Monoethanolamine

Other names:	1-Amino-2-hydroxyethane
	2-Amino-1-Ethanol
	2-Aminoaethanol
	2-Aminoetanolo
	2-Aminoethan-1-ol
	2-Aminoethanol
	2-Aminoethyl alcohol
	2-Ethanolamine
	2-Hydroxyethanamine
	2-Hydroxyethylamine
	Aethanolamin
	Aminoethanol
	Colamine
	Etanolamina
	Ethanol, 2-amino-
	Ethanolamine
	Ethylolamine
	Glycinol
	Kolamin
	MEA
	Monoaethanolamin
	NH2CH2CH2OH
	Olamine
	Thiofaco M-50
	UN 2491
	USAF EK-1597
	«beta»-Aminoethanol
	«beta»-Aminoethyl alcohol
	«beta»-Ethanolamine
	«beta»-Hydroxyethylamine
	Â«betaÂ»-Aminoethanol
	Â«betaÂ»-Aminoethyl alcohol
	Â«betaÂ»-Ethanolamine
	Â«betaÂ»-Hydroxyethylamine
Inchi:	InChI=1S/C2H7NO/c3-1-2-4/h4H,1-3H2
InchiKey:	HZAXFHJVJLSVMW-UHFFFAOYSA-N
Formula:	C2H7NO
SMILES:	NCCO
Mol. weight [g/mol]:	61.08
CAS:	141-43-5

Physical Properties

Property code	Value	Unit	Source
affp	930.30	kJ/mol	NIST Webbook
basg	896.80	kJ/mol	NIST Webbook
gf	-104.41	kJ/mol	Joback Method
hf	-203.05	kJ/mol	Joback Method
hfl	-507.50	kJ/mol	NIST Webbook
hfus	10.22	kJ/mol	Joback Method
hvap	47.37	kJ/mol	Joback Method
ie	8.90	eV	NIST Webbook
ie	8.96	eV	NIST Webbook
ie	8.96	eV	NIST Webbook
ie	8.96	eV	NIST Webbook
ie	9.88	eV	NIST Webbook
ie	9.87 ± 0.06	eV	NIST Webbook
ie	9.88	eV	NIST Webbook
log10ws	0.64		Crippen Method
logp	-1.063		Crippen Method
mcvol	54.890	ml/mol	McGowan Method
pc	8030.00 ± 40.00	kPa	NIST Webbook
rinpol	643.00		NIST Webbook
rinpol	680.00		NIST Webbook
rinpol	643.00		NIST Webbook
rinpol	698.00		NIST Webbook
ripol	1427.00		NIST Webbook
ripol	1470.00		NIST Webbook
ripol	1450.00		NIST Webbook
ripol	1402.00		NIST Webbook
ripol	1413.00		NIST Webbook
ripol	1427.00		NIST Webbook
ripol	1402.00		NIST Webbook
ripol	1411.00		NIST Webbook
tb	443.20	K	NIST Webbook
tb	444.15 ± 1.50	K	NIST Webbook
tb	443.55	K	NIST Webbook
tb	444.00 ± 0.50	K	NIST Webbook
tb	444.00	K	NIST Webbook
tc	670.00 ± 50.00	K	NIST Webbook
tc	671.40 ± 1.50	K	NIST Webbook

tf	283.46 ± 0.05	K	NIST Webbook
tf	283.45	K	NIST Webbook
tf	283.70 ± 0.60	K	NIST Webbook
vc	0.196	m3/kmol	Joback Method

Temperature Dependent Properties

Property code	Value	Unit	Temperature [K]	Source
cpg	107.21	J/mol×K	409.87	Joback Method
cpg	112.35	J/mol×K	439.58	Joback Method
cpg	117.29	J/mol×K	469.30	Joback Method
cpg	122.02	J/mol×K	499.01	Joback Method
cpg	126.57	J/mol×K	528.72	Joback Method
cpg	130.93	J/mol×K	558.43	Joback Method
cpg	135.10	J/mol×K	588.15	Joback Method
cpl	182.00	J/mol×K	353.15	Molar Heat Capacity of Various Aqueous Alkanolamine Solutions from 303.15 K to 353.15 K
cpl	179.00	J/mol×K	343.15	Molar Heat Capacity of Various Aqueous Alkanolamine Solutions from 303.15 K to 353.15 K
cpl	180.00	J/mol×K	348.15	Molar Heat Capacity of Various Aqueous Alkanolamine Solutions from 303.15 K to 353.15 K
cpl	178.00	J/mol×K	338.15	Molar Heat Capacity of Various Aqueous Alkanolamine Solutions from 303.15 K to 353.15 K
cpl	163.70	J/mol×K	303.15	Density, Viscosity, Heat Capacity, Surface Tension, and Solubility of CO2 in Aqueous Solutions of Potassium Serinate

cpl	167.98	J/mol×K	313.15	Density, Viscosity, Heat Capacity, Surface Tension, and Solubility of CO2 in Aqueous Solutions of Potassium Serinate
cpl	172.25	J/mol×K	323.15	Density, Viscosity, Heat Capacity, Surface Tension, and Solubility of CO2 in Aqueous Solutions of Potassium Serinate
cpl	175.00	J/mol×K	333.15	Molar Heat Capacity of Various Aqueous Alkanolamine Solutions from 303.15 K to 353.15 K
cpl	174.00	J/mol×K	328.15	Molar Heat Capacity of Various Aqueous Alkanolamine Solutions from 303.15 K to 353.15 K
cpl	173.00	J/mol×K	323.15	Molar Heat Capacity of Various Aqueous Alkanolamine Solutions from 303.15 K to 353.15 K
cpl	171.00	J/mol×K	318.15	Molar Heat Capacity of Various Aqueous Alkanolamine Solutions from 303.15 K to 353.15 K
cpl	169.00	J/mol×K	313.15	Molar Heat Capacity of Various Aqueous Alkanolamine Solutions from 303.15 K to 353.15 K
cpl	169.00	J/mol×K	308.15	Molar Heat Capacity of Various Aqueous Alkanolamine Solutions from 303.15 K to 353.15 K

cpl	168.00	J/mol×K	303.15	Molar Heat Capacity of Various Aqueous Alkanolamine Solutions from 303.15 K to 353.15 K
dvisc	0.0096100	Pa×s	313.15	Density and Viscosity of Monoethanolamine + Water + Carbon Dioxide from (25 to 80) deg C
dvisc	0.0013760	Pa×s	393.15	Viscometric and volumetric behaviour of binary mixtures of sulfolane and N-methylpyrrolidone with monoethanolamine and diethanolamine in the range 303 373 K
dvisc	0.0015880	Pa×s	383.15	Viscometric and volumetric behaviour of binary mixtures of sulfolane and N-methylpyrrolidone with monoethanolamine and diethanolamine in the range 303 373 K
dvisc	0.0067200	Pa×s	323.15	Density and Viscosity of Monoethanolamine + Water + Carbon Dioxide from (25 to 80) deg C
dvisc	0.0028500	Pa×s	353.15	Density and Viscosity of Monoethanolamine + Water + Carbon Dioxide from (25 to 80) deg C
dvisc	0.0144230	Pa×s	303.15	Viscometric and volumetric behaviour of binary mixtures of sulfolane and N-methylpyrrolidone with monoethanolamine and diethanolamine in the range 303 373 K

dvisc	0.0095620	Pa×s	313.15	Viscometric and volumetric behaviour of binary mixtures of sulfolane and N-methylpyrrolidone with monoethanolamine and diethanolamine in the range 303 373 K
dvisc	0.0066650	Pa×s	323.15	Viscometric and volumetric behaviour of binary mixtures of sulfolane and N-methylpyrrolidone with monoethanolamine and diethanolamine in the range 303 373 K
dvisc	0.0048140	Pa×s	333.15	Viscometric and volumetric behaviour of binary mixtures of sulfolane and N-methylpyrrolidone with monoethanolamine and diethanolamine in the range 303 373 K
dvisc	0.0035510	Pa×s	343.15	Viscometric and volumetric behaviour of binary mixtures of sulfolane and N-methylpyrrolidone with monoethanolamine and diethanolamine in the range 303 373 K
dvisc	0.0036900	Pa×s	343.15	Density and Viscosity of Monoethanolamine + Water + Carbon Dioxide from (25 to 80) deg C

