

# 3,5-Dimethylthiophenol

<b>Other names:</b>	3,5-dimethylbenzenethiol Benzenethiol, 3,5-dimethyl-
<b>Inchi:</b>	InChI=1S/C8H10S/c1-6-3-7(2)5-8(9)4-6/h3-5,9H,1-2H3
<b>InchiKey:</b>	CESBAYSBPMAEI-UHFFFAOYSA-N
<b>Formula:</b>	C8H10S
<b>SMILES:</b>	<chem>Cc1cc(C)cc(S)c1</chem>
<b>Mol. weight [g/mol]:</b>	138.23
<b>CAS:</b>	38360-81-5

## Physical Properties

Property code	Value	Unit	Source
gf	139.02	kJ/mol	Joback Method
hf	43.62	kJ/mol	Joback Method
hfus	13.78	kJ/mol	Joback Method
hvap	43.74	kJ/mol	Joback Method
log10ws	-2.94		Crippen Method
logp	2.592		Crippen Method
mvol	116.170	ml/mol	McGowan Method
pc	3805.69	kPa	Joback Method
tb	481.94	K	Joback Method
tc	719.76	K	Joback Method
tf	267.84	K	Joback Method
vc	0.429	m <sup>3</sup> /kmol	Joback Method

## Temperature Dependent Properties

Property code	Value	Unit	Temperature [K]	Source
cpg	221.55	J/molxK	481.94	Joback Method
cpg	233.86	J/molxK	521.58	Joback Method
cpg	245.45	J/molxK	561.21	Joback Method
cpg	256.34	J/molxK	600.85	Joback Method
cpg	266.57	J/molxK	640.49	Joback Method
cpg	276.15	J/molxK	680.13	Joback Method
cpg	285.11	J/molxK	719.76	Joback Method

# Correlations

Information	Value
Property code	pvap
Equation	$\ln(P_{vp}) = A + B/(T + C)$
Coeff. A	1.47964e+01
Coeff. B	-4.16901e+03
Coeff. C	-7.72170e+01
Temperature range (K), min.	364.56
Temperature range (K), max.	516.76

## Sources

<b>NIST Webbook:</b>	<a href="http://webbook.nist.gov/cgi/cbook.cgi?ID=C38360815&amp;Units=SI">http://webbook.nist.gov/cgi/cbook.cgi?ID=C38360815&amp;Units=SI</a>
<b>The Yaws Handbook of Vapor Pressure: Crippen Method:</b>	<a href="https://www.sciencedirect.com/book/9780128029992/the-yaws-handbook-of-vapor-pressure">https://www.sciencedirect.com/book/9780128029992/the-yaws-handbook-of-vapor-pressure</a> <a href="http://pubs.acs.org/doi/abs/10.1021/ci9903071">http://pubs.acs.org/doi/abs/10.1021/ci9903071</a>
<b>Crippen Method:</b>	<a href="https://www.chemeo.com/doc/models/crippen_log10ws">https://www.chemeo.com/doc/models/crippen_log10ws</a>
<b>Joback Method:</b>	<a href="https://en.wikipedia.org/wiki/Joback_method">https://en.wikipedia.org/wiki/Joback_method</a>
<b>McGowan Method:</b>	<a href="http://link.springer.com/article/10.1007/BF02311772">http://link.springer.com/article/10.1007/BF02311772</a>

## Legend

<b>cp<sub>g</sub>:</b>	Ideal gas heat capacity
<b>g<sub>f</sub>:</b>	Standard Gibbs free energy of formation
<b>h<sub>f</sub>:</b>	Enthalpy of formation at standard conditions
<b>h<sub>fus</sub>:</b>	Enthalpy of fusion at standard conditions
<b>h<sub>vap</sub>:</b>	Enthalpy of vaporization at standard conditions
<b>log<sub>10</sub>ws:</b>	Log <sub>10</sub> of Water solubility in mol/l
<b>log<sub>p</sub>:</b>	Octanol/Water partition coefficient
<b>mc<sub>vol</sub>:</b>	McGowan's characteristic volume
<b>pc:</b>	Critical Pressure
<b>pvap:</b>	Vapor pressure
<b>tb:</b>	Normal Boiling Point Temperature
<b>tc:</b>	Critical Temperature

**tf:** Normal melting (fusion) point

**vc:** Critical Volume

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