluoromethylsulfonyl)imide Anion with 1-Propanol
rhol	799.60	kg/m3	298.15	Vapor-Liquid Equilibrium, Volumetric, and Compressibility Properties of 1-Propanol + Poly(ethylene glycol) Dimethyl Ether 250 and 500 Binary Mixtures
rhol	791.40	kg/m3	308.15	Vapor-Liquid Equilibrium, Volumetric, and Compressibility Properties of 1-Propanol + Poly(ethylene glycol) Dimethyl Ether 250 and 500 Binary Mixtures
rhol	783.12	kg/m3	318.15	Vapor-Liquid Equilibrium, Volumetric, and Compressibility Properties of 1-Propanol + Poly(ethylene glycol) Dimethyl Ether 250 and 500 Binary Mixtures

rhol	803.50	kg/m3	293.15	Density and Viscosity of 2-Butanol + (1-Propanol, 2-Propanol, or 3-Amino-1-propanol) Mixtures at Temperatures of (293.15 to 323.15) K: Application of the ERAS Model
rhol	799.40	kg/m3	298.15	Density and Viscosity of 2-Butanol + (1-Propanol, 2-Propanol, or 3-Amino-1-propanol) Mixtures at Temperatures of (293.15 to 323.15) K: Application of the ERAS Model
rhol	795.40	kg/m3	303.15	Density and Viscosity of 2-Butanol + (1-Propanol, 2-Propanol, or 3-Amino-1-propanol) Mixtures at Temperatures of (293.15 to 323.15) K: Application of the ERAS Model
rhol	791.30	kg/m3	308.15	Density and Viscosity of 2-Butanol + (1-Propanol, 2-Propanol, or 3-Amino-1-propanol) Mixtures at Temperatures of (293.15 to 323.15) K: Application of the ERAS Model
rhol	787.20	kg/m3	313.15	Density and Viscosity of 2-Butanol + (1-Propanol, 2-Propanol, or 3-Amino-1-propanol) Mixtures at Temperatures of (293.15 to 323.15) K: Application of the ERAS Model

rhol	783.10	kg/m3	318.15	Density and Viscosity of 2-Butanol + (1-Propanol, 2-Propanol, or 3-Amino-1-propanol) Mixtures at Temperatures of (293.15 to 323.15) K: Application of the ERAS Model	
rhol	778.90	kg/m3	323.15	Density and Viscosity of 2-Butanol + (1-Propanol, 2-Propanol, or 3-Amino-1-propanol) Mixtures at Temperatures of (293.15 to 323.15) K: Application of the ERAS Model	
rhol	803.60	kg/m3	293.15	Liquid Densities and Speed of Sound for Ionic Liquid (2-HEAA and 2-HDEAA) + Alcohol (1-Propanol and 2-Propanol) Mixtures at T = (293.15-323.15 K) and Atmospheric Pressure	
rhol	795.70	kg/m3	303.15	Liquid Densities and Speed of Sound for Ionic Liquid (2-HEAA and 2-HDEAA) + Alcohol (1-Propanol and 2-Propanol) Mixtures at T = (293.15-323.15 K) and Atmospheric Pressure	
rhol	787.50	kg/m3	313.15	Liquid Densities and Speed of Sound for Ionic Liquid (2-HEAA and 2-HDEAA) + Alcohol (1-Propanol and 2-Propanol) Mixtures at T = (293.15-323.15 K) and Atmospheric Pressure	

rhol	779.20	kg/m3	323.15	Liquid Densities and Speed of Sound for Ionic Liquid (2-HEAA and 2-HDEAA) + Alcohol (1-Propanol and 2-Propanol) Mixtures at T = (293.15-323.15 K) and Atmospheric Pressure
rhol	803.63	kg/m3	293.15	Excess Molar Enthalpies for Binary Mixtures of Ethanol + Acetone, + Octane, + Cyclohexane and 1-Propanol + Acetone, + Octane, + Heptane at 323.15
rhol	799.50	kg/m3	298.15	Dynamic Viscosities of Diethyl Carbonate with Linear and Secondary Alcohols at Several Temperatures
rhol	799.75	kg/m3	298.15	Measurement and Correlation of Vapor-Liquid Equilibria at T) 333.15 K and Excess Molar Volumes at T) 298.15 K for Ethanol + Dimethyl Carbonate (DMC), DMC + 1-Propanol, and DMC + 1-Butanol
rhol	803.50	kg/m3	293.00	Fluid-Phase Behavior of {1-Hexyl-3-methylimidazolium Bis(trifluoromethylsulfonyl) Imide, [C6mim][NTf2], + C2-C8 n-Alcohol} Mixtures: Liquid-Liquid Equilibrium and Excess Volumes

rhol	803.50	kg/m3	293.00	Fluid-Phase Behavior of {1-Hexyl-3-methylimidazolium Bis(trifluoromethylsulfonyl) Imide, [C6mim][NTf2], + C2-C8 n-Alcohol} Mixtures: Liquid-Liquid Equilibrium and Excess Volumes
rhol	799.50	kg/m3	298.00	Fluid-Phase Behavior of {1-Hexyl-3-methylimidazolium Bis(trifluoromethylsulfonyl) Imide, [C6mim][NTf2], + C2-C8 n-Alcohol} Mixtures: Liquid-Liquid Equilibrium and Excess Volumes
rhol	799.50	kg/m3	298.00	Fluid-Phase Behavior of {1-Hexyl-3-methylimidazolium Bis(trifluoromethylsulfonyl) Imide, [C6mim][NTf2], + C2-C8 n-Alcohol} Mixtures: Liquid-Liquid Equilibrium and Excess Volumes
rhol	795.50	kg/m3	303.00	Fluid-Phase Behavior of {1-Hexyl-3-methylimidazolium Bis(trifluoromethylsulfonyl) Imide, [C6mim][NTf2], + C2-C8 n-Alcohol} Mixtures: Liquid-Liquid Equilibrium and Excess Volumes
rhol	795.50	kg/m3	303.00	Fluid-Phase Behavior of {1-Hexyl-3-methylimidazolium Bis(trifluoromethylsulfonyl) Imide, [C6mim][NTf2], + C2-C8 n-Alcohol} Mixtures: Liquid-Liquid Equilibrium and Excess Volumes

rhol	799.60	kg/m3	298.15	Densities and Excess Molar Volumes for Binary Glycerol + 1-Propanol, + 2-Propanol, + 1,2-Propanediol, and + 1,3-Propanediol Mixtures at Different Temperatures
rhol	795.50	kg/m3	303.15	Densities and Excess Molar Volumes for Binary Glycerol + 1-Propanol, + 2-Propanol, + 1,2-Propanediol, and + 1,3-Propanediol Mixtures at Different Temperatures
rhol	791.50	kg/m3	308.15	Densities and Excess Molar Volumes for Binary Glycerol + 1-Propanol, + 2-Propanol, + 1,2-Propanediol, and + 1,3-Propanediol Mixtures at Different Temperatures
rhol	787.30	kg/m3	313.15	Densities and Excess Molar Volumes for Binary Glycerol + 1-Propanol, + 2-Propanol, + 1,2-Propanediol, and + 1,3-Propanediol Mixtures at Different Temperatures
rhol	783.30	kg/m3	318.15	Densities and Excess Molar Volumes for Binary Glycerol + 1-Propanol, + 2-Propanol, + 1,2-Propanediol, and + 1,3-Propanediol Mixtures at Different Temperatures

rhol	779.00	kg/m3	323.15	Densities and Excess Molar Volumes for Binary Glycerol + 1-Propanol, + 2-Propanol, + 1,2-Propanediol, and + 1,3-Propanediol Mixtures at Different Temperatures
rhol	774.90	kg/m3	328.15	Densities and Excess Molar Volumes for Binary Glycerol + 1-Propanol, + 2-Propanol, + 1,2-Propanediol, and + 1,3-Propanediol Mixtures at Different Temperatures
rhol	770.80	kg/m3	333.15	Densities and Excess Molar Volumes for Binary Glycerol + 1-Propanol, + 2-Propanol, + 1,2-Propanediol, and + 1,3-Propanediol Mixtures at Different Temperatures
rhol	809.01	kg/m3	288.15	Excess Molar Volumes of 1,3-Diethyl Propanedioate with Methanol, Ethanol, Propan-1-ol, Propan-2-ol, Butan-2-ol, 2-Methyl-propan-1-ol, and Pentan-1-ol at T = (288.15, 298.15, 313.15, and 328.15) K
rhol	800.97	kg/m3	298.15	Excess Molar Volumes of 1,3-Diethyl Propanedioate with Methanol, Ethanol, Propan-1-ol, Propan-2-ol, Butan-2-ol, 2-Methyl-propan-1-ol, and Pentan-1-ol at T = (288.15, 298.15, 313.15, and 328.15) K

rhol	788.69	kg/m3	313.15	Excess Molar Volumes of 1,3-Diethyl Propanedioate with Methanol, Ethanol, Propan-1-ol, Propan-2-ol, Butan-2-ol, 2-Methyl-propan-1-ol, and Pentan-1-ol at T = (288.15, 298.15, 313.15, and 328.15) K
rhol	775.97	kg/m3	328.15	Excess Molar Volumes of 1,3-Diethyl Propanedioate with Methanol, Ethanol, Propan-1-ol, Propan-2-ol, Butan-2-ol, 2-Methyl-propan-1-ol, and Pentan-1-ol at T = (288.15, 298.15, 313.15, and 328.15) K
rhol	797.60	kg/m3	298.15	Volumetric Properties for (Ionic Liquid + Methanol or Ethanol or 1-Propanol + Nitromethane) at 298.15 K and Atmospheric Pressure
rhol	804.27	kg/m3	298.15	Phase Equilibria for Reactive Distillation of Propyl Propanoate. Pure Component Property Data, Vapor-Liquid Equilibria, and Liquid-Liquid Equilibria
rhol	799.56	kg/m3	298.15	ACSExcess Molar Enthalpies of Mixtures of (+-)-Linalool with Several Alkanols

rhol	799.70	kg/m3	298.15	Excess Volumes of Ternary Mixtures 2,2,4-Trimethylpentane + Diisopropyl Ether or Methyl tert-Butyl Ether + Methanol, Ethanol, or 1-Propanol at 298.15 K
rhol	799.67	kg/m3	298.15	Thermophysical Properties of the Pure Ionic Liquid 1-Butyl-1-methylpyrrolidinium Dicyanamide and Its Binary Mixtures with Alcohols
rhol	803.57	kg/m3	293.15	Density and Viscosity Measurements of Binary Alkanol Mixtures from (293.15 to 333.15) K at Atmospheric Pressure
rhol	795.53	kg/m3	303.15	Density and Viscosity Measurements of Binary Alkanol Mixtures from (293.15 to 333.15) K at Atmospheric Pressure
rhol	787.35	kg/m3	313.15	Density and Viscosity Measurements of Binary Alkanol Mixtures from (293.15 to 333.15) K at Atmospheric Pressure
rhol	778.98	kg/m3	323.15	Density and Viscosity Measurements of Binary Alkanol Mixtures from (293.15 to 333.15) K at Atmospheric Pressure
rhol	770.38	kg/m3	333.15	Density and Viscosity Measurements of Binary Alkanol Mixtures from (293.15 to 333.15) K at Atmospheric Pressure

rhol	803.84	kg/m3	293.15	Volumetric and Transport Properties of Binary Mixtures of n-Octane + Ethanol, + 1-Propanol, + 1-Butanol, and + 1-Pentanol from (293.15 to 323.15) K at Atmospheric Pressure	
rhol	799.85	kg/m3	298.15	Volumetric and Transport Properties of Binary Mixtures of n-Octane + Ethanol, + 1-Propanol, + 1-Butanol, and + 1-Pentanol from (293.15 to 323.15) K at Atmospheric Pressure	
rhol	795.81	kg/m3	303.15	Volumetric and Transport Properties of Binary Mixtures of n-Octane + Ethanol, + 1-Propanol, + 1-Butanol, and + 1-Pentanol from (293.15 to 323.15) K at Atmospheric Pressure	
rhol	791.74	kg/m3	308.15	Volumetric and Transport Properties of Binary Mixtures of n-Octane + Ethanol, + 1-Propanol, + 1-Butanol, and + 1-Pentanol from (293.15 to 323.15) K at Atmospheric Pressure	

rhol	787.64	kg/m3	313.15	Volumetric and Transport Properties of Binary Mixtures of n-Octane + Ethanol, + 1-Propanol, + 1-Butanol, and + 1-Pentanol from (293.15 to 323.15) K at Atmospheric Pressure
rhol	783.55	kg/m3	318.15	Volumetric and Transport Properties of Binary Mixtures of n-Octane + Ethanol, + 1-Propanol, + 1-Butanol, and + 1-Pentanol from (293.15 to 323.15) K at Atmospheric Pressure
rhol	779.32	kg/m3	323.15	Volumetric and Transport Properties of Binary Mixtures of n-Octane + Ethanol, + 1-Propanol, + 1-Butanol, and + 1-Pentanol from (293.15 to 323.15) K at Atmospheric Pressure
rhol	804.00	kg/m3	293.15	Composition and Temperature Dependence of Density, Surface Tension, and Viscosity of EMIM DEP/MMIM DMP + Water + 1-Propanol/2-Propanol Ternary Mixtures and Their Mathematical Representation Using the Jouyban Acree Model

rhol	800.10	kg/m3	298.15	Composition and Temperature Dependence of Density, Surface Tension, and Viscosity of EMIM DEP/MMIM DMP + Water + 1-Propanol/2-Propanol Ternary Mixtures and Their Mathematical Representation Using the Jouyban Acree Model
rhol	796.00	kg/m3	303.15	Composition and Temperature Dependence of Density, Surface Tension, and Viscosity of EMIM DEP/MMIM DMP + Water + 1-Propanol/2-Propanol Ternary Mixtures and Their Mathematical Representation Using the Jouyban Acree Model
rhol	788.00	kg/m3	313.15	Composition and Temperature Dependence of Density, Surface Tension, and Viscosity of EMIM DEP/MMIM DMP + Water + 1-Propanol/2-Propanol Ternary Mixtures and Their Mathematical Representation Using the Jouyban Acree Model

rhol	780.00	kg/m3	323.15	Composition and Temperature Dependence of Density, Surface Tension, and Viscosity of EMIM DEP/MMIM DMP + Water + 1-Propanol/2-Propanol Ternary Mixtures and Their Mathematical Representation Using the Jouyban Acree Model
rhol	764.00	kg/m3	333.15	Composition and Temperature Dependence of Density, Surface Tension, and Viscosity of EMIM DEP/MMIM DMP + Water + 1-Propanol/2-Propanol Ternary Mixtures and Their Mathematical Representation Using the Jouyban Acree Model
rhol	803.86	kg/m3	293.15	Densities and Viscosities of Diaminotoluene with Water, Ethanol, Propan-1-ol, and Butan-1-ol from (293.15 to 333.15) K
rhol	795.54	kg/m3	303.15	Densities and Viscosities of Diaminotoluene with Water, Ethanol, Propan-1-ol, and Butan-1-ol from (293.15 to 333.15) K
rhol	787.37	kg/m3	313.15	Densities and Viscosities of Diaminotoluene with Water, Ethanol, Propan-1-ol, and Butan-1-ol from (293.15 to 333.15) K

rhol	779.01	kg/m3	323.15	Densities and Viscosities of Diaminotoluene with Water, Ethanol, Propan-1-ol, and Butan-1-ol from (293.15 to 333.15) K	
rhol	770.42	kg/m3	333.15	Densities and Viscosities of Diaminotoluene with Water, Ethanol, Propan-1-ol, and Butan-1-ol from (293.15 to 333.15) K	
rhol	803.52	kg/m3	293.15	Measurement and Correlation of the Excess Properties of Ternary Mixture of {x1[Hmim][BF4] + x21-Propanol + x32-Propanol} at Different Temperatures	
rhol	795.47	kg/m3	303.15	Measurement and Correlation of the Excess Properties of Ternary Mixture of {x1[Hmim][BF4] + x21-Propanol + x32-Propanol} at Different Temperatures	
rhol	787.30	kg/m3	313.15	Measurement and Correlation of the Excess Properties of Ternary Mixture of {x1[Hmim][BF4] + x21-Propanol + x32-Propanol} at Different Temperatures	
rhol	778.93	kg/m3	323.15	Measurement and Correlation of the Excess Properties of Ternary Mixture of {x1[Hmim][BF4] + x21-Propanol + x32-Propanol} at Different Temperatures	

rhol	770.32	kg/m3	333.15	Measurement and Correlation of the Excess Properties of Ternary Mixture of {x1[Hmim][BF4] + x21-Propanol + x32-Propanol} at Different Temperatures
rhol	803.65	kg/m3	293.15	Densities and Viscosities of Binary Mixtures Containing 1,3-Dimethylimidazolium Dimethylphosphate and Alcohols
rhol	799.67	kg/m3	298.15	Densities and Viscosities of Binary Mixtures Containing 1,3-Dimethylimidazolium Dimethylphosphate and Alcohols
rhol	795.87	kg/m3	303.15	Densities and Viscosities of Binary Mixtures Containing 1,3-Dimethylimidazolium Dimethylphosphate and Alcohols
rhol	791.83	kg/m3	308.15	Densities and Viscosities of Binary Mixtures Containing 1,3-Dimethylimidazolium Dimethylphosphate and Alcohols
rhol	787.72	kg/m3	313.15	Densities and Viscosities of Binary Mixtures Containing 1,3-Dimethylimidazolium Dimethylphosphate and Alcohols
rhol	783.56	kg/m3	318.15	Densities and Viscosities of Binary Mixtures Containing 1,3-Dimethylimidazolium Dimethylphosphate and Alcohols
rhol	779.36	kg/m3	323.15	Densities and Viscosities of Binary Mixtures Containing 1,3-Dimethylimidazolium Dimethylphosphate and Alcohols

rhol	807.75	kg/m3	288.15	Volumetric and Viscometric Study of Binary Systems of Ethyl Butyrate with Alcohols
rhol	803.77	kg/m3	293.15	Volumetric and Viscometric Study of Binary Systems of Ethyl Butyrate with Alcohols
rhol	799.76	kg/m3	298.15	Volumetric and Viscometric Study of Binary Systems of Ethyl Butyrate with Alcohols
rhol	795.72	kg/m3	303.15	Volumetric and Viscometric Study of Binary Systems of Ethyl Butyrate with Alcohols
rhol	791.64	kg/m3	308.15	Volumetric and Viscometric Study of Binary Systems of Ethyl Butyrate with Alcohols
rhol	787.52	kg/m3	313.15	Volumetric and Viscometric Study of Binary Systems of Ethyl Butyrate with Alcohols
rhol	783.36	kg/m3	318.15	Volumetric and Viscometric Study of Binary Systems of Ethyl Butyrate with Alcohols
rhol	779.14	kg/m3	323.15	Volumetric and Viscometric Study of Binary Systems of Ethyl Butyrate with Alcohols
rhol	803.48	kg/m3	293.15	Speed of sound, density and related thermodynamic excess properties of binary mixtures of butan-2-one with C1-C4 nalkanols and chloroform

rhol	799.47	kg/m3	298.15	Speed of sound, density and related thermodynamic excess properties of binary mixtures of butan-2-one with C1-C4 nalkanols and chloroform
rhol	795.44	kg/m3	303.15	Speed of sound, density and related thermodynamic excess properties of binary mixtures of butan-2-one with C1-C4 nalkanols and chloroform
rhol	791.37	kg/m3	308.15	Speed of sound, density and related thermodynamic excess properties of binary mixtures of butan-2-one with C1-C4 nalkanols and chloroform
rhol	787.26	kg/m3	313.15	Speed of sound, density and related thermodynamic excess properties of binary mixtures of butan-2-one with C1-C4 nalkanols and chloroform
rhol	783.10	kg/m3	318.15	Speed of sound, density and related thermodynamic excess properties of binary mixtures of butan-2-one with C1-C4 nalkanols and chloroform
rhol	778.90	kg/m3	323.15	Speed of sound, density and related thermodynamic excess properties of binary mixtures of butan-2-one with C1-C4 nalkanols and chloroform

rhol	804.97	kg/m3	293.15	Densities of Ionic Liquids, 1-Butyl-3-methylimidazolium Hexafluorophosphate and 1-Butyl-3-methylimidazolium Tetrafluoroborate, with Benzene, Acetonitrile, and 1-Propanol at T = (293.15 to 343.15) K
rhol	801.01	kg/m3	298.15	Densities of Ionic Liquids, 1-Butyl-3-methylimidazolium Hexafluorophosphate and 1-Butyl-3-methylimidazolium Tetrafluoroborate, with Benzene, Acetonitrile, and 1-Propanol at T = (293.15 to 343.15) K
rhol	796.89	kg/m3	303.15	Densities of Ionic Liquids, 1-Butyl-3-methylimidazolium Hexafluorophosphate and 1-Butyl-3-methylimidazolium Tetrafluoroborate, with Benzene, Acetonitrile, and 1-Propanol at T = (293.15 to 343.15) K
rhol	792.83	kg/m3	308.15	Densities of Ionic Liquids, 1-Butyl-3-methylimidazolium Hexafluorophosphate and 1-Butyl-3-methylimidazolium Tetrafluoroborate, with Benzene, Acetonitrile, and 1-Propanol at T = (293.15 to 343.15) K
rhol	789.07	kg/m3	313.15	Densities of Ionic Liquids, 1-Butyl-3-methylimidazolium Hexafluorophosphate and 1-Butyl-3-methylimidazolium Tetrafluoroborate, with Benzene, Acetonitrile, and 1-Propanol at T = (293.15 to 343.15) K

rhol	780.74	kg/m3	323.15	Densities of Ionic Liquids, 1-Butyl-3-methylimidazolium Hexafluorophosphate and 1-Butyl-3-methylimidazolium Tetrafluoroborate, with Benzene, Acetonitrile, and 1-Propanol at T = (293.15 to 343.15) K
rhol	771.75	kg/m3	333.15	Densities of Ionic Liquids, 1-Butyl-3-methylimidazolium Hexafluorophosphate and 1-Butyl-3-methylimidazolium Tetrafluoroborate, with Benzene, Acetonitrile, and 1-Propanol at T = (293.15 to 343.15) K
rhol	774.73	kg/m3	328.15	Volumetric properties of the boldine + alcohol mixtures at atmospheric pressure from 283.15 to 333.15K A new method for the determination of the density of pure boldine
rhol	799.70	kg/m3	298.15	Thermodynamic Properties of 1-Butyl-3-methylpyridinium Tetrafluoroborate and Its Mixtures with Water and Alkanols
rhol	783.16	kg/m3	318.15	Thermodynamic Properties of 1-Butyl-3-methylpyridinium Tetrafluoroborate and Its Mixtures with Water and Alkanols
rhol	796.55	kg/m3	303.15	Isothermal VLE and VE at 303.15 K for the Binary and Ternary Mixtures of Di-isopropyl Ether (DIPE) + 1-Propanol + 2,2,4-Trimethylpentane

rhol	799.52	kg/m3	298.15	Densities and Excess Molar Volumes of N-Methylmorpholine + 1-Alkanol Systems at 298.15 K
rhol	799.94	kg/m3	298.15	Liquid Liquid Equilibrium for Ternary Systems of Propyl Vinyl Ether + C3 or C4 Alcohols + Water at 298.15 K and Excess Molar Enthalpies for Ternary and Constituent Binary Systems of Propyl Vinyl Ether + Ethanol + Isooctane at 303.15 K
rhol	803.59	kg/m3	293.15	Volumetric Properties of Binary and Ternary Liquid Mixtures of 1-Propanol (1) + 2-Propanol (2) + Water (3) at Different Temperatures and Ambient Pressure (81.5 kPa)
rhol	795.53	kg/m3	303.15	Volumetric Properties of Binary and Ternary Liquid Mixtures of 1-Propanol (1) + 2-Propanol (2) + Water (3) at Different Temperatures and Ambient Pressure (81.5 kPa)
rhol	787.38	kg/m3	313.15	Volumetric Properties of Binary and Ternary Liquid Mixtures of 1-Propanol (1) + 2-Propanol (2) + Water (3) at Different Temperatures and Ambient Pressure (81.5 kPa)

rhol	778.92	kg/m3	323.15	Volumetric Properties of Binary and Ternary Liquid Mixtures of 1-Propanol (1) + 2-Propanol (2) + Water (3) at Different Temperatures and Ambient Pressure (81.5 kPa)	
rhol	799.98	kg/m3	298.15	Binary Liquid-Liquid Equilibrium (LLE) for Dibutyl Ether (DBE) + Water from (288.15 to 318.15) K and Ternary LLE for Systems of DBE + C1 !less thanless than C4 Alcohols + Water at 298.15 K	
rhol	799.89	kg/m3	298.15	Binary Liquid-Liquid Equilibrium (LLE) for Methyl tert-Amyl Ether (TAME) + Water from (288.15 to 313.15) K and Ternary LLE for Systems of TAME + C1-C4 Alcohols + Water at 298.15 K	
rhol	799.81	kg/m3	298.15	Densities and Excess Molar Volumes of Cyclopentane (1) + 1-Alkanol (2) Systems at (298.15 and 308.15) K	
rhol	791.69	kg/m3	308.15	Densities and Excess Molar Volumes of Cyclopentane (1) + 1-Alkanol (2) Systems at (298.15 and 308.15) K	

rhol	800.61	kg/m3	298.15	Binary and Ternary Vapor-Liquid Equilibrium at 323.15 K and Excess Molar Volumes at 298.15 K for the Mixtures of Propyl Vinyl Ether + 1-Propanol + Toluene
rhol	799.70	kg/m3	298.15	Solubility and Liquid-Liquid Equilibrium of Aqueous Systems of Iodoethane with Methanol, Ethanol, or 1-Propanol at Temperature 298.15 K and Pressure 101.2 kPa
rhol	779.30	kg/m3	323.15	Liquid-Liquid Equilibria, Density, Viscosity, and Surface and Interfacial Tension of the System Water + n-Butyl Acetate + 1-Propanol at 323.15 K and Atmospheric Pressure
rhol	800.70	kg/m3	298.15	Apparent Molal Volumes and Viscosity B-Coefficients of Acetyl Salicylic Acid (2-Acetoxy Benzoic Acid) Solutions in Higher Alcohols at Different Temperatures
rhol	799.56	kg/m3	298.15	Refractive Indices and Deviations in Refractive Indices for Binary Mixtures of 1-Ethyl-3-methylimidazolium Trifluoromethanesulfonate with Methanol, Ethanol, 1-Propanol, and 2-Propanol at Several Temperatures

rhol	804.00	kg/m3	293.00	KDB
rhol	800.62	kg/m3	298.15	Speeds of Sound and Isentropic Compressibilities in Binary Mixtures of 2-Propanol with Several 1-Alkanols at 298.15K
rhol	803.71	kg/m3	293.15	Densities and Volumetric Properties of Binary Mixtures of Aniline with 1-Propanol, 2-Propanol, 2-Methyl-1-Propanol, and 2-Methyl-2-Propanol at Temperatures from 293.15 to 318.15 K
rhol	799.74	kg/m3	298.15	Densities and Volumetric Properties of Binary Mixtures of Aniline with 1-Propanol, 2-Propanol, 2-Methyl-1-Propanol, and 2-Methyl-2-Propanol at Temperatures from 293.15 to 318.15 K
rhol	795.77	kg/m3	303.15	Densities and Volumetric Properties of Binary Mixtures of Aniline with 1-Propanol, 2-Propanol, 2-Methyl-1-Propanol, and 2-Methyl-2-Propanol at Temperatures from 293.15 to 318.15 K
rhol	791.80	kg/m3	308.15	Densities and Volumetric Properties of Binary Mixtures of Aniline with 1-Propanol, 2-Propanol, 2-Methyl-1-Propanol, and 2-Methyl-2-Propanol at Temperatures from 293.15 to 318.15 K

rhol	787.83	kg/m3	313.15	Densities and Volumetric Properties of Binary Mixtures of Aniline with 1-Propanol, 2-Propanol, 2-Methyl-1-Propanol, and 2-Methyl-2-Propanol at Temperatures from 293.15 to 318.15 K
rhol	783.86	kg/m3	318.15	Densities and Volumetric Properties of Binary Mixtures of Aniline with 1-Propanol, 2-Propanol, 2-Methyl-1-Propanol, and 2-Methyl-2-Propanol at Temperatures from 293.15 to 318.15 K
rhol	799.28	kg/m3	298.15	Topological Investigations of Excess Molar Volumes and Excess Isentropic Compressibilities of Ternary Mixtures Containing Pyrrolidin-2-one at 308.15 K
rhol	799.60	kg/m3	298.15	Experimental Determination of Densities and Isobaric Vapor Liquid Equilibria of Methyl Acetate and Ethyl Acetate with Alcohols (C3 and C4) at 0.3 MPa
rhol	799.51	kg/m3	298.15	Fluid Phase Topology of Benzene + Cyclohexane + 1-Propanol at 101.3 kPa
rhol	795.80	kg/m3	303.15	Viscous synergy and antagonism and isentropic compressibility of ternary mixtures containing 1,3-dioxolane, water and monoalkanols at 303.15K

rhol	811.57	kg/m3	283.15	Volumetric properties of the boldine + alcohol mixtures at atmospheric pressure from 283.15 to 333.15K A new method for the determination of the density of pure boldine	
rhol	807.62	kg/m3	288.15	Volumetric properties of the boldine + alcohol mixtures at atmospheric pressure from 283.15 to 333.15K A new method for the determination of the density of pure boldine	
rhol	803.63	kg/m3	293.15	Volumetric properties of the boldine + alcohol mixtures at atmospheric pressure from 283.15 to 333.15K A new method for the determination of the density of pure boldine	
rhol	799.62	kg/m3	298.15	Volumetric properties of the boldine + alcohol mixtures at atmospheric pressure from 283.15 to 333.15K A new method for the determination of the density of pure boldine	
rhol	795.58	kg/m3	303.15	Volumetric properties of the boldine + alcohol mixtures at atmospheric pressure from 283.15 to 333.15K A new method for the determination of the density of pure boldine	

rhol	791.49	kg/m3	308.15	Volumetric properties of the boldine + alcohol mixtures at atmospheric pressure from 283.15 to 333.15K A new method for the determination of the density of pure boldine	
rhol	787.37	kg/m3	313.15	Volumetric properties of the boldine + alcohol mixtures at atmospheric pressure from 283.15 to 333.15K A new method for the determination of the density of pure boldine	
rhol	783.22	kg/m3	318.15	Volumetric properties of the boldine + alcohol mixtures at atmospheric pressure from 283.15 to 333.15K A new method for the determination of the density of pure boldine	
rhol	799.46	kg/m3	298.15	Thermophysical and volumetric study of mixtures {p-cymene + propan-1-ol} at several temperatures and atmospheric pressure. Modeling with COSMO-RS	
rhol	787.73	kg/m3	313.15	Temperature and composition dependence of the volumetric and acoustic properties of ionic liquid [emim][HSO4] with polar protic and aprotic co-solvents	

rhol	791.82	kg/m3	308.15	Temperature and composition dependence of the volumetric and acoustic properties of ionic liquid [emim][HSO4] with polar protic and aprotic co-solvents
rhol	795.92	kg/m3	303.15	Temperature and composition dependence of the volumetric and acoustic properties of ionic liquid [emim][HSO4] with polar protic and aprotic co-solvents
rhol	799.96	kg/m3	298.15	Temperature and composition dependence of the volumetric and acoustic properties of ionic liquid [emim][HSO4] with polar protic and aprotic co-solvents
rhol	766.20	kg/m3	338.15	Thermophysical properties of binary mixtures of 1-butyl-1-methylpyrrolidinium trifluoromethanesulfonate ionic liquid with alcohols at several temperatures
rhol	774.90	kg/m3	328.15	Thermophysical properties of binary mixtures of 1-butyl-1-methylpyrrolidinium trifluoromethanesulfonate ionic liquid with alcohols at several temperatures
rhol	783.30	kg/m3	318.15	Thermophysical properties of binary mixtures of 1-butyl-1-methylpyrrolidinium trifluoromethanesulfonate ionic liquid with alcohols at several temperatures

rhol	791.60	kg/m3	308.15	Thermophysical properties of binary mixtures of 1-butyl-1-methylpyrrolidinium trifluoromethanesulfonate ionic liquid with alcohols at several temperatures
rhol	799.70	kg/m3	298.15	Thermophysical properties of binary mixtures of 1-butyl-1-methylpyrrolidinium trifluoromethanesulfonate ionic liquid with alcohols at several temperatures
rhol	807.70	kg/m3	288.15	Thermophysical properties of binary mixtures of 1-butyl-1-methylpyrrolidinium trifluoromethanesulfonate ionic liquid with alcohols at several temperatures
rhol	815.60	kg/m3	278.15	Thermophysical properties of binary mixtures of 1-butyl-1-methylpyrrolidinium trifluoromethanesulfonate ionic liquid with alcohols at several temperatures
rhol	783.20	kg/m3	318.15	Excess molar volume and excess Gibbs energy of activation for viscous flow for the binary mixtures of N-ethylpyridinium dicyanamide [C2py][DCA] with alcohols
rhol	787.36	kg/m3	313.15	Excess molar volume and excess Gibbs energy of activation for viscous flow for the binary mixtures of N-ethylpyridinium dicyanamide [C2py][DCA] with alcohols

rhol	791.47	kg/m3	308.15	Excess molar volume and excess Gibbs energy of activation for viscous flow for the binary mixtures of N-ethylpyridinium dicyanamide [C2py][DCA] with alcohols	
rhol	795.55	kg/m3	303.15	Excess molar volume and excess Gibbs energy of activation for viscous flow for the binary mixtures of N-ethylpyridinium dicyanamide [C2py][DCA] with alcohols	
rhol	799.58	kg/m3	298.15	Excess molar volume and excess Gibbs energy of activation for viscous flow for the binary mixtures of N-ethylpyridinium dicyanamide [C2py][DCA] with alcohols	
rhol	803.59	kg/m3	293.15	Excess molar volume and excess Gibbs energy of activation for viscous flow for the binary mixtures of N-ethylpyridinium dicyanamide [C2py][DCA] with alcohols	
rhol	807.55	kg/m3	288.15	Excess molar volume and excess Gibbs energy of activation for viscous flow for the binary mixtures of N-ethylpyridinium dicyanamide [C2py][DCA] with alcohols	

