

# Diosgenin

## Other names:

(20R,25R)-spirost-5-en-3«beta»-ol  
(20R,25R)-spirost-5-en-3Â«betaÂ»-ol  
(25R)-Spirost-5-en-3beta-ol  
(25R)-Spirost-5-en-3«beta»-ol  
(25R)-Spirost-5-en-3Â«betaÂ»-ol  
(3.beta.,25R)-spirost-5-en-3-ol  
22«alpha»-Spirost-5-en-3«beta»-ol  
22Â«alphaÂ»-Spirost-5-en-3Â«betaÂ»-ol  
25D-Spirost-5-en-3beta-ol  
25D-spirost-5-en-3«beta»-ol  
25D-spirost-5-en-3Â«betaÂ»-ol  
NSC 33396  
Nitogenin  
SP 37  
Spiro(8H-naphth(2',1':4,5)-indeno(2,1-b)furan-8,2'-(2H)-pyran)-2-ol,  
1,2,3,3',4,4',4a,4b,5,5',6,6',6a,6b,7,9a,10,10a,10b,11-eicosahydro-4a,5',6a,7-tetramethyl-  
Spiro[8H-naphth[2',1':4,5]indeno[2,1-b]furan-8,2-[2H]pyran], spirost-5-en-3-ol  
deriv  
Spirost-5-en-3-ol, (3beta,25R)-  
Spirost-5-en-3-ol, (3«beta»,25R)-  
Spirost-5-en-3-ol, (3Â«betaÂ»,25R)-  
Spirost-5-en-3beta-ol, (25R)-  
Spirost-5-en-3«beta»-ol, (25R)-  
Spirost-5-en-3Â«betaÂ»-ol, (25R)-  
**Inchi:** InChI=1S/C27H42O3/c1-16-7-12-27(29-15-16)17(2)24-23(30-27)14-22-20-6-5-18-13-19(20-21)  
**InchiKey:** WQLVFSAGQJTQCK-ICPPQKBESA-N  
**Formula:** C27H42O3  
**SMILES:** CC1CCC2(OC1)OC1CC3C4CC=C5CC(O)CCC5(C)C4CCC3(C)C1C2C  
**Mol. weight [g/mol]:** 414.62  
**CAS:** 512-04-9

## Physical Properties

Property code	Value	Unit	Source
gf	112.51	kJ/mol	Joback Method
hf	-632.83	kJ/mol	Joback Method
hfus	47.14	kJ/mol	Joback Method
hvap	98.03	kJ/mol	Joback Method
log10ws	-5.37		Aqueous Solubility Prediction Method

log10ws	-2.62		Aqueous and cosolvent solubility data for drug-like organic compounds
log10ws	-7.32		Estimated Solubility Method
logp	5.714		Crippen Method
mvol	339.440	ml/mol	McGowan Method
pc	1284.67	kPa	Joback Method
rinpol	3220.00		NIST Webbook
rinpol	3220.00		NIST Webbook
tb	1015.08	K	Joback Method
tc	1258.05	K	Joback Method
tf	478.65	K	Aqueous Solubility Prediction Method
tf	480.24	K	Application of the NRTL method to correlate solubility of diosgenin
tf	474.40	K	Experimental solubility of diosgenin and estriol in various solvents between T = (293.2-328.2) K
vc	1.270	m <sup>3</sup> /kmol	Joback Method

## Temperature Dependent Properties

Property code	Value	Unit	Temperature [K]	Source
cpg	1396.38	J/mol×K	1015.08	Joback Method
cpg	1437.02	J/mol×K	1055.58	Joback Method
cpg	1480.32	J/mol×K	1096.07	Joback Method
cpg	1526.88	J/mol×K	1136.57	Joback Method
cpg	1577.23	J/mol×K	1177.06	Joback Method
cpg	1631.97	J/mol×K	1217.56	Joback Method
cpg	1691.64	J/mol×K	1258.05	Joback Method

## Sources

Solubility of diosgenin in different solvents:  
 Solubility of diosgenin in several imidazolium-based ionic liquids:  
 Determination and correlation of the solubility for diosgenin in alcohols  
 Crippen Method:

<https://www.doi.org/10.1016/j.jct.2011.11.009>

<https://www.doi.org/10.1021/je5004324>

<https://www.doi.org/10.1016/j.jct.2012.02.002>

<http://pubs.acs.org/doi/abs/10.1021/ci9903071>

McGowan Method:

<http://link.springer.com/article/10.1007/BF02311772>

<b>Application of the NRTL method to correlate solubility of diosgenin:</b>	<a href="https://www.doi.org/10.1016/j.jct.2013.12.020">https://www.doi.org/10.1016/j.jct.2013.12.020</a>
<b>Experimental solubility for betulin and estrone in various solvents within the temperature range T = (293.2 to 328.2) K:</b>	<a href="https://www.doi.org/10.1016/j.jct.2016.02.006">https://www.doi.org/10.1016/j.jct.2016.02.006</a>
<b>Joback Method:</b>	<a href="https://en.wikipedia.org/wiki/Joback_method">https://en.wikipedia.org/wiki/Joback_method</a>
<b>Estimated Solubility Method:</b>	<a href="http://pubs.acs.org/doi/suppl/10.1021/ci034243x/suppl_file/ci034243xsi20040112_053635.txt">http://pubs.acs.org/doi/suppl/10.1021/ci034243x/suppl_file/ci034243xsi20040112_053635.txt</a>
<b>Experimental solubility of diosgenin and estriol in various solvents between 15 (296.15 to 328.2) K:</b>	<a href="https://www.doi.org/10.1016/j.jct.2016.11.017">https://www.doi.org/10.1016/j.jct.2016.11.017</a>
<b>NIST Webbook:</b>	<a href="http://webbook.nist.gov/cgi/cbook.cgi?ID=C512049&amp;Units=SI">http://webbook.nist.gov/cgi/cbook.cgi?ID=C512049&amp;Units=SI</a>
<b>Aqueous and cosolvent solubility data for drug-like organic compounds:</b>	<a href="https://www.ncbi.nlm.nih.gov/pmc/articles/PMC2751500/">https://www.ncbi.nlm.nih.gov/pmc/articles/PMC2751500/</a>
<b>Aqueous Solubility Prediction Method:</b>	<a href="http://onschallenge.wikispaces.com/file/view/AqueousDataset002.xlsx/351826032/AqueousDa">http://onschallenge.wikispaces.com/file/view/AqueousDataset002.xlsx/351826032/AqueousDa</a>

## Legend

<b>cp<sub>g</sub>:</b>	Ideal gas heat capacity
<b>g<sub>f</sub>:</b>	Standard Gibbs free energy of formation
<b>h<sub>f</sub>:</b>	Enthalpy of formation at standard conditions
<b>h<sub>fus</sub>:</b>	Enthalpy of fusion at standard conditions
<b>h<sub>vap</sub>:</b>	Enthalpy of vaporization at standard conditions
<b>log<sub>10</sub>w<sub>s</sub>:</b>	Log <sub>10</sub> of Water solubility in mol/l
<b>log<sub>p</sub>:</b>	Octanol/Water partition coefficient
<b>mc<sub>vol</sub>:</b>	McGowan's characteristic volume
<b>pc:</b>	Critical Pressure
<b>rin<sub>pol</sub>:</b>	Non-polar retention indices
<b>tb:</b>	Normal Boiling Point Temperature
<b>tc:</b>	Critical Temperature
<b>tf:</b>	Normal melting (fusion) point
<b>vc:</b>	Critical Volume

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