dvisc	0.0027640	Pa×s	353.15	Viscometric and volumetric behaviour of binary mixtures of sulfolane and N-methylpyrrolidone with monoethanolamine and diethanolamine in the range 303 373 K
dvisc	0.0023370	Pa×s	363.15	Viscometric and volumetric behaviour of binary mixtures of sulfolane and N-methylpyrrolidone with monoethanolamine and diethanolamine in the range 303 373 K
dvisc	0.0019010	Pa×s	373.15	Viscometric and volumetric behaviour of binary mixtures of sulfolane and N-methylpyrrolidone with monoethanolamine and diethanolamine in the range 303 373 K
dvisc	0.0199000	Pa×s	298.15	Density and Viscosity of Monoethanolamine + Water + Carbon Dioxide from (25 to 80) deg C
hvapt	55.90	kJ/mol	396.00	NIST Webbook
hvapt	61.70	kJ/mol	377.00	NIST Webbook
hvapt	54.70	kJ/mol	411.00	NIST Webbook
hvapt	58.90	kJ/mol	390.50	NIST Webbook
pvap	51.97	kPa	423.15	Vapor Liquid Equilibrium for Several Compounds Relevant to the Biofuels Industry Modeled with the Wilson Equation
pvap	2.98	kPa	357.46	Ebulliometric Determination of Vapor
pvap	3.98	kPa	362.99	Ebulliometric Determination of Vapor

pvap	5.48	kPa	369.33	Ebulliometric Determination of Vapor
pvap	6.98	kPa	374.29	Ebulliometric Determination of Vapor
pvap	8.47	kPa	378.41	Ebulliometric Determination of Vapor
pvap	10.97	kPa	384.07	Ebulliometric Determination of Vapor
pvap	14.96	kPa	391.14	Ebulliometric Determination of Vapor
pvap	19.97	kPa	398.01	Ebulliometric Determination of Vapor
pvap	6.64	kPa	373.15	Vapor Liquid Equilibrium for Several Compounds Relevant to the Biofuels Industry Modeled with the Wilson Equation
pvap	29.97	kPa	408.19	Ebulliometric Determination of Vapor
pvap	39.98	kPa	415.78	Ebulliometric Determination of Vapor
pvap	39.95	kPa	415.81	Ebulliometric Determination of Vapor
pvap	54.95	kPa	424.61	Ebulliometric Determination of Vapor
pvap	64.95	kPa	429.44	Ebulliometric Determination of Vapor
pvap	69.95	kPa	431.61	Ebulliometric Determination of Vapor
pvap	79.95	kPa	435.61	Ebulliometric Determination of Vapor
pvap	0.02	kPa	283.34	Isothermal Vapor-Liquid Equilibria of (Monoethanolamine + Water) and (4-Methylmorpholine + Water) Binary Systems at Several Temperatures

pvap	0.03	kPa	288.34	Isothermal Vapor-Liquid Equilibria of (Monoethanolamine + Water) and (4-Methylmorpholine + Water) Binary Systems at Several Temperatures
pvap	0.04	kPa	293.34	Isothermal Vapor-Liquid Equilibria of (Monoethanolamine + Water) and (4-Methylmorpholine + Water) Binary Systems at Several Temperatures
pvap	0.06	kPa	298.34	Isothermal Vapor-Liquid Equilibria of (Monoethanolamine + Water) and (4-Methylmorpholine + Water) Binary Systems at Several Temperatures
pvap	0.09	kPa	303.34	Isothermal Vapor-Liquid Equilibria of (Monoethanolamine + Water) and (4-Methylmorpholine + Water) Binary Systems at Several Temperatures
pvap	0.13	kPa	308.34	Isothermal Vapor-Liquid Equilibria of (Monoethanolamine + Water) and (4-Methylmorpholine + Water) Binary Systems at Several Temperatures
pvap	0.19	kPa	313.34	Isothermal Vapor-Liquid Equilibria of (Monoethanolamine + Water) and (4-Methylmorpholine + Water) Binary Systems at Several Temperatures

pvap	0.27	kPa	318.34	Isothermal Vapor-Liquid Equilibria of (Monoethanolamine + Water) and (4-Methylmorpholine + Water) Binary Systems at Several Temperatures
pvap	0.38	kPa	323.34	Isothermal Vapor-Liquid Equilibria of (Monoethanolamine + Water) and (4-Methylmorpholine + Water) Binary Systems at Several Temperatures
pvap	0.53	kPa	328.34	Isothermal Vapor-Liquid Equilibria of (Monoethanolamine + Water) and (4-Methylmorpholine + Water) Binary Systems at Several Temperatures
pvap	0.74	kPa	333.34	Isothermal Vapor-Liquid Equilibria of (Monoethanolamine + Water) and (4-Methylmorpholine + Water) Binary Systems at Several Temperatures
pvap	1.01	kPa	338.34	Isothermal Vapor-Liquid Equilibria of (Monoethanolamine + Water) and (4-Methylmorpholine + Water) Binary Systems at Several Temperatures
pvap	1.36	kPa	343.34	Isothermal Vapor-Liquid Equilibria of (Monoethanolamine + Water) and (4-Methylmorpholine + Water) Binary Systems at Several Temperatures

pvap	1.83	kPa	348.34	Isothermal Vapor-Liquid Equilibria of (Monoethanolamine + Water) and (4-Methylmorpholine + Water) Binary Systems at Several Temperatures
pvap	2.43	kPa	353.34	Isothermal Vapor-Liquid Equilibria of (Monoethanolamine + Water) and (4-Methylmorpholine + Water) Binary Systems at Several Temperatures
pvap	3.19	kPa	358.34	Isothermal Vapor-Liquid Equilibria of (Monoethanolamine + Water) and (4-Methylmorpholine + Water) Binary Systems at Several Temperatures
pvap	4.11	kPa	363.09	Isothermal Vapor-Liquid Equilibria of (Monoethanolamine + Water) and (4-Methylmorpholine + Water) Binary Systems at Several Temperatures
pvap	6.16	kPa	373.15	Vapor Liquid Equilibrium for Several Compounds Relevant to the Biofuels Industry Modeled with the Wilson Equation
pvap	52.71	kPa	423.15	Vapor Liquid Equilibrium for Several Compounds Relevant to the Biofuels Industry Modeled with the Wilson Equation
pvap	6.17	kPa	373.15	Vapor Liquid Equilibrium for Several Compounds Relevant to the Biofuels Industry Modeled with the Wilson Equation

pvap	50.62	kPa	423.15	Vapor Liquid Equilibrium for Several Compounds Relevant to the Biofuels Industry Modeled with the Wilson Equation
pvap	95.30	kPa	442.28	Activity Coefficients and Excess Gibbs Energies for Binary Mixtures of N-Methyl-2-pyrrolidone with Some Substituted Ethanols
pvap	97.50	kPa	441.90	Vapor Pressures of Several Commercially Used Alkanolamines
pvap	71.80	kPa	432.60	Vapor Pressures of Several Commercially Used Alkanolamines
pvap	24.90	kPa	403.60	Vapor Pressures of Several Commercially Used Alkanolamines
pvap	11.70	kPa	385.70	Vapor Pressures of Several Commercially Used Alkanolamines
pvap	6.40	kPa	373.00	Vapor Pressures of Several Commercially Used Alkanolamines
pvap	3.02	kPa	358.60	Vapor Pressures of Several Commercially Used Alkanolamines
pvap	0.31	kPa	324.50	Vapor Pressures and Vaporization Enthalpies of a Series of Ethanolamines
pvap	0.26	kPa	321.40	Vapor Pressures and Vaporization Enthalpies of a Series of Ethanolamines

pvap	0.20	kPa	318.40	Vapor Pressures and Vaporization Enthalpies of a Series of Ethanolamines
pvap	0.17	kPa	315.40	Vapor Pressures and Vaporization Enthalpies of a Series of Ethanolamines
pvap	0.13	kPa	312.40	Vapor Pressures and Vaporization Enthalpies of a Series of Ethanolamines
pvap	0.11	kPa	309.40	Vapor Pressures and Vaporization Enthalpies of a Series of Ethanolamines
pvap	0.08	kPa	306.30	Vapor Pressures and Vaporization Enthalpies of a Series of Ethanolamines
pvap	0.06	kPa	303.20	Vapor Pressures and Vaporization Enthalpies of a Series of Ethanolamines
pvap	0.05	kPa	300.20	Vapor Pressures and Vaporization Enthalpies of a Series of Ethanolamines
pvap	0.04	kPa	297.20	Vapor Pressures and Vaporization Enthalpies of a Series of Ethanolamines
pvap	0.03	kPa	294.00	Vapor Pressures and Vaporization Enthalpies of a Series of Ethanolamines
pvap	0.01	kPa	281.10	Vapor Pressures and Vaporization Enthalpies of a Series of Ethanolamines
pvap	8.48e-03	kPa	279.00	Vapor Pressures and Vaporization Enthalpies of a Series of Ethanolamines
pvap	24.97	kPa	403.43	Ebulliometric Determination of Vapor

