Isopropyl Alcohol

Other names:	1-Methylethyl Alcohol 1-methylethanol 2-Hydroxypropane 2-PROPANOL 2-Propyl alcohol Alcojel Alcolo Alcool isopropilico Alcool isopropylique Alcosolve Alcosolve 2 Alkolave Arquad DMCB Avantin Avantine Combi-Schutz Dimethylcarbinol Hartosol IPA IPS 1 Imsol A Isohol Isopropanol Isopropyl alcohol, rubbing Lavacol Lutosol PRO Petrohol Propan-2-ol Propane, 2-hydroxy- Propanol-2 Propol SEC-PROPYL ALCOHOL Spectrar Sterisol hand disinfectant Takineocol UN 1219 Virahol Visco 1152 i-Propanol

	i-Propylalkohol
	iso-C3H7OH
	iso-Propylalkohol
	n-Propan-2-ol
	sec-Propanol
Inchi:	InChI=1S/C3H8O/c1-3(2)4/h3-4H,1-2H3
InchiKey:	KFZMGEQAYNKOFK-UHFFFAOYSA-N
Formula:	C3H8O
SMILES:	CC(C)O
Mol. weight [g/mol]:	60.09
CAS:	67-63-0

Physical Properties

Property code	Value	Unit	Source
af	0.6650		KDB
affp	793.00	kJ/mol	NIST Webbook
affp	796.00 ± 6.00	kJ/mol	NIST Webbook
aigt	672.04	K	KDB
basg	762.60	kJ/mol	NIST Webbook
chl	-2005.80 ± 0.40	kJ/mol	NIST Webbook
chl	-2005.10	kJ/mol	NIST Webbook
chl	-2006.90 ± 0.20	kJ/mol	NIST Webbook
dm	1.70	debye	KDB
dvisc	0.0020980	Pa×s	Speeds of sound, isentropic compressibilities, viscosities and excess molar volumes of binary mixtures of methylcyclohexane + 2-alkanols or ethanol at T = 298.15 K
dvisc	0.0020630	Pa×s	Thermodynamic properties of binary liquid mixtures of diethylenetriamine with alcohols at different temperatures
dvisc	0.0020443	Pa×s	Densities and Viscosities of Binary Liquid Mixtures of Trichloroethylene and Tetrachloroethylene with Some Polar and Nonpolar Solvents
fll	2.30	% in Air	KDB
flu	12.70	% in Air	KDB
fpc	291.48	K	KDB

fpo	284.82	K	KDB
gf	-173.50	kJ/mol	KDB
gyrad	2.7260		KDB
hf	-272.80	kJ/mol	NIST Webbook
hf	-271.10	kJ/mol	NIST Webbook
hf	-272.30 ± 0.92	kJ/mol	NIST Webbook
hf	-272.80	kJ/mol	NIST Webbook
hf	-272.60	kJ/mol	KDB
hfl	-317.00 ± 0.30	kJ/mol	NIST Webbook
hfl	-318.70	kJ/mol	NIST Webbook
hfl	-318.20 ± 0.71	kJ/mol	NIST Webbook
hfus	4.09	kJ/mol	Joback Method
hvap	45.51	kJ/mol	NIST Webbook
hvap	44.00	kJ/mol	NIST Webbook
hvap	45.90 ± 0.20	kJ/mol	NIST Webbook
hvap	44.40	kJ/mol	NIST Webbook
hvap	45.48	kJ/mol	NIST Webbook
hvap	45.20 ± 0.10	kJ/mol	NIST Webbook
hvap	45.34 ± 0.02	kJ/mol	NIST Webbook
hvap	45.34 ± 0.02	kJ/mol	NIST Webbook
hvap	44.00	kJ/mol	NIST Webbook
hvap	45.20 ± 0.10	kJ/mol	NIST Webbook
hvap	40.00 ± 0.04	kJ/mol	NIST Webbook
ie	10.10 ± 0.02	eV	NIST Webbook
ie	10.15 ± 0.05	eV	NIST Webbook
ie	10.44	eV	NIST Webbook
ie	10.15 ± 0.07	eV	NIST Webbook
ie	10.17 ± 0.02	eV	NIST Webbook
ie	10.42	eV	NIST Webbook
ie	10.49 ± 0.03	eV	NIST Webbook
ie	10.42	eV	NIST Webbook
ie	10.36	eV	NIST Webbook
ie	10.29 ± 0.02	eV	NIST Webbook
ie	10.12 ± 0.03	eV	NIST Webbook
ie	10.18	eV	NIST Webbook
log10ws	0.43		Aqueous Solubility Prediction Method
log10ws	0.43		Estimated Solubility Method
logp	0.387		Crippen Method
mcvol	59.000	ml/mol	McGowan Method
nfpaf	%!d(float64=3)		KDB
nfpah	%!d(float64=1)		KDB
pc	4760.00 ± 20.00	kPa	NIST Webbook
pc	4764.00	kPa	KDB

pc	4761.00 ± 20.00	kPa	NIST Webbook
pc	4755.00	kPa	Isothermal Vapor-Liquid Equilibria for the 2-Propanol + n-Heptane System near the Critical Region
pc	5380.00	kPa	NIST Webbook
pc	5370.23 ± 101.32	kPa	NIST Webbook
pc	4764.00	kPa	NIST Webbook
pc	4762.00 ± 4.75	kPa	NIST Webbook
pc	4716.00 ± 8.00	kPa	NIST Webbook
pc	4762.00	kPa	NIST Webbook
pc	4761.00 ± 20.00	kPa	NIST Webbook
pc	4770.00	kPa	NIST Webbook
rhoc	271.03 ± 1.20	kg/m3	NIST Webbook
rhoc	272.83	kg/m3	NIST Webbook
rhoc	272.71	kg/m3	NIST Webbook
rinpol	508.00		NIST Webbook
rinpol	490.00		NIST Webbook
rinpol	508.00		NIST Webbook
rinpol	491.00		NIST Webbook
rinpol	453.00		NIST Webbook
rinpol	480.00		NIST Webbook
rinpol	491.00		NIST Webbook
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rinpol	508.00		NIST Webbook
rinpol	491.00		NIST Webbook
rinpol	453.00		NIST Webbook
rinpol	481.00		NIST Webbook
rinpol	508.00		NIST Webbook
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ripol	884.00	NIST Webbook
ripol	926.00	NIST Webbook
ripol	925.00	NIST Webbook
ripol	884.00	NIST Webbook
ripol	923.00	NIST Webbook
ripol	948.00	NIST Webbook
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ripol	935.00		NIST Webbook
ripol	963.00		NIST Webbook
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ripol	935.00		NIST Webbook
ripol	903.00		NIST Webbook
sl	190.80	J/mol×K	NIST Webbook
sl	192.90	J/mol×K	NIST Webbook
sl	179.90	J/mol×K	NIST Webbook
sl	180.58	J/mol×K	NIST Webbook
tb	356.00 ± 0.50	K	NIST Webbook
tb	355.65 ± 1.50	K	NIST Webbook
tb	356.00 ± 0.50	K	NIST Webbook
tb	356.15 ± 1.00	K	NIST Webbook

tb	355.19 ± 0.20	K	NIST Webbook
tb	355.59 ± 0.20	K	NIST Webbook
tb	355.15 ± 1.50	K	NIST Webbook
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tb	355.75 ± 0.30	K	NIST Webbook
tb	351.65 ± 3.00	K	NIST Webbook
tb	355.55 ± 0.20	K	NIST Webbook
tb	355.41 ± 0.05	K	NIST Webbook
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tb	355.50 ± 0.30	K	NIST Webbook
tb	355.55 ± 0.30	K	NIST Webbook
tb	353.85 ± 0.50	K	NIST Webbook
tb	355.60 ± 0.20	K	NIST Webbook
tb	355.15 ± 0.20	K	NIST Webbook
tb	355.60 ± 0.50	K	NIST Webbook
tb	355.60 ± 0.50	K	NIST Webbook
tb	82.21 ± 0.20	K	NIST Webbook
tb	355.40	K	KDB
tb	355.37	K	Physical properties and phase equilibria of the system isopropyl acetate + isopropanol + 1-octyl-3-methyl-imidazolium bis(trifluoromethylsulfonyl)imide
tb	355.15	K	Multiphase equilibria for mixtures containing water, isopropanol, propionic acid, and isopropyl propionate
tb	355.53	K	Isobaric VLE at 0.6 MPa for binary systems isobutyl acetate + ethanol, + 1-propanol or + 2-propanol
tb	355.35	K	Isobaric vapor-liquid equilibrium data for the binary system methyl acetate + isopropyl acetate and the quaternary system methyl acetate + methanol + isopropanol + isopropyl acetate at 101.3 kPa

tb	355.60	K	Isobaric vapor-liquid equilibria of the binary mixtures propylene glycol methyl ether + propylene glycol methyl ether acetate, methyl acetate + propylene glycol methyl ether and methanol + propylene glycol methyl ether acetate at 101.3 kPa
tb	355.40	K	Experimental studies and thermodynamic analysis of isobaric vapor-liquid-liquid equilibria of 2-propanol + water system using n-propyl acetate and isopropyl acetate as entrainers
tb	355.61	K	VLE of the binary systems (dimethyl carbonate with 2-propanol or 2-butanol) and (diethyl carbonate with methylcyclohexane) at 101.3 kPa
tb	355.45	K	Isobaric (vapour + liquid) equilibrium for (2-propanol + water + ammonium thiocyanate): Fitting the data by an empirical equation
tb	355.25	K	Isobaric (vapour + liquid + liquid) equilibrium data for (di-n-propyl ether + n-propyl alcohol + water) and (diisopropyl ether + isopropyl alcohol + water) systems at 100 kPa
tb	355.27	K	Vapour-liquid equilibrium and extractive distillation for separation of azeotrope isopropyl alcohol and diisopropyl ether
tb	355.33	K	Separation of the mixture (isopropyl alcohol + diisopropyl ether + n-propanol): Entrainer selection, interaction exploration and vapour-liquid equilibrium measurements
tb	355.71	K	A new analysis method for improving collection of vapor-liquid equilibrium (VLE) data of binary mixtures using differential scanning calorimetry (DSC)
tb	355.33	K	Vapor Liquid Equilibria for Ternary Mixtures of Isopropyl Alcohol, Isopropyl Acetate, and DMSO at 101.3 kPa

tb	355.60	K	Acetonitrile Dehydration via Extractive Distillation Using Low Transition Temperature Mixtures as Entrainers
tb	355.37	K	Vapor Liquid Equilibria for Ethanol + 2,4,4-Trimethyl-1-pentene and 2-Propanol + 2,4,4-Trimethyl-1-pentene at 101 kPa
tb	355.35	K	Measurement and Correlation of Excess Molar Enthalpies for Ethylene Glycol + Alkanol Systems at the Temperatures (298.15, 308.15, and 323.15) K
tb	355.15	K	Vapor-Liquid Equilibria for the Binary Systems of Dimethoxymethane with Some Fuel Oxygenates
tb	355.41	K	Vapor-Liquid Equilibrium Behavior of Tolan in Alcohol
tb	355.39	K	Vapor-Liquid Equilibrium Behaviors of Coumarin and Vanillin in Ethanol, 1-Propanol, and 2-Propanol
tb	355.35	K	Isobaric Vapor-Liquid Equilibrium for Binary Systems of Toluene + Ethanol, Toluene + Isopropanol at (101.3, 121.3, 161.3 and 201.3) kPa
tb	355.40	K	Vapor-Liquid Equilibrium Behaviors of 3-Ethoxy-4-hydroxybenzaldehyde in Alcohol
tb	355.60 ± 0.50	K	NIST Webbook
tb	355.40	K	Vapor Liquid Equilibrium Behaviors of 5-Methyl-2-(1-methylethyl)phenol in Alcohol
tb	355.95 ± 1.00	K	NIST Webbook
tb	355.45 ± 1.00	K	NIST Webbook
tb	355.55 ± 0.50	K	NIST Webbook
tb	355.50 ± 0.20	K	NIST Webbook
tb	355.46 ± 0.20	K	NIST Webbook
tb	355.55 ± 0.10	K	NIST Webbook
tb	355.44 ± 0.10	K	NIST Webbook
tb	355.54 ± 0.10	K	NIST Webbook

tb	355.65	K	Isobaric Vapor-Liquid Equilibrium for (Propan-2-ol + Water + 1-Butyl-3-methylimidazolium Tetrafluoroborate)
tb	355.42 ± 0.20	K	NIST Webbook
tb	355.42 ± 0.30	K	NIST Webbook
tb	355.42 ± 0.07	K	NIST Webbook
tb	355.65 ± 0.20	K	NIST Webbook
tb	355.45 ± 0.50	K	NIST Webbook
tb	354.55 ± 0.50	K	NIST Webbook
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tb	355.60 ± 0.25	K	NIST Webbook
tb	355.55 ± 0.50	K	NIST Webbook
tb	355.45 ± 0.10	K	NIST Webbook
tb	354.70 ± 1.00	K	NIST Webbook
tb	3.50 ± 2.00	K	NIST Webbook
tb	355.55 ± 0.50	K	NIST Webbook
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tc	508.30 ± 0.30	K	NIST Webbook
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tc	508.00 ± 0.60	K	NIST Webbook

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tc	508.00	K	NIST Webbook
tc	507.36 ± 0.20	K	NIST Webbook
tc	508.30	K	NIST Webbook
tc	508.30 ± 0.05	K	NIST Webbook
tc	508.30 ± 1.00	K	NIST Webbook
tc	508.30 ± 0.20	K	NIST Webbook
tc	508.30	K	KDB
tc	507.80	K	NIST Webbook
tc	511.20 ± 3.00	K	NIST Webbook
tc	516.60	K	NIST Webbook
tc	508.80	K	NIST Webbook
tc	508.40 ± 0.30	K	NIST Webbook
tc	508.15 ± 0.50	K	NIST Webbook
tc	511.50	K	NIST Webbook
tc	508.26	K	NIST Webbook
tc	508.30 ± 0.30	K	NIST Webbook
tc	508.72	K	One- and two-phase isochoric heat capacities and saturated densities of 2-propanol in the critical and supercritical regions
tc	508.30 ± 0.30	K	NIST Webbook
tf	183.60	K	KDB
tf	183.90	K	Aqueous Solubility Prediction Method
tf	184.65	K	NIST Webbook
tf	185.35	K	NIST Webbook
tf	185.75 ± 0.50	K	NIST Webbook
tt	185.26 ± 0.05	K	NIST Webbook
tt	185.25 ± 0.06	K	NIST Webbook
tt	185.20 ± 0.05	K	NIST Webbook
tt	184.67 ± 0.10	K	NIST Webbook
tt	184.60 ± 0.30	K	NIST Webbook
tt	184.60 ± 0.15	K	NIST Webbook
vc	0.223 ± 0.003	m3/kmol	NIST Webbook
vc	0.222	m3/kmol	NIST Webbook
vc	0.222	m3/kmol	KDB
zc	0.2502460		KDB
zra	0.25		KDB

Temperature Dependent Properties

Property code	Value	Unit	Temperature [K]	Source
cpg	110.80 ± 1.60	J/mol×K	393.65	NIST Webbook
cpg	109.20 ± 1.60	J/mol×K	384.95	NIST Webbook
cpg	148.10 ± 1.60	J/mol×K	597.25	NIST Webbook
cpg	142.60 ± 1.60	J/mol×K	567.05	NIST Webbook
cpg	137.50 ± 1.60	J/mol×K	539.05	NIST Webbook
cpg	132.90 ± 1.60	J/mol×K	513.95	NIST Webbook
cpg	130.30 ± 1.60	J/mol×K	499.75	NIST Webbook
cpg	127.01	J/mol×K	473.15	NIST Webbook
cpg	124.20 ± 1.60	J/mol×K	466.75	NIST Webbook
cpg	121.70 ± 1.60	J/mol×K	453.15	NIST Webbook
cpg	122.80	J/mol×K	451.15	NIST Webbook
cpg	108.10 ± 1.60	J/mol×K	378.85	NIST Webbook
cpg	118.70	J/mol×K	431.15	NIST Webbook
cpg	117.02	J/mol×K	423.15	NIST Webbook
cpg	114.35	J/mol×K	411.15	NIST Webbook
cpg	113.00 ± 1.60	J/mol×K	405.35	NIST Webbook
cpg	106.29	J/mol×K	373.15	NIST Webbook
cpg	105.77	J/mol×K	371.15	NIST Webbook
cpg	105.70 ± 1.60	J/mol×K	365.75	NIST Webbook
cpg	103.06	J/mol×K	358.72	NIST Webbook
cpg	110.08	J/mol×K	391.15	NIST Webbook
cpg	122.10	J/mol×K	448.15	NIST Webbook
cpg	126.70 ± 1.60	J/mol×K	480.55	NIST Webbook
cpg	111.65	J/mol×K	398.15	NIST Webbook
cpl	173.13	J/mol×K	521.87	Influence of nanofluid instability on thermodynamic properties near the critical point
cpl	180.16	J/mol×K	509.45	Influence of nanofluid instability on thermodynamic properties near the critical point
cpl	178.84	J/mol×K	509.55	Influence of nanofluid instability on thermodynamic properties near the critical point
cpl	179.98	J/mol×K	509.64	Influence of nanofluid instability on thermodynamic properties near the critical point

cpl	177.46	J/mol×K	509.79	Influence of nanofluid instability on thermodynamic properties near the critical point
cpl	177.28	J/mol×K	509.99	Influence of nanofluid instability on thermodynamic properties near the critical point
cpl	178.12	J/mol×K	510.14	Influence of nanofluid instability on thermodynamic properties near the critical point
cpl	177.94	J/mol×K	510.23	Influence of nanofluid instability on thermodynamic properties near the critical point
cpl	176.68	J/mol×K	510.33	Influence of nanofluid instability on thermodynamic properties near the critical point
cpl	178.60	J/mol×K	510.48	Influence of nanofluid instability on thermodynamic properties near the critical point
cpl	176.74	J/mol×K	516.90	Influence of nanofluid instability on thermodynamic properties near the critical point
cpl	176.80	J/mol×K	516.99	Influence of nanofluid instability on thermodynamic properties near the critical point
cpl	176.80	J/mol×K	517.08	Influence of nanofluid instability on thermodynamic properties near the critical point
cpl	172.17	J/mol×K	517.27	Influence of nanofluid instability on thermodynamic properties near the critical point

cpl	179.56	J/mol×K	509.35	Influence of nanofluid instability on thermodynamic properties near the critical point
cpl	168.57	J/mol×K	521.96	Influence of nanofluid instability on thermodynamic properties near the critical point
cpl	186.90	J/mol×K	508.76	Influence of nanofluid instability on thermodynamic properties near the critical point
cpl	168.63	J/mol×K	522.33	Influence of nanofluid instability on thermodynamic properties near the critical point
cpl	171.63	J/mol×K	522.42	Influence of nanofluid instability on thermodynamic properties near the critical point
cpl	173.85	J/mol×K	522.51	Influence of nanofluid instability on thermodynamic properties near the critical point
cpl	192.91	J/mol×K	508.67	Influence of nanofluid instability on thermodynamic properties near the critical point
cpl	168.39	J/mol×K	530.01	Influence of nanofluid instability on thermodynamic properties near the critical point
cpl	168.27	J/mol×K	532.04	Influence of nanofluid instability on thermodynamic properties near the critical point
cpl	168.33	J/mol×K	535.11	Influence of nanofluid instability on thermodynamic properties near the critical point

cpl	168.39	J/mol×K	540.06	Influence of nanofluid instability on thermodynamic properties near the critical point
cpl	168.63	J/mol×K	545.13	Influence of nanofluid instability on thermodynamic properties near the critical point
cpl	169.77	J/mol×K	560.04	Influence of nanofluid instability on thermodynamic properties near the critical point
cpl	173.85	J/mol×K	570.09	Influence of nanofluid instability on thermodynamic properties near the critical point
cpl	174.22	J/mol×K	580.16	Influence of nanofluid instability on thermodynamic properties near the critical point
cpl	176.38	J/mol×K	590.07	Influence of nanofluid instability on thermodynamic properties near the critical point
cpl	178.60	J/mol×K	600.10	Influence of nanofluid instability on thermodynamic properties near the critical point
cpl	161.20	J/mol×K	298.15	NIST Webbook
cpl	154.75	J/mol×K	298.15	NIST Webbook
cpl	165.60	J/mol×K	311.60	NIST Webbook
cpl	154.43	J/mol×K	298.15	NIST Webbook
cpl	162.80	J/mol×K	298.20	NIST Webbook
cpl	180.30	J/mol×K	324.00	NIST Webbook
cpl	154.00	J/mol×K	298.00	NIST Webbook
cpl	159.99	J/mol×K	298.04	NIST Webbook
cpl	186.90	J/mol×K	508.57	Influence of nanofluid instability on thermodynamic properties near the critical point
cpl	163.60	J/mol×K	298.00	NIST Webbook
cpl	149.75	J/mol×K	292.84	NIST Webbook

cpl	180.30	J/mol×K	298.10	NIST Webbook
cpl	185.69	J/mol×K	508.47	Influence of nanofluid instability on thermodynamic properties near the critical point
cpl	152.30	J/mol×K	293.10	NIST Webbook
cpl	169.90	J/mol×K	303.00	NIST Webbook
cpl	187.80	J/mol×K	508.37	Influence of nanofluid instability on thermodynamic properties near the critical point
cpl	195.91	J/mol×K	508.28	Influence of nanofluid instability on thermodynamic properties near the critical point
cpl	208.53	J/mol×K	508.18	Influence of nanofluid instability on thermodynamic properties near the critical point
cpl	218.14	J/mol×K	508.16	Influence of nanofluid instability on thermodynamic properties near the critical point
cpl	172.40	J/mol×K	303.20	NIST Webbook
cpl	175.18	J/mol×K	522.60	Influence of nanofluid instability on thermodynamic properties near the critical point
cpl	151.00	J/mol×K	293.10	NIST Webbook
cpl	173.19	J/mol×K	522.14	Influence of nanofluid instability on thermodynamic properties near the critical point
dvisc	0.0015630	Pa×s	308.15	Densities and viscosities of binary mixtures of ethylmethylketone and 2-alkanols; application of the ERAS model and cubic EOS

dvisc	0.0023820	Pa×s	293.15	Viscosities, Densities, and Speed of Sound of the Cycloalkanes with Secondary Alcohols at T = (293.15, 298.15, and 303.15) K: New UNIFAC-VISCO Interaction Parameters
dvisc	0.0020450	Pa×s	298.15	Viscosities, Densities, and Speed of Sound of the Cycloalkanes with Secondary Alcohols at T = (293.15, 298.15, and 303.15) K: New UNIFAC-VISCO Interaction Parameters
dvisc	0.0017630	Pa×s	303.15	Viscosities, Densities, and Speed of Sound of the Cycloalkanes with Secondary Alcohols at T = (293.15, 298.15, and 303.15) K: New UNIFAC-VISCO Interaction Parameters
dvisc	0.0020550	Pa×s	298.15	Densities and Viscosities of Binary Mixtures of Cyclohexanone and 2-Alkanols
dvisc	0.0017880	Pa×s	303.15	Densities and Viscosities of Binary Mixtures of Cyclohexanone and 2-Alkanols
dvisc	0.0015630	Pa×s	308.15	Densities and Viscosities of Binary Mixtures of Cyclohexanone and 2-Alkanols
dvisc	0.0013780	Pa×s	313.15	Densities and Viscosities of Binary Mixtures of Cyclohexanone and 2-Alkanols

dvisc	0.0023840	Pa×s	293.15	Viscosities and Densities of Binary Mixtures of (N-Acetylmorpholine + Alkanols) from (293.15 to 323.15) K
dvisc	0.0017520	Pa×s	303.15	Viscosities and Densities of Binary Mixtures of (N-Acetylmorpholine + Alkanols) from (293.15 to 323.15) K
dvisc	0.0010145	Pa×s	323.15	Experimental excess molar properties of binary mixtures of (3-amino-1-propanol + isobutanol, 2-propanol) at T = (293.15 to 333.15) K and modelling the excess molar volume by Prigogine-Flory-Patterson theory
dvisc	0.0013350	Pa×s	313.15	Viscosities and Densities of Binary Mixtures of (N-Acetylmorpholine + Alkanols) from (293.15 to 323.15) K
dvisc	0.0010380	Pa×s	323.15	Viscosities and Densities of Binary Mixtures of (N-Acetylmorpholine + Alkanols) from (293.15 to 323.15) K
dvisc	0.0022256	Pa×s	293.15	Densities and Viscosities of Binary Liquid Mixtures of 2-Butanone with Branched Alcohols at (293.15 to 313.15) K

