

# Diethylmalonic acid, di(5-methoxy-3-methylpentyl) ester

Inchi:	InChI=1S/C21H40O6/c1-7-21(8-2,19(22)26-15-11-17(3)9-13-24-5)20(23)27-16-12-18(4)1
InchiKey:	SPSYZNSKQBSKDA-UHFFFAOYSA-N
Formula:	C21H40O6
SMILES:	CCC(CC)(C(=O)OCCC(C)CCOC)C(=O)OCCC(C)CCOC
Mol. weight [g/mol]:	388.54

## Physical Properties

Property code	Value	Unit	Source
gf	-553.94	kJ/mol	Joback Method
hf	-1250.12	kJ/mol	Joback Method
hfus	43.64	kJ/mol	Joback Method
hvap	83.40	kJ/mol	Joback Method
log10ws	-3.79		Crippen Method
logp	4.005		Crippen Method
mcvol	333.370	ml/mol	McGowan Method
pc	1013.59	kPa	Joback Method
rinpol	2277.00		NIST Webbook
tb	873.19	K	Joback Method
tc	1069.93	K	Joback Method
tf	487.63	K	Joback Method
vc	1.272	m3/kmol	Joback Method

## Temperature Dependent Properties

Property code	Value	Unit	Temperature [K]	Source
cpg	1095.07	J/molxK	873.19	Joback Method
cpg	1113.06	J/molxK	905.98	Joback Method
cpg	1129.72	J/molxK	938.77	Joback Method
cpg	1145.06	J/molxK	971.56	Joback Method
cpg	1159.11	J/molxK	1004.35	Joback Method
cpg	1171.87	J/molxK	1037.14	Joback Method
cpg	1183.36	J/molxK	1069.93	Joback Method
dvisc	0.0004285	Paxs	487.63	Joback Method
dvisc	0.0001820	Paxs	551.89	Joback Method

dvisc	0.0000925	Paxs	616.15	Joback Method
dvisc	0.0000534	Paxs	680.41	Joback Method
dvisc	0.0000339	Paxs	744.67	Joback Method
dvisc	0.0000231	Paxs	808.93	Joback Method
dvisc	0.0000167	Paxs	873.19	Joback Method

## Sources

<b>NIST Webbook:</b>	<a href="http://webbook.nist.gov/cgi/cbook.cgi?ID=U370773&amp;Units=SI">http://webbook.nist.gov/cgi/cbook.cgi?ID=U370773&amp;Units=SI</a>
<b>Crippen Method:</b>	<a href="http://pubs.acs.org/doi/abs/10.1021/ci9903071">http://pubs.acs.org/doi/abs/10.1021/ci9903071</a>
<b>Crippen Method:</b>	<a href="https://www.chemeo.com/doc/models/crippen_log10ws">https://www.chemeo.com/doc/models/crippen_log10ws</a>
<b>Joback Method:</b>	<a href="https://en.wikipedia.org/wiki/Joback_method">https://en.wikipedia.org/wiki/Joback_method</a>
<b>McGowan Method:</b>	<a href="http://link.springer.com/article/10.1007/BF02311772">http://link.springer.com/article/10.1007/BF02311772</a>

## Legend

<b>cpg:</b>	Ideal gas heat capacity
<b>dvisc:</b>	Dynamic viscosity
<b>gf:</b>	Standard Gibbs free energy of formation
<b>hf:</b>	Enthalpy of formation at standard conditions
<b>hfus:</b>	Enthalpy of fusion at standard conditions
<b>hvap:</b>	Enthalpy of vaporization at standard conditions
<b>log10ws:</b>	Log10 of Water solubility in mol/l
<b>logp:</b>	Octanol/Water partition coefficient
<b>mcvol:</b>	McGowan's characteristic volume
<b>pc:</b>	Critical Pressure
<b>rinpol:</b>	Non-polar retention indices
<b>tb:</b>	Normal Boiling Point Temperature
<b>tc:</b>	Critical Temperature
<b>tf:</b>	Normal melting (fusion) point
<b>vc:</b>	Critical Volume

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