pvap	51.10	kPa	422.70	Vapor Pressures of Several Commercially Used Alkanolamines
rfi	1.45420		293.15	Experimental solubility for betulin and estrone in various solvents within the temperature range T = (293.2 to 328.2) K
rfi	1.45250		298.15	Vapor-Liquid Equilibrium and Excess Gibbs Energies of Hexane + N,N-Dimethyl Formamide, 2-Methylpropan-2-ol + 2-Aminophenol, N,N-Dimethyl Formamide, and 2-Propanol + Diisopropyl Amine at 94.4 kPa
rfi	1.43472		353.15	Densities, Refractive Indices and Excess Properties of Binary Mixtures of 1-Butyl-3-Methylimidazolium Tetraflouroborate with Water and Monoethanolamine
rfi	1.43862		343.15	Densities, Refractive Indices and Excess Properties of Binary Mixtures of 1-Butyl-3-Methylimidazolium Tetraflouroborate with Water and Monoethanolamine

rfi	1.44213	333.15	Densities, Refractive Indices and Excess Properties of Binary Mixtures of 1-Butyl-3-Methylimidazolium Tetraflouroborate with Water and Monoethanolamine
rfi	1.44561	323.15	Densities, Refractive Indices and Excess Properties of Binary Mixtures of 1-Butyl-3-Methylimidazolium Tetraflouroborate with Water and Monoethanolamine
rfi	1.44913	313.15	Densities, Refractive Indices and Excess Properties of Binary Mixtures of 1-Butyl-3-Methylimidazolium Tetraflouroborate with Water and Monoethanolamine
rfi	1.45273	303.15	Densities, Refractive Indices and Excess Properties of Binary Mixtures of 1-Butyl-3-Methylimidazolium Tetraflouroborate with Water and Monoethanolamine
rfi	1.45432	298.15	Densities, Refractive Indices and Excess Properties of Binary Mixtures of 1-Butyl-3-Methylimidazolium Tetraflouroborate with Water and Monoethanolamine

rfi	1.45601	293.15	Densities, Refractive Indices and Excess Properties of Binary Mixtures of 1-Butyl-3-Methylimidazolium Tetraflouroborate with Water and Monoethanolamine
rfi	1.44170	323.15	Density, Speed of Sound, Viscosity, Refractive Index, and Excess Volume of N-Methyl-2-pyrrolidone + Ethanol (or Water or Ethanolamine) from T = (293.15 to 323.15) K
rfi	1.44490	313.15	Density, Speed of Sound, Viscosity, Refractive Index, and Excess Volume of N-Methyl-2-pyrrolidone + Ethanol (or Water or Ethanolamine) from T = (293.15 to 323.15) K
rfi	1.44880	303.15	Density, Speed of Sound, Viscosity, Refractive Index, and Excess Volume of N-Methyl-2-pyrrolidone + Ethanol (or Water or Ethanolamine) from T = (293.15 to 323.15) K
rfi	1.45250	293.15	Density, Speed of Sound, Viscosity, Refractive Index, and Excess Volume of N-Methyl-2-pyrrolidone + Ethanol (or Water or Ethanolamine) from T = (293.15 to 323.15) K

rhol	1004.65	kg/m3	308.15	Volumetric, acoustic and spectroscopic properties of 3-chloroaniline with substituted ethanols at various temperatures
rhol	1008.80	kg/m3	303.15	Monoethanolamine+2-methoxyethanol mixtures for CO2 capture: Density, viscosity and CO2 solubility
rhol	1004.90	kg/m3	308.15	Monoethanolamine+2-methoxyethanol mixtures for CO2 capture: Density, viscosity and CO2 solubility
rhol	1000.90	kg/m3	313.15	Monoethanolamine+2-methoxyethanol mixtures for CO2 capture: Density, viscosity and CO2 solubility
rhol	996.90	kg/m3	318.15	Monoethanolamine+2-methoxyethanol mixtures for CO2 capture: Density, viscosity and CO2 solubility
rhol	992.90	kg/m3	323.15	Monoethanolamine+2-methoxyethanol mixtures for CO2 capture: Density, viscosity and CO2 solubility
rhol	988.80	kg/m3	328.15	Monoethanolamine+2-methoxyethanol mixtures for CO2 capture: Density, viscosity and CO2 solubility
rhol	984.80	kg/m3	333.15	Monoethanolamine+2-methoxyethanol mixtures for CO2 capture: Density, viscosity and CO2 solubility
rhol	980.70	kg/m3	338.15	Monoethanolamine+2-methoxyethanol mixtures for CO2 capture: Density, viscosity and CO2 solubility

rhol	976.60	kg/m3	343.15	Monoethanolamine+2-methoxyethanol mixtures for CO2 capture: Density, viscosity and CO2 solubility
rhol	972.50	kg/m3	348.15	Monoethanolamine+2-methoxyethanol mixtures for CO2 capture: Density, viscosity and CO2 solubility
rhol	968.40	kg/m3	353.15	Monoethanolamine+2-methoxyethanol mixtures for CO2 capture: Density, viscosity and CO2 solubility
rhol	1016.47	kg/m3	293.15	Volumetric properties of monoethanolamine and alcohol binary mixtures at different temperatures and 0.1 MPa
rhol	1008.72	kg/m3	303.15	Volumetric, acoustic and spectroscopic properties of 3-chloroaniline with substituted ethanols at various temperatures
rhol	1008.47	kg/m3	303.15	Volumetric properties of monoethanolamine and alcohol binary mixtures at different temperatures and 0.1 MPa
rhol	1004.49	kg/m3	308.15	Volumetric properties of monoethanolamine and alcohol binary mixtures at different temperatures and 0.1 MPa
rhol	1000.50	kg/m3	313.15	Volumetric properties of monoethanolamine and alcohol binary mixtures at different temperatures and 0.1 MPa

rhol	996.39	kg/m3	318.15	Volumetric properties of monoethanolamine and alcohol binary mixtures at different temperatures and 0.1 MPa
rhol	992.41	kg/m3	323.15	Volumetric properties of monoethanolamine and alcohol binary mixtures at different temperatures and 0.1 MPa
rhol	1016.30	kg/m3	293.15	Volumetric and viscometric properties of ternary solution of (N-methyldiethanolamine + monoethanolamine + ethanol)
rhol	1008.02	kg/m3	303.15	Volumetric and viscometric properties of ternary solution of (N-methyldiethanolamine + monoethanolamine + ethanol)
rhol	999.98	kg/m3	313.15	Volumetric and viscometric properties of ternary solution of (N-methyldiethanolamine + monoethanolamine + ethanol)
rhol	991.97	kg/m3	323.15	Volumetric and viscometric properties of ternary solution of (N-methyldiethanolamine + monoethanolamine + ethanol)
rhol	1008.80	kg/m3	303.15	Solubility of N2O and CO2 in non-aqueous systems of monoethanolamine and glycol ethers: Measurements and model representation

rhol	1000.90	kg/m3	313.15	Solubility of N2O and CO2 in non-aqueous systems of monoethanolamine and glycol ethers: Measurements and model representation
rhol	1016.90	kg/m3	293.15	Volumetric and viscometric properties of binary and ternary mixtures of monoethanolamine, 2-(diethylamino) ethanol and water from (293.15 to 333.15) K
rhol	959.54	kg/m3	363.15	Volumetric properties of binary mixtures of dimethyl sulfoxide with amines from (293.15 to 363.15) K
rhol	1001.10	kg/m3	313.15	Volumetric and viscometric properties of binary and ternary mixtures of monoethanolamine, 2-(diethylamino) ethanol and water from (293.15 to 333.15) K
rhol	992.70	kg/m3	323.15	Volumetric and viscometric properties of binary and ternary mixtures of monoethanolamine, 2-(diethylamino) ethanol and water from (293.15 to 333.15) K

rhol	984.30	kg/m3	333.15	Volumetric and viscometric properties of binary and ternary mixtures of monoethanolamine, 2-(diethylamino) ethanol and water from (293.15 to 333.15) K	
rhol	1025.55	kg/m3	281.15	Volumetric properties of the monoethanolamine methanol mixture at atmospheric pressure from 283.15 to 353.15K	
rhol	1023.99	kg/m3	283.15	Volumetric properties of the monoethanolamine methanol mixture at atmospheric pressure from 283.15 to 353.15K	
rhol	1022.42	kg/m3	285.15	Volumetric properties of the monoethanolamine methanol mixture at atmospheric pressure from 283.15 to 353.15K	
rhol	1020.84	kg/m3	287.15	Volumetric properties of the monoethanolamine methanol mixture at atmospheric pressure from 283.15 to 353.15K	
rhol	1019.26	kg/m3	289.15	Volumetric properties of the monoethanolamine methanol mixture at atmospheric pressure from 283.15 to 353.15K	
rhol	1017.68	kg/m3	291.15	Volumetric properties of the monoethanolamine methanol mixture at atmospheric pressure from 283.15 to 353.15K	

