

# 2,3-Dihydrothiophene 1,1-dioxide

<b>Other names:</b>	Thiophene, 2,3-dihydro-, 1,1-dioxide Sulfol-2-ene «alpha»-Sulfolene 2-Sulfolene
<b>Inchi:</b>	InChI=1S/C4H6O2S/c5-7(6)3-1-2-4-7/h1,3H,2,4H2
<b>InchiKey:</b>	DYGJDTCGUUMUBL-UHFFFAOYSA-N
<b>Formula:</b>	C4H6O2S
<b>SMILES:</b>	O=S1(=O)C=CCC1
<b>Mol. weight [g/mol]:</b>	118.15
<b>CAS:</b>	1192-16-1

## Physical Properties

Property code	Value	Unit	Source
chs	-2710.40 ± 1.60	kJ/mol	NIST Webbook
gf	-404.78	kJ/mol	Joback Method
hf	-262.00 ± 3.10	kJ/mol	NIST Webbook
hfs	-323.20 ± 1.70	kJ/mol	NIST Webbook
hfus	11.11	kJ/mol	Joback Method
hsub	61.50	kJ/mol	NIST Webbook
hvap	42.99	kJ/mol	Joback Method
ie	10.28	eV	NIST Webbook
ie	10.28	eV	NIST Webbook
ie	10.00	eV	NIST Webbook
log10ws	-0.57		Crippen Method
logp	0.319		Crippen Method
mcvol	80.150	ml/mol	McGowan Method
pc	6707.62	kPa	Joback Method
tb	336.86	K	Joback Method
tc	523.01	K	Joback Method
tf	238.35	K	Joback Method
vc	0.305	m3/kmol	Joback Method

## Temperature Dependent Properties

Property code	Value	Unit	Temperature [K]	Source
cpg	120.43	J/mol×K	336.86	Joback Method
cpg	130.30	J/mol×K	367.88	Joback Method
cpg	139.66	J/mol×K	398.91	Joback Method
cpg	148.52	J/mol×K	429.93	Joback Method
cpg	156.90	J/mol×K	460.96	Joback Method
cpg	164.80	J/mol×K	491.98	Joback Method
cpg	172.25	J/mol×K	523.01	Joback Method

## Sources

<b>Crippen Method:</b>	<a href="http://pubs.acs.org/doi/abs/10.1021/ci9903071">http://pubs.acs.org/doi/abs/10.1021/ci9903071</a>
<b>Crippen Method:</b>	<a href="https://www.chemeo.com/doc/models/crippen_log10ws">https://www.chemeo.com/doc/models/crippen_log10ws</a>
<b>Joback Method:</b>	<a href="https://en.wikipedia.org/wiki/Joback_method">https://en.wikipedia.org/wiki/Joback_method</a>
<b>McGowan Method:</b>	<a href="http://link.springer.com/article/10.1007/BF02311772">http://link.springer.com/article/10.1007/BF02311772</a>
<b>NIST Webbook:</b>	<a href="http://webbook.nist.gov/cgi/cbook.cgi?ID=C1192161&amp;Units=SI">http://webbook.nist.gov/cgi/cbook.cgi?ID=C1192161&amp;Units=SI</a>

## Legend

<b>chs:</b>	Standard solid enthalpy of combustion
<b>cpg:</b>	Ideal gas heat capacity
<b>gf:</b>	Standard Gibbs free energy of formation
<b>hf:</b>	Enthalpy of formation at standard conditions
<b>hfs:</b>	Solid phase enthalpy of formation at standard conditions
<b>hfus:</b>	Enthalpy of fusion at standard conditions
<b>hsub:</b>	Enthalpy of sublimation at standard conditions
<b>hvap:</b>	Enthalpy of vaporization at standard conditions
<b>ie:</b>	Ionization energy
<b>log10ws:</b>	Log10 of Water solubility in mol/l
<b>logp:</b>	Octanol/Water partition coefficient
<b>mcvol:</b>	McGowan's characteristic volume
<b>pc:</b>	Critical Pressure
<b>tb:</b>	Normal Boiling Point Temperature
<b>tc:</b>	Critical Temperature
<b>tf:</b>	Normal melting (fusion) point
<b>vc:</b>	Critical Volume

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