Glutaric acid, 2-methyl-4-chlorophenyl tetradecyl ester

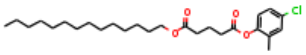
InChI: InChI=1S/C26H41ClO4/c1-3-4-5-6-7-8-9-10-11-12-13-14-20-30-25(28)16-15-17-26(29)31-24-19-18-23(27)22-21-24/h18-19,21H,3-17,20H2,1-2H3

InChI Key: GFEKUYGJNVPYEF-UHFFFAOYSA-N

Formula: C26H41ClO4

SMILES: CCCCCCCCCCCCCCOC(=O)CCCC(=O)Oc1ccc(Cl)cc1C

Molecular Weight: 453.05



Physical Properties

Property	Value	Unit	Source
$\Delta_f G°$	-218.58	kJ/mol	Joback Method
$\Delta_f H°_{gas}$	-871.72	kJ/mol	Joback Method
$\Delta_{fus} H°$	66.13	kJ/mol	Joback Method
$\Delta_{vap} H°$	99.77	kJ/mol	Joback Method
$logP_{oct/wat}$	7.97		Crippen Method
P_c	886.83	kPa	Joback Method
T_{boil}	1020.93	K	Joback Method
T_c	1251.86	K	Joback Method
T_{fus}	608.48	K	Joback Method
V_c	1.48	m^3/kg-mol	Joback Method

Temperature Dependent Properties

Property	Value	Unit	Temperature (K)	Source
$C_{p,gas}$	1268.95	J/mol×K	1020.93	Joback Method
η	0.00	Pa×s	1020.93	Joback Method

Sources

Joback Method: https://en.wikipedia.org/wiki/Joback_method

NIST Webbook: http://webbook.nist.gov/cgi/inchi/InChI=1S/C26H41ClO4/c1-3-4-5-6-7-8-9-10-11-12-13-14-20-30-25(28)16-15-17-26(29)31-24-19-18-23(27)21-22(24)2/h18-19,21H,3-17,20H2,1-2H3

Crippen Method: http://pubs.acs.org/doi/abs/10.1021/ci990307l

Legend

$\mathbf{C_{p,gas}}$: Ideal gas heat capacity (J/mol×K).
η: Dynamic viscosity (Pa×s).
$\Delta_f\mathbf{G°}$: Standard Gibbs free energy of formation (kJ/mol).
$\Delta_f\mathbf{H°_{gas}}$: Enthalpy of formation at standard conditions (kJ/mol).
$\Delta_{fus}\mathbf{H°}$: Enthalpy of fusion at standard conditions (kJ/mol).
$\Delta_{vap}\mathbf{H°}$: Enthalpy of vaporization at standard conditions (kJ/mol).
$\boldsymbol{log}\mathbf{P_{oct/wat}}$: Octanol/Water partition coefficient .
$\mathbf{P_c}$: Critical Pressure (kPa).
$\mathbf{T_{boil}}$: Normal Boiling Point Temperature (K).
$\mathbf{T_c}$: Critical Temperature (K).
$\mathbf{T_{fus}}$: Normal melting (fusion) point (K).
$\mathbf{V_c}$: Critical Volume (m^3/kg-mol).