

Sucrose

Other names:

(+)-Sucrose
(2R,3R,4S,5S,6R)-2-(((2S,3S,4S,5R)-3,4-dihydroxy-2,5-bis(hydroxymethyl)tetrahydrofuran-2-ylidene)-4,6-dihydroxy-2H-pyran-2-ylidene)-D-glucopyranose
(«alpha»-D-Glucosido)-«beta»-D-fructofuranoside
(Â«alphaÂ»-D-Glucosido)-Â«betaÂ»-D-fructofuranoside
.alpha.-trehalose
.beta.-D-fructofuranosyl .alpha.-D-glucopyranoside
4-O-.beta.-D-galactopyranosyl-D-glucose
Amerfond
Beet sugar
Cane sugar
Confectioner's sugar
D-(+)-Saccharose
D-(+)-Sucrose
D-(+)-lactose
D-Sucrose
D-trehalose
Fructofuranoside, «alpha»-D-glucopyranosyl, «beta»-D-fructofuranosyl
Fructofuranoside, Â«alphaÂ»-D-glucopyranosyl, Â«betaÂ»-D-fructofuranosyl
Glucopyranoside, «beta»-D-fructofuranosyl, «alpha»-D-glucopyranoside
Glucopyranoside, Â«betaÂ»-D-fructofuranosyl, Â«alphaÂ»-D-glucopyranoside
Granulated sugar
Microse
NCI-C56597
NSC 406942
Rock candy
Saccharose
Saccharum
Sugar
Table sugar
White sugar
alpha,alpha-trehalose
lactose
«alpha»-D-Glucopyranoside, «beta»-D-fructofuranosyl
«alpha»-D-Glucopyranosyl «beta»-D-fructofuranoside
«beta»-D-Fructofuranoside, «alpha»-D-glucopyranosyl
«beta»-D-Fructofuranosyl «alpha»-D-glucopyranoside
Â«alphaÂ»-D-Glucopyranoside, Â«betaÂ»-D-fructofuranosyl
Â«alphaÂ»-D-Glucopyranosyl Â«betaÂ»-D-fructofuranoside
Â«betaÂ»-D-Fructofuranoside, Â«alphaÂ»-D-glucopyranosyl
Â«betaÂ»-D-Fructofuranosyl Â«alphaÂ»-D-glucopyranoside

Inchi: InChI=1S/C12H22O11/c13-1-4-6(16)8(18)9(19)11(21-4)23-12(3-15)10(20)7(17)5(2-14)22-12
InchiKey: CZMRCDWAGMRECN-SFOFJGFUSA-N
Formula: C12H22O11
SMILES: OCC1OC(OC2(CO)OC(CO)C(O)C2O)C(O)C(O)C1O
Mol. weight [g/mol]: 342.30
CAS: 57-50-1

Physical Properties

Property code	Value	Unit	Source
chs	-5643.40 ± 1.80	kJ/mol	NIST Webbook
chs	-5644.17	kJ/mol	NIST Webbook
chs	-5637.40 ± 1.70	kJ/mol	NIST Webbook
chs	-5664.38 ± 0.69	kJ/mol	NIST Webbook
gf	-1320.10	kJ/mol	Joback Method
hf	-1917.41	kJ/mol	Joback Method
hfs	-2221.20	kJ/mol	NIST Webbook
hfus	63.65	kJ/mol	Joback Method
hvap	184.54	kJ/mol	Joback Method
log10ws	0.79		Aqueous Solubility Prediction Method
log10ws	0.79		Estimated Solubility Method
logp	-5.396		Crippen Method
mcvol	222.790	ml/mol	McGowan Method
pc	4627.70	kPa	Joback Method
ss	392.40	J/molxK	NIST Webbook
ss	360.20	J/molxK	NIST Webbook
tb	1290.10	K	Joback Method
tc	1782.75	K	Joback Method
tf	462.00 ± 3.00	K	NIST Webbook
tf	461.00 ± 6.00	K	NIST Webbook
tf	464.05	K	Artificial neural networks as a supporting tool for compatibility study based on thermogravimetric data
tf	424.40	K	Heat capacity and transition behavior of sucrose by standard, fast scanning and temperature-modulated calorimetry
tf	458.65	K	Aqueous Solubility Prediction Method
vc	0.784	m3/kmol	Joback Method

Temperature Dependent Properties

Property code	Value	Unit	Temperature [K]	Source
cpg	1178.26	J/molxK	1782.75	Joback Method
cpg	970.71	J/molxK	1290.10	Joback Method
cpg	998.85	J/molxK	1372.21	Joback Method
cpg	1028.30	J/molxK	1454.32	Joback Method
cpg	1060.01	J/molxK	1536.42	Joback Method
cpg	1094.94	J/molxK	1618.53	Joback Method
cpg	1134.04	J/molxK	1700.64	Joback Method
cps	422.50	J/molxK	297.00	NIST Webbook
cps	408.50	J/molxK	288.15	Temperature dependence of the heat capacities in the solid state of 18 mono-, di-, and poly-saccharides
cps	416.60	J/molxK	293.15	Temperature dependence of the heat capacities in the solid state of 18 mono-, di-, and poly-saccharides
cps	424.30	J/molxK	298.15	Temperature dependence of the heat capacities in the solid state of 18 mono-, di-, and poly-saccharides
cps	429.40	J/molxK	303.15	Temperature dependence of the heat capacities in the solid state of 18 mono-, di-, and poly-saccharides
cps	437.50	J/molxK	308.15	Temperature dependence of the heat capacities in the solid state of 18 mono-, di-, and poly-saccharides