rhol	1016.10	kg/m3	293.15	Volumetric properties of the monoethanolamine methanol mixture at atmospheric pressure from 283.15 to 353.15K
rhol	1014.52	kg/m3	295.15	Volumetric properties of the monoethanolamine methanol mixture at atmospheric pressure from 283.15 to 353.15K
rhol	1012.94	kg/m3	297.15	Volumetric properties of the monoethanolamine methanol mixture at atmospheric pressure from 283.15 to 353.15K
rhol	1011.35	kg/m3	299.15	Volumetric properties of the monoethanolamine methanol mixture at atmospheric pressure from 283.15 to 353.15K
rhol	1009.76	kg/m3	301.15	Volumetric properties of the monoethanolamine methanol mixture at atmospheric pressure from 283.15 to 353.15K
rhol	1008.17	kg/m3	303.15	Volumetric properties of the monoethanolamine methanol mixture at atmospheric pressure from 283.15 to 353.15K
rhol	1006.58	kg/m3	305.15	Volumetric properties of the monoethanolamine methanol mixture at atmospheric pressure from 283.15 to 353.15K

rhol	1004.99	kg/m3	307.15	Volumetric properties of the monoethanolamine methanol mixture at atmospheric pressure from 283.15 to 353.15K
rhol	1003.40	kg/m3	309.15	Volumetric properties of the monoethanolamine methanol mixture at atmospheric pressure from 283.15 to 353.15K
rhol	1001.80	kg/m3	311.15	Volumetric properties of the monoethanolamine methanol mixture at atmospheric pressure from 283.15 to 353.15K
rhol	1000.21	kg/m3	313.15	Volumetric properties of the monoethanolamine methanol mixture at atmospheric pressure from 283.15 to 353.15K
rhol	998.60	kg/m3	315.15	Volumetric properties of the monoethanolamine methanol mixture at atmospheric pressure from 283.15 to 353.15K
rhol	997.00	kg/m3	317.15	Volumetric properties of the monoethanolamine methanol mixture at atmospheric pressure from 283.15 to 353.15K
rhol	995.40	kg/m3	319.15	Volumetric properties of the monoethanolamine methanol mixture at atmospheric pressure from 283.15 to 353.15K

rhol	993.79	kg/m3	321.15	Volumetric properties of the monoethanolamine methanol mixture at atmospheric pressure from 283.15 to 353.15K
rhol	992.19	kg/m3	323.15	Volumetric properties of the monoethanolamine methanol mixture at atmospheric pressure from 283.15 to 353.15K
rhol	990.58	kg/m3	325.15	Volumetric properties of the monoethanolamine methanol mixture at atmospheric pressure from 283.15 to 353.15K
rhol	988.96	kg/m3	327.15	Volumetric properties of the monoethanolamine methanol mixture at atmospheric pressure from 283.15 to 353.15K
rhol	987.34	kg/m3	329.15	Volumetric properties of the monoethanolamine methanol mixture at atmospheric pressure from 283.15 to 353.15K
rhol	985.72	kg/m3	331.15	Volumetric properties of the monoethanolamine methanol mixture at atmospheric pressure from 283.15 to 353.15K
rhol	984.10	kg/m3	333.15	Volumetric properties of the monoethanolamine methanol mixture at atmospheric pressure from 283.15 to 353.15K

rhol	982.48	kg/m3	335.15	Volumetric properties of the monoethanolamine methanol mixture at atmospheric pressure from 283.15 to 353.15K
rhol	980.85	kg/m3	337.15	Volumetric properties of the monoethanolamine methanol mixture at atmospheric pressure from 283.15 to 353.15K
rhol	979.21	kg/m3	339.15	Volumetric properties of the monoethanolamine methanol mixture at atmospheric pressure from 283.15 to 353.15K
rhol	977.58	kg/m3	341.15	Volumetric properties of the monoethanolamine methanol mixture at atmospheric pressure from 283.15 to 353.15K
rhol	975.94	kg/m3	343.15	Volumetric properties of the monoethanolamine methanol mixture at atmospheric pressure from 283.15 to 353.15K
rhol	974.30	kg/m3	345.15	Volumetric properties of the monoethanolamine methanol mixture at atmospheric pressure from 283.15 to 353.15K
rhol	972.65	kg/m3	347.15	Volumetric properties of the monoethanolamine methanol mixture at atmospheric pressure from 283.15 to 353.15K

rhol	970.99	kg/m3	349.15	Volumetric properties of the monoethanolamine methanol mixture at atmospheric pressure from 283.15 to 353.15K
rhol	969.34	kg/m3	351.15	Volumetric properties of the monoethanolamine methanol mixture at atmospheric pressure from 283.15 to 353.15K
rhol	967.68	kg/m3	353.15	Volumetric properties of the monoethanolamine methanol mixture at atmospheric pressure from 283.15 to 353.15K
rhol	1012.94	kg/m3	297.15	Volumetric Properties of Water + Monoethanolamine + Methanol Mixtures at Atmospheric Pressure from 283.15 to 353.15 K
rhol	1010.67	kg/m3	303.15	Densities and Excess Molar Volumes of Binary Mixtures of Bis(2-hydroxyethyl)ammonium Acetate + Water and Monoethanolamine + Bis(2-hydroxyethyl)ammonium Acetate at Temperatures from (303.15 to 353.15) K
rhol	1002.77	kg/m3	313.15	Densities and Excess Molar Volumes of Binary Mixtures of Bis(2-hydroxyethyl)ammonium Acetate + Water and Monoethanolamine + Bis(2-hydroxyethyl)ammonium Acetate at Temperatures from (303.15 to 353.15) K

rhol	994.80	kg/m3	323.15	Densities and Excess Molar Volumes of Binary Mixtures of Bis(2-hydroxyethyl)ammonium Acetate + Water and Monoethanolamine + Bis(2-hydroxyethyl)ammonium Acetate at Temperatures from (303.15 to 353.15) K
rhol	986.75	kg/m3	333.15	Densities and Excess Molar Volumes of Binary Mixtures of Bis(2-hydroxyethyl)ammonium Acetate + Water and Monoethanolamine + Bis(2-hydroxyethyl)ammonium Acetate at Temperatures from (303.15 to 353.15) K
rhol	978.62	kg/m3	343.15	Densities and Excess Molar Volumes of Binary Mixtures of Bis(2-hydroxyethyl)ammonium Acetate + Water and Monoethanolamine + Bis(2-hydroxyethyl)ammonium Acetate at Temperatures from (303.15 to 353.15) K
rhol	970.40	kg/m3	353.15	Densities and Excess Molar Volumes of Binary Mixtures of Bis(2-hydroxyethyl)ammonium Acetate + Water and Monoethanolamine + Bis(2-hydroxyethyl)ammonium Acetate at Temperatures from (303.15 to 353.15) K

rhol	1011.98	kg/m3	298.15	Density, Speed of Sound, and Viscosity of N-Methyl-2-pyrrolidone + Ethanolamine + Water from T = (293.15 to 323.15) K
rhol	1007.90	kg/m3	303.15	Densities, Viscosities and Excess Properties of Binary Mixtures of 1,1,3,3-Tetramethylguanidinium Lactate + Water at T = (303.15 to 328.15) K
rhol	1004.20	kg/m3	308.15	Densities, Viscosities and Excess Properties of Binary Mixtures of 1,1,3,3-Tetramethylguanidinium Lactate + Water at T = (303.15 to 328.15) K
rhol	1000.50	kg/m3	313.15	Densities, Viscosities and Excess Properties of Binary Mixtures of 1,1,3,3-Tetramethylguanidinium Lactate + Water at T = (303.15 to 328.15) K
rhol	996.20	kg/m3	318.15	Densities, Viscosities and Excess Properties of Binary Mixtures of 1,1,3,3-Tetramethylguanidinium Lactate + Water at T = (303.15 to 328.15) K
rhol	992.70	kg/m3	323.15	Densities, Viscosities and Excess Properties of Binary Mixtures of 1,1,3,3-Tetramethylguanidinium Lactate + Water at T = (303.15 to 328.15) K

rhol	988.70	kg/m3	328.15	Densities, Viscosities and Excess Properties of Binary Mixtures of 1,1,3,3-Tetramethylguanidinium Lactate + Water at T = (303.15 to 328.15) K
rhol	1024.11	kg/m3	283.15	Densities and Excess Properties of Primary Amines in Alcoholic Solutions
rhol	1016.23	kg/m3	293.15	Densities and Excess Properties of Primary Amines in Alcoholic Solutions
rhol	1008.32	kg/m3	303.15	Densities and Excess Properties of Primary Amines in Alcoholic Solutions
rhol	1000.37	kg/m3	313.15	Densities and Excess Properties of Primary Amines in Alcoholic Solutions
rhol	992.36	kg/m3	323.15	Densities and Excess Properties of Primary Amines in Alcoholic Solutions
rhol	984.30	kg/m3	333.15	Densities and Excess Properties of Primary Amines in Alcoholic Solutions
rhol	976.15	kg/m3	343.15	Densities and Excess Properties of Primary Amines in Alcoholic Solutions