dvisc	0.0019102	Pa×s	298.15	Densities and Viscosities of Binary Liquid Mixtures of 2-Butanone with Branched Alcohols at (293.15 to 313.15) K	
dvisc	0.0017370	Pa×s	303.15	Densities and Viscosities of Binary Liquid Mixtures of 2-Butanone with Branched Alcohols at (293.15 to 313.15) K	
dvisc	0.0010020	Pa×s	323.15	Densities and Viscosities of Diethyl Carbonate + Toluene, + Methanol, and + 2-Propanol from (293.15 to 363.15) K	
dvisc	0.0012880	Pa×s	313.15	Densities and Viscosities of Diethyl Carbonate + Toluene, + Methanol, and + 2-Propanol from (293.15 to 363.15) K	
dvisc	0.0017430	Pa×s	303.15	Densities and Viscosities of Diethyl Carbonate + Toluene, + Methanol, and + 2-Propanol from (293.15 to 363.15) K	
dvisc	0.0020690	Pa×s	298.15	Densities and Viscosities of Diethyl Carbonate + Toluene, + Methanol, and + 2-Propanol from (293.15 to 363.15) K	
dvisc	0.0023820	Pa×s	293.15	Densities and Viscosities of Diethyl Carbonate + Toluene, + Methanol, and + 2-Propanol from (293.15 to 363.15) K	

dvisc	0.0013790	Pa×s	313.15	Densities and viscosities of binary mixtures of ethylmethylketone and 2-alkanols; application of the ERAS model and cubic EOS
dvisc	0.0015242	Pa×s	308.15	Densities and Viscosities of Binary Liquid Mixtures of 2-Butanone with Branched Alcohols at (293.15 to 313.15) K
dvisc	0.0017880	Pa×s	303.15	Densities and viscosities of binary mixtures of ethylmethylketone and 2-alkanols; application of the ERAS model and cubic EOS
dvisc	0.0020550	Pa×s	298.15	Densities and viscosities of binary mixtures of ethylmethylketone and 2-alkanols; application of the ERAS model and cubic EOS
dvisc	0.0006241	Pa×s	333.15	Experimental excess molar properties of binary mixtures of (3-amino-1-propanol + isobutanol, 2-propanol) at T = (293.15 to 333.15) K and modelling the excess molar volume by Prigogine-Flory-Patterson theory
dvisc	0.0007890	Pa×s	333.15	Densities and Viscosities of Diethyl Carbonate + Toluene, + Methanol, and + 2-Propanol from (293.15 to 363.15) K

dvisc	0.0013297	Pa×s	313.15	Experimental excess molar properties of binary mixtures of (3-amino-1-propanol + isobutanol, 2-propanol) at T = (293.15 to 333.15) K and modelling the excess molar volume by Prigogine-Flory-Patterson theory
dvisc	0.0017694	Pa×s	303.15	Experimental excess molar properties of binary mixtures of (3-amino-1-propanol + isobutanol, 2-propanol) at T = (293.15 to 333.15) K and modelling the excess molar volume by Prigogine-Flory-Patterson theory
dvisc	0.0023621	Pa×s	293.15	Experimental excess molar properties of binary mixtures of (3-amino-1-propanol + isobutanol, 2-propanol) at T = (293.15 to 333.15) K and modelling the excess molar volume by Prigogine-Flory-Patterson theory
dvisc	0.0015730	Pa×s	308.15	Studies of mixing properties of binary systems of 2-propanol with hexadecane and squalane at T = (298.15, 303.15, and 308.15) K
dvisc	0.0017580	Pa×s	303.15	Studies of mixing properties of binary systems of 2-propanol with hexadecane and squalane at T = (298.15, 303.15, and 308.15) K

dvisc	0.0020610	Pa×s	298.15	Studies of mixing properties of binary systems of 2-propanol with hexadecane and squalane at T = (298.15, 303.15, and 308.15) K
dvisc	0.0024050	Pa×s	293.15	Viscosities and densities for binary mixtures of N-methylpiperazine with methanol, ethanol, n-propanol, iso-propanol, n-butanol and iso-butanol at 293.15, 298.15 and 303.15K
dvisc	0.0010411	Pa×s	323.15	Densities and viscosities of binary mixtures of {dimethylsulfoxide + aliphatic lower alkanols (C1 C3)} at temperatures from T = 303.15 K to T = 323.15 K
dvisc	0.0011822	Pa×s	318.15	Densities and viscosities of binary mixtures of {dimethylsulfoxide + aliphatic lower alkanols (C1 C3)} at temperatures from T = 303.15 K to T = 323.15 K
dvisc	0.0013554	Pa×s	313.15	Densities and viscosities of binary mixtures of {dimethylsulfoxide + aliphatic lower alkanols (C1 C3)} at temperatures from T = 303.15 K to T = 323.15 K
dvisc	0.0015529	Pa×s	308.15	Densities and viscosities of binary mixtures of {dimethylsulfoxide + aliphatic lower alkanols (C1 C3)} at temperatures from T = 303.15 K to T = 323.15 K

dvisc	0.0017630	Pa×s	303.15	Viscosity, density, and speed of sound of methylcyclopentane with primary and secondary alcohols at T = (293.15, 298.15, and 303.15) K
dvisc	0.0020450	Pa×s	298.15	Viscosity, density, and speed of sound of methylcyclopentane with primary and secondary alcohols at T = (293.15, 298.15, and 303.15) K
dvisc	0.0023820	Pa×s	293.15	Viscosity, density, and speed of sound of methylcyclopentane with primary and secondary alcohols at T = (293.15, 298.15, and 303.15) K
dvisc	0.0011300	Pa×s	313.00	Ultrasonic velocity, viscosity and excess properties of binary mixture of tetrahydrofuran with 1-propanol and 2-propanol
dvisc	0.0013472	Pa×s	313.15	Densities and Viscosities of Binary Liquid Mixtures of 2-Butanone with Branched Alcohols at (293.15 to 313.15) K
dvisc	0.0014900	Pa×s	303.00	Ultrasonic velocity, viscosity and excess properties of binary mixture of tetrahydrofuran with 1-propanol and 2-propanol
dvisc	0.0006260	Pa×s	343.15	Densities and Viscosities of Diethyl Carbonate + Toluene, + Methanol, and + 2-Propanol from (293.15 to 363.15) K

dvisc	0.0021000	Pa×s	293.00	Ultrasonic velocity, viscosity and excess properties of binary mixture of tetrahydrofuran with 1-propanol and 2-propanol
dvisc	0.0017820	Pa×s	303.15	Viscosities and densities for binary mixtures of N-methylpiperazine with methanol, ethanol, n-propanol, iso-propanol, n-butanol and iso-butanol at 293.15, 298.15 and 303.15K
dvisc	0.0020620	Pa×s	298.15	Viscosities and densities for binary mixtures of N-methylpiperazine with methanol, ethanol, n-propanol, iso-propanol, n-butanol and iso-butanol at 293.15, 298.15 and 303.15K
dvisc	0.0017919	Pa×s	303.15	Densities and viscosities of binary mixtures of {dimethylsulfoxide + aliphatic lower alkanols (C1 C3)} at temperatures from T = 303.15 K to T = 323.15 K
econd	8.08e-05	S/m	313.15	Electrical Conductivity of Caprolactam Tetrabutylammonium Bromide Ionic Liquids in Aqueous and Alcohol Binary Systems
econd	7.50e-05	S/m	308.15	Electrical Conductivity of Caprolactam Tetrabutylammonium Bromide Ionic Liquids in Aqueous and Alcohol Binary Systems

econd	7.00e-05	S/m	303.15	Electrical Conductivity of Caprolactam Tetrabutylammonium Bromide Ionic Liquids in Aqueous and Alcohol Binary Systems
econd	5.20e-05	S/m	298.15	Electrical Conductivity of Caprolactam Tetrabutylammonium Bromide Ionic Liquids in Aqueous and Alcohol Binary Systems
econd	8.43e-05	S/m	318.15	Electrical Conductivity of Caprolactam Tetrabutylammonium Bromide Ionic Liquids in Aqueous and Alcohol Binary Systems
hfust	5.41	kJ/mol	185.20	NIST Webbook
hfust	5.30	kJ/mol	184.60	NIST Webbook
hfust	5.30	kJ/mol	184.60	NIST Webbook
hfust	5.41	kJ/mol	185.20	NIST Webbook
hfust	5.37	kJ/mol	184.67	NIST Webbook
hfust	5.41	kJ/mol	185.20	NIST Webbook
hfust	5.37	kJ/mol	184.70	NIST Webbook
hvapt	45.70	kJ/mol	323.50	NIST Webbook
hvapt	39.10	kJ/mol	451.50	NIST Webbook
hvapt	38.90 ± 0.10	kJ/mol	363.00	NIST Webbook
hvapt	39.80 ± 0.10	kJ/mol	355.00	NIST Webbook
hvapt	41.00 ± 0.10	kJ/mol	346.00	NIST Webbook
hvapt	42.80	kJ/mol	346.00	NIST Webbook
hvapt	42.70 ± 0.10	kJ/mol	330.00	NIST Webbook
hvapt	45.50	kJ/mol	318.00	NIST Webbook
hvapt	43.20	kJ/mol	324.00	NIST Webbook
hvapt	43.10	kJ/mol	343.50	NIST Webbook
hvapt	39.20	kJ/mol	420.00	NIST Webbook
hvapt	41.30	kJ/mol	366.50	NIST Webbook
hvapt	42.00	kJ/mol	357.50	NIST Webbook
hvapt	50.30	kJ/mol	211.50	NIST Webbook
hvapt	44.80	kJ/mol	327.50	NIST Webbook
hvapt	10.50	kJ/mol	503.00	NIST Webbook
hvapt	16.50	kJ/mol	483.00	NIST Webbook
hvapt	23.70	kJ/mol	453.00	NIST Webbook

hvapt	29.70	kJ/mol	423.00	NIST Webbook
hvapt	39.80	kJ/mol	355.00	NIST Webbook
hvapt	43.20	kJ/mol	338.50	NIST Webbook
hvapt	39.85	kJ/mol	355.40	NIST Webbook
hvapt	39.83	kJ/mol	355.50	KDB
hvapt	39.80	kJ/mol	355.00	NIST Webbook
hvapt	43.40 ± 0.08	kJ/mol	324.11	NIST Webbook
hvapt	41.10	kJ/mol	387.00	NIST Webbook
hvapt	41.70	kJ/mol	339.00	NIST Webbook
hvapt	35.30	kJ/mol	480.50	NIST Webbook
pvap	3677.00	kPa	493.15	Measurement and correlation of vapor-liquid equilibria for the 2-propanol + n-hexane system near the critical
pvap	177.97	kPa	370.40	Vapor-Liquid Equilibrium of Binary Mixtures Containing Isopropyl Acetate and Alkanols at 101.32 kPa
pvap	152.55	kPa	366.10	Vapor-Liquid Equilibrium of Binary Mixtures Containing Isopropyl Acetate and Alkanols at 101.32 kPa
pvap	203.40	kPa	374.20	Vapor-Liquid Equilibrium of Binary Mixtures Containing Isopropyl Acetate and Alkanols at 101.32 kPa
pvap	228.82	kPa	377.60	Vapor-Liquid Equilibrium of Binary Mixtures Containing Isopropyl Acetate and Alkanols at 101.32 kPa
pvap	254.25	kPa	380.80	Vapor-Liquid Equilibrium of Binary Mixtures Containing Isopropyl Acetate and Alkanols at 101.32 kPa

pvap	101.30	kPa	355.33	Vapor Liquid Equilibria for Ternary Mixtures of Isopropyl Alcohol, Isopropyl Acetate, and DMSO at 101.3 kPa
pvap	96.15	kPa	354.17	Vapor Liquid Equilibrium Data for Binary Mixtures of Acetic Acid + Anisole, Acetone + Anisole, and Isopropanol + Anisole at Pressure 96.15 kPa
pvap	101.06	kPa	355.40	Ammonium-Based Ionic Liquid as an Entrainer for the Separation of n-Propanol + Water and Isopropanol + Water Mixtures
pvap	1610.00	kPa	450.00	Pressure-Drop Method for Detecting Bubble and Dew Points of Multicomponent Mixtures at Temperatures of up to 573 K
pvap	1640.00	kPa	451.50	Pressure-Drop Method for Detecting Bubble and Dew Points of Multicomponent Mixtures at Temperatures of up to 573 K
pvap	2030.00	kPa	462.50	Pressure-Drop Method for Detecting Bubble and Dew Points of Multicomponent Mixtures at Temperatures of up to 573 K

pvap	2040.00	kPa	461.50	Pressure-Drop Method for Detecting Bubble and Dew Points of Multicomponent Mixtures at Temperatures of up to 573 K
pvap	2420.00	kPa	469.00	Pressure-Drop Method for Detecting Bubble and Dew Points of Multicomponent Mixtures at Temperatures of up to 573 K
pvap	2460.00	kPa	469.00	Pressure-Drop Method for Detecting Bubble and Dew Points of Multicomponent Mixtures at Temperatures of up to 573 K
pvap	2740.00	kPa	477.00	Pressure-Drop Method for Detecting Bubble and Dew Points of Multicomponent Mixtures at Temperatures of up to 573 K
pvap	2780.00	kPa	477.00	Pressure-Drop Method for Detecting Bubble and Dew Points of Multicomponent Mixtures at Temperatures of up to 573 K
pvap	3060.00	kPa	482.00	Pressure-Drop Method for Detecting Bubble and Dew Points of Multicomponent Mixtures at Temperatures of up to 573 K
pvap	3090.00	kPa	482.00	Pressure-Drop Method for Detecting Bubble and Dew Points of Multicomponent Mixtures at Temperatures of up to 573 K

pvap	3460.00	kPa	489.50	Pressure-Drop Method for Detecting Bubble and Dew Points of Multicomponent Mixtures at Temperatures of up to 573 K
pvap	3500.00	kPa	489.00	Pressure-Drop Method for Detecting Bubble and Dew Points of Multicomponent Mixtures at Temperatures of up to 573 K
pvap	3970.00	kPa	494.50	Pressure-Drop Method for Detecting Bubble and Dew Points of Multicomponent Mixtures at Temperatures of up to 573 K
pvap	101.32	kPa	355.60	Acetonitrile Dehydration via Extractive Distillation Using Low Transition Temperature Mixtures as Entrainers
pvap	39.54	kPa	333.72	Vapor-Liquid Equilibrium Data at 343 K and Excess Molar Enthalpy Data at 298 K for the Binary Systems of Ethanol + 2,4,4-Trimethyl-1-pentene and 2-Propanol + 2,4,4-Trimethyl-1-pentene
pvap	55.03	kPa	340.98	Vapor-Liquid Equilibrium Data at 343 K and Excess Molar Enthalpy Data at 298 K for the Binary Systems of Ethanol + 2,4,4-Trimethyl-1-pentene and 2-Propanol + 2,4,4-Trimethyl-1-pentene

pvap	60.66	kPa	343.19	Vapor-Liquid Equilibrium Data at 343 K and Excess Molar Enthalpy Data at 298 K for the Binary Systems of Ethanol + 2,4,4-Trimethyl-1-pentene and 2-Propanol + 2,4,4-Trimethyl-1-pentene
pvap	61.30	kPa	343.44	Vapor-Liquid Equilibrium Data at 343 K and Excess Molar Enthalpy Data at 298 K for the Binary Systems of Ethanol + 2,4,4-Trimethyl-1-pentene and 2-Propanol + 2,4,4-Trimethyl-1-pentene
pvap	65.23	kPa	344.86	Vapor-Liquid Equilibrium Data at 343 K and Excess Molar Enthalpy Data at 298 K for the Binary Systems of Ethanol + 2,4,4-Trimethyl-1-pentene and 2-Propanol + 2,4,4-Trimethyl-1-pentene
pvap	69.47	kPa	346.33	Vapor-Liquid Equilibrium Data at 343 K and Excess Molar Enthalpy Data at 298 K for the Binary Systems of Ethanol + 2,4,4-Trimethyl-1-pentene and 2-Propanol + 2,4,4-Trimethyl-1-pentene
pvap	75.27	kPa	348.21	Vapor-Liquid Equilibrium Data at 343 K and Excess Molar Enthalpy Data at 298 K for the Binary Systems of Ethanol + 2,4,4-Trimethyl-1-pentene and 2-Propanol + 2,4,4-Trimethyl-1-pentene

pvap	79.34	kPa	349.48	Vapor-Liquid Equilibrium Data at 343 K and Excess Molar Enthalpy Data at 298 K for the Binary Systems of Ethanol + 2,4,4-Trimethyl-1-pentene and 2-Propanol + 2,4,4-Trimethyl-1-pentene
pvap	90.34	kPa	352.63	Vapor-Liquid Equilibrium Data at 343 K and Excess Molar Enthalpy Data at 298 K for the Binary Systems of Ethanol + 2,4,4-Trimethyl-1-pentene and 2-Propanol + 2,4,4-Trimethyl-1-pentene
pvap	93.41	kPa	353.46	Vapor-Liquid Equilibrium Data at 343 K and Excess Molar Enthalpy Data at 298 K for the Binary Systems of Ethanol + 2,4,4-Trimethyl-1-pentene and 2-Propanol + 2,4,4-Trimethyl-1-pentene
pvap	101.07	kPa	355.42	Vapor-Liquid Equilibrium Data at 343 K and Excess Molar Enthalpy Data at 298 K for the Binary Systems of Ethanol + 2,4,4-Trimethyl-1-pentene and 2-Propanol + 2,4,4-Trimethyl-1-pentene
pvap	3081.00	kPa	483.15	Measurement and correlation of vapor-liquid equilibria for the 2-propanol + n-hexane system near the critical
pvap	4362.00	kPa	503.15	Measurement and correlation of vapor-liquid equilibria for the 2-propanol + n-hexane system near the critical

pvap	36.47	kPa	332.03	Vapor Liquid Equilibrium for the Trans-2-Butene + Methanol, + Ethanol, + 2-Propanol, + 2-Butanol and + 2-Methyl-2-Propanol Systems at 332 K
pvap	23.47	kPa	323.16	Isothermal Vapor Liquid Equilibrium for Binary 2-Methylpropene - C1-C4 Alcohol-Systems
pvap	5.80	kPa	298.15	Vapor-Pressure Measurements of Liquid Solutions at Different Temperatures: Apparatus for Use over an Extended Temperature Range and Some New Data
pvap	7.86	kPa	303.15	Vapor-Pressure Measurements of Liquid Solutions at Different Temperatures: Apparatus for Use over an Extended Temperature Range and Some New Data
pvap	10.52	kPa	308.15	Vapor-Pressure Measurements of Liquid Solutions at Different Temperatures: Apparatus for Use over an Extended Temperature Range and Some New Data
pvap	13.94	kPa	313.15	Vapor-Pressure Measurements of Liquid Solutions at Different Temperatures: Apparatus for Use over an Extended Temperature Range and Some New Data

pvap	127.12	kPa	361.21	Vapor-Liquid Equilibrium of Binary Mixtures Containing Isopropyl Acetate and Alkanols at 101.32 kPa
pvap	101.54	kPa	355.41	Vapor-Liquid Equilibrium of Binary Mixtures Containing Isopropyl Acetate and Alkanols at 101.32 kPa
pvap	63.92	kPa	344.70	An experimental investigation and modelling of the thermal and caloric properties of nanofluids isopropyl alcohol - Al2O3 nanoparticles
pvap	36.06	kPa	332.10	An experimental investigation and modelling of the thermal and caloric properties of nanofluids isopropyl alcohol - Al2O3 nanoparticles
pvap	16.22	kPa	315.80	An experimental investigation and modelling of the thermal and caloric properties of nanofluids isopropyl alcohol - Al2O3 nanoparticles
pvap	8.00	kPa	303.10	An experimental investigation and modelling of the thermal and caloric properties of nanofluids isopropyl alcohol - Al2O3 nanoparticles
pvap	5.78	kPa	297.80	An experimental investigation and modelling of the thermal and caloric properties of nanofluids isopropyl alcohol - Al2O3 nanoparticles

pvap	101.30	kPa	355.71	A new analysis method for improving collection of vapor-liquid equilibrium (VLE) data of binary mixtures using differential scanning calorimetry (DSC)
pvap	133.65	kPa	362.52	Isobaric vapor-liquid equilibria for the extractive distillation of 2-propanol + water mixtures using 1-ethyl-3-methylimidazolium dicyanamide ionic liquid
pvap	132.35	kPa	362.26	Isobaric vapor-liquid equilibria for the extractive distillation of 2-propanol + water mixtures using 1-ethyl-3-methylimidazolium dicyanamide ionic liquid
pvap	130.93	kPa	361.98	Isobaric vapor-liquid equilibria for the extractive distillation of 2-propanol + water mixtures using 1-ethyl-3-methylimidazolium dicyanamide ionic liquid
pvap	129.08	kPa	361.62	Isobaric vapor-liquid equilibria for the extractive distillation of 2-propanol + water mixtures using 1-ethyl-3-methylimidazolium dicyanamide ionic liquid

pvap	127.36	kPa	361.26	Isobaric vapor-liquid equilibria for the extractive distillation of 2-propanol + water mixtures using 1-ethyl-3-methylimidazolium dicyanamide ionic liquid
pvap	125.62	kPa	360.90	Isobaric vapor-liquid equilibria for the extractive distillation of 2-propanol + water mixtures using 1-ethyl-3-methylimidazolium dicyanamide ionic liquid
pvap	18.24	kPa	318.15	Vapor-Pressure Measurements of Liquid Solutions at Different Temperatures: Apparatus for Use over an Extended Temperature Range and Some New Data
pvap	122.11	kPa	360.16	Isobaric vapor-liquid equilibria for the extractive distillation of 2-propanol + water mixtures using 1-ethyl-3-methylimidazolium dicyanamide ionic liquid
pvap	120.24	kPa	359.77	Isobaric vapor-liquid equilibria for the extractive distillation of 2-propanol + water mixtures using 1-ethyl-3-methylimidazolium dicyanamide ionic liquid

pvap	118.64	kPa	359.43	Isobaric vapor-liquid equilibria for the extractive distillation of 2-propanol + water mixtures using 1-ethyl-3-methylimidazolium dicyanamide ionic liquid
pvap	116.64	kPa	358.98	Isobaric vapor-liquid equilibria for the extractive distillation of 2-propanol + water mixtures using 1-ethyl-3-methylimidazolium dicyanamide ionic liquid
pvap	115.10	kPa	358.65	Isobaric vapor-liquid equilibria for the extractive distillation of 2-propanol + water mixtures using 1-ethyl-3-methylimidazolium dicyanamide ionic liquid
pvap	113.79	kPa	358.35	Isobaric vapor-liquid equilibria for the extractive distillation of 2-propanol + water mixtures using 1-ethyl-3-methylimidazolium dicyanamide ionic liquid
pvap	111.88	kPa	357.93	Isobaric vapor-liquid equilibria for the extractive distillation of 2-propanol + water mixtures using 1-ethyl-3-methylimidazolium dicyanamide ionic liquid

pvap	110.13	kPa	357.52	Isobaric vapor-liquid equilibria for the extractive distillation of 2-propanol + water mixtures using 1-ethyl-3-methylimidazolium dicyanamide ionic liquid
pvap	108.63	kPa	357.16	Isobaric vapor-liquid equilibria for the extractive distillation of 2-propanol + water mixtures using 1-ethyl-3-methylimidazolium dicyanamide ionic liquid
pvap	106.91	kPa	356.76	Isobaric vapor-liquid equilibria for the extractive distillation of 2-propanol + water mixtures using 1-ethyl-3-methylimidazolium dicyanamide ionic liquid
pvap	105.38	kPa	356.40	Isobaric vapor-liquid equilibria for the extractive distillation of 2-propanol + water mixtures using 1-ethyl-3-methylimidazolium dicyanamide ionic liquid
pvap	103.44	kPa	355.93	Isobaric vapor-liquid equilibria for the extractive distillation of 2-propanol + water mixtures using 1-ethyl-3-methylimidazolium dicyanamide ionic liquid

