

# 3,3-dimethylcyclohexanone

<b>Inchi:</b>	InChI=1S/C8H14O/c1-8(2)5-3-4-7(9)6-8/h3-6H2,1-2H3
<b>InchiKey:</b>	ZVJQBBAVPAFLX-UHFFFAOYSA-N
<b>Formula:</b>	C8H14O
<b>SMILES:</b>	CC1(C)CCCC(=O)C1
<b>Mol. weight [g/mol]:</b>	126.20
<b>CAS:</b>	2979-19-3

## Physical Properties

Property code	Value	Unit	Source
gf	-87.15	kJ/mol	Joback Method
hf	-276.59	kJ/mol	Joback Method
hfus	1.52	kJ/mol	Joback Method
hvap	36.93	kJ/mol	Joback Method
log10ws	-2.10		Crippen Method
logp	2.156		Crippen Method
mvol	114.290	ml/mol	McGowan Method
pc	3456.14	kPa	Joback Method
rinpol	1036.00		NIST Webbook
tb	470.05	K	Joback Method
tc	700.41	K	Joback Method
tf	279.42	K	Joback Method
vc	0.421	m <sup>3</sup> /kmol	Joback Method

## Temperature Dependent Properties

Property code	Value	Unit	Temperature [K]	Source
cpg	247.36	J/mol×K	470.05	Joback Method
cpg	264.46	J/mol×K	508.44	Joback Method
cpg	280.52	J/mol×K	546.84	Joback Method
cpg	295.62	J/mol×K	585.23	Joback Method
cpg	309.86	J/mol×K	623.62	Joback Method
cpg	323.33	J/mol×K	662.01	Joback Method
cpg	336.13	J/mol×K	700.41	Joback Method

# Correlations

Information	Value
Property code	pvap
Equation	$\ln(P_{vp}) = A + B/(T + C)$
Coeff. A	1.49459e+01
Coeff. B	-3.92833e+03
Coeff. C	-6.71260e+01
Temperature range (K), min.	335.12
Temperature range (K), max.	474.87

## Sources

<b>NIST Webbook:</b>	<a href="http://webbook.nist.gov/cgi/cbook.cgi?ID=C2979193&amp;Units=SI">http://webbook.nist.gov/cgi/cbook.cgi?ID=C2979193&amp;Units=SI</a>
<b>The Yaws Handbook of Vapor Pressure:</b>	<a href="https://www.sciencedirect.com/book/9780128029992/the-yaws-handbook-of-vapor-pressure">https://www.sciencedirect.com/book/9780128029992/the-yaws-handbook-of-vapor-pressure</a>
<b>Crippen Method:</b>	<a href="http://pubs.acs.org/doi/abs/10.1021/ci9903071">http://pubs.acs.org/doi/abs/10.1021/ci9903071</a>
<b>Crippen Method:</b>	<a href="https://www.chemeo.com/doc/models/crippen_log10ws">https://www.chemeo.com/doc/models/crippen_log10ws</a>
<b>Joback Method:</b>	<a href="https://en.wikipedia.org/wiki/Joback_method">https://en.wikipedia.org/wiki/Joback_method</a>
<b>McGowan Method:</b>	<a href="http://link.springer.com/article/10.1007/BF02311772">http://link.springer.com/article/10.1007/BF02311772</a>

## Legend

<b>cp<sub>g</sub>:</b>	Ideal gas heat capacity
<b>gf:</b>	Standard Gibbs free energy of formation
<b>hf:</b>	Enthalpy of formation at standard conditions
<b>hfus:</b>	Enthalpy of fusion at standard conditions
<b>hvap:</b>	Enthalpy of vaporization at standard conditions
<b>log<sub>10</sub>ws:</b>	Log <sub>10</sub> of Water solubility in mol/l
<b>log<sub>p</sub>:</b>	Octanol/Water partition coefficient
<b>mcvol:</b>	McGowan's characteristic volume
<b>pc:</b>	Critical Pressure
<b>pvap:</b>	Vapor pressure
<b>rinpol:</b>	Non-polar retention indices
<b>tb:</b>	Normal Boiling Point Temperature

**tc:** Critical Temperature  
**tf:** Normal melting (fusion) point  
**vc:** Critical Volume

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