

# Ethane, 1-bromo-2-fluoro-

<b>Other names:</b>	1-Bromo-2-fluoroethane 1-Fluoro-2-bromoethane CH <sub>2</sub> FCH <sub>2</sub> Br
<b>Inchi:</b>	InChI=1S/C2H4BrF/c3-1-2-4/h1-2H2
<b>InchiKey:</b>	JTLAIKFGRHDNQM-UHFFFAOYSA-N
<b>Formula:</b>	C <sub>2</sub> H <sub>4</sub> BrF
<b>SMILES:</b>	FCCBr
<b>Mol. weight [g/mol]:</b>	126.96
<b>CAS:</b>	762-49-2

## Physical Properties

Property code	Value	Unit	Source
gf	-214.53	kJ/mol	Joback Method
hf	-254.39	kJ/mol	Joback Method
hfus	9.30	kJ/mol	Joback Method
hvap	25.66	kJ/mol	Joback Method
ie	10.57	eV	NIST Webbook
log10ws	-0.94		Crippen Method
logp	1.351		Crippen Method
mcvol	58.310	ml/mol	McGowan Method
pc	5205.63	kPa	Joback Method
tb	344.50 ± 0.50	K	NIST Webbook
tb	345.00	K	NIST Webbook
tc	485.18	K	Joback Method
tf	172.69	K	Joback Method
vc	0.228	m <sup>3</sup> /kmol	Joback Method

## Temperature Dependent Properties

Property code	Value	Unit	Temperature [K]	Source
cpg	75.97	J/mol×K	310.59	Joback Method
cpg	80.28	J/mol×K	339.69	Joback Method
cpg	84.38	J/mol×K	368.79	Joback Method
cpg	88.29	J/mol×K	397.89	Joback Method

cpg	92.02	J/mol×K	426.98	Joback Method
cpg	95.56	J/mol×K	456.08	Joback Method
cpg	98.94	J/mol×K	485.18	Joback Method

## Correlations

Information	Value
Property code	pvap
Equation	$\ln(P_{vp}) = A + B/(T + C)$
Coeff. A	1.50012e+01
Coeff. B	-3.57845e+03
Temperature range (K), min.	243.21
Temperature range (K), max.	369.30

## Sources

<b>McGowan Method:</b>	<a href="http://link.springer.com/article/10.1007/BF02311772">http://link.springer.com/article/10.1007/BF02311772</a>
<b>NIST Webbook:</b>	<a href="http://webbook.nist.gov/cgi/cbook.cgi?ID=C762492&amp;Units=SI">http://webbook.nist.gov/cgi/cbook.cgi?ID=C762492&amp;Units=SI</a>
<b>The Yaws Handbook of Vapor Pressure:</b>	<a href="https://www.sciencedirect.com/book/9780128029992/the-yaws-handbook-of-vapor-pressure">https://www.sciencedirect.com/book/9780128029992/the-yaws-handbook-of-vapor-pressure</a>
<b>Crippen Method:</b>	<a href="http://pubs.acs.org/doi/abs/10.1021/ci9903071">http://pubs.acs.org/doi/abs/10.1021/ci9903071</a>
<b>Crippen Method:</b>	<a href="https://www.chemeo.com/doc/models/crippen_log10ws">https://www.chemeo.com/doc/models/crippen_log10ws</a>
<b>Joback Method:</b>	<a href="https://en.wikipedia.org/wiki/Joback_method">https://en.wikipedia.org/wiki/Joback_method</a>

## Legend

<b>cpg:</b>	Ideal gas heat capacity
<b>gf:</b>	Standard Gibbs free energy of formation
<b>hf:</b>	Enthalpy of formation at standard conditions
<b>hfus:</b>	Enthalpy of fusion at standard conditions
<b>hvap:</b>	Enthalpy of vaporization at standard conditions
<b>ie:</b>	Ionization energy
<b>log10ws:</b>	Log10 of Water solubility in mol/l
<b>logp:</b>	Octanol/Water partition coefficient
<b>mcvol:</b>	McGowan's characteristic volume
<b>pc:</b>	Critical Pressure

<b>pvap:</b>	Vapor pressure
<b>tb:</b>	Normal Boiling Point Temperature
<b>tc:</b>	Critical Temperature
<b>tf:</b>	Normal melting (fusion) point
<b>vc:</b>	Critical Volume

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