pvap	101.28	kPa	355.38	Isobaric vapor-liquid equilibria for the extractive distillation of 2-propanol + water mixtures using 1-ethyl-3-methylimidazolium dicyanamide ionic liquid
pvap	100.03	kPa	355.09	Isobaric vapor-liquid equilibria for the extractive distillation of 2-propanol + water mixtures using 1-ethyl-3-methylimidazolium dicyanamide ionic liquid
pvap	100.00	kPa	355.09	Isobaric vapor-liquid equilibria for the extractive distillation of 2-propanol + water mixtures using 1-ethyl-3-methylimidazolium dicyanamide ionic liquid
pvap	98.05	kPa	354.59	Isobaric vapor-liquid equilibria for the extractive distillation of 2-propanol + water mixtures using 1-ethyl-3-methylimidazolium dicyanamide ionic liquid
pvap	96.49	kPa	354.20	Isobaric vapor-liquid equilibria for the extractive distillation of 2-propanol + water mixtures using 1-ethyl-3-methylimidazolium dicyanamide ionic liquid

pvap	94.46	kPa	353.66	Isobaric vapor-liquid equilibria for the extractive distillation of 2-propanol + water mixtures using 1-ethyl-3-methylimidazolium dicyanamide ionic liquid
pvap	92.57	kPa	353.16	Isobaric vapor-liquid equilibria for the extractive distillation of 2-propanol + water mixtures using 1-ethyl-3-methylimidazolium dicyanamide ionic liquid
pvap	90.24	kPa	352.54	Isobaric vapor-liquid equilibria for the extractive distillation of 2-propanol + water mixtures using 1-ethyl-3-methylimidazolium dicyanamide ionic liquid
pvap	88.63	kPa	352.09	Isobaric vapor-liquid equilibria for the extractive distillation of 2-propanol + water mixtures using 1-ethyl-3-methylimidazolium dicyanamide ionic liquid
pvap	87.17	kPa	351.69	Isobaric vapor-liquid equilibria for the extractive distillation of 2-propanol + water mixtures using 1-ethyl-3-methylimidazolium dicyanamide ionic liquid

pvap	85.47	kPa	351.21	Isobaric vapor-liquid equilibria for the extractive distillation of 2-propanol + water mixtures using 1-ethyl-3-methylimidazolium dicyanamide ionic liquid
pvap	83.94	kPa	350.75	Isobaric vapor-liquid equilibria for the extractive distillation of 2-propanol + water mixtures using 1-ethyl-3-methylimidazolium dicyanamide ionic liquid
pvap	81.58	kPa	350.08	Isobaric vapor-liquid equilibria for the extractive distillation of 2-propanol + water mixtures using 1-ethyl-3-methylimidazolium dicyanamide ionic liquid
pvap	79.96	kPa	349.61	Isobaric vapor-liquid equilibria for the extractive distillation of 2-propanol + water mixtures using 1-ethyl-3-methylimidazolium dicyanamide ionic liquid
pvap	77.81	kPa	348.95	Isobaric vapor-liquid equilibria for the extractive distillation of 2-propanol + water mixtures using 1-ethyl-3-methylimidazolium dicyanamide ionic liquid

pvap	75.96	kPa	348.37	Isobaric vapor-liquid equilibria for the extractive distillation of 2-propanol + water mixtures using 1-ethyl-3-methylimidazolium dicyanamide ionic liquid
pvap	74.18	kPa	347.79	Isobaric vapor-liquid equilibria for the extractive distillation of 2-propanol + water mixtures using 1-ethyl-3-methylimidazolium dicyanamide ionic liquid
pvap	72.04	kPa	347.11	Isobaric vapor-liquid equilibria for the extractive distillation of 2-propanol + water mixtures using 1-ethyl-3-methylimidazolium dicyanamide ionic liquid
pvap	70.12	kPa	346.46	Isobaric vapor-liquid equilibria for the extractive distillation of 2-propanol + water mixtures using 1-ethyl-3-methylimidazolium dicyanamide ionic liquid
pvap	69.14	kPa	346.13	Isobaric vapor-liquid equilibria for the extractive distillation of 2-propanol + water mixtures using 1-ethyl-3-methylimidazolium dicyanamide ionic liquid

pvap	68.03	kPa	345.75	Isobaric vapor-liquid equilibria for the extractive distillation of 2-propanol + water mixtures using 1-ethyl-3-methylimidazolium dicyanamide ionic liquid
pvap	66.84	kPa	345.35	Isobaric vapor-liquid equilibria for the extractive distillation of 2-propanol + water mixtures using 1-ethyl-3-methylimidazolium dicyanamide ionic liquid
pvap	65.97	kPa	345.05	Isobaric vapor-liquid equilibria for the extractive distillation of 2-propanol + water mixtures using 1-ethyl-3-methylimidazolium dicyanamide ionic liquid
pvap	64.96	kPa	344.68	Isobaric vapor-liquid equilibria for the extractive distillation of 2-propanol + water mixtures using 1-ethyl-3-methylimidazolium dicyanamide ionic liquid
pvap	63.92	kPa	344.32	Isobaric vapor-liquid equilibria for the extractive distillation of 2-propanol + water mixtures using 1-ethyl-3-methylimidazolium dicyanamide ionic liquid

pvap	62.98	kPa	343.96	Isobaric vapor-liquid equilibria for the extractive distillation of 2-propanol + water mixtures using 1-ethyl-3-methylimidazolium dicyanamide ionic liquid
pvap	61.94	kPa	343.59	Isobaric vapor-liquid equilibria for the extractive distillation of 2-propanol + water mixtures using 1-ethyl-3-methylimidazolium dicyanamide ionic liquid
pvap	60.91	kPa	343.20	Isobaric vapor-liquid equilibria for the extractive distillation of 2-propanol + water mixtures using 1-ethyl-3-methylimidazolium dicyanamide ionic liquid
pvap	59.93	kPa	342.82	Isobaric vapor-liquid equilibria for the extractive distillation of 2-propanol + water mixtures using 1-ethyl-3-methylimidazolium dicyanamide ionic liquid
pvap	59.04	kPa	342.48	Isobaric vapor-liquid equilibria for the extractive distillation of 2-propanol + water mixtures using 1-ethyl-3-methylimidazolium dicyanamide ionic liquid

pvap	58.05	kPa	342.09	Isobaric vapor-liquid equilibria for the extractive distillation of 2-propanol + water mixtures using 1-ethyl-3-methylimidazolium dicyanamide ionic liquid
pvap	56.96	kPa	341.69	Isobaric vapor-liquid equilibria for the extractive distillation of 2-propanol + water mixtures using 1-ethyl-3-methylimidazolium dicyanamide ionic liquid
pvap	55.92	kPa	341.25	Isobaric vapor-liquid equilibria for the extractive distillation of 2-propanol + water mixtures using 1-ethyl-3-methylimidazolium dicyanamide ionic liquid
pvap	55.03	kPa	340.89	Isobaric vapor-liquid equilibria for the extractive distillation of 2-propanol + water mixtures using 1-ethyl-3-methylimidazolium dicyanamide ionic liquid
pvap	53.94	kPa	340.43	Isobaric vapor-liquid equilibria for the extractive distillation of 2-propanol + water mixtures using 1-ethyl-3-methylimidazolium dicyanamide ionic liquid

pvap	53.01	kPa	340.04	Isobaric vapor-liquid equilibria for the extractive distillation of 2-propanol + water mixtures using 1-ethyl-3-methylimidazolium dicyanamide ionic liquid
pvap	52.33	kPa	339.72	Isobaric vapor-liquid equilibria for the extractive distillation of 2-propanol + water mixtures using 1-ethyl-3-methylimidazolium dicyanamide ionic liquid
pvap	51.26	kPa	339.30	Isobaric vapor-liquid equilibria for the extractive distillation of 2-propanol + water mixtures using 1-ethyl-3-methylimidazolium dicyanamide ionic liquid
pvap	50.25	kPa	338.86	Isobaric vapor-liquid equilibria for the extractive distillation of 2-propanol + water mixtures using 1-ethyl-3-methylimidazolium dicyanamide ionic liquid
pvap	49.34	kPa	338.45	Isobaric vapor-liquid equilibria for the extractive distillation of 2-propanol + water mixtures using 1-ethyl-3-methylimidazolium dicyanamide ionic liquid

pvap	48.35	kPa	337.98	Isobaric vapor-liquid equilibria for the extractive distillation of 2-propanol + water mixtures using 1-ethyl-3-methylimidazolium dicyanamide ionic liquid
pvap	47.28	kPa	337.50	Isobaric vapor-liquid equilibria for the extractive distillation of 2-propanol + water mixtures using 1-ethyl-3-methylimidazolium dicyanamide ionic liquid
pvap	46.34	kPa	337.07	Isobaric vapor-liquid equilibria for the extractive distillation of 2-propanol + water mixtures using 1-ethyl-3-methylimidazolium dicyanamide ionic liquid
pvap	45.32	kPa	336.55	Isobaric vapor-liquid equilibria for the extractive distillation of 2-propanol + water mixtures using 1-ethyl-3-methylimidazolium dicyanamide ionic liquid
pvap	44.38	kPa	336.11	Isobaric vapor-liquid equilibria for the extractive distillation of 2-propanol + water mixtures using 1-ethyl-3-methylimidazolium dicyanamide ionic liquid

pvap	43.46	kPa	335.64	Isobaric vapor-liquid equilibria for the extractive distillation of 2-propanol + water mixtures using 1-ethyl-3-methylimidazolium dicyanamide ionic liquid
pvap	42.44	kPa	335.12	Isobaric vapor-liquid equilibria for the extractive distillation of 2-propanol + water mixtures using 1-ethyl-3-methylimidazolium dicyanamide ionic liquid
pvap	41.44	kPa	334.60	Isobaric vapor-liquid equilibria for the extractive distillation of 2-propanol + water mixtures using 1-ethyl-3-methylimidazolium dicyanamide ionic liquid
pvap	40.41	kPa	334.10	Isobaric vapor-liquid equilibria for the extractive distillation of 2-propanol + water mixtures using 1-ethyl-3-methylimidazolium dicyanamide ionic liquid
pvap	39.33	kPa	333.50	Isobaric vapor-liquid equilibria for the extractive distillation of 2-propanol + water mixtures using 1-ethyl-3-methylimidazolium dicyanamide ionic liquid

pvap	38.41	kPa	332.99	Isobaric vapor-liquid equilibria for the extractive distillation of 2-propanol + water mixtures using 1-ethyl-3-methylimidazolium dicyanamide ionic liquid
pvap	37.45	kPa	332.44	Isobaric vapor-liquid equilibria for the extractive distillation of 2-propanol + water mixtures using 1-ethyl-3-methylimidazolium dicyanamide ionic liquid
pvap	36.37	kPa	331.83	Isobaric vapor-liquid equilibria for the extractive distillation of 2-propanol + water mixtures using 1-ethyl-3-methylimidazolium dicyanamide ionic liquid
pvap	35.48	kPa	331.32	Isobaric vapor-liquid equilibria for the extractive distillation of 2-propanol + water mixtures using 1-ethyl-3-methylimidazolium dicyanamide ionic liquid
pvap	34.52	kPa	330.73	Isobaric vapor-liquid equilibria for the extractive distillation of 2-propanol + water mixtures using 1-ethyl-3-methylimidazolium dicyanamide ionic liquid

pvap	33.77	kPa	330.26	Isobaric vapor-liquid equilibria for the extractive distillation of 2-propanol + water mixtures using 1-ethyl-3-methylimidazolium dicyanamide ionic liquid
pvap	32.76	kPa	329.64	Isobaric vapor-liquid equilibria for the extractive distillation of 2-propanol + water mixtures using 1-ethyl-3-methylimidazolium dicyanamide ionic liquid
pvap	31.86	kPa	329.06	Isobaric vapor-liquid equilibria for the extractive distillation of 2-propanol + water mixtures using 1-ethyl-3-methylimidazolium dicyanamide ionic liquid
pvap	30.96	kPa	328.47	Isobaric vapor-liquid equilibria for the extractive distillation of 2-propanol + water mixtures using 1-ethyl-3-methylimidazolium dicyanamide ionic liquid
pvap	30.08	kPa	327.86	Isobaric vapor-liquid equilibria for the extractive distillation of 2-propanol + water mixtures using 1-ethyl-3-methylimidazolium dicyanamide ionic liquid

pvap	29.45	kPa	327.46	Isobaric vapor-liquid equilibria for the extractive distillation of 2-propanol + water mixtures using 1-ethyl-3-methylimidazolium dicyanamide ionic liquid
pvap	28.83	kPa	326.99	Isobaric vapor-liquid equilibria for the extractive distillation of 2-propanol + water mixtures using 1-ethyl-3-methylimidazolium dicyanamide ionic liquid
pvap	28.19	kPa	326.53	Isobaric vapor-liquid equilibria for the extractive distillation of 2-propanol + water mixtures using 1-ethyl-3-methylimidazolium dicyanamide ionic liquid
pvap	27.69	kPa	326.18	Isobaric vapor-liquid equilibria for the extractive distillation of 2-propanol + water mixtures using 1-ethyl-3-methylimidazolium dicyanamide ionic liquid
pvap	27.08	kPa	325.72	Isobaric vapor-liquid equilibria for the extractive distillation of 2-propanol + water mixtures using 1-ethyl-3-methylimidazolium dicyanamide ionic liquid

pvap	26.52	kPa	325.30	Isobaric vapor-liquid equilibria for the extractive distillation of 2-propanol + water mixtures using 1-ethyl-3-methylimidazolium dicyanamide ionic liquid
pvap	26.04	kPa	324.94	Isobaric vapor-liquid equilibria for the extractive distillation of 2-propanol + water mixtures using 1-ethyl-3-methylimidazolium dicyanamide ionic liquid
pvap	25.48	kPa	324.49	Isobaric vapor-liquid equilibria for the extractive distillation of 2-propanol + water mixtures using 1-ethyl-3-methylimidazolium dicyanamide ionic liquid
pvap	25.03	kPa	324.13	Isobaric vapor-liquid equilibria for the extractive distillation of 2-propanol + water mixtures using 1-ethyl-3-methylimidazolium dicyanamide ionic liquid
pvap	24.51	kPa	323.71	Isobaric vapor-liquid equilibria for the extractive distillation of 2-propanol + water mixtures using 1-ethyl-3-methylimidazolium dicyanamide ionic liquid

pvap	24.02	kPa	323.32	Isobaric vapor-liquid equilibria for the extractive distillation of 2-propanol + water mixtures using 1-ethyl-3-methylimidazolium dicyanamide ionic liquid
pvap	23.52	kPa	322.91	Isobaric vapor-liquid equilibria for the extractive distillation of 2-propanol + water mixtures using 1-ethyl-3-methylimidazolium dicyanamide ionic liquid
pvap	23.03	kPa	322.47	Isobaric vapor-liquid equilibria for the extractive distillation of 2-propanol + water mixtures using 1-ethyl-3-methylimidazolium dicyanamide ionic liquid
pvap	22.55	kPa	322.07	Isobaric vapor-liquid equilibria for the extractive distillation of 2-propanol + water mixtures using 1-ethyl-3-methylimidazolium dicyanamide ionic liquid
pvap	22.03	kPa	321.58	Isobaric vapor-liquid equilibria for the extractive distillation of 2-propanol + water mixtures using 1-ethyl-3-methylimidazolium dicyanamide ionic liquid

pvap	21.49	kPa	321.11	Isobaric vapor-liquid equilibria for the extractive distillation of 2-propanol + water mixtures using 1-ethyl-3-methylimidazolium dicyanamide ionic liquid
pvap	21.00	kPa	320.65	Isobaric vapor-liquid equilibria for the extractive distillation of 2-propanol + water mixtures using 1-ethyl-3-methylimidazolium dicyanamide ionic liquid
pvap	20.52	kPa	320.19	Isobaric vapor-liquid equilibria for the extractive distillation of 2-propanol + water mixtures using 1-ethyl-3-methylimidazolium dicyanamide ionic liquid
pvap	20.03	kPa	319.72	Isobaric vapor-liquid equilibria for the extractive distillation of 2-propanol + water mixtures using 1-ethyl-3-methylimidazolium dicyanamide ionic liquid
pvap	19.48	kPa	319.19	Isobaric vapor-liquid equilibria for the extractive distillation of 2-propanol + water mixtures using 1-ethyl-3-methylimidazolium dicyanamide ionic liquid

pvap	19.01	kPa	318.73	Isobaric vapor-liquid equilibria for the extractive distillation of 2-propanol + water mixtures using 1-ethyl-3-methylimidazolium dicyanamide ionic liquid
pvap	18.39	kPa	318.10	Isobaric vapor-liquid equilibria for the extractive distillation of 2-propanol + water mixtures using 1-ethyl-3-methylimidazolium dicyanamide ionic liquid
pvap	17.98	kPa	317.65	Isobaric vapor-liquid equilibria for the extractive distillation of 2-propanol + water mixtures using 1-ethyl-3-methylimidazolium dicyanamide ionic liquid
pvap	17.46	kPa	317.07	Isobaric vapor-liquid equilibria for the extractive distillation of 2-propanol + water mixtures using 1-ethyl-3-methylimidazolium dicyanamide ionic liquid
pvap	101.30	kPa	355.45	Isobaric (vapour + liquid) equilibrium for (2-propanol + water + ammonium thiocyanate): Fitting the data by an empirical equation

pvap	101.00	kPa	355.40	Experimental studies and thermodynamic analysis of isobaric vapor-liquid-liquid equilibria of 2-propanol + water system using n-propyl acetate and isopropyl acetate as entrainers
pvap	593.80	kPa	408.30	Isothermal vapor liquid equilibrium for binary mixtures containing furfural and its derivatives
pvap	200.10	kPa	373.50	Isothermal vapor liquid equilibrium for binary mixtures containing furfural and its derivatives
pvap	93.80	kPa	353.30	Isothermal vapor liquid equilibrium for binary mixtures containing furfural and its derivatives
pvap	101.30	kPa	355.15	Multiphase equilibria for mixtures containing water, isopropanol, propionic acid, and isopropyl propionate
pvap	144.10	kPa	364.50	Vapour liquid equilibrium for the systems butane + methanol, +2-propanol, +1-butanol, +2-butanol, +2-methyl-2-propanol at 364.5K
pvap	13.77	kPa	313.11	Vapour liquid equilibrium for the 2-methylpropane + methanol, +ethanol, +2-propanol, +2-butanol and +2-methyl-2-propanol systems at 313.15K

pvap	23.64	kPa	323.15	Vapor-Pressure Measurements of Liquid Solutions at Different Temperatures: Apparatus for Use over an Extended Temperature Range and Some New Data	
pvap	30.35	kPa	328.15	Vapor-Pressure Measurements of Liquid Solutions at Different Temperatures: Apparatus for Use over an Extended Temperature Range and Some New Data	
pvap	38.59	kPa	333.15	Vapor-Pressure Measurements of Liquid Solutions at Different Temperatures: Apparatus for Use over an Extended Temperature Range and Some New Data	
pvap	48.66	kPa	338.15	Vapor-Pressure Measurements of Liquid Solutions at Different Temperatures: Apparatus for Use over an Extended Temperature Range and Some New Data	
pvap	60.66	kPa	343.15	Vapor-Pressure Measurements of Liquid Solutions at Different Temperatures: Apparatus for Use over an Extended Temperature Range and Some New Data	

pvap	75.22	kPa	348.15	Vapor-Pressure Measurements of Liquid Solutions at Different Temperatures: Apparatus for Use over an Extended Temperature Range and Some New Data
pvap	92.44	kPa	353.15	Vapor-Pressure Measurements of Liquid Solutions at Different Temperatures: Apparatus for Use over an Extended Temperature Range and Some New Data
pvap	144.50	kPa	364.50	Isothermal Vapor Liquid Equilibrium for 2-Methylpropene + Methanol, + 1-Propanol, + 2-Propanol, + 2-Butanol, and + 2-Methyl-2-propanol Binary Systems at 364.5 K
pvap	7.93	kPa	303.15	Solubility of Carbonyl Sulfide in Aqueous Solutions of Ethylene Glycol at Temperatures from (308.15 K to 323.15) K
pvap	10.72	kPa	308.15	Solubility of Carbonyl Sulfide in Aqueous Solutions of Ethylene Glycol at Temperatures from (308.15 K to 323.15) K
pvap	101.30	kPa	355.35	Isobaric Vapor-Liquid Equilibrium for Binary Systems of Toluene + Ethanol, Toluene + Isopropanol at (101.3, 121.3, 161.3 and 201.3) kPa

pvap	121.30	kPa	359.95	Isobaric Vapor-Liquid Equilibrium for Binary Systems of Toluene + Ethanol, Toluene + Isopropanol at (101.3, 121.3, 161.3 and 201.3) kPa
pvap	161.30	kPa	367.65	Isobaric Vapor-Liquid Equilibrium for Binary Systems of Toluene + Ethanol, Toluene + Isopropanol at (101.3, 121.3, 161.3 and 201.3) kPa
pvap	201.30	kPa	373.65	Isobaric Vapor-Liquid Equilibrium for Binary Systems of Toluene + Ethanol, Toluene + Isopropanol at (101.3, 121.3, 161.3 and 201.3) kPa
pvap	60.00	kPa	342.87	Isobaric Vapor Liquid Equilibria for the 2-Propanol + Ethylene Glycol Monopropyl Ether and 2-Butanol + Ethylene Glycol Monopropyl Ether Systems at 60 kPa, 80 kPa, and 100 kPa
pvap	80.00	kPa	349.60	Isobaric Vapor Liquid Equilibria for the 2-Propanol + Ethylene Glycol Monopropyl Ether and 2-Butanol + Ethylene Glycol Monopropyl Ether Systems at 60 kPa, 80 kPa, and 100 kPa

pvap	100.00	kPa	355.06	Isobaric Vapor Liquid Equilibria for the 2-Propanol + Ethylene Glycol Monopropyl Ether and 2-Butanol + Ethylene Glycol Monopropyl Ether Systems at 60 kPa, 80 kPa, and 100 kPa	
pvap	23.60	kPa	323.19	Vapor Liquid Equilibrium for Butane + Methanol, + Ethanol, + 2-Propanol, + 2-Butanol, and + 2-Methyl-2-Propanol (TBA) at 323 K	
pvap	144.10	kPa	364.51	Vapor Liquid Equilibrium for the Systems trans-2-Butene + Methanol, + 1-Propanol, + 2-Propanol, + 2-Butanol, and + 2-Methyl-2-propanol at 364.5 K	
pvap	144.20	kPa	364.51	Vapor Liquid Equilibrium for the Systems 2-Methylpropane + Methanol, + 2-Propanol, + 2-Butanol, and + 2-Methyl-2-propanol at 364.5 K	
pvap	143.90	kPa	364.52	Vapor-Liquid Equilibrium for the cis-2-Butene + Methanol, + 2-Propanol, + 2-Butanol, + 2-Methyl-2-propanol Systems at 364.5 K	
pvap	40.00	kPa	333.94	Isobaric Vapor-Liquid Equilibria for Tetrahydropyran and Alcohol Systems	
pvap	53.33	kPa	340.18	Isobaric Vapor-Liquid Equilibria for Tetrahydropyran and Alcohol Systems	

pvap	66.66	kPa	345.25	Isobaric Vapor-Liquid Equilibria for Tetrahydropyran and Alcohol Systems
pvap	79.99	kPa	349.55	Isobaric Vapor-Liquid Equilibria for Tetrahydropyran and Alcohol Systems
pvap	93.32	kPa	353.30	Isobaric Vapor-Liquid Equilibria for Tetrahydropyran and Alcohol Systems
pvap	98.66	kPa	354.67	Isobaric Vapor-Liquid Equilibria for Tetrahydropyran and Alcohol Systems
pvap	101.32	kPa	355.65	Isobaric Vapor-Liquid Equilibrium for (Propan-2-ol + Water + 1-Butyl-3-methylimidazolium Tetrafluoroborate)
pvap	85.15	kPa	351.19	Vapor-Liquid Equilibrium Data at 343 K and Excess Molar Enthalpy Data at 298 K for the Binary Systems of Ethanol + 2,4,4-Trimethyl-1-pentene and 2-Propanol + 2,4,4-Trimethyl-1-pentene
pvap	124.03	kPa	360.57	Isobaric vapor-liquid equilibria for the extractive distillation of 2-propanol + water mixtures using 1-ethyl-3-methylimidazolium dicyanamide ionic liquid
rfi	1.37450		298.15	Densities, Viscosities, and Refractive Indices of Binary Mixtures of Acetophenone and 2-Alkanols

