

# Benzaldehyde, 3-ethoxy-

<b>Other names:</b>	Benzaldehyde, m-ethoxy- m-Ethoxybenzaldehyde 3-Ethoxybenzaldehyde
<b>Inchi:</b>	InChI=1S/C9H10O2/c1-2-11-9-5-3-4-8(6-9)7-10/h3-7H,2H2,1H3
<b>InchiKey:</b>	QZMGMXBYJZVAJN-UHFFFAOYSA-N
<b>Formula:</b>	C9H10O2
<b>SMILES:</b>	CCOc1cccc(C=O)c1
<b>Mol. weight [g/mol]:</b>	150.17
<b>CAS:</b>	22924-15-8

## Physical Properties

Property code	Value	Unit	Source
gf	-76.84	kJ/mol	Joback Method
hf	-221.83	kJ/mol	Joback Method
hfus	16.20	kJ/mol	Joback Method
hvap	47.70	kJ/mol	Joback Method
log10ws	-2.25		Crippen Method
logp	1.898		Crippen Method
mcvol	121.350	ml/mol	McGowan Method
pc	3423.86	kPa	Joback Method
tb	508.06	K	Joback Method
tc	719.99	K	Joback Method
tf	294.36	K	Joback Method
vc	0.467	m3/kmol	Joback Method

## Temperature Dependent Properties

Property code	Value	Unit	Temperature [K]	Source
cpg	257.62	J/molxK	508.06	Joback Method
cpg	269.51	J/molxK	543.38	Joback Method
cpg	280.78	J/molxK	578.70	Joback Method
cpg	291.44	J/molxK	614.02	Joback Method
cpg	301.51	J/molxK	649.35	Joback Method
cpg	310.99	J/molxK	684.67	Joback Method

cpg	319.90	J/mol×K	719.99	Joback Method
dvisc	0.0019322	Paxs	294.36	Joback Method
dvisc	0.0011334	Paxs	329.98	Joback Method
dvisc	0.0007377	Paxs	365.59	Joback Method
dvisc	0.0005181	Paxs	401.21	Joback Method
dvisc	0.0003855	Paxs	436.83	Joback Method
dvisc	0.0002999	Paxs	472.44	Joback Method
dvisc	0.0002417	Paxs	508.06	Joback Method

## Sources

<b>NIST Webbook:</b>	<a href="http://webbook.nist.gov/cgi/cbook.cgi?ID=C22924158&amp;Units=SI">http://webbook.nist.gov/cgi/cbook.cgi?ID=C22924158&amp;Units=SI</a>
<b>Crippen Method:</b>	<a href="http://pubs.acs.org/doi/abs/10.1021/ci9903071">http://pubs.acs.org/doi/abs/10.1021/ci9903071</a>
<b>Crippen Method:</b>	<a href="https://www.chemeo.com/doc/models/crippen_log10ws">https://www.chemeo.com/doc/models/crippen_log10ws</a>
<b>Joback Method:</b>	<a href="https://en.wikipedia.org/wiki/Joback_method">https://en.wikipedia.org/wiki/Joback_method</a>
<b>McGowan Method:</b>	<a href="http://link.springer.com/article/10.1007/BF02311772">http://link.springer.com/article/10.1007/BF02311772</a>

## Legend

<b>cpg:</b>	Ideal gas heat capacity
<b>dvisc:</b>	Dynamic viscosity
<b>gf:</b>	Standard Gibbs free energy of formation
<b>hf:</b>	Enthalpy of formation at standard conditions
<b>hfus:</b>	Enthalpy of fusion at standard conditions
<b>hvap:</b>	Enthalpy of vaporization at standard conditions
<b>log10ws:</b>	Log10 of Water solubility in mol/l
<b>logp:</b>	Octanol/Water partition coefficient
<b>mcvol:</b>	McGowan's characteristic volume
<b>pc:</b>	Critical Pressure
<b>tb:</b>	Normal Boiling Point Temperature
<b>tc:</b>	Critical Temperature
<b>tf:</b>	Normal melting (fusion) point
<b>vc:</b>	Critical Volume

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