rhol	1016.10	kg/m3	293.15	Density and Viscosity for Binary Mixtures of Diethylene Glycol Monobutyl Ether with Monoethanolamine, Diethanolamine, and Triethanolamine from (293.15 to 333.15) K
rhol	1008.40	kg/m3	303.15	Density and Viscosity for Binary Mixtures of Diethylene Glycol Monobutyl Ether with Monoethanolamine, Diethanolamine, and Triethanolamine from (293.15 to 333.15) K
rhol	1000.10	kg/m3	313.15	Density and Viscosity for Binary Mixtures of Diethylene Glycol Monobutyl Ether with Monoethanolamine, Diethanolamine, and Triethanolamine from (293.15 to 333.15) K
rhol	991.90	kg/m3	323.15	Density and Viscosity for Binary Mixtures of Diethylene Glycol Monobutyl Ether with Monoethanolamine, Diethanolamine, and Triethanolamine from (293.15 to 333.15) K
rhol	983.60	kg/m3	333.15	Density and Viscosity for Binary Mixtures of Diethylene Glycol Monobutyl Ether with Monoethanolamine, Diethanolamine, and Triethanolamine from (293.15 to 333.15) K

rhol	1019.87	kg/m3	288.15	Excess Enthalpy and Excess Volume for Pyridine + Methyldiethanolamine and Pyridine + Ethanolamine Mixtures
rhol	1012.12	kg/m3	298.15	Excess Enthalpy and Excess Volume for Pyridine + Methyldiethanolamine and Pyridine + Ethanolamine Mixtures
rhol	992.38	kg/m3	323.15	Volumetric properties of binary mixtures of dimethyl sulfoxide with amines from (293.15 to 363.15) K
rhol	1011.98	kg/m3	298.15	Density, Speed of Sound, and Viscosity of Monoethanolamine + Water + N-Ethyl-2-pyrrolidone from T = (293.15 to 323.15) K
rhol	1012.38	kg/m3	298.15	Physicochemical Properties of Aqueous Solutions of Sodium l-Prolinate as an Absorbent for CO2 Removal
rhol	1009.35	kg/m3	303.15	Physicochemical Properties of Aqueous Solutions of Sodium l-Prolinate as an Absorbent for CO2 Removal
rhol	1002.16	kg/m3	313.15	Physicochemical Properties of Aqueous Solutions of Sodium l-Prolinate as an Absorbent for CO2 Removal

rhol	994.40	kg/m3	323.15	Physicochemical Properties of Aqueous Solutions of Sodium l-Prolinate as an Absorbent for CO2 Removal
rhol	986.24	kg/m3	333.15	Physicochemical Properties of Aqueous Solutions of Sodium l-Prolinate as an Absorbent for CO2 Removal
rhol	967.89	kg/m3	353.14	Volumetric properties of binary mixtures of dimethyl sulfoxide with amines from (293.15 to 363.15) K
rhol	976.14	kg/m3	343.14	Volumetric properties of binary mixtures of dimethyl sulfoxide with amines from (293.15 to 363.15) K
rhol	984.30	kg/m3	333.14	Volumetric properties of binary mixtures of dimethyl sulfoxide with amines from (293.15 to 363.15) K
rhol	1000.41	kg/m3	313.14	Volumetric properties of binary mixtures of dimethyl sulfoxide with amines from (293.15 to 363.15) K
rhol	1008.40	kg/m3	303.14	Volumetric properties of binary mixtures of dimethyl sulfoxide with amines from (293.15 to 363.15) K

rhol	1016.34	kg/m3	293.16	Volumetric properties of binary mixtures of dimethyl sulfoxide with amines from (293.15 to 363.15) K
rhol	1012.10	kg/m3	293.15	Viscosity of aqueous solutions of 2-methoxyethanol, 2-ethoxyethanol, and ethanolamine
rhol	1004.50	kg/m3	308.15	Thermodynamic properties and CO2 solubility of monoethanolamine + diethylenetriamine/aminoethylethanolamine mixtures: Experimental measurements and thermodynamic modeling
rhol	1008.40	kg/m3	303.15	Thermodynamic properties and CO2 solubility of monoethanolamine + diethylenetriamine/aminoethylethanolamine mixtures: Experimental measurements and thermodynamic modeling
rhol	1012.40	kg/m3	298.15	Thermodynamic properties and CO2 solubility of monoethanolamine + diethylenetriamine/aminoethylethanolamine mixtures: Experimental measurements and thermodynamic modeling
rhol	1004.50	kg/m3	308.15	Thermodynamic properties and CO2 solubility of monoethanolamine + diethylenetriamine/aminoethylethanolamine mixtures: Experimental measurements and thermodynamic modeling

rhol	1012.80	kg/m3	298.15	Monoethanolamine+2-methoxyethanol mixtures for CO2 capture: Density, viscosity and CO2 solubility
rhol	1016.70	kg/m3	293.15	Monoethanolamine+2-methoxyethanol mixtures for CO2 capture: Density, viscosity and CO2 solubility
rhol	1012.00	kg/m3	298.15	Volumetric, acoustic and transport properties of mixtures containing dimethyl sulfoxide and some amines or alkanolamines: Measurement and correlation
rhol	992.00	kg/m3	323.15	Volumetric, acoustic and transport properties of mixtures containing dimethyl sulfoxide and some amines or alkanolamines: Measurement and correlation
rhol	1000.00	kg/m3	313.15	Volumetric, acoustic and transport properties of mixtures containing dimethyl sulfoxide and some amines or alkanolamines: Measurement and correlation
rhol	1008.00	kg/m3	303.15	Volumetric, acoustic and transport properties of mixtures containing dimethyl sulfoxide and some amines or alkanolamines: Measurement and correlation

rhol	1016.00	kg/m3	293.15	Volumetric, acoustic and transport properties of mixtures containing dimethyl sulfoxide and some amines or alkanolamines: Measurement and correlation
rhol	983.60	kg/m3	333.15	Volumetric and viscometric properties of binary and ternary mixtures of 1-butyl-3-methylimidazolium tetrafluoroborate, monoethanolamine and water
rhol	991.70	kg/m3	323.15	Volumetric and viscometric properties of binary and ternary mixtures of 1-butyl-3-methylimidazolium tetrafluoroborate, monoethanolamine and water
rhol	999.70	kg/m3	313.15	Volumetric and viscometric properties of binary and ternary mixtures of 1-butyl-3-methylimidazolium tetrafluoroborate, monoethanolamine and water
rhol	1007.60	kg/m3	303.15	Volumetric and viscometric properties of binary and ternary mixtures of 1-butyl-3-methylimidazolium tetrafluoroborate, monoethanolamine and water
rhol	1015.50	kg/m3	293.15	Volumetric and viscometric properties of binary and ternary mixtures of 1-butyl-3-methylimidazolium tetrafluoroborate, monoethanolamine and water

rhol	996.04	kg/m3	318.15	Volumetric, acoustic and spectroscopic properties of 3-chloroaniline with substituted ethanols at various temperatures
rhol	1000.74	kg/m3	313.15	Volumetric, acoustic and spectroscopic properties of 3-chloroaniline with substituted ethanols at various temperatures
rhol	1008.40	kg/m3	303.15	Thermodynamic properties and CO2 solubility of monoethanolamine + diethylenetriamine/aminoethylethanolamine mixtures: Experimental measurements and thermodynamic modeling
rhol	1012.40	kg/m3	298.15	Thermodynamic properties and CO2 solubility of monoethanolamine + diethylenetriamine/aminoethylethanolamine mixtures: Experimental measurements and thermodynamic modeling
rhol	998.15	kg/m3	318.15	Hydrogen bond interactions in the blends of 1,4-dioxane with some 1, 2-disubstituted ethanes at T = (298.15, 308.15 and 318.15) K
rhol	1004.68	kg/m3	308.15	Hydrogen bond interactions in the blends of 1,4-dioxane with some 1, 2-disubstituted ethanes at T = (298.15, 308.15 and 318.15) K

rhol	1011.79	kg/m3	298.15	Hydrogen bond interactions in the blends of 1,4-dioxane with some 1, 2-disubstituted ethanes at T = (298.15, 308.15 and 318.15) K
rhol	1006.80	kg/m3	308.15	Liquid-liquid phase equilibrium for ternary mixtures of formamide (or ethylene glycol, or monoethanolamine) + indole + 2-methylnaphthalene at 308.15 K
rhol	983.98	kg/m3	333.15	Solubility of CO2 in aqueous mixtures of monoethanolamine and dicyanamide-based ionic liquids
rhol	992.07	kg/m3	323.15	Solubility of CO2 in aqueous mixtures of monoethanolamine and dicyanamide-based ionic liquids
rhol	1000.09	kg/m3	313.15	Solubility of CO2 in aqueous mixtures of monoethanolamine and dicyanamide-based ionic liquids
rhol	1008.05	kg/m3	303.15	Solubility of CO2 in aqueous mixtures of monoethanolamine and dicyanamide-based ionic liquids
rhol	1015.96	kg/m3	293.15	Solubility of CO2 in aqueous mixtures of monoethanolamine and dicyanamide-based ionic liquids
rhol	1017.90	kg/m3	293.15	Experiment and model for the viscosity of carbonated MDEA - MEA aqueous solutions

