

# 2-Propenoic acid, ethenyl ester

<b>Other names:</b>	Acrylic acid, vinyl ester ethenyl 2-propenoate ethenyl acrylate vinyl 2-propenoate vinyl acrylate
<b>Inchi:</b>	InChI=1S/C5H6O2/c1-3-5(6)7-4-2/h3-4H,1-2H2
<b>InchiKey:</b>	BLCTWBJQROOONQ-UHFFFAOYSA-N
<b>Formula:</b>	C5H6O2
<b>SMILES:</b>	C=COC(=O)C=C
<b>Mol. weight [g/mol]:</b>	98.10
<b>CAS:</b>	2177-18-6

## Physical Properties

Property code	Value	Unit	Source
gf	-67.02	kJ/mol	Joback Method
hf	-140.47	kJ/mol	Joback Method
hfus	8.93	kJ/mol	Joback Method
hvap	34.54	kJ/mol	Joback Method
log10ws	-0.98		Crippen Method
logp	0.859		Crippen Method
mcvol	80.150	ml/mol	McGowan Method
pc	4067.32	kPa	Joback Method
tb	383.45	K	Joback Method
tc	569.81	K	Joback Method
tf	191.20 ± 0.60	K	NIST Webbook
vc	0.301	m3/kmol	Joback Method

## Temperature Dependent Properties

Property code	Value	Unit	Temperature [K]	Source
cpg	135.44	J/mol×K	383.45	Joback Method
cpg	167.81	J/mol×K	538.75	Joback Method
cpg	161.87	J/mol×K	507.69	Joback Method
cpg	155.66	J/mol×K	476.63	Joback Method

cpg	149.19	J/molxK	445.57	Joback Method
cpg	142.45	J/molxK	414.51	Joback Method
cpg	173.50	J/molxK	569.81	Joback Method
dvisc	0.0002597	Paxs	383.45	Joback Method
dvisc	0.0003205	Paxs	355.33	Joback Method
dvisc	0.0004100	Paxs	327.22	Joback Method
dvisc	0.0005495	Paxs	299.10	Joback Method
dvisc	0.0007826	Paxs	270.98	Joback Method
dvisc	0.0012096	Paxs	242.87	Joback Method
dvisc	0.0020954	Paxs	214.75	Joback Method

## Sources

<b>Crippen Method:</b>	<a href="http://pubs.acs.org/doi/abs/10.1021/ci990307I">http://pubs.acs.org/doi/abs/10.1021/ci990307I</a>
<b>Crippen Method:</b>	<a href="https://www.chemeo.com/doc/models/crippen_log10ws">https://www.chemeo.com/doc/models/crippen_log10ws</a>
<b>Bubble-Point Measurement for CO2 + Vinyl Acetate and CO2 + Vinyl Acrylate Systems at High Pressures.:</b>	<a href="https://www.doi.org/10.1021/je0201121">https://www.doi.org/10.1021/je0201121</a>
<b>Joback Method</b>	<a href="https://en.wikipedia.org/wiki/Joback_method">https://en.wikipedia.org/wiki/Joback_method</a>
<b>McGowan Method:</b>	<a href="http://link.springer.com/article/10.1007/BF02311772">http://link.springer.com/article/10.1007/BF02311772</a>
<b>NIST Webbook:</b>	<a href="http://webbook.nist.gov/cgi/cbook.cgi?ID=C2177186&amp;Units=SI">http://webbook.nist.gov/cgi/cbook.cgi?ID=C2177186&amp;Units=SI</a>

## Legend

<b>cpg:</b>	Ideal gas heat capacity
<b>dvisc:</b>	Dynamic viscosity
<b>gf:</b>	Standard Gibbs free energy of formation
<b>hf:</b>	Enthalpy of formation at standard conditions
<b>hfus:</b>	Enthalpy of fusion at standard conditions
<b>hvap:</b>	Enthalpy of vaporization at standard conditions
<b>log10ws:</b>	Log10 of Water solubility in mol/l
<b>logp:</b>	Octanol/Water partition coefficient
<b>mcvol:</b>	McGowan's characteristic volume
<b>pc:</b>	Critical Pressure
<b>tb:</b>	Normal Boiling Point Temperature
<b>tc:</b>	Critical Temperature
<b>tf:</b>	Normal melting (fusion) point
<b>vc:</b>	Critical Volume

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