

# 1,3-Butylene glycol dimethacrylate

<b>Other names:</b>	1,3-Butanediol, bis(2-methylpropenoate) 1,3-butanediol dimethacrylate 1,3-butanediyl dimethacrylate 1-methyltrimethylene dimethacrylate 2-Propenoic acid, 2-methyl-, 1-methyl-1,3-propanediyl ester 2-methyl-2-propenoic acid 1-methyl-1,3-propanediyl ester butane-1,3-diyl bis(2-methylacrylate)
<b>Inchi:</b>	InChI=1S/C12H18O4/c1-8(2)11(13)15-7-6-10(5)16-12(14)9(3)4/h10H,1,3,6-7H2,2,4-5H3
<b>InchiKey:</b>	VDYWHVQKENANGY-UHFFFAOYSA-N
<b>Formula:</b>	C12H18O4
<b>SMILES:</b>	<chem>C=C(C)C(=O)OCCC(C)OC(=O)C(=C)C</chem>
<b>Mol. weight [g/mol]:</b>	226.27
<b>CAS:</b>	1189-08-8

## Physical Properties

Property code	Value	Unit	Source
gf	-261.54	kJ/mol	Joback Method
hf	-554.61	kJ/mol	Joback Method
hfus	23.71	kJ/mol	Joback Method
hvap	59.05	kJ/mol	Joback Method
log10ws	-2.39		Crippen Method
logp	2.004		Crippen Method
mcvol	186.220	ml/mol	McGowan Method
pc	2111.94	kPa	Joback Method
tb	619.22	K	Joback Method
tc	810.79	K	Joback Method
tf	322.88	K	Joback Method
vc	0.714	m <sup>3</sup> /kmol	Joback Method

## Temperature Dependent Properties

Property code	Value	Unit	Temperature [K]	Source
cpg	472.43	J/mol×K	619.22	Joback Method
cpg	486.48	J/mol×K	651.15	Joback Method

cpg	499.84	J/mol×K	683.08	Joback Method
cpg	512.51	J/mol×K	715.01	Joback Method
cpg	524.50	J/mol×K	746.93	Joback Method
cpg	535.82	J/mol×K	778.86	Joback Method
cpg	546.47	J/mol×K	810.79	Joback Method

## Sources

**NIST Webbook:** <http://webbook.nist.gov/cgi/cbook.cgi?ID=C1189088&Units=SI>

**Crippen Method:** <http://pubs.acs.org/doi/abs/10.1021/ci9903071>

**Crippen Method:** [https://www.chemeo.com/doc/models/crippen\\_log10ws](https://www.chemeo.com/doc/models/crippen_log10ws)

**Phase behaviour for the (carbon dioxide + 1,3-butanediol diacrylate) and (carbon dioxide + 1,3-butanediol dimethacrylate) systems at elevated pressures and temperatures:** <https://www.doi.org/10.1016/j.jct.2013.11.030>

**Joback Method:** [https://en.wikipedia.org/wiki/Joback\\_method](https://en.wikipedia.org/wiki/Joback_method)

**McGowan Method:** <http://link.springer.com/article/10.1007/BF02311772>

## Legend

<b>cpg:</b>	Ideal gas heat capacity
<b>gf:</b>	Standard Gibbs free energy of formation
<b>hf:</b>	Enthalpy of formation at standard conditions
<b>hfus:</b>	Enthalpy of fusion at standard conditions
<b>hvap:</b>	Enthalpy of vaporization at standard conditions
<b>log10ws:</b>	Log10 of Water solubility in mol/l
<b>logp:</b>	Octanol/Water partition coefficient
<b>mcvol:</b>	McGowan's characteristic volume
<b>pc:</b>	Critical Pressure
<b>tb:</b>	Normal Boiling Point Temperature
<b>tc:</b>	Critical Temperature
<b>tf:</b>	Normal melting (fusion) point
<b>vc:</b>	Critical Volume

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