rhol	1012.53	kg/m3	298.15	Volumetric properties of monoethanolamine and alcohol binary mixtures at different temperatures and 0.1 MPa
rhol	1009.50	kg/m3	303.15	Volumetric and viscometric properties of binary and ternary mixtures of monoethanolamine, 2-(diethylamino) ethanol and water from (293.15 to 333.15) K
rhol	968.33	kg/m3	353.15	Excess Enthalpy and Excess Volume for Pyridine + Methyldiethanolamine and Pyridine + Ethanolamine Mixtures
speedsl	1654.19	m/s	318.15	Density, Speed of Sound, Isentropic Compressibility, and Excess Volume of (Monoethanolamine + 2-Amino-2-methyl-1-propanol), (Monoethanolamine + Triethanolamine), and (Monoethanolamine + N-Methyldiethanolamine) at Temperatures from (293.15 to 323.15) K

speedsl	1670.54	m/s	313.15	Density, Speed of Sound, Isentropic Compressibility, and Excess Volume of (Monoethanolamine + 2-Amino-2-methyl-1-propanol), (Monoethanolamine + Triethanolamine), and (Monoethanolamine + N-Methyldiethanolamine) at Temperatures from (293.15 to 323.15) K
speedsl	1686.79	m/s	308.15	Density, Speed of Sound, Isentropic Compressibility, and Excess Volume of (Monoethanolamine + 2-Amino-2-methyl-1-propanol), (Monoethanolamine + Triethanolamine), and (Monoethanolamine + N-Methyldiethanolamine) at Temperatures from (293.15 to 323.15) K
speedsl	1703.09	m/s	303.15	Density, Speed of Sound, Isentropic Compressibility, and Excess Volume of (Monoethanolamine + 2-Amino-2-methyl-1-propanol), (Monoethanolamine + Triethanolamine), and (Monoethanolamine + N-Methyldiethanolamine) at Temperatures from (293.15 to 323.15) K

speedsl	1719.23	m/s	298.15	Density, Speed of Sound, Isentropic Compressibility, and Excess Volume of (Monoethanolamine + 2-Amino-2-methyl-1-propanol), (Monoethanolamine + Triethanolamine), and (Monoethanolamine + N-Methyldiethanolamine) at Temperatures from (293.15 to 323.15) K
speedsl	1735.46	m/s	293.15	Density, Speed of Sound, Isentropic Compressibility, and Excess Volume of (Monoethanolamine + 2-Amino-2-methyl-1-propanol), (Monoethanolamine + Triethanolamine), and (Monoethanolamine + N-Methyldiethanolamine) at Temperatures from (293.15 to 323.15) K
speedsl	1637.20	m/s	323.15	Density, Speed of Sound, Viscosity, Surface Tension, and Excess Volume of N-Ethyl-2-pyrrolidone + Ethanolamine (or Diethanolamine or Triethanolamine) from T = (293.15 to 323.15) K

speedsl	1670.10	m/s	313.15	Density, Speed of Sound, Viscosity, Surface Tension, and Excess Volume of N-Ethyl-2-pyrrolidone + Ethanolamine (or Diethanolamine or Triethanolamine) from T = (293.15 to 323.15) K
speedsl	1702.10	m/s	303.15	Density, Speed of Sound, Viscosity, Surface Tension, and Excess Volume of N-Ethyl-2-pyrrolidone + Ethanolamine (or Diethanolamine or Triethanolamine) from T = (293.15 to 323.15) K
speedsl	1734.40	m/s	293.15	Density, Speed of Sound, Viscosity, Surface Tension, and Excess Volume of N-Ethyl-2-pyrrolidone + Ethanolamine (or Diethanolamine or Triethanolamine) from T = (293.15 to 323.15) K
speedsl	1637.86	m/s	323.15	Density, Speed of Sound, Isentropic Compressibility, and Excess Volume of (Monoethanolamine + 2-Amino-2-methyl-1-propanol), (Monoethanolamine + Triethanolamine), and (Monoethanolamine + N-Methyldiethanolamine) at Temperatures from (293.15 to 323.15) K

srf	0.05	N/m	303.20	Investigation of surface tension and viscosity for aqueous solutions of MEA-MeOH and DEA-MeOH
srf	0.05	N/m	313.15	Densities and Surface Tensions of CO2 Loaded Aqueous Monoethanolamine Solutions with r = (0.2 to 0.7) at T = (303.15 to 333.15) K
srf	0.04	N/m	323.20	Investigation of surface tension and viscosity for aqueous solutions of MEA-MeOH and DEA-MeOH
srf	0.05	N/m	303.15	Density of Water (1) + Monoethanolamine (2) + CO2 (3) from (298.15 to 413.15) K and Surface Tension of Water (1) + Monoethanolamine (2) from (303.15 to 333.15) K
srf	0.05	N/m	313.15	Density of Water (1) + Monoethanolamine (2) + CO2 (3) from (298.15 to 413.15) K and Surface Tension of Water (1) + Monoethanolamine (2) from (303.15 to 333.15) K
srf	0.05	N/m	323.15	Density of Water (1) + Monoethanolamine (2) + CO2 (3) from (298.15 to 413.15) K and Surface Tension of Water (1) + Monoethanolamine (2) from (303.15 to 333.15) K

srf	0.04	N/m	333.15	Density of Water (1) + Monoethanolamine (2) + CO2 (3) from (298.15 to 413.15) K and Surface Tension of Water (1) + Monoethanolamine (2) from (303.15 to 333.15) K
srf	0.05	N/m	293.15	Density and Surface Tension Measurements of Partially Carbonated Aqueous Monoethanolamine Solutions
srf	0.05	N/m	303.15	Density and Surface Tension Measurements of Partially Carbonated Aqueous Monoethanolamine Solutions
srf	0.05	N/m	313.15	Density and Surface Tension Measurements of Partially Carbonated Aqueous Monoethanolamine Solutions
srf	0.05	N/m	293.15	Densities and Surface Tensions of CO2 Loaded Aqueous Monoethanolamine Solutions with r = (0.2 to 0.7) at T = (303.15 to 333.15) K
srf	0.05	N/m	303.15	Densities and Surface Tensions of CO2 Loaded Aqueous Monoethanolamine Solutions with r = (0.2 to 0.7) at T = (303.15 to 333.15) K
srf	0.05	N/m	313.20	Investigation of surface tension and viscosity for aqueous solutions of MEA-MeOH and DEA-MeOH

Correlations

Information	Value
Property code	pvap
Equation	ln(Pvp) = A + B/T + C*ln(T) + D*T^2
Coeff. A	1.62291e+02
Coeff. B	-1.33605e+04
Coeff. C	-2.13543e+01
Coeff. D	1.31668e-05
Temperature range (K), min.	283.65
Temperature range (K), max.	638.00

Datasets

Viscosity, Pa*s

Temperature, K - Liquid	Pressure, kPa - Liquid	Viscosity, Pa*s - Liquid
298.15	100.00	0.0187400

Reference *https://www.doi.org/10.1021/acs.jced.5b00447*

Mass density, kg/m3

Temperature, K - Liquid	Pressure, kPa - Liquid	Mass density, kg/m3 - Liquid
298.15	100.00	1012.0

Reference *https://www.doi.org/10.1021/acs.jced.7b01101*

Sources

Solubility of Carbon Dioxide in Aqueous Solutions of Monoethanolamine in the Low and High Gas Loading Regions:	https://www.doi.org/10.1021/je301030z

