

# 1,2-Ethanediamine, N,N-diethyl-N'-methyl-

<b>Other names:</b>	2-(Diethylamino)-N-methylethylamine Ethylenediamine, N,N-diethyl-N'-methyl- N,N-Diethyl-N'-methylenediamine N,N-Diethyl-N'-methylethylenediamine [2-(Diethylamino)ethyl]methylamine diethyl(2-methylaminoethyl)amine
<b>Inchi:</b>	InChI=1S/C7H18N2/c1-4-9(5-2)7-6-8-3/h8H,4-7H2,1-3H3
<b>InchiKey:</b>	MKDYQLJYEBWUIG-UHFFFAOYSA-N
<b>Formula:</b>	C7H18N2
<b>SMILES:</b>	CCN(CC)CCNC
<b>Mol. weight [g/mol]:</b>	130.23
<b>CAS:</b>	104-79-0

## Physical Properties

Property code	Value	Unit	Source
gf	208.23	kJ/mol	Joback Method
hf	-66.81	kJ/mol	Joback Method
hfus	22.01	kJ/mol	Joback Method
hvap	39.66	kJ/mol	Joback Method
log10ws	-0.51		Crippen Method
logp	0.548		Crippen Method
mcvol	129.450	ml/mol	McGowan Method
pc	2832.35	kPa	Joback Method
tb	431.70	K	NIST Webbook
tc	589.99	K	Joback Method
tf	253.78	K	Joback Method
vc	0.480	m3/kmol	Joback Method

## Temperature Dependent Properties

Property code	Value	Unit	Temperature [K]	Source
cpg	268.14	J/molxK	422.17	Joback Method
cpg	281.72	J/molxK	450.14	Joback Method
cpg	294.74	J/molxK	478.11	Joback Method

cpg	307.22	J/mol×K	506.08	Joback Method
cpg	319.17	J/mol×K	534.05	Joback Method
cpg	330.62	J/mol×K	562.02	Joback Method
cpg	341.57	J/mol×K	589.99	Joback Method

## Correlations

Information	Value
Property code	pvap
Equation	$\ln(P_{vp}) = A + B/(T + C)$
Coeff. A	1.56618e+01
Coeff. B	-4.09217e+03
Coeff. C	-6.11490e+01
Temperature range (K), min.	302.05
Temperature range (K), max.	456.52

## Sources

<b>Joback Method:</b>	<a href="https://en.wikipedia.org/wiki/Joback_method">https://en.wikipedia.org/wiki/Joback_method</a>
<b>McGowan Method:</b>	<a href="http://link.springer.com/article/10.1007/BF02311772">http://link.springer.com/article/10.1007/BF02311772</a>
<b>NIST Webbook:</b>	<a href="http://webbook.nist.gov/cgi/cbook.cgi?ID=C104790&amp;Units=SI">http://webbook.nist.gov/cgi/cbook.cgi?ID=C104790&amp;Units=SI</a>
<b>The Yaws Handbook of Vapor Pressure:</b>	<a href="https://www.sciencedirect.com/book/9780128029992/the-yaws-handbook-of-vapor-pressure">https://www.sciencedirect.com/book/9780128029992/the-yaws-handbook-of-vapor-pressure</a>
<b>Crippen Method:</b>	<a href="http://pubs.acs.org/doi/abs/10.1021/ci9903071">http://pubs.acs.org/doi/abs/10.1021/ci9903071</a>
<b>Crippen Method:</b>	<a href="https://www.chemeo.com/doc/models/crippen_log10ws">https://www.chemeo.com/doc/models/crippen_log10ws</a>

## Legend

<b>cpg:</b>	Ideal gas heat capacity
<b>gf:</b>	Standard Gibbs free energy of formation
<b>hf:</b>	Enthalpy of formation at standard conditions
<b>hfus:</b>	Enthalpy of fusion at standard conditions
<b>hvap:</b>	Enthalpy of vaporization at standard conditions
<b>log10ws:</b>	Log10 of Water solubility in mol/l
<b>logp:</b>	Octanol/Water partition coefficient
<b>mcvol:</b>	McGowan's characteristic volume

<b>pc:</b>	Critical Pressure
<b>pvap:</b>	Vapor pressure
<b>tb:</b>	Normal Boiling Point Temperature
<b>tc:</b>	Critical Temperature
<b>tf:</b>	Normal melting (fusion) point
<b>vc:</b>	Critical Volume

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