cps	445.50	J/molxK	313.15	Temperature dependence of the heat capacities in the solid state of 18 mono-, di-, and poly-saccharides
cps	451.00	J/molxK	318.15	Temperature dependence of the heat capacities in the solid state of 18 mono-, di-, and poly-saccharides
cps	466.20	J/molxK	323.15	Temperature dependence of the heat capacities in the solid state of 18 mono-, di-, and poly-saccharides
cps	472.60	J/molxK	328.15	Temperature dependence of the heat capacities in the solid state of 18 mono-, di-, and poly-saccharides
cps	482.50	J/molxK	333.15	Temperature dependence of the heat capacities in the solid state of 18 mono-, di-, and poly-saccharides
cps	490.30	J/molxK	338.15	Temperature dependence of the heat capacities in the solid state of 18 mono-, di-, and poly-saccharides
cps	498.80	J/molxK	343.15	Temperature dependence of the heat capacities in the solid state of 18 mono-, di-, and poly-saccharides
cps	506.70	J/molxK	348.15	Temperature dependence of the heat capacities in the solid state of 18 mono-, di-, and poly-saccharides
cps	513.90	J/molxK	353.15	Temperature dependence of the heat capacities in the solid state of 18 mono-, di-, and poly-saccharides

cps	522.00	J/molxK	358.15	Temperature dependence of the heat capacities in the solid state of 18 mono-, di-, and poly-saccharides
cps	424.30	J/molxK	298.15	NIST Webbook
cps	430.00	J/molxK	300.00	NIST Webbook
cps	425.50	J/molxK	298.15	NIST Webbook
hfust	46.20	kJ/mol	459.00	NIST Webbook

Sources

Volumetric Properties for Ionic Liquid Sucrose Water Systems:

Effect of Citrate Salts on the Volumetric and Ultrasonic Properties of Sucrose in Aqueous Solutions at Temperatures = 288.15-318.15 K

Partial molar volumes and viscosity B-coefficients of arginine in aqueous solution at 288.15 K

Nonclassical Hydrogen Bonding in the Sucrose + Mannitol + Sorbitol System: Spectroscopic and Thermodynamic Studies in Binary and Ternary Subsystems at 298.15 K and Ambient Pressure

Polysaccharide-Based Ionic Liquid Salts: Coefficients for Six Sugars at 0.1 MPa

Interactions in Glycylglycine (2 to 65.2) aqueous glucose / 1 M aqueous

Volume and Heat Capacity Properties of MgSO4/CuSO4 in Sucrose + Water

Volume Properties of L-alanine, and L-valine in aqueous sucrose solutions at 298.15 K

Properties of Sugars and Polysaccharide Water-Ethanol

Heat Capacity and transition behavior of sucrose by standard, fast scanning

Study of solute-solute and solute-solvent interactions of L-methionine in aqueous sucrose solutions at different

Apparent molar volumes and apparent molar heat capacities of aqueous

Polysaccharides, D-glucose, and D-fructose in aqueous sucrose solutions at 298.15 K

Volume and Heat Capacity Properties of Sugars and Polysaccharide Water-Ethanol

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[illegible]

https://en.wikipedia.org/wiki/Joback_method

<https://www.doi.org/10.1016/j.ijct.2013.06.016>

Effect of ionic liquid,
1-hexyl-3-methylimidazolium bromide
on the solubility of sucrose in water
Estimated Solubility Method
viscometric behavior of aqueous
sucrose solutions at different
temperatures
Ammonium Phosphate Salts at
Different Temperatures through
Density and Speed of Sound
Measurements

<https://www.doi.org/10.1016/j.jct.2015.09.021>

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<https://www.doi.org/10.1021/acs.jced.8b01157>

Legend

chs:	Standard solid enthalpy of combustion
cpg:	Ideal gas heat capacity
cps:	Solid phase heat capacity
gf:	Standard Gibbs free energy of formation
hf:	Enthalpy of formation at standard conditions
hfs:	Solid phase enthalpy of formation at standard conditions
hfus:	Enthalpy of fusion at standard conditions
hfust:	Enthalpy of fusion at a given temperature
hvap:	Enthalpy of vaporization at standard conditions
log10ws:	Log10 of Water solubility in mol/l
logp:	Octanol/Water partition coefficient
mcvol:	McGowan's characteristic volume
pc:	Critical Pressure
ss:	Solid phase molar entropy at standard conditions
tb:	Normal Boiling Point Temperature
tc:	Critical Temperature
tf:	Normal melting (fusion) point
vc:	Critical Volume

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