Absorption performance of (CO2 + N2) gas mixtures in amino acid ionic liquids promoted N-methyldiethanolamine aqueous solutions:	https://www.doi.org/10.1016/j.jct.2017.06.010
Experimental solubility for betulin and estrone in various solvents within the temperature range T = (293.2 to 328.2) K:	https://www.doi.org/10.1016/j.jct.2016.02.006
Viscosity of aqueous solutions of 2-methoxyethanol, 2-ethoxyethanol, and ethanolamine:	https://www.doi.org/10.1016/j.jct.2004.06.004
NIST Webbook:	http://webbook.nist.gov/cgi/cbook.cgi?ID=C141435&Units=SI
Density, Viscosity, Heat Capacity, Surface Tension, and Solubility of CO2 in Aqueous Solutions of Potassium Serinate :	https://www.doi.org/10.1021/je101144k
Joback Method:	https://en.wikipedia.org/wiki/Joback_method
Crippen Method:	http://pubs.acs.org/doi/abs/10.1021/ci990307l
Volumetric, acoustic and spectroscopic properties of 3-chloroaniline with substituted ethanols at various temperatures:	https://www.doi.org/10.1016/j.jct.2015.11.012
Vapor Pressures of Several Commercially Used Alkanolamines:	https://www.doi.org/10.1021/je101259r
Experiment and model for the surface tension of amine - ionic liquids aqueous solutions:	https://www.doi.org/10.1016/j.jct.2014.03.032
Density, Speed of Sound, and Viscosity of Monoethanolamine + Water + N-Ethyl-2-pyrrolidone from T = (293.15 to 323.15) K:	https://www.doi.org/10.1021/je4005883
Densities and Excess Molar Volumes of Binary Mixtures of Bis(2-hydroxyethyl)ammonium Acetate + Water and Monoethanolamine + Bis(2-hydroxyethyl)ammonium Acetate at Temperatures from (303.15 to 353.15) K:	https://www.doi.org/10.1021/je1008125
Experimental solubility of diosgenin and estriol in various solvents between T = (293.2 and 328.2) K:	https://www.doi.org/10.1016/j.jct.2016.11.017
Physical Properties of Partially CO2 Loaded Aqueous Monoethanolamine (MEA):	https://www.doi.org/10.1021/je401081e
Density, Viscosity, and Performances of Carbon Dioxide Capture in 16 Absorbents of Amine + Ionic Liquid + H2O, Ionic Liquid + H2O, and Amine + H2O Systems:	https://www.doi.org/10.1021/je100078w
Density and Viscosity of Aqueous Solutions of (N-Methyldiethanolamine + Monoethanolamine), (N-Methyldiethanolamine + Diethanolamine), (2-Amino-2-methyl-1-propanol + Monoethanolamine), and (2-Amino-2-methyl-1-propanol + Diethanolamine):	https://www.doi.org/10.1021/je020206a
Densities, Viscosities and Excess Properties of Binary Mixtures of 1,1,3,3-Tetramethylguanidinium Lactate + Water at T = (303.15 to 328.15) K:	https://www.doi.org/10.1021/je3009073
Bubble Point Determination for CO2 + Ethanol + Alkanolamines (Monoethanolamine, Diethanolamine, or Triethanolamine) at High Pressures:	https://www.doi.org/10.1021/je5002957
Volumetric, acoustic and transport properties of mixtures containing dimethyl sulfoxide and some amines or alkanolamines: Measurement and correlation:	https://www.doi.org/10.1016/j.jct.2018.02.018
Vapor Pressures and Vaporization Enthalpies of a Series of Ethanolamines:	https://www.doi.org/10.1021/je049761y
Experiment and model for the surface tension of carbonated MEA-MDEA aqueous solutions:	https://www.doi.org/10.1016/j.fluid.2012.10.013
McGowan Method:	http://link.springer.com/article/10.1007/BF02311772
Hydrogen bond interactions in the blends of 1,4-dioxane with some 1, 2-disubstituted ethanes at T = (298.15, 308.15 and 318.15) K:	https://www.doi.org/10.1016/j.fluid.2015.06.041
Density and Surface Tension Measurements of Partially Carbonated Aqueous Monoethanolamine Solutions:	https://www.doi.org/10.1021/je300920t
Specific Heat Capacities of Two Functional Ionic Liquids and Two Functional Deep Eutectic Solvents for the Absorption of SO2:	https://www.doi.org/10.1021/acs.jced.7b00102
Solubility of N2O in Density and Viscosity of Aqueous Solutions of 1,4-Butanediamine, 2-(Diethylamino)-ethanol, and Their Mixtures from (298.15 to 333.15) K:	https://www.doi.org/10.1021/je301371p
Density, Speed of Sound, Viscosity, Refractive Index, and Excess Volume of N-Methyl-2-pyrrolidone + Ethanol (or Water or Ethanolamine) from T = (293.15 to 323.15) K:	https://www.doi.org/10.1021/je100967k
Activity Coefficients and Excess Gibbs Energies for Binary Mixtures of N-Methyl-2-pyrrolidone with Some Substituted Ethanols:	https://www.doi.org/10.1021/je3002078
Solubility data for toluene in various aqueous alkanolamine solutions:	https://www.doi.org/10.1016/j.jct.2006.07.027
Surface thermodynamics of DMA2P, DMA2P-MEA and DMA2P-PZ aqueous solutions:	https://www.doi.org/10.1016/j.jct.2016.12.022
High pressure solubility of carbon dioxide (CO2) in aqueous piperazine solutions:	https://www.doi.org/10.1016/j.fluid.2010.01.009
Monoethanolamine-based deep eutectic solvents, their synthesis and characterization:	https://www.doi.org/10.1016/j.fluid.2017.03.008
Carbon dioxide solubility in aqueous potassium salt solutions of L-proline and DL-.alpha.-aminobutyric acid at high pressures:	https://www.doi.org/10.1016/j.jct.2014.12.010
KDB:	https://www.cheric.org/files/research/kdb/mol/mol1317.mol
Crippen Method:	https://www.chemeo.com/doc/models/crippen_log10ws
Density, Speed of Sound, Viscosity, and Surface Tension of Dimethylethylenediamine + Water and (Ethanolamine + Dimethylethanolamine) + Water from T = (298.15 to 323.15) K:	https://www.doi.org/10.1021/acs.jced.5b00447
Density and Viscosity of Partially Carbonated Aqueous Tertiary Alkanolamine Solutions at Temperatures between (298.15 and 353.15) K:	https://www.doi.org/10.1021/acs.jced.5b00282
KDB Vapor Pressure Data:	https://www.cheric.org/research/kdb/hcprop/showprop.php?cmpid=1317
CO2 Solubility in Hybrid Solvents Containing 1-Butyl-3-methylimidazolium tetrafluoroborate and Mixtures of Alkanolamines:	https://www.doi.org/10.1021/acs.jced.5b00273
Liquid-Liquid Equilibrium Measurements and Correlation for Ternary Systems (Butyl Acetate + 1-Butanol + Ethylene Glycol/1,3-Propanediol/Ethanolamine) at 298.15 K:	https://www.doi.org/10.1021/acs.jced.8b01219
Ebulliometric Determination of Vapor:	https://www.doi.org/10.1021/je800290k
Density, Speed of Sound, and Viscosity of N-Methyl-2-pyrrolidone + Ethanolamine + Water from T = (293.15 to 323.15) K:	https://www.doi.org/10.1021/je300742t
CO2 absorption using aqueous potassium lysinate solutions: Vapor - liquid equilibrium data and modelling:	https://www.doi.org/10.1016/j.jct.2017.07.041