rfi	1.37740	293.15	Vapor liquid equilibria and excess volumes of the binary systems ethanol + ethyl lactate, isopropanol + isopropyl lactate and n-butanol + n-butyl lactate at 101.325 kPa
rfi	1.37520	298.15	Effect of anion fluorination in 1-ethyl-3-methylimidazolium as solvent for the liquid extraction of ethanol from ethyl tert-butyl ether
rfi	1.37520	298.15	Activity coefficients of the binary mixtures of a-cresol or p-cresol with C l-C4 aliphatic alcohols near ambient pressure
rfi	1.37500	298.15	Excess volumes and excess enthalpies of N-methyl-2-pyrrolidone with branched alcohols
rfi	1.37540	298.15	Effect of pressure and the capability of 2-methoxyethanol as a solvent in the behaviour of a diisopropyl ether isopropyl alcohol azeotropic mixture
rfi	1.37504	298.15	Isothermal vapour liquid equilibria in the binary and ternary systems composed of 2-propanol, 3-methyl-2-butanone and 2,2,4-trimethylpentane
rfi	1.37497	298.15	Isobaric vapour liquid equilibria for binary systems of 2-butanone with ethanol, 1-propanol, and 2-propanol at 20 and 101.3 kPa

rfi	1.37502	298.15	Isobaric vapour liquid equilibria and physical properties for isopropyl acetate + isopropanol + 1-butyl-3-methyl-imidazolium bis(trifluoromethylsulfonyl)imide mixtures
rfi	1.37760	298.15	Investigation on vapor liquid equilibrium for 2-propanol + 1-butanol + 1-pentanol at 101.3 kPa
rfi	1.37740	293.15	Vapour-liquid equilibrium of carboxylic acid alcohol binary systems:2-Propanol + butyric acid, 2-butanol + butyric acid and 2-methyl-1-propanol + butyric acid
rfi	1.37480	298.15	Excess molar volumes and partial molar volumes for (propionitrile + an alkanol) at T = 298.15 K and p = 0.1 MPa
rfi	1.37650	298.15	Physico-chemical and excess properties of the binary mixtures of methylcyclohexane + ethanol, + propan-1-ol, + propan-2-ol, + butan-1-ol, + 2-methyl-1-propanol, or 3-methyl-1-butanol at T = (298.15, 303.15, and 308.15) K

rfi	1.37410	303.15	Physico-chemical and excess properties of the binary mixtures of methylcyclohexane + ethanol, + propan-1-ol, + propan-2-ol, + butan-1-ol, + 2-methyl-1-propanol, or 3-methyl-1-butanol at T = (298.15, 303.15, and 308.15) K
rfi	1.37160	308.15	Physico-chemical and excess properties of the binary mixtures of methylcyclohexane + ethanol, + propan-1-ol, + propan-2-ol, + butan-1-ol, + 2-methyl-1-propanol, or 3-methyl-1-butanol at T = (298.15, 303.15, and 308.15) K
rfi	1.37450	298.15	Thermodynamic properties of (tetradecane + benzene, + toluene, + chlorobenzene, + bromobenzene, + anisole) binary mixtures at T = (298.15, 303.15, and 308.15) K
rfi	1.37210	303.15	Thermodynamic properties of (tetradecane + benzene, + toluene, + chlorobenzene, + bromobenzene, + anisole) binary mixtures at T = (298.15, 303.15, and 308.15) K
rfi	1.37020	308.15	Thermodynamic properties of (tetradecane + benzene, + toluene, + chlorobenzene, + bromobenzene, + anisole) binary mixtures at T = (298.15, 303.15, and 308.15) K

rfi	1.37520	298.15	(Vapor + liquid) equilibria of the binary mixtures of m-cresol with C1 C4 aliphatic alcohols at 95.5 kPa	
rfi	1.37450	298.15	Thermodynamic interactions in binary mixtures of anisole with ethanol, propan-1-ol, propan-2-ol, butan-1-ol, pentan-1-ol, and 3-methylbutan-1-ol at T = (298.15, 303.15, and 308.15) K	
rfi	1.37210	303.15	Thermodynamic interactions in binary mixtures of anisole with ethanol, propan-1-ol, propan-2-ol, butan-1-ol, pentan-1-ol, and 3-methylbutan-1-ol at T = (298.15, 303.15, and 308.15) K	
rfi	1.37020	308.15	Thermodynamic interactions in binary mixtures of anisole with ethanol, propan-1-ol, propan-2-ol, butan-1-ol, pentan-1-ol, and 3-methylbutan-1-ol at T = (298.15, 303.15, and 308.15) K	
rfi	1.37707	293.15	Mixing properties of binary mixtures presenting azeotropes at several temperatures	
rfi	1.37496	298.15	Mixing properties of binary mixtures presenting azeotropes at several temperatures	

rfi	1.37278	303.15	Mixing properties of binary mixtures presenting azeotropes at several temperatures
rfi	1.37510	298.15	(Vapor + liquid) equilibrium of the binary mixtures formed by acetonitrile with selected compounds at 95.5 kPa
rfi	1.37496	298.15	Ternary (liquid + liquid) equilibria of the azeotrope (ethyl acetate + 2-propanol) with different ionic liquids at T = 298.15 K
rfi	1.37500	298.15	Isobaric (vapour + liquid) equilibrium for N-methyl-2-pyrrolidone with branched alcohols
rfi	1.37410	298.15	Application of the ERAS model to volumetric properties of binary mixtures of banana oil with primary and secondary alcohols (C1- C4) at different temperatures
rfi	1.37500	298.15	Effect of the temperature on the physical properties of pure 1-propyl-3-methylimidazolium bis(trifluoromethylsulfonyl)imide and characterization of its binary mixtures with alcohols
rfi	1.37504	298.15	Isothermal (vapour + liquid) equilibria in the binary and ternary systems composed of 2-propanol, 2,2,4-trimethylpentane, and 2,4-dimethyl-3-pentanone

rfi	1.37518	298.15	(Vapour + liquid) equilibria for binary and ternary mixtures of 2-propanol, tetrahydropyran, and 2,2,4-trimethylpentane at P = 101.3 kPa
rfi	1.37410	303.15	Experimental study on the calorimetric data of cyclohexanol with alkanols (C1-C4) and correlation with Wilson, NRTL and UNIQUAC models at 300 K
rfi	1.37527	298.15	Densities, speeds of sound, and refractive indices for binary mixtures of 1-butyl-3-methylimidazolium methyl sulphate ionic liquid with alcohols at T = (298.15, 303.15, 308.15, and 313.15) K
rfi	1.37314	303.15	Densities, speeds of sound, and refractive indices for binary mixtures of 1-butyl-3-methylimidazolium methyl sulphate ionic liquid with alcohols at T = (298.15, 303.15, 308.15, and 313.15) K
rfi	1.37093	308.15	Densities, speeds of sound, and refractive indices for binary mixtures of 1-butyl-3-methylimidazolium methyl sulphate ionic liquid with alcohols at T = (298.15, 303.15, 308.15, and 313.15) K

rfi	1.36870	313.15	Densities, speeds of sound, and refractive indices for binary mixtures of 1-butyl-3-methylimidazolium methyl sulphate ionic liquid with alcohols at T = (298.15, 303.15, 308.15, and 313.15) K
rfi	1.37523	298.15	Effect of the temperature on the physical properties of the pure ionic liquid 1-ethyl-3-methylimidazolium methylsulfate and characterization of its binary mixtures with alcohols
rfi	1.37560	298.15	Ternary (liquid + liquid) equilibria for the extraction of ethanol, or 2-propanol from aqueous solutions with 1,1'-oxybis(butane) at different temperatures
rfi	1.37660	293.15	Densities, viscosities, excess molar volumes, and refractive indices of acetonitrile and 2-alkanols binary mixtures at different temperatures: Experimental results and application of the Prigogine Flory Patterson theory
rfi	1.37450	298.15	Densities, viscosities, excess molar volumes, and refractive indices of acetonitrile and 2-alkanols binary mixtures at different temperatures: Experimental results and application of the Prigogine Flory Patterson theory

rfi	1.37240	303.15	Densities, viscosities, excess molar volumes, and refractive indices of acetonitrile and 2-alkanols binary mixtures at different temperatures: Experimental results and application of the Prigogine Flory Patterson theory
rfi	1.37030	308.15	Densities, viscosities, excess molar volumes, and refractive indices of acetonitrile and 2-alkanols binary mixtures at different temperatures: Experimental results and application of the Prigogine Flory Patterson theory
rfi	1.37360	303.15	Solute-Solvent and Solvent-Solvent Interactions of Menthol in Isopropyl Alcohol and its Binary Mixtures with Methyl Salicylate by Volumetric, Viscometric, Interferometric and Refractive Index Techniques.
rfi	1.37530	298.15	Molar excess enthalpies and molar excess volumes of formamide + 1-propanol or 2-propanol and thermodynamic modeling by Prigogine-Flory-Patterson theory and Treszczanowicz-Benson association model

rfi	1.37710	293.20	Isobaric Vapor Liquid Equilibria for Binary Mixtures of .gamma.-Valerolactone + Methanol, Ethanol, and 2-Propanol
rfi	1.37510	298.20	Isobaric Vapor Liquid Equilibria for Binary Mixtures of .gamma.-Valerolactone + Methanol, Ethanol, and 2-Propanol
rfi	1.37718	293.15	Vapor Liquid Equilibrium for Methyl Isobutyl Ketone (MIBK) + (1-Propanol or 2-Propanol) Binary Mixtures
rfi	1.37740	293.15	Excess Properties and Phase Equilibria for the Potential Biofuel System of Propan-2-ol and 2-Methyl-propan-1-ol at 333.15, 343.15, and 353.15 K
rfi	1.37580	298.00	Determination of Physicochemical Parameters of Sodium Dodecyl Sulfate in Aqueous Micellar Solutions Containing Short-Chain Alcohols
rfi	1.37740	293.15	Solid-Liquid Equilibrium Measurements for Posaconazole and Voriconazole in Several Solvents between T = 278.2 and 323.2 K Using Differential Thermal Analysis/Thermal Gravimetric Analysis

rfi	1.38000	298.15	Bubble Temperatures of the Binary Mixtures of Dimethylcarbonate with Some Alcohols at 95.8 kPa
rfi	1.37500	298.15	Vapor-Liquid Equilibria Data for Methanol + 2-Propanol+ 2-Methyl-2-butanol and Constituent Binary Systems at 101.3 kPa
rfi	1.37707	293.15	Thermodynamic Properties of Ionic Liquids in Organic Solvents from (293.15 to 303.15) K
rfi	1.37496	298.15	Thermodynamic Properties of Ionic Liquids in Organic Solvents from (293.15 to 303.15) K
rfi	1.37278	303.15	Thermodynamic Properties of Ionic Liquids in Organic Solvents from (293.15 to 303.15) K
rfi	1.37500	298.15	Densities, Excess Molar Volumes, and Refractive Properties of the Binary Mixtures of the Amino Acid Ionic Liquid [bmim][Gly] with 1-Butanol or Isopropanol at T = (298.15 to 313.15) K
rfi	1.37400	303.15	Densities, Excess Molar Volumes, and Refractive Properties of the Binary Mixtures of the Amino Acid Ionic Liquid [bmim][Gly] with 1-Butanol or Isopropanol at T = (298.15 to 313.15) K

rfi	1.37300	308.15	Densities, Excess Molar Volumes, and Refractive Properties of the Binary Mixtures of the Amino Acid Ionic Liquid [bmim][Gly] with 1-Butanol or Isopropanol at T = (298.15 to 313.15) K
rfi	1.37200	313.15	Densities, Excess Molar Volumes, and Refractive Properties of the Binary Mixtures of the Amino Acid Ionic Liquid [bmim][Gly] with 1-Butanol or Isopropanol at T = (298.15 to 313.15) K
rfi	1.37709	293.15	Liquid Liquid Equilibria of Methanol, Ethanol, and Propan-2-ol with Water and Dodecane
rfi	1.37510	298.15	Physical Properties of Binary and Ternary Mixtures of 2-Propanol, Water, and 1-Butyl-3-methylimidazolium Tetrafluoroborate Ionic Liquid
rfi	1.36040	323.15	Physical Properties of Binary and Ternary Mixtures of 2-Propanol, Water, and 1-Butyl-3-methylimidazolium Tetrafluoroborate Ionic Liquid
rfi	1.37710	293.20	Vapor Liquid Equilibrium Data for Binary Systems of 1-Methyl-4-(1-methylethenyl)-cyclohexene + {Ethanol, Propan-1-ol, Propan-2-ol, Butan-1-ol, Pentan-1-ol, or Hexan-1-ol} at 40 kPa

rfi	1.37880	295.15	Isobaric Vapor Liquid Equilibrium Data for Binary Mixtures of 1-Phenylethanone with a Few Alcohols at 95.2 kPa
rfi	1.35540	298.15	Vapor Liquid Equilibrium for Ternary and Binary Mixtures of 2-Isopropoxypropane, 2-Propanol, and N,N-Dimethylacetamide at 101.3 kPa
rfi	1.37495	298.15	Physical Properties of Binary Mixtures of the Ionic Liquid 1-Ethyl-3-methylimidazolium Ethyl Sulfate with Several Alcohols at T = (298.15, 313.15, and 328.15) K and Atmospheric Pressure
rfi	1.37496	298.15	Ternary Liquid-Liquid Equilibria Ethanol + 2-Butanone + 1-Butyl-3-methylimidazolium Hexafluorophosphate, 2-Propanol + 2-Butanone + 1-Butyl-3-methylimidazolium Hexafluorophosphate, and 2-Butanone + 2-Propanol + 1,3-Dimethylimidazolium Methyl Sulfate at 298.15 K
rfi	1.37520	298.15	Vapor-Liquid Equilibrium and Excess Gibbs Energies of Hexane + N,N-Dimethyl Formamide, 2-Methylpropan-2-ol + 2-Aminophenol, N,N-Dimethyl Formamide, and 2-Propanol + Diisopropyl Amine at 94.4 kPa

rfi	1.37500	298.15	Density and Viscosity Experimental Data of the Ternary Mixtures 1-Propanol or 2-Propanol + Water + 1-Ethyl-3-methylimidazolium Ethylsulfate. Correlation and Prediction of Physical Properties of the Ternary Systems
rfi	1.37063	308.15	Densities, Viscosities, Refractive Indices, and Surface Tensions for the Ternary Mixtures of 2-Propanol + Benzyl Alcohol + 2-Phenylethanol at T = 308.15 K
rfi	1.37540	298.15	Isobaric Vapor-Liquid Equilibria for Binary and Ternary Mixtures of Diisopropyl Ether, 2-Propyl Alcohol, and 3-Methyl-1-Butanol
rfi	1.37800	293.15	Measurement and Correlation of the Solubilities of m-Phthalic Acid in Monobasic Alcohols
rfi	1.37540	298.15	sobaric Vapor-Liquid Equilibria for Binary and Ternary Mixtures of Diisopropyl Ether, 2-Propyl Alcohol, and n-Butyl Propionate at 101.3 kPa
rfi	1.37509	298.15	Vapor-Liquid Equilibria for Binary and Ternary Mixtures of 1,3-Dioxolane, 2-Propanol, and 2,2,4-Trimethylpentane at 101.3 kPa

rfi	1.37540		298.15	Isobaric Vapor-Liquid Equilibria for Binary and Ternary Mixtures of Ethanol and 2-Propanol with 2-Butanone and Butyl Propionate at 101.3 kPa
rhol	780.99	kg/m3	298.15	Volumetric properties of binary liquid mixtures of alcohols with 1,2-dichloroethane at different temperatures and atmospheric pressure
rhol	758.60	kg/m3	323.15	Densities and Excess Molar Volumes for Binary Glycerol + 1-Propanol, + 2-Propanol, + 1,2-Propanediol, and + 1,3-Propanediol Mixtures at Different Temperatures
rhol	753.80	kg/m3	328.15	Densities and Excess Molar Volumes for Binary Glycerol + 1-Propanol, + 2-Propanol, + 1,2-Propanediol, and + 1,3-Propanediol Mixtures at Different Temperatures
rhol	748.90	kg/m3	333.15	Densities and Excess Molar Volumes for Binary Glycerol + 1-Propanol, + 2-Propanol, + 1,2-Propanediol, and + 1,3-Propanediol Mixtures at Different Temperatures

rhol	789.34	kg/m3	288.15	Excess Molar Volumes of 1,3-Diethyl Propanedioate with Methanol, Ethanol, Propan-1-ol, Propan-2-ol, Butan-2-ol, 2-Methyl-propan-1-ol, and Pentan-1-ol at T = (288.15, 298.15, 313.15, and 328.15) K
rhol	780.99	kg/m3	298.15	Excess Molar Volumes of 1,3-Diethyl Propanedioate with Methanol, Ethanol, Propan-1-ol, Propan-2-ol, Butan-2-ol, 2-Methyl-propan-1-ol, and Pentan-1-ol at T = (288.15, 298.15, 313.15, and 328.15) K
rhol	767.92	kg/m3	313.15	Excess Molar Volumes of 1,3-Diethyl Propanedioate with Methanol, Ethanol, Propan-1-ol, Propan-2-ol, Butan-2-ol, 2-Methyl-propan-1-ol, and Pentan-1-ol at T = (288.15, 298.15, 313.15, and 328.15) K
rhol	754.00	kg/m3	328.15	Excess Molar Volumes of 1,3-Diethyl Propanedioate with Methanol, Ethanol, Propan-1-ol, Propan-2-ol, Butan-2-ol, 2-Methyl-propan-1-ol, and Pentan-1-ol at T = (288.15, 298.15, 313.15, and 328.15) K
rhol	781.05	kg/m3	298.15	Densities, Viscosities, and Speeds of Sound of Binary Liquid Mixtures of Ethylenediamine with Alcohols at T = (293.15 to 313.15) K

rhol	781.23	kg/m3	298.15	ACSExcess Molar Enthalpies of Mixtures of (+-)-Linalool with Several Alkanols
rhol	781.60	kg/m3	298.15	Thermophysical Properties of the Pure Ionic Liquid 1-Butyl-1-methylpyrrolidinium Dicyanamide and Its Binary Mixtures with Alcohols
rhol	785.35	kg/m3	293.15	Density and Viscosity Measurements of Binary Alkanol Mixtures from (293.15 to 333.15) K at Atmospheric Pressure
rhol	776.86	kg/m3	303.15	Density and Viscosity Measurements of Binary Alkanol Mixtures from (293.15 to 333.15) K at Atmospheric Pressure
rhol	768.07	kg/m3	313.15	Density and Viscosity Measurements of Binary Alkanol Mixtures from (293.15 to 333.15) K at Atmospheric Pressure
rhol	758.89	kg/m3	323.15	Density and Viscosity Measurements of Binary Alkanol Mixtures from (293.15 to 333.15) K at Atmospheric Pressure
rhol	749.25	kg/m3	333.15	Density and Viscosity Measurements of Binary Alkanol Mixtures from (293.15 to 333.15) K at Atmospheric Pressure

rhol	794.00	kg/m3	283.15	Densities and Excess Properties of Primary Amines in Alcoholic Solutions	
rhol	785.76	kg/m3	293.15	Densities and Excess Properties of Primary Amines in Alcoholic Solutions	
rhol	777.28	kg/m3	303.15	Densities and Excess Properties of Primary Amines in Alcoholic Solutions	
rhol	768.48	kg/m3	313.15	Densities and Excess Properties of Primary Amines in Alcoholic Solutions	
rhol	759.31	kg/m3	323.15	Densities and Excess Properties of Primary Amines in Alcoholic Solutions	
rhol	749.68	kg/m3	333.15	Densities and Excess Properties of Primary Amines in Alcoholic Solutions	
rhol	739.53	kg/m3	343.15	Densities and Excess Properties of Primary Amines in Alcoholic Solutions	
rhol	786.00	kg/m3	293.15	Composition and Temperature Dependence of Density, Surface Tension, and Viscosity of EMIM DEP/MMIM DMP + Water + 1-Propanol/2-Propanol Ternary Mixtures and Their Mathematical Representation Using the Jouyban Acree Model	

rhol	781.00	kg/m3	298.15	Composition and Temperature Dependence of Density, Surface Tension, and Viscosity of EMIM DEP/MMIM DMP + Water + 1-Propanol/2-Propanol Ternary Mixtures and Their Mathematical Representation Using the Jouyban Acree Model
rhol	777.00	kg/m3	303.15	Composition and Temperature Dependence of Density, Surface Tension, and Viscosity of EMIM DEP/MMIM DMP + Water + 1-Propanol/2-Propanol Ternary Mixtures and Their Mathematical Representation Using the Jouyban Acree Model
rhol	768.00	kg/m3	313.15	Composition and Temperature Dependence of Density, Surface Tension, and Viscosity of EMIM DEP/MMIM DMP + Water + 1-Propanol/2-Propanol Ternary Mixtures and Their Mathematical Representation Using the Jouyban Acree Model

rhol	759.00	kg/m3	323.15	Composition and Temperature Dependence of Density, Surface Tension, and Viscosity of EMIM DEP/MMIM DMP + Water + 1-Propanol/2-Propanol Ternary Mixtures and Their Mathematical Representation Using the Jouyban Acree Model
rhol	750.00	kg/m3	333.15	Composition and Temperature Dependence of Density, Surface Tension, and Viscosity of EMIM DEP/MMIM DMP + Water + 1-Propanol/2-Propanol Ternary Mixtures and Their Mathematical Representation Using the Jouyban Acree Model
rhol	779.36	kg/m3	298.15	Vapor Liquid Equilibrium of Mixtures Containing the Following Higher Alcohols: 2-Propanol, 2-Methyl-1-propanol, and 3-Methyl-1-butanol
rhol	781.10	kg/m3	298.15	Thermodynamic Properties of Binary Mixtures Containing N,N-Dimethylacetamide + 2-Alkanol: Experimental Data and Modeling
rhol	776.80	kg/m3	303.15	Thermodynamic Properties of Binary Mixtures Containing N,N-Dimethylacetamide + 2-Alkanol: Experimental Data and Modeling

rhol	772.40	kg/m3	308.15	Thermodynamic Properties of Binary Mixtures Containing N,N-Dimethylacetamide + 2-Alkanol: Experimental Data and Modeling
rhol	767.80	kg/m3	313.15	Thermodynamic Properties of Binary Mixtures Containing N,N-Dimethylacetamide + 2-Alkanol: Experimental Data and Modeling
rhol	785.24	kg/m3	293.15	Measurement and Correlation of the Excess Properties of Ternary Mixture of {x1[Hmim][BF4] + x21-Propanol + x32-Propanol} at Different Temperatures
rhol	776.76	kg/m3	303.15	Measurement and Correlation of the Excess Properties of Ternary Mixture of {x1[Hmim][BF4] + x21-Propanol + x32-Propanol} at Different Temperatures
rhol	767.96	kg/m3	313.15	Measurement and Correlation of the Excess Properties of Ternary Mixture of {x1[Hmim][BF4] + x21-Propanol + x32-Propanol} at Different Temperatures
rhol	758.79	kg/m3	323.15	Measurement and Correlation of the Excess Properties of Ternary Mixture of {x1[Hmim][BF4] + x21-Propanol + x32-Propanol} at Different Temperatures

rhol	749.16	kg/m3	333.15	Measurement and Correlation of the Excess Properties of Ternary Mixture of {x1[Hmim][BF4] + x21-Propanol + x32-Propanol} at Different Temperatures
rhol	785.30	kg/m3	293.15	Densities and Viscosities of Binary Mixtures Containing 1,3-Dimethylimidazolium Dimethylphosphate and Alcohols
rhol	781.09	kg/m3	298.15	Densities and Viscosities of Binary Mixtures Containing 1,3-Dimethylimidazolium Dimethylphosphate and Alcohols
rhol	776.80	kg/m3	303.15	Densities and Viscosities of Binary Mixtures Containing 1,3-Dimethylimidazolium Dimethylphosphate and Alcohols
rhol	772.43	kg/m3	308.15	Densities and Viscosities of Binary Mixtures Containing 1,3-Dimethylimidazolium Dimethylphosphate and Alcohols
rhol	767.99	kg/m3	313.15	Densities and Viscosities of Binary Mixtures Containing 1,3-Dimethylimidazolium Dimethylphosphate and Alcohols
rhol	763.47	kg/m3	318.15	Densities and Viscosities of Binary Mixtures Containing 1,3-Dimethylimidazolium Dimethylphosphate and Alcohols
rhol	758.85	kg/m3	323.15	Densities and Viscosities of Binary Mixtures Containing 1,3-Dimethylimidazolium Dimethylphosphate and Alcohols

