

# Sucrose

**Other names:**

(+)-Sucrose

(2R,3R,4S,5S,6R)-2-(((2S,3S,4S,5R)-3,4-dihydroxy-2,5-bis(hydroxymethyl)tetrahydrofura

(«alpha»-D-Glucosido)-«beta»-D-fructofuranoside

(Î«alphaÎ»-D-Glucosido)-Î«betaÎ»-D-fructofuranoside

.alpha.-trehalose

.beta.-D-fructofuranosyl .alpha.-D-glucopyranoside

4-O-.beta.-D-galactopyranosyl-D-glucose

Amerfond

## Beet sugar

## Cane sugar

Confectioner's sugar

## D-(+)-Saccharose

D-(+)-Sucrose

D-(+)-lactose

## D-Sucrose

D-trehalose

Fructofuranoside, «alpha»-D-glucopyranosyl, «beta»-D

Fructofuranoside,  $\alpha$ -D-glucopyranosyl,  $\beta$ -D-

Glucopyranoside, «beta»-D-fructofuranosyl, «alpha»-D

Glucopyranoside, Â«betaÂ»-D-fructofuranosyl, Â«alphaÂ»-D-

Granulated sugar

## Microse

NCI-C56597

NSC 406942

## Rock candy

## Saccharose

## Saccharum

## Sugar

## Table sugar

White sugar

alpha,alpha-trehalose

lactose

«alpha»-D-Glucopyranoside, «beta»-D-fructofuranosyl

«alpha»-D-Glucopyranosyl «beta»-D-fructofuranoside

«beta»-D-Fructofuranoside, «alpha»-D-glucopyranosyl

«beta»-D-Fructofuranosyl «alpha»-D-glucopyranoside

Â«alphaÂ»-D-Glucopyranoside, Â«betaÂ»-D-fructofuranosyl

 $\alpha$ -D-Glucopyranosyl  $\beta$ -D-fructofuranoside

Â«betaÂ»-D-Fructofuranoside, Â«alphaÂ»-D-glucopyranosyl

Â«betaÂ»-D-Fructofuranosyl Â«alphaÂ»-D-glucopyranoside

Inchi:

InchiKey:

Formula:

SMILES:

Mol. weight [g/mol]:

CAS:

InChI=1S/C12H22O11/c13-1-4-6(16)8(18)9(19)11(21-4)23-12(3-15)10(20)7(17)5(2-14)22

CZMRCDWAGMRECN-SFOFJGFUSA-N

C12H22O11

OCC1OC(OC2(CO)OC(CO)C(O)C2O)C(O)C(O)C1O

342.30

57-50-1

Physical Properties

Property code	Value	Unit	Source
chs	-5664.38 ± 0.69	kJ/mol	NIST Webbook
chs	-5637.40 ± 1.70	kJ/mol	NIST Webbook
chs	-5644.17	kJ/mol	NIST Webbook
chs	-5643.40 ± 1.80	kJ/mol	NIST Webbook
gf	-1320.10	kJ/mol	Joback Method
hf	-1917.41	kJ/mol	Joback Method
hfs	-2221.20	kJ/mol	NIST Webbook
hfus	63.65	kJ/mol	Joback Method
hvap	184.54	kJ/mol	Joback Method
log10ws	0.79		Aqueous Solubility Prediction Method
log10ws	0.79		Estimated Solubility Method
logp	-5.396		Crippen Method
mcvol	222.790	ml/mol	McGowan Method
pc	4627.70	kPa	Joback Method
ss	392.40	J/molxK	NIST Webbook
ss	360.20	J/molxK	NIST Webbook
tb	1290.10	K	Joback Method
tc	1782.75	K	Joback Method
tf	464.05	K	Artificial neural networks as a supporting tool for compatibility study based on thermogravimetric data
tf	461.00 ± 6.00	K	NIST Webbook
tf	462.00 ± 3.00	K	NIST Webbook
tf	424.40	K	Heat capacity and transition behavior of sucrose by standard, fast scanning and temperature-modulated calorimetry
tf	458.65	K	Aqueous Solubility Prediction Method
vc	0.784	m3/kmol	Joback Method

# Temperature Dependent Properties

Property code	Value	Unit	Temperature [K]	Source
cpg	970.71	J/molxK	1290.10	Joback Method
cpg	998.85	J/molxK	1372.21	Joback Method
cpg	1028.30	J/molxK	1454.32	Joback Method
cpg	1060.01	J/molxK	1536.42	Joback Method
cpg	1094.94	J/molxK	1618.53	Joback Method
cpg	1134.04	J/molxK	1700.64	Joback Method
cpg	1178.26	J/molxK	1782.75	Joback Method
cps	451.00	J/molxK	318.15	Temperature dependence of the heat capacities in the solid state of 18 mono-, di-, and poly-saccharides
cps	416.60	J/molxK	293.15	Temperature dependence of the heat capacities in the solid state of 18 mono-, di-, and poly-saccharides
cps	424.30	J/molxK	298.15	Temperature dependence of the heat capacities in the solid state of 18 mono-, di-, and poly-saccharides
cps	429.40	J/molxK	303.15	Temperature dependence of the heat capacities in the solid state of 18 mono-, di-, and poly-saccharides
cps	437.50	J/molxK	308.15	Temperature dependence of the heat capacities in the solid state of 18 mono-, di-, and poly-saccharides
cps	445.50	J/molxK	313.15	Temperature dependence of the heat capacities in the solid state of 18 mono-, di-, and poly-saccharides