rhol	803.70	kg/m3	293.15	Measurement and correlation of (vapour-liquid) equilibrium for binary mixtures composed of 1-(ethoxymethoxy)-propane with ethanol and 1-propanol at 101.33 kPa
rhol	789.00	kg/m3	313.00	Comparative study of physical properties of binary mixtures of halogen free ionic liquids with alcohols
rhol	792.00	kg/m3	308.00	Comparative study of physical properties of binary mixtures of halogen free ionic liquids with alcohols
rhol	796.00	kg/m3	303.00	Comparative study of physical properties of binary mixtures of halogen free ionic liquids with alcohols
rhol	800.00	kg/m3	298.00	Comparative study of physical properties of binary mixtures of halogen free ionic liquids with alcohols
rhol	799.92	kg/m3	298.20	A green process for recovery of 1-propanol/2-propanol from their aqueous solutions: Experimental and MD simulation studies
rhol	799.50	kg/m3	298.15	Fully automatized apparatus for determining speed of sound for liquids in the temperature and pressure interval (283.15-343.15) K and (0.1-95) MPa

rhol	787.41	kg/m3	313.15	Thermodynamic and spectroscopic properties of binary mixtures of n-butylammonium butanoate ionic liquid with alcohols at T = (293.15-313.15) K	
rhol	791.51	kg/m3	308.15	Thermodynamic and spectroscopic properties of binary mixtures of n-butylammonium butanoate ionic liquid with alcohols at T = (293.15-313.15) K	
rhol	795.58	kg/m3	303.15	Thermodynamic and spectroscopic properties of binary mixtures of n-butylammonium butanoate ionic liquid with alcohols at T = (293.15-313.15) K	
rhol	799.62	kg/m3	298.15	Thermodynamic and spectroscopic properties of binary mixtures of n-butylammonium butanoate ionic liquid with alcohols at T = (293.15-313.15) K	
rhol	803.62	kg/m3	293.15	Thermodynamic and spectroscopic properties of binary mixtures of n-butylammonium butanoate ionic liquid with alcohols at T = (293.15-313.15) K	

rhol	799.61	kg/m3	298.15	Experimental study on the calorimetric data of 2-butoxyethanol with aliphatic alcohols (C1-C4) and correlation with the Wilson, NRTL and UNIQUAC models at T = 298 K
rhol	803.85	kg/m3	293.15	(Liquid + liquid) equilibria of four alcohol-water systems containing 1,8-cineole at T = 298.15 K
rhol	799.70	kg/m3	298.15	Solubility and solution thermodynamics of thymol in six pure organic solvents
rhol	799.56	kg/m3	298.15	Measurements and equation-of-state modelling of thermodynamic properties of binary mixtures of 1-butyl-1-methylpyrrolidinium tetracyanoborate ionic liquid with molecular compounds
rhol	799.56	kg/m3	298.15	Excess molar enthalpies of R-fenchone + propan-1-ol or +propan-2-ol. Modeling with COSMO-RS and UNIFAC
rhol	787.42	kg/m3	313.15	Properties of pure 1,1,3,3-tetramethylguanidine imidazole ionic liquid and its binary mixtures with alcohols at T = (293.15 to 313.15) K

rhol	791.56	kg/m3	308.15	Properties of pure 1,1,3,3-tetramethylguanidine imidazole ionic liquid and its binary mixtures with alcohols at T = (293.15 to 313.15) K
rhol	795.65	kg/m3	303.15	Properties of pure 1,1,3,3-tetramethylguanidine imidazole ionic liquid and its binary mixtures with alcohols at T = (293.15 to 313.15) K
rhol	799.71	kg/m3	298.15	Properties of pure 1,1,3,3-tetramethylguanidine imidazole ionic liquid and its binary mixtures with alcohols at T = (293.15 to 313.15) K
rhol	803.75	kg/m3	293.15	Properties of pure 1,1,3,3-tetramethylguanidine imidazole ionic liquid and its binary mixtures with alcohols at T = (293.15 to 313.15) K
rhol	803.85	kg/m3	293.15	(Liquid + liquid) equilibria for (water + 1-propanol or acetone + .beta.-citronellol) at different temperatures
rhol	799.56	kg/m3	298.15	A combined experimental and computational investigation of excess molar enthalpies of (nitrobenzene + alkanol) mixtures
rhol	787.88	kg/m3	313.15	Molecular interactions in binary mixtures of 1-butoxy-2-propanol with alcohols at different temperatures: A thermophysical and spectroscopic approach

rhol	792.00	kg/m3	308.15	Molecular interactions in binary mixtures of 1-butoxy-2-propanol with alcohols at different temperatures: A thermophysical and spectroscopic approach
rhol	796.08	kg/m3	303.15	Molecular interactions in binary mixtures of 1-butoxy-2-propanol with alcohols at different temperatures: A thermophysical and spectroscopic approach
rhol	800.13	kg/m3	298.15	Molecular interactions in binary mixtures of 1-butoxy-2-propanol with alcohols at different temperatures: A thermophysical and spectroscopic approach
rhol	804.14	kg/m3	293.15	Molecular interactions in binary mixtures of 1-butoxy-2-propanol with alcohols at different temperatures: A thermophysical and spectroscopic approach
rhol	799.52	kg/m3	298.15	Extraction desulfurization process of fuels with ionic liquids
rhol	799.52	kg/m3	298.15	Effect of the alkyl side chain of the 1-alkylpiperidinium-based ionic liquids on desulfurization of fuels

rhol	779.09	kg/m3	323.15	Osmotic coefficients and apparent molar volumes of 1-hexyl-3-methylimidazolium trifluoromethanesulfonate ionic liquid in alcohols
rhol	799.52	kg/m3	298.15	Separation of sulfur compounds from alkanes with 1-alkylcyanopyridinium-based ionic liquids
rhol	799.66	kg/m3	298.15	Thermodynamics of (ketone + amine) mixtures. Part XI. Excess molar enthalpies at T = 298.15 K for the (1-propanol + N,N,N-triethylamine + 2-butanone) system
rhol	798.00	kg/m3	303.15	Study of molecular interactions in the mixtures of some primary alcohols with equimolar mixture of 1-propanol and alkylbenzoates at T = 303.15 K
rhol	791.78	kg/m3	308.15	Physical properties of the pure 1-methyl-1-propylpyrrolidinium bis(trifluoromethylsulfonyl)imide ionic liquid and its binary mixtures with alcohols
rhol	795.94	kg/m3	303.15	Physical properties of the pure 1-methyl-1-propylpyrrolidinium bis(trifluoromethylsulfonyl)imide ionic liquid and its binary mixtures with alcohols
rhol	799.87	kg/m3	298.15	Physical properties of the pure 1-methyl-1-propylpyrrolidinium bis(trifluoromethylsulfonyl)imide ionic liquid and its binary mixtures with alcohols

rhol	791.65	kg/m3	308.15	FT-IR studies on excess thermodynamic properties of binary liquid mixtures o-chlorotoluene with 1-propanol, 1-butanol, 1-pentanol, 1-hexanol and 1-heptanol at different temperatures
rhol	795.65	kg/m3	303.15	FT-IR studies on excess thermodynamic properties of binary liquid mixtures o-chlorotoluene with 1-propanol, 1-butanol, 1-pentanol, 1-hexanol and 1-heptanol at different temperatures
rhol	799.65	kg/m3	298.15	FT-IR studies on excess thermodynamic properties of binary liquid mixtures o-chlorotoluene with 1-propanol, 1-butanol, 1-pentanol, 1-hexanol and 1-heptanol at different temperatures
rhol	779.33	kg/m3	323.15	Osmotic and apparent molar properties of binary mixtures alcohol + 1-butyl-3-methylimidazolium trifluoromethanesulfonate ionic liquid
rhol	791.84	kg/m3	308.15	Osmotic and apparent molar properties of binary mixtures alcohol + 1-butyl-3-methylimidazolium trifluoromethanesulfonate ionic liquid

rhol	803.92	kg/m3	293.15	Osmotic and apparent molar properties of binary mixtures alcohol + 1-butyl-3-methylimidazolium trifluoromethanesulfonate ionic liquid	
rhol	783.16	kg/m3	318.15	Volume effects for binary mixtures of propane-1,2-diol with methanol, propan-1-ol, hexan-1-ol, octan-1-ol, or nonan-1-ol at temperatures (293.15 to 318.15) K	
rhol	787.32	kg/m3	313.15	Volume effects for binary mixtures of propane-1,2-diol with methanol, propan-1-ol, hexan-1-ol, octan-1-ol, or nonan-1-ol at temperatures (293.15 to 318.15) K	
rhol	791.43	kg/m3	308.15	Volume effects for binary mixtures of propane-1,2-diol with methanol, propan-1-ol, hexan-1-ol, octan-1-ol, or nonan-1-ol at temperatures (293.15 to 318.15) K	
rhol	795.49	kg/m3	303.15	Volume effects for binary mixtures of propane-1,2-diol with methanol, propan-1-ol, hexan-1-ol, octan-1-ol, or nonan-1-ol at temperatures (293.15 to 318.15) K	

rhol	799.53	kg/m3	298.15	Volume effects for binary mixtures of propane-1,2-diol with methanol, propan-1-ol, hexan-1-ol, octan-1-ol, or nonan-1-ol at temperatures (293.15 to 318.15) K
rhol	803.53	kg/m3	293.15	Volume effects for binary mixtures of propane-1,2-diol with methanol, propan-1-ol, hexan-1-ol, octan-1-ol, or nonan-1-ol at temperatures (293.15 to 318.15) K
rhol	799.45	kg/m3	298.15	Apparent molar volumes and compressibilities of tetrabutyl-ammonium bromide in organic solvents
rhol	788.60	kg/m3	313.15	Excess molar volumes and isentropic compressibility of binary systems {trioctylmethylammonium bis(trifluoromethysulfonyl)imide + methanol or ethanol or 1-propanol} at different temperatures
rhol	796.40	kg/m3	303.15	Excess molar volumes and isentropic compressibility of binary systems {trioctylmethylammonium bis(trifluoromethysulfonyl)imide + methanol or ethanol or 1-propanol} at different temperatures

rhol	799.40	kg/m3	298.15	Excess molar volumes and isentropic compressibility of binary systems {trioctylmethylammonium bis(trifluoromethysulfonyl)imide + methanol or ethanol or 1-propanol} at different temperatures
rhol	799.67	kg/m3	298.15	Volumetric, acoustic, and viscometric studies of molecular interactions in binary mixtures of dipropylene glycol dimethyl ether with 1-alkanols at 298.15 K
rhol	778.97	kg/m3	323.15	Densities and volumetric properties of (N-(2-hydroxyethyl)morpholine + ethanol, + 1-propanol, + 2-propanol, + 1-butanol, and + 2-butanol) at (293.15, 298.15, 303.15, 313.15, and 323.15) K
rhol	787.34	kg/m3	313.15	Densities and volumetric properties of (N-(2-hydroxyethyl)morpholine + ethanol, + 1-propanol, + 2-propanol, + 1-butanol, and + 2-butanol) at (293.15, 298.15, 303.15, 313.15, and 323.15) K
rhol	795.58	kg/m3	303.15	Densities and volumetric properties of (N-(2-hydroxyethyl)morpholine + ethanol, + 1-propanol, + 2-propanol, + 1-butanol, and + 2-butanol) at (293.15, 298.15, 303.15, 313.15, and 323.15) K

rhol	799.71	kg/m3	298.15	Densities and volumetric properties of (N-(2-hydroxyethyl)morpholine + ethanol, + 1-propanol, + 2-propanol, + 1-butanol, and + 2-butanol) at (293.15, 298.15, 303.15, 313.15, and 323.15) K
rhol	803.56	kg/m3	293.15	Densities and volumetric properties of (N-(2-hydroxyethyl)morpholine + ethanol, + 1-propanol, + 2-propanol, + 1-butanol, and + 2-butanol) at (293.15, 298.15, 303.15, 313.15, and 323.15) K
rhol	787.36	kg/m3	313.15	Excess molar volumes of binary mixtures of 1,3-dimethylimidazolidin-2-one with an alkan-1-ol at the temperatures 283.15 K, 298.15 K, and 313.15 K
rhol	799.59	kg/m3	298.15	Excess molar volumes of binary mixtures of 1,3-dimethylimidazolidin-2-one with an alkan-1-ol at the temperatures 283.15 K, 298.15 K, and 313.15 K
rhol	811.54	kg/m3	283.15	Excess molar volumes of binary mixtures of 1,3-dimethylimidazolidin-2-one with an alkan-1-ol at the temperatures 283.15 K, 298.15 K, and 313.15 K
rhol	761.51	kg/m3	343.15	Densities and derived thermodynamic properties of (2-methoxyethanol + 1-propanol, or 2-propanol, or 1,2-propandiol) at temperatures from T = (293.15 to 343.15) K

rhol	770.42	kg/m3	333.15	Densities and derived thermodynamic properties of (2-methoxyethanol + 1-propanol, or 2-propanol, or 1,2-propandiol) at temperatures from T = (293.15 to 343.15) K
rhol	779.02	kg/m3	323.15	Densities and derived thermodynamic properties of (2-methoxyethanol + 1-propanol, or 2-propanol, or 1,2-propandiol) at temperatures from T = (293.15 to 343.15) K
rhol	787.39	kg/m3	313.15	Densities and derived thermodynamic properties of (2-methoxyethanol + 1-propanol, or 2-propanol, or 1,2-propandiol) at temperatures from T = (293.15 to 343.15) K
rhol	795.57	kg/m3	303.15	Densities and derived thermodynamic properties of (2-methoxyethanol + 1-propanol, or 2-propanol, or 1,2-propandiol) at temperatures from T = (293.15 to 343.15) K
rhol	803.61	kg/m3	293.15	Densities and derived thermodynamic properties of (2-methoxyethanol + 1-propanol, or 2-propanol, or 1,2-propandiol) at temperatures from T = (293.15 to 343.15) K

rhol	783.22	kg/m3	318.15	Volumetric properties of binary mixtures of ionic liquid 1-butyl-3-methylimidazolium octylsulfate with water or propanol in the temperature range of 278.15 K to 328.15 K
rhol	791.44	kg/m3	308.15	Volumetric properties of binary mixtures of ionic liquid 1-butyl-3-methylimidazolium octylsulfate with water or propanol in the temperature range of 278.15 K to 328.15 K
rhol	799.54	kg/m3	298.15	Volumetric properties of binary mixtures of ionic liquid 1-butyl-3-methylimidazolium octylsulfate with water or propanol in the temperature range of 278.15 K to 328.15 K
rhol	807.54	kg/m3	288.15	Volumetric properties of binary mixtures of ionic liquid 1-butyl-3-methylimidazolium octylsulfate with water or propanol in the temperature range of 278.15 K to 328.15 K
rhol	815.98	kg/m3	278.15	Volumetric properties of binary mixtures of ionic liquid 1-butyl-3-methylimidazolium octylsulfate with water or propanol in the temperature range of 278.15 K to 328.15 K
rhol	799.60	kg/m3	298.15	Bubble point temperatures of the binary mixtures of nitrobenzene with C1 C4 aliphatic alcohols at 94.95 kPa

rhol	799.74	kg/m3	298.15	(Vapor + liquid) equilibria for the binary mixtures (1-propanol + dibromomethane, or + bromochloromethane, or + 1,2-dichloroethane or +1-bromo-2-chloroethane) at T = 313.15 K.	
rhol	791.30	kg/m3	308.15	Temperature dependence of the volumetric properties of some alkoxypropanols + n-alkanol mixtures	
rhol	799.40	kg/m3	298.15	Temperature dependence of the volumetric properties of some alkoxypropanols + n-alkanol mixtures	
rhol	807.30	kg/m3	288.15	Temperature dependence of the volumetric properties of some alkoxypropanols + n-alkanol mixtures	
rhol	800.06	kg/m3	298.15	Solid-liquid equilibria for selected binary systems containing diphenyl carbonate	
rhol	791.82	kg/m3	308.15	A systematic study on volumetric and transport properties of binary systems 1-propanol + n-hexadecane, 1-butanol + n-hexadecane and 1-propanol + ethyl oleate at different temperatures: Experimental and modeling	

rhol	795.89	kg/m3	303.15	A systematic study on volumetric and transport properties of binary systems 1-propanol + n-hexadecane, 1-butanol + n-hexadecane and 1-propanol + ethyl oleate at different temperatures: Experimental and modeling
rhol	799.93	kg/m3	298.15	A systematic study on volumetric and transport properties of binary systems 1-propanol + n-hexadecane, 1-butanol + n-hexadecane and 1-propanol + ethyl oleate at different temperatures: Experimental and modeling
rhol	799.52	kg/m3	298.15	Separation of pyridine from heptane with tricyanomethanide-based ionic liquids
rhol	779.14	kg/m3	323.15	Measurement and modeling of volumetric properties and speeds of sound of several mixtures of alcohol liquids containing 1-propanol and 2-propanol at T= (298.15 - 323.15) K and ambient pressure
rhol	783.36	kg/m3	318.15	Measurement and modeling of volumetric properties and speeds of sound of several mixtures of alcohol liquids containing 1-propanol and 2-propanol at T= (298.15 - 323.15) K and ambient pressure

rhol	787.52	kg/m3	313.15	Measurement and modeling of volumetric properties and speeds of sound of several mixtures of alcohol liquids containing 1-propanol and 2-propanol at T= (298.15 - 323.15) K and ambient pressure	
rhol	791.63	kg/m3	308.15	Measurement and modeling of volumetric properties and speeds of sound of several mixtures of alcohol liquids containing 1-propanol and 2-propanol at T= (298.15 - 323.15) K and ambient pressure	
rhol	795.70	kg/m3	303.15	Measurement and modeling of volumetric properties and speeds of sound of several mixtures of alcohol liquids containing 1-propanol and 2-propanol at T= (298.15 - 323.15) K and ambient pressure	
rhol	799.71	kg/m3	298.15	Measurement and modeling of volumetric properties and speeds of sound of several mixtures of alcohol liquids containing 1-propanol and 2-propanol at T= (298.15 - 323.15) K and ambient pressure	

rhol	799.62	kg/m3	298.15	Liquid liquid equilibria of 4-methyl-2-pentanone + 1-propanol or 2-propanol + water ternary systems: Measurements and correlation at different temperatures
rhol	800.13	kg/m3	298.15	Thermodynamic and spectral investigations of binary liquid mixturesof 2-butoxy ethanol with alcohols at temperature range of 293.15-313.15 K
rhol	800.61	kg/m3	298.15	Isothermal vapor liquid equilibrium at 333.15K and excess molar volumes at 298.15K for the ternary system di-isopropyl ether + n-propyl alcohol + toluene and its binary subsystems
rhol	799.96	kg/m3	298.15	Liquid liquid equilibria for the binary system of di-isopropyl ether (DIPE) +water in between 288.15 and 323.15K and the ternary systems of DIPE +water + C1 C4 alcohols at 298.15K
rhol	791.76	kg/m3	308.15	Thermodynamic properties of mixtures containing alkoxypropanol and n-alkanol
rhol	799.67	kg/m3	298.15	Thermodynamic properties of mixtures containing alkoxypropanol and n-alkanol
rhol	807.90	kg/m3	288.15	Thermodynamic properties of mixtures containing alkoxypropanol and n-alkanol

rhol	787.47	kg/m3	313.15	Influence of chain length and degree of branching of alcohol + chlorobenzene mixtures on determination and modelling of VE by CEOS and CEOS/GE mixing rules	
rhol	791.58	kg/m3	308.15	Influence of chain length and degree of branching of alcohol + chlorobenzene mixtures on determination and modelling of VE by CEOS and CEOS/GE mixing rules	
rhol	795.65	kg/m3	303.15	Influence of chain length and degree of branching of alcohol + chlorobenzene mixtures on determination and modelling of VE by CEOS and CEOS/GE mixing rules	
rhol	799.69	kg/m3	298.15	Influence of chain length and degree of branching of alcohol + chlorobenzene mixtures on determination and modelling of VE by CEOS and CEOS/GE mixing rules	
rhol	803.70	kg/m3	293.15	Influence of chain length and degree of branching of alcohol + chlorobenzene mixtures on determination and modelling of VE by CEOS and CEOS/GE mixing rules	

rhol	807.69	kg/m3	288.15	Influence of chain length and degree of branching of alcohol + chlorobenzene mixtures on determination and modelling of VE by CEOS and CEOS/GE mixing rules
rhol	799.65	kg/m3	298.15	Excess molar enthalpy of 1-alkanol + 1-octene mixtures at 298.15K Experimental results and theoretical description by means of the ERAS and TB models
rhol	770.39	kg/m3	333.15	Volumetric properties of the boldine + alcohol mixtures at atmospheric pressure from 283.15 to 333.15K A new method for the determination of the density of pure boldine
rhol	762.86	kg/m3	343.15	Densities of Ionic Liquids, 1-Butyl-3-methylimidazolium Hexafluorophosphate and 1-Butyl-3-methylimidazolium Tetrafluoroborate, with Benzene, Acetonitrile, and 1-Propanol at T = (293.15 to 343.15) K
rhol	779.01	kg/m3	323.15	Volumetric properties of the boldine + alcohol mixtures at atmospheric pressure from 283.15 to 333.15K A new method for the determination of the density of pure boldine

rhol	799.65	kg/m3	298.15	Thermodynamic evidence for nano-heterogeneity in solutions of the macrocycle C-butylresorcin[4]arene in non-aqueous solvents
sfust	36.11	J/mol×K	148.75	NIST Webbook
sfust	35.30	J/mol×K	147.00	NIST Webbook
speedsl	1104.40	m/s	328.15	Volumetric and Ultrasonic Studies of 1-Ethyl-3-methylimidazolium Trifluoromethanesulfonate Ionic Liquid with Methanol, Ethanol, 1-Propanol, and Water at Several Temperatures
speedsl	1205.37	m/s	298.15	Volumetric and Ultrasonic Studies of 1-Ethyl-3-methylimidazolium Trifluoromethanesulfonate Ionic Liquid with Methanol, Ethanol, 1-Propanol, and Water at Several Temperatures
speedsl	1239.86	m/s	288.15	Volumetric and Ultrasonic Studies of 1-Ethyl-3-methylimidazolium Trifluoromethanesulfonate Ionic Liquid with Methanol, Ethanol, 1-Propanol, and Water at Several Temperatures
speedsl	1275.27	m/s	278.15	Volumetric and Ultrasonic Studies of 1-Ethyl-3-methylimidazolium Trifluoromethanesulfonate Ionic Liquid with Methanol, Ethanol, 1-Propanol, and Water at Several Temperatures
speedsl	1207.00	m/s	298.15	Compressibility Studies of Binary Solutions Involving Water as a Solute in Nonaqueous Solvents at T) 298.15 K

speedsl	1155.10	m/s	313.15	Excess Enthalpy, Excess Volume, and Speed of Sound Deviation for the Mixtures ss-Pinene + Ethanol and ss-Pinene + 1-propanol at (283.15, 298.15 and, 313.15) K
speedsl	1206.00	m/s	298.15	Excess Enthalpy, Excess Volume, and Speed of Sound Deviation for the Mixtures ss-Pinene + Ethanol and ss-Pinene + 1-propanol at (283.15, 298.15 and, 313.15) K
speedsl	1257.90	m/s	283.15	Excess Enthalpy, Excess Volume, and Speed of Sound Deviation for the Mixtures ss-Pinene + Ethanol and ss-Pinene + 1-propanol at (283.15, 298.15 and, 313.15) K
speedsl	1121.80	m/s	323.15	Acoustic, volumetric and osmotic properties of binary mixtures containing the ionic liquid 1-butyl-3-methylimidazolim dicyanamide mixed with primary and secondary alcohols
speedsl	1172.30	m/s	308.15	Acoustic, volumetric and osmotic properties of binary mixtures containing the ionic liquid 1-butyl-3-methylimidazolim dicyanamide mixed with primary and secondary alcohols

speedsl	1223.40	m/s	293.15	Acoustic, volumetric and osmotic properties of binary mixtures containing the ionic liquid 1-butyl-3-methylimidazolim dicyanamide mixed with primary and secondary alcohols
speedsl	1205.80	m/s	298.15	Ultrasonic speeds and isentropic compressibilities of {difurylmethane + (C1 C6) n-alkanol} binary mixtures at T = 298.15 K
speedsl	1172.04	m/s	308.15	Densities, excess molar volumes, speeds of sound and isothermal compressibilities for {2-(2-hexyloxyethoxy)ethanol + n-alkanol} systems at temperatures between (288.15 and 308.15) K
speedsl	1189.86	m/s	303.15	Densities, excess molar volumes, speeds of sound and isothermal compressibilities for {2-(2-hexyloxyethoxy)ethanol + n-alkanol} systems at temperatures between (288.15 and 308.15) K
speedsl	1206.47	m/s	298.15	Densities, excess molar volumes, speeds of sound and isothermal compressibilities for {2-(2-hexyloxyethoxy)ethanol + n-alkanol} systems at temperatures between (288.15 and 308.15) K

speedsl	1223.17	m/s	293.15	Densities, excess molar volumes, speeds of sound and isothermal compressibilities for {2-(2-hexyloxyethoxy)ethanol + n-alkanol} systems at temperatures between (288.15 and 308.15) K
speedsl	1240.01	m/s	288.15	Densities, excess molar volumes, speeds of sound and isothermal compressibilities for {2-(2-hexyloxyethoxy)ethanol + n-alkanol} systems at temperatures between (288.15 and 308.15) K
speedsl	1171.41	m/s	308.15	Volumetric and Ultrasonic Studies of 1-Ethyl-3-methylimidazolium Trifluoromethanesulfonate Ionic Liquid with Methanol, Ethanol, 1-Propanol, and Water at Several Temperatures
speedsl	1137.83	m/s	318.15	Volumetric and Ultrasonic Studies of 1-Ethyl-3-methylimidazolium Trifluoromethanesulfonate Ionic Liquid with Methanol, Ethanol, 1-Propanol, and Water at Several Temperatures
speedsl	1070.97	m/s	338.15	Volumetric and Ultrasonic Studies of 1-Ethyl-3-methylimidazolium Trifluoromethanesulfonate Ionic Liquid with Methanol, Ethanol, 1-Propanol, and Water at Several Temperatures
srf	0.02	N/m	298.15	Surface Tension of Binary Mixtures of 2,2,4-Trimethylpentane +1-Alkanols from 298.15 to 323.15 K

srf	0.02	N/m	313.15	Surface Tension of Dilute Solutions of Linear Alcohols in Benzyl Alcohol
srf	0.02	N/m	308.15	Surface Tension of Dilute Solutions of Linear Alcohols in Benzyl Alcohol
srf	0.02	N/m	303.15	Surface Tension of Dilute Solutions of Linear Alcohols in Benzyl Alcohol
srf	0.02	N/m	298.15	Surface Tension of Dilute Solutions of Linear Alcohols in Benzyl Alcohol
srf	0.02	N/m	293.15	Surface Tension of Dilute Solutions of Linear Alcohols in Benzyl Alcohol
srf	0.02	N/m	323.15	Surface Tension of Binary Mixtures of 2,2,4-Trimethylpentane +1-Alkanols from 298.15 to 323.15 K
srf	0.02	N/m	318.15	Surface Tension of Binary Mixtures of 2,2,4-Trimethylpentane +1-Alkanols from 298.15 to 323.15 K
srf	0.02	N/m	313.15	Surface Tension of Binary Mixtures of 2,2,4-Trimethylpentane +1-Alkanols from 298.15 to 323.15 K
srf	0.02	N/m	308.15	Surface Tension of Binary Mixtures of 2,2,4-Trimethylpentane +1-Alkanols from 298.15 to 323.15 K
srf	0.02	N/m	303.15	Surface Tension of Binary Mixtures of 2,2,4-Trimethylpentane +1-Alkanols from 298.15 to 323.15 K

srf	0.02	N/m	318.15	Surface Tension of Dilute Solutions of Linear Alcohols in Benzyl Alcohol	
srf	0.02	N/m	333.20	Surface tension and interfacial compositions of binary glycerol/alcohol mixtures	
srf	0.02	N/m	323.20	Surface tension and interfacial compositions of binary glycerol/alcohol mixtures	
srf	0.02	N/m	313.20	Surface tension and interfacial compositions of binary glycerol/alcohol mixtures	
srf	0.02	N/m	303.20	Surface tension and interfacial compositions of binary glycerol/alcohol mixtures	
srf	0.02	N/m	293.20	Surface tension and interfacial compositions of binary glycerol/alcohol mixtures	
srf	0.02	N/m	318.15	The molar surface Gibbs energy and its application to the binary mixtures of N-butylpyridinium dicyanamide [C4py][DCA] with alcohols	
srf	0.02	N/m	313.15	The molar surface Gibbs energy and its application to the binary mixtures of N-butylpyridinium dicyanamide [C4py][DCA] with alcohols	

srf	0.02	N/m	308.15	The molar surface Gibbs energy and its application to the binary mixtures of N-butylpyridinium dicyanamide [C4py][DCA] with alcohols	
srf	0.02	N/m	303.15	The molar surface Gibbs energy and its application to the binary mixtures of N-butylpyridinium dicyanamide [C4py][DCA] with alcohols	
srf	0.02	N/m	298.15	The molar surface Gibbs energy and its application to the binary mixtures of N-butylpyridinium dicyanamide [C4py][DCA] with alcohols	
srf	0.02	N/m	293.15	The molar surface Gibbs energy and its application to the binary mixtures of N-butylpyridinium dicyanamide [C4py][DCA] with alcohols	
srf	0.02	N/m	288.15	The molar surface Gibbs energy and its application to the binary mixtures of N-butylpyridinium dicyanamide [C4py][DCA] with alcohols	
srf	0.02	N/m	328.15	Study of surface tension and surface properties of binary systems of DMSO with long chain alcohols at various temperatures	

srf	0.02	N/m	318.15	Study of surface tension and surface properties of binary systems of DMSO with long chain alcohols at various temperatures	
srf	0.02	N/m	308.15	Study of surface tension and surface properties of binary systems of DMSO with long chain alcohols at various temperatures	
srf	0.02	N/m	298.15	Study of surface tension and surface properties of binary systems of DMSO with long chain alcohols at various temperatures	
srf	0.02	N/m	288.15	Study of surface tension and surface properties of binary systems of DMSO with long chain alcohols at various temperatures	
srf	0.02	N/m	318.15	Surface thermodynamics of binary mixtures of aliphatic alcohols in heavy water	
srf	0.02	N/m	308.15	Surface thermodynamics of binary mixtures of aliphatic alcohols in heavy water	
srf	0.02	N/m	298.15	Surface thermodynamics of binary mixtures of aliphatic alcohols in heavy water	
srf	0.02	N/m	288.15	Surface thermodynamics of binary mixtures of aliphatic alcohols in heavy water	

srf	0.02	N/m	313.15	Application of the Extended Langmuir Model for the Determination of Lyophobicity of 1-Propanol in Acetonitrile	
srf	0.02	N/m	308.15	Application of the Extended Langmuir Model for the Determination of Lyophobicity of 1-Propanol in Acetonitrile	
srf	0.02	N/m	303.15	Application of the Extended Langmuir Model for the Determination of Lyophobicity of 1-Propanol in Acetonitrile	
srf	0.02	N/m	323.15	Surface Tension of Dilute Solutions of Linear Alcohols in Benzyl Alcohol	
srf	0.02	N/m	293.15	Application of the Extended Langmuir Model for the Determination of Lyophobicity of 1-Propanol in Acetonitrile	
srf	0.02	N/m	288.15	Application of the Extended Langmuir Model for the Determination of Lyophobicity of 1-Propanol in Acetonitrile	
srf	0.02	N/m	283.15	Application of the Extended Langmuir Model for the Determination of Lyophobicity of 1-Propanol in Acetonitrile	
srf	0.02	N/m	278.15	Application of the Extended Langmuir Model for the Determination of Lyophobicity of 1-Propanol in Acetonitrile	
srf	0.02	N/m	293.20	KDB	

srf	0.02	N/m	293.15	Density and Surface Tension of Binary Mixtures of Acetonitrile + 1-Alkanol at 293.15 K
srf	0.02	N/m	298.00	Surface Tension and Density of Pure Ionic Liquids and Some Binary Mixtures with 1-Propanol and 1-Butanol
srf	0.02	N/m	298.15	Application of the Extended Langmuir Model for the Determination of Lyophobicity of 1-Propanol in Acetonitrile

Pressure Dependent Properties

Property code	Value	Unit	Pressure [kPa]	Source
tbp	360.86	K	69.79	A novel dynamic recirculating apparatus for vapour-liquid equilibrium measurements at moderate pressures and temperatures
tbp	362.60	K	74.92	A novel dynamic recirculating apparatus for vapour-liquid equilibrium measurements at moderate pressures and temperatures
tbp	364.11	K	79.60	A novel dynamic recirculating apparatus for vapour-liquid equilibrium measurements at moderate pressures and temperatures

tbp	365.70	K	84.80	A novel dynamic recirculating apparatus for vapour-liquid equilibrium measurements at moderate pressures and temperatures	
tbp	367.14	K	89.78	A novel dynamic recirculating apparatus for vapour-liquid equilibrium measurements at moderate pressures and temperatures	
tbp	368.52	K	94.77	A novel dynamic recirculating apparatus for vapour-liquid equilibrium measurements at moderate pressures and temperatures	
tbp	369.82	K	99.70	A novel dynamic recirculating apparatus for vapour-liquid equilibrium measurements at moderate pressures and temperatures	
tbp	371.30	K	105.14	A novel dynamic recirculating apparatus for vapour-liquid equilibrium measurements at moderate pressures and temperatures	
tbp	372.42	K	110.10	A novel dynamic recirculating apparatus for vapour-liquid equilibrium measurements at moderate pressures and temperatures	
tbp	374.88	K	120.75	A novel dynamic recirculating apparatus for vapour-liquid equilibrium measurements at moderate pressures and temperatures	

tbp	373.15	K	113.19	A novel dynamic recirculating apparatus for vapour-liquid equilibrium measurements at moderate pressures and temperatures	
tbp	376.91	K	130.10	A novel dynamic recirculating apparatus for vapour-liquid equilibrium measurements at moderate pressures and temperatures	
tbp	384.57	K	171.00	A novel dynamic recirculating apparatus for vapour-liquid equilibrium measurements at moderate pressures and temperatures	
tbp	389.14	K	199.90	A novel dynamic recirculating apparatus for vapour-liquid equilibrium measurements at moderate pressures and temperatures	
tbp	393.15	K	228.50	A novel dynamic recirculating apparatus for vapour-liquid equilibrium measurements at moderate pressures and temperatures	
tbp	393.35	K	229.90	A novel dynamic recirculating apparatus for vapour-liquid equilibrium measurements at moderate pressures and temperatures	
tbp	397.13	K	260.00	A novel dynamic recirculating apparatus for vapour-liquid equilibrium measurements at moderate pressures and temperatures	

tbp	400.57	K	290.00	A novel dynamic recirculating apparatus for vapour-liquid equilibrium measurements at moderate pressures and temperatures	
tbp	402.71	K	310.00	A novel dynamic recirculating apparatus for vapour-liquid equilibrium measurements at moderate pressures and temperatures	
tbp	404.74	K	330.00	A novel dynamic recirculating apparatus for vapour-liquid equilibrium measurements at moderate pressures and temperatures	
tbp	406.69	K	350.00	A novel dynamic recirculating apparatus for vapour-liquid equilibrium measurements at moderate pressures and temperatures	
tbp	409.43	K	380.00	A novel dynamic recirculating apparatus for vapour-liquid equilibrium measurements at moderate pressures and temperatures	
tbp	411.18	K	400.00	A novel dynamic recirculating apparatus for vapour-liquid equilibrium measurements at moderate pressures and temperatures	
tbp	412.84	K	420.00	A novel dynamic recirculating apparatus for vapour-liquid equilibrium measurements at moderate pressures and temperatures	

tbp	415.23	K	450.00	A novel dynamic recirculating apparatus for vapour-liquid equilibrium measurements at moderate pressures and temperatures	
tbp	417.52	K	480.00	A novel dynamic recirculating apparatus for vapour-liquid equilibrium measurements at moderate pressures and temperatures	
tbp	418.96	K	500.00	A novel dynamic recirculating apparatus for vapour-liquid equilibrium measurements at moderate pressures and temperatures	
tbp	420.37	K	520.00	A novel dynamic recirculating apparatus for vapour-liquid equilibrium measurements at moderate pressures and temperatures	
tbp	421.06	K	530.00	A novel dynamic recirculating apparatus for vapour-liquid equilibrium measurements at moderate pressures and temperatures	
tbp	422.41	K	550.00	A novel dynamic recirculating apparatus for vapour-liquid equilibrium measurements at moderate pressures and temperatures	
tbp	423.16	K	561.50	A novel dynamic recirculating apparatus for vapour-liquid equilibrium measurements at moderate pressures and temperatures	