rhol	777.50	kg/m3	303.15	Density, Viscosities, and Excess Properties for Binary Mixtures of Sulfolane + Alcohols and Sulfolane + Glycols at Different Temperatures	
rhol	768.80	kg/m3	313.15	Density, Viscosities, and Excess Properties for Binary Mixtures of Sulfolane + Alcohols and Sulfolane + Glycols at Different Temperatures	
rhol	759.70	kg/m3	323.15	Density, Viscosities, and Excess Properties for Binary Mixtures of Sulfolane + Alcohols and Sulfolane + Glycols at Different Temperatures	
rhol	750.00	kg/m3	333.15	Density, Viscosities, and Excess Properties for Binary Mixtures of Sulfolane + Alcohols and Sulfolane + Glycols at Different Temperatures	
rhol	739.50	kg/m3	343.15	Density, Viscosities, and Excess Properties for Binary Mixtures of Sulfolane + Alcohols and Sulfolane + Glycols at Different Temperatures	
rhol	785.40	kg/m3	293.15	Densities and Viscosities of Binary Mixtures Containing Ethyl Formate and 2-Alkanols: Friction Theory and Free Volume Theory	

rhol	781.10	kg/m3	298.15	Densities and Viscosities of Binary Mixtures Containing Ethyl Formate and 2-Alkanols: Friction Theory and Free Volume Theory	
rhol	776.80	kg/m3	303.15	Densities and Viscosities of Binary Mixtures Containing Ethyl Formate and 2-Alkanols: Friction Theory and Free Volume Theory	
rhol	772.40	kg/m3	308.15	Densities and Viscosities of Binary Mixtures Containing Ethyl Formate and 2-Alkanols: Friction Theory and Free Volume Theory	
rhol	768.00	kg/m3	313.15	Densities and Viscosities of Binary Mixtures Containing Ethyl Formate and 2-Alkanols: Friction Theory and Free Volume Theory	
rhol	763.40	kg/m3	318.15	Densities and Viscosities of Binary Mixtures Containing Ethyl Formate and 2-Alkanols: Friction Theory and Free Volume Theory	
rhol	758.80	kg/m3	323.15	Densities and Viscosities of Binary Mixtures Containing Ethyl Formate and 2-Alkanols: Friction Theory and Free Volume Theory	

rhol	781.32	kg/m3	298.15	Liquid Liquid Equilibrium for Ternary Systems of Propyl Vinyl Ether + C3 or C4 Alcohols + Water at 298.15 K and Excess Molar Enthalpies for Ternary and Constituent Binary Systems of Propyl Vinyl Ether + Ethanol + Isooctane at 303.15 K
rhol	781.40	kg/m3	298.15	Liquid-Liquid Equilibrium for 2,2,2-Trifluoroethanol + Ethanol + Cyclohexane from (288.15 to 308.15) K
rhol	785.10	kg/m3	293.15	Volumetric Properties of Binary and Ternary Liquid Mixtures of 1-Propanol (1) + 2-Propanol (2) + Water (3) at Different Temperatures and Ambient Pressure (81.5 kPa)
rhol	776.63	kg/m3	303.15	Volumetric Properties of Binary and Ternary Liquid Mixtures of 1-Propanol (1) + 2-Propanol (2) + Water (3) at Different Temperatures and Ambient Pressure (81.5 kPa)
rhol	767.84	kg/m3	313.15	Volumetric Properties of Binary and Ternary Liquid Mixtures of 1-Propanol (1) + 2-Propanol (2) + Water (3) at Different Temperatures and Ambient Pressure (81.5 kPa)

rhol	758.68	kg/m3	323.15	Volumetric Properties of Binary and Ternary Liquid Mixtures of 1-Propanol (1) + 2-Propanol (2) + Water (3) at Different Temperatures and Ambient Pressure (81.5 kPa)
rhol	781.40	kg/m3	298.15	Liquid-Liquid Equilibrium of (Cyclohexane + 2,2,2-Trifluoroethanol) and (Cyclohexane + Methanol) from (278.15 to 318.15) K
rhol	781.40	kg/m3	298.15	Effect of Temperature on Phase Equilibrium of the Mixed-Solvent System of (2,2,2-Trifluoroethanol + Methanol + Cyclohexane)
rhol	781.35	kg/m3	298.15	Binary Liquid-Liquid Equilibrium (LLE) for Dibutyl Ether (DBE) + Water from (288.15 to 318.15) K and Ternary LLE for Systems of DBE + C1 !less thanless than C4 Alcohols + Water at 298.15 K
rhol	781.34	kg/m3	298.15	Binary Liquid-Liquid Equilibrium (LLE) for Methyl tert-Amyl Ether (TAME) + Water from (288.15 to 313.15) K and Ternary LLE for Systems of TAME + C1-C4 Alcohols + Water at 298.15 K

rhol	781.50	kg/m3	298.15	Solubility and Liquid-Liquid Equilibrium of Aqueous Systems of Iodoethane with Methanol, Ethanol, or 1-Propanol at Temperature 298.15 K and Pressure 101.2 kPa
rhol	780.90	kg/m3	298.15	Apparent Molal Volumes and Viscosity B-Coefficients of Acetyl Salicylic Acid (2-Acetoxy Benzoic Acid) Solutions in Higher Alcohols at Different Temperatures
rhol	780.89	kg/m3	298.15	Refractive Indices and Deviations in Refractive Indices for Binary Mixtures of 1-Ethyl-3-methylimidazolium Trifluoromethanesulfonate with Methanol, Ethanol, 1-Propanol, and 2-Propanol at Several Temperatures
rhol	780.98	kg/m3	298.15	Topological investigations of the molecular species and molecular interactions that characterize pyrrolidin-2-one + lower alkanol mixtures
rhol	780.35	kg/m3	298.15	Excess enthalpies of binary mixtures of some propylamines + some propanols at 298.15K

rhol	749.39	kg/m3	333.15	Properties of pure 1-(3-(trimethylammonio)prop-1-yl)-3-methylimidazolium bis(dicyanamide) asymmetrical gemini ionic liquid and its binary mixture with isopropanol at T = (283.15 333.15) K
rhol	754.26	kg/m3	328.15	Properties of pure 1-(3-(trimethylammonio)prop-1-yl)-3-methylimidazolium bis(dicyanamide) asymmetrical gemini ionic liquid and its binary mixture with isopropanol at T = (283.15 333.15) K
rhol	759.01	kg/m3	323.15	Properties of pure 1-(3-(trimethylammonio)prop-1-yl)-3-methylimidazolium bis(dicyanamide) asymmetrical gemini ionic liquid and its binary mixture with isopropanol at T = (283.15 333.15) K
rhol	763.65	kg/m3	318.15	Properties of pure 1-(3-(trimethylammonio)prop-1-yl)-3-methylimidazolium bis(dicyanamide) asymmetrical gemini ionic liquid and its binary mixture with isopropanol at T = (283.15 333.15) K
rhol	768.18	kg/m3	313.15	Properties of pure 1-(3-(trimethylammonio)prop-1-yl)-3-methylimidazolium bis(dicyanamide) asymmetrical gemini ionic liquid and its binary mixture with isopropanol at T = (283.15 333.15) K

rhol	772.62	kg/m3	308.15	Properties of pure 1-(3-(trimethylammonio)prop-1-yl)-3-methylimidazolium bis(dicyanamide) asymmetrical gemini ionic liquid and its binary mixture with isopropanol at T = (283.15 333.15) K
rhol	776.97	kg/m3	303.15	Properties of pure 1-(3-(trimethylammonio)prop-1-yl)-3-methylimidazolium bis(dicyanamide) asymmetrical gemini ionic liquid and its binary mixture with isopropanol at T = (283.15 333.15) K
rhol	781.24	kg/m3	298.15	Properties of pure 1-(3-(trimethylammonio)prop-1-yl)-3-methylimidazolium bis(dicyanamide) asymmetrical gemini ionic liquid and its binary mixture with isopropanol at T = (283.15 333.15) K
rhol	785.44	kg/m3	293.15	Properties of pure 1-(3-(trimethylammonio)prop-1-yl)-3-methylimidazolium bis(dicyanamide) asymmetrical gemini ionic liquid and its binary mixture with isopropanol at T = (283.15 333.15) K
rhol	789.58	kg/m3	288.15	Properties of pure 1-(3-(trimethylammonio)prop-1-yl)-3-methylimidazolium bis(dicyanamide) asymmetrical gemini ionic liquid and its binary mixture with isopropanol at T = (283.15 333.15) K

rhol	793.67	kg/m3	283.15	Properties of pure 1-(3-(trimethylammonio)prop-1-yl)-3-methylimidazolium bis(dicyanamide) asymmetrical gemini ionic liquid and its binary mixture with isopropanol at T = (283.15 333.15) K
rhol	780.95	kg/m3	298.15	Excess volumes and partial molar volumes of binary liquid mixtures of furfural or 2-methylfuran with alcohols at 298.15 K
rhol	758.55	kg/m3	323.15	Volumetric properties of monoethanolamine and alcohol binary mixtures at different temperatures and 0.1 MPa
rhol	763.19	kg/m3	318.15	Volumetric properties of monoethanolamine and alcohol binary mixtures at different temperatures and 0.1 MPa
rhol	767.82	kg/m3	313.15	Volumetric properties of monoethanolamine and alcohol binary mixtures at different temperatures and 0.1 MPa
rhol	772.24	kg/m3	308.15	Volumetric properties of monoethanolamine and alcohol binary mixtures at different temperatures and 0.1 MPa
rhol	776.57	kg/m3	303.15	Volumetric properties of monoethanolamine and alcohol binary mixtures at different temperatures and 0.1 MPa

rhol	780.91	kg/m3	298.15	Volumetric properties of monoethanolamine and alcohol binary mixtures at different temperatures and 0.1 MPa
rhol	785.10	kg/m3	293.15	Volumetric properties of monoethanolamine and alcohol binary mixtures at different temperatures and 0.1 MPa
rhol	753.91	kg/m3	328.15	Mass density, sound velocity, mixing enthalpy, 1H NMR, Ab initio calculations and intermolecular interactions in binary mixtures of N-methylimidazole + water, +methanol, +ethanol, +1-propanol, +2-propanol
rhol	763.30	kg/m3	318.15	Mass density, sound velocity, mixing enthalpy, 1H NMR, Ab initio calculations and intermolecular interactions in binary mixtures of N-methylimidazole + water, +methanol, +ethanol, +1-propanol, +2-propanol
rhol	772.26	kg/m3	308.15	Mass density, sound velocity, mixing enthalpy, 1H NMR, Ab initio calculations and intermolecular interactions in binary mixtures of N-methylimidazole + water, +methanol, +ethanol, +1-propanol, +2-propanol

rhol	780.86	kg/m3	298.15	Mass density, sound velocity, mixing enthalpy, 1H NMR, Ab initio calculations and intermolecular interactions in binary mixtures of N-methylimidazole + water, +methanol, +ethanol, +1-propanol, +2-propanol	
rhol	789.18	kg/m3	288.15	Mass density, sound velocity, mixing enthalpy, 1H NMR, Ab initio calculations and intermolecular interactions in binary mixtures of N-methylimidazole + water, +methanol, +ethanol, +1-propanol, +2-propanol	
rhol	780.69	kg/m3	298.15	Liquid-liquid equilibria and density data for pseudoternary systems of refined soybean oil + (hexanal, or heptanal, or butyric acid, or valeric acid, or caproic acid, or caprylic acid) + dimethyl sulfoxide at 298.15 K	
rhol	781.00	kg/m3	298.15	Density, speed of sound and refractive index of mixtures containing 2-phenoxyethanol with propanol or butanol at various temperatures	

rhol	759.00	kg/m3	323.15	Density, speed of sound and refractive index of mixtures containing 2-phenoxyethanol with propanol or butanol at various temperatures
rhol	768.00	kg/m3	313.15	Density, speed of sound and refractive index of mixtures containing 2-phenoxyethanol with propanol or butanol at various temperatures
rhol	777.00	kg/m3	303.15	Density, speed of sound and refractive index of mixtures containing 2-phenoxyethanol with propanol or butanol at various temperatures
rhol	785.00	kg/m3	293.15	Density, speed of sound and refractive index of mixtures containing 2-phenoxyethanol with propanol or butanol at various temperatures
rhol	754.10	kg/m3	328.15	Thermophysical properties of binary mixtures of 1-butyl-1-methylpyrrolidinium trifluoromethanesulfonate ionic liquid with alcohols at several temperatures
rhol	763.40	kg/m3	318.15	Thermophysical properties of binary mixtures of 1-butyl-1-methylpyrrolidinium trifluoromethanesulfonate ionic liquid with alcohols at several temperatures

rhol	772.40	kg/m3	308.15	Thermophysical properties of binary mixtures of 1-butyl-1-methylpyrrolidinium trifluoromethanesulfonate ionic liquid with alcohols at several temperatures
rhol	781.00	kg/m3	298.15	Thermophysical properties of binary mixtures of 1-butyl-1-methylpyrrolidinium trifluoromethanesulfonate ionic liquid with alcohols at several temperatures
rhol	789.40	kg/m3	288.15	Thermophysical properties of binary mixtures of 1-butyl-1-methylpyrrolidinium trifluoromethanesulfonate ionic liquid with alcohols at several temperatures
rhol	797.50	kg/m3	278.15	Thermophysical properties of binary mixtures of 1-butyl-1-methylpyrrolidinium trifluoromethanesulfonate ionic liquid with alcohols at several temperatures
rhol	763.48	kg/m3	318.15	Excess molar volume and excess Gibbs energy of activation for viscous flow for the binary mixtures of N-ethylpyridinium dicyanamide [C2py][DCA] with alcohols
rhol	768.02	kg/m3	313.15	Excess molar volume and excess Gibbs energy of activation for viscous flow for the binary mixtures of N-ethylpyridinium dicyanamide [C2py][DCA] with alcohols

rhol	772.46	kg/m3	308.15	Excess molar volume and excess Gibbs energy of activation for viscous flow for the binary mixtures of N-ethylpyridinium dicyanamide [C2py][DCA] with alcohols	
rhol	776.81	kg/m3	303.15	Excess molar volume and excess Gibbs energy of activation for viscous flow for the binary mixtures of N-ethylpyridinium dicyanamide [C2py][DCA] with alcohols	
rhol	781.08	kg/m3	298.15	Excess molar volume and excess Gibbs energy of activation for viscous flow for the binary mixtures of N-ethylpyridinium dicyanamide [C2py][DCA] with alcohols	
rhol	785.29	kg/m3	293.15	Excess molar volume and excess Gibbs energy of activation for viscous flow for the binary mixtures of N-ethylpyridinium dicyanamide [C2py][DCA] with alcohols	
rhol	789.41	kg/m3	288.15	Excess molar volume and excess Gibbs energy of activation for viscous flow for the binary mixtures of N-ethylpyridinium dicyanamide [C2py][DCA] with alcohols	

rhol	753.12	kg/m3	328.15	The study of physico-chemical properties of binary systems consisting of N-Methylcyclohexylamine with 2-alkanols at T = (298.15-328.15) K
rhol	763.41	kg/m3	318.15	The study of physico-chemical properties of binary systems consisting of N-Methylcyclohexylamine with 2-alkanols at T = (298.15-328.15) K
rhol	772.41	kg/m3	308.15	The study of physico-chemical properties of binary systems consisting of N-Methylcyclohexylamine with 2-alkanols at T = (298.15-328.15) K
rhol	781.05	kg/m3	298.15	The study of physico-chemical properties of binary systems consisting of N-Methylcyclohexylamine with 2-alkanols at T = (298.15-328.15) K
rhol	780.85	kg/m3	298.20	A green process for recovery of 1-propanol/2-propanol from their aqueous solutions: Experimental and MD simulation studies
rhol	769.70	kg/m3	313.15	Investigation on molecular interactions of antibiotics in alcohols using volumetric and acoustic studies at different temperatures

rhol	772.02	kg/m3	308.15	Investigation on molecular interactions of antibiotics in alcohols using volumetric and acoustic studies at different temperatures
rhol	777.06	kg/m3	303.15	Investigation on molecular interactions of antibiotics in alcohols using volumetric and acoustic studies at different temperatures
rhol	781.34	kg/m3	298.15	Investigation on molecular interactions of antibiotics in alcohols using volumetric and acoustic studies at different temperatures
rhol	785.81	kg/m3	293.15	Investigation on molecular interactions of antibiotics in alcohols using volumetric and acoustic studies at different temperatures
rhol	781.11	kg/m3	298.15	Experimental study on the calorimetric data of 2-butoxyethanol with aliphatic alcohols (C1-C4) and correlation with the Wilson, NRTL and UNIQUAC models at T = 298 K
rhol	781.20	kg/m3	298.15	Measurement and correlation of solubility and solution thermodynamics of 1,3-dimethylurea in different solvents from T = (288.15 to 328.15) K

rhol	781.10	kg/m3	298.15	Solubility and solution thermodynamics of thymol in six pure organic solvents
rhol	781.20	kg/m3	298.15	Excess molar enthalpies of R-fenchone + propan-1-ol or +propan-2-ol. Modeling with COSMO-RS and UNIFAC
rhol	758.80	kg/m3	323.15	Thermodynamic and transport properties of binary mixtures; friction theory coupled with PC-SAFT model
rhol	763.40	kg/m3	318.15	Thermodynamic and transport properties of binary mixtures; friction theory coupled with PC-SAFT model
rhol	768.00	kg/m3	313.15	Thermodynamic and transport properties of binary mixtures; friction theory coupled with PC-SAFT model
rhol	772.40	kg/m3	308.15	Thermodynamic and transport properties of binary mixtures; friction theory coupled with PC-SAFT model
rhol	776.80	kg/m3	303.15	Thermodynamic and transport properties of binary mixtures; friction theory coupled with PC-SAFT model
rhol	781.10	kg/m3	298.15	Thermodynamic and transport properties of binary mixtures; friction theory coupled with PC-SAFT model
rhol	785.40	kg/m3	293.15	Thermodynamic and transport properties of binary mixtures; friction theory coupled with PC-SAFT model

rhol	780.60	kg/m3	298.15	Solubility and solution thermodynamics of sorbic acid in eight pure organic solvents
rhol	780.80	kg/m3	298.15	Measurements and modeling of LLE and HE for (methanol + 2,4,4-trimethyl-1-pentene), and LLE for (water + methanol + 2,4,4-trimethyl-1-pentene)
rhol	780.90	kg/m3	298.15	(Liquid + liquid) equilibria for mixtures of dodecane and ethanol with alkylsulfate-based ionic liquids
rhol	780.90	kg/m3	298.15	A combined experimental and computational investigation of excess molar enthalpies of (nitrobenzene + alkanol) mixtures
rhol	768.45	kg/m3	313.15	Molecular interactions in binary mixtures of 1-butoxy-2-propanol with alcohols at different temperatures: A thermophysical and spectroscopic approach
rhol	773.15	kg/m3	308.15	Molecular interactions in binary mixtures of 1-butoxy-2-propanol with alcohols at different temperatures: A thermophysical and spectroscopic approach

rhol	777.47	kg/m3	303.15	Molecular interactions in binary mixtures of 1-butoxy-2-propanol with alcohols at different temperatures: A thermophysical and spectroscopic approach
rhol	781.59	kg/m3	298.15	Molecular interactions in binary mixtures of 1-butoxy-2-propanol with alcohols at different temperatures: A thermophysical and spectroscopic approach
rhol	785.74	kg/m3	293.15	Molecular interactions in binary mixtures of 1-butoxy-2-propanol with alcohols at different temperatures: A thermophysical and spectroscopic approach
rhol	767.76	kg/m3	313.15	Densities and viscosities of the mixtures (formamide + 2-alkanol): Experimental and theoretical approaches
rhol	772.39	kg/m3	308.15	Densities and viscosities of the mixtures (formamide + 2-alkanol): Experimental and theoretical approaches
rhol	776.79	kg/m3	303.15	Densities and viscosities of the mixtures (formamide + 2-alkanol): Experimental and theoretical approaches

rhol	780.95	kg/m3	298.15	Densities and viscosities of the mixtures (formamide + 2-alkanol): Experimental and theoretical approaches
rhol	758.63	kg/m3	323.15	Osmotic coefficients and apparent molar volumes of 1-hexyl-3-methylimidazolium trifluoromethanesulfonate ionic liquid in alcohols
rhol	758.70	kg/m3	323.15	Osmotic and apparent molar properties of binary mixtures alcohol + 1-butyl-3-methylimidazolium trifluoromethanesulfonate ionic liquid
rhol	772.31	kg/m3	308.15	Osmotic and apparent molar properties of binary mixtures alcohol + 1-butyl-3-methylimidazolium trifluoromethanesulfonate ionic liquid
rhol	785.10	kg/m3	293.15	Osmotic and apparent molar properties of binary mixtures alcohol + 1-butyl-3-methylimidazolium trifluoromethanesulfonate ionic liquid
rhol	781.40	kg/m3	298.15	Phase diagrams of (hexane + methanol + 2,2,2-trifluoroethanol) at three temperatures: Measurement and correlation
rhol	758.71	kg/m3	323.15	Densities and volumetric properties of (N-(2-hydroxyethyl)morpholine + ethanol, + 1-propanol, + 2-propanol, + 1-butanol, and + 2-butanol) at (293.15, 298.15, 303.15, 313.15, and 323.15) K

rhol	767.83	kg/m3	313.15	Densities and volumetric properties of (N-(2-hydroxyethyl)morpholine + ethanol, + 1-propanol, + 2-propanol, + 1-butanol, and + 2-butanol) at (293.15, 298.15, 303.15, 313.15, and 323.15) K
rhol	776.66	kg/m3	303.15	Densities and volumetric properties of (N-(2-hydroxyethyl)morpholine + ethanol, + 1-propanol, + 2-propanol, + 1-butanol, and + 2-butanol) at (293.15, 298.15, 303.15, 313.15, and 323.15) K
rhol	781.31	kg/m3	298.15	Densities and volumetric properties of (N-(2-hydroxyethyl)morpholine + ethanol, + 1-propanol, + 2-propanol, + 1-butanol, and + 2-butanol) at (293.15, 298.15, 303.15, 313.15, and 323.15) K
rhol	785.18	kg/m3	293.15	Densities and volumetric properties of (N-(2-hydroxyethyl)morpholine + ethanol, + 1-propanol, + 2-propanol, + 1-butanol, and + 2-butanol) at (293.15, 298.15, 303.15, 313.15, and 323.15) K
rhol	780.80	kg/m3	298.15	Density and speed of sound of lithium bromide with organic solvents: Measurement and correlation

rhol	738.91	kg/m3	343.15	Densities and derived thermodynamic properties of (2-methoxyethanol + 1-propanol, or 2-propanol, or 1,2-propandiol) at temperatures from T = (293.15 to 343.15) K	
rhol	749.06	kg/m3	333.15	Densities and derived thermodynamic properties of (2-methoxyethanol + 1-propanol, or 2-propanol, or 1,2-propandiol) at temperatures from T = (293.15 to 343.15) K	
rhol	758.68	kg/m3	323.15	Densities and derived thermodynamic properties of (2-methoxyethanol + 1-propanol, or 2-propanol, or 1,2-propandiol) at temperatures from T = (293.15 to 343.15) K	
rhol	767.85	kg/m3	313.15	Densities and derived thermodynamic properties of (2-methoxyethanol + 1-propanol, or 2-propanol, or 1,2-propandiol) at temperatures from T = (293.15 to 343.15) K	
rhol	776.63	kg/m3	303.15	Densities and derived thermodynamic properties of (2-methoxyethanol + 1-propanol, or 2-propanol, or 1,2-propandiol) at temperatures from T = (293.15 to 343.15) K	

rhol	785.10	kg/m3	293.15	Densities and derived thermodynamic properties of (2-methoxyethanol + 1-propanol, or 2-propanol, or 1,2-propandiol) at temperatures from T = (293.15 to 343.15) K
rhol	781.30	kg/m3	298.15	Bubble point temperatures of the binary mixtures of nitrobenzene with C1 C4 aliphatic alcohols at 94.95 kPa
rhol	780.79	kg/m3	298.15	Solid-liquid equilibria for selected binary systems containing diphenyl carbonate
rhol	786.00	kg/m3	293.15	Effect of imidazolium-based ionic liquid on vapor-liquid equilibria of 2-propanol + acetonitrile binary system at 101.3 kPa
rhol	782.45	kg/m3	298.15	Experimental vapour - liquid equilibrium data of the quaternary system Methanol (1) + Isopropyl Alcohol (2) + Water (3) + Glycerol (4) along with Isopropyl Alcohol (2) + Glycerol (4) and Isopropyl Alcohol (2) + Water (3) binary data at atmospheric and sub-atmospheric pressures.

