

# Hydrazine, 1,1-dibutyl-

<b>Other names:</b>	1,1-DBH 1,1-Di-n-butylhydrazine 1,1-Dibutylhydrazine N,N-Dibutylhydrazine
<b>Inchi:</b>	InChI=1S/C8H20N2/c1-3-5-7-10(9)8-6-4-2/h3-9H2,1-2H3
<b>InchiKey:</b>	DJTIANXYHUNHQV-UHFFFAOYSA-N
<b>Formula:</b>	C8H20N2
<b>SMILES:</b>	CCCCN(N)CCCC
<b>Mol. weight [g/mol]:</b>	144.26
<b>CAS:</b>	7422-80-2

## Physical Properties

Property code	Value	Unit	Source
gf	193.71	kJ/mol	Joback Method
hf	-107.13	kJ/mol	Joback Method
hfus	24.69	kJ/mol	Joback Method
hvap	46.09	kJ/mol	Joback Method
ie	7.47 ± 0.05	eV	NIST Webbook
ie	7.75 ± 0.05	eV	NIST Webbook
log10ws	-2.17		Crippen Method
logp	1.762		Crippen Method
mcvol	143.540	ml/mol	McGowan Method
pc	2681.86	kPa	Joback Method
tb	467.41	K	Joback Method
tc	642.94	K	Joback Method
tf	295.65	K	Joback Method
vc	0.530	m <sup>3</sup> /kmol	Joback Method

## Temperature Dependent Properties

Property code	Value	Unit	Temperature [K]	Source
cpg	324.85	J/mol×K	467.41	Joback Method
cpg	339.33	J/mol×K	496.67	Joback Method
cpg	353.17	J/mol×K	525.92	Joback Method

cpg	366.40	J/mol×K	555.18	Joback Method
cpg	379.04	J/mol×K	584.43	Joback Method
cpg	391.11	J/mol×K	613.69	Joback Method
cpg	402.62	J/mol×K	642.94	Joback Method

## Correlations

Information	Value
Property code	pvap
Equation	$\ln(P_{vp}) = A + B/(T + C)$
Coeff. A	1.62135e+01
Coeff. B	-4.22869e+03
Coeff. C	-6.08030e+01
Temperature range (K), min.	326.33
Temperature range (K), max.	448.68

## Sources

<b>NIST Webbook:</b>	<a href="http://webbook.nist.gov/cgi/cbook.cgi?ID=C7422802&amp;Units=SI">http://webbook.nist.gov/cgi/cbook.cgi?ID=C7422802&amp;Units=SI</a>
<b>The Yaws Handbook of Vapor Pressure:</b>	<a href="https://www.sciencedirect.com/book/9780128029992/the-yaws-handbook-of-vapor-pressure">https://www.sciencedirect.com/book/9780128029992/the-yaws-handbook-of-vapor-pressure</a>
<b>Crippen Method:</b>	<a href="http://pubs.acs.org/doi/abs/10.1021/ci9903071">http://pubs.acs.org/doi/abs/10.1021/ci9903071</a>
<b>Crippen Method:</b>	<a href="https://www.chemeo.com/doc/models/crippen_log10ws">https://www.chemeo.com/doc/models/crippen_log10ws</a>
<b>Joback Method:</b>	<a href="https://en.wikipedia.org/wiki/Joback_method">https://en.wikipedia.org/wiki/Joback_method</a>
<b>McGowan Method:</b>	<a href="http://link.springer.com/article/10.1007/BF02311772">http://link.springer.com/article/10.1007/BF02311772</a>

## Legend

<b>cpg:</b>	Ideal gas heat capacity
<b>gf:</b>	Standard Gibbs free energy of formation
<b>hf:</b>	Enthalpy of formation at standard conditions
<b>hfus:</b>	Enthalpy of fusion at standard conditions
<b>hvap:</b>	Enthalpy of vaporization at standard conditions
<b>ie:</b>	Ionization energy
<b>log10ws:</b>	Log10 of Water solubility in mol/l
<b>logp:</b>	Octanol/Water partition coefficient

<b>mcvol:</b>	McGowan's characteristic volume
<b>pc:</b>	Critical Pressure
<b>pvap:</b>	Vapor pressure
<b>tb:</b>	Normal Boiling Point Temperature
<b>tc:</b>	Critical Temperature
<b>tf:</b>	Normal melting (fusion) point
<b>vc:</b>	Critical Volume

Latest version available from:

<https://www.chemeo.com/cid/17-566-5/Hydrazine-1-1-dibutyl.pdf>

Generated by Cheméo on 2024-04-18 03:15:08.983386346 +0000 UTC m=+15699357.903963657.

Cheméo (<https://www.chemeo.com>) is the biggest free database of chemical and physical data for the process industry.