Correlations

Information	Value
Property code	pvap
Equation	ln(Pvp) = A + B/(T + C)
Coeff. A	1.58913e+01
Coeff. B	-3.49716e+03
Coeff. C	-6.00740e+01
Temperature range (K), min.	284.20
Temperature range (K), max.	536.78

Information	Value
Property code	pvap
Equation	ln(Pvp) = A + B/T + C*ln(T) + D*T^2
Coeff. A	7.17170e+01
Coeff. B	-8.00269e+03
Coeff. C	-7.70223e+00
Coeff. D	3.95045e-07
Temperature range (K), min.	146.95
Temperature range (K), max.	536.71

Datasets

Viscosity, Pa*s

Temperature, K - Liquid	Pressure, kPa - Liquid	Viscosity, Pa*s - Liquid
298.15	81.50	0.0019430

Reference *https://www.doi.org/10.1016/j.jct.2016.12.036*

Pressure, kPa	Temperature, K	Viscosity, Pa*s
100.00	293.15	0.0022040
100.00	313.15	0.0013830

100.00	333.15	0.0009130
100.00	353.15	0.0006300
20000.00	293.15	0.0025450
20000.00	313.15	0.0016030
20000.00	333.15	0.0010860
20000.00	353.15	0.0007340
40000.00	293.15	0.0028550
40000.00	313.15	0.0018710
40000.00	333.15	0.0012630
40000.00	353.15	0.0008450
60000.00	293.15	0.0031910
60000.00	313.15	0.0021460
60000.00	333.15	0.0014390
60000.00	353.15	0.0009620
80000.00	293.15	0.0036130
80000.00	313.15	0.0024010
80000.00	333.15	0.0016140
80000.00	353.15	0.0010920
100000.00	293.15	0.0041490
100000.00	313.15	0.0026180
100000.00	333.15	0.0017860
100000.00	353.15	0.0012420

Reference *https://www.doi.org/10.1021/je9002477*

Refractive index (Na D-line)

Pressure, kPa - Liquid	Temperature, K - Liquid	Refractive index (Na D-line) - Liquid
100.00	298.15	1.3836

Reference *https://www.doi.org/10.1016/j.jct.2016.01.012*

Temperature, K	Pressure, kPa	Refractive index (Na D-line)
294.15	100.00	1.3846

Reference *https://www.doi.org/10.1021/acs.jced.5b01024*

Mass density, kg/m3

Pressure, kPa - Liquid	Temperature, K - Liquid	Mass density, kg/m3 - Liquid
100.00	293.15	803.5
100.00	303.15	795.2
100.00	313.15	787.6
100.00	323.15	779.3
100.00	333.15	770.5
100.00	343.15	761.0
100.00	353.15	752.2
5000.00	293.15	807.3
5000.00	303.15	799.2
5000.00	313.15	791.9
5000.00	323.15	783.7
5000.00	333.15	775.4
5000.00	343.15	766.0
5000.00	353.15	757.6
10000.00	293.15	810.9
10000.00	303.15	802.9
10000.00	313.15	795.9
10000.00	323.15	788.0
10000.00	333.15	780.0
10000.00	343.15	770.7
10000.00	353.15	762.7
15000.00	293.15	814.3
15000.00	303.15	806.7
15000.00	313.15	799.7
15000.00	323.15	792.0
15000.00	333.15	784.3
15000.00	343.15	775.3
15000.00	353.15	767.6
20000.00	293.15	817.7
20000.00	303.15	810.1
20000.00	313.15	803.3
20000.00	323.15	795.8
20000.00	333.15	788.4
20000.00	343.15	779.6
20000.00	353.15	772.1
25000.00	293.15	821.0
25000.00	303.15	813.6
25000.00	313.15	806.8
25000.00	323.15	799.7

25000.00	333.15	792.0
25000.00	343.15	783.7
25000.00	353.15	776.3
30000.00	293.15	824.0
30000.00	303.15	816.7
30000.00	313.15	810.2
30000.00	323.15	803.2
30000.00	333.15	795.7
30000.00	343.15	787.6
30000.00	353.15	780.4
35000.00	293.15	827.0
35000.00	303.15	819.8
35000.00	313.15	813.5
35000.00	323.15	806.6
35000.00	333.15	799.2
35000.00	343.15	791.2
35000.00	353.15	784.1
40000.00	293.15	829.9
40000.00	303.15	822.9
40000.00	313.15	816.6
40000.00	323.15	809.9
40000.00	333.15	802.6
40000.00	343.15	794.7
40000.00	353.15	787.7
45000.00	293.15	832.7
45000.00	303.15	825.8
45000.00	313.15	819.7
45000.00	323.15	813.0
45000.00	333.15	805.9
45000.00	343.15	798.3
45000.00	353.15	791.8
50000.00	293.15	835.4
50000.00	303.15	828.8
50000.00	313.15	822.7
50000.00	323.15	816.1
50000.00	333.15	809.1
50000.00	343.15	801.6
50000.00	353.15	795.1
55000.00	293.15	838.1
55000.00	303.15	831.6
55000.00	313.15	825.5
55000.00	323.15	819.0
55000.00	333.15	812.3
55000.00	343.15	804.6

55000.00	353.15	798.5
60000.00	293.15	840.7
60000.00	303.15	834.2
60000.00	313.15	828.3
60000.00	323.15	821.8
60000.00	333.15	815.2
60000.00	343.15	807.7
60000.00	353.15	801.7
65000.00	293.15	843.2
65000.00	303.15	837.0
65000.00	313.15	831.0
65000.00	323.15	824.6
65000.00	333.15	818.1
65000.00	343.15	810.6
65000.00	353.15	805.0

Reference *https://www.doi.org/10.1016/j.fluid.2006.05.032*

Temperature, K	Pressure, kPa	Mass density, kg/m3
298.15	110.00	799.54
298.15	5270.00	803.6
298.15	9870.00	807.01
298.15	14700.00	810.46
298.15	20530.00	814.46
298.15	25030.00	817.44
298.15	30100.00	820.7
298.15	35010.00	823.76
298.15	39560.00	826.52
323.15	590.00	779.78
323.15	5940.00	784.71
323.15	10600.00	788.81
323.15	16000.00	793.2
323.15	20640.00	796.73
323.15	24880.00	799.83
323.15	30800.00	803.99
323.15	35220.00	807.01
323.15	39790.00	810.02
348.15	140.00	757.11
348.15	5170.00	762.68
348.15	10260.00	767.5
348.15	15790.00	772.52
348.15	20560.00	776.41
348.15	24810.00	780.32

348.15	29300.00	784.09
348.15	35560.00	789.01
348.15	39950.00	792.23
373.15	360.00	733.24
373.15	5480.00	739.29
373.15	10290.00	745.11
373.15	15700.00	750.9
373.15	19960.00	755.19
373.15	25170.00	760.17
373.15	30510.00	765.02
373.15	35920.00	769.7
373.15	39780.00	772.91
398.15	620.00	705.55
398.15	5670.00	713.23
398.15	10620.00	720.1
398.15	15580.00	726.46
398.15	20350.00	732.35
398.15	25090.00	737.77
398.15	31150.00	744.07
398.15	35410.00	748.41
398.15	39390.00	752.32
423.15	1570.00	674.63
423.15	5870.00	682.79
423.15	10160.00	690.93
423.15	15320.00	699.32
423.15	20500.00	707.2
423.15	26280.00	715.01
423.15	31600.00	721.36
423.15	35780.00	726.03
423.15	39700.00	730.17

Reference *https://www.doi.org/10.1016/j.fluid.2008.03.009*

Temperature, K	Pressure, kPa	Mass density, kg/m3
293.15	100.00	803.8
293.15	10000.00	811.3
293.15	20000.00	818.1
293.15	30000.00	824.4
293.15	40000.00	830.0
293.15	50000.00	835.6
293.15	60000.00	841.0
293.15	70000.00	845.9
293.15	80000.00	850.7

293.15	90000.00	855.2
293.15	100000.00	859.6
293.15	110000.00	863.7
293.15	120000.00	867.8
293.15	130000.00	871.6
293.15	140000.00	875.3
313.15	100.00	787.1
313.15	10000.00	795.5
313.15	20000.00	803.0
313.15	30000.00	810.0
313.15	40000.00	816.4
313.15	50000.00	822.3
313.15	60000.00	827.9
313.15	70000.00	833.2
313.15	80000.00	838.3
313.15	90000.00	843.0
313.15	100000.00	847.9
313.15	110000.00	852.3
313.15	120000.00	856.6
313.15	130000.00	860.7
313.15	140000.00	864.6
333.15	100.00	769.9
333.15	10000.00	780.0
333.15	20000.00	788.3
333.15	30000.00	795.7
333.15	40000.00	802.5
333.15	50000.00	809.0
333.15	60000.00	815.0
333.15	70000.00	820.6
333.15	80000.00	826.2
333.15	90000.00	831.2
333.15	100000.00	836.3
333.15	110000.00	840.9
333.15	120000.00	845.3
333.15	130000.00	849.6
333.15	140000.00	853.6
353.15	100.00	752.0
353.15	10000.00	762.5
353.15	20000.00	771.8
353.15	30000.00	780.2
353.15	40000.00	787.9
353.15	50000.00	795.0
353.15	60000.00	801.4
353.15	70000.00	807.7

353.15	80000.00	813.4
353.15	90000.00	818.7
353.15	100000.00	823.8
353.15	110000.00	828.8
353.15	120000.00	833.5
353.15	130000.00	837.9
353.15	140000.00	842.1

Reference *https://www.doi.org/10.1016/j.fluid.2013.11.031*

Temperature, K	Pressure, kPa	Mass density, kg/m3
278.15	100.00	815.7
278.15	1000.00	816.4
278.15	5000.00	819.4
278.15	10000.00	822.8
278.15	15000.00	826.2
278.15	20000.00	829.4
278.15	25000.00	832.6
278.15	30000.00	835.7
278.15	35000.00	838.6
278.15	40000.00	841.5
278.15	45000.00	844.3
278.15	50000.00	847.1
278.15	55000.00	849.7
278.15	60000.00	852.1
288.15	100.00	807.8
288.15	1000.00	808.5
288.15	5000.00	811.6
288.15	10000.00	815.3
288.15	15000.00	818.9
288.15	20000.00	822.3
288.15	25000.00	825.5
288.15	30000.00	828.7
288.15	35000.00	831.8
288.15	40000.00	834.7
288.15	45000.00	837.6
288.15	50000.00	840.4
288.15	55000.00	843.1
288.15	60000.00	845.6
298.15	100.00	799.8
298.15	1000.00	800.5
298.15	5000.00	803.8
298.15	10000.00	807.6

298.15	15000.00	811.4
298.15	20000.00	814.9
298.15	25000.00	818.3
298.15	30000.00	821.6
298.15	35000.00	824.8
298.15	40000.00	827.9
298.15	45000.00	830.9
298.15	50000.00	833.8
298.15	55000.00	836.6
298.15	60000.00	839.2
308.15	100.00	791.6
308.15	1000.00	792.4
308.15	5000.00	795.8
308.15	10000.00	800.1
308.15	15000.00	803.8
308.15	20000.00	807.5
308.15	25000.00	811.1
308.15	30000.00	814.5
308.15	35000.00	817.8
308.15	40000.00	821.0
308.15	45000.00	824.1
308.15	50000.00	827.1
308.15	55000.00	830.0
308.15	60000.00	832.7
318.15	100.00	783.4
318.15	1000.00	784.2
318.15	5000.00	787.9
318.15	10000.00	792.2
318.15	15000.00	796.3
318.15	20000.00	800.2
318.15	25000.00	803.9
318.15	30000.00	807.5
318.15	35000.00	810.9
318.15	40000.00	814.2
318.15	45000.00	817.5
318.15	50000.00	820.5
318.15	55000.00	823.6
318.15	60000.00	825.8
328.15	100.00	775.0
328.15	1000.00	775.9
328.15	5000.00	779.7
328.15	10000.00	784.2
328.15	15000.00	788.5
328.15	20000.00	792.6

328.15	25000.00	796.5
328.15	30000.00	800.2
328.15	35000.00	803.8
328.15	40000.00	807.2
328.15	45000.00	810.6
328.15	50000.00	813.8
328.15	55000.00	816.9
328.15	60000.00	819.8
338.15	100.00	766.3
338.15	1000.00	767.2
338.15	5000.00	771.3
338.15	10000.00	776.1
338.15	15000.00	780.6
338.15	20000.00	784.9
338.15	25000.00	789.0
338.15	30000.00	792.9
338.15	35000.00	796.6
338.15	40000.00	800.2
338.15	45000.00	803.7
338.15	50000.00	807.0
338.15	55000.00	810.3
338.15	60000.00	813.4
348.15	100.00	757.2
348.15	1000.00	758.3
348.15	5000.00	762.6
348.15	10000.00	767.7
348.15	15000.00	772.5
348.15	20000.00	777.0
348.15	25000.00	781.3
348.15	30000.00	785.4
348.15	35000.00	789.3
348.15	40000.00	793.0
348.15	45000.00	796.6
348.15	50000.00	800.1
348.15	55000.00	803.5
348.15	60000.00	806.5
358.15	100.00	747.7
358.15	1000.00	748.8
358.15	5000.00	753.5
358.15	10000.00	758.8
358.15	15000.00	764.0
358.15	20000.00	768.7
358.15	25000.00	773.3
358.15	30000.00	777.5

358.15	35000.00	781.6
358.15	40000.00	785.5
358.15	45000.00	789.3
358.15	50000.00	792.9
358.15	55000.00	796.4
358.15	60000.00	799.6

Reference *https://www.doi.org/10.1016/j.jct.2011.10.023*

Temperature, K	Pressure, kPa	Mass density, kg/m3
303.20	1000.00	796.7
303.20	2000.00	797.5
303.20	3000.00	798.3
303.20	4000.00	799.2
303.20	5000.00	799.9
303.20	6000.00	800.7
303.20	7000.00	801.5
303.20	8000.00	802.3
303.20	9000.00	803.1
303.20	10000.00	803.9
323.20	1000.00	779.0
323.20	2000.00	779.9
323.20	3000.00	780.8
323.20	4000.00	781.7
323.20	5000.00	782.6
323.20	6000.00	783.5
323.20	7000.00	784.4
323.20	8000.00	785.3
323.20	9000.00	786.2
323.20	10000.00	787.0

Reference *https://www.doi.org/10.1016/j.jct.2012.10.002*

Temperature, K	Pressure, kPa	Mass density, kg/m3
293.15	100.00	803.85

Reference *https://www.doi.org/10.1016/j.jct.2016.12.011*

Temperature, K	Pressure, kPa	Mass density, kg/m3
298.15	100.00	799.66

Temperature, K	Pressure, kPa	Mass density, kg/m3
288.15	100.00	807.62
288.15	5000.00	811.32
288.15	10000.00	814.7
288.15	15000.00	817.95
288.15	20000.00	821.27
288.15	25000.00	824.24
288.15	30000.00	827.24
288.15	35000.00	830.04
288.15	40000.00	832.73
293.15	100.00	803.57
293.15	5000.00	807.24
293.15	10000.00	810.8
293.15	15000.00	814.22
293.15	20000.00	817.46
293.15	25000.00	820.74
293.15	30000.00	823.7
293.15	35000.00	826.5
293.15	40000.00	829.09
298.15	100.00	799.61
298.15	5000.00	803.49
298.15	10000.00	807.15
298.15	15000.00	810.5
298.15	20000.00	813.78
298.15	25000.00	816.92
298.15	30000.00	820.04
298.15	35000.00	822.87
298.15	40000.00	825.81
303.15	100.00	795.62
303.15	5000.00	799.57
303.15	10000.00	803.13
303.15	15000.00	806.69
303.15	20000.00	810.14
303.15	25000.00	813.48
303.15	30000.00	816.64
303.15	35000.00	819.75
303.15	40000.00	822.6
308.15	100.00	791.31
308.15	5000.00	795.4
308.15	10000.00	799.21

308.15	15000.00	802.79
308.15	20000.00	806.37
308.15	25000.00	809.72
308.15	30000.00	813.03
308.15	35000.00	816.12
308.15	40000.00	819.06

Reference *https://www.doi.org/10.1016/j.jct.2019.02.011*

Temperature, K	Pressure, kPa	Mass density, kg/m3
298.15	81.50	799.54

Reference *https://www.doi.org/10.1021/acs.jced.5b00162*

Temperature, K	Pressure, kPa	Mass density, kg/m3
293.15	100.00	803.17
293.15	5000.00	807.25
293.15	10000.00	811.02
293.15	15000.00	814.54
293.15	20000.00	817.87
293.15	25000.00	821.05
293.15	30000.00	824.15
293.15	35000.00	827.21
303.15	100.00	795.52
303.15	5000.00	799.83
303.15	10000.00	803.8
303.15	15000.00	807.49
303.15	20000.00	810.95
303.15	25000.00	814.23
303.15	30000.00	817.38
303.15	35000.00	820.48
313.15	100.00	787.3
313.15	5000.00	791.61
313.15	10000.00	795.67
313.15	15000.00	799.52
313.15	20000.00	803.19
313.15	25000.00	806.72
313.15	30000.00	810.15
313.15	35000.00	813.5

Reference *https://www.doi.org/10.1021/acs.jced.8b00975*

Temperature, K	Pressure, kPa	Mass density, kg/m3
303.15	100.00	795.6
303.15	5000.00	799.5
303.15	10000.00	803.4
303.15	15000.00	807.0
303.15	20000.00	810.5
303.15	25000.00	813.9
303.15	30000.00	817.2
313.15	100.00	787.6
313.15	5000.00	791.6
313.15	10000.00	795.7
313.15	15000.00	799.5
313.15	20000.00	803.2
313.15	25000.00	806.7
313.15	30000.00	810.1
323.15	100.00	779.1
323.15	5000.00	783.4
323.15	10000.00	787.7
323.15	15000.00	791.7
323.15	20000.00	795.6
323.15	25000.00	799.3
323.15	30000.00	802.9
333.15	100.00	770.6
333.15	5000.00	775.1
333.15	10000.00	779.6
333.15	15000.00	783.9
333.15	20000.00	787.9
333.15	25000.00	791.8
333.15	30000.00	795.5

Reference *https://www.doi.org/10.1021/je049685z*

Temperature, K	Pressure, kPa	Mass density, kg/m3
283.15	100.00	810.9
283.15	5000.00	814.7
283.15	10000.00	818.1
283.15	15000.00	821.7
283.15	20000.00	824.9
298.15	100.00	799.0
298.15	5000.00	803.0
298.15	10000.00	806.9
298.15	15000.00	810.6
298.15	20000.00	814.1

313.15	100.00	786.8
313.15	5000.00	791.2
313.15	10000.00	795.3
313.15	15000.00	799.3
313.15	20000.00	803.1
328.15	100.00	774.5
328.15	5000.00	779.1
328.15	10000.00	783.7
328.15	15000.00	787.8
328.15	20000.00	791.9

Reference *https://www.doi.org/10.1021/je100581m*

Speed of sound, m/s

Temperature, K - Liquid	Pressure, kPa - Liquid	Speed of sound, m/s - Liquid
253.15	102.00	1368.034
253.15	5133.00	1393.016
253.15	10109.00	1416.633
253.15	15198.00	1439.865
253.15	20092.00	1461.442
253.15	25341.00	1483.73
253.15	30423.00	1504.606
273.15	101.00	1293.525
273.15	5181.00	1320.802
273.15	10163.00	1346.251
273.15	15287.00	1371.412
273.15	20261.00	1394.803
273.15	25076.00	1416.612
273.15	30164.00	1438.924
293.15	105.00	1222.539
293.15	5042.00	1251.176
293.15	10141.00	1279.332
293.15	15272.00	1306.262
293.15	20136.00	1330.72
293.15	25068.00	1354.567
293.15	30231.00	1378.484
313.15	108.00	1154.108
313.15	5139.00	1185.752
313.15	10136.00	1215.403
313.15	15128.00	1243.591

313.15	20151.00	1270.557
313.15	25136.00	1296.129
313.01	30237.00	1321.177
333.15	101.00	1086.768
333.15	5018.00	1120.458
333.15	10050.00	1152.816
333.15	15186.00	1183.885
333.15	20183.00	1212.543
333.15	25121.00	1239.487
333.15	30028.00	1265.042
353.15	104.00	1018.903
353.15	5034.00	1055.962
353.15	10009.00	1090.768
353.15	15132.00	1124.22
353.15	20099.00	1154.783
353.15	25054.00	1183.626
353.15	30493.00	1213.605

Reference *https://www.doi.org/10.1016/j.jct.2015.10.006*

Temperature, K	Pressure, kPa	Speed of sound, m/s
249.99	100.00	1382.6
218.20	102.00	1512.0
299.97	138.00	1199.6
249.99	177.00	1383.0
350.05	270.00	1030.8
400.05	307.00	847.6
218.26	875.00	1515.2
400.05	947.00	854.9
350.07	997.00	1036.3
450.12	1081.00	636.4
249.99	1128.00	1387.7
299.96	1174.00	1206.0
350.07	2073.00	1044.4
218.21	2111.00	1520.8
400.05	2155.00	867.0
450.12	2181.00	653.9
250.01	2296.00	1393.4
299.96	2560.00	1214.3
500.31	4798.00	438.4
400.03	4879.00	893.3
218.26	5018.00	1533.1
299.98	5042.00	1228.8

450.11	5172.00	697.0
350.07	5244.00	1067.5
249.97	5362.00	1408.3
249.94	9322.00	1427.1
218.32	9465.00	1551.0
299.98	9515.00	1254.1
500.32	9558.00	542.3
350.06	9697.00	1098.3
400.07	9768.00	936.6
450.11	9782.00	754.2
218.35	19810.00	1591.6
299.96	19849.00	1308.7
500.32	19968.00	690.2
400.07	19974.00	1015.3
350.08	20259.00	1164.3
450.10	20385.00	860.0
249.90	20903.00	1478.7
400.06	27823.00	1068.3
218.34	29071.00	1626.6
450.11	30113.00	938.6
399.98	30784.00	1087.2
500.31	30851.00	799.1
299.96	39521.00	1400.6
249.93	39564.00	1553.9
500.28	39800.00	870.0
450.13	40189.00	1008.7
399.96	40232.00	1142.9
350.04	40537.00	1273.0
218.31	49499.00	1698.1
299.96	49960.00	1444.6
450.13	50091.00	1069.5
399.98	50284.00	1196.8
500.30	50346.00	946.6
350.03	52186.00	1327.6
249.94	59448.00	1625.8
399.97	59616.00	1242.9
500.31	60509.00	1009.7
350.04	60588.00	1364.3
218.32	61345.00	1736.4
500.31	77533.00	1102.4

Reference *https://www.doi.org/10.1021/acs.jced.8b00938*

Pressure, kPa	**Temperature, K**	**Speed of sound, m/s**

100.00	293.19	1222.98
100.00	298.12	1205.93
100.00	303.12	1188.72
100.00	308.20	1171.37
100.00	313.17	1154.51
100.00	318.20	1137.48
15200.00	293.12	1308.47
15200.00	298.11	1292.86
15200.00	303.14	1277.31
15200.00	308.20	1263.59
15200.00	313.16	1248.7
30400.00	293.22	1379.93
30400.00	298.15	1365.6
30400.00	303.24	1350.89
30400.00	308.13	1336.85
30400.00	313.16	1322.58
30400.00	318.16	1308.73
45590.00	293.21	1444.18
45590.00	298.14	1430.65
45590.00	303.13	1417.04
45590.00	308.13	1403.65
45590.00	313.16	1390.41
45590.00	318.14	1377.4
60790.00	293.21	1502.84
60790.00	298.14	1490.03
60790.00	303.20	1477.02
60790.00	308.13	1464.46
60790.00	313.17	1451.85
60790.00	318.12	1439.64
75990.00	293.20	1557.22
75990.00	298.14	1544.87
75990.00	303.15	1532.53
75990.00	308.12	1520.4
75990.00	313.16	1508.49
75990.00	318.12	1496.85
91190.00	293.20	1607.57
91190.00	298.14	1595.85
91190.00	303.17	1584.03
91190.00	308.20	1572.36
91190.00	313.15	1561.08
91190.00	318.12	1549.89
101320.00	293.18	1639.47
101320.00	298.14	1628.0
101320.00	303.17	1616.49

101320.00	308.19	1605.09
101320.00	313.15	1594.14
101320.00	318.11	1583.33
111450.00	293.17	1669.92
111450.00	298.14	1658.78
111450.00	303.19	1647.55
111450.00	308.16	1636.69
111450.00	313.22	1625.76
111450.00	318.17	1615.27
121580.00	293.16	1699.12
121580.00	298.14	1688.24
121580.00	303.18	1677.36
121580.00	308.11	1666.98
121580.00	313.17	1656.22
121580.00	318.17	1645.95

Reference *https://www.doi.org/10.1021/je030136n*

Molar volume, m3/mol

Temperature, K - Liquid	Pressure, kPa - Liquid	Molar volume, m3/mol - Liquid
298.15	100.00	0.0001
298.15	10000.00	0.0001
313.15	100.00	0.0001
313.15	10000.00	0.0001
328.15	100.00	0.0001
328.15	10000.00	0.0001

Reference *https://www.doi.org/10.1021/je800334m*

Molar heat capacity at constant volume, J/K/mol

Temperature, K - Liquid	Pressure, kPa - Liquid	Molar heat capacity at constant volume, J/K/mol - Liquid
314.15	4353.00	130.947
315.15	5217.00	131.0071
316.15	6081.00	131.7283
317.15	6946.00	132.6297

318.15	7808.00	133.2307
319.15	8666.00	133.2908
320.15	9526.00	133.6513
321.15	10392.00	134.6128
322.15	11260.00	134.9133
323.15	12130.00	135.8147
324.15	13000.00	135.8748
325.15	13871.00	136.3556
326.15	14740.00	137.4974
327.15	15599.00	136.9566
328.15	16461.00	139.1801
329.15	17326.00	138.519
330.15	18193.00	138.6993
331.15	19066.00	139.9012
332.15	19945.00	139.6007
333.15	20825.00	140.1416
334.15	21705.00	140.8026
335.15	22585.00	141.4637
336.15	23467.00	141.644
337.15	24350.00	141.2834
338.15	25234.00	142.5454
339.15	26119.00	142.7858
340.15	27006.00	143.7473
341.15	27893.00	143.0862
342.15	28774.00	145.0093
343.15	3874.00	148.615
343.15	29657.00	143.9276
344.15	4675.00	146.9924
345.15	5478.00	149.3361
346.15	6283.00	148.1943
347.15	7088.00	149.5164
348.15	7896.00	148.8554
349.15	8704.00	149.9972
350.15	9514.00	150.7784
351.15	10326.00	149.9972
352.15	11138.00	151.3193
353.15	11953.00	151.7399
354.15	12769.00	152.4611
355.15	13587.00	151.6798
356.15	14406.00	152.5813
357.15	15227.00	152.9418
358.15	16050.00	153.4827
359.15	16870.00	153.4226
360.15	17670.00	155.3456

361.15	18489.00	154.1437
362.15	19314.00	155.5259
363.15	20144.00	154.9851
364.15	20977.00	154.9851
365.15	21811.00	155.4057
366.15	22647.00	156.6677
367.15	23484.00	156.0067
368.15	24324.00	156.0668
369.15	25165.00	156.9682
370.15	26008.00	156.3071
371.15	26853.00	157.0884
372.15	2876.00	161.3551
372.15	27699.00	156.5475
373.15	3622.00	161.9561
373.15	28548.00	157.5691
374.15	4369.00	162.1965
374.15	29398.00	157.8095
375.15	5118.00	162.4969
376.15	5868.00	163.158
377.15	6620.00	163.158
378.15	7374.00	162.8575
379.15	8129.00	163.2181
380.15	8887.00	163.6387
381.15	9646.00	163.7589
382.15	10407.00	162.8575
383.15	11170.00	164.6003
384.15	11935.00	162.7373
385.15	12702.00	164.4801
386.15	13470.00	163.819
387.15	14240.00	164.2397
388.15	15013.00	164.1796
389.15	15786.00	164.2998
390.15	16562.00	164.2397
391.15	17338.00	164.8406
392.15	18116.00	163.4585
393.15	18896.00	164.3599
394.15	19677.00	164.2998
395.15	20459.00	164.3599
396.15	21244.00	164.5402
397.15	22030.00	164.6604
398.15	22817.00	164.7204
399.15	23606.00	164.8406
400.15	24395.00	163.9993
401.15	25186.00	164.3599

402.15	25978.00	164.2397
403.15	26772.00	163.7589
404.15	27566.00	163.819
405.15	28362.00	163.0979
406.15	29159.00	163.5786
414.15	10151.00	171.2708
415.15	10841.00	173.7948
416.15	11533.00	172.5929
417.15	12227.00	172.7732
418.15	12922.00	173.0737
419.15	13618.00	171.6314
420.15	14316.00	172.5328
421.15	15015.00	172.2924
422.15	15715.00	172.052
423.15	16417.00	172.3525
424.15	17119.00	172.4126
427.15	19232.00	172.1121
428.15	19938.00	171.4511
429.15	20645.00	171.4511
430.15	21353.00	172.052

Reference *https://www.doi.org/10.1021/je034101z*

Sources

Thermodynamic Properties of Mixtures Containing Ionic Liquids. 9. Activity	https://www.doi.org/10.1021/je050440b
Coefficients at Infinite Dilution of	https://www.doi.org/10.1021/acs.jced.5b00823
Hydrocarbons, Alcohols, Esters, and	https://www.doi.org/10.1021/acs.jced.7b00714
Alkanes in	https://www.doi.org/10.1021/acs.jced.9b00190
Trimethylbutylammonium	https://www.doi.org/10.1016/j.jct.2016.01.022
Solubility Determination and	https://www.doi.org/10.1021/acs.jced.8b01062
Measurement and Correlation of the	https://www.doi.org/10.1016/j.jct.2013.09.005
Solubility of Mifepristone in Eight Pure	https://www.doi.org/10.1021/je100483d
Liquid–Liquid Equilibrium Data for the	https://www.doi.org/10.1016/j.jct.2016.10.029
Solubility and thermodynamic	https://www.doi.org/10.1016/j.jct.2016.11.032
Liquid–Liquid Equilibrium Data for the	https://www.doi.org/10.1021/acs.jced.8b00776
Excess Molar Enthalpies of Propyl	https://www.doi.org/10.1021/je050052+
Thermodynamic Properties of Mixtures	https://www.doi.org/10.1021/je0503554
Measurement and Correlation of	https://www.doi.org/10.1021/je400813d
Solubilities of Indole-2-carboxylic Acid	https://www.doi.org/10.1021/je501045s
Solubility of 3,4-Dimethoxybenzoic	https://www.doi.org/10.1016/j.jct.2018.05.018

