

# Benzene, 1,2-dichloro-3-methyl-

<b>Other names:</b>	2,3-Dichlorotoluene Toluene, 2,3-dichloro-
<b>Inchi:</b>	InChI=1S/C7H6Cl2/c1-5-3-2-4-6(8)7(5)9/h2-4H,1H3
<b>InchiKey:</b>	GWLKCPXYBLCEKC-UHFFFAOYSA-N
<b>Formula:</b>	C7H6Cl2
<b>SMILES:</b>	Cc1cccc(Cl)c1Cl
<b>Mol. weight [g/mol]:</b>	161.03
<b>CAS:</b>	32768-54-0

## Physical Properties

Property code	Value	Unit	Source
gf	77.35	kJ/mol	Joback Method
hf	-5.70	kJ/mol	Joback Method
hfus	15.54	kJ/mol	Joback Method
hvap	43.55	kJ/mol	Joback Method
log10ws	-3.32		Crippen Method
logp	3.302		Crippen Method
mvol	110.210	ml/mol	McGowan Method
pc	3615.89	kPa	Joback Method
tb	480.70	K	NIST Webbook
tc	700.48	K	Joback Method
tf	279.95	K	Joback Method
vc	0.417	m <sup>3</sup> /kmol	Joback Method

## Temperature Dependent Properties

Property code	Value	Unit	Temperature [K]	Source
cpg	186.67	J/molxK	471.06	Joback Method
cpg	195.96	J/molxK	509.30	Joback Method
cpg	204.68	J/molxK	547.53	Joback Method
cpg	212.85	J/molxK	585.77	Joback Method
cpg	220.49	J/molxK	624.01	Joback Method
cpg	227.64	J/molxK	662.24	Joback Method
cpg	234.30	J/molxK	700.48	Joback Method

dvisc	0.0016666	Paxs	279.95	Joback Method
dvisc	0.0010588	Paxs	311.80	Joback Method
dvisc	0.0007317	Paxs	343.65	Joback Method
dvisc	0.0005384	Paxs	375.50	Joback Method
dvisc	0.0004156	Paxs	407.36	Joback Method
dvisc	0.0003331	Paxs	439.21	Joback Method
dvisc	0.0002751	Paxs	471.06	Joback Method

## Correlations

Information	Value
Property code	pvap
Equation	$\ln(P_{vp}) = A + B/(T + C)$
Coeff. A	1.39492e+01
Coeff. B	-3.78766e+03
Coeff. C	-7.47710e+01
Temperature range (K), min.	352.02
Temperature range (K), max.	513.27

## Sources

<b>NIST Webbook:</b>	<a href="http://webbook.nist.gov/cgi/cbook.cgi?ID=C32768540&amp;Units=SI">http://webbook.nist.gov/cgi/cbook.cgi?ID=C32768540&amp;Units=SI</a>
<b>The Yaws Handbook of Vapor Pressure:</b>	<a href="https://www.sciencedirect.com/book/9780128029992/the-yaws-handbook-of-vapor-pressure">https://www.sciencedirect.com/book/9780128029992/the-yaws-handbook-of-vapor-pressure</a>
<b>Crippen Method:</b>	<a href="http://pubs.acs.org/doi/abs/10.1021/ci9903071">http://pubs.acs.org/doi/abs/10.1021/ci9903071</a>
<b>Crippen Method:</b>	<a href="https://www.chemeo.com/doc/models/crippen_log10ws">https://www.chemeo.com/doc/models/crippen_log10ws</a>
<b>Joback Method:</b>	<a href="https://en.wikipedia.org/wiki/Joback_method">https://en.wikipedia.org/wiki/Joback_method</a>
<b>McGowan Method:</b>	<a href="http://link.springer.com/article/10.1007/BF02311772">http://link.springer.com/article/10.1007/BF02311772</a>

## Legend

<b>cpg:</b>	Ideal gas heat capacity
<b>dvisc:</b>	Dynamic viscosity
<b>gf:</b>	Standard Gibbs free energy of formation
<b>hf:</b>	Enthalpy of formation at standard conditions
<b>hfus:</b>	Enthalpy of fusion at standard conditions

<b>h<sub>vap</sub>:</b>	Enthalpy of vaporization at standard conditions
<b>log<sub>10</sub>w<sub>s</sub>:</b>	Log10 of Water solubility in mol/l
<b>log<sub>p</sub>:</b>	Octanol/Water partition coefficient
<b>mc<sub>vol</sub>:</b>	McGowan's characteristic volume
<b>p<sub>c</sub>:</b>	Critical Pressure
<b>p<sub>vap</sub>:</b>	Vapor pressure
<b>t<sub>b</sub>:</b>	Normal Boiling Point Temperature
<b>t<sub>c</sub>:</b>	Critical Temperature
<b>t<sub>f</sub>:</b>	Normal melting (fusion) point
<b>v<sub>c</sub>:</b>	Critical Volume

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