rhol	786.91	kg/m3	293.15	Experimental vapour - liquid equilibrium data of the quaternary system Methanol (1) + Isopropyl Alcohol (2) + Water (3) + Glycerol (4) along with Isopropyl Alcohol (2) + Glycerol (4) and Isopropyl Alcohol (2) + Water (3) binary data at atmospheric and sub-atmospheric pressures.
rhol	780.80	kg/m3	298.15	Determination and prediction of solubilities of active pharmaceutical ingredients in selected organic solvents
rhol	782.45	kg/m3	298.15	Isobaric ternary vapour-liquid equilibrium of methanol(1) + diisopropyl ether(2) + isopropyl alcohol(3) along with methanol + isopropyl alcohol binary data at atmospheric and sub-atmospheric pressures
rhol	754.84	kg/m3	328.15	Densities, surface tensions, and isobaric vapor-liquid equilibria for the mixtures of 2-propanol, water, and 1,2-propanediol
rhol	763.61	kg/m3	318.15	Densities, surface tensions, and isobaric vapor-liquid equilibria for the mixtures of 2-propanol, water, and 1,2-propanediol

rhol	772.32	kg/m3	308.15	Densities, surface tensions, and isobaric vapor-liquid equilibria for the mixtures of 2-propanol, water, and 1,2-propanediol	
rhol	781.18	kg/m3	298.15	Densities, surface tensions, and isobaric vapor-liquid equilibria for the mixtures of 2-propanol, water, and 1,2-propanediol	
rhol	782.45	kg/m3	298.15	Experimental isobaric vapor-liquid equilibrium at atmospheric and sub-atmospheric pressures, excess molar volumes and deviations in molar refractivity from 293.15 K to 318.15 K of diisopropyl ether with methanol and isopropyl alcohol.	
rhol	758.84	kg/m3	323.15	Measurement and modeling of volumetric properties and speeds of sound of several mixtures of alcohol liquids containing 1-propanol and 2-propanol at T= (298.15 - 323.15) K and ambient pressure	
rhol	763.49	kg/m3	318.15	Measurement and modeling of volumetric properties and speeds of sound of several mixtures of alcohol liquids containing 1-propanol and 2-propanol at T= (298.15 - 323.15) K and ambient pressure	

rhol	768.02	kg/m3	313.15	Measurement and modeling of volumetric properties and speeds of sound of several mixtures of alcohol liquids containing 1-propanol and 2-propanol at T= (298.15 - 323.15) K and ambient pressure	
rhol	772.45	kg/m3	308.15	Measurement and modeling of volumetric properties and speeds of sound of several mixtures of alcohol liquids containing 1-propanol and 2-propanol at T= (298.15 - 323.15) K and ambient pressure	
rhol	776.80	kg/m3	303.15	Measurement and modeling of volumetric properties and speeds of sound of several mixtures of alcohol liquids containing 1-propanol and 2-propanol at T= (298.15 - 323.15) K and ambient pressure	
rhol	781.08	kg/m3	298.15	Measurement and modeling of volumetric properties and speeds of sound of several mixtures of alcohol liquids containing 1-propanol and 2-propanol at T= (298.15 - 323.15) K and ambient pressure	

rhol	781.18	kg/m3	298.15	Isobaric vapor-liquid equilibria for the binary and ternary mixtures of 2-propanol, water, and 1,3-propanediol at p = 101.3 kPa: Effect of the 1,3-propanediol addition
rhol	782.70	kg/m3	298.15	Solubility of androstenedione in lower alcohols
rhol	782.70	kg/m3	298.15	Liquid liquid equilibria of 4-methyl-2-pentanone + 1-propanol or 2-propanol + water ternary systems: Measurements and correlation at different temperatures
rhol	781.59	kg/m3	298.15	Thermodynamic and spectral investigations of binary liquid mixturesof 2-butoxy ethanol with alcohols at temperature range of 293.15-313.15 K
rhol	785.90	kg/m3	293.15	Measurement and correlation of liquid liquid equilibrium data for 2-methyl-1-propanol + 2-propanol + water at several temperatures
rhol	781.21	kg/m3	298.15	Liquid liquid equilibria for the binary system of di-isopropyl ether (DIPE) +water in between 288.15 and 323.15K and the ternary systems of DIPE +water + C1 C4 alcohols at 298.15K

rhol	749.01	kg/m3	333.15	Volumetric properties of the boldine + alcohol mixtures at atmospheric pressure from 283.15 to 333.15K A new method for the determination of the density of pure boldine
rhol	753.88	kg/m3	328.15	Volumetric properties of the boldine + alcohol mixtures at atmospheric pressure from 283.15 to 333.15K A new method for the determination of the density of pure boldine
rhol	758.64	kg/m3	323.15	Volumetric properties of the boldine + alcohol mixtures at atmospheric pressure from 283.15 to 333.15K A new method for the determination of the density of pure boldine
rhol	763.28	kg/m3	318.15	Volumetric properties of the boldine + alcohol mixtures at atmospheric pressure from 283.15 to 333.15K A new method for the determination of the density of pure boldine
rhol	767.81	kg/m3	313.15	Volumetric properties of the boldine + alcohol mixtures at atmospheric pressure from 283.15 to 333.15K A new method for the determination of the density of pure boldine

rhol	772.25	kg/m3	308.15	Volumetric properties of the boldine + alcohol mixtures at atmospheric pressure from 283.15 to 333.15K A new method for the determination of the density of pure boldine
rhol	776.60	kg/m3	303.15	Volumetric properties of the boldine + alcohol mixtures at atmospheric pressure from 283.15 to 333.15K A new method for the determination of the density of pure boldine
rhol	780.87	kg/m3	298.15	Volumetric properties of the boldine + alcohol mixtures at atmospheric pressure from 283.15 to 333.15K A new method for the determination of the density of pure boldine
rhol	785.08	kg/m3	293.15	Volumetric properties of the boldine + alcohol mixtures at atmospheric pressure from 283.15 to 333.15K A new method for the determination of the density of pure boldine
rhol	789.22	kg/m3	288.15	Volumetric properties of the boldine + alcohol mixtures at atmospheric pressure from 283.15 to 333.15K A new method for the determination of the density of pure boldine

rhol	793.31	kg/m3	283.15	Volumetric properties of the boldine + alcohol mixtures at atmospheric pressure from 283.15 to 333.15K A new method for the determination of the density of pure boldine
rhol	777.30	kg/m3	303.15	Viscous synergy and antagonism and isentropic compressibility of ternary mixtures containing 1,3-dioxolane, water and monoalkanols at 303.15K
rhol	780.91	kg/m3	298.15	Isothermal vapour liquid equilibria in the binary and ternary systems composed of 2-propanol, diisopropyl ether and 4-methyl-2-pentanone
rhol	741.00	kg/m3	343.30	Effect of Al2O3 Nanoparticles Additives on the Density, Saturated Vapor Pressure, Surface Tension and Viscosity of Isopropyl Alcohol
rhol	760.10	kg/m3	323.40	Effect of Al2O3 Nanoparticles Additives on the Density, Saturated Vapor Pressure, Surface Tension and Viscosity of Isopropyl Alcohol
rhol	777.50	kg/m3	303.60	Effect of Al2O3 Nanoparticles Additives on the Density, Saturated Vapor Pressure, Surface Tension and Viscosity of Isopropyl Alcohol

rhol	784.30	kg/m3	295.30	Effect of Al2O3 Nanoparticles Additives on the Density, Saturated Vapor Pressure, Surface Tension and Viscosity of Isopropyl Alcohol
rhol	795.30	kg/m3	283.10	Effect of Al2O3 Nanoparticles Additives on the Density, Saturated Vapor Pressure, Surface Tension and Viscosity of Isopropyl Alcohol
rhol	810.90	kg/m3	263.10	Effect of Al2O3 Nanoparticles Additives on the Density, Saturated Vapor Pressure, Surface Tension and Viscosity of Isopropyl Alcohol
rhol	763.90	kg/m3	318.15	Densities and Volumetric Properties of Binary Mixtures of Aniline with 1-Propanol, 2-Propanol, 2-Methyl-1-Propanol, and 2-Methyl-2-Propanol at Temperatures from 293.15 to 318.15 K
rhol	768.11	kg/m3	313.15	Densities and Volumetric Properties of Binary Mixtures of Aniline with 1-Propanol, 2-Propanol, 2-Methyl-1-Propanol, and 2-Methyl-2-Propanol at Temperatures from 293.15 to 318.15 K

rhol	772.32	kg/m3	308.15	Densities and Volumetric Properties of Binary Mixtures of Aniline with 1-Propanol, 2-Propanol, 2-Methyl-1-Propanol, and 2-Methyl-2-Propanol at Temperatures from 293.15 to 318.15 K
rhol	776.54	kg/m3	303.15	Densities and Volumetric Properties of Binary Mixtures of Aniline with 1-Propanol, 2-Propanol, 2-Methyl-1-Propanol, and 2-Methyl-2-Propanol at Temperatures from 293.15 to 318.15 K
rhol	780.75	kg/m3	298.15	Densities and Volumetric Properties of Binary Mixtures of Aniline with 1-Propanol, 2-Propanol, 2-Methyl-1-Propanol, and 2-Methyl-2-Propanol at Temperatures from 293.15 to 318.15 K
rhol	784.96	kg/m3	293.15	Densities and Volumetric Properties of Binary Mixtures of Aniline with 1-Propanol, 2-Propanol, 2-Methyl-1-Propanol, and 2-Methyl-2-Propanol at Temperatures from 293.15 to 318.15 K
rhol	781.91	kg/m3	298.15	Speeds of Sound and Isentropic Compressibilities in Binary Mixtures of 2-Propanol with Several 1-Alkanols at 298.15K
rhol	786.00	kg/m3	293.00	KDB

rhol	763.30	kg/m3	318.15	Densities and Excess Molar Volumes for Binary Glycerol + 1-Propanol, + 2-Propanol, + 1,2-Propanediol, and + 1,3-Propanediol Mixtures at Different Temperatures
rhol	767.90	kg/m3	313.15	Densities and Excess Molar Volumes for Binary Glycerol + 1-Propanol, + 2-Propanol, + 1,2-Propanediol, and + 1,3-Propanediol Mixtures at Different Temperatures
rhol	772.30	kg/m3	308.15	Densities and Excess Molar Volumes for Binary Glycerol + 1-Propanol, + 2-Propanol, + 1,2-Propanediol, and + 1,3-Propanediol Mixtures at Different Temperatures
rhol	776.60	kg/m3	303.15	Densities and Excess Molar Volumes for Binary Glycerol + 1-Propanol, + 2-Propanol, + 1,2-Propanediol, and + 1,3-Propanediol Mixtures at Different Temperatures
rhol	780.60	kg/m3	298.15	Densities and Excess Molar Volumes for Binary Glycerol + 1-Propanol, + 2-Propanol, + 1,2-Propanediol, and + 1,3-Propanediol Mixtures at Different Temperatures

rhol	780.90	kg/m3	298.15	Dynamic Viscosities of Diethyl Carbonate with Linear and Secondary Alcohols at Several Temperatures
rhol	759.00	kg/m3	323.15	Liquid Densities and Speed of Sound for Ionic Liquid (2-HEAA and 2-HDEAA) + Alcohol (1-Propanol and 2-Propanol) Mixtures at T = (293.15-323.15 K) and Atmospheric Pressure
rhol	768.20	kg/m3	313.15	Liquid Densities and Speed of Sound for Ionic Liquid (2-HEAA and 2-HDEAA) + Alcohol (1-Propanol and 2-Propanol) Mixtures at T = (293.15-323.15 K) and Atmospheric Pressure
rhol	777.00	kg/m3	303.15	Liquid Densities and Speed of Sound for Ionic Liquid (2-HEAA and 2-HDEAA) + Alcohol (1-Propanol and 2-Propanol) Mixtures at T = (293.15-323.15 K) and Atmospheric Pressure
rhol	791.50	kg/m3	293.15	Liquid Densities and Speed of Sound for Ionic Liquid (2-HEAA and 2-HDEAA) + Alcohol (1-Propanol and 2-Propanol) Mixtures at T = (293.15-323.15 K) and Atmospheric Pressure

rhol	758.50	kg/m3	323.15	Density and Viscosity of 2-Butanol + (1-Propanol, 2-Propanol, or 3-Amino-1-propanol) Mixtures at Temperatures of (293.15 to 323.15) K: Application of the ERAS Model
rhol	763.10	kg/m3	318.15	Density and Viscosity of 2-Butanol + (1-Propanol, 2-Propanol, or 3-Amino-1-propanol) Mixtures at Temperatures of (293.15 to 323.15) K: Application of the ERAS Model
rhol	767.60	kg/m3	313.15	Density and Viscosity of 2-Butanol + (1-Propanol, 2-Propanol, or 3-Amino-1-propanol) Mixtures at Temperatures of (293.15 to 323.15) K: Application of the ERAS Model
rhol	772.10	kg/m3	308.15	Density and Viscosity of 2-Butanol + (1-Propanol, 2-Propanol, or 3-Amino-1-propanol) Mixtures at Temperatures of (293.15 to 323.15) K: Application of the ERAS Model
rhol	776.40	kg/m3	303.15	Density and Viscosity of 2-Butanol + (1-Propanol, 2-Propanol, or 3-Amino-1-propanol) Mixtures at Temperatures of (293.15 to 323.15) K: Application of the ERAS Model

rhol	780.70	kg/m3	298.15	Density and Viscosity of 2-Butanol + (1-Propanol, 2-Propanol, or 3-Amino-1-propanol) Mixtures at Temperatures of (293.15 to 323.15) K: Application of the ERAS Model
rhol	784.90	kg/m3	293.15	Density and Viscosity of 2-Butanol + (1-Propanol, 2-Propanol, or 3-Amino-1-propanol) Mixtures at Temperatures of (293.15 to 323.15) K: Application of the ERAS Model
rhol	781.40	kg/m3	298.15	Determination and Correlation of Liquid-Liquid Equilibria Data for Ternary System Isopropyl Acetate + Isopropanol + Water at Different Temperatures
rhol	768.00	kg/m3	313.15	Total Pressure Phase Equilibrium Measurements for the Binary Systems of n-Pentane + Cyclohexane and 1-Hexene + 2-Propanol
rhol	743.66	kg/m3	338.15	Density, Speed of Sound, and Refractive Index Measurements for Binary Mixtures of Pentan-2-one with Propan-2-ol and Butan-2-ol
rhol	754.19	kg/m3	328.15	Density, Speed of Sound, and Refractive Index Measurements for Binary Mixtures of Pentan-2-one with Propan-2-ol and Butan-2-ol

rhol	763.51	kg/m3	318.15	Density, Speed of Sound, and Refractive Index Measurements for Binary Mixtures of Pentan-2-one with Propan-2-ol and Butan-2-ol
rhol	772.51	kg/m3	308.15	Density, Speed of Sound, and Refractive Index Measurements for Binary Mixtures of Pentan-2-one with Propan-2-ol and Butan-2-ol
rhol	781.10	kg/m3	298.15	Density, Speed of Sound, and Refractive Index Measurements for Binary Mixtures of Pentan-2-one with Propan-2-ol and Butan-2-ol
rhol	763.88	kg/m3	318.15	Densities and Excess Molar Volumes for the Binary and Ternary Systems of (1,4-Dioxane, 1-Propanol or 2-Propanol, and 1,2-Dichloroethane) at T = (288.15 to 318.15) K. Experimental Measurements and Prigogine-Flory-Patterson Modeling
rhol	772.48	kg/m3	308.15	Densities and Excess Molar Volumes for the Binary and Ternary Systems of (1,4-Dioxane, 1-Propanol or 2-Propanol, and 1,2-Dichloroethane) at T = (288.15 to 318.15) K. Experimental Measurements and Prigogine-Flory-Patterson Modeling

rhol	781.08	kg/m3	298.15	Densities and Excess Molar Volumes for the Binary and Ternary Systems of (1,4-Dioxane, 1-Propanol or 2-Propanol, and 1,2-Dichloroethane) at T = (288.15 to 318.15) K. Experimental Measurements and Prigogine-Flory-Patterson Modeling
rhol	789.68	kg/m3	288.15	Densities and Excess Molar Volumes for the Binary and Ternary Systems of (1,4-Dioxane, 1-Propanol or 2-Propanol, and 1,2-Dichloroethane) at T = (288.15 to 318.15) K. Experimental Measurements and Prigogine-Flory-Patterson Modeling
rhol	781.10	kg/m3	298.15	Isobaric Vapor-Liquid Phase Equilibrium Measurements, Correlation, and Prediction for Separation of the Mixtures of Cyclohexanone and Alcohols
rhol	768.00	kg/m3	313.15	Experimental Phase Equilibrium for the Binary System of n-Pentane +2-Propanol Using a New Equilibrium Cell and the Static Total Pressure Method
rhol	782.00	kg/m3	298.15	Isobaric Vapor-Liquid Equilibrium for the Binary Systems of Sec-butyl Acetate and Ethanol, 1-Propanol, or 2-Propanol at 101.3 kPa

rhol	758.80	kg/m3	323.15	Influence of Temperature and Carbon Chain on Thermophysical Properties of Benzaldehyde/Alkan-2-ol Binary Mixtures
rhol	763.40	kg/m3	318.15	Influence of Temperature and Carbon Chain on Thermophysical Properties of Benzaldehyde/Alkan-2-ol Binary Mixtures
rhol	768.00	kg/m3	313.15	Influence of Temperature and Carbon Chain on Thermophysical Properties of Benzaldehyde/Alkan-2-ol Binary Mixtures
rhol	772.40	kg/m3	308.15	Influence of Temperature and Carbon Chain on Thermophysical Properties of Benzaldehyde/Alkan-2-ol Binary Mixtures
rhol	776.80	kg/m3	303.15	Influence of Temperature and Carbon Chain on Thermophysical Properties of Benzaldehyde/Alkan-2-ol Binary Mixtures
rhol	781.10	kg/m3	298.15	Influence of Temperature and Carbon Chain on Thermophysical Properties of Benzaldehyde/Alkan-2-ol Binary Mixtures
rhol	785.40	kg/m3	293.15	Influence of Temperature and Carbon Chain on Thermophysical Properties of Benzaldehyde/Alkan-2-ol Binary Mixtures
rhol	758.80	kg/m3	323.15	Studies on Thermodynamic and Transport Properties of Binary Mixtures Containing Alcohols and Aniline

rhol	763.40	kg/m3	318.15	Studies on Thermodynamic and Transport Properties of Binary Mixtures Containing Alcohols and Aniline	
rhol	768.00	kg/m3	313.15	Studies on Thermodynamic and Transport Properties of Binary Mixtures Containing Alcohols and Aniline	
rhol	772.40	kg/m3	308.15	Studies on Thermodynamic and Transport Properties of Binary Mixtures Containing Alcohols and Aniline	
rhol	776.80	kg/m3	303.15	Studies on Thermodynamic and Transport Properties of Binary Mixtures Containing Alcohols and Aniline	
rhol	781.10	kg/m3	298.15	Studies on Thermodynamic and Transport Properties of Binary Mixtures Containing Alcohols and Aniline	
rhol	789.34	kg/m3	288.15	Volumetric properties of binary liquid mixtures of alcohols with 1,2-dichloroethane at different temperatures and atmospheric pressure	

rhol	781.18	kg/m3	298.15	Liquid Liquid Equilibria, Equilibrium Phase Densities, and Refractive Indices for the Quaternary Mixtures Containing 2-Propanol or 2-Methyl-2-Propanol of Fuel Oxygenate at T = 298.15 and 318.15 K
rhol	744.02	kg/m3	338.15	Thermophysical Characterization of the Mixtures of the Ionic Liquid 1-Ethyl-3-Methylimidazolium Acetate with 1-Propanol or 2-Propanol
rhol	753.89	kg/m3	328.15	Thermophysical Characterization of the Mixtures of the Ionic Liquid 1-Ethyl-3-Methylimidazolium Acetate with 1-Propanol or 2-Propanol
rhol	763.27	kg/m3	318.15	Thermophysical Characterization of the Mixtures of the Ionic Liquid 1-Ethyl-3-Methylimidazolium Acetate with 1-Propanol or 2-Propanol
rhol	772.23	kg/m3	308.15	Thermophysical Characterization of the Mixtures of the Ionic Liquid 1-Ethyl-3-Methylimidazolium Acetate with 1-Propanol or 2-Propanol
rhol	780.84	kg/m3	298.15	Thermophysical Characterization of the Mixtures of the Ionic Liquid 1-Ethyl-3-Methylimidazolium Acetate with 1-Propanol or 2-Propanol
rhol	789.18	kg/m3	288.15	Thermophysical Characterization of the Mixtures of the Ionic Liquid 1-Ethyl-3-Methylimidazolium Acetate with 1-Propanol or 2-Propanol

rhol	798.00	kg/m3	298.15	Liquid Liquid Equilibrium data for the ternary systems of Water, Isopropyl alcohol, and selected entrainers	
rhol	767.99	kg/m3	313.15	Density, Speed of Sound, Viscosity, Excess Properties, and Prigogine Flory Patterson (PFP) Theory of Binary Mixtures of Amine and Alcohols	
rhol	772.42	kg/m3	308.15	Density, Speed of Sound, Viscosity, Excess Properties, and Prigogine Flory Patterson (PFP) Theory of Binary Mixtures of Amine and Alcohols	
rhol	776.78	kg/m3	303.15	Density, Speed of Sound, Viscosity, Excess Properties, and Prigogine Flory Patterson (PFP) Theory of Binary Mixtures of Amine and Alcohols	
rhol	781.05	kg/m3	298.15	Density, Speed of Sound, Viscosity, Excess Properties, and Prigogine Flory Patterson (PFP) Theory of Binary Mixtures of Amine and Alcohols	
rhol	785.25	kg/m3	293.15	Density, Speed of Sound, Viscosity, Excess Properties, and Prigogine Flory Patterson (PFP) Theory of Binary Mixtures of Amine and Alcohols	

rhol	781.21	kg/m3	298.15	Modified Method for Measuring the Solubility of Pharmaceutical Compounds in Organic Solvents by Visual Camera	
rhol	763.29	kg/m3	318.15	The excess molar volume and the molar surface Gibbs energy of the binary of the ether-functionalized ionic liquids [C22O1IM][TfO] with ethanol and isomeric propanols at T = (288.15-318.15) K	
rhol	767.83	kg/m3	313.15	The excess molar volume and the molar surface Gibbs energy of the binary of the ether-functionalized ionic liquids [C22O1IM][TfO] with ethanol and isomeric propanols at T = (288.15-318.15) K	
rhol	772.27	kg/m3	308.15	The excess molar volume and the molar surface Gibbs energy of the binary of the ether-functionalized ionic liquids [C22O1IM][TfO] with ethanol and isomeric propanols at T = (288.15-318.15) K	
rhol	776.62	kg/m3	303.15	The excess molar volume and the molar surface Gibbs energy of the binary of the ether-functionalized ionic liquids [C22O1IM][TfO] with ethanol and isomeric propanols at T = (288.15-318.15) K	

rhol	780.89	kg/m3	298.15	The excess molar volume and the molar surface Gibbs energy of the binary of the ether-functionalized ionic liquids [C22O1IM][TfO] with ethanol and isomeric propanols at T = (288.15-318.15) K	
rhol	785.09	kg/m3	293.15	The excess molar volume and the molar surface Gibbs energy of the binary of the ether-functionalized ionic liquids [C22O1IM][TfO] with ethanol and isomeric propanols at T = (288.15-318.15) K	
rhol	789.22	kg/m3	288.15	The excess molar volume and the molar surface Gibbs energy of the binary of the ether-functionalized ionic liquids [C22O1IM][TfO] with ethanol and isomeric propanols at T = (288.15-318.15) K	
rhol	758.80	kg/m3	323.15	Evaluation of thermodynamic properties of fluid mixtures by PC-SAFT model	
rhol	763.40	kg/m3	318.15	Evaluation of thermodynamic properties of fluid mixtures by PC-SAFT model	
rhol	768.00	kg/m3	313.15	Evaluation of thermodynamic properties of fluid mixtures by PC-SAFT model	
rhol	772.40	kg/m3	308.15	Evaluation of thermodynamic properties of fluid mixtures by PC-SAFT model	