Thermodynamic Functions for the Solubility of 3-Nitrobenzonitrile in 12 Organic Solvents from T/K = (278.15 to 313.15)	https://www.doi.org/10.1021/acs.jced.7b00615
Activity coefficients at infinite dilution measurements for organic solutes and water in the ionic liquid	https://www.doi.org/10.1016/j.jct.2011.02.012
Limiting activity coefficient measurements in binary mixtures of [Emim][MeSO4] and correlation of Abraham model solute parameters	https://www.doi.org/10.1016/j.fluid.2015.11.037
Measurement and Correlation of Solubilities of (Z)-2-(2-Aminothiazol-4-yl)-2-methoxyiminoacetic Acid in Different Pure Solvents and Binary Mixtures of Water + (Ethanol, Methanol, or Acetone)	https://www.doi.org/10.1021/je101351w
1-Butanol: Webbook.	http://webbook.nist.gov/cgi/cbook.cgi?ID=C71238&Units=SI
The liquid-liquid phase separation and crystallization of vanillin in 1-propanol/water solution.	https://www.doi.org/10.1016/j.fluid.2015.09.011
Liquid-Liquid Equilibrium Data of Propan-1-ol + 2,2,4-Trimethylpentane and Butan-1-ol + 2,2,4-Trimethylpentane at 318.15 K	https://www.doi.org/10.1021/je200279q
Physical Properties of the Pure Deep Eutectic Solvent [ChCl][ev] (1:2) and Its Binary Mixtures with Alcohols	https://www.doi.org/10.1021/acs.jced.6b00563
Liquid Phase Equilibrium and Thermodynamic Analysis of (Vapour + liquid) equilibria for the binary mixtures of propan-1-ol +	https://www.doi.org/10.1021/acs.jced.9b00854
1,1-Diethylurea in Different Solvent Systems at dibromomethane, or +	https://www.doi.org/10.1016/j.jct.2004.07.012
Measurement and Correlation of Solubility of Clemizole Hydrogen Sulfate (Metastable Form) in Lower	https://www.doi.org/10.1021/je200067e
Vapor−Liquid Equilibria for the ternary system dipropyl ether + 1-propanol +	https://www.doi.org/10.1016/j.fluid.2006.07.003
Volumetric and Viscometric Study of Binary Mixtures of Dipropyl Ether + 1-Propanol, + 1-Pentanol at (303.15, 308.15, and 313.15) K	https://www.doi.org/10.1021/je049912x
Solubility and Preferential Solvation of Carbamazepine in Propan-1-ol + Water Mixtures at 298.15 K	https://www.doi.org/10.1021/acs.jced.7b00982
Solubility and Data Correlation of Methanol, Ethanol, 1-Propanol, and 2-Propanol in Different Monosolvents and Binary Mixtures of	https://www.doi.org/10.1021/acs.jced.9b00703
Thermodynamic properties of Mixtures Containing Ionic Liquids. 6. Activity Coefficients at Infinite Dilution of	https://www.doi.org/10.1021/je050134y
Thermodynamic solubility modeling, solvent effect and preferential solvation	https://www.doi.org/10.1016/j.jct.2018.11.022
Measurement and Correlation of Solubility of Resorcinol in Pure and Binary Solvent Mixtures of Ethanol + Water, Ethanol + Ethyl Acetate, and Water + Ethyl Acetate from T = 278.15 K to T = 323.15 K	https://www.doi.org/10.1021/acs.jced.7b00695
Solubility of Bis(oxalato)boric Acid and Solubility Equilibrium Constant in Water + Ethanol Mixtures at Different Temperatures	https://www.doi.org/10.1016/j.fluid.2004.11.018
Enthalpy of solution of 1,2-dichloroethane in 1-propanol, 2-propanol, 1-butanol, 2-methyl-1-propanol, and 2-methyl-2-propanol at T = 298.15 K	https://www.doi.org/10.1016/j.jct.2015.04.034
Modeling with COSMO-RS and UNIFAC: 1-Ethylpyridinium Ethylsulfate and its Binary Mixtures with Ethanol and	https://www.doi.org/10.1021/je800981d
Thermochemistry of Etherified Ethers: Experimental Study of Chemical Equilibrium in the Reacting System of	https://www.doi.org/10.1021/je034172y
Compressibility Studies of Binary Solutions Involving Water as a Solute	https://www.doi.org/10.1021/je0601098
Measurement of infinite dilution activity coefficients of 1-alkanols (C1 C5) in	https://www.doi.org/10.1016/j.fluid.2005.08.018
Measurement and modeling of high-pressure (vapour + liquid)	https://www.doi.org/10.1016/j.jct.2009.11.015
Solubility Measurement and Correlation of Phenacetin in 12 Pure Organic Solvents and Thermodynamic	https://www.doi.org/10.1021/acs.jced.8b00863
Properties of Phenformin in 14 Pure Solvents at Temperatures Ranging	https://www.doi.org/10.1021/acs.jced.9b00844
Solubility and thermodynamic properties of vortioxetine in nine pure solvents and (ethanol + water) binary	https://www.doi.org/10.1021/acs.jced.8b00663
Naphthalene Solubility in Binary Solvent Mixtures of	https://www.doi.org/10.1016/j.fluid.2015.03.027
Estimation and confirmation of the thermodynamic stability relationships of the enantiotropic polymorphs of	https://www.doi.org/10.1016/j.jct.2018.03.007
Atmospheric Ternary Liquid-Liquid Equilibrium for the Diethyl Carbonate + 1-Propanol + Water System at	https://www.doi.org/10.1021/je7005049
Separation of thiophene from heptane with ionic liquids.	https://www.doi.org/10.1016/j.jct.2017.10.011
Solubility of 2-(4-Isobutylphenyl)propionic Acid in 12 Pure Solvents from T = (278.15 to 323.15) K	https://www.doi.org/10.1021/je0499519
Physicochemical properties of systems of interest to the edible oil industry: Measurement, correlation, and	https://www.doi.org/10.1021/acs.jced.8b00937
Viscosities and densities of model systems for the determination of vapor pressure for binary	https://www.doi.org/10.1016/j.jct.2015.01.003
Solubility Determination and Modeling of Gliclazide in Fatty Acid + Solvents	https://www.doi.org/10.1016/j.jct.2013.01.033
Correlation of Glibenclamide in 11 Monosolvents and Modeling of Its Solubility in Ethanol +	https://www.doi.org/10.1021/acs.jced.8b00139
Solubility and Modeling of Hesperidin in Cosolvent Mixtures of Ethanol, Methylbutanol, and Acetonitrile	https://www.doi.org/10.1016/j.jct.2017.06.012
Measurement and Correlation of Solubility of Ethyl Paraben in Ethanol + Propan-1-ol or + Propan-2-ol.	https://www.doi.org/10.1016/j.fluid.2014.10.028
Binary Mixtures of Ethanol + Water: Measurement and Correlation	https://www.doi.org/10.1021/acs.jced.8b00717
Solubility and Modeling of Azoxystrobin in Seven Monosolvents and Two Different Binary Mixed Solvents	https://www.doi.org/10.1021/acs.jced.7b00948
Solubilities of Organic Semiconductors and Nonsteroidal Anti-inflammatory Drugs in Pure and Mixed Organic Solvents: Measurement and Modeling with Hansen Solubility Parameter:	https://www.doi.org/10.1021/acs.jced.7b00669
	https://www.doi.org/10.1021/acs.jced.8b00536

Solubility determination and thermodynamic modeling of 2,4-dinitroaniline in nine organic solvents from 278.15 to 313.15 K.	https://www.doi.org/10.1016/j.jct.2016.07.009
Experimental and Computational Thermodynamic Properties of Pyrene, 9,10-Dimethylanthracene, and 1,2,3,4-Tetrahydrocarbazole in Pure Solvents.	https://www.doi.org/10.1021/acs.jced.5b00162
Thermodynamic investigation of the mixing properties of solutions: ATPSs composed of some (aliphatic alcohol + sodium carbonate + water) ternary systems.	https://www.doi.org/10.1016/j.jct.2012.07.021
Liquid-Liquid Equilibria for Aliphatic Alcohols + Dipotassium Oxalate + Water.	https://www.doi.org/10.1021/je9008188
Volumetric properties of the boldine + alcohol mixtures at atmospheric pressure from 283.15 to 333.15K. A new method for the determination of the density of pure boldine. Several parameters of ionic liquid 1-propyl-3-methylimidazolium bromide in 1-hexadecyl-3-methylimidazolium tetrafluoroborate ionic liquid.	https://www.doi.org/10.1016/j.fluid.2007.04.030
Volumetric, acoustic and viscometric investigation of ceftriaxone disodium in aqueous solutions of 1-propanol and 2-propanol.	https://www.doi.org/10.1021/acs.jced.7b00299
Thermodynamic interactions in binary mixtures of anisole with ethanol, propan-1-ol, propan-2-ol, butan-1-ol, pentan-1-ol, and 3-methylbutan-1-ol at T = (298.15, 308.15, and 318.15) K.	https://www.doi.org/10.1016/j.jct.2018.09.023
Solubility of Phenanthrene in Binary Mixtures of C1-C4 Alcohols at 298.2 K:	https://www.doi.org/10.1016/j.jct.2007.01.004
Solubility of oxcarbazepine in eight solvents within the temperature range of (288.15 to 323.15) K.	https://www.doi.org/10.1016/j.jct.2012.09.006
Vapor-Liquid Equilibrium Behaviors of 3-Ethoxy-4-hydroxybenzaldehyde in Alcohol.	https://www.doi.org/10.1016/j.jct.2019.105880
Measurement and Correlation of the Solubility of Baicalin in Several Mixed Solvents.	https://www.doi.org/10.1021/je900599q
Prediction of different thermodynamic properties for systems of alcohols and hydrocarbons in pure alkanols at several temperatures.	https://www.doi.org/10.1016/j.jct.2006.03.018
Solubility of triclosan and iodosuccinyl sulfate based anionic ionic liquids using n-alkanol as a probe.	https://www.doi.org/10.1016/j.jct.2017.05.037
Liquid-Liquid Equilibria of Quaternary Systems Composed of 1,3-Propanediol, Short-Chain Alcohol, Water, and Salt:	https://www.doi.org/10.1016/j.jct.2016.09.011
Solubility determination and thermodynamic modelling of 3,5-dinitropyrazole in nine organic solvents from T = (283.15 to 313.15) K.	https://www.doi.org/10.1021/je700316s
Solubility and Thermodynamic Behavior of Veramoss in Different Pure and Mixing Solvents.	https://www.doi.org/10.1021/acs.jced.8b01186
Phase Diagrams of Some Aliphatic Alcohols + Potassium Citrate + Water at 25 deg C.	https://www.doi.org/10.1016/j.fluid.2012.11.003
Liquid-Liquid Equilibrium for Ternary Systems of Propyl Vinyl Ether + C3 or C4 Alcohols + Water at 298.15 K and Atmospheric Pressure.	https://www.doi.org/10.1016/j.fluid.2012.05.020
Density and Viscosity of Binary Mixture of the Ionic Liquid 1-Butyl-3-methylimidazolium Tetrafluoroborate with Methanol, Ethanol, 1-Propanol, and 1-Butanol at (293.15, 303.15, 313.15, and 323.15) K.	https://www.doi.org/10.1021/acs.jced.6b00472
Excess Molar Enthalpies for Ternary Mixtures Containing Dimethyl Carbonate + Methanol + Ethanol at 298.15 K.	https://www.doi.org/10.1016/j.jct.2017.02.011
Liquid-Liquid Equilibria for Ternary Systems of Propanol with Methanol, Ethanol, and Water at 298.15 K.	https://www.doi.org/10.1021/acs.jced.6b00646
Isobaric Vapor-Liquid Equilibrium for Binary Mixtures of Ethanol + 1-Propanol and Methanol + 1-Propanol at 101.3 kPa.	https://www.doi.org/10.1021/je050086h
Refractive Index and Viscosity of (Propyl Vinyl Ether + Methyl tert-Amyl Ether (TAME) + Methanol) Ternary Mixtures at (293.15 to 313.15) K.	https://www.doi.org/10.1021/je7005743
Measurement and Correlation of Solubility of Nitrofurantoin in Different Pure Solvents from 283.15 to 323.15 K.	https://www.doi.org/10.1021/je100848r
Solubility Measurement and Correlation of Nitrofurazone in Different Pure Solvents from 283.15 to 323.15 K.	https://www.doi.org/10.1016/j.jct.2017.05.004
Solubility Measurement and Correlation of Nitrofurazone in Different Pure Solvents from 283.15 to 323.15 K.	https://www.doi.org/10.1016/j.tca.2013.03.025
Isobaric Vapor-Liquid Equilibrium for Binary Mixtures of Propyl Vinyl Ether + Ethanol and Propyl Vinyl Ether + 1-Propanol at 101.3 kPa.	https://www.doi.org/10.1021/je8006698
Refractive Index and Viscosity of (Propyl Vinyl Ether + Methyl tert-Amyl Ether (TAME) + Methanol) Ternary Mixtures at (293.15 to 313.15) K.	https://www.doi.org/10.1016/j.jct.2019.07.017
Solubility determination and thermodynamic properties of metronidazole benzoate in different pure solvents.	https://www.doi.org/10.1016/j.jct.2014.08.026
Solubility of Imidazole Antifungal Drugs in Different Pure Solvents from 278.15 to 323.15 K: Determination, Correlation, and Thermodynamic Dissolution Properties.	https://www.doi.org/10.1021/je100577v
Solubility determination and thermodynamic modeling of 6-chloro-3-hydroxy-2-pyrazinecarboxamide (favipiravir) in pure solvents.	https://www.doi.org/10.1016/j.jct.2019.06.007
Measurement and correlation of critical properties for binary mixtures and ternary mixtures containing gasoline additives.	https://www.doi.org/10.1016/j.jct.2014.01.025
Isobaric Vapor-Liquid Equilibria of the 1,3,3-Trimethyl-2-oxabicyclo[2.2.2]octane + Ethanol + 1-Propanol Binary Systems.	https://www.doi.org/10.1016/j.jct.2005.05.004
Measurement and correlation of critical properties for binary mixtures and ternary mixtures containing gasoline additives.	https://www.doi.org/10.1016/j.jct.2008.12.021
Solubility of Hydrogen Sulfide in Ethanol, 1-Propanol, and 2-Propanol.	https://www.doi.org/10.1021/acs.jced.5b00680
Excess Molar Enthalpies of Ternary and Binary Mixtures Containing 2-Methylpropan-2-ol, Methylacetate, and Methylpropanoate.	https://www.doi.org/10.1021/je8007606
Experimental Measurement and Modeling of Volumetric Properties and Speeds of Sound of Several Mixtures of Alcohol + Water.	https://www.doi.org/10.1021/je700690q
Activity coefficients at infinite dilution of organic solutes in 1-hexyl-3-methylimidazolium trifluoroacetate and refractive indices for binary mixtures of 1-hexyl-3-methylimidazolium trifluoroacetate with 1-propanol, 2-propanol, 1-butanol, and 2-butanol at T = (298.15 to 323.15) K.	https://www.doi.org/10.1016/j.fluid.2014.06.005
Solubility of Acenaphthene in Different Solvents from (283.00 to 323.00) K: Measurement and Correlation.	https://www.doi.org/10.1021/je8006869
Activity coefficients at infinite dilution of organic solutes in 1-hexyl-3-methylimidazolium trifluoroacetate and refractive indices for binary mixtures of 1-hexyl-3-methylimidazolium trifluoroacetate with 1-propanol, 2-propanol, 1-butanol, and 2-butanol at T = (298.15 to 323.15) K and Ambient Pressure.	https://www.doi.org/10.1016/j.jct.2013.10.038
Influence of interfacial adsorption during gas-liquid chromatography: Solubility of n-alkanes in 1-alkanol solvents at T = (298.15, 308.15, 318.15, and 328.15) K: Binary Systems of 1H-Pyrrole with Butan-1-ol, Propan-1-ol, or Pentan-1-ol :	https://www.doi.org/10.1016/j.jct.2012.08.030
	https://www.doi.org/10.1016/j.jct.2005.08.011
	https://www.doi.org/10.1021/je3004194

Vapor-Liquid Equilibrium of Binary Mixtures Containing Isopropyl Acetate and Alkanols at 101.32 kPa.	https://www.doi.org/10.1021/acs.jced.5b00360
Vapor Pressure Measurements of Liquid Solutions at Different Temperatures: Apparatus for Use over an Extended Temperature Range and Some Data.	https://www.doi.org/10.1021/je034197x
Thermodynamic behavior of binary mixtures of 1-ethyl-3-methylimidazolium ionic liquids with 1-propanol and 1-butanol.	https://www.doi.org/10.1016/j.tca.2012.09.009
Thermodynamic Functions of Solution of Benzoic Acid in Ethanol + 1-Propanol between 293 K and 333 K.	https://www.doi.org/10.1021/je0499768
Determination of Activity Coefficients at Infinite Dilution of Solutes in the Ionic Liquid, 1-Ethyl-3-methylimidazolium Ethyl Sulfate, between 313 K and 353 K by Inverse Gas-Liquid Chromatography.	https://www.doi.org/10.1021/je800105r
Solubility Modeling and Solvent Effect for Ethambutol Dihydrochloride in 12 Neat Solvents: Determination and Correlation.	https://www.doi.org/10.1021/acs.jced.8b01126
Solid-Liquid Equilibrium of Bis(trifluoromethylsulfonyl) Imide, Rebaudioside A in Pure and Binary Mixed Solvents at T = (278.15 to 328.15) K.	https://www.doi.org/10.1021/acs.jced.8b00780
Liquid Chromatography at T = (283.15, 298.15, 313.15, and 328.15) K.	https://www.doi.org/10.1021/je7006633
Solubilities of cerium(III) nitrate in (methanol, + ethanol, + 1-propanol, and + 2-propanol) at different temperatures and Pressures: Isothermal (vapor + liquid) equilibria for the binary mixtures of (propylene oxide + ethanol) and (propylene oxide + 1-propanol) at several temperatures.	https://www.doi.org/10.1016/j.fluid.2010.08.006
Solubility of cerium(III) nitrate in aqueous alcohol mixtures at atmospheric pressure and different temperatures.	https://www.doi.org/10.1016/j.jct.2015.02.006
Solubility Measurement of Lauric, Palmitic, and Stearic Acids in Ethanol, 1-Propanol, and 2-Propanol Using Differential Scanning Calorimetry.	https://www.doi.org/10.1021/acs.jced.8b01044
Solubility of Beauvericin in pure solvent alcohols and water with methanol mixture solvents at atmospheric pressure and different temperatures.	https://www.doi.org/10.1016/j.fluid.2017.12.024
Thermodynamic Properties of Mixtures Containing Ionic Liquids.4.LLE of Binary Mixtures of [bzMIM][NTf2] with 2-Propanol, 1-Butanol, and 1-Hexanol.	https://www.doi.org/10.1021/je020228x
Isothermal vapor-liquid equilibria at 303.15K and excess molar volumes at 298.15K for the ternary mixtures 1-propanol + 1-butanol + 1-ethyl-3-methylimidazolium ethylsulfate ionic liquid:	https://www.doi.org/10.1021/je049564z
Solubility of Chlorpheniramine Maleate in Ethanol, 1-Propanol, 1-Butanol, Benzene, Ethyl Acetate, Ethyl Formate, and Butyl Acetate between 288.15 K and 323.15 K.	https://www.doi.org/10.1016/j.fluid.2007.07.002
Measurement and correlation of solubility of ciclesonide in seven pure organic solvents.	https://www.doi.org/10.1016/j.jct.2018.07.013
Solubility of ternary mixtures of (oxygenated additives + cycloalkane) in fuels and biofuels: (diisobutyl ether + methylcyclohexane) at T = 298.15 and 313.15 K.	https://www.doi.org/10.1016/j.jct.2013.01.032
Thermodynamics and limiting activity coefficients measurements for organic solutes and water in 1-(2-methoxyethyl)-3-methylimidazolium dicyanamide and 1-(2-methoxyethyl)-1-methylpyrrolidinium dicyanamide ionic liquids.	https://www.doi.org/10.1021/acs.jced.8b01080
Solubility Modeling of Ionic Liquid Clotrimazole Form I.	https://www.doi.org/10.1021/je7001682
Measurement and correlation of solubility of cyclesonide in seven pure organic solvents.	https://www.doi.org/10.1016/j.jct.2016.10.014
Solubility of ternary mixtures of (oxygenated additives + cycloalkane) in fuels and biofuels: (diisobutyl ether + methylcyclohexane) at T = 298.15 and 313.15 K.	https://www.doi.org/10.1016/j.jct.2016.10.005
Thermodynamics and limiting activity coefficients measurements for organic solutes and water in 1-(2-methoxyethyl)-3-methylimidazolium dicyanamide and 1-(2-methoxyethyl)-1-methylpyrrolidinium dicyanamide ionic liquids.	https://www.doi.org/10.1016/j.jct.2016.08.008
Solubility Modeling of Ionic Liquid 4-(Dimethylamino)benzaldehyde in Nine Monosolvents and in Some Binary Solvent Mixtures.	https://www.doi.org/10.1021/acs.jced.5b01053
Solubility and transport properties of the ionic liquid 1-ethyl-3-methylimidazolium bis(trifluoromethylsulfonyl)imide.	https://www.doi.org/10.1016/j.tca.2008.11.007
Liquid-liquid phase equilibrium of ternary mixtures containing 1-propanol + ethyl butanoate + cyclohexane or benzene at 303.15 K: Experimental data, correlation and prediction by COSMO-SAC model.	https://www.doi.org/10.1016/j.fluid.2012.05.003
Solubility of glyphosate in selected solvents: Measurements and derived properties.	https://www.doi.org/10.1016/j.jct.2011.10.023
Isobaric (vapor-liquid) equilibria for binary systems of methanol + 1-(methoxymethoxy)-propane and 2-propanol + 1-(methoxymethoxy)-propane at 101.33 kPa.	https://www.doi.org/10.1016/j.fluid.2015.07.017
Solubility of Benzoin in Six Monosolvents and in Some Binary Solvent Mixtures of Ethanol + Ethyl Acetate at 101.33 kPa from 278.15 to 323.15 K.	https://www.doi.org/10.1021/acs.jced.7b00238
Isobaric Vapor-Liquid Equilibrium for Methyl Isobutyl Ketone (MIBK) + (1-Propanol or 2-Propanol) Binary Mixtures.	https://www.doi.org/10.1021/acs.jced.7b00026
Liquid-Liquid Phase Equilibrium of Some Aliphatic Alcohols + Cesium Sulfate + Water.	https://www.doi.org/10.1021/je049914h
Thermodynamics of 1,3-dimethylurea in eight alcohols:	https://www.doi.org/10.1021/acs.jced.5b00882
Densities and Viscosities of Binary Mixtures of 2,2,4-Trimethylpentane + 1-Propanol, + 1-Pentanol, + 1-Hexanol, and + 1-Heptanol from (298.15 to 323.15) K.	https://www.doi.org/10.1021/je501041r
Volumetric, isentropic compressibility and electrical conductivity of solutions of aromatic amino acids in aqueous 1-propanol at 298.15 K: Experimental and computational investigation of excess properties.	https://www.doi.org/10.1016/j.fluid.2008.01.004
Determination and correlation of the solubility of four Brønsted-acidic ionic liquids based on benzotriazolium cations in six alcohols.	https://www.doi.org/10.1016/j.jct.2014.09.001
Equation of State with Association: 3-chloro-N-phenylphthalimide in ten organic solvents from 278.15 to 323.15 K: Determination and Correlation of Solubility.	https://www.doi.org/10.1016/j.jct.2013.12.031
Measurement of the Solubility of Diclazuril and Discussion of the Effect of Different Solvent.	https://www.doi.org/10.1021/je100595s
Solubilities of (S)-Indoline-2-carboxylic Acid in Six Different Solvents from 283.15 to 338.15 K.	https://www.doi.org/10.1016/j.jct.2016.01.003
Vapor-Liquid Equilibria Measurements for the Dehydration of Isopropanol, and n-Propanol via Azeotropic Distillation Using DIPE and Isooctane as Entrainers.	https://www.doi.org/10.1021/acs.jced.9b00353
Activity coefficient at infinite dilution measurements for organic solutes in the ionic liquid tributylmethylphosphonium methylsulfate from T = (293.15 to 343.15) K.	https://www.doi.org/10.1021/je400188h
Vapor-Liquid Equilibria for Mixtures of Polar and Non-Polar n-Fatty Compounds: Saturated fatty acids + Ethanol or 1-Propanol.	https://www.doi.org/10.1016/j.jct.2012.06.009
Solubility of Tris(hydroxymethyl)aminomethane in Water, Methanol, Ethanol, and 1-Propanol at Various Temperatures.	https://www.doi.org/10.1021/je300847v
Thermodynamics and activity coefficients at infinite dilution for organic solutes, water and diols in the ionic liquid choline bis(trifluoromethylsulfonyl)imide:	https://www.doi.org/10.1021/je500620m
	https://www.doi.org/10.1016/j.jct.2014.04.024

A 1-alkylcyanopyridinium-based ionic liquid in the separation processes:	https://www.doi.org/10.1016/j.jct.2016.01.017
Paramagnetic Ionic Liquid	https://www.doi.org/10.1021/je800846j
1-Butyl-3-methylimidazolium Tetrabromidocobaltate(II): Activity	https://www.doi.org/10.1016/j.fluid.2006.01.008
Coefficients at Infinite Dilution of Solutes and Crystal Structure!	https://www.doi.org/10.1016/j.jct.2016.09.038
Ethyl tert-butyl ether +	https://www.doi.org/10.1016/j.jct.2016.10.020
pharmaceutical in nine organic solvents from = 278.15 to 318.15 K and	https://www.doi.org/10.1021/je700092d
Isobaric Vapor–Liquid Equilibria for	https://www.doi.org/10.1021/je300401c
three ternary systems of Water + 1-Propanol +	https://www.doi.org/10.1021/acs.jced.8b00288
2-Propanol +	https://www.doi.org/10.1021/je201129y
Ethyl 3-methylimidazolium	https://www.doi.org/10.1016/j.fluid.2011.03.017
Tetrafluoroborate, and Water +	https://www.doi.org/10.1016/j.jct.2017.02.008
1-Propanol + 2-Propanol at	https://www.doi.org/10.1021/je500222s
1-Butyl-3-methylimidazolium	https://www.doi.org/10.1016/j.fluid.2018.04.025
Tetrafluoroborate + 2-nitroaniline in	https://www.doi.org/10.1016/j.jct.2014.01.030
fourteen organic solvents from T =	https://www.doi.org/10.1021/acs.jced.8b00600
(278.15 to 313.15) K and Mixing	https://www.doi.org/10.1016/j.fluid.2006.04.010
Properties of Solutions:	https://www.doi.org/10.1021/acs.jced.9b00458
Molecular dynamic simulation,	https://www.doi.org/10.1016/j.tca.2019.178383
molecular interactions and structural	https://www.doi.org/10.1016/j.jct.2011.11.025
Effect of the temperature on the	https://www.doi.org/10.1021/acs.jced.7b00301
properties of	https://www.doi.org/10.1016/j.fluid.2011.03.012
1-Butyl-3-methylimidazolium	https://www.doi.org/10.1016/j.jct.2005.11.008
Tetrafluoroborate + Ethanol + Water	https://www.doi.org/10.1016/j.jct.2017.07.022
Activity Coefficients at Infinite Dilution	https://www.doi.org/10.1016/j.jct.2013.03.014
of Various Solutes in the Ionic Liquid	https://www.doi.org/10.1016/j.fluid.2008.03.009
1-Butyl-3-methylimidazolium	https://www.cheric.org/research/kdb/hcprop/showprop.php?cmpid=819
Tetrafluoroborate + Ethanol +	https://www.doi.org/10.1016/j.jct.2013.06.018
Measurements of activity coefficients	https://www.doi.org/10.1021/je900504e
at infinite dilution for organic solutes	https://www.doi.org/10.1021/je050474j
in 1-hexadecyl-3-methylimidazolium	https://www.doi.org/10.1016/j.fluid.2013.11.031
tetrafluoroborate ionic liquid from	https://www.doi.org/10.1021/je0495942
293.15 K to 353.15 K:	https://www.doi.org/10.1016/j.jct.2008.09.004
Measurements and Correlation of	https://www.doi.org/10.1016/j.jct.2016.10.006
Solubility of Two Isomers of	https://www.doi.org/10.1016/j.fluid.2014.12.029
1-Butyl-3-methylimidazolium	https://www.doi.org/10.1016/j.fluid.2009.12.009
Tetrafluoroborate and Pure Solvents	https://www.doi.org/10.1016/j.jct.2012.12.003
from 283.18 K to 313.18 K:	https://www.doi.org/10.1021/acs.jced.9b00930
Measurements of activity coefficients	https://www.doi.org/10.1021/je800355e
at infinite dilution of organic solutes	https://www.doi.org/10.1016/j.jct.2011.02.018
in 1-hexyl-3-methylimidazolium	https://www.doi.org/10.1016/j.jct.2012.10.008
bis(trifluoromethylsulfonyl)imide +	https://www.doi.org/10.1021/acs.jced.7b00665
Measurement and Correlation of	https://www.doi.org/10.1016/j.fluid.2014.03.006
Solubility of Pravastatin Sodium in	https://www.doi.org/10.1021/je800196k
Water + Ethanol + 2-Propanol,	
1-Propanol, and 1-Butanol from (278 to	
333) K:	

Vapor Liquid Equilibrium Data for Binary Systems of n-Dodecane + {Propan-1-ol, Butan-1-ol,	https://www.doi.org/10.1021/je500237g
Thermodynamics and selectivity of separation based on activity	https://www.doi.org/10.1016/j.jct.2016.06.028
Study of the liquid-liquid equilibrium of various solutes in	https://www.doi.org/10.1016/j.fluid.2011.10.004
Measurement and correlation of solubility of trans-resveratrol in 11	https://www.doi.org/10.1016/j.jct.2007.10.006
Solubility of Ibuprofen in Different Pure Solvents and Binary	https://www.doi.org/10.1021/acs.jced.7b00743
Evaluation of the solvent structural effect upon the vapor-liquid equilibrium	https://www.doi.org/10.1016/j.fluid.2017.02.016
Determination and Correlation of Solubility of Cefradine Form I in Five	https://www.doi.org/10.1021/je400218d
Isothermal Vapor-Liquid Equilibrium Measurements for the Pentan-2-one +	https://www.doi.org/10.1021/acs.jced.9b00100
Solubility of N-phenylanthranilic acid in nine organic solvents from T = (283.15	https://www.doi.org/10.1016/j.jct.2016.08.017
Phase equilibrium in systems with ionic liquids: An example for the	https://www.doi.org/10.1016/j.fluid.2011.05.027
Solubility and preferential solvation of 2-methyl-4-nitroaniline in mixed	https://www.doi.org/10.1016/j.jct.2017.05.044
Acetic Acid Using Ionic Liquids	https://www.doi.org/10.1016/j.fluid.2019.01.005
Isothermal vapour-liquid equilibria for	https://www.doi.org/10.1016/j.jct.2009.06.023
1-butyl-3-methylimidazolium acetate	https://www.doi.org/10.1016/j.fluid.2010.03.008
Solubility Measurements at 296 and 310 K and Physicochemical	https://www.doi.org/10.1021/acs.jced.8b00566
Solubility and Thermodynamic Characterization of Abiraterone and	https://www.doi.org/10.1021/acs.jced.9b00411
Modeling of Sulfanilamide in 12 Mono Solvents and 4 Binary Solvent Mixtures	https://www.doi.org/10.1021/acs.jced.7b00110
Liquid-Phase Equilibria of Ternary (Solvent + Alcohol + Entrainer) Systems:	https://www.doi.org/10.1021/je200074c
Isobaric Vapor-Liquid Equilibria for the	https://www.doi.org/10.1021/je2008236
Thermodynamic functions for solubility	https://www.doi.org/10.1016/j.jct.2017.05.013
Correlation of solvents at temperatures	https://www.doi.org/10.1016/j.fluid.2013.05.008
Thermophysical properties of the binary mixtures (1,8-cineole +	https://www.doi.org/10.1016/j.jct.2009.08.016
Speed of Sound in Propan-1-ol + Heptane Mixtures under Elevated	https://www.doi.org/10.1021/je030136n
Binary Solid-Liquid Solubility Determination and Model Correlation of	https://www.doi.org/10.1021/acs.jced.8b01181
Measurement and Correlation of Solubility of Loratadine in Different Pure Solvents	https://www.doi.org/10.1021/acs.jced.6b00721
Boiling Temperatures of the Binary Mixtures of Dimethylcarbonate with	https://www.doi.org/10.1021/je030131q
Solubility of 2-isopropylimidazole in nine pure organic solvents and liquid	https://www.doi.org/10.1016/j.jct.2016.12.028
Measurement and Correlation of the Solubility of Ternary Mixture of	https://www.doi.org/10.1021/je401101u
Excess Properties of Ternary Mixture of Methanol + Ethyl Acetate +	https://www.doi.org/10.1021/acs.jced.8b00555
Solubility, Thermodynamic Modeling and Molecular Interactions of	https://www.doi.org/10.1021/acs.jced.8b00235
Excess molar enthalpy of 1-alkanol +	https://www.doi.org/10.1016/j.fluid.2007.07.063
Solubility of Vanillin in Different Solvents,	https://www.doi.org/10.1021/je8003244
Correlation, and Solvent Effect in the	https://www.doi.org/10.1016/j.jct.2013.06.002
Liquid-Liquid Extraction of Alcohols from Their Azeotropic Mixtures with	https://www.doi.org/10.1021/acs.jced.8b00362
Excess molar enthalpies of methyl isobutyl ketone (MIBK) with alkan-1-ols	https://www.doi.org/10.1016/j.tca.2004.09.012
Thermophysical and volumetric study of mixtures {p-cymene + propan-1-ol} at	https://www.doi.org/10.1016/j.jct.2018.03.021
Activity Coefficients at Infinite Dilution in Methylimidazolium Nitrate Ionic Liquids:	https://www.doi.org/10.1021/je100998r

Activity coefficients at infinite dilution measurements for organic solutes and water in the	https://www.doi.org/10.1016/j.fluid.2009.08.017
Tetramethylammonium chloride + glycerol deep eutectic solvents	https://www.doi.org/10.1016/j.fluid.2018.06.003
Isobaric Vapor–Liquid Equilibria at 101.32 kPa and Densities, Speeds of Sound, and Refractive Indices at 298.15 K for Binary Mixtures of 1-Propanol + 2,6-Dichloro-4-nitroaniline in 12 Pure Solvents at Temperatures from (293 to 328.15) K	https://www.doi.org/10.1021/je900274n
Measurement and correlation of solubility equilibrium for binary solvent mixtures	https://www.doi.org/10.1016/j.jct.2017.03.007
Thermodynamic properties of mixtures containing ionic liquids. Vapor	https://www.doi.org/10.1016/j.fluid.2005.07.008
Measurement and correlation of the solubility of 2,5-Dihydroxybenzoic Acid in alcohols and benzene at	https://www.doi.org/10.1021/acs.jced.6b00957
Separation of binary azeotropic mixtures with binary solvents	https://www.doi.org/10.1016/j.jct.2019.05.011
Isobaric Vapor–Liquid Equilibria for Systems Containing Alcohol or Acetic Acid and Water	https://www.doi.org/10.1021/acs.jced.6b00725
Viscosities and Densities of Binary Mixtures of (N-Acetylmorpholine + Ethanol, 1-Propanol, or 2-Propanol) at	https://www.doi.org/10.1021/je8003194
Measurement and correlation of the solubility of	https://www.doi.org/10.1016/j.tca.2016.01.009
Determination and modeling of ethanol (2-Ethoxyethyl) Phosphate + Water Binary Systems at 101.23 kPa:	https://www.doi.org/10.1021/acs.jced.8b00975
Measurement and correlation of the solubility of	https://www.doi.org/10.1016/j.jct.2016.09.015
Activity Coefficients at Infinite Dilution of Polar Solutes in Solvents at Temperatures from 292.15 K to 313.15 K:	https://www.doi.org/10.1021/je700560s
Isobaric vapor–liquid equilibrium data for binary systems of 1-propanol + 2-methylpentane and 1-propanol + 2,2,4-trimethylpentane at 101.3 kPa	https://www.doi.org/10.1016/j.fluid.2017.12.029
Activity Coefficients at Infinite Dilution Determined by Gas-Liquid Chromatography. Propanol + Squalane, Squalane + Sulfolane, and the	https://www.doi.org/10.1021/je100033s
Density, Viscosity, and Electrical Conductivity in Binary Potassium Sulfate + Water + (Ethanol, 1-Propanol, or 2-Propanol) from (278.15 to 308.15) K:	https://www.doi.org/10.1021/je800673j
Solubilities of Cefotaxime Hydrochloride in Water + 1-Propanol System at Different Temperatures	https://www.doi.org/10.1021/je100301p
Vapor Pressures and Activity Coefficients of n-Alcohols and Benzene in Binary Mixtures with	https://www.doi.org/10.1021/je050394f
Density, Viscosity, and Speed of Sound of Pure and Binary Mixtures of Ionic	https://www.doi.org/10.1021/acs.jced.8b00537
Activity Coefficients at Infinite Dilution of Organic Solutes and Water in the Ionic Liquid	https://www.doi.org/10.1021/je900704b
Measurement and correlation of solubility of Dihydrocholesterol in alcohol solvents and thermodynamic analysis of dissolution process	https://www.doi.org/10.1016/j.jct.2019.03.021
Activity Coefficients at Infinite Dilution Determined by Gas–Liquid Chromatography. Squalane as the Liquid Phase	https://www.doi.org/10.1021/je0502142
Determination and correlation of solubility and solution thermodynamics of Propanil in	https://www.doi.org/10.1016/j.jct.2019.02.005
Surface Tension and Density of Pure Ionic Liquids and Some Binary Mixtures with 1-Propanol and 2-Methyl-1-propanol:	https://www.doi.org/10.1021/je700621d
Densities and Viscosities of Binary Mixtures of 2-Ethyl-1,1,3,3-tetramethylguanidinium Ionic Liquids with Ethanol and 1-Propanol	https://www.doi.org/10.1021/acs.jced.5b00259
Solubility Measurement, Correlation, and Molecular Interactions of 3-Methyl-4-nitroimidazolium in Different Neat Solvents and Mixed Solvents from T = 278.15 to 328.15 K:	https://www.doi.org/10.1021/acs.jced.8b01256
Excess Molar Enthalpies of Diethyl Malonate + Methanol, + Ethanol, + 1-Propanol, and + 2-Propanol at 298.15 K	https://www.doi.org/10.1021/je800547k
Solubility of Dimethyl Fumarate in Methanol, Ethanol, 1-Propanol, 2-Propanol, and Water from 289.95 to 341.15 K	https://www.doi.org/10.1021/je9003806
Measurement and correlation of the solubility of Vanillic Acid in Eight Pure Solvents at T = 278.15 to 328.15 K	https://www.doi.org/10.1021/acs.jced.5b00619
Thermodynamic Functions of Solvents at Different Temperatures	https://www.doi.org/10.1021/acs.jced.7b00316
Solubility of 1-Hydroxybenzotriazole in Sixteen Solvents at Temperatures from (278.15 to 313.15) K and	https://www.doi.org/10.1021/je100964a
Phase Equilibrium, Excess Enthalpies, and Densities of Binary Mixtures of Propionic Acid + 1-propanol, and	https://www.doi.org/10.1016/j.fluid.2007.01.021
Density, viscosity, and derived thermodynamic properties of binary alkanol + holding solvent (1-propanol or	https://www.doi.org/10.1016/j.jct.2012.03.003
Activity coefficients at infinite dilution measurements for organic solutes and water in the ionic liquid	https://www.doi.org/10.1016/j.jct.2011.06.007
Solubilities of Ibuprofen in Different Pure Solvents at different temperatures:	https://www.doi.org/10.1016/j.fluid.2018.05.005
Solubility of Ethanol in 1-Propanol + Glycerol + Water Ternary Mixtures at Different Temperatures:	https://www.doi.org/10.1021/je100255z
Excess molar volumes and isentropic compressibility of binary systems (1-octyl-1-methylpyrrolidinium bis(trifluoromethylsulfonyl)imide +	https://www.doi.org/10.1016/j.fluid.2011.01.019
Activity Coefficients at Infinite Dilution of Organic Solutes in Tetraoctylammonium Bromide or Trihexyltetradecylphosphonium Chloride Ionic Liquids	https://www.doi.org/10.1016/j.jct.2008.03.007
Activity Coefficients at Infinite Dilution Determined by Gas-Liquid Chromatography:	https://www.doi.org/10.1016/j.fluid.2012.05.006
Solubility of Ethanol in 1-Propanol at Different Temperatures:	https://www.doi.org/10.1021/je1005517

