lithium bromide

Inchi: InChl=1S/BrH.Li/h1H;/q;+1/p-1

InchiKey: AMXOYNBUYSYVKV-UHFFFAOYSA-M

Formula: BrLi
SMILES: [Li]Br
Mol. weight [g/mol]: 86.84

CAS: 7550-35-8

Physical Properties

Property code	Value	Unit	Source
affp	819.00	kJ/mol	NIST Webbook
basg	792.50	kJ/mol	NIST Webbook
ea	0.66 ± 0.04	eV	NIST Webbook
ie	8.70	eV	NIST Webbook
ie	9.40	eV	NIST Webbook
ie	9.43 ± 0.05	eV	NIST Webbook
ie	10.00	eV	NIST Webbook
ie	9.30	eV	NIST Webbook
tf	824.00	К	An anion effect on the separation of Agl-containing melts using sound waves

Correlations

Information	Value	
Property code	pvap	
Equation	ln(Pvp) = A + B/(T + C)	
Coeff. A	1.54385e+01	
Coeff. B	-1.59969e+04	
Coeff. C	-1.04720e+02	
Temperature range (K), min.	1021.15	
Temperature range (K), max.	1583.15	

Sources

Thermophysical Properties of Ternary https://www.doi.org/10.1021/acs.jced.8b00742 Systems Potassium Formate + Phopywe dayebpolyee Vapowater: Pressure: https://www.sciencedirect.com/book/9780128029992/the-yaws-handbook-of-vapor-pressure An anion effect on the separation of https://www.doi.org/10.1016/j.jct.2015.03.022 Agl-containing melts using sound Waysis and Chemistry of Lithium Raysigs and Chemistry of Lithium Halides in 1,3-Dioxolane and Its Binary โดงในเมื่อง in 1,3-Dioxolane https://www.doi.org/10.1016/j.tca.2012.08.009 https://www.doi.org/10.1016/j.fluid.2017.12.034 https://www.doi.org/10.1021/acs.jced.7b00951 https://www.doi.org/10.1021/je1009202 https://www.doi.org/10.1021/je900709n https://www.doi.org/10.1016/j.fluid.2005.07.002 https://www.doi.org/10.1016/j.jct.2005.04.007 https://www.doi.org/10.1016/j.jct.2003.09.005 aame elestrolutga in storlene + sodium salusilities es evalutlities Salts in 1-Butyl-3-methylimidazolium https://www.doi.org/10.1021/acs.jced.8b00618 DRINGERITHMENT AND A STREET OF THE STREET OF https://www.doi.org/10.1016/j.jct.2012.05.026 http://webbook.nist.gov/cgi/cbook.cgi?ID=C7550358&Units=SI The experimental study on the https://www.doi.org/10.1016/j.fluid.2018.11.036 MARAE Processure Manney on the first some the sure of the control influence of crown ethers and glycols https://www.doi.org/10.1021/acs.jced.8b01217 thermodynamic properties for the adjactive graph in the properties of the pro solutions by experiment andmolecular https://www.doi.org/10.1016/j.fluid.2016.08.036 https://www.doi.org/10.1021/acs.jced.5b00329 https://www.doi.org/10.1021/je900656c https://www.doi.org/10.1021/acs.jced.8b01078 https://www.doi.org/10.1016/j.tca.2009.01.008 https://www.doi.org/10.1016/j.fluid.2006.02.012 https://www.doi.org/10.1016/j.jct.2004.09.015 Thermodynamic properties of (LiCl + Thermodynamic properties of (LICI+N,N-dimethylacetamide) and (LiBr+N,N-dimethylacetamide) and (LiBr+N,N-dimethylacetamide) are alkali tental-halides indelimethylasticasides)+K: acomonariasy tast decourses and viscosities of the (Water + Lithium + Belightleiguis of the (Water + Lithium + Belightleiguis of the Water + Lithium https://www.doi.org/10.1016/j.jct.2009.03.005 https://www.doi.org/10.1021/je010312x https://www.doi.org/10.1021/acs.jced.9b00517

Density of Methanolic Alkali Halide Salt https://www.doi.org/10.1021/je5009944 Solutions by Experiment and Molecular Samparature Dependence of the Density of Aqueous Alkali Halide Salt Samparature Dependence of the Density of Aqueous Alkali Halide Salt Santimental Experimental Molecular syndaling of solubility of LiBr and LiNO3 in methanol, ethanol, 1-propanol, 2-propanol and 1-butanol:

Legend

affp: Proton affinity Gas basicity basg: Electron affinity ea: Ionization energy ie: Vapor pressure pvap:

tf: Normal melting (fusion) point

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