rhol	785.20	kg/m3	293.15	Volumetric properties of binary liquid mixtures of alcohols with 1,2-dichloroethane at different temperatures and atmospheric pressure
rhol	781.10	kg/m3	298.15	Evaluation of thermodynamic properties of fluid mixtures by PC-SAFT model
rhol	785.40	kg/m3	293.15	Evaluation of thermodynamic properties of fluid mixtures by PC-SAFT model
rhol	767.93	kg/m3	313.15	Volumetric properties of binary liquid mixtures of alcohols with 1,2-dichloroethane at different temperatures and atmospheric pressure
rhol	772.37	kg/m3	308.15	Volumetric properties of binary liquid mixtures of alcohols with 1,2-dichloroethane at different temperatures and atmospheric pressure
rhol	776.72	kg/m3	303.15	Volumetric properties of binary liquid mixtures of alcohols with 1,2-dichloroethane at different temperatures and atmospheric pressure
rhol	776.80	kg/m3	303.15	Evaluation of thermodynamic properties of fluid mixtures by PC-SAFT model
rhol	785.40	kg/m3	293.15	Studies on Thermodynamic and Transport Properties of Binary Mixtures Containing Alcohols and Aniline

sfust	29.21	J/mol×K	185.20	NIST Webbook
sfust	28.72	J/mol×K	184.60	NIST Webbook
sfust	29.09	J/mol×K	184.67	NIST Webbook
sfust	28.70	J/mol×K	184.60	NIST Webbook
speedsl	1138.16	m/s	298.15	Ultrasonic and Volumetric Properties of 1-Ethyl-3-methylimidazolium Trifluoromethanesulfonate Ionic Liquid with 2-Propanol or Tetrahydrofuran at Several Temperatures
speedsl	1086.62	m/s	313.15	Thermophysical study of the binary mixtures of N,N-dimethylacetamide with 2-propanol and 2-butanol
speedsl	1139.17	m/s	298.15	Thermophysical study of the binary mixtures of N,N-dimethylacetamide with 2-propanol and 2-butanol
speedsl	1173.07	m/s	288.15	Ultrasonic and Volumetric Properties of 1-Ethyl-3-methylimidazolium Trifluoromethanesulfonate Ionic Liquid with 2-Propanol or Tetrahydrofuran at Several Temperatures
speedsl	1191.88	m/s	283.15	Thermophysical study of the binary mixtures of N,N-dimethylacetamide with 2-propanol and 2-butanol
speedsl	1103.25	m/s	308.15	Ultrasonic and Volumetric Properties of 1-Ethyl-3-methylimidazolium Trifluoromethanesulfonate Ionic Liquid with 2-Propanol or Tetrahydrofuran at Several Temperatures

speedsl	1068.09	m/s	318.15	Ultrasonic and Volumetric Properties of 1-Ethyl-3-methylimidazolium Trifluoromethanesulfonate Ionic Liquid with 2-Propanol or Tetrahydrofuran at Several Temperatures
speedsl	1032.57	m/s	328.15	Ultrasonic and Volumetric Properties of 1-Ethyl-3-methylimidazolium Trifluoromethanesulfonate Ionic Liquid with 2-Propanol or Tetrahydrofuran at Several Temperatures
speedsl	1157.30	m/s	293.15	Acoustic, volumetric and osmotic properties of binary mixtures containing the ionic liquid 1-butyl-3-methylimidazolim dicyanamide mixed with primary and secondary alcohols
speedsl	1051.60	m/s	323.15	Acoustic, volumetric and osmotic properties of binary mixtures containing the ionic liquid 1-butyl-3-methylimidazolim dicyanamide mixed with primary and secondary alcohols
speedsl	1104.80	m/s	308.15	Acoustic, volumetric and osmotic properties of binary mixtures containing the ionic liquid 1-butyl-3-methylimidazolim dicyanamide mixed with primary and secondary alcohols

speedsl	1208.50	m/s	278.15	Ultrasonic and Volumetric Properties of 1-Ethyl-3-methylimidazolium Trifluoromethanesulfonate Ionic Liquid with 2-Propanol or Tetrahydrofuran at Several Temperatures
srf	0.02	N/m	303.15	The molar surface Gibbs energy and its application to the binary mixtures of N-butylpyridinium dicyanamide [C4py][DCA] with alcohols
srf	0.02	N/m	333.20	Surface tension and interfacial compositions of binary glycerol/alcohol mixtures
srf	0.02	N/m	323.20	Surface tension and interfacial compositions of binary glycerol/alcohol mixtures
srf	0.02	N/m	313.20	Surface tension and interfacial compositions of binary glycerol/alcohol mixtures
srf	0.02	N/m	303.20	Surface tension and interfacial compositions of binary glycerol/alcohol mixtures
srf	0.02	N/m	293.20	Surface tension and interfacial compositions of binary glycerol/alcohol mixtures
srf	0.02	N/m	318.15	The molar surface Gibbs energy and its application to the binary mixtures of N-butylpyridinium dicyanamide [C4py][DCA] with alcohols

srf	0.02	N/m	313.15	The molar surface Gibbs energy and its application to the binary mixtures of N-butylpyridinium dicyanamide [C4py][DCA] with alcohols
srf	0.02	N/m	308.15	The molar surface Gibbs energy and its application to the binary mixtures of N-butylpyridinium dicyanamide [C4py][DCA] with alcohols
srf	0.02	N/m	293.15	Surface Tension of Dilute Solutions of Linear Alcohols in Benzyl Alcohol
srf	0.02	N/m	298.15	The molar surface Gibbs energy and its application to the binary mixtures of N-butylpyridinium dicyanamide [C4py][DCA] with alcohols
srf	0.02	N/m	293.15	The molar surface Gibbs energy and its application to the binary mixtures of N-butylpyridinium dicyanamide [C4py][DCA] with alcohols
srf	0.02	N/m	288.15	The molar surface Gibbs energy and its application to the binary mixtures of N-butylpyridinium dicyanamide [C4py][DCA] with alcohols

srf	0.02	N/m	328.15	Effect of temperature and composition on the surface tension and surface properties of binary mixtures containing DMSO and short chain alcohols
srf	0.02	N/m	318.15	Effect of temperature and composition on the surface tension and surface properties of binary mixtures containing DMSO and short chain alcohols
srf	0.02	N/m	308.15	Effect of temperature and composition on the surface tension and surface properties of binary mixtures containing DMSO and short chain alcohols
srf	0.02	N/m	298.15	Effect of temperature and composition on the surface tension and surface properties of binary mixtures containing DMSO and short chain alcohols
srf	0.02	N/m	293.20	KDB
srf	0.02	N/m	298.15	Surface Tension of Dilute Solutions of Linear Alcohols in Benzyl Alcohol
srf	0.02	N/m	303.15	Surface Tension of Dilute Solutions of Linear Alcohols in Benzyl Alcohol
srf	0.02	N/m	308.15	Surface Tension of Dilute Solutions of Linear Alcohols in Benzyl Alcohol

srf	0.02	N/m	313.15	Surface Tension of Dilute Solutions of Linear Alcohols in Benzyl Alcohol	
srf	0.02	N/m	318.15	Surface Tension of Dilute Solutions of Linear Alcohols in Benzyl Alcohol	
srf	0.02	N/m	323.15	Surface Tension of Dilute Solutions of Linear Alcohols in Benzyl Alcohol	
srf	0.02	N/m	298.15	Experimental Data and Correlation of Surface Tensions of the Binary and Ternary Systems of Water + Acetonitrile + 2-Propanol at 298.15 K and Atmospheric Pressure	
srf	0.02	N/m	298.15	Densities, Viscosities, Refractive Indexes, and Surface Tensions for Binary Mixtures of 2-Propanol + Benzyl Alcohol, + 2-Phenylethanol and Benzyl Alcohol + 2-Phenylethanol at T) (298.15, 308.15, and 318.15) K	
srf	0.02	N/m	308.15	Densities, Viscosities, Refractive Indexes, and Surface Tensions for Binary Mixtures of 2-Propanol + Benzyl Alcohol, + 2-Phenylethanol and Benzyl Alcohol + 2-Phenylethanol at T) (298.15, 308.15, and 318.15) K	

srf	0.02	N/m	318.15	Densities, Viscosities, Refractive Indexes, and Surface Tensions for Binary Mixtures of 2-Propanol + Benzyl Alcohol, + 2-Phenylethanol and Benzyl Alcohol + 2-Phenylethanol at T) (298.15, 308.15, and 318.15) K
srf	0.02	N/m	288.15	Densities, Viscosities, Refractive Indexes, and Surface Tensions for Binary and Ternary Mixtures of Tetrahydofuran, 2-Propanol, and 2,2,4-Trimethylpentane
srf	0.02	N/m	298.15	Densities, Viscosities, Refractive Indexes, and Surface Tensions for Binary and Ternary Mixtures of Tetrahydofuran, 2-Propanol, and 2,2,4-Trimethylpentane
srf	0.02	N/m	308.15	Densities, Viscosities, Refractive Indexes, and Surface Tensions for Binary and Ternary Mixtures of Tetrahydofuran, 2-Propanol, and 2,2,4-Trimethylpentane
srf	0.02	N/m	288.15	Densities, Viscosities, Refractive Indices, and Surface Tensions for the Mixtures of 1,3-Dioxolane + 2-Propanol or + 2,2,4-Trimethylpentane at (288.15, 298.15, and 308.15) K and 1,3-Dioxolane + 2-Propanol + 2,2,4-Trimethylpentane at 298.15 K

srf	0.02	N/m	298.15	Densities, Viscosities, Refractive Indices, and Surface Tensions for the Mixtures of 1,3-Dioxolane + 2-Propanol or + 2,2,4-Trimethylpentane at (288.15, 298.15, and 308.15) K and 1,3-Dioxolane + 2-Propanol + 2,2,4-Trimethylpentane at 298.15 K
srf	0.02	N/m	308.15	Densities, Viscosities, Refractive Indices, and Surface Tensions for the Mixtures of 1,3-Dioxolane + 2-Propanol or + 2,2,4-Trimethylpentane at (288.15, 298.15, and 308.15) K and 1,3-Dioxolane + 2-Propanol + 2,2,4-Trimethylpentane at 298.15 K

Pressure Dependent Properties

Property code	Value	Unit	Pressure [kPa]	Source
tbp	355.01	K	100.00	Isobaric Vapor-Liquid Equilibria for Water + 2-Propanol + 1-Butyl-3-methylimidazolium Tetrafluoroborate

Correlations

Information	Value
Property code	pvap
Equation	ln(Pvp) = A + B/(T + C)
Coeff. A	1.60031e+01

Coeff. B	-3.41274e+03
Coeff. C	-5.57370e+01
Temperature range (K), min.	272.89
Temperature range (K), max.	508.31

Information	Value
Property code	pvap
Equation	ln(Pvp) = A + B/T + C*ln(T) + D*T^2
Coeff. A	7.13411e+01
Coeff. B	-7.69090e+03
Coeff. C	-7.69405e+00
Coeff. D	7.65635e-07
Temperature range (K), min.	185.28
Temperature range (K), max.	508.31

Datasets

Speed of sound, m/s

Temperature, K - Liquid	Pressure, kPa - Liquid	Speed of sound, m/s - Liquid
253.15	99.00	1300.967
253.15	5138.00	1328.819
253.15	10106.00	1354.875
253.15	15136.00	1380.021
253.15	20143.00	1404.015
253.15	25229.00	1427.462
253.15	30036.00	1448.868
273.15	100.00	1227.88
273.15	5275.00	1258.803
273.15	10173.00	1286.447
273.15	15164.00	1313.265
273.15	20289.00	1339.577
273.15	25102.00	1363.26
273.15	30264.00	1387.65
293.15	100.00	1157.361
293.15	5085.00	1189.739
293.15	10183.00	1220.822
293.15	15273.00	1250.197

293.15	20239.00	1277.428
293.15	25220.00	1303.56
293.15	30278.00	1328.856
313.15	100.00	1087.156
313.15	5177.00	1123.06
313.15	10104.00	1155.744
313.15	15187.00	1187.465
313.15	20126.00	1216.587
313.15	25138.00	1244.664
313.15	30110.00	1271.13
333.15	100.00	1015.957
333.15	5048.00	1054.493
333.15	10038.00	1090.587
333.15	15065.00	1124.563
333.15	20102.00	1156.534
333.15	25093.00	1186.433
333.15	30014.00	1214.42
353.15	100.00	940.503
353.15	5080.00	983.885
353.15	10091.00	1023.857
353.15	15114.00	1060.82
353.15	20132.00	1095.253
353.15	25164.00	1127.605
353.15	30380.00	1159.239

Reference *https://www.doi.org/10.1016/j.jct.2016.06.001*

Temperature, K	Pressure, kPa	Speed of sound, m/s
299.79	86.00	1132.8
349.93	138.00	950.4
217.77	143.00	1445.1
249.84	159.00	1317.8
399.64	492.00	745.0
299.80	865.00	1138.2
249.86	992.00	1322.3
217.84	1031.00	1449.0
399.63	1090.00	752.7
349.94	1099.00	958.9
450.06	1639.00	499.2
349.97	1915.00	966.0
399.63	1986.00	764.1
249.87	2039.00	1328.0
299.81	2097.00	1146.6

450.04	2190.00	512.0
217.84	2256.00	1454.8
249.85	4937.00	1343.6
349.97	5044.00	992.5
299.81	5133.00	1166.5
399.62	5275.00	802.7
217.88	6084.00	1472.2
500.17	7568.00	371.1
217.00	8951.00	1489.1
249.85	9128.00	1365.3
500.15	9582.00	424.0
299.79	9677.00	1195.1
349.97	9682.00	1029.2
399.62	9782.00	850.0
249.86	10803.00	1373.6
249.85	19905.00	1417.6
299.80	20194.00	1256.1
399.61	20985.00	948.9
349.98	21490.00	1111.7
450.02	22115.00	792.1
500.11	22653.00	636.6
216.95	24689.00	1556.1
399.63	38384.00	1071.2
299.80	39488.00	1353.7
500.25	40483.00	812.4
216.87	41140.00	1619.7
249.85	41768.00	1512.2
349.98	42396.00	1232.7
449.97	42419.00	957.3
249.85	59235.00	1579.4
216.89	60566.00	1688.1
299.81	60935.00	1447.1
349.97	61179.00	1324.0
449.98	61580.00	1076.2
500.27	61769.00	961.8
399.63	61965.00	1203.3
216.90	78557.00	1746.2
249.85	80375.00	1653.1
500.27	80461.00	1067.0
399.61	80647.00	1291.3
349.97	81377.00	1409.9
299.89	81993.00	1527.9
449.99	82381.00	1183.6
249.85	96644.00	1705.5

349.97	96752.00	1469.0
299.88	97237.00	1581.4
399.60	97607.00	1362.4
450.00	97646.00	1252.6
500.27	98305.00	1153.8
216.97	98400.00	1805.5
216.96	121241.00	1868.7
399.66	122092.00	1454.2
450.00	122429.00	1352.4
299.85	123509.00	1665.9
349.94	123510.00	1562.0
500.28	123715.00	1261.3
249.85	124462.00	1787.6

Reference *https://www.doi.org/10.1021/acs.jced.8b00938*

Molar volume, m3/mol

Temperature, K - Liquid	Pressure, kPa - Liquid	Molar volume, m3/mol - Liquid
298.15	100.00	0.0001
298.15	10000.00	0.0001
313.15	100.00	0.0001
313.15	10000.00	0.0001
328.15	100.00	0.0001
328.15	10000.00	0.0001

Reference *https://www.doi.org/10.1021/je800334m*

Molar heat capacity at constant pressure, J/K/mol

Temperature, K - Liquid	Pressure, kPa - Liquid	Molar heat capacity at constant pressure, J/K/mol - Liquid
253.15	100.00	129.40
253.15	20000.00	126.10
253.15	40000.00	125.80
258.15	100.00	130.60
258.15	20000.00	128.30
258.15	40000.00	127.30
263.15	100.00	132.60

263.15	20000.00	130.70
263.15	40000.00	129.20
268.15	100.00	135.00
268.15	20000.00	133.40
268.15	40000.00	131.50
273.15	100.00	138.00
273.15	20000.00	136.30
273.15	40000.00	134.00
278.15	100.00	141.40
278.15	20000.00	139.40
278.15	40000.00	136.80
283.15	100.00	145.10
283.15	20000.00	142.60
283.15	40000.00	140.00
288.15	100.00	148.90
288.15	20000.00	146.00
288.15	40000.00	143.30
293.15	100.00	152.90
293.15	20000.00	149.60
293.15	40000.00	146.90
298.15	100.00	157.00
298.15	20000.00	153.30
298.15	40000.00	150.70
303.15	100.00	161.10
303.15	20000.00	157.10
303.15	40000.00	154.70
308.15	100.00	165.00
308.15	20000.00	161.10
308.15	40000.00	159.00
313.15	100.00	168.60
313.15	20000.00	165.00
313.15	40000.00	163.30
318.15	100.00	171.90
318.15	20000.00	169.10
318.15	40000.00	167.80
323.15	100.00	174.90
323.15	20000.00	173.20
323.15	40000.00	172.40

Reference *https://www.doi.org/10.1016/j.tca.2010.05.012*

Viscosity, Pa*s

Temperature, K - Liquid	Pressure, kPa - Liquid	Viscosity, Pa*s - Liquid
298.15	81.50	0.0020670

Reference *https://www.doi.org/10.1016/j.jct.2016.12.036*

Refractive index (Na D-line)

Temperature, K - Liquid	Pressure, kPa - Liquid	Refractive index (Na D-line) - Liquid
298.15	101.33	1.3752
293.15	101.00	1.37724
303.15	101.00	1.37286
313.15	101.00	1.36852
323.15	101.00	1.36363

Reference *https://www.doi.org/10.1016/j.jct.2010.08.019*

Temperature, K	Pressure, kPa	Refractive index (Na D-line)
298.15	100.00	1.37503

Reference *https://www.doi.org/10.1016/j.jct.2018.02.005*

Mass density, kg/m3

Pressure, kPa - Liquid	Temperature, K - Liquid	Mass density, kg/m3 - Liquid
100.00	313.15	767.67
5000.00	313.15	772.3
10000.00	313.15	776.83
15000.00	313.15	781.04
20000.00	313.15	785.09
25000.00	313.15	788.94
30000.00	313.15	792.51
35000.00	313.15	796.05

40000.00	313.15	799.39
50000.00	313.15	805.89
60000.00	313.15	811.8
70000.00	313.15	817.44
80000.00	313.15	822.71
90000.00	313.15	827.67
100000.00	313.15	832.64
110000.00	313.15	837.15
120000.00	313.15	841.57
130000.00	313.15	845.79
140000.00	313.15	849.83
100.00	323.15	758.87
5000.00	323.15	763.96
10000.00	323.15	768.84
15000.00	323.15	773.34
20000.00	323.15	777.63
25000.00	323.15	781.69
30000.00	323.15	785.6
35000.00	323.15	789.17
40000.00	323.15	792.74
50000.00	323.15	799.4
60000.00	323.15	805.57
70000.00	323.15	811.34
80000.00	323.15	816.78
90000.00	323.15	822.03
100000.00	323.15	826.88
110000.00	323.15	831.63
120000.00	323.15	836.11
130000.00	323.15	840.39
140000.00	323.15	844.6
100.00	333.15	749.24
5000.00	333.15	754.45
10000.00	333.15	759.66
15000.00	333.15	764.42
20000.00	333.15	768.95
25000.00	333.15	773.18
30000.00	333.15	777.21
35000.00	333.15	781.18
40000.00	333.15	784.88
50000.00	333.15	791.92
60000.00	333.15	798.25
70000.00	333.15	804.32
80000.00	333.15	810.15
90000.00	333.15	815.49

100000.00	333.15	820.8
110000.00	333.15	825.59
120000.00	333.15	830.23
130000.00	333.15	834.5
140000.00	333.15	838.74
100.00	343.15	738.88
5000.00	343.15	744.75
10000.00	343.15	750.17
15000.00	343.15	755.4
20000.00	343.15	760.17
25000.00	343.15	764.81
30000.00	343.15	769.03
35000.00	343.15	773.28
40000.00	343.15	777.23
50000.00	343.15	784.53
60000.00	343.15	791.22
70000.00	343.15	797.48
80000.00	343.15	803.44
90000.00	343.15	808.87
100000.00	343.15	814.1
110000.00	343.15	819.16
120000.00	343.15	823.87
130000.00	343.15	828.44
140000.00	343.15	832.84
100.00	353.15	728.44
5000.00	353.15	734.66
10000.00	353.15	740.52
15000.00	353.15	745.99
20000.00	353.15	751.21
25000.00	353.15	756.02
30000.00	353.15	760.43
35000.00	353.15	764.79
40000.00	353.15	768.87
50000.00	353.15	777.01
60000.00	353.15	783.94
70000.00	353.15	790.65
80000.00	353.15	796.56
90000.00	353.15	802.3
100000.00	353.15	807.6
110000.00	353.15	812.87
120000.00	353.15	817.76
130000.00	353.15	822.37
140000.00	353.15	826.85
5000.00	363.15	723.93

10000.00	363.15	730.44
15000.00	363.15	736.42
20000.00	363.15	741.95
25000.00	363.15	747.05
30000.00	363.15	751.95
35000.00	363.15	756.37
40000.00	363.15	760.76
50000.00	363.15	769.06
60000.00	363.15	776.43
70000.00	363.15	783.33
80000.00	363.15	789.52
90000.00	363.15	795.58
100000.00	363.15	801.38
110000.00	363.15	806.5
120000.00	363.15	811.57
130000.00	363.15	816.39
140000.00	363.15	820.9
5000.00	373.15	712.57
10000.00	373.15	719.75
15000.00	373.15	726.13
20000.00	373.15	732.08
25000.00	373.15	737.65
30000.00	373.15	742.73
35000.00	373.15	747.59
40000.00	373.15	752.29
50000.00	373.15	760.75
60000.00	373.15	768.5
70000.00	373.15	775.82
80000.00	373.15	782.24
90000.00	373.15	788.49
100000.00	373.15	794.34
110000.00	373.15	799.85
120000.00	373.15	805.17
130000.00	373.15	810.06
140000.00	373.15	814.94
5000.00	383.15	700.32
10000.00	383.15	708.06
15000.00	383.15	715.03
20000.00	383.15	721.43
25000.00	383.15	727.35
30000.00	383.15	732.77
35000.00	383.15	737.95
40000.00	383.15	742.76
50000.00	383.15	751.85

60000.00	383.15	760.06
70000.00	383.15	767.53
80000.00	383.15	774.49
90000.00	383.15	780.94
100000.00	383.15	786.9
110000.00	383.15	792.77
120000.00	383.15	798.27
130000.00	383.15	803.3
140000.00	383.15	808.26
5000.00	393.15	687.82
10000.00	393.15	696.15
15000.00	393.15	703.77
20000.00	393.15	710.73
25000.00	393.15	717.07
30000.00	393.15	722.91
35000.00	393.15	728.4
40000.00	393.15	733.66
50000.00	393.15	743.39
60000.00	393.15	751.87
70000.00	393.15	759.71
80000.00	393.15	766.86
90000.00	393.15	773.67
100000.00	393.15	779.96
110000.00	393.15	785.9
120000.00	393.15	791.54
130000.00	393.15	796.8
140000.00	393.15	801.97
5000.00	403.15	674.03
10000.00	403.15	683.54
15000.00	403.15	691.9
20000.00	403.15	699.4
25000.00	403.15	706.26
30000.00	403.15	712.66
35000.00	403.15	718.38
40000.00	403.15	724.01
50000.00	403.15	734.72
60000.00	403.15	743.58
70000.00	403.15	751.89
80000.00	403.15	759.17
90000.00	403.15	766.2
100000.00	403.15	772.7
110000.00	403.15	778.85
120000.00	403.15	784.7
130000.00	403.15	790.2