Measurement and correlation of aqueous biphasic systems composed of alcohol (1-propanol, 2-propanol, 1-butanol, or 2-butanol) + H2O at 298.15 K and ammonium sulfate or potassium citrate tribasic monohydrate	https://www.doi.org/10.1016/j.fluid.2018.03.009
Determination of Physicochemical Parameters of Sodium Dodecyl Sulfate Micelles in Ethanol + Water Solutions	https://www.doi.org/10.1021/acs.jced.8b00181
Solubility determination and modelling of 4-nitro-1,2-phenylenediamine in eleven organic solvents from T = (283.15 to 313.15) K and thermodynamic properties of solutions	https://www.doi.org/10.1016/j.jct.2016.11.014
Solubility, Solvent Effect, and Thermodynamic Properties of 2-Amino-5-chloro-3-nitropyridine in Different Pure Solvents at Temperatures from 278.15 to 323.15 K	https://www.doi.org/10.1016/j.jct.2012.09.033
Solubility Determination and Thermodynamic Properties of 4-Hydroxybenzaldehyde in Different Pure Solvents from 278.15 to 323.15 K	https://www.doi.org/10.1016/j.jct.2012.02.007
Thermodynamic properties of binary mixtures containing 2-methoxyethanol and 1-alkanols at 298.15 K	https://www.doi.org/10.1016/j.fluid.2013.05.031
Measurement and Correlation of Solubility and Dissolution Thermodynamic Properties of 1H-1,2,4-Triazole-3-carboxylic Acid in Pure and Binary Solvents from 283.15 to 323.15 K	https://www.doi.org/10.1021/je500012b
Excess Volumes of Ternary Mixtures 2,2,4-Trimethylpentane + Diisopropyl Ether or Methyl tert-Butyl Ether + Methanol, Ethanol, or 1-Propanol at 298.15 K	https://www.doi.org/10.1016/j.jct.2010.02.001
Volumetric and Ultrasonic Studies of 1-Ethyl-3-methylimidazolium Ethylsulfate Ionic Liquid with Methanol, Ethanol, 1-Propanol, and Water at Several Temperatures	https://www.doi.org/10.1021/je050242k
Quantitative NMR spectroscopy of binary liquid mixtures (aldehyde + alcohol) Part I: Acetaldehyde + (methanol or ethanol or 1-propanol)	https://www.doi.org/10.1021/je2011314
Activity coefficients at infinite dilution of organic solutes in N-alkylpyridinium bis(trifluoromethylsulfonyl)amide ionic liquids by inverse gas–liquid chromatography	https://www.doi.org/10.1021/je8004485
Separation of water/butan-1-ol based on activity coefficients at infinite dilution: Validation of the Mintsa Method	https://www.doi.org/10.1021/je7001804
Liquid–Liquid Equilibrium Behavior of 1-Methyl-2-(1-methylethyl)benzene + 1-Propanol + Water at 298.15 K	https://www.doi.org/10.1016/j.jct.2017.10.003
Determination and correlation of solubility of 11alpha,17alpha-dihydroxy-4-pregnene-3,20-dione in different solvents	http://link.springer.com/article/10.1007/BF02311772
Densities and volumetric properties of (formamide + ethanol, or 1-propanol, or 1,2-ethanediol, or 1,2-propanediol) mixtures at temperatures between 293.15 K and 313.15 K	https://www.doi.org/10.1016/j.jct.2013.03.022
Volumetric and transport properties of ternary mixtures containing 1-propanol, n-heptane and 2-methylbutan-2-ol or 3-methylbutan-1-ol at 298.15 K	https://www.doi.org/10.1016/j.jct.2010.06.009
Liquid–Liquid Equilibrium Behavior of 1-Methyl-2-(1-methylethyl)benzene + 1-Propanol + Water at 303.15 K	https://www.doi.org/10.1021/je7007582
Determination and correlation of solubility of 11alpha,17alpha-dihydroxy-4-pregnene-3,20-dione in (formamide + ethanol, or 1-propanol, or 1,2-ethanediol, or 1,2-propanediol) mixtures at temperatures between 293.15 K and 313.15 K	https://www.doi.org/10.1016/j.jct.2016.10.046
Volumetric and transport properties of ternary mixtures containing 1-propanol, n-heptane and 2-methylbutan-2-ol or 3-methylbutan-1-ol at 298.15 K	https://www.doi.org/10.1016/j.jct.2006.07.021
Liquid–Liquid Equilibrium Behavior of 1-Methyl-2-(1-methylethyl)benzene + 1-Propanol + Water at 303.15 K	https://www.doi.org/10.1016/j.tca.2008.09.006
Solid–Liquid Phase Equilibrium of Binary Systems of Isomeric Dihydroxybenzaldehyde and Their Thermodynamic Modeling	https://www.doi.org/10.1021/je400622g
Measurement and Correlation of Solubility of 2-Amino-5-chloro-3-nitropyridine in Different Pure Solvents at Several Temperatures	https://www.doi.org/10.1021/acs.jced.7b00206
Thermodynamic Properties, and Molecular Simulation of Solubility of Paracetamol in Pure Solvents	https://www.doi.org/10.1016/j.fluid.2017.09.007
Effect of the Alkyl Chain on the Physical Properties of Acid in 15 Pure Solvents at Different Temperatures	https://www.doi.org/10.1021/acs.jced.7b00118
Thermodynamic modelling for solubility of 4-nitrobenzaldehyde in 1-Propanol, 1-Butanol, and 1-Pentanol	https://www.doi.org/10.1021/je0341413
Solubility of 4-Nitrobenzaldehyde in Different Pure Solvents at Temperature Range from (278.15 to 318.15) K and Thermodynamic Modeling	https://www.doi.org/10.1021/acs.jced.9b00362
Measurement and Correlation of Solubility of 2-Amino-5-chloro-3-nitropyridine in 15 Pure Solvents at Several Temperatures	https://www.doi.org/10.1021/acs.jced.8b00948
Thermodynamic and spectroscopic study of binary mixtures containing dimethyl carbonate (DMC) + alcohols at T = (288.15 to 308.15) K and P = 0.1 MPa: Experimental study and ERAS-model calculation	https://www.doi.org/10.1016/j.jct.2016.09.012
Solubility of Organic Systems Containing 1,3-Dioxan-2-one: Measurement and Correlation of Activity, Density, and Speed of Sound for Binary Mixtures of 1-Propanol + Poly(propylene Glycol) 400, 1-Butyl-3-methylimidazolium Tetrafluoroborate, and Diethyl Carbonate with Application to Solubility and Preferential Solvation of Paracetamol in Pure Solvents: Predictive Results by UNIFAC-VISCO Group Contribution Method	https://www.doi.org/10.1021/je401044h
Solubility Measurement and Phase Equilibrium Modeling of 2-Aminobenzamide in 15 Pure Solvents	https://www.doi.org/10.1021/acs.jced.9b00350
Measurement and Correlation of Solubility of 2-Chloro-4-ethylamino-6-isopropylamino-1,3,5-triazine in Different Organic Solvents: Dimethyl Carbonate (DMC) + alcohols} at T = (288.15 to 328.15) K and P = 0.1 MPa: Experimental study and PC-SAFT Model Calculation	https://www.doi.org/10.1021/je400639m
Solubility of 1-Butyl-3-methylimidazolium Tetrafluoroborate in 1-Propanol at 323.15 K and Atmospheric Pressure	https://www.doi.org/10.1016/j.jct.2019.02.011
Measurement and Correlation of Solubility of Organic Systems Containing 1,3-Dioxan-2-one: Activity, Density, and Speed of Sound for Binary Mixtures of 1-Propanol + Poly(propylene Glycol) 400, 1-Butyl-3-methylimidazolium Tetrafluoroborate, and Diethyl Carbonate with Application to Solubility and Preferential Solvation of Paracetamol in Pure Solvents: Predictive Results by UNIFAC-VISCO Group Contribution Method	https://www.doi.org/10.1021/je900337a
Solubility Measurement and Phase Equilibrium Modeling of 4-Nitrobenzaldehyde in 15 Pure Solvents	https://www.doi.org/10.1016/j.jct.2013.02.006
Measurement and Correlation of Solubility of Organic Systems Containing 1,3-Dioxan-2-one: Activity, Density, and Speed of Sound for Binary Mixtures of 1-Propanol + Poly(propylene Glycol) 400, 1-Butyl-3-methylimidazolium Tetrafluoroborate, and Diethyl Carbonate with Application to Solubility and Preferential Solvation of Paracetamol in Pure Solvents: Predictive Results by UNIFAC-VISCO Group Contribution Method	https://www.doi.org/10.1021/je050406x
Measurement and Correlation of Activity, Density, and Speed of Sound for Binary Mixtures of 1-Propanol + Poly(propylene Glycol) 400, 1-Butyl-3-methylimidazolium Tetrafluoroborate, and Diethyl Carbonate with Application to Solubility and Preferential Solvation of Paracetamol in Pure Solvents: Predictive Results by UNIFAC-VISCO Group Contribution Method	https://www.doi.org/10.1021/acs.jced.7b00581
Density and Viscosity of Binary Mixtures of Diethyl Carbonate with 1-Propanol and 2-Propanol at (293.15 to 313.15) K	https://www.doi.org/10.1021/je0600808
Solubility and preferential solvation of paracetamol in (1-propanol + water) at 298.15 K	https://www.doi.org/10.1016/j.jct.2018.03.010
Density and Viscosity of Binary Mixtures of Ethyl 2-Methylbutyrate and 1-Propanol or 2-Propanol at (293.15, 298.15, 313.15, 328.15, and 343.15) K	https://www.doi.org/10.1021/je0500633
Experimental Measurement and Modeling of Vapor–Liquid Equilibrium for the Ternary System Water + 1-Propanol + Glycerol	https://www.doi.org/10.1016/j.fluid.2007.07.030
Solid–Liquid Phase Equilibrium of trans-Cinnamic Acid in Several Alcohols: Measurements and Thermodynamic Modeling	https://www.doi.org/10.1021/acs.jced.5b01015
Densities and Excess Molar Volumes of Cycloalkanone (1) + 1-Alkanol (2) Systems at (298.15 and 308.15) K:	https://www.doi.org/10.1021/acs.jced.5b00814
	https://www.doi.org/10.1021/je800831d

Experimental measurements and predictions of density, viscosity, and carbon dioxide solubility in methanol, ethanol, and 1-propanol:	https://www.doi.org/10.1016/j.jct.2012.10.002
Solubility of Phenanthrene in Binary Mixtures of C1-C4 Alcohols + 2-Propanol and Ethanol + Methanol at 298.15 K	https://www.doi.org/10.1021/je900027h
Isobaric Vapor Liquid Equilibrium Data for Binary Mixtures of 1-Phenylethanone with a Few Alcohols at 95.2 kPa	https://www.doi.org/10.1021/je300733a
Thermodynamics and activity coefficients at infinite dilution for organic solutes in the ionic liquid 1-butyl-1-methylpyrrolidinium tricyanomethanide:	https://www.doi.org/10.1016/j.fluid.2018.06.013
Densities and Viscosities of Diaminotoluene with Water, Ethanol, Propan-1-ol, and Butan-1-ol from (293.15 to 333.15) K	https://www.doi.org/10.1021/je400968v
Determination and correlation of solubility of N-methyl-3,4,5-trinitropyrazole (MTNP) in ten pure solvents from 283.15 K to 323.15 K:	https://www.doi.org/10.1016/j.fluid.2017.04.008
Thermodynamics and selectivity of separation based on activity coefficients at infinite dilution of various solutes in ionic liquid [HMIM][BF4]:	https://www.doi.org/10.1016/j.jct.2018.05.017
Solubility of Thiotriazinone in Ethanol, 1-Propanol, 2-Propanol, 1-Butanol, Acetone, Ethyl Acetate, and Water from 293 K to 343 K	https://www.doi.org/10.1021/je900177h
Activity coefficients at infinite dilution, physicochemical and thermodynamic properties for organic solutes and water in the ionic liquid ethyl-dimethyl-(2-methoxyethyl)ammonium tris(pentafluoroethyl)trifluorophosphate:	https://www.doi.org/10.1016/j.jct.2015.05.022
Solubilities of some organic solutes in 1-ethyl-3-methylimidazolium acetate. Chromatographic measurements and predictions from COSMO-RS:	https://www.doi.org/10.1016/j.jct.2011.11.005
Measurement and Correlation of Solubility of Theobromine, Theophylline, and Caffeine in Water and Organic Solvents at Various Temperatures	https://www.doi.org/10.1021/acs.jced.7b00065
Solubility of 2-Methyl-1,4-naphthoquinone in Water + (Methanol, Ethanol, 1-Propanol, 2-Propanol, 1,2-Propanediol, and Glycerin) Respectively from (293.15 to 337.92) K	https://www.doi.org/10.1021/je7002989
Solubility Measurement and Modeling of 3-Hydroxy-2-nitropyridine in Ten Pure Solvents and Two Binary Mixed Solvents for T = (278.15–318.15) K:	https://www.doi.org/10.1021/acs.jced.9b00564
Vapor-liquid equilibrium, densities, and interfacial tensions for the system ethyl 1,1-dimethylethyl ether (ETBE) + propan-1-ol:	https://www.doi.org/10.1016/j.fluid.2007.04.009
Physico-chemical and excess properties of the binary mixtures of methylcyclohexane + ethanol, + propan-1-ol, + propan-2-ol, + butan-1-ol, + 2-methyl-1-propanol, or 3-methyl-1-butanol at T = (298.15, 303.15, and 308.15) K:	https://www.doi.org/10.1016/j.jct.2005.06.011
A green process for recovery of 1-propanol/2-propanol from their aqueous solutions: Experimental and simulation studies	https://www.doi.org/10.1016/j.jct.2016.10.001
Isobaric vapour liquid equilibria for binary systems of 2-butanone with ethanol, 1-propanol, and 2-propanol at 20 and 101.3 kPa:	https://www.doi.org/10.1016/j.fluid.2008.06.004
Densities and Kinematic Viscosities of Ten Binary 1-Alkanol Liquid Systems at Temperatures of (293.15 and 298.15) K:	https://www.doi.org/10.1021/je800497u
Solubility of 1-methyl-4-nitropyrazole in seventeen pure solvents at temperatures from 283.15 K to 323.15 K:	https://www.doi.org/10.1016/j.fluid.2018.05.023
Experimental Determination, Correlation, and Prediction of Physical Properties of the Ternary Mixtures Ethanol and 1-Propanol + Water + 1-Ethyl-3-methylpyridinium Ethylsulfate at 298.15 K	https://www.doi.org/10.1021/je8009732
Solid-liquid equilibria for selected binary systems containing diphenyl carbonate:	https://www.doi.org/10.1016/j.fluid.2018.09.023
Liquid-liquid equilibrium of (water + 1-propanol + 1-pentanol) system at 298.15 and 323.15 K:	https://www.doi.org/10.1016/j.fluid.2010.02.026
Solubility of Trimethoprim (TMP) in Different Organic Solvents from (278 to 333) K	https://www.doi.org/10.1021/je700497h
Isothermal Vapor-Liquid Equilibrium Measurements for Alcohol + Water/n-Hexane Azeotropic Systems Using Both Dynamic and Automated Static-Synthetic Methods	https://www.doi.org/10.1021/acs.jced.9b00102
Excess molar volumes of binary mixtures of 1,3-dimethylimidazolidin-2-one with an alkan-1-ol at the temperatures 288.15 K, 298.15 K, and 313.15 K:	https://www.doi.org/10.1016/j.jct.2006.09.017
Activity coefficients at infinite dilution of organic solutes in diethylene glycol from gas-liquid chromatography:	https://www.doi.org/10.1016/j.jct.2013.05.011
Solubility of N-Chloro-4-toluenesulfonamide Sodium and N-Chlorobenzenesulfonamide Sodium in Ethanol, Water, 1-Propanol, and Butan-2-ol	https://www.doi.org/10.1021/je300721p
Application of the Extended Langmuir Model for the Determination of Activity of 1-Propanol in Acetonitrile	https://www.doi.org/10.1007/s10765-009-0581-7
Determination of Activity Coefficients at Infinite Dilution of Organic Solutes in the Ionic Liquid 1-(2-Hydroxyethyl)-3-methylimidazolium Nonafluoro-1-butanesulfonate Using Gas-Liquid Chromatography	https://www.doi.org/10.1021/acs.jced.8b00080
The buffering-out effect and phase separation in aqueous solutions of EPPS buffer with 1-propanol, 2-propanol, and 2-methyl-2-propanol at T = 298.15 K:	https://www.doi.org/10.1016/j.jct.2011.10.024
Solubility of dimethyl fumarate in water + (methanol, ethanol, 1-propanol) from (278.15 to 333.15) K:	https://www.doi.org/10.1016/j.fluid.2012.12.031
Densities and Viscosities of Binary Mixtures Containing 1,3-Dimethylimidazolium Dimethylphosphate and Alcohols:	https://www.doi.org/10.1021/je401124t
(Vapor + liquid) equilibria of alcohol + 1-methyl-1-propylpiperidinium triflate ionic liquid: VPO measurements and modeling:	https://www.doi.org/10.1016/j.jct.2016.01.026
Solubilities of (?Z)-2-Amino-?-(methoxyimino)-4-thiazole-ethanethioic Acid S-2-Benzothiazolyl Ester in Pure Solvents and a Mixture of Acetonitrile and Dichloromethane:	https://www.doi.org/10.1021/je900784v
Measurement of activity coefficients at infinite dilution of organic solutes in the ionic liquid 1-hexyl-1,4-diaza[2.2.2]bicyclooctanium bis(trifluoromethylsulfonyl)imide using gas-liquid chromatography:	https://www.doi.org/10.1016/j.jct.2013.10.026
Vapor-liquid equilibrium and excess properties of the binary mixtures formed by ethyl isobutyrate and n-alkanols:	https://www.doi.org/10.1016/j.jct.2019.105884
Experimental Study of Thermodynamic Properties of Mixtures Containing Ionic Liquid 1-Ethyl-3-methylimidazolium Ethyl Sulfate Using Gas-Liquid Chromatography and Transpiration Method	https://www.doi.org/10.1021/je0602723
Liquid-liquid equilibria for the ternary systems water + 1-propanol + methyl methacrylate, + butyl methacrylate and + isobutyl methacrylate:	https://www.doi.org/10.1016/j.fluid.2007.05.018
Vapor-Liquid Equilibrium for 1-Butene + Methanol, + 1-Propanol, + 2-Propanol, + 2-Butanol, and 2-Methyl-2-propanol (TBA) at 364.5 K:	https://www.doi.org/10.1021/je800160q
Thermodynamic study of aqueous two phase systems for some aliphatic alcohols + sodium thiosulfate + water:	https://www.doi.org/10.1016/j.fluid.2012.02.015
High pressure vapor liquid equilibria for CO2 + alkanol systems and densities of n-dodecane and n-tridecane:	https://www.doi.org/10.1016/j.fluid.2007.04.023
Densities, excess molar volumes, speeds of sound and isothermal compressibilities for {2-(2-hexyloxyethoxy)ethanol + n-alkanol} systems at temperatures between (288.15 and 308.15) K:	https://www.doi.org/10.1016/j.jct.2008.01.015
Excess Molar Enthalpies of N,N-Dimethylethanolamine with (Methanol, Ethanol, 1-Propanol, and 2-Propanol) at T) (298.2, 313.2, and 328.2) K and p) (0.1 and 10.0) MPa:	https://www.doi.org/10.1021/je800334m

Partition Coefficients of Organic Compounds in New Imidazolium and Tetraalkylammonium Based Ionic Liquids Using 1-Octanol/Water Partition Coefficients and Hydrogen-Bonding Descriptors.	https://www.doi.org/10.1021/je9003178
Influence of chain length and degree of branching of alcohols + chlorobenzene mixtures on excess enthalpies.	https://www.doi.org/10.1016/j.fluid.2007.09.028
Oxygenated Compounds Chromatographic Determination and Modeling of VLE of CES and Hydrocarbon Mixtures of Fuels and Determination and Correlation of Activity Coefficients at Infinite Dilution of Organic Solutes in Ionic Liquid 1-Hexyl-3-methylimidazolium Bis(trifluoromethylsulfonyl)imide Using Gas-Liquid Chromatography.	https://www.doi.org/10.1021/je5004912
Densities and Excess Enthalpies of Ternary Mixtures Containing 1-Butoxybutane + Propan-1-ol + 1-Hexene, or Heptane, or Methylcyclohexane at 298.15 and 313.15 K.	https://www.doi.org/10.1021/acs.jced.7b00825
Isothermal Vapor–Liquid Equilibrium Measurements for Binary Systems Containing 1-Butoxybutane + Propan-1-ol + 1-Hexene, or Heptane, or Methylcyclohexane at 313.15 K.	https://www.doi.org/10.1021/acs.jced.7b00740
Isothermal vapor–liquid equilibrium and excess properties of the ternary mixture methanol + 1,2-dimethoxyethane + 1-butanol at 313.15 K.	https://www.doi.org/10.1016/j.fluid.2005.11.014
Activity Coefficients at Infinite Dilution of Organic Solutes in Ionic Liquid 1-Hexyl-3-methylimidazolium Bis(trifluoromethylsulfonyl)imide Using Gas-Liquid Chromatography.	https://www.doi.org/10.1021/je9008443
Isothermal Data, Correlation, and Prediction for Binary Systems Ethanol + 1-Ethyl-3-methylimidazolium Ethyl Sulfate and 1-Propanol + 1-Ethyl-3-methylimidazolium Ethyl Sulfate.	https://www.doi.org/10.1021/je500191j
Isothermal vapor-liquid equilibrium measurements for binary systems of 1-propanol + 2-butanol + 2-methyl-2-propanol at 101.33 kPa.	https://www.doi.org/10.1016/j.jct.2016.03.011
Activity coefficients at infinite dilution for organic solutes in 1-hexadecyl-3-methylimidazolium tetrafluoroborate ionic liquid by inverse gas chromatography.	https://www.doi.org/10.1016/j.fluid.2012.10.019
Activity coefficients at infinite dilution measurements for organic solutes in 1-(2-methoxyethyl)-1-methylpiperidinium bis(trifluoromethylsulfonyl)imide ionic liquid by GLC.	https://www.doi.org/10.1016/j.jct.2012.01.019
Activity coefficients at infinite dilution for organic compounds dissolved in 1-(2-methoxyethyl)-1-methylpiperidinium bis(trifluoromethylsulfonyl)imide ionic liquid by gas-liquid chromatography.	https://www.doi.org/10.1021/je300692s
Isothermal vapor-liquid equilibrium measurements for binary systems of 1-propanol + 1-butanol + 2-butanol + 2-methyl-2-propanol at 101.33 kPa.	https://www.doi.org/10.1016/j.jct.2016.07.013
Isobaric vapor-liquid equilibrium for binary systems of 1-propanol + 2-butanol + 2-methyl-2-propanol at 101.33 kPa.	https://www.doi.org/10.1021/je049852v
Vapor-liquid equilibrium for binary mixtures of 1-propanol + 2-propanol + 2-methyl-2-propanol at 101.33 kPa.	https://www.doi.org/10.1016/j.jct.2003.11.003
Measurement and Correlation for Solubility of 21-acetoxy-16alpha,17alpha-epoxyprogesterone and 16alpha,17alpha-Epoxyprogesterone in Several Solvent Systems.	https://www.doi.org/10.1021/je101072d
Vapour-liquid and chemical equilibria in the ethyl ethanoate + ethanol + propyl ethanoate + propanol system accompanied with transesterification reaction.	https://www.doi.org/10.1016/j.jct.2015.10.006
Measurements and correlations of density and viscosity for short chain 1-alkanol + water mixtures in the temperature range from 300.15 K to 340.15 K at pressures up to 40 MPa.	https://www.doi.org/10.1016/j.jct.2016.07.049
Solubility of methylphosphonic acid in selected organic solvents:	https://www.doi.org/10.1016/j.fluid.2012.05.016
Excess Molar Enthalpies of Methyl Acetate and (1-Propanol, 2-Propanol, 1-Butanol, 2-Butanol, and 2-Methyl-2-propanol) at T = (298.15 and 308.15) K and P = (5.0 and 10.0) MPa.	https://www.doi.org/10.1016/j.fluid.2015.07.012
Vapour-liquid equilibria, azeotropic data, excess enthalpies, activity coefficients at infinite dilution and solid liquid equilibria for binary alcohol + ketone systems.	https://www.doi.org/10.1016/j.jct.2013.12.029
Thermophysical Properties of the Pure Ionic Liquid 1-Butyl-3-methylpyrrolidinium Dicyanamide and Its Mixtures with Alcohols.	https://www.doi.org/10.1016/j.fluid.2015.01.009
Solubility of d-Aspartic Acid in Several Neat Solvents: Determination, Correlation, and Solvent Effect Analysis: Nonaqueous Solvent Mixtures at 298.2 K.	https://www.doi.org/10.1021/je050175u
Measurement and Correlation of Solubility of Marbofloxacin in 12 Pure Solvents from 288.15 to 328.15 K.	https://www.doi.org/10.1016/j.fluid.2008.02.021
Isobaric vapor-liquid equilibria for 1-propanol +water + copper(II) chloride at 100 kPa.	https://en.wikipedia.org/wiki/Joback_method
Thermodynamic evidence for nano-heterogeneity in solutions of the macrocycle 0-butyresorc[4]arene in non-aqueous solvents.	https://www.doi.org/10.1021/je300384g
Solubility of Topiramic Sulfamate in Different Solvents from (278.35 to 313.15) K.	https://www.doi.org/10.1021/acs.jced.9b00320
Measurement and Correlation of Excess Molar Enthalpies for Ethylene Glycol + Water Systems at Temperatures (298.15, 303.15, and 323.15) K.	https://www.doi.org/10.1021/acs.jced.5b00355
Thermodynamic properties of quantitropic polymorphs of tolcolide:	https://www.doi.org/10.1021/acs.jced.9b00490
Mutual diffusion behavior of short chain alcohols + n-octane mixtures:	https://www.doi.org/10.1016/j.fluid.2004.11.014
Isobaric Vapor-Liquid Equilibrium of the Acetonitrile + 1-Propanol + Ionic Liquids at an Atmospheric Pressure:	https://www.doi.org/10.1016/j.jct.2018.06.006
Application of the ERAS model to volumetric properties of binary mixtures of benzyl oil with primary and secondary alcohols C1 - C4 at different temperatures.	https://www.doi.org/10.1021/je800689a
Measurement of critical temperatures and critical pressures for binary systems containing 1-propanol and 1-butyl ether	https://www.doi.org/10.1021/je030207i
Solubility modeling and thermodynamic dissolution functions of phthalimide in ten organic solvents:	https://www.doi.org/10.1016/j.jct.2017.03.011
Chemical calibration of isothermal Titration Calorimeters: An evaluation of the dilution of propan-1-ol into water as a test reaction using different calorimeters, concentrations, and	https://www.doi.org/10.1016/j.tca.2015.11.024
Determination and Correlation of Solubility of Pharmaceutical Compounds in Propan-1-ol + Water Cosolvent Mixtures.	https://www.doi.org/10.1021/acs.jced.9b00015
Measurement of activity coefficients at infinite dilution of organic solutes in the ionic liquid 1-methyl-2-pentanone + 1-propanol or 2-propanol + water ternary systems at different temperatures.	https://www.doi.org/10.1016/j.jct.2010.12.009
Measurements and correlation of (vapor + liquid) equilibria of 1-methyl-3-pentylimidazolium 2-(2-methoxyethoxy) ethylsulfate at T = (308.15, 313.15, 323.15 and 333.15) K using gas + liquid chromatography:	https://www.doi.org/10.1016/j.jct.2013.03.002
	https://www.doi.org/10.1016/j.jct.2015.10.024
	https://www.doi.org/10.1016/j.jct.2011.12.018
	https://www.doi.org/10.1021/je301029j
	https://www.doi.org/10.1016/j.jct.2013.10.017
	https://www.doi.org/10.1016/j.fluid.2013.10.034

Isobaric Vapor-Liquid Equilibria for Binary and Ternary Mixtures of	https://www.doi.org/10.1021/je100912t
Solubility of	https://www.doi.org/10.1021/acs.jced.6b00617
Propanol, Propanol, and Propanoic	
4-Methylsulfonylacetophenone in Nine	https://www.doi.org/10.1021/je800246v
Measurement of Infinite Dilution	
Pure Solvents and (Ethyl Acetate +	https://www.doi.org/10.1016/j.fluid.2008.04.010
Activity Coefficients of Alcohols	
(1-Hexene) Binary Solvent Mixtures	https://www.doi.org/10.1016/j.tca.2005.06.011
Liquid-liquid equilibria for binary	
systems of isopropyl ether (DIPE)	https://www.doi.org/10.1021/je0201931
+ ternary excess models and	
1-Ethyl-3-methylimidazolium	https://www.doi.org/10.1016/j.fluid.2014.09.008
Tetrafluoroborate and	
1-Butyl-3-methylimidazolium	https://www.doi.org/10.1021/acs.jced.8b00763
Thermodynamic Properties of Mixtures	
Containing Ionic Liquids. 3.	https://www.doi.org/10.1016/j.jct.2006.03.010
Liquid-Liquid Equilibria of Binary	
mixture solvents and binary solvent	https://www.doi.org/10.1021/je100581m
mixtures	
1-Ethyl-3-methylimidazolium	https://www.doi.org/10.1021/acs.jced.9b00445
Bis(trifluoromethylsulfonyl)imide with	
Binary Vapor-Liquid Equilibrium	https://www.doi.org/10.1016/j.tca.2005.03.009
Data for 1-Butanol + 1-Butyl	
Viscosity studies of binary mixtures of	https://www.doi.org/10.1016/j.fluid.2010.03.004
1-butanol with trichloromethane and	
Measurements of the (Ethanol +	https://www.doi.org/10.1016/j.fluid.2007.06.007
Linalool) + (Propan-1-ol and Linalool)	
Solubility Measurement and	https://www.doi.org/10.1016/j.fluid.2009.01.011
Thermodynamic Modeling for	
Molar excess enthalpies of ethyl	https://www.doi.org/10.1021/acs.jced.5b00903
acetate + ethanol at 313.15K,	
Analysis of 1-Butyl-1-methylpyrrolidinium	https://www.doi.org/10.1016/j.jct.2019.01.015
trifluoromethanesulfonate using GLC:	
Ebulliometric determination and	https://www.doi.org/10.1016/j.jct.2011.08.006
prediction of (vapour + liquid) equilibria	
Solubility and solution	https://www.doi.org/10.1016/j.jct.2016.05.027
thermodynamics of (C1 C4) and	
Aqueous Solubility Prediction Method:	http://onschallenge.wikispaces.com/file/view/AqueousDataset002.xlsx/351826032/AqueousDa
Solubility Determination and Modeling	https://www.doi.org/10.1021/acs.jced.9b00065
of p-Nitrobenzamide Dissolved in	
Solubility measurement and	https://www.doi.org/10.1016/j.jct.2016.10.040
thermodynamic functions of	
Solubility and Density of	https://www.doi.org/10.1021/je049967z
2,6-Dimethylnaphthalene in C1 C7	
Excess molar volumes and partial	https://www.doi.org/10.1016/j.jct.2005.05.007
molar volumes for (propionitrile + an	
Isobaric vapor-liquid equilibria for	https://www.doi.org/10.1016/j.fluid.2010.06.004
water + n-propanol + n-butanol ternary	
Solubility of Phenacetin in	https://www.doi.org/10.1021/je700209v
Methanol, Ethanol, 1-Propanol,	
Activity coefficient at infinite dilution	https://www.doi.org/10.1016/j.jct.2012.12.009
measurements for organic solutes	
Molecular interactions in binary	https://www.doi.org/10.1016/j.jct.2014.07.019
mixtures of 1-butoxy-2-propanol with	
Vapor-liquid equilibria for alcohol +	https://www.doi.org/10.1016/j.jct.2009.06.024
Measurement of vapour-liquid equilibria	https://www.doi.org/10.1016/j.fluid.2007.06.001
Solubility of Cesium Nitrate in Alcohol	https://www.doi.org/10.1021/je050072b
Aqueous Solutions at Between	
Excess properties of binary mixtures of	https://www.doi.org/10.1016/j.tca.2006.08.006
Solubility of alpha- form and beta-form	https://www.doi.org/10.1016/j.fluid.2010.10.020
Measurement and Correlation of	https://www.doi.org/10.1021/je101208p
Solubilities of Asiaticoside in Water,	
Solubility of N-Ethylcarbazole in	https://www.doi.org/10.1016/j.fluid.2014.06.030
different organic solvent at 279.15-	
Solubility Correlation and	https://www.doi.org/10.1021/acs.jced.6b00630
Thermodynamic Analysis of Sorafenib	https://www.doi.org/10.1021/acs.jced.8b00560
Solubility Determination of	
Nicotinamide and Its Application for the	https://www.doi.org/10.1016/j.jct.2018.12.038
Solubility Modeling and	
Measured	https://www.doi.org/10.1016/j.jct.2016.11.019
thermodynamic aspect of	
Solubility determination and	https://www.doi.org/10.1021/je020149l
thermodynamic modelling of several	
Excess Molar Enthalpies for Binary	https://www.doi.org/10.1016/j.tca.2006.02.017
Mixtures of Ethanol + Acetone +	
Solubility of 4-chlorobenzoic acid in eleven	https://www.doi.org/10.1021/acs.jced.9b00243
Solubility of Solid-Liquid Phase	
Equilibrium and Thermodynamic	https://www.doi.org/10.1021/je700200b
Properties of 1-Octanol + Heptane at	
Solubility of Benzocaine in Several Pure Solvents: Measurement,	
Correlation, and Solvent Effect	
Solubility of 1 H-Tetrazole-4-acetic Acid in Different Solvents between 283 K and 323 K:	
Analysis:	