cps	408.50	J/mol×K	288.15	Temperature dependence of the heat capacities in the solid state of 18 mono-, di-, and poly-saccharides
cps	466.20	J/mol×K	323.15	Temperature dependence of the heat capacities in the solid state of 18 mono-, di-, and poly-saccharides
cps	472.60	J/mol×K	328.15	Temperature dependence of the heat capacities in the solid state of 18 mono-, di-, and poly-saccharides
cps	482.50	J/mol×K	333.15	Temperature dependence of the heat capacities in the solid state of 18 mono-, di-, and poly-saccharides
cps	490.30	J/mol×K	338.15	Temperature dependence of the heat capacities in the solid state of 18 mono-, di-, and poly-saccharides
cps	498.80	J/mol×K	343.15	Temperature dependence of the heat capacities in the solid state of 18 mono-, di-, and poly-saccharides
cps	506.70	J/mol×K	348.15	Temperature dependence of the heat capacities in the solid state of 18 mono-, di-, and poly-saccharides
cps	513.90	J/mol×K	353.15	Temperature dependence of the heat capacities in the solid state of 18 mono-, di-, and poly-saccharides
cps	522.00	J/mol×K	358.15	Temperature dependence of the heat capacities in the solid state of 18 mono-, di-, and poly-saccharides



Viscosities of Some Saccharides in Aqueous Solutions of Temperature Dependence of the Heat Capacities in the Solid State of 18 Monosaccharides and Polysaccharides: of some saccharides in aqueous solutions. *Journal of Chemical Thermodynamics* 15: 115-120 (1983) [15] K: Molar Volumes and Partial Molar Volumes of Simple and Complex Sugars in Methanol and Aqueous Solutions. Studies on Saccharide-Dipotassium Tetraborate (Borax) Interactions in Aqueous Solutions of Sucrose, Ascorbic Acid, and NaF in (Glucose + Water) and (Sucrose + Water) Mixtures at 298.15 K. Equilibrium of Cellobiose, Sucrose, and Maltose Monohydrate in Measurments and Determination of Water Activity in Aqueous Solutions. Chemical study of solute-solute and solute-solvent interactions of drug molecules in aqueous protic liquids and in aqueous binary mixtures of the solubility of different temperatures of solutions of sucrose by standard, fast scanning calorimetric interactions of N-methylformamide in aqueous glucose and sucrose concentrations and compressibilities of glyglyglycine in aqueous solutions and in the ternary systems containing sucrose in aqueous solutions of formamide. Temperature interactions in bromide, sodium, and potassium glycolate in aqueous solutions and polymer-polymer systems. Systems from 0-1 coefficients of arginine in aqueous glucose, sucrose, and L-ascorbic acid + D-mannose + methylglycine + sodium chloride in aqueous solutions of sucrose and apparent solubility capacities of aqueous liquid binary mixtures of formamide and sucrose at temperatures of 278.15 K to 303.15 K and at the pressure 0.35 MPa. Fluoroborate + Sucrose/Maltose + Water.

Artificial neural networks as a supporting tool for compatibility study. Acoustic and volumetric properties of betaine hydrochloride drug in aqueous solutions of D-mannose. Influence of NaBr on Solvation Behavior of Polyhydroxy Solutes in Aqueous Solutions. Apparent Viscosity and Apparent Viscosity Behavior of Carbohydrates in Aqueous Solutions. Bromide and Chloride Salts in Aqueous Solutions. Properties of Sucrose in Aqueous Solutions at Temperatures of 278.15 to 303.15 K. Enthalpy and Entropy of Solvation of Sucrose in Aqueous Solutions. Class of Aqueous Binary Mixtures of Polyhydroxy Compounds. Diffusion Coefficients of Sucrose in Aqueous Solutions of Sodium Chloride and Sodium Chloride Solutions at 298.15 to 303.15 K. Aqueous two-phase systems based on methanol and water. Acoustic and volumetric studies of aqueous solutions of sucrose and glucose. Diffusion coefficients of sucrose in aqueous solutions of water, methanol, and ethanol. Diffusion coefficients of arginine in aqueous solutions. Volumetric and Transport Behavior of Different Carbohydrates in Aqueous Solutions. Diffusion Coefficients of Sucrose in Aqueous Solutions of Water + Protic Ionic Binary Solutions. Diffusion Coefficients of Aqueous Solutions of a-Cyclodextrin at 298.15 to 312.15 K: McGowan Method.

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# Legend

<b>chs:</b>	Standard solid enthalpy of combustion
<b>cpg:</b>	Ideal gas heat capacity
<b>cps:</b>	Solid phase heat capacity
<b>gf:</b>	Standard Gibbs free energy of formation
<b>hf:</b>	Enthalpy of formation at standard conditions
<b>hfs:</b>	Solid phase enthalpy of formation at standard conditions
<b>hfus:</b>	Enthalpy of fusion at standard conditions
<b>hfust:</b>	Enthalpy of fusion at a given temperature
<b>hvap:</b>	Enthalpy of vaporization at standard conditions
<b>log10ws:</b>	Log10 of Water solubility in mol/l
<b>logp:</b>	Octanol/Water partition coefficient
<b>mcvol:</b>	McGowan's characteristic volume
<b>pc:</b>	Critical Pressure
<b>ss:</b>	Solid phase molar entropy at standard conditions
<b>tb:</b>	Normal Boiling Point Temperature
<b>tc:</b>	Critical Temperature
<b>tf:</b>	Normal melting (fusion) point
<b>vc:</b>	Critical Volume

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