140000.00	403.15	795.51

Reference *https://www.doi.org/10.1016/j.fluid.2013.08.007*

Pressure, kPa	Temperature, K	Mass density, kg/m3
117.90	280.21	797.61
186.70	300.09	780.22
202.70	325.12	757.05
220.40	350.14	731.92
349.50	375.01	702.23
407.60	392.57	679.26
439.60	280.21	797.89
609.40	325.12	757.49
610.50	300.09	780.62
663.40	350.14	732.49
703.30	375.01	702.82
888.70	392.57	680.18
938.70	280.21	798.29
953.40	300.09	780.94
960.00	325.12	757.87
1046.80	375.01	703.38
1071.20	350.14	733.03
1171.40	392.57	680.72
1253.90	325.12	758.21
1328.40	280.21	798.62
1339.50	300.09	781.31
1412.60	350.14	733.49
1437.50	375.01	704.01
1526.60	392.57	681.39
1638.80	325.12	758.63
1656.60	280.21	798.89
1722.10	300.09	781.66
1791.60	375.01	704.58
1844.70	350.14	734.07
1897.40	392.57	682.08
1944.90	280.21	799.14
2021.00	325.12	759.06
2044.80	300.09	781.96
2161.70	375.01	705.16
2226.70	350.14	734.58
2272.30	392.57	682.77
2328.80	325.12	759.39
2440.60	280.21	799.56

2446.40	300.09	782.34
2545.70	375.01	705.76
2619.10	392.57	683.41
2644.70	350.14	735.11
2655.80	325.15	759.75
2834.60	300.09	782.7
2842.70	350.14	735.41
2850.60	280.21	799.88
2909.90	375.01	706.34
3014.90	325.12	760.15
3050.40	392.57	684.2
3198.00	280.21	800.17
3224.90	300.09	783.06
3286.90	350.14	735.99
3321.40	325.12	760.48
3364.10	375.01	707.04
3411.20	392.57	684.86
3630.70	350.14	736.45
3659.40	325.12	760.85
3671.10	300.09	783.47
3677.60	375.01	707.53
3679.10	280.21	800.56
3835.00	392.57	685.63
4004.10	325.12	761.22
4012.20	350.14	736.94
4071.10	280.21	800.9
4074.10	300.09	783.84
4083.20	375.01	708.16
4095.70	392.57	686.12
4362.50	300.09	784.11
4394.60	350.14	737.43
4399.20	325.12	761.65
4437.70	280.21	801.19
4515.70	375.01	708.82
4524.90	392.57	686.89
4691.20	325.12	761.97
4757.60	350.14	737.91
4771.80	300.09	784.48
4832.60	280.21	801.5
4853.50	375.01	709.33
4902.80	392.57	687.57
5010.70	325.12	762.32
5182.10	280.21	801.8
5197.90	350.14	738.48

5226.50	300.09	784.89
5280.20	375.01	709.98
5283.50	392.57	688.24
5383.40	325.12	762.71
5559.10	280.21	802.1
5577.90	350.14	738.96
5579.20	300.09	785.21
5646.90	375.01	710.53
5730.50	325.12	763.08
5775.90	392.57	689.09
5910.00	280.21	802.39
5939.30	350.14	739.42
6001.60	300.09	785.59
6039.60	392.57	689.55
6138.10	375.01	711.26
6143.60	325.12	763.52
6321.30	350.14	739.89
6334.90	280.21	802.73
6427.20	300.09	785.98
6447.10	375.01	711.73
6476.60	392.57	690.3
6525.60	325.12	763.91
6730.90	350.14	740.41
6750.90	280.21	803.06
6780.30	300.09	786.29
6880.90	375.01	712.37
6889.70	392.57	691.01
6903.20	325.12	764.31
7119.70	350.14	740.89
7166.50	281.21	803.39
7214.70	300.09	786.68
7269.80	375.01	712.94
7293.80	325.12	764.72
7348.00	392.57	691.78
7564.40	350.14	741.45
7592.00	280.21	803.74
7629.90	300.09	787.05
7661.90	325.12	765.11
7728.00	392.57	692.42
7772.40	375.01	713.68
7918.60	325.12	765.38
7988.50	350.14	741.97
7999.00	280.21	804.07
8020.30	300.09	787.4

8119.20	375.01	714.18
8172.70	392.57	693.17
8254.40	325.12	765.72
8329.00	280.21	804.33
8417.50	300.09	787.75
8472.50	350.14	742.55
8557.20	375.01	714.81
8598.30	392.57	693.88
8755.50	325.12	766.24
8805.70	300.09	788.1
8817.00	280.21	804.7
8925.80	350.14	743.13
8987.30	375.01	715.42
9018.50	392.57	694.57
9212.90	300.09	788.45
9280.80	350.14	743.56
9350.60	280.21	805.11
9405.00	392.57	695.2
9423.80	375.01	716.04
9621.50	300.09	788.81
9658.50	280.21	805.38
9676.30	350.14	744.05
9837.90	392.57	695.89
9842.20	375.01	716.64

Reference *https://www.doi.org/10.1016/j.jct.2008.12.014*

Temperature, K	Pressure, kPa	Mass density, kg/m3
293.15	100.00	785.0
293.15	10000.00	793.2
293.15	20000.00	800.6
293.15	30000.00	807.4
293.15	40000.00	813.7
293.15	50000.00	819.6
293.15	60000.00	825.1
293.15	70000.00	830.4
293.15	80000.00	835.3
293.15	90000.00	840.1
293.15	100000.00	844.6
293.15	110000.00	848.9
293.15	120000.00	853.1
293.15	130000.00	857.1
293.15	140000.00	861.0

303.15	100.00	776.6
303.15	10000.00	785.3
303.15	20000.00	793.3
303.15	30000.00	800.2
303.15	40000.00	806.9
303.15	50000.00	813.1
303.15	60000.00	818.9
303.15	70000.00	824.3
303.15	80000.00	829.1
303.15	90000.00	834.2
303.15	100000.00	838.8
303.15	110000.00	843.1
303.15	120000.00	847.3
303.15	130000.00	851.4
303.15	140000.00	855.2
313.15	100.00	767.8
313.15	10000.00	777.1
313.15	20000.00	785.4
313.15	30000.00	792.9
313.15	40000.00	799.8
313.15	50000.00	806.2
313.15	60000.00	812.1
313.15	70000.00	817.6
313.15	80000.00	822.8
313.15	90000.00	827.8
313.15	100000.00	832.6
313.15	110000.00	837.3
313.15	120000.00	841.5
313.15	130000.00	845.8
313.15	140000.00	849.7
323.15	100.00	758.7
323.15	10000.00	768.6
323.15	20000.00	777.3
323.15	30000.00	785.2
323.15	40000.00	792.4
323.15	50000.00	799.0
323.15	60000.00	805.2
323.15	70000.00	811.0
323.15	80000.00	816.4
323.15	90000.00	821.6
323.15	100000.00	826.6
323.15	110000.00	831.3
323.15	120000.00	835.7
323.15	130000.00	840.0

323.15	140000.00	844.2
333.15	100.00	749.2
333.15	10000.00	759.7
333.15	20000.00	769.1
333.15	30000.00	777.4
333.15	40000.00	785.0
333.15	50000.00	792.0
333.15	60000.00	798.5
333.15	70000.00	804.5
333.15	80000.00	810.1
333.15	90000.00	815.6
333.15	100000.00	820.6
333.15	110000.00	825.4
333.15	120000.00	830.0
333.15	130000.00	834.4
333.15	140000.00	838.5
343.15	100.00	738.8
343.15	10000.00	750.2
343.15	20000.00	760.1
343.15	30000.00	768.8
343.15	40000.00	776.8
343.15	50000.00	784.1
343.15	60000.00	790.9
343.15	70000.00	797.1
343.15	80000.00	803.0
343.15	90000.00	808.6
343.15	100000.00	813.9
343.15	110000.00	818.9
343.15	120000.00	823.8
343.15	130000.00	828.3
343.15	140000.00	832.7
353.15	100.00	728.3
353.15	10000.00	740.5
353.15	20000.00	751.1
353.15	30000.00	760.6
353.15	40000.00	768.9
353.15	50000.00	776.6
353.15	60000.00	783.7
353.15	70000.00	790.3
353.15	80000.00	796.4
353.15	90000.00	802.1
353.15	100000.00	807.5
353.15	110000.00	812.6
353.15	120000.00	817.5

353.15	130000.00	822.3
353.15	140000.00	826.7
363.15	10000.00	730.2
363.15	20000.00	741.7
363.15	30000.00	751.7
363.15	40000.00	760.6
363.15	50000.00	768.7
363.15	60000.00	776.0
363.15	70000.00	782.8
363.15	80000.00	789.2
363.15	90000.00	795.2
363.15	100000.00	800.8
363.15	110000.00	806.2
363.15	120000.00	811.3
363.15	130000.00	816.2
363.15	140000.00	820.8
373.15	10000.00	719.5
373.15	20000.00	731.9
373.15	30000.00	742.6
373.15	40000.00	752.1
373.15	50000.00	760.5
373.15	60000.00	768.2
373.15	70000.00	775.4
373.15	80000.00	782.1
373.15	90000.00	788.4
373.15	100000.00	794.2
373.15	110000.00	799.8
373.15	120000.00	805.0
373.15	130000.00	810.0
373.15	140000.00	814.7
383.15	10000.00	708.0
383.15	20000.00	721.5
383.15	30000.00	732.9
383.15	40000.00	743.1
383.15	50000.00	752.0
383.15	60000.00	760.1
383.15	70000.00	767.6
383.15	80000.00	774.6
383.15	90000.00	781.0
383.15	100000.00	787.1
383.15	110000.00	792.8
383.15	120000.00	798.3
383.15	130000.00	803.5
383.15	140000.00	808.5

393.15	10000.00	696.3
393.15	20000.00	710.9
393.15	30000.00	723.2
393.15	40000.00	733.7
393.15	50000.00	743.1
393.15	60000.00	751.7
393.15	70000.00	759.6
393.15	80000.00	766.6
393.15	90000.00	773.6
393.15	100000.00	779.9
393.15	110000.00	785.9
393.15	120000.00	791.5
393.15	130000.00	796.8
393.15	140000.00	801.6
403.15	10000.00	683.7
403.15	20000.00	699.5
403.15	30000.00	712.8
403.15	40000.00	724.1
403.15	50000.00	734.2
403.15	60000.00	743.2
403.15	70000.00	751.4
403.15	80000.00	758.9
403.15	90000.00	766.0
403.15	100000.00	772.5
403.15	110000.00	778.7
403.15	120000.00	784.5
403.15	130000.00	790.1
403.15	140000.00	795.4

Reference *https://www.doi.org/10.1016/j.jct.2012.05.016*

Temperature, K	Pressure, kPa	Mass density, kg/m3
302.90	1040.00	778.1
302.90	2000.00	778.9
302.90	2990.00	779.8
302.90	4060.00	780.8
302.90	5010.00	781.6
302.90	6000.00	782.5
302.90	7020.00	783.4
302.90	8020.00	784.3
302.90	9000.00	785.1
302.90	10000.00	786.0
323.00	1000.00	760.6

323.00	2020.00	761.6
323.00	3000.00	762.7
323.00	4070.00	763.7
323.00	5000.00	764.7
323.00	6010.00	765.7
323.00	7020.00	766.7
323.00	8010.00	767.7
323.00	9010.00	768.6
323.00	10020.00	769.6

Reference *https://www.doi.org/10.1016/j.jct.2013.05.040*

Temperature, K	Pressure, kPa	Mass density, kg/m3
293.15	100.00	785.4
293.15	1000.00	786.2
293.15	5000.00	789.6
293.15	10000.00	793.6
293.15	15000.00	797.4
293.15	20000.00	801.1
293.15	25000.00	804.5
293.15	30000.00	807.9
293.15	35000.00	811.2
293.15	40000.00	814.2
293.15	45000.00	817.2
293.15	50000.00	820.1
293.15	55000.00	822.9
293.15	60000.00	825.6
293.15	65000.00	828.2
293.15	70000.00	830.8
293.15	80000.00	835.8
293.15	90000.00	840.5
293.15	100000.00	845.0
293.15	110000.00	849.4
293.15	120000.00	853.5
293.15	130000.00	857.6
293.15	140000.00	861.2
298.15	100.00	781.4
298.15	1000.00	782.1
298.15	5000.00	785.6
298.15	10000.00	789.8
298.15	15000.00	793.7
298.15	20000.00	797.3
298.15	25000.00	800.9

298.15	30000.00	804.3
298.15	35000.00	807.7
298.15	40000.00	810.8
298.15	45000.00	813.8
298.15	50000.00	816.7
298.15	55000.00	819.7
298.15	60000.00	822.5
298.15	65000.00	825.2
298.15	70000.00	827.6
298.15	80000.00	832.8
298.15	90000.00	837.5
298.15	100000.00	842.3
298.15	110000.00	846.4
298.15	120000.00	850.8
298.15	130000.00	854.8
298.15	140000.00	858.7
313.15	100.00	768.2
313.15	1000.00	769.1
313.15	5000.00	773.0
313.15	10000.00	777.4
313.15	15000.00	781.8
313.15	20000.00	785.7
313.15	25000.00	789.5
313.15	30000.00	793.3
313.15	35000.00	796.9
313.15	40000.00	800.1
313.15	45000.00	803.4
313.15	50000.00	806.5
313.15	55000.00	809.6
313.15	60000.00	812.5
313.15	65000.00	815.4
313.15	70000.00	818.2
313.15	80000.00	823.4
313.15	90000.00	828.5
313.15	100000.00	833.1
313.15	110000.00	837.8
313.15	120000.00	842.1
313.15	130000.00	846.4
313.15	140000.00	850.4
333.15	100.00	749.4
333.15	1000.00	750.5
333.15	5000.00	754.9
333.15	10000.00	760.0
333.15	15000.00	764.9

333.15	20000.00	769.3
333.15	25000.00	773.6
333.15	30000.00	777.7
333.15	35000.00	781.5
333.15	40000.00	785.3
333.15	45000.00	788.9
333.15	50000.00	792.3
333.15	55000.00	795.6
333.15	60000.00	798.7
333.15	65000.00	801.8
333.15	70000.00	804.8
333.15	80000.00	810.5
333.15	90000.00	815.9
333.15	100000.00	820.9
333.15	110000.00	825.8
333.15	120000.00	830.4
333.15	130000.00	834.8
333.15	140000.00	839.2
353.15	1000.00	729.8
353.15	5000.00	735.0
353.15	10000.00	741.0
353.15	15000.00	746.5
353.15	20000.00	751.7
353.15	25000.00	756.5
353.15	30000.00	761.0
353.15	35000.00	765.4
353.15	40000.00	769.3
353.15	45000.00	773.3
353.15	50000.00	777.0
353.15	55000.00	780.6
353.15	60000.00	784.0
353.15	65000.00	787.4
353.15	70000.00	790.6
353.15	80000.00	796.7
353.15	90000.00	802.6
353.15	100000.00	808.0
353.15	110000.00	813.2
353.15	120000.00	818.1
353.15	130000.00	822.8
353.15	140000.00	827.3
373.15	1000.00	706.6
373.15	5000.00	712.8
373.15	10000.00	719.9
373.15	15000.00	726.3

373.15	20000.00	732.0
373.15	25000.00	737.7
373.15	30000.00	742.9
373.15	35000.00	747.8
373.15	40000.00	752.2
373.15	45000.00	756.6
373.15	50000.00	760.7
373.15	55000.00	764.7
373.15	60000.00	768.6
373.15	65000.00	772.1
373.15	70000.00	775.6
373.15	80000.00	782.3
373.15	90000.00	788.5
373.15	100000.00	794.4
373.15	110000.00	799.9
373.15	120000.00	805.3
373.15	130000.00	810.2
373.15	140000.00	815.1
393.15	1000.00	680.4
393.15	5000.00	688.2
393.15	10000.00	696.7
393.15	15000.00	704.4
393.15	20000.00	711.4
393.15	25000.00	717.6
393.15	30000.00	723.5
393.15	35000.00	729.0
393.15	40000.00	734.2
393.15	45000.00	739.0
393.15	50000.00	743.5
393.15	55000.00	747.9
393.15	60000.00	752.1
393.15	65000.00	756.1
393.15	70000.00	760.0
393.15	80000.00	767.2
393.15	90000.00	773.9
393.15	100000.00	780.2
393.15	110000.00	786.2
393.15	120000.00	791.8
393.15	130000.00	797.2
393.15	140000.00	802.3
303.15	100.00	777.0
293.15	100.00	785.46
313.15	100.00	768.21
323.15	100.00	759.07

298.15	100.00	781.27
333.15	100.00	749.42

Reference https://www.doi.org/10.1016/j.jct.2018.12.018

Temperature, K	Pressure, kPa	Mass density, kg/m3
298.15	81.50	780.99

Reference https://www.doi.org/10.1021/acs.jced.5b00162

Temperature, K	Pressure, kPa	Mass density, kg/m3
298.15	100.00	781.3

Reference https://www.doi.org/10.1021/acs.jced.7b00141

Temperature, K	Pressure, kPa	Mass density, kg/m3
283.15	100.00	793.7
283.15	5000.00	797.5
283.15	10000.00	801.3
283.15	20000.00	808.9
283.18	15000.00	805.2
298.15	100.00	781.2
298.15	5000.00	785.5
298.15	10000.00	789.7
298.15	15000.00	793.8
298.15	20000.00	797.6
313.15	100.00	768.1
313.15	5000.00	773.0
313.15	10000.00	777.5
313.15	15000.00	781.9
313.15	20000.00	786.1
328.15	100.00	754.4
328.15	5000.00	759.6
328.15	15000.00	769.5
328.15	20000.00	774.0
328.18	10000.00	764.8

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Solubility of Captopril in [illegible] 2-Propanol, [illegible] Acetone, Ethyl Acetate, [illegible]	https://www.doi.org/10.1021/je900401z
[illegible] binary systems (dimethyl carbonate with 2-propanol or 2-butanol) and (diethyl carbonate with [illegible])	https://www.doi.org/10.1016/j.jct.2004.09.009
[illegible] ionic liquids for the separation processes: [illegible] (methylcyclohexane) at 101.3 kPa	https://www.doi.org/10.1016/j.jct.2016.01.017
Densities, Viscosities, Refractive Indexes, and Surface Tensions for Binary Mixtures of 2-Propanol + Benzyl Alcohol, [illegible] + Benzyl Alcohol [illegible]	https://www.doi.org/10.1021/je700140j
Solubility of N-(Phosphonomethyl)iminodiacetic [illegible]	https://www.doi.org/10.1021/je900764f
Isobaric Vapor–Liquid Equilibria for Three Ternary Systems: Water + 2-Propanol + [illegible]	https://www.doi.org/10.1021/je700092d
[illegible] for different thermodynamic properties for systems of alcohols and [illegible]	https://www.doi.org/10.1016/j.fluid.2012.11.003
Viscosities and densities for binary mixtures of N-methylpiperazine with [illegible]	https://www.doi.org/10.1016/j.fluid.2005.05.012
[illegible] of Mixtures [illegible] n-butanol and iso-butanol	https://www.doi.org/10.1021/je050440b
[illegible] Solubility of [illegible] Bioactive [illegible] Alcohols, Esters, and Aldehydes at (298.15 to 338.15) K	https://www.doi.org/10.1021/je500893g
[illegible] (1R)-fenchone + propan-1-ol or + propan-2-ol	https://www.doi.org/10.1016/j.jct.2015.04.034
[illegible] and UNIFAC: [illegible] N,N-Dimethylethanolamine with [illegible]	https://www.doi.org/10.1021/je800334m
[illegible] Using Gas-Liquid Chromatography and [illegible] Solvents from (275 to 343) [illegible]	https://www.doi.org/10.1021/je8002157
Solubility Measurements at 296 and 310 K and Physicochemical Characterization of [illegible]	https://www.doi.org/10.1021/acs.jced.8b00566
Excess Molar Enthalpies of Methyl Acetate and (1-Propanol, 2-Propanol, 1-Butanol, 2-Butanol, and 1-Pentanol) [illegible]	https://www.doi.org/10.1021/je050175u
Measurement and Correlation of the Solubility of Pyrimethanil in Seven [illegible] and Two Different Binary Mixed Solvents	https://www.doi.org/10.1021/acs.jced.8b00124
Volumetric properties of binary liquid mixtures of alcohols with 1,2-dichloroethane at different [illegible]	https://www.doi.org/10.1016/j.tca.2012.05.004
Experiment and Computation of Solubility and Dissolution Properties [illegible] Maleate and Its Intermediate in Pure Solvents	https://www.doi.org/10.1021/acs.jced.8b00333
[illegible] models of drug Norfloxacin in eight pure organic and binary solvents at T = (278.15 to 328.15) K	https://www.doi.org/10.1016/j.fluid.2016.12.012
Solubility measurement and modelling of ethyl 5-amino-4-cyano-3-(2-ethoxy-2-oxoethyl)-2-thiophenecarboxylate [illegible]	https://www.doi.org/10.1016/j.jct.2016.09.001
Solubility of Phenanthrene in Binary Mixtures of C1–C4 Alcohols [illegible]	https://www.doi.org/10.1021/je900027h
[illegible] Solvents: 2-Propanol and Ethanol + Methanol at 298.15 K	https://www.doi.org/10.1021/je600526x
Solubilities and Phase Equilibria for Ternary Solutions of alpha, alpha, alpha-Trifluorotoluene, Water, and 2-Propanol at Three Temperatures and Pressure of 101.2 kPa	https://www.doi.org/10.1016/j.fluid.2013.02.007
Measurement and correlation of solubility of [illegible] Solubility of L-Aspartic Acid in Several Neat Solvents [illegible]	https://www.doi.org/10.1021/acs.jced.9b00320
Thermochemistry of Branched Ethers: [illegible] Experimental Study of Chemical Equilibrium in the Reacting System of tert-Amyl Alkyl Ether Synthesis	https://www.doi.org/10.1021/je034172y
[illegible] mixtures from 278.00 K to 323.00 K: solubility and solution thermodynamics [illegible]	https://www.doi.org/10.1016/j.fluid.2014.03.022
Measurement and Correlation of [illegible] hydrogen tartrate in [illegible]	https://www.doi.org/10.1021/acs.jced.8b00578
Solubility of 3-Hydroxy-2-naphthoic Acid in Ten Pure and Binary Mixed Organic Solvents from T = (278.15 to 323.15) K	https://www.doi.org/10.1021/acs.jced.9b00696
Measurement and Correlation of the Solubility of Diacerein in Twelve Pure and [illegible] Binary Solvents at Temperatures from 278.15 to 318.15 K	https://www.doi.org/10.1021/acs.jced.9b00229
[illegible] Ammonium-Based Ionic Liquid as an Entrainer for the Separation of n-Propanol + Water and Isopropanol + Water Mixtures	https://www.doi.org/10.1021/acs.jced.7b00523
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Isobaric vapor-liquid equilibria for the binary and ternary mixtures of [illegible] bis(trifluoromethylsulfonyl)imide ionic liquids at 298.15 K [illegible]	https://www.doi.org/10.1016/j.fluid.2014.02.006
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Legend

af:	Acentric Factor

affp: Proton affinity
aigt: Autoignition Temperature
basg: Gas basicity
chl: Standard liquid enthalpy of combustion
cpg: Ideal gas heat capacity
cpl: Liquid phase heat capacity
dm: Dipole Moment
dvisc: Dynamic viscosity
econd: Electrical conductivity
fll: Lower Flammability Limit
flu: Upper Flammability Limit
fpc: Flash Point (Closed Cup Method)
fpo: Flash Point (Open Cup Method)
gf: Standard Gibbs free energy of formation
gyrad: Radius of Gyration
hf: Enthalpy of formation at standard conditions
hfl: Liquid phase enthalpy of formation at standard conditions
hfus: Enthalpy of fusion at standard conditions
hfust: Enthalpy of fusion at a given temperature
hvap: Enthalpy of vaporization at standard conditions
hvapt: Enthalpy of vaporization at a given temperature
ie: Ionization energy
log10ws: Log10 of Water solubility in mol/l
logp: Octanol/Water partition coefficient
mcvol: McGowan's characteristic volume
nfpaf: NFPA Fire Rating
nfpah: NFPA Health Rating
pc: Critical Pressure
pvap: Vapor pressure
rfi: Refractive Index
rhoc: Critical density
rhol: Liquid Density
rinpol: Non-polar retention indices
ripol: Polar retention indices
sfust: Entropy of fusion at a given temperature
sl: Liquid phase molar entropy at standard conditions
speedsl: Speed of sound in fluid
srf: Surface Tension
tb: Normal Boiling Point Temperature
tbp: Boiling point at given pressure
tc: Critical Temperature
tf: Normal melting (fusion) point
tt: Triple Point Temperature

vc:	Critical Volume
volm:	Molar Volume
zc:	Critical Compressibility
zra:	Rackett Parameter