Liquid-Liquid Phase Equilibria of (1-Propanol or 2-Propanol + Water) Containing Dipotassium Hydrogen Phosphate	https://www.doi.org/10.1021/je060232z
Measurement and correlation of liquid-liquid equilibria at different temperatures of water + 1-propanol + 1-heptanol	https://www.doi.org/10.1016/j.fluid.2016.08.038
Physical Properties of Binary Mixtures of the Ionic Liquid 1-Ethyl-3-methylimidazolium Ethyl Sulfate with Water or Ethanol at Several Temperatures	https://www.doi.org/10.1021/je060123k
Excess Molar Volumes, Viscosities, and Refractive Indexes for Binary Mixtures of 1-Ethyl-3-methylimidazolium Ethyl Sulfate + Water + 1-Propanol at (298.15, 308.15, and 318.15) K	https://www.doi.org/10.1021/je030226s
Activity coefficients at infinite dilution measurements for organic solutes in the ionic liquid 1-(2-Hydroxyethyl)-3-methylimidazolium Ethyl Sulfate using GLC	https://www.doi.org/10.1016/j.jct.2012.08.016
Activity coefficients at infinite dilution for solutes in the ionic liquid 1-(2-hydroxyethyl)-3-methylimidazolium ethyl sulfate using GLC	https://www.doi.org/10.1016/j.jct.2009.08.012
Research on the application of 7-chloro-quinaldine adducts in ionic liquid extraction separation of Vanadium(V)	https://www.doi.org/10.1016/j.jct.2019.02.002
Volumetric Properties of 1-Butyl-4-methylpyridinium tetrafluoroborate [bmpy][BF4] with Water and with an Alkan-1-ol (Methanol to 1-Pentanol) at Several Temperatures	https://www.doi.org/10.1016/j.jct.2008.02.019
Chromatography: Validity Model Correlation, and Solvent Effect	https://www.doi.org/10.1021/acs.jced.8b01265
Liquid-liquid equilibria for water + 1-propanol or 1-butanol + potassium chloride + ammonium chloride at 298.15 K	https://www.doi.org/10.1016/j.fluid.2015.03.050
Density and Surface Tension of Binary Mixtures of Acetonitrile and Alkanol at 298.15 K:	https://www.doi.org/10.1021/je050519g
Quaternary systems at 298.15 K: Excess enthalpies and molar excess volumes of formamide + 1-propanol or 2-propanol and	https://www.doi.org/10.1016/j.tca.2013.02.010
Solubility of L-proline in aqueous solutions of methanol, ethanol, 1-propanol or 2-propanol at	https://www.doi.org/10.1016/j.jct.2018.05.011
Thermodynamic modeling of Pregnane Poly 4-alkenyl and Progesterone Derivatives in Alkanol	https://www.doi.org/10.1016/j.jct.2017.04.019
Solubility modeling and dissolution thermodynamic analysis of Aclacinomycin A in different pure solvents	https://www.doi.org/10.1016/j.jct.2017.07.012
Infinite dilution activity coefficients of alkanols + ethyl acetate)	https://www.doi.org/10.1021/je0201727
Solubility of Acetazolamide in Binary and Ternary Mixtures of Ethanol + Water + Propylene Glycol at Different Temperatures	https://www.doi.org/10.1016/j.fluid.2019.112263
Alcohol + Acetonitrile Solvent Mixtures	https://www.doi.org/10.1021/acs.jced.8b01226
Measurement of the thermal conductivity of biofuel mixtures: Methyl 1-butyl-3-methylimidazolium Saccharinate	https://www.doi.org/10.1021/je0301904
Solubility Model, Correlation, and Solvent Effect	https://www.doi.org/10.1007/s10765-014-1740-z
Separate components of biodiesel and alcohols:	https://www.doi.org/10.1021/je060248p
Solubility in Binary Solvent Mixtures: 2-Amino-3-methylbenzoic Acid in 12 Pure Solvents	https://www.doi.org/10.1021/acs.jced.8b00124
Solubility of Baicalein in Different Solvents from (287 to 323) K:	https://www.doi.org/10.1021/acs.jced.9b00659
Effect of Pressure on the Static Relative Permittivities of Alkan-1-ols at 298.15 K	https://www.doi.org/10.1016/j.jct.2005.08.004
Measurement and Correlation of the Solubility of Pyrimethanil in Seven Monosolvents and Two Different Binary Mixed Solvents	https://www.doi.org/10.1021/acs.jced.9b00308
Determination and Modeling of Histidine Solubility in Several Pure Solvents from 293.15 to 333.15 K:	https://www.doi.org/10.1016/j.jct.2010.05.017
Henry's law constants and infinite dilution activity coefficients of cis-2-butene, dimethylether, chloroethane, and 1,1-difluoroethane in	http://pubs.acs.org/doi/abs/10.1021/ci990307l
Measurements of activity coefficients at infinite dilution of organic solutes in 1-propanol, isobutanol, tert-butanol, 1-pentanol, 2-pentanol,	https://www.doi.org/10.1016/j.fluid.2006.05.032
1-propyl-1-methylpiperidinium bis(trifluoromethylsulfonyl)imide + 1-propanol + toluene: A discussion of	https://www.doi.org/10.1016/j.jct.2013.02.004
Solubility and SAFT-VR Thermodynamic Models for Solid−Liquid Equilibrium of 1,2-Propanediol + Alkan-1-ols (C1−C6)	https://www.doi.org/10.1021/acs.jced.8b00165
Excess Molar Enthalpies of 1-Propanol + Dimethyl Carbonate + Alkan-1-ol Ternary Systems at 298.15 K	https://www.doi.org/10.1021/je060086x
Solubility and solution thermodynamics of Pemetrexed Disodium Heptahydrate in Different Pure Solvents:	https://www.doi.org/10.1016/j.fluid.2015.03.022
Modeling of Aniline Hydrochloride Solubility in Water + Methanol, Ethanol, 1-Propanol, and Their Mixed Systems:	https://www.doi.org/10.1021/je201047m
Surface Tension of Dilute Solutions of Linear Alcohols in Benzyl Alcohol	https://www.doi.org/10.1021/je0500431
Thermodynamics of fenofibrate and solubility in pure organic solvents:	https://www.doi.org/10.1016/j.fluid.2014.01.029
Solubility and Thermodynamic Modeling of Dimethyl Terephthalate in Pure Solvents and the Evaluation of the Mixing Properties of the Solutions	https://www.doi.org/10.1021/acs.jced.9b00658
VLE and VLLE Measurements of Dimethyl Ether Containing Systems. 2:	https://www.doi.org/10.1021/je025599s
Henry's constants and infinite dilution activity coefficients of propane, propene, butane, isobutene, 1-butene,	https://www.doi.org/10.1016/j.jct.2003.08.018
Influence of the temperature on the (liquid + liquid) phase equilibria of 3-(1-butyl-3-imidazolio)-1-propanesulfonate in	https://www.doi.org/10.1016/j.jct.2016.12.011
Capability study of 1-butyl-3-methylimidazolium bis(trifluoromethylsulfonyl)imide and	https://www.doi.org/10.1016/j.fluid.2018.04.011
Excess molar volumes of (C1−C5) Dialkylamine + (C1−C5) Alkan-1-ol Mixtures at 298.15 K	https://www.doi.org/10.1016/j.jct.2011.05.006
Thermodynamic and Spectroscopic Characterization of the Ionic Liquid Tributyl(tetradecyl)phosphonium 2,4,4-Trimethylpentylphosphinate	https://www.doi.org/10.1021/acs.jced.5b01023
Liquid-liquid equilibria of 1-propanol (or 2-propanol) + water + 1-butyl-3-methylimidazolium Acetate with 1-propanol or 2-propanol:	https://www.doi.org/10.1016/j.fluid.2013.12.022
Measurement and correlation of the solubility of 1-propanol in 1-butyl-3-methylimidazolium hexafluorophosphate with 1-propanol or tert-butanol in the critical region:	

Liquid Liquid Equilibrium for Ternary Systems Containing Water, Oleic Acid, and Alcohols at 313.15 K.	https://www.doi.org/10.1021/je501174y
Solubility Study of (2E)-1-(2-Pyridyl)-3-(dimethylamino)-2-propen-1-one in Different Pure Solvents and Binary Solvent Mixtures from 278.15 to 328.15 K: Correlation with Different Thermodynamic Models.	https://www.doi.org/10.1021/acs.jced.9b00047
Solubility of Gastrodin in Pure and Mixed Solvents at 273.15–313.15 K and Dissolution Properties of Binary Mixtures of Cyclopropanecarboxylic Acid + Methanol, Ethanol, 1-Propanol, and 2-Propanol at Different Temperatures.	https://www.doi.org/10.1021/acs.jced.9b00286
Phase Equilibrium Properties of Binary and Ternary Mixtures Containing Methyl Ethyl Ether, Ethanol, and Water at Different Temperatures.	https://www.doi.org/10.1021/je800930v
Measurement and Correlation of the Solubility of Diprogestin in Pure and Binary Solvent Mixtures of Ethanol + Water and 1-Propanol + Water at Temperatures from 278.15 to 318.15 K.	https://www.doi.org/10.1021/je060253t
The Solid–Liquid Equilibrium and Crystal Habit of Lamivudine–Fumaric Acid Cocrystal in Different Solvents.	https://www.doi.org/10.1021/acs.jced.9b00696
Measurement and correlation solubility and mixing properties of dimethyl succinate in pure and mixture organic solvents from (278.15 to 333.15) K.	https://www.doi.org/10.1021/acs.jced.7b00768
Measurement and correlation of solubility of thiourea in two solvent mixtures from (278.15 to 333.15) K.	https://www.doi.org/10.1016/j.fluid.2013.06.037
Solubility of Sodium Acetate in Binary Mixtures of Methanol, 1-Propanol, Acetonitrile, and Water at 298.2 K.	https://www.doi.org/10.1016/j.jct.2015.11.004
Phase equilibria involved in extractive distillation of dipropyl ether + 1-propyl alcohol using 2-ethoxyethanol as entrainer.	https://www.doi.org/10.1021/je400625f
Excess Enthalpy, Excess Volume, and Speed of Sound Deviation for the Mixtures of Pinene + Ethanol and 1-Propanol at (283.15, 298.15, and 313.15) K.	https://www.doi.org/10.1016/j.fluid.2007.03.030
Activity coefficients at Infinite Dilution of Organic Compounds in 1-Propanol at (283.15, 298.15, and 313.15) K.	https://www.doi.org/10.1021/je050006+
Excess molar enthalpies and heat capacities of dimethyl sulfoxide + chloroform or + 1,2-dichloroethane at 298.15 K and atmospheric pressure:	https://www.doi.org/10.1021/je800658v
Solubility of carvacrol in different organic solvents:	https://www.doi.org/10.1016/j.tca.2006.10.025
Excess molar enthalpies of {diethyl oxalate + (methanol, + ethanol, + 1-propanol, and + 2-propanol)} at T = (288.2, 298.2, 313.2, and 328.2) K and p = 101.5 kPa.	https://www.doi.org/10.1016/j.fluid.2011.12.004
Densities and Viscosities of Binary Liquid Mixtures of 2-Butoxyethanol with 1-Propanol, 2-Propanol, 2-Methyl-1-propanol, and 2-Methyl-2-propanol at 298.15 K.	https://www.doi.org/10.1016/j.jct.2013.05.024
The Yaws Handbook of Vapor Pressure:	https://www.doi.org/10.1016/j.jct.2016.11.029
Densities and Excess Molar Volumes of the Binary 1-Propanol + Chloroform and 1-Propanol + Benzene and Ternary 1-Propanol + Chloroform + Benzene Mixtures at 288.15, 293.15, 298.15, 303.15, and 308.15 K.	https://www.doi.org/10.1021/je034203p
Osmotic coefficients and apparent molar volumes of thiamine dihydrochloride in aqueous solutions at (298.15 to 318.15) K	https://www.sciencedirect.com/book/9780128029992/the-yaws-handbook-of-vapor-pressure
Thermodynamic functions for solubility and fusion of triclocarban in (1-propanol + water) cosolvent mixtures	https://www.doi.org/10.1021/je700254t
Measurement and correlation of vapor pressure of benzene and thiophene in the presence of [BMIM][BF4] ionic liquid.	https://www.doi.org/10.1016/j.jct.2013.09.041
Activity coefficients at infinite dilution measurements for various organic solutes in 1-propanol + benzene, + toluene, and + p-xylene using inverse gas chromatography.	https://www.doi.org/10.1016/j.jct.2017.02.017
Measurements and correlation of vapor pressure of 1-propanol + 1-butanol + 1-pentanol + 1-hexanol + 1-heptanol + 1-octanol at (293.15 to 318.15) K	https://www.doi.org/10.1016/j.fluid.2008.09.018
Activity coefficients at infinite dilution measurements for various organic solutes in 1-propanol + benzene, + toluene, and + p-xylene using inverse gas chromatography.	https://www.doi.org/10.1021/je200457q
Measurements of activity coefficients at infinite dilution of various organic solutes in 1-propanol + 1-hexanol + 1-octanol at 298.15 K.	https://www.doi.org/10.1016/j.jct.2010.01.004
Activity coefficients at infinite dilution and physicochemical properties for 1-butyl-1-methylpyrrolidinium tetracyanoborate ionic liquid with water and 1-propanol.	https://www.doi.org/10.1016/j.jct.2018.01.003
Measurements of activity coefficients at infinite dilution for organic solutes in 1-butyl-1-methylpyrrolidinium tetracyanoborate ionic liquid and correlation with various solvation models.	https://www.doi.org/10.1016/j.jct.2013.07.004
Measurement and correlation of viscosity and excess properties of binary mixture of 1-butyl-2-methylpyrrolidinium tetracyanoborate ionic liquid and 1-propanol at 298.15 K.	https://www.doi.org/10.1016/j.fluid.2005.07.013
Thermodynamic Properties of the Binary and Ternary Mixtures of 2-Propanol + Diisopropyl Ether (DIPE) + 1-Propanol or 2,2,4-Trimethylpentane (Isooctane) + 1-Propanol at T = (293.15 to 313.15) K.	https://www.doi.org/10.1021/je700460w
Measurement and Modelization of VLE for Butyl Acetate with Methanol, Ethanol, 1-Propanol, and 1-Butanol.	https://www.doi.org/10.1021/acs.jced.8b00182
Measurement and Correlation of the Solubility of 5-Fluorouracil in Pure and Binary Solvents.	https://www.doi.org/10.1021/acs.jced.8b00425
Measurements and equation-of-state modelling of thermodynamic properties of binary mixtures of 1-butyl-1-methylpyrrolidinium tetracyanoborate ionic liquid with 1-propanol and 2-propanol	https://www.doi.org/10.1016/j.jct.2015.06.004
Solubility, liquid–liquid equilibrium and critical states for the system acetic acid + water + 1-propanol (or 2-propanol)	https://www.doi.org/10.1016/j.fluid.2014.04.034
Measurement and Modeling of Solid–Liquid Equilibria of l-Glutamic Acid in Pure Solvents and Aqueous Binary Mixtures.	https://www.doi.org/10.1016/j.jct.2012.12.019
Equilibrium solubility of dinitolmide in several neat solvents and binary aqueous co-solvent mixtures.	https://www.doi.org/10.1021/acs.jced.8b01084
Modeling of Densities and Refractive Index Measurements under Pressure for the Binary Systems 1-Propanol + Diethyl Ether at T = (293.15 to 328.15) K.	https://www.doi.org/10.1016/j.jct.2019.01.013
Thermodynamic Properties of 1-Butyl-3-methylpyridinium Tetrafluoroborate and Its Mixtures with Water and Alkanols at Various Temperatures.	https://www.doi.org/10.1016/j.tca.2011.08.002
Measurement and prediction of tie-line data for mixtures of (water + 1-propanol + 4-isopropyl-1-methylcyclohexane) at T = (298.15 and 313.15) K.	https://www.doi.org/10.1021/je049685z
Activity coefficients at infinite dilution and physicochemical properties for organic solutes and water in the ionic liquid 4-(3-Hydroxypropyl)-4-methylmorpholinium bis(trifluoromethylsulfonyl)-amide:	https://www.doi.org/10.1021/je700294p
Solubility of Candesartan Cilexetil in Different Solvents at Various Temperatures.	https://www.doi.org/10.1021/je101020m
Water and Alkanols.	https://www.doi.org/10.1016/j.fluid.2008.11.014
Azeotropy in Ethyl Ether + 1-Propanol at Various Temperatures.	https://www.doi.org/10.1016/j.jct.2015.02.024
Isothermal Vapor–Liquid Equilibrium Data for the Propan-1-ol + Dodecane System at (323.0, 343.4, 353.2, 363.1, and 369.2) K.	https://www.doi.org/10.1021/je201032t

Experimental Determination of Densities and Isobaric Vapor Liquid Equilibria of Methyl Acetate and Ethyl Acetate with Alcohols (C1 and C2) at 0.3 MPa:	https://www.doi.org/10.1007/s10765-013-1483-2
Fluid Phase Topology of Benzene + Cyclohexane + 1-Propanol at 1013 kPa: Measurement and Correlation of	https://www.doi.org/10.1007/s10765-015-1922-3
Solubility of Calcium Formate (Form alpha) in Different Binary Solvent Mixtures at Temperatures from 283.15 to 323.15 K:	https://www.doi.org/10.1021/acs.jced.9b00009
Fully automatized apparatus for determining speed of sound for liquids at temperatures and pressures of the	https://www.doi.org/10.1016/j.jct.2016.09.022
Densities and Viscosities of 1-Octyl-3-methylimidazolium Nitrate and 1-Octylpyridinium Nitrate in Binary Mixtures with Alcohols at	https://www.doi.org/10.1021/je901058s
Interfacial thermodynamics and thermolar surface Gibbs energy and its application to the binary mixtures of	https://www.doi.org/10.1016/j.jct.2018.08.004
Surface tension and interfacial properties of binary glycerol/alcohol mixtures	https://www.doi.org/10.1016/j.jct.2019.03.014
N-Butyl-N-methylpyrrolidinium dicyanamide ([bmpyr][DCA]) with alcohols:	https://www.doi.org/10.1016/j.jct.2018.08.001
Solubility, modelling, solvent effect and preferential solvation of carbendazim in aqueous co-solvent mixtures of	https://www.doi.org/10.1021/acs.jced.8b01214
Solubility Determination of 1,5-Naphthalenediamine and 1,8-Naphthalenediamine in Different Solvents and Modified	https://www.doi.org/10.1021/je900838a
Study of Ether + Alcohol on Synthesized Tricationic Ionic Liquids	https://www.doi.org/10.1021/je2002539
Liquid−Liquid Phase Equilibria of Solutions Use Gas Chromatography:	https://www.doi.org/10.1021/acs.jced.8b00601
1-Propanol + Water + n-Alcohol Ternary Systems at 298.15 K and Atmospheric Pressure.	https://www.doi.org/10.1016/j.jct.2017.07.027
Thermodynamic Parameters of a New Synthesized Tricationic Ionic Liquid	https://www.doi.org/10.1016/j.jct.2016.03.007
Stationary Phase in Inverse Gas Chromatography.	https://www.doi.org/10.1021/je5009685
Thermodynamic modeling of solubility and preferential solvation for ribavirin in (1-propanol + water) cosolvent	https://www.doi.org/10.1021/je900038m
Determination and thermodynamic modeling of solid−liquid phase equilibrium for 3,5-dichloroaniline in pure solvents and ternary	https://www.doi.org/10.1016/j.jct.2015.09.010
Solubility of Tris(hydroxymethyl)aminomethane in Methanol, 1-Propanol, 2-Propanol Mixtures at	https://www.doi.org/10.1016/j.jct.2013.07.025
3,5-Dinitrobenzamide in Different Solvents from (283.15 to 323.15) K.	https://www.doi.org/10.1021/je060217u
Thermodynamics of thymol in six pure organic solvents.	https://www.doi.org/10.1016/j.jct.2013.05.035
Biocompatible choline based ionic salts: Solubility in short-chain alcohols.	https://www.doi.org/10.1021/acs.jced.9b00097
Measurement and Correlation of Vapor-Liquid Equilibria at T) 333.15 K and Excess Molar Volumes at T) 298.15	https://www.doi.org/10.1016/j.fluid.2014.06.009
Infinite dilution activity coefficients of volatile organic compounds in two ionic liquids composed of four	https://www.doi.org/10.1021/je800158z
K for Ethanol + Dimethyl Carbonate (DMC), Methanol + DMC, 1-Propanol + DMC, and Ethanol + DMC +	https://www.doi.org/10.1016/j.fluid.2012.01.019
Solubility Measurements and Solvent Effect Correlation of Trifluoroethoxy-	https://www.doi.org/10.1016/j.fluid.2009.03.006
Modification properties and related thermodynamic parameters ofaqueous	https://www.doi.org/10.1021/acs.jced.7b00523
Volumetric Properties of Binary and Ternary Liquid Mixtures of 1-Propanol	https://www.doi.org/10.1016/j.tca.2012.04.033
Determination of the solubility, dissolution enthalpy and entropy of	https://www.doi.org/10.1021/je050442w
Solubility and activity coefficient of simvastatin in different pure solvents.	https://www.doi.org/10.1016/j.fluid.2017.05.012
Ammonium Based Ionic Liquid as an Entrainer for the Separation of	https://www.doi.org/10.1016/j.jct.2005.04.010
Solubilities of 1,3-dimethyl-1,1,3,3-tetraphenyldisiloxane in different solvents from (288.15 to	https://www.doi.org/10.1021/je4004806
Vapor−Liquid Equilibria for the Ternary 1-Propanol + Water + Isopropyl Acetate System at	https://www.doi.org/10.1021/je7007236
n-Pentane + 1-Propanol and n-Pentane + Ethanol Systems near the Critical Region.	https://www.doi.org/10.1021/je700502x
Heterogeneous azeotropic distillation for the separation of n-propanol + water systems with four liquids	https://www.doi.org/10.1016/j.jct.2006.06.001
Measurement of VLE and c1 data and prediction of their thermodynamic consistency for binary mixtures of	https://www.doi.org/10.1021/je060133l
Solubility of Isatin Potassium in Different Pure Solvents from (278.15 to 323.15 K) and Thermodynamic	https://www.doi.org/10.1021/je1012839
Solubility of Anthracene in Binary Propyl Acetate + Alcohol Solvent Mixtures at 298.15 K.	https://www.doi.org/10.1021/acs.jced.9b00661
Densities and Derived thermodynamic properties of (2-methoxyethanol + 1-propanol, or 2-propanol, or	https://www.doi.org/10.1016/j.jct.2011.11.021
Solubilities of Ureylene in Organic Solvents from T = (283.19 to 313.19) K	https://www.doi.org/10.1021/acs.jced.7b00638
Solubility of Vinpocetine in Methanol, Ethanol, 1-Propanol, 2-Propanol, 1-Butanol, and Water from (278.15 to	https://www.doi.org/10.1021/je1006482
Solubility Determination of 2-Chloronicotinic Acid and Analysis of Solvent Effect	https://www.doi.org/10.1016/j.jct.2013.04.002
Solvent coefficients at infinite dilution and physicochemical properties for	https://www.doi.org/10.1021/je1000582
Solubility Measurement and Correlation of Fosfomycin Sodium in Six Organic Solvents at	https://www.doi.org/10.1021/je100331h
Excess Molar Enthalpies of Methyl Acetoacetate + Methanol + 2-Methoxyethyl 4-methylbenzenesulfonate +	
Solubility of Ethyl 4-(Methylamino)benzoate in Binary (Methanol + 1-Propanol) at Temperatures from (288.2, 298.2, 313.2, and 328.2) K and p	
Measurements for Organic Solutes and Solubility of 3,5-Dinitrobenzoic Acid in Supercritical Carbon Dioxide with Ethanol as Cosolvent at Temperatures from (308 to 328) K and Pressures from (10.0 to 21.0) MPa:	

Solubility of Potassium Clavulanate in Ethanol, 1-Propanol, 1-Butanol, 2-Propanol, and 2-Methyl-1-propanol	https://www.doi.org/10.1021/je050145r
Phase equilibrium measurements and thermodynamic modeling of methane + 1-propanol + water at 298.15 K	https://www.doi.org/10.1016/j.jct.2012.03.028
Solubility determination and thermodynamic modelling of 4-hydroxybenzoic acid in monosolvents and binary mixtures of ethanol, ethyl acetate and 1,4-dioxane from 278.15 to 318.15 K	https://www.doi.org/10.1016/j.jct.2016.08.007
Solubility of 2,4-Dihydro-5-methyl-2-(4-methylphenyl)-3H-pyrazol-3-one in Twelve Organic Solvents from (278.15 to 313.15) K and Mixing Properties of Solutions	https://www.doi.org/10.1021/acs.jced.6b00929
Solubility of 4,5,7-Triacetoxyflavanone in Fourteen Neat Organic Solvents from (293.15 to 313.15) K and Their Aqueous Mixtures:	https://www.doi.org/10.1021/je034101z
Henry's constants and activity coefficients of some organic solutes in 1-butyl-3-methylimidazolium hydrogen sulfate and	https://www.doi.org/10.1016/j.jct.2012.08.022
Thermophysical properties of {1-hexanol + propan-1-ol}: A first measurement and correlation of	https://www.doi.org/10.1016/j.jct.2010.11.007
Solubilities of apigenin and apigenin 7-O-methyl-beta-cyclodextrin in seven	https://www.doi.org/10.1016/j.jct.2010.09.013
Solubility and Solvent Effects of Allopurinol in 15 Neat Solvents:	https://www.doi.org/10.1021/acs.jced.8b00430
(Liquid + liquid) equilibria of four alcohol-water systems containing	https://www.doi.org/10.1016/j.jct.2016.06.010
Activity coefficients at infinite dilution of organic solutes in	https://www.doi.org/10.1016/j.jct.2015.02.023
Determination and correlation of solubility of spironolactone form II in pure solvents and binary solvent	https://www.doi.org/10.1016/j.jct.2014.07.011
Measurement and correlation of vapor pressure for liquid equilibrium apparatus:	https://www.doi.org/10.1016/j.fluid.2008.05.002
Equilibrium Phase Behavior of Water + Propan-1-ol or Propan-2-ol + Cesium Chloride at (298.15, 308.15, and 318.15)	https://www.doi.org/10.1021/je050265z
Volumetric and Viscometric Study of Binary Systems of Ethyl Butyrate with Alcohols:	https://www.doi.org/10.1021/je5005752
Molar excess enthalpies of binary mixtures of [tributyl phosphate + (CnH2n+1OH (n = 1 to 3)] at 298.15 K	https://www.doi.org/10.1016/j.jct.2006.08.001
Excess molar enthalpies of methylformate + (1-propanol, 2-propanol, 1-butanol, 2-butanol and 2-methyl-2-propanol) at 298.15 K	https://www.doi.org/10.1016/j.tca.2005.11.041
Solubility of CO2 in Alcohols, Glycols, Ethers, and Ketones at High Pressures	https://www.doi.org/10.1021/je101344v
Vapor pressures and osmotic coefficients of binary mixtures of 1-propanol, 2-propanol, 1-butanol, and methylformate +	https://www.doi.org/10.1016/j.jct.2013.05.027
Thermodynamic Study of Solubility for Lithium Acetate in Nine Monosolvents and Two Binary Solvent	https://www.doi.org/10.1021/acs.jced.8b00551
Distillation of Dipropyl Ether + 1-Propyl Mixtures from 278.15 to 338.15 K	https://www.doi.org/10.1021/je0604479
Measurement and Correlation for the Solubility of Dimethyl 1,4-Cyclohexanedione-2,5-dicarboxylate	https://www.doi.org/10.1021/je2000292
Thermodynamic Properties of Mixtures Containing Ionic Liquids. 5. Activity	https://www.doi.org/10.1021/je050125p
Measurement and modeling of osmotic coefficients of binary mixtures (alcohol	https://www.doi.org/10.1016/j.jct.2011.01.010
Volumetric and Viscometric Study of Binary Mixtures of Ethyl Acetate with	https://www.doi.org/10.1016/j.fluid.2019.01.028
Solubility of Tetraalkylammonium Bis(trifluoromethylsulfonyl)imide [N1444][Tf2N]:	https://www.doi.org/10.1016/j.jct.2016.07.023
Liquid Mixtures of Ethanol + Propan-1-ol at Temperatures	https://www.doi.org/10.1021/je300632p
Vapour + liquid equilibria (VLE) for 1-propanol + 2-methylpropan-1-ol binary mixtures at 313.15 K and 0.1 MPa :	https://www.doi.org/10.1016/j.jct.2015.06.023
Chloroform, Acetone, 1-Propanol, Dimethyl Sulfoxide, and Their Binary and Ternary Systems at	https://www.doi.org/10.1021/je800676s
Solubility of Benzoic Acid in Mixed Solvents	https://www.doi.org/10.1021/acs.jced.8b00067
Solubility of Pyrazinamide in Ethanol, 1-Propanol, 2-Propanol, and Their Binary Mixtures from	https://www.doi.org/10.1016/j.jct.2016.09.033
Solubility of 4-Hydroxybenzoic Acid in Water + Ethanol, + 1-Propanol, + 2-Propanol, + 1-Butanol, + 2-Butanol, + 2-Methyl-1-propanol, + 2-Methyl-2-propanol Binary	https://www.doi.org/10.1016/j.jct.2015.04.009
Solubility of Sorbic Acid in Ethanol, 1-Propanol, 2-Propanol, 1-Butanol, 2-Butanol, and 2-Methyl-2-propanol at	https://www.doi.org/10.1016/j.jct.2013.12.002
Solubility of Sulfamethoxazole in Ethanol, 1-Propanol, 2-Propanol, 1-Butanol, 2-Butanol, 2-Methyl-1-propanol, and 2-Methyl-2-propanol at	https://www.doi.org/10.1016/j.jct.2016.04.018
Thermodynamic modeling of 5-nitro-8-hydroxyquinoline in ten organic	https://www.doi.org/10.1016/j.fluid.2013.06.021
Solubility of PhCOOH in lower alcohols:	https://www.doi.org/10.1016/j.fluid.2011.01.003
Solubility Measurement and Modeling of 4,4′-Dihydroxydiphenyl Sulfone in Nine Organic Solvents from	https://www.doi.org/10.1021/acs.jced.5b00714
Binary and Ternary Vapor-Liquid Equilibrium at 328.15 K and Excess Molar Property of Dissolution for the	https://www.doi.org/10.1021/je800972b
Isothermal Vapor-Liquid Equilibrium for Methanol + Propan-1-ol + 1-Propanol + Toluene or + 2-Butanol,	https://www.doi.org/10.1021/je050404c
Thermodynamic Analysis of the Experimental Equilibria for the	https://www.doi.org/10.1021/acs.jced.5b00557
Solubility of Esterification of 4-nitrotoluene-2-sulfonate in (propanol,	https://www.doi.org/10.1016/j.jct.2011.08.010
Measurements and Correlation of Solubility of Isobutene with C1 to C4 Linear Primary Alcohols.	https://www.doi.org/10.1016/j.jct.2009.12.001
Density, Viscosity, and Refractive Index Properties for the Binary Mixtures of (Ethylene glycol + water) containing 1-ethylpyridinium	https://www.doi.org/10.1021/je200707b
Activity coefficients for binary mixtures of several alcohols at 323.15 K:	
Solubility and Thermodynamic Properties of n-Butylammonium Acetate Ionic Liquid + Alkanols at Several Temperatures:	

