



hvap	60.40 ± 1.20		kJ/mol	NIST Webbook
hvap	60.40 ± 1.20		kJ/mol	NIST Webbook
ie	11.30		eV	NIST Webbook
ie	11.70		eV	NIST Webbook
ie	11.30		eV	NIST Webbook
log10ws	-9.71			Crippen Method
logp	8.566			Crippen Method
mcvol	237.710		ml/mol	McGowan Method
pc	886.30		kPa	Joback Method
tb	451.00		K	NIST Webbook
tb	452.20		K	NIST Webbook
tb	451.95 ± 1.00		K	NIST Webbook
tc	528.64		K	Joback Method
tf	302.44		K	Joback Method
vc	1.079		m <sup>3</sup> /kmol	Joback Method

## Temperature Dependent Properties

Property code	Value	Unit	Temperature [K]	Source
cpg	732.72	J/mol×K	495.07	Joback Method
cpg	744.93	J/mol×K	511.85	Joback Method
cpg	675.18	J/mol×K	427.93	Joback Method
cpg	690.94	J/mol×K	444.71	Joback Method
cpg	705.77	J/mol×K	461.50	Joback Method
cpg	719.68	J/mol×K	478.28	Joback Method
cpg	756.33	J/mol×K	528.64	Joback Method
hvapt	57.40	kJ/mol	374.00	NIST Webbook
hvapt	51.10	kJ/mol	457.50	NIST Webbook

## Correlations

Information	Value
Property code	pvap
Equation	$\ln(P_{vp}) = A + B/(T + C)$
Coeff. A	1.50242e+01
Coeff. B	-3.87093e+03
Coeff. C	-7.91550e+01
Temperature range (K), min.	341.83

## Sources

<b>Crippen Method:</b>	<a href="http://pubs.acs.org/doi/abs/10.1021/ci9903071">http://pubs.acs.org/doi/abs/10.1021/ci9903071</a>
<b>Crippen Method:</b>	<a href="https://www.chemeo.com/doc/models/crippen_log10ws">https://www.chemeo.com/doc/models/crippen_log10ws</a>
<b>Joback Method:</b>	<a href="https://en.wikipedia.org/wiki/Joback_method">https://en.wikipedia.org/wiki/Joback_method</a>
<b>McGowan Method:</b>	<a href="http://link.springer.com/article/10.1007/BF02311772">http://link.springer.com/article/10.1007/BF02311772</a>
<b>NIST Webbook:</b>	<a href="http://webbook.nist.gov/cgi/cbook.cgi?ID=C311897&amp;Units=SI">http://webbook.nist.gov/cgi/cbook.cgi?ID=C311897&amp;Units=SI</a>
<b>The Yaws Handbook of Vapor Pressure:</b>	<a href="https://www.sciencedirect.com/book/9780128029992/the-yaws-handbook-of-vapor-pressure">https://www.sciencedirect.com/book/9780128029992/the-yaws-handbook-of-vapor-pressure</a>

## Legend

<b>chl:</b>	Standard liquid enthalpy of combustion
<b>cp<sub>g</sub>:</b>	Ideal gas heat capacity
<b>gf:</b>	Standard Gibbs free energy of formation
<b>hf:</b>	Enthalpy of formation at standard conditions
<b>hfl:</b>	Liquid phase enthalpy of formation at standard conditions
<b>hfus:</b>	Enthalpy of fusion at standard conditions
<b>h<sub>vap</sub>:</b>	Enthalpy of vaporization at standard conditions
<b>h<sub>vapt</sub>:</b>	Enthalpy of vaporization at a given temperature
<b>ie:</b>	Ionization energy
<b>log<sub>10</sub>ws:</b>	Log <sub>10</sub> of Water solubility in mol/l
<b>log<sub>p</sub>:</b>	Octanol/Water partition coefficient
<b>mcvol:</b>	McGowan's characteristic volume
<b>pc:</b>	Critical Pressure
<b>p<sub>vap</sub>:</b>	Vapor pressure
<b>tb:</b>	Normal Boiling Point Temperature
<b>tc:</b>	Critical Temperature
<b>tf:</b>	Normal melting (fusion) point
<b>vc:</b>	Critical Volume

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