Freezing Point Depressions of Aqueous MEA, MDEA, and MEA-MDEA Measured with a New Apparatus.	https://www.doi.org/10.1021/je100994v
Vapor−Liquid Equilibrium and Excess Gibbs Energies of Hexane + N,N-Dimethylformamide.	https://www.doi.org/10.1021/je7001336
Density, Viscosity, and N2O Solubility of Aqueous Solutions of MEA + MDEA and MEA + DEA	https://www.doi.org/10.1021/acs.jced.8b00070
Densities and Viscosities of 2-Methylpiperidine + 2-Aminophenol, + 3-Aminophenol, and + 4-Aminophenol	https://www.doi.org/10.1021/je060108f
Solubility and Diffusivity of Nitrous Oxide in Aqueous Solutions of 2-Piperidineethanol and 2-(Diethylamino)-ethanol	https://www.doi.org/10.1016/j.jct.2016.03.021
Density and Viscosity Measurements of (2-Piperidineethanol + Diethanolamine) and (2-Piperidineethanol + Monoethanolamine)	https://www.doi.org/10.1021/je4000372
Solubility of Carbon Dioxide in Aqueous Blends of 2-Amino-2-methyl-1-propanol and Monoethanolamine at T = (313.15 to 393.15) K	https://www.doi.org/10.1021/je101234t
Density and Viscosity for Binary Mixtures of Ethylene Glycol Monobutyl Ether with Monoethanolamine, Diethanolamine, and Triethanolamine from (293.15 to 333.15) K	https://www.doi.org/10.1021/je049691v
Density and Viscosity of Monoethanolamine + Water + Carbon Dioxide from (25 to 80) Deg C.	https://www.doi.org/10.1021/acs.jced.5b00710
Effect of the Presence of MEA on the CO2 Capture Ability of Superbase Ionic Liquids.	https://www.cheric.org/research/kdb/hcprop/showprop.php?cmpid=1317
KDB Pure (Korean Thermophysical Properties Databank):	
Density, Viscosity, Refractive Index, and Electrical Conductivity of Aqueous Monoethanolamine Solutions in Equilibrium with Carbon Dioxide	https://www.doi.org/10.1021/acs.jced.7b01101
Degree of Monoethanolamine Degradation in Aqueous Solutions at Standard Temperatures.	https://www.doi.org/10.1021/je2010038
Molar Heat Capacity of Various Aqueous Alkanolamine Solutions from (298.15 to 423.15) K and Surface Tension of Pure Monoethanolamine (2) from (303.15 to 353.15) K.	https://www.doi.org/10.1021/je0604232
Volumetric properties of the monoethanolamine methanol mixture at atmospheric pressure from 283.15 to 353.15K.	https://www.doi.org/10.1016/j.tca.2004.11.015
Henry's law solubility of carbon dioxide in aqueous solutions of diethanolamine + piperazine.	https://www.doi.org/10.1016/j.jct.2013.05.005
Isothermal vapor-liquid equilibria of the aqueous solutions of simple alkanolamines. Influence of polarities (Monoethanolamine + Water) and (4-Methylmorpholine + Water) Binary Solutions at Several Temperatures.	https://www.doi.org/10.1016/j.tca.2015.02.020
Solubility of Ethane in Aqueous Solutions of Monoethanolamine and Diethanolamine.	https://www.doi.org/10.1021/je800530u
Vapor-Liquid Equilibrium for Binary Mixtures of 1,4-Diazabicyclo[2.2.2]octane with Ethylene Glycol, Diethanolamine, and Triethanolamine.	https://www.doi.org/10.1021/je060031v
Density, Speed of Sound, Viscosity, Surface Tension, and Excess Volume for Ethylene Glycol + Monoethanolamine + Water at 298.15 K	https://www.doi.org/10.1007/s10765-009-0570-x
Density, Speed of Sound, Isentropic Compressibility, and Excess Volume of Diethanolamine + Ethylene Glycol at T = (293.15 to 323.15) K	https://www.doi.org/10.1021/je301123j
Volumetric and viscometric properties of ternary solution (1-propanol), (Monoethanolamine + Methanol), and (Monoethanolamine + Methanol) at T = (293.15 to 323.15) K	https://www.doi.org/10.1021/je900437b
Experimental and model for the surface tension of MEA-PEG400 and MDEA-PEG400 aqueous solutions.	https://www.doi.org/10.1016/j.jct.2019.02.019
Volumetric properties of binary and ternary mixtures of N-methylpyrrolidone with monoethanolamine and diethanolamine at temperatures from 283.15 to 323.15 K.	https://www.doi.org/10.1016/j.jct.2013.10.016
Experimental and model for the surface tension of DEAE-PZ and DEAE-MEA aqueous solutions.	https://www.doi.org/10.1016/j.jct.2019.06.032
Experiment and model for the viscosities of MEA-PEG400, DEA-PEG400 and MEA-PEG400 aqueous solutions.	https://www.doi.org/10.1016/j.jct.2016.10.007
Volumetric properties of monoethanolamine and alcohol binary mixtures at different temperatures and 0.1 MPa.	https://www.doi.org/10.1016/j.jct.2014.06.017
Solubility of N2O and CO2 in aqueous systems of monoethanolamine and diethanolamine in the range 303.3-343K.	https://www.doi.org/10.1016/j.jct.2019.01.025
Effective extraction of benzene and thiophene by novel deep eutectic solvents: Experimental and COSMO-RS study.	https://www.doi.org/10.1016/j.fluid.2008.02.023
Vapor-liquid equilibrium measurement and ENRTL-Modeling of CO2 absorption in aqueous hexamethylenediamine.	https://www.doi.org/10.1016/j.jct.2019.05.017
Volumetric properties and CO2 solubility of monoethanolamine + diethylenetriamine + 2-(Diethylamino)ethanol mixtures: Experimental measurements and thermodynamic modeling.	https://www.doi.org/10.1016/j.fluid.2018.11.025
Densities, Refractive Indices and Excess Properties of Binary Mixtures of 1-Ethyl-3-methylimidazolium Tetrafluoroborate Ionic Liquid with Monoethanolamine and Diethanolamine at (293.15 to 313.15) K	https://www.doi.org/10.1016/j.fluid.2015.05.033
Viscosity, Density, and Volatility of Binary Mixtures of Imidazole and Monoethanolamine.	https://www.doi.org/10.1016/j.fluid.2017.06.018
Solubility of CO2 in aqueous mixtures of monoethanolamine and 2-(Ethylamino)ethanol.	https://www.doi.org/10.1016/j.jct.2016.07.041
Densities and Excess Properties of 1,2,4,5-Tetramethylimidazole with Primary Amines in Alcoholic Solutions:	https://www.doi.org/10.1021/je301077y
Experiment and model for the viscosity of carbonated MDEA - MEA aqueous solutions.	https://www.doi.org/10.1021/je2007204
Investigation of Thermodynamic Properties on CO2 Absorbents Blended with Aminoethyl Amino Acids and Monoethanolamine.	https://www.doi.org/10.1021/acs.jced.8b00674
Excess Enthalpy and Excess Volume for Pyridine + Methyldiethanolamine, + Ethanolamine, and + Diethanolamine Mixtures:	https://www.doi.org/10.1016/j.fluid.2013.12.020
Corrosion Inhibitors at 318.15, 338.15, and 353.15 K.	https://www.doi.org/10.1021/je3013205
Liquid−Liquid Equilibria for Benzene/Thiophene + Cyclohexane/Hexadecane + Deep Eutectic Solvents: Data and Correlation:	https://www.doi.org/10.1016/j.fluid.2012.01.029
	https://www.doi.org/10.1021/acs.jced.8b00175
	https://www.doi.org/10.1021/je400184t
	https://www.doi.org/10.1021/acs.jced.9b00313

CO2 solubility measurement and thermodynamic modeling for 1-methylpiperazine/water/CO2	https://www.doi.org/10.1016/j.fluid.2015.03.021
Monoethanolamine-2-methoxyethanol mixtures for CO2 capture: Density, viscosity and CO2 solubility.	https://www.doi.org/10.1016/j.jct.2018.12.028
Investigation of surface tension and viscosity for aqueous solutions of MEA-MeOH and DEA-MeOH.	https://www.doi.org/10.1016/j.jct.2017.08.024
Vapor-Liquid Equilibrium for Several Compounds Relevant to the Biofuels Industry Modeled with the Wilson Equation.	https://www.doi.org/10.1021/je400885z
CO2 solubility in aqueous potassium lysinate solutions at absorber conditions.	https://www.doi.org/10.1016/j.jct.2017.03.024
Liquid-liquid phase equilibrium for ternary mixtures of formamide (or ethylene glycol, or monoethanolamine) + water + 2-methylnaphthalene at 298.15 K.	https://www.doi.org/10.1016/j.fluid.2015.04.001
Characteristics of Potassium Prolinate + Water + Ethanol Solution as a Phase Splitting Absorbent for CO2 Capture: Equilibrium Total Pressure and CO2 Solubility in Binary and Ternary Aqueous Solutions.	https://www.doi.org/10.1021/acs.jced.7b00267
Vapour-liquid equilibrium study of tertiary amines, single and in blend 2-(Diethylamino)-ethanol (DEEA) and 3-Methylamino-propylamine (MAPA): Physicochemical Properties of the Aqueous Solutions for CO2 Capture.	https://www.doi.org/10.1021/je400886w
Vapour-liquid equilibrium study of tertiary amines, single and in blend 2-(Diethylamino)-ethanol (DEEA) and 3-Methylamino-propylamine (MAPA): Low-Temperature CO2 Capture.	https://www.doi.org/10.1016/j.jct.2019.06.017
Physicochemical Properties of the Aqueous Solutions of 2-(Diethylamino)-ethanol (DEEA) and 3-Methylamino-propylamine (MAPA) for Low-Temperature CO2 Capture.	https://www.doi.org/10.1021/je400830w
Volumetric properties of binary mixtures of dimethyl sulfoxide with amines from (293.15 to 363.15) K: Partial, Excess, and Surface Tensions of CO2 Loaded Aqueous Monoethanolamine Solutions with r = (0.2 to 0.7) at T = (303.15 to 333.15) K:	https://www.doi.org/10.1016/j.jct.2012.08.021
1-Formate as an Absorbent for CO2 Removal:	https://www.doi.org/10.1021/je301279x

Legend

affp:	Proton affinity
basg:	Gas basicity
cpg:	Ideal gas heat capacity
cpl:	Liquid phase heat capacity
dvisc:	Dynamic viscosity
gf:	Standard Gibbs free energy of formation
hf:	Enthalpy of formation at standard conditions
hfl:	Liquid phase enthalpy of formation at standard conditions
hfus:	Enthalpy of fusion at standard conditions
hvap:	Enthalpy of vaporization at standard conditions
hvapt:	Enthalpy of vaporization at a given temperature
ie:	Ionization energy
log10ws:	Log10 of Water solubility in mol/l
logp:	Octanol/Water partition coefficient
mcvol:	McGowan's characteristic volume
pc:	Critical Pressure
pvap:	Vapor pressure
rfi:	Refractive Index
rhol:	Liquid Density
rinpol:	Non-polar retention indices
ripol:	Polar retention indices
speedsl:	Speed of sound in fluid
srf:	Surface Tension
tb:	Normal Boiling Point Temperature
tc:	Critical Temperature
tf:	Normal melting (fusion) point

vc: Critical Volume