Separation of the mixture (isopropyl alcohol + diisopropyl ether + n-propanol): Entrainer selection,	https://www.doi.org/10.1016/j.jct.2019.03.018
Correlation of solubility and calculation of thermodynamic properties of	https://www.doi.org/10.1016/j.fluid.2014.12.037
Extraction of Butanol and Solubilities of Fatty Liquid Fatty Alcohols in	https://www.doi.org/10.1021/je4009816
Measurements of Infinite Activity Coefficients of Volatile Organic	https://www.doi.org/10.1016/j.jct.2012.09.017
Density and Viscosity Measurements of Binary Alkanol Mixtures from (293.15 to	https://www.doi.org/10.1021/je3011634
Solubility Measurement and Correlation of Salicylic Acid in	https://www.doi.org/10.1021/acs.jced.8b00720
Thermodynamic Properties of Levetiracetam in Pure and Mixed	https://www.doi.org/10.1021/acs.jced.7b00542
Measurement and Correlation of Solubility of 2-Hydroxy-2-naphthoic	https://www.doi.org/10.1021/acs.jced.8b00578
Thermodynamic properties of (tetradecane + benzene + toluene, +	https://www.doi.org/10.1016/j.jct.2005.10.022
Solubility of Butanedioic Acid in Different Solvents at Temperatures	https://www.doi.org/10.1021/je900021g
Densities and Excess Molar Volumes of N-Methylmorpholine + 1-Alkanol	https://www.doi.org/10.1021/je7005055
Thermodynamic Properties of Polymorphs of 2,2'-Thiodiethylene	https://www.doi.org/10.1021/je500869p
Thermodynamics of Fluconazole	https://www.doi.org/10.1021/je4010257
Solubility in Various Solvents at Different Temperatures:	https://www.doi.org/10.1016/j.jct.2014.12.023
Thermodynamic study of solubility for pyrazinamide in ten solvents from T =	https://www.doi.org/10.1016/j.jct.2017.04.014
Solubility of Anthracene in C1-C3 Alcohols from (298.2 to 333.2) K and	https://www.doi.org/10.1021/je1001945
Study on the Solubilities of Three Mixtures with 2-Propanone at	https://www.doi.org/10.1021/acs.jced.7b00431
Mononitro-Substituted Products of Nitration of Phenol at Several	https://www.doi.org/10.1016/j.fluid.2015.04.009
Isothermal vapor−liquid equilibrium for ethanol + 1-propanol	https://www.doi.org/10.1016/j.jct.2017.02.022
Comparative study of physical properties of binary mixtures of	https://www.doi.org/10.1021/je2004066
Measurement and Calculation of Liquid−Liquid Equilibria of Ternary and	https://www.doi.org/10.1021/acs.jced.8b00023
Quaternary Systems Containing Water, Propan-1-ol and	https://www.doi.org/10.1016/j.jct.2013.05.030
Activity coefficients at infinite dilution of organic solutes in the ionic liquid	https://www.doi.org/10.1021/acs.jced.9b00179
Solid−Liquid Equilibrium Measurements for Isosorbide and Isomannide in	https://www.doi.org/10.1016/j.jct.2008.01.008
Measurement and correlation of solubility of Florfenicol form A in	https://www.doi.org/10.1021/acs.jced.8b00043
Solubility of Florfenicol form A in Twelve Pure and Binary Solvents:	https://www.doi.org/10.1021/je0600956
Surface thermodynamics of binary mixtures of aliphatic alcohols in heavy	https://www.doi.org/10.1016/j.jct.2012.02.017
Activity coefficients at infinite dilution and physicochemical properties for	https://www.doi.org/10.1016/j.jct.2013.01.007
Solubilities of Betulinic Acid in Thirteen Organic Solvents at Different	https://www.doi.org/10.1021/je200531k
Properties of pure 1-ethyl-3-methylimidazolium ethylsulfate and its binary	https://www.doi.org/10.1016/j.jct.2015.04.027
Density and Viscosity Measurements of 1,1,3,3-Tetramethylguanidine Imidazole	https://www.doi.org/10.1021/je700700f
Data of the Ternary Mixtures 1-Propanol or 2-Propanol + Water +	https://www.doi.org/10.1021/je200195q
Measurement, Correlation, and Prediction of Physical Properties of the Ternary	https://www.doi.org/10.1016/j.fluid.2012.02.011
Solubility Modeling and Solvent Effect of 2-Amino-6-chloropurine in Twelve	https://www.doi.org/10.1021/acs.jced.8b01014
Solubilities of Betulin in Fourteen Organic Solvents at Different	https://www.doi.org/10.1021/je700069g
Determination and Modeling of Solubility of 4-Aminobenzamide in	https://www.doi.org/10.1021/acs.jced.8b01144
KDB:	https://www.cheric.org/files/research/kdb/mol/mol819.mol
Excess Molar Enthalpies, Molar Heat Capacities, Densities, Viscosities, and	https://www.doi.org/10.1021/je0601513
Determination and thermodynamic modelling for 2-methyl-6-nitroaniline	https://www.doi.org/10.1016/j.jct.2016.10.043
Solute Vaporization Enthalpies of Pyrene, Benzo[ghi]perylene, and	https://www.doi.org/10.1021/acs.jced.8b00033
Phase Measurements and other derived separation properties of the binary	https://www.doi.org/10.1016/j.jct.2016.10.023

1-Butyl-1-methylpyrrolidinium
Solubility of Butanedioic Acid in Different Solvents at Temperatures
Solubility of Florfenicol form A in
Solubility of Florfenicol form A in Twelve Pure and Binary Solvents:
Solubilities of Betulin in Fourteen Organic Solvents at Different
Solubilities of Betulinic Acid in Thirteen Organic Solvents at Different
Solubilities of Betulin in Fourteen Organic Solvents at Different

Study of the effect of increasing the chain length of the alkane in excess molar enthalpies of mixtures containing diethyl ether, 1-propanol, alkane:	https://www.doi.org/10.1016/j.fluid.2008.06.011
Solubility of Sebacic Acid in Binary Mixtures Containing Alcohols:	https://www.doi.org/10.1021/je800549f
Determination of Solubility and Thermodynamic Properties of Benzophenone in Different Pure Solvents:	https://www.doi.org/10.1021/acs.jced.8b00196
Solubility of lansoprazole in different solvents:	https://www.doi.org/10.1016/j.fluid.2012.06.011
Partition Coefficients of Organic Compounds in Four New Tetraalkylammonium Bis(trifluoromethylsulfonyl)imide Ionic Liquids Using Inverse Gas Chromatography:	https://www.doi.org/10.1021/je200637v
Solubility of Clopidogrel Hydrogen Sulfate (Form II) in Different Solvents:	https://www.doi.org/10.1021/je100022w
Solubility and Data Correlation of Isoniazid in Different Pure and Binary Mixed Solvent Systems from 283.15 K to 323.15 K:	https://www.doi.org/10.1021/acs.jced.8b00785
Density and related thermodynamic excess properties of binary mixtures of butan-2-ol + (Methanol, Ethanol, 1-Propanol, and 2-Propanol) at T = (298.15, 308.15, and 318.15) K:	https://www.doi.org/10.1021/je5007604
Binary Mixtures of 2,4-Pentanedione + C1-C2 Alkanols and Chloroform: Temperature Effect on the Liquid–Liquid Equilibria of Some Ternary Systems and Correlation with UNIQUAC and NRTL Models:	https://www.doi.org/10.1021/je700492k
Solubility Determination and Thermodynamic Modeling of Sodium Thiocyanate in Pure and Binary Solvent Mixtures from T = (293.15 to 338.15) K:	https://www.doi.org/10.1021/je050515b
Study of Surface and Interfacial Properties of Binary Systems of DMSO + Alcohols + Water + K2CO3 Liquid–Liquid Equilibria of Ternary Systems with Salting-Out Ability at 298.15 K:	https://www.doi.org/10.1021/acs.jced.7b00245
KDB Pure (Korean Thermophysical Properties Databank):	https://www.doi.org/10.1016/j.fluid.2011.05.016
Solubility of climbazole in various alcohols at different temperatures:	https://www.doi.org/10.1016/j.jct.2013.04.009
Determination and correlation of solubility of linezolid form II in different pure and binary solvents:	https://www.cheric.org/research/kdb/hcprop/showprop.php?cmpid=819
Refractive Indices and Deviations in Refractive Indices for Binary Mixtures of 1-Ethyl-3-methylimidazolium Ethyl Sulfate with Methanol, Ethanol, 1-Propanol, and 2-Propanol at T = (293.15 to 323.15) K:	https://www.doi.org/10.1016/j.jct.2014.04.011
Liquid–liquid equilibria of the systems dimethyl ether + propanol + water and Measurement of Ternary Liquid–Liquid Equilibrium Data for Water + 1-Propanol + Ethers at 298.15 K:	https://www.doi.org/10.1016/j.fluid.2016.10.021
Solubility Data for Boflumilast and Measurement of Several Cosolvents	https://www.doi.org/10.1021/je9006037
Measurements of Activity Coefficients at Infinite Dilution of Organic Solutes in 4-Methyl-N-butylpyridinium Bis(trifluoromethylsulfonyl)imide:	https://www.doi.org/10.1016/j.fluid.2007.08.008
Solubility of Vitamin A in Water, Ethanol, Propan-1-ol, Propan-2-ol, Butan-1-ol, and Butan-2-ol from (278.2 to 323.2) K:	https://www.doi.org/10.1021/acs.jced.9b00710
Solubility of 2-Hydroxybenzoic Acid in Water from (278.15 to 323.15) K:	https://www.doi.org/10.1016/j.jct.2009.06.011
Measurements of activity coefficient at infinite dilution for organic solutes in water in the ionic liquid 1-Ethyl-3-methylimidazolium Ethyl Sulfate:	https://www.doi.org/10.1016/j.fluid.2018.01.019
Solubility of Vitamin A in Water, Ethanol, Propan-1-ol, Propan-2-ol, Butan-1-ol, and Butan-2-ol from (278.2 to 323.2) K:	https://www.doi.org/10.1021/je900687y
Solubility of 2-Hydroxybenzoic Acid in Water from (278.15 to 323.15) K:	https://www.doi.org/10.1021/je300980k
Effect of temperature on the phase separation ability of KCl + 2-Propanol + water (salt + alcohol + water) aqueous two phase systems composed of propanol-2:	https://www.doi.org/10.1016/j.jct.2019.04.016
Determination of the critical temperature and extension of the liquid–liquid coexistence curve for binary mixtures containing hydrocarbons and sulfur compounds:	https://www.doi.org/10.1016/j.jct.2016.07.017
Activity coefficients at infinite dilution measurements for alkanols in 1-ethyl-3-methylimidazolium ethyl sulfate:	https://www.doi.org/10.1016/j.jct.2010.02.006
Experimental Studies on the Thermal Conductivity of Methanol + 1-Propanol + Water Ternary Mixtures:	https://www.doi.org/10.1016/j.jct.2019.105881
Isobaric Vapor–Liquid Equilibrium for the Binary Systems of Sec-butyl Acetate and Ethanol, 1-Propanol, or 2-Propanol at 101.3 kPa:	https://www.doi.org/10.1016/j.jct.2019.02.027
Correlations at 298.15 K and Ambient Pressure:	https://www.doi.org/10.1021/je700099w
Activity Coefficients at Infinite Dilution of Polar Solutes in 1-Hexadecyl-3-methylimidazolium Tetrafluoroborate Using Gas Liquid Chromatography:	https://www.doi.org/10.1021/acs.jced.7b00445
Excess Molar Enthalpies of Benzyl Alcohol + Alkanols (C1–C6) and Their Correlations at 298.15 K and Ambient Pressure:	https://www.doi.org/10.1021/je100095p
Activity Coefficients at Infinite Dilution for 2-Methyl-2-propanol + Methanol or Ethanol or 1-Propanol + Nitromethane) at 298.15 K and Atmospheric Pressure:	https://www.doi.org/10.1021/je100830b
Densities and Viscosities of Binary Mixtures of Aniline with Chloroalkanols at Temperatures from 293.15 to 313.15 K:	https://www.doi.org/10.1021/je700640r
Thermodynamics and activity coefficients at infinite dilution for 2-methyl-2-propanol at temperatures from 293.15 to 313.15 K:	https://www.doi.org/10.1007/s10765-007-0204-0
Comparative Tensiometry of 1-Propanol + Water and 1-Propanol + Ethylene Glycol Mixtures: Liquid–Liquid Equilibria and Interfacial Properties:	https://www.doi.org/10.1016/j.jct.2018.08.028
Thermodynamics of biofuels. Excess enthalpies for binary mixtures of 2-methoxy-2-methylpropane + 1-propanol + 1-butanol at 293.15 K:	https://www.doi.org/10.1021/je060307z
Investigation of the solvent effect on the calorimetric data of 2-butoxyethanol with aliphatic and hydrocarbons at different temperatures using a new flow microcalorimeter and UNIQUAC model:	https://www.doi.org/10.1021/je0302295
Gas–liquid chromatographic determination of activity coefficients at infinite dilution of hydrocarbons and alcohols at 298 K in 1-alkyl-3-methylimidazolium Hexafluorophosphate and bis(oxalato)borate:	https://www.doi.org/10.1021/je200822w
Activity coefficients at infinite dilution measurements for alkanols in 1-Butyl-3-methylimidazolium Tetrafluoroborate, with Benzene, Acetonitrile, and 1-Propanol at T = (293.15 to 343.15) K:	https://www.doi.org/10.1016/j.fluid.2006.02.023
Thermodynamics of biofuels. Excess enthalpies for binary mixtures of 2,5-dioxolane, 2-methoxy-2-methylpropane, and 1-propanol at 293.15K:	https://www.doi.org/10.1016/j.jct.2008.12.004
Investigation of the solvent effect on the calorimetric data of 2-butoxyethanol with aliphatic and hydrocarbons at different temperatures using a new flow microcalorimeter and UNIQUAC model:	https://www.doi.org/10.1016/j.jct.2016.08.003
Measurements of activity coefficients at infinite dilution of hydrocarbons and alcohols in 1-Alkyl-3-methylimidazolium Ionic Liquids:	https://www.doi.org/10.1016/j.fluid.2010.08.016
	https://www.doi.org/10.1021/je700266n

Solubility of 1H-1,2,4-Triazole in Ethanol, 1-Propanol, 2-Propanol, 1,2-Propanediol, Ethyl Formate, Methyl Acetate, Ethyl Acetate, and Butyl Acetate at (283 to 363) K	https://www.doi.org/10.1021/je060452c
Thermodynamic solubility modelling, solvent effect and preferential solvation of curcumin in aqueous co-solvent solutions of (n-propanol, ethanol, isopropanol and dimethyl sulfoxide):	https://www.doi.org/10.1016/j.jct.2019.02.016
(Vapour + liquid) equilibria of the binary mixtures of m-cresol with C1-C4 aliphatic alcohols at [illegible] kPa:	https://www.doi.org/10.1016/j.jct.2006.03.006
Indirect Analysis of a Homogeneous Ternary Mixture:	https://www.doi.org/10.1021/acs.jced.5b01024
Isobaric vapor-liquid equilibria for the 1-propanol + water + 1-ethyl-3-methylimidazolium dicyanamide system at 100 kPa:	https://www.doi.org/10.1016/j.jct.2017.06.005
Solubilities of Menadione Sodium Bisulfite in Water + (Methanol, Ethanol, 1-Propanol, 2-Propanol, 1,2-Propanediol, and Glycerin, Respectively) from (297.87 to 337.76) K:	https://www.doi.org/10.1021/je800022j
Experimental determination and correlation of the solubility of [illegible] (DMHP) in six different solvents:	https://www.doi.org/10.1016/j.jct.2015.08.028
Isobaric Vapor-Liquid Equilibria of 1-Propanol + Water + Trifluoromethanesulfonate-Based Ionic Liquid Ternary Systems at 100 kPa:	https://www.doi.org/10.1021/je2003447
Liquid-liquid equilibrium for the quaternary system propionic acid + n-propanol + n-propyl propionate + water at 293.15, 313.15 and 333.15 K:	https://www.doi.org/10.1016/j.fluid.2016.05.033
Liquid–Liquid Equilibrium of Aqueous Two-Phase Systems Composed of Hydrophilic Alcohols (Ethanol/2-Propanol/1-Propanol) and MgSO4/ZnSO4 at (303.15 and 313.15) K	https://www.doi.org/10.1016/j.tca.2012.07.022
Abraham model correlations for describing the thermodynamic properties of solute transfer into [illegible] acetate based on headspace chromatographic and solubility measurements:	https://www.doi.org/10.1016/j.jct.2018.05.003
Activity Coefficients at Infinite Dilution of Organic Solutes in 1-Butyl-3-methylimidazolium Nitrate Using Gas–Liquid Chromatography:	https://www.doi.org/10.1021/je200050q
Vapor Pressures and Activity Coefficients of (1-Propanol + 1,8-Cineole) at 10 Temperatures between 278.15 K and 323.15 K:	https://www.doi.org/10.1021/je300736k
Solubility Measurement and Modeling of 1-(3-nitrophenyl)Ethanone and 1-(4-nitrophenyl)Ethenone in Nine Pure Organic Solvents from T = (278.15 to 318.15) K and Mixing Properties of Solutions:	https://www.doi.org/10.1021/acs.jced.8b00192
Measurement and Correlation of Solubilities of Cefpirome Sulfate in Aqueous Alcohol Solution between (278.15 and 308.15) K:	https://www.doi.org/10.1021/je100870x
Solubility of citric acid in water, ethanol, n-propanol and in mixtures of ethanol + water:	https://www.doi.org/10.1016/j.fluid.2013.05.014
Measurement and Correlation of Solubility of -Aminobutyric Acid in Different Binary Solvents:	https://www.doi.org/10.1021/acs.jced.5b00829
Solubility and mixing thermodynamic properties of levamisole hydrochloride in twelve pure solvents at various temperatures:	https://www.doi.org/10.1016/j.jct.2019.105882
Liquid-Liquid Phase Behavior of Solutions of 1,3-Dimethylimidazolium- and 1-Methyl-3-propylimidazolium Bis((trifluoromethyl)sulfonyl)amide (C1,3mimNTf2) with n-Alkyl Alcohols:	https://www.doi.org/10.1021/je400253e
Activity coefficients at infinite dilution for organic solutes and water in [illegible] acetate:	https://www.doi.org/10.1016/j.jct.2015.05.014
Excess molar enthalpies of the ternary mixture n-propanol + acetic acid + water at 313.15 K and atmospheric pressure:	https://www.doi.org/10.1016/j.fluid.2015.07.053
Isobaric (vapour + liquid + liquid) equilibrium data for (di-n-propyl ether + n-propyl alcohol + water) and (diisopropyl ether + isopropyl alcohol + water) systems at 100 kPa:	https://www.doi.org/10.1016/j.jct.2008.01.002
Measurement and Correlation of the Solubility of Tetramethylpyrazine in Nine Monosolvents and Two Binary Solvent Systems:	https://www.doi.org/10.1021/acs.jced.8b00888
Solubility of Ethyl p-Aminobenzoate in Six Alcohols within (283.15 to 327.15) K:	https://www.doi.org/10.1021/acs.jced.6b00037
Experimental determination and modeling of excess molar volumes, viscosities and refractive indices of the binary systems (pyridine + 1-propanol, +1,2-propanediol, +1,3-propanediol, and +glycerol). New UNIFAC-VISCO parameters determination:	https://www.doi.org/10.1016/j.jct.2012.06.031
Solubility of 2-nitro-p-phenylenediamine in nine pure solvents and mixture of (methanol + N-methyl-2-pyrrolidone) from T = (283.15 to 318.15) K: Determination and modelling:	https://www.doi.org/10.1016/j.jct.2017.01.006
Solubility of antioxidant 1010 in pure alkanols [illegible]	https://www.doi.org/10.1016/j.fluid.2008.11.005
Activity Coefficients at Infinite Dilution for Organic Solutes Dissolved in Three 1-Alkyl-1-methylpyrrolidinium Bis(trifluoromethylsulfonyl)imide Ionic Liquids Bearing Short Linear Alkyl Side Chains of Three to Five Carbons:	https://www.doi.org/10.1021/je4001894
A systematic study on volumetric and transport properties of binary systems 1-propanol + [illegible] methyl oleate at different temperatures:	https://www.doi.org/10.1016/j.fluid.2018.05.028
Topological investigations of the molecular species and molecular interactions [illegible] Experimental and modelling:	https://www.doi.org/10.1016/j.tca.2009.03.014
Properties of pure 1-methylimidazolium acetate ionic liquid and its binary mixtures with alcohols:	https://www.doi.org/10.1016/j.jct.2012.01.013
Thermodynamic and spectroscopic properties of binary mixtures of n-butylammonium butanoate ionic liquid with alcohols at T = (293.15-313.15) K:	https://www.doi.org/10.1016/j.jct.2016.09.007
4-Nitrobenzylcyanide Solubility in Several Nonaqueous Solvent Mixtures: Determination and Modeling:	https://www.doi.org/10.1021/acs.jced.9b00005
Binary Liquid-Liquid Equilibrium (LLE) for Dibutyl Ether (DBE) + Water from (288.15 to 318.15) K and Ternary LLE for Systems of DBE + C1-less thanless than C4 Alcohols + Water at 298.15 K:	https://www.doi.org/10.1021/je800446c
Evaluation of the 2-Methoxyethanol as Entrainer in Ethanol-Water and 1-Propanol-Water Mixtures:	https://www.doi.org/10.1021/je400791n
The study of activity coefficients at infinite dilution for organic solutes and water in 1-butyl-4-methylpyridinium dicyanamide, [B4MPy][DCA] using GLC:	https://www.doi.org/10.1016/j.jct.2013.09.007
Excess molar enthalpies of ethylformate and (1-propanol, 2-propanol, 1-butanol, 2-butanol, and 1-pentanol) at T = (298.15 and 308.15) K and p = (5 and 10) MPa:	https://www.doi.org/10.1016/j.jct.2005.04.015
Cloud point measurements of 1-butyl-2,3-dimethylimidazolium tetrafluoroborate with alcohols:	https://www.doi.org/10.1016/j.jct.2010.07.005
Assessment of Pyrrolidinium-Based Ionic Liquid for the Separation of Binary Mixtures Based on Activity Coefficients at Infinite Dilution:	https://www.doi.org/10.1021/acs.jced.9b00341
Phase equilibria of glycerol containing systems and their description with the Cubic-Plus-Association (CPA) Equation of State:	https://www.doi.org/10.1016/j.fluid.2009.03.011
Excess Molar Volumes of 1,3-Diethyl Propanedioate with Methanol, Ethanol, Propan-1-ol, Propan-2-ol, Butan-2-ol, 2-Methyl-propan-1-ol, and Pentan-1-ol at T = (288.15, 298.15, 313.15, and 328.15) K:	https://www.doi.org/10.1021/je100100n
Solubility of Fumaric Acid in Propan-2-ol, Ethanol, Acetone, Propan-1-ol, and Water:	https://www.doi.org/10.1021/je9001637
New ionic liquid [P4,4,4,4][NTf2] in bio-butanol extraction on investigation of limiting activity coefficients:	https://www.doi.org/10.1016/j.fluid.2018.07.028

Determination of Ternary Liquid Liquid Equilibria for Dimethyl Carbonate + 2-Methyl-1-propanol or 1-Methyl-2-propanol + Water Systems at 303.15 and 313.15 K.	https://www.doi.org/10.1021/acs.jced.6b00763
Activity Coefficients at Infinite Dilution Measurements for Organic Solutes and Water in 1-Hexyl-3-methylimidazolium Bis(trifluoromethylsulfonyl)imide by GLC at Several Temperatures.	https://www.doi.org/10.1021/je900890u
Measurement and Correlation of Solubility of Amorphous Cefmetazole Sodium in Pure Solvents and Binary Solvent Mixtures.	https://www.doi.org/10.1021/acs.jced.6b00741
Activity Coefficients at Infinite Dilution Measurements for Alkanediamines as Solutes in Ionic Liquids by GLC.	https://www.doi.org/10.1021/je200628n
Determination and correlation of the solubility for diosgenin in alcohols solvents.	https://www.doi.org/10.1016/j.jct.2012.02.002
Isobaric Vapor-Liquid Equilibria for 1-Propanol + Water + 1-Ethyl-3-methylimidazolium Trifluoromethanesulfonate at 101.3 kPa.	https://www.doi.org/10.1021/je800436u
Experimental Determination and Correlation of Phase Equilibria for the Ethanol + 1-Propanol and Propan-2-ol + 1-Propanol Systems.	https://www.doi.org/10.1021/je060112r
Isobaric vapor-liquid equilibria for the binary systems 1-propyl alcohol + 2-methyl-1-propanol and 1-butanol + 2-methyl-1-propanol at 20 and 101.3 kPa.	https://www.doi.org/10.1016/j.fluid.2006.06.017
Measurement and Correlation of the Solubility of 2-Cyanoacetamide in 14 Pure Solvents and the Mixing Properties of Solutions.	https://www.doi.org/10.1021/acs.jced.8b01205
Solubility of L-tryptophan and L-tyrosine in Several Organic Solvents: Determination and Solvent Effect.	https://www.doi.org/10.1016/j.fluid.2013.06.047
Measurement and Correlation of the Solubility of e-CL-20 in 12 Organic Solvents at Temperatures Ranging from 278.18 to 313.15 K.	https://www.doi.org/10.1021/acs.jced.9b00258
Experimental Measurement and Modeling of Vapor Liquid Equilibrium for the Ternary Systems Water + 2-Propanol + Ethylene Glycol, Water + 2-Propanol + Diethylene Glycol, and Water + 2-Propanol + Triethylene Glycol:	https://www.doi.org/10.1021/acs.jced.6b00761
Physical Properties of Binary Mixtures of Ethanol + Ionic Liquid 1-Ethyl-3-methylimidazolium Ethylsulfate and 1-Propanol + Ionic Liquid 1-Ethyl-3-methylimidazolium Ethylsulfate at Several Temperatures.	https://www.doi.org/10.1021/acs.jced.6b00264
Isobaric Vapor-Liquid Equilibrium Data of Diethyl Succinate + Alcohols at 7.5 kPa: 1-Propanol, 2-Butanol, 2-Methyl-1-propanol, and 2-Methyl-2-propanol.	https://www.doi.org/10.1021/je700029q
Isobaric Vapor-Liquid Equilibrium Data of Diethyl Succinate + Alcohols at 7.5 kPa: 1-Propanol, 2-Butanol, 2-Methyl-1-propanol, and 2-Methyl-2-propanol.	https://www.doi.org/10.1021/je300696c
Vapor-Liquid Equilibrium Data of Diethyl Succinate + Alcohols at 7.5 kPa: 1-Propanol, 2-Butanol, 2-Methyl-1-propanol, and 2-Methyl-2-propanol.	https://www.doi.org/10.1021/je0505164
Solubility of Coumarin and Vanillin in Ethanol, Isopropanol, Propan-1-ol, Butan-1-ol, Butan-2-ol, and Pentan-1-ol from 283.15 to 323.15 K.	https://www.doi.org/10.1016/j.fluid.2014.12.034
Isobaric Vapor-Liquid Equilibrium Data of Diethyl Succinate + Alcohols at 7.5 kPa: 1-Propanol, 2-Butanol, 2-Methyl-1-propanol, and 2-Methyl-2-propanol.	https://www.doi.org/10.1021/je0505164
Prediction of Activity Coefficients at Infinite Dilution for Solutes in Polymer Solvents: Improved Method Based on Functional-Group Interaction Parameters for the Entropy Part of the Flory-Huggins Theory.	https://www.doi.org/10.1016/j.jct.2013.09.026
Measurements of activity coefficients at infinite dilution for organic solutes in 1-hexadecyl-3-methylimidazolium tetrafluoroborate and 1-octadecyl-3-methylimidazolium tetrafluoroborate ionic liquids by inverse gas chromatography.	https://www.doi.org/10.1016/j.jct.2012.03.005
Vapor-liquid equilibrium for binary mixtures of N-methylpiperazine with 1-propanol, 2-propanol and 2-methyl-2-propanol at 298.15 and 323.15 K.	https://www.doi.org/10.1016/j.fluid.2005.05.012
Isobaric Vapor-Liquid Equilibria for Ternary Mixtures of 1-Propanol + 2-Propanol + 2-Butanol at 20.0 kPa.	https://www.doi.org/10.1021/je060314n
Liquid-liquid equilibrium of 1-propanol, 2-propanol or 2-methyl-2-propanol + water + 2-butanol + sodium sulfite at 298.15, 308.15 and 318.15 K.	https://www.doi.org/10.1016/j.fluid.2012.05.012
Solubility of Glyphosate in Ethanol + Water, 1-Propanol + Water and 2-Propanol + Water from (293 to 333) K:	https://www.doi.org/10.1021/je901006c
Solubilities of Phloretin in 22 Solvents at Different Temperatures:	https://www.doi.org/10.1021/je101168w
Liquid Liquid Equilibria of the Water + 1-Propanol + 1-Ethyl-3-methylpyridinium Ethylsulfate Ionic Liquid System.	https://www.doi.org/10.1021/acs.jced.6b00524
Binodal Curves and Tie-Lines of Aliphatic Alcohols + Diammonium Hydrogen Citrate Aqueous Two-Phase Systems at T = 298.15 K.	https://www.doi.org/10.1021/je3001574
Dynamic Viscosity of the Binary System 1-Propanol + Solvent Naphtha: Experimental Study and Modeling.	https://www.doi.org/10.1021/je9002477
Ternary System Study of the Ability of 1-Propanol, 2-Propanol, and Their Mixtures to Dissolve Cholesterol: Function of Temperature and Pressure.	https://www.doi.org/10.1016/j.fluid.2016.02.004
Activity coefficients at infinite dilution of organic solvents and water in 1-butyl-3-methylimidazolium dicyanamide: A literature review.	https://www.doi.org/10.1016/j.fluid.2010.10.008
Activity coefficients at infinite dilution of organic solutes in the ionic liquid 1-hexyl-1-ole separation.	https://www.doi.org/10.1016/j.jct.2017.03.004
Separation of binary mixtures methanol/ethyl acetate, cyclohexane/cyclohexene and hydrophilic Alcohols with three ionic liquids based on limiting activity coefficients:	https://www.doi.org/10.1021/acs.jced.5b00771
A novel dynamic recirculating apparatus for vapour-liquid equilibrium measurements at moderate pressures and temperatures.	https://www.doi.org/10.1016/j.fluid.2013.07.044
Binary Mixtures of NTf2 Ionic Liquids with 1-Propanol: Volumetric, Viscosities, and Excess Gibbs Energy of Activation for Viscous Flow, for Binary and ternary solubility of carbendazim in lower dimethylformamid + chloroform at different temperatures.	https://www.doi.org/10.1021/acs.jced.6b00556
Measurement and correlation of solubility of carbendazim in lower dimethylformamid + chloroform at different temperatures.	https://www.doi.org/10.1016/j.tca.2014.05.004
Solubility Determination, Correlation, and Solute-Solvent Molecular Interactions of 6-Aminotetrazole in Various Pure Solvents.	https://www.doi.org/10.1016/j.tca.2017.12.007
Viscosity of didecyldimethylammonium nitrate ionic liquid + 1-propanol or + 2-propanol at several temperatures.	https://www.doi.org/10.1021/acs.jced.9b00385
Density and Speed of Sound Measurements and Theoretical Modeling of Binary Mixtures of Ionic Liquids with Water and Organic Solvents:	https://www.doi.org/10.1016/j.jct.2006.10.009
Compressibilities in Binary Mixtures of 2-Propanol with Several 1-Alkanols at 298.15 K.	https://www.doi.org/10.1007/s10765-005-8101-x
Volumetric and viscometric study with ESD analysis of binary systems with diethyl succinate and alcohols.	https://www.doi.org/10.1016/j.jct.2016.06.016
The use of ionic liquids for separation of binary hydrocarbons mixtures based on gamma infinity data measurements:	https://www.doi.org/10.1016/j.jct.2018.07.024
Solubility of Benzoic Acid in Aqueous Solutions Containing Ethanol or 1-Propanol.	https://www.doi.org/10.1021/je800507m
Vapor-Liquid Equilibrium Data for Binary Systems of 1-Propanol + 1-(1-methylethenyl)-cyclohexene at 40 kPa.	https://www.doi.org/10.1021/je300347z
Research and comparison of solid-liquid phase equilibrium and ternary phase diagrams of {Butan-1-ol + Pentan-1-ol or Hexan-1-ol}	https://www.doi.org/10.1016/j.jct.2019.03.028
1-(3-nitrophenyl)ethanone + 1-(4-nitrophenyl)ethanone in different solvents:	

Excess properties of binary mixtures containing	https://www.doi.org/10.1016/j.jct.2011.11.007
Separation of binary mixtures	https://www.doi.org/10.1016/j.jct.2017.11.017
Densities and Viscosities of Binary	https://www.doi.org/10.1021/je101173w
Solutions of Benzene + 1,3-dioxane + Water,	
Excess molar volumes, excess molar	https://www.doi.org/10.1016/j.fluid.2012.06.031
enthalpies and refractive index	
deviations for binary mixtures of	https://www.doi.org/10.1016/j.jct.2012.03.015
propan-1-ol, butan-1-ol and pentan-1-ol	https://www.doi.org/10.1016/j.fluid.2005.11.024
(Liquid + liquid) equilibria of water +	
(propanol + solvent) at T = 298.2K:	https://www.doi.org/10.1016/j.fluid.2014.09.027
Experimental study of the density and	
derived partial molar volumes of	https://www.doi.org/10.1021/acs.jced.8b01250
complex ternary solution	
Thermodynamics of	https://www.doi.org/10.1021/acs.jced.7b00098
Separation of Isobutyric Acid from	
Dilute Aqueous Solutions by Reactive	https://www.doi.org/10.1021/je300533r
Extraction and Some	
Activity Coefficients at Infinite Dilution	https://www.doi.org/10.1021/je800754w
of Organic Compounds in	
Measurement and Modeling	https://www.doi.org/10.1021/acs.jced.8b00292
of 2-Amino-4,6-dichloropyrimidine in	
Pure Solvent and Binary	https://www.doi.org/10.1021/acs.jced.8b00554
Solvent Mixtures: n-Nonane or	
Determination and Correlation of	https://www.doi.org/10.1021/je300517q
Solubility in Eight	
Solubility measurement and	https://www.doi.org/10.1021/acs.jced.9b00693
thermodynamic correlation of	
Dynamic + Activity Coefficients of	https://www.doi.org/10.1016/j.jct.2006.06.008
Measurement and Correlation of	https://www.doi.org/10.1021/acs.jced.6b00415
Solubility or Reactivity Crystallization of	
Excess Molar Enthalpies of Binary	https://www.doi.org/10.1021/je400709f
Mixtures Containing	
Solubilities of Prostaglandin	https://www.doi.org/10.1021/je101230a
Measurement and Modeling of Vapor	https://www.doi.org/10.1021/je301374c
Liquid Equilibria for Systems	
Determination and Correlation of	https://www.doi.org/10.1021/acs.jced.9b00381
Solubilities of	
Liquid–Liquid Equilibria of Alcohols +	https://www.doi.org/10.1021/acs.jced.5b00200
Liquid–Liquid Equilibrium Data of Acetate	
Esters + Water Systems: Experimental and	https://www.doi.org/10.1016/j.fluid.2012.06.013
Liquid–Liquid Equilibria of Three	https://www.doi.org/10.1021/je500939v
Ternary Systems: {Glycerol + Benzene	
Thermodynamic Investigation of the	https://www.doi.org/10.1021/je0342668
Azeotropic Binary Mixture Water +	
Densities and Viscosities of Binary	https://www.doi.org/10.1016/j.jct.2006.10.016
Solubilities of	
Solubility of 3,5-Dinitro-2-methylbenzoic Acid in 13	https://www.doi.org/10.1021/acs.jced.6b00349
Experimental Measurement and	https://www.doi.org/10.1021/je3007502
Modeling of Ternary Vapor–Liquid	
Thermodynamic and quantum	https://www.doi.org/10.1016/j.jct.2013.08.030
Solubility Determination and Correlation	https://www.doi.org/10.1016/j.fluid.2018.07.024
Thermodynamic properties of solutions	https://www.doi.org/10.1016/j.jct.2008.01.018
Temperature and Composition	https://www.doi.org/10.1021/je8008254
Dependence of the Density and	https://www.doi.org/10.1016/j.fluid.2008.03.012
Measurement and Correlation of	https://www.doi.org/10.1021/acs.jced.8b00902
Solubility of Benzoic Acid and Pure	https://www.doi.org/10.1021/acs.jced.5b00746
High selective water/butan-1-ol	https://www.doi.org/10.1016/j.fluid.2017.06.001
separation on investigation of limiting	
Activity coefficients at infinite dilution	https://www.doi.org/10.1016/j.jct.2008.12.005
measurements for organic solutes and	
Solubility and Dissolution	https://www.doi.org/10.1021/acs.jced.5b00607
Thermodynamics for 1,3,5-	
Equilibrium Solubility of Piperazine in	https://www.doi.org/10.1021/acs.jced.8b01051

Solubilities of Proxetine Hydrochloride Hemihydrate between 286 K and 363 K:	https://www.doi.org/10.1021/je0498762
Separation of binary mixtures based on gamma infinity data using	https://www.doi.org/10.1016/j.jct.2017.12.012
[OMMIM][NTf2] Ionic Liquid and	https://www.doi.org/10.1016/j.jct.2010.05.016
Modeling of Thermodynamic Functions:	https://www.doi.org/10.1016/j.tca.2011.10.030
(C1 C6) 1-Alkanol Binary Mixtures at T	https://www.doi.org/10.1021/acs.jced.8b01193
= 298.15 K.	https://www.doi.org/10.1016/j.jct.2011.04.018
Mixture of x1[C6mim][BF4] + (1 - x1) Ionic Liquid Equilibrium and	https://www.doi.org/10.1016/j.jct.2018.09.017
Molecular Modeling.	https://www.doi.org/10.1016/j.fluid.2013.10.029
Activity coefficients at infinite dilution and physicochemical properties for	https://www.doi.org/10.1021/je4003515
Solubility of dibenzothiophene in nine organic solvents: Experimental	https://www.doi.org/10.1016/j.jct.2013.09.032
Thermodynamic Modeling.: The Ionic Liquid 1-(3-Hydroxypropyl)-3-methylimidazolium	https://www.doi.org/10.1021/je101302p
Bis(trifluoromethylsulfonyl)imide.	https://www.doi.org/10.1021/je500050p
Measurement and Modeling.	https://www.doi.org/10.1021/je600525e
Thermodynamic properties and phase equilibria study of the binary systems	https://www.doi.org/10.1021/je800824h
(n-octane + hydrocarbon, or an alcohol) and Modeling.	https://www.doi.org/10.1016/j.jct.2010.12.019
Separation of sulfur compounds from alkanes with	https://www.doi.org/10.1016/j.fluid.2007.09.021
Phase Equilibria for Reactive Distillation of Propyl Propanoate. Pure	https://www.doi.org/10.1016/j.fluid.2012.05.027
Component Property Data, Solubility Coefficients of Solutes Dissolved in Two	https://www.doi.org/10.1021/je101031j
Solubility of Alkylbenzenes in Binary Liquid-Liquid Equilibria.	https://www.doi.org/10.1021/acs.jced.7b00078
Diisopropyl Ether + Alcohol Solvent Mixtures at 298.15 K.	https://www.doi.org/10.1021/je200481u
Solubilities of Nicotinic Acid in (Methanol, Ethanol, 1-Propanol, 2-Propanol, and 1,2-Propanediol)	https://www.doi.org/10.1021/je700234e
Measurement and correlation of solubility of levofloxacin in six solvents	https://www.doi.org/10.1021/je060496l
at temperatures from 288.15 to 328.15 K.	https://www.doi.org/10.1021/acs.jced.7b00549
2-Methyl-2-butanol (1) + 1-Alkanols (C1 to C5) (2) at 298.15 K.	https://www.doi.org/10.1016/j.jct.2011.05.014
Isobaric Vapor-Liquid Equilibrium for the Binary Systems Dimethyl Disulfide + C1-C4 1-Alkanol at 40.000 and 101.32 kPa.	https://www.doi.org/10.1016/j.jct.2013.05.008
Solid-Liquid Equilibria of Dibenzothiophene and	https://www.doi.org/10.1016/j.fluid.2011.09.027
Dibenzothiophene Sulfone in Organic Solvents.	https://www.doi.org/10.1021/je100502b
Solid-Liquid Equilibria for the Glycine-Alcohol-NaCl-H2O Systems	https://www.doi.org/10.1016/j.jct.2018.02.014
Application of the extended real associated solution (ERAS) theory to	https://www.doi.org/10.1016/j.jct.2011.08.028
excess molar enthalpies of	https://www.doi.org/10.1021/acs.jced.8b00625
benzaldehyde + 1-alkanols (C1 to C5) at T = 298.15 K.	https://www.doi.org/10.1021/je4000197
Solubility of Citric and Oxalic Acid in Different Pure Solvents and Their	https://www.doi.org/10.1016/j.fluid.2005.01.013
Binary Mixtures.	https://www.doi.org/10.1016/j.fluid.2015.09.003
Solubility of Potassium Sodium Tartrate Tetrahydrate in Binary Mixtures of	https://www.doi.org/10.1021/je0603574
Water + Ethanol, + 1-Propanol, + 2-Propanol, + 1-Butanol, + 2-Butanol, + 2-Methyl-2-propanol, + 1,2-Propanediol, and + 2-Methyl-2-propanol + 1,2-Propanediol Mixtures at	https://www.doi.org/10.1021/acs.jced.9b00232
Temperatures	https://www.doi.org/10.1021/je400576e
Application of the extended real associated solution (ERAS) theory to	https://www.doi.org/10.1021/je1008284
excess molar enthalpies of	https://www.doi.org/10.1016/j.fluid.2012.04.008
benzaldehyde + 1-alkanols (C1 to C5) at T = 298.15 K.	https://www.doi.org/10.1021/je100410k
Solubility and Thermodynamic Properties of Sulfamethoxypyridazine in Different	https://www.doi.org/10.1021/acs.jced.8b00938
Pure and Binary Solvents.	https://www.doi.org/10.1016/j.tca.2008.02.015
Activity coefficients at infinite dilution of organic solutes in the ionic liquid	https://www.doi.org/10.1016/j.jct.2012.08.007
1-butyl-3-methylimidazolium	https://www.doi.org/10.1021/je2011659
Determination of Activity Coefficients at Infinite Dilution Using Gas Liquid Chromatography.	
Measurement and correlation of solubility of Tetracycline hydrochloride in six organic solvents.	
Solubility and Solution Thermodynamics of Betaine in Different Pure Solvents and Binary Mixtures:	

Solubility of Avermectin B1a in Some Pure and Mixed Solvents from (278.2 to 318.2) K:	https://www.doi.org/10.1021/je900772h
Activity Coefficients at Infinite Dilution of Polar Solutes in 1-Butyl-3-methylimidazolium Tetrafluoroborate Using Gas Liquid Chromatography:	https://www.doi.org/10.1021/je060305e
Densities and Viscosities for Binary Mixtures of the Ionic Liquid N-Ethyl Morpholinium Ethylsulfate with (Methanol, Ethanol, 1-Propanol, and 2-Propanol) at Several Temperatures:	https://www.doi.org/10.1021/je201116t
Physicochemical investigations of Complexes in Binary Solvent Mixtures:	https://www.doi.org/10.1007/s10765-010-0861-2
Solubility of Anthracene Dissolved in Alcohol + 2-Pyrrolidinone Binary Solvent Mixtures at 298.15 K:	https://www.doi.org/10.1021/je030175z
Solubility of 1-Methyl-1-butanol in Water at 298.2 K:	https://www.doi.org/10.1016/j.fluid.2011.08.009
Ammonium ionic liquids in extraction of bio-butan-1-ol from water phase using activity coefficients at infinite dilution:	https://www.doi.org/10.1016/j.fluid.2018.09.024
Solubility and Density of Potassium Iodide in a Binary Propan-1-ol-Water Solvent Mixture at (298.15, 303.15, 308.15, and 313.15) K:	https://www.doi.org/10.1021/je9006426
Thermochemical properties of binary mixtures of 1,1,3,3-tetramethylguanidine + 1-alcanol (1-propanol, 1-butanol, 1-pentanol, 1-hexanol, 1-heptanol, 1-octanol, 1-nonanol, and 1-decanol) at different temperatures:	https://www.doi.org/10.1016/j.jct.2016.01.012
Enthalpies of dilution of formamide in aqueous alcohol solutions at 298.15K:	https://www.doi.org/10.1016/j.tca.2007.10.006
Densities and volumetric properties of binary mixtures of 1-hexanol with 1-propanol, 2-propanol, 1-butanol, and 2-butanol at (293.15, 298.15, 303.15, 308.15, 313.15, 318.15, and 323.15) K:	https://www.doi.org/10.1016/j.jct.2007.11.010
Densities and volumetric properties of (N-(2-hydroxyethyl)morpholine + ethanol, + 1-propanol, + 2-propanol, + 1-butanol, and + 2-butanol) at (293.15, 298.15, 303.15, 308.15, 313.15, 318.15, and 323.15) K:	https://www.doi.org/10.1021/acs.jced.8b00257
Cosolvency and Mathematical Modeling Analysis of Chloroxine in Some Binary Green Cosolvent + Water + 1-propanol or acetone + Lactic Acid) at different Temperatures:	https://www.doi.org/10.1016/j.jct.2015.02.013
Apparent Molar Volume and Compressibility of Tetrabutylphosphonium Bromide in Lactate Solutions:	https://www.doi.org/10.1021/je900888e
Extraction desulfurization process of fuels with ionic liquids:	https://www.doi.org/10.1016/j.jct.2014.04.025
Experimental and theoretically study of interaction between organic compounds and triacetonamine based ionic liquids:	https://www.doi.org/10.1016/j.jct.2014.12.027
Quantitative NMR spectroscopy of binary liquid mixtures (aldehyde + alcohol). Part III. (1-decanal or 2-chlorobenzaldehyde) + (methanol or 1-propanol or 2-propanol or 1-butanol or 2-butanol) mixtures:	https://www.doi.org/10.1016/j.jct.2014.01.027
Determination and correlation of solubility of 2-(Methylamino)-1-phenylethanol, Ephedrine, and Pseudoephedrine in Pure Solvents:	https://www.doi.org/10.1016/j.jct.2016.06.032
Determination of the crystallization thermodynamics and kinetics of Ethylparaben in Ethanol + Propanol Binary Solvent Mixtures at Different Temperatures:	https://www.doi.org/10.1016/j.fluid.2011.09.028
Solubility of Butylparaben in Methanol, Ethanol, Propanol, Ethyl Acetate, and Dimethyl Sulfoxide at Various Temperatures and Atmospheric Pressure:	http://pubs.acs.org/doi/suppl/10.1021/ci034243x/suppl_file/ci034243xsi20040112_053635.txt
Liquid Densities and Speed of Sound for Ionic Liquid (2-HEAA and 2-HDEAA) + Alcohol (1-Propanol and 2-Propanol) Mixtures at T = (293.15 to 323.15) K and Atmospheric Pressure:	https://www.doi.org/10.1021/acs.jced.9b00147
Solubility of Butyl Paraben in Methanol, Ethanol, Propanol, Ethyl Acetate, and Dimethyl Sulfoxide at Various Temperatures:	https://www.doi.org/10.1021/je1006289
Measurement and correlation of phase equilibria for ternary systems of water + (ethanol + 1-propanol) + 1-decyl-3-methylimidazolium chloride at 298.15 K:	https://www.doi.org/10.1016/j.fluid.2016.07.024
Measurements of activity coefficients at infinite dilution of aromatic and chloroalkane solutes in the new ionic liquid 1-hexyl-3-methylimidazolium bis(trifluoromethylsulfonyl)imide at T = (298.15 to 328.15) K:	https://www.doi.org/10.1016/j.jct.2008.01.004
Binary Mixtures of Alcohols and 1-Propanol or 2-Propanol or 1-Butanol or 2-Butanol at 298.15 K:	https://www.doi.org/10.1021/acs.jced.6b00954
Solubility and mixing thermodynamic properties of 2-Methoxy-4-(2-propenyl) phenol (Eugenol) in Pure Solvents and Binary Solvent Mixtures of (Ethanol + Water) at Different Temperatures:	https://www.doi.org/10.1016/j.fluid.2018.01.015
Measurement and Correlation of the Solubilities of m-Phthalic Acid in Monobasic Alcohols:	https://www.doi.org/10.1021/je800520w
Activity coefficients at infinite dilution and physicochemical properties for organic solutes and water in the ionic liquid 3-(2-Methoxyethyl)-1-ethylmorpholinium Bis(trifluoromethylsulfonyl)imide:	https://www.doi.org/10.1016/j.jct.2012.05.017
Solubility Determination and Thermodynamic Mixing Properties of Ethyl Paraben in Binary (Water + Isopropyl Alcohol) Mixtures at T = (293.15 to 318.15) K:	https://www.doi.org/10.1021/acs.jced.9b00406
Solubility of p-Aminobenzoic Acid in Potassium Hydroxide Ethanolic Solutions and Thermodynamic Properties of Solution:	https://www.doi.org/10.1021/acs.jced.7b00978
Solubility of L-Serine in Several Pure Solvents from (283.15 to 328.15) K:	https://www.doi.org/10.1016/j.jct.2016.08.021
Solubility of Fisetin in Aqueous Solutions from T = 278.15 to 318.15 K:	https://www.doi.org/10.1016/j.fluid.2008.05.008
Solid−Liquid Equilibrium of 1-hydroxyethane-1,1-diphosphonic acid in Water at Several Temperatures:	https://www.doi.org/10.1016/j.fluid.2012.11.001
Viscosity, density, and speed of sound of methylcyclopentane with primary and secondary alcohols at T = (293.15, 298.15, and 303.15) K:	https://www.doi.org/10.1016/j.jct.2005.12.010
Density, Surface Tension, and Refractive Index of Octane + 1-Alkanol Mixtures at T = 298.15 K:	https://www.doi.org/10.1021/acs.jced.6b01058
Physical properties of the pure 1-methyl-1-propylpyrrolidinium bis(trifluoromethylsulfonyl)imide ionic liquid at T = (293.15 to 343.15) K:	https://www.doi.org/10.1021/je034053i
Study of interaction between organic compounds and bromo or dicationic ionic liquids using gas chromatography:	https://www.doi.org/10.1016/j.fluid.2010.02.007
Solubility of Diphenyl Carbonate in Pure Alcohols from (283 to 333) K:	https://www.doi.org/10.1016/j.jct.2013.08.021
Measurement and Correlation of Solubility of Hydrochlorothiazide in Monosolvents and Binary Solvent Mixtures at 283.15 to 323.15 K:	https://www.doi.org/10.1016/j.fluid.2014.11.020
Solubility and solution thermodynamics of 2-methyl-4-nitroaniline in eleven organic solvents at elevated temperatures:	https://www.doi.org/10.1016/j.jct.2018.08.029
	https://www.doi.org/10.1021/je800569v
	https://www.doi.org/10.1021/acs.jced.9b00220
	https://www.doi.org/10.1016/j.jct.2016.10.037

Solubility Determination and Thermodynamic Modeling of Ibuprofen in Different Solvents and Mixing Properties of Solutions.	https://www.doi.org/10.1021/acs.jced.8b01099
Measurement and correlation solubility and mixing properties of oxalic acid in pure and mixture organic solvents.	https://www.doi.org/10.1016/j.fluid.2013.10.011
Thermodynamic properties of binary mixtures of 2,2,2-Trifluoroethanol with Water or Alkanols at T = 298.15 K.	https://www.doi.org/10.1016/j.tca.2007.04.012
SLE and SLLVE of the water + NH4Cl + 1-propanol system at 101.3 kPa:	https://www.doi.org/10.1016/j.fluid.2018.03.003
Density, Viscosity, and Sound Speed of Bis(trifluoromethylsulfonyl)imide-Based Ionic Liquids + 1-Propanol Mixtures.	https://www.doi.org/10.1021/acs.jced.5b00181
Experimental and Predicted Results of Anomeric Equilibrium of Glucose in Alcohols.	https://www.doi.org/10.1021/je050193b
Measurements of Infinite Dilution Activity Coefficients of Alkanols in the Ionic Liquid Tributylmethylammonium Methylsulfate Using RS-SPME/GC-FID:	https://www.doi.org/10.1021/acs.jced.5b00919
Solid-liquid equilibrium of diphenyl amine/phosphonate in the different organic solvents.	https://www.doi.org/10.1016/j.fluid.2015.03.036
Solubility of Carbamazepine (Form III) in Different Solvents from (275 to 343) K.	https://www.doi.org/10.1021/je8002157
Measurement and calculation of solid liquid equilibrium for ternary systems of 3,4 dichloronitrobenzene + 2,3 dichloronitrobenzene + ethanol/1-propanol	https://www.doi.org/10.1016/j.jct.2016.03.024
Isobaric Vapor-Liquid Equilibria of Ternary Amine and n-Alkane/Alkanol Binary Mixtures.	https://www.doi.org/10.1021/je300568h
Measurement, correlation and dissolution thermodynamics of 2,3-dichloroaniline in 1-propanol + 2-propanol + ethanol + ethyl acetate binary solvent mixtures at different temperatures.	https://www.doi.org/10.1016/j.jct.2014.03.026
Measurements and Modeling with GC-PC-SAFT of Activity Coefficients in Binary Mixtures of 2-diethylethanolamine with 1-propanol and 1-butanol at Different Temperatures.	https://www.doi.org/10.1016/j.jct.2018.06.007
Mass density, sound velocity, mixing enthalpy, 1H NMR, Ab initio calculations and intermolecular interaction in binary mixtures of 2-methoxyethanol with 1-propanol, 2-propanol, 2-butanol and 2-methyl-2-propanol at different temperatures.	https://www.doi.org/10.1016/j.jct.2018.12.019
The dissolution behavior and apparent thermodynamic analysis of 4-methylbenzenesulfonamide in pure and binary solvent mixtures of 1-propanol + 2-propanol + ethanol + methanol at different temperatures.	https://www.doi.org/10.1016/j.jct.2018.11.026
Measurements of activity coefficients at infinite dilution of aliphatic and aromatic hydrocarbons, alcohols, ketones, and water in 1-propanol, 2-propanol, 1-butanol, and 2-butanol by inverse gas chromatography.	https://www.doi.org/10.1016/j.jct.2008.12.018
Ethanol/1-Propanol by Ionic Liquid Entrainer Extractive Distillation.	https://www.doi.org/10.1021/acs.jced.7b00643
Phase equilibria study of the binary systems of 1-butyl-1-methylpyrrolidinium bis(trifluoromethylsulfonyl)imide ionic liquid and (1-propanol, 2-propanol, or 2-methyl-2-propanol).	https://www.doi.org/10.1016/j.jct.2009.03.008
Solubility Determination, Modeling, and Thermodynamic Dissolution Properties of Benzenesulfonamide in 16 Neat Solvents from 273.15 to 324.45 K.	https://www.doi.org/10.1021/acs.jced.9b00360
Vapor-liquid equilibria for the ternary mixture of carbon dioxide + 1-propanol + propyl acetate at elevated pressures:	https://www.doi.org/10.1016/j.fluid.2006.07.016
N-(Phosphonomethyl)iminodiacetic Acid + Different Binary Mixtures:	https://www.doi.org/10.1021/je900764f
Thermodynamics of risperidone and solubility in pure organic solvents:	https://www.doi.org/10.1016/j.fluid.2014.04.028
Apparent molar volumes and compressibilities of tetrabutyl-ammonium bromide in organic solvents.	https://www.doi.org/10.1016/j.jct.2011.01.005
Isobaric Phase Equilibria of Diethyl Carbonate with Five Alcohols at 101.3 kPa.	https://www.doi.org/10.1021/je0201018
(Vapour + liquid) equilibria in the ternary system (acetonitrile + 1-propanol + ethylene glycol) and its binary mixtures at 101.3 kPa.	https://www.doi.org/10.1016/j.jct.2013.08.011
Excess molar enthalpies for binary mixtures of n-propanol + acetic acid and + propyl acetate at 313.15 K and atmospheric pressure:	https://www.doi.org/10.1016/j.fluid.2014.08.022
Thermophysical properties of binary mixtures of 1-Butyl-1-methylpyrrolidinium trifluoromethanesulfonate ionic liquid and 1-propanol at several temperatures.	https://www.doi.org/10.1016/j.jct.2017.12.009
Activity Coefficients at Infinite Dilution in 1-Alkyl-3-methylimidazolium Tetrafluoroborate Ionic Liquids Using Inverse Gas Chromatography.	https://www.doi.org/10.1021/je050544m
Thermodynamic Equilibrium of Azeotropes (Ester + Alcohol) with Different Ionic Liquids at T = 298.15 K:	https://www.doi.org/10.1021/acs.jced.6b00811
Measurement and correlation of density and viscosity of 1-alcohol + water mixtures at	https://www.doi.org/10.1016/j.fluid.2017.09.005
Measurement and Correlation of the Solubility of Coumaric Acid in Nine Solvents at Temperatures from 278.15 to 318.15 K.	https://www.doi.org/10.1021/acs.jced.6b00361
Solubility of 2-(2-Ethoxybenzylidene) Acid in Eleven Organic Solvents at Temperatures from 293.15 to 333.15 K and Thermodynamic Properties of Solution.	https://www.doi.org/10.1021/je700426k
Measurement and correlation of compressibility and conductivity of Sodium Tripolyphosphate in water and in aqueous solutions of 1-propanol.	https://www.doi.org/10.1016/j.fluid.2007.07.066
Thermodynamic Model Correlation and Volumetric and Compressibility behaviour of poly(propylene glycol) - Amino acid aqueous solutions at Different Temperatures.	https://www.doi.org/10.1021/acs.jced.8b00931
Solubility of 2-Chloro-5-nitroaniline in Different Solvents from (263.15 to 293.15) K:	https://www.doi.org/10.1016/j.jct.2015.06.024
THERMODYNAMICS OF MIXTURES CONTAINING AMINES. XIV. CpEm OF BENZYLAMINE WITH HEPTANE AT 293.15K OR WITH METHANOL, 1-PROPANOL OR 1-PENTANOL AT 283.15K AND EXCESS MOLAR VOLUMES AT 293.15K OR WITH METHANOL AT 283.15K.	https://www.doi.org/10.1016/j.tca.2014.03.027
Isothermal vapor-liquid equilibria at 293.15K and excess molar volumes at 298.15K for methanol + ethanol + 1-propanol + 1-pentanol.	https://www.doi.org/10.1016/j.fluid.2008.06.019
Excess molar enthalpies of ethane + heptane plus ethanol and + 1-propanol at 298.15 K.	https://www.doi.org/10.1016/j.jct.2016.12.036
Vapor–Liquid Equilibria of Ethanol + Toluene and Its Binary Subsystems: (1-Propanol with 2-Butanol and Methyl Isobutyl Ketone with 2-Butanol) at 101.3 kPa.	https://www.doi.org/10.1021/je900212d
Correlation with Redlich–Kister, Wilson, NRTL and UNIQUAC equations of thermodynamic properties of binary mixtures containing (1-propanol, 2-propanol, 2-butanol or 2-methyl-2-propanol + 1,2-dichloroethane) at T = 298 K	https://www.doi.org/10.1016/j.jct.2013.08.013
Benzaldehyde + 1,2-Dichloroethane and (1-Propanol + 1,2-Dichloroethane + Benzaldehyde) Ternary Mixtures: Measurement and Correlation of Isobaric Vapor–Liquid Equilibrium Data at 101.3 kPa.	https://www.doi.org/10.1021/je9004646
Solubility and Thermodynamic Properties of 5-Acetamido-2-hydroxy Benzoic Acid (5A2HBA) in Pure and Binary Solvent Mixtures of 1-Propanol + Water at Different Temperatures.	https://www.doi.org/10.1016/j.jct.2015.11.025
Solubility Measurement and Thermodynamic Modeling of Methyl 4-cyano-3-(2-ethoxy-2-oxoethyl)-2-thiophenecarboxylate in nine organic solvents at evaluated temperatures:	

Density and Viscosity of 2-Butanol + (1-Propanol, 2-Propanol, or 2-Amino-1-propanol) Mixtures at Temperatures from (293.15 to 323.15) K:	https://www.doi.org/10.1021/acs.jced.8b01097
Solubilities of Sulfanilamide in Methanol, Ethanol, 1-Propanol, 2-Propanol, 1-Butanol, and 2-Butanol at Temperatures from (278.15 to 323.15) K:	https://www.doi.org/10.1021/je900742p
Co-solvency phenomenon and Application of the ERAS Model (294.15 to 318.15) K:	https://www.doi.org/10.1016/j.jct.2019.06.018
Molar Volumes and Viscosity, Excess Molar Volumes and Viscosity, and thermodynamic properties of ethanol + propan-2-ol + cetrimide in pure and mixed solvents:	https://www.doi.org/10.1021/je049556i
Deviations of Binary Mixtures of Isobaric vapor–liquid equilibria for 2,3,3-trimethylbutane + ethanol + 1-propanol and 1-butanol at 200.00, 300.15, and 500.15 kPa:	https://www.doi.org/10.1016/j.jct.2017.06.014
Extractive distillation of 1-propanol + water with 1-propanol and 1-butanol as solvents:	https://www.doi.org/10.1021/acs.jced.5b00803
Liquid–Liquid Equilibria for Ternary Systems of Water + Methoxycyclopentane (CPME) + Alcohol (Methanol, Ethanol, 1-Propanol, or 2-Propanol) at 298.15 K:	https://www.doi.org/10.1021/acs.jced.5b00122
Modified Method for Measuring the Solubility of Pharmaceutical Compounds in Organic Solvents by Visual Camera:	https://www.doi.org/10.1016/j.fluid.2013.05.038
Application of ionic liquids for determination of activity coefficients at infinite dilution by gas–liquid chromatography:	https://www.doi.org/10.1016/j.fluid.2005.05.019
Fraction and experimental verification of the effect of salt on the vapour–liquid equilibrium of tetrahydrofuran + 1-propanol + water:	https://www.doi.org/10.1016/j.fluid.2019.03.023
Molar excess enthalpies at T = 298.15 K for 1-alkanol + dibutylether) systems:	https://www.doi.org/10.1016/j.jct.2009.07.001
Solubility of Salol in Pure Alcohols from (283.15 to 308.15) K:	https://www.doi.org/10.1021/je800668j
Experimental study of the density and derived volumetric (excess, apparent, and partial molar volumes) properties and measurements and correlation for temperatures from 298K to 362K and Pressure 0.1 to 40MPa:	https://www.doi.org/10.1016/j.jct.2013.11.036
Liquid–liquid equilibrium measurements and correlation for 1-propanol + water + dicyanamide ionic liquid for various temperatures	https://www.doi.org/10.1016/j.jct.2016.11.026
Salt Effect on Vapor–Liquid Equilibria for Binary Systems of 1-Propanol/CaCl2 and Butanol/CaCl2:	https://www.doi.org/10.1016/j.fluid.2005.07.023
Solubility and theoretical study on infinite dilution activity coefficients of various solutes in imidazolium ionic liquids:	https://www.doi.org/10.1016/j.jct.2013.01.005
Solubility and Liquid–Liquid Equilibrium of Aqueous Systems of 1-Propanol at Temperature 298.15 K and Pressure 101.2 kPa:	https://www.doi.org/10.1021/je900059e
Measurement of Diffusion Coefficients in Thermodynamically Nonideal Interactions of volatile organic compounds with the ionic liquid 1-ethyl-3-methylimidazolium tetracyanoborate:	https://www.doi.org/10.1021/je0497303
Measurement, correlation, and prediction of vapor pressure for binary and ternary systems containing an ionic liquid:	https://www.doi.org/10.1016/j.jct.2011.09.028
Experimental study on the calorimetric and volumetric behavior of binary mixtures of 1-propanol with alkanols (C1–C4) and correlation with Wilson, ERAS, and Modified ERAS models:	https://www.doi.org/10.1016/j.fluid.2008.02.005
Measurement and correlation of the solubility of 2,3,4,6-tetrabromothiophene in different solvents:	https://www.doi.org/10.1016/j.fluid.2015.03.043
Measurements of activity coefficients at infinite dilution of organic solutes in 1-propyl-3-methyl imidazolium based ionic liquid by inverse gas chromatography:	https://www.doi.org/10.1016/j.jct.2012.02.037
Application of ionic liquid [1-hexyl-3-methylimidazolium][tris(pentafluoroethyl)trifluorophosphate] to the removal of alcohol from mixtures with water:	https://www.doi.org/10.1021/je800145h
Vapor–Liquid Equilibria for Tetrahydropyran and Alcohol Systems:	https://www.doi.org/10.1016/j.jct.2012.06.005
Measurement and correlation of the solubility of 2,3,4,6-tetrabromothiophene in different solvents:	https://www.doi.org/10.1016/j.fluid.2008.09.005
Measurements of activity coefficients at infinite dilution of organic solutes in 1-propyl-3-methyl imidazolium based ionic liquid by inverse gas chromatography:	https://www.doi.org/10.1016/j.jct.2010.10.026
Application of 1-propanol + benzene and its binary mixtures with 1-hexyl-3-methylimidazolium tetrafluoroborate ionic liquid at (303.15 to 313.15) K:	https://www.doi.org/10.1016/j.fluid.2017.04.005
Solubilities, Densities, and Ultrasonic Velocity Studies of Binary Mixtures of Chloroform with Propan-1-ol and Butan-1-ol at (303.15 and 313.15) K:	https://www.doi.org/10.1021/je050169y
Determination of the Solubility, Dissolution Enthalpy, and Entropy of 1-(3-Amino-2,2-dimethylpropyl)-5-(trifluoromethyl)-2(1H)-pyridinone in Different Pure Solvents:	https://www.doi.org/10.1021/acs.jced.5b00758
Determination and correlation of solubility and solution thermodynamics of 1-(2-Bromo-phenyl)-pyrrole-2,5-dione in 14 Pure Solvents from 278.15 to 323.15 K:	https://www.doi.org/10.1016/j.fluid.2012.06.012
Solubilities of Lauric Acid in n-Hexane, 1-Propanol, 2-Propanol, 1-Bromopropane, and 1-Chloropropane at Pressure:	https://www.doi.org/10.1016/j.fluid.2016.08.019
Liquid–liquid equilibria of hydrophilic alcohol + sodium hydroxide + water + 1-chloropropane systems at 298.15 K:	https://www.doi.org/10.1016/j.fluid.2010.11.026
Vapor–Liquid Equilibrium Behavior of Tolan in Alcohol:	https://www.doi.org/10.1016/j.tca.2013.06.008
Molar Volumes and Refractive Indexes of Hexane-1,2,3,4,5,6-hexol in Aqueous Solutions of 1-Propanol and 2-Propanol:	https://www.doi.org/10.1021/je049778g
Densities and Viscosities of Four Binary Diethyl Carbonate + 1-Alcohol Systems from (288.15 to 313.15) K:	https://www.doi.org/10.1021/je900966r
Densities and Viscosities of Naphthalen-1-ol, Naphthalen-2-ol, and 1-Chloronaphthalene in the Mixtures of Ethanol, Propan-1-ol, and Butan-1-ol:	https://www.doi.org/10.1021/je1013476
Solubility in Binary Solvent Mixtures: Pyrene Dissolved in Alcohol + Acetonitrile Mixtures at 299.2 K:	https://www.doi.org/10.1021/je201222n
Vapor–Liquid Equilibrium of 1-Propanol + 1-Butene, + cis-2-Butene, + 2-Methyl-propene, + trans-2-Butene, + n-Butane, and + 2-Methyl-propane:	https://www.doi.org/10.1021/je020226c
	https://www.doi.org/10.1021/je049959i

Densities and Viscosities for Binary Liquid Mixtures of n-Undecane + 1-Propanol, + 1-Butanol, + 1-Pentanol, + 1-Hexanol, + 1-Octanol, + 1-Decanol, and + 1-Dodecanol at T = (288.15 to 313.15) K at 0.1 MPa
https://www.doi.org/10.1021/acs.jced.6b00121

Determination of activity coefficients at infinite dilution of organic solutes in trihexyl(tetradecyl) phosphonium tris(pentafluoroethyl) trifluorophosphate by gas liquid chromatography
https://www.doi.org/10.1016/j.fluid.2005.04.021

Thermodynamic Properties of Mixtures Containing Ionic Liquids. Activity Coefficients at Infinite Dilution of Organic Compounds in 1-Propyl Boronic Acid-3-Alkylimidazolium Bromide and 1-Propenyl-3-alkylimidazolium Bromide Using Inverse Gas Chromatography:
https://www.doi.org/10.1021/je060033f

Legend

af:	Acentric Factor
affp:	Proton affinity
aigt:	Autoignition Temperature
basg:	Gas basicity
chl:	Standard liquid enthalpy of combustion
cpg:	Ideal gas heat capacity
cpl:	Liquid phase heat capacity
cps:	Solid phase heat capacity
dm:	Dipole Moment
dvisc:	Dynamic viscosity
fll:	Lower Flammability Limit
flu:	Upper Flammability Limit
fpc:	Flash Point (Closed Cup Method)
fpo:	Flash Point (Open Cup Method)
gf:	Standard Gibbs free energy of formation
gyrad:	Radius of Gyration
hf:	Enthalpy of formation at standard conditions
hfl:	Liquid phase enthalpy of formation at standard conditions
hfus:	Enthalpy of fusion at standard conditions
hfust:	Enthalpy of fusion at a given temperature
hvap:	Enthalpy of vaporization at standard conditions
hvapt:	Enthalpy of vaporization at a given temperature
ie:	Ionization energy
log10ws:	Log10 of Water solubility in mol/l
logp:	Octanol/Water partition coefficient
mcvol:	McGowan's characteristic volume
nfpaf:	NFPA Fire Rating
nfpah:	NFPA Health Rating
pc:	Critical Pressure
pvap:	Vapor pressure
rfi:	Refractive Index
rhol:	Liquid Density
rinpol:	Non-polar retention indices
ripol:	Polar retention indices
sfust:	Entropy of fusion at a given temperature

sg:	Molar entropy at standard conditions
sl:	Liquid phase molar entropy at standard conditions
speedsl:	Speed of sound in fluid
srf:	Surface Tension
ss:	Solid phase molar entropy at standard conditions
tb:	Normal Boiling Point Temperature
tbp:	Boiling point at given pressure
tc:	Critical Temperature
tf:	Normal melting (fusion) point
tt:	Triple Point Temperature
vc:	Critical Volume
volm:	Molar Volume
zc:	Critical Compressibility
zra:	Rackett Parameter