

# nitric acid

<b>Other names:</b>	AZOTIC ACID HYDROGEN NITRATE
<b>Inchi:</b>	InChI=1S/HNO3/c2-1(3)4/h(H,2,3,4)
<b>InchiKey:</b>	GRYLNZFGIOXLOG-UHFFFAOYSA-N
<b>Formula:</b>	HNO3
<b>SMILES:</b>	O=[N+](O-)O
<b>Mol. weight [g/mol]:</b>	63.01
<b>CAS:</b>	7697-37-2

## Physical Properties

Property code	Value	Unit	Source
affp	751.40	kJ/mol	NIST Webbook
affp	670.30	kJ/mol	NIST Webbook
basg	731.50	kJ/mol	NIST Webbook
ea	0.57 ± 0.15	eV	NIST Webbook
ea	0.56 ± 0.17	eV	NIST Webbook
gf	-152.15	kJ/mol	Joback Method
hf	-206.32	kJ/mol	Joback Method
hfus	11.21	kJ/mol	Joback Method
hvap	48.86	kJ/mol	Joback Method
ie	11.03 ± 0.01	eV	NIST Webbook
ie	12.20	eV	NIST Webbook
ie	11.96	eV	NIST Webbook
ie	11.95 ± 0.01	eV	NIST Webbook
log10ws	0.44		Crippen Method
logp	-0.348		Crippen Method
mvol	34.150	ml/mol	McGowan Method
nfpah	%!d(float64=2)		KDB
nfpas	%!d(float64=1)		KDB
pc	8058.04	kPa	Joback Method
tb	443.42	K	Joback Method
tc	648.46	K	Joback Method
tf	294.19	K	Joback Method
vc	0.137	m3/kmol	Joback Method

# Temperature Dependent Properties

Property code	Value	Unit	Temperature [K]	Source
cpg	59.38	J/molxK	443.42	Joback Method
cpg	61.79	J/molxK	477.59	Joback Method
cpg	64.07	J/molxK	511.77	Joback Method
cpg	66.21	J/molxK	545.94	Joback Method
cpg	68.23	J/molxK	580.12	Joback Method
cpg	70.12	J/molxK	614.29	Joback Method
cpg	71.89	J/molxK	648.46	Joback Method
hvapt	38.60	kJ/mol	314.50	NIST Webbook

## Correlations

Information	Value
Property code	pvap
Equation	$\ln(P_{vp}) = A + B/(T + C)$
Coeff. A	1.54030e+01
Coeff. B	-3.29672e+03
Coeff. C	-5.04600e+01
Temperature range (K), min.	231.55
Temperature range (K), max.	376.10

## Sources

NIST Webbook:

<http://webbook.nist.gov/cgi/cbook.cgi?ID=C7697372&Units=SI>

Thermodynamic Properties of Ternary Solutions in the Water Nitric Acid Rare Earth Nitrate (Pawliszko) Systems at High Pressure  
 Equilibrium Data for the N2O5-HNO3-N2O4 System at 158, 265, 365, 473, and 581 K  
 Joback Method:

<https://www.doi.org/10.1021/acs.jced.6b00357>

<https://www.doi.org/10.1021/je800878b>

[https://en.wikipedia.org/wiki/Joback\\_method](https://en.wikipedia.org/wiki/Joback_method)

The Yaws Handbook of Vapor Pressure:  
 Crippen Method:

<https://www.sciencedirect.com/book/9780128029992/the-yaws-handbook-of-vapor-pressure>

<http://pubs.acs.org/doi/abs/10.1021/ci9903071>

Thermochemical Analysis on Rare Earth Complex of Gadolinium with Salicylic acid and 8-hydroxyquinoline:

<https://www.doi.org/10.1016/j.tca.2012.08.027>

<https://www.cheric.org/research/kdb/hcprop/showprop.php?cmpid=1911>

Apparent molar volumes and apparent molar heat capacities of aqueous lead(II) nitrate solutions from 273.15 to 323.15 K  
 Equilibrium Data for the N2O5-HNO3-N2O4 System at 158, 265, 365, 473, and 581 K  
 Key and 25 °C liquid equilibrium: Nitric acid - water - anisole/4-methyl anisole:

<https://www.doi.org/10.1016/j.jct.2003.09.013>

<https://www.doi.org/10.1021/je800971f>

<https://www.doi.org/10.1016/j.fluid.2011.10.022>

Densities and Excess Molar Volumes of the Ternary System  $\text{N}_2\text{O}_4 + \text{H}_2\text{O} + \text{HNO}_3$  at 270, 310, and 350 K  
<https://www.doi.org/10.1021/je101357s>

Effect of 17 $\beta$ -estradiol on the phase behavior of systems containing 5-hydroxymethylfurfural, water, organic solvent in the absence and presence of sodium chloride  
<https://www.doi.org/10.1016/j.jct.2019.01.001>

Distillation Separation of Hydrofluoric Acid and Nitric Acid From Acid Waste using the Salt Effect on Vapor-Liquid Equilibrium :  
[https://www.chemeo.com/doc/models/crippen\\_log10ws](https://www.chemeo.com/doc/models/crippen_log10ws)  
<http://link.springer.com/article/10.1007/BF02311772>  
<https://www.doi.org/10.1007/s10765-010-0904-8>

## Legend

<b>affp:</b>	Proton affinity
<b>basg:</b>	Gas basicity
<b>cpg:</b>	Ideal gas heat capacity
<b>ea:</b>	Electron affinity
<b>gf:</b>	Standard Gibbs free energy of formation
<b>hf:</b>	Enthalpy of formation at standard conditions
<b>hfus:</b>	Enthalpy of fusion at standard conditions
<b>hvap:</b>	Enthalpy of vaporization at standard conditions
<b>hvapt:</b>	Enthalpy of vaporization at a given temperature
<b>ie:</b>	Ionization energy
<b>log10ws:</b>	Log10 of Water solubility in mol/l
<b>logp:</b>	Octanol/Water partition coefficient
<b>mcvol:</b>	McGowan's characteristic volume
<b>nfpah:</b>	NFPA Health Rating
<b>nfpas:</b>	NFPA Safety Rating
<b>pc:</b>	Critical Pressure
<b>pvap:</b>	Vapor pressure
<b>tb:</b>	Normal Boiling Point Temperature
<b>tc:</b>	Critical Temperature
<b>tf:</b>	Normal melting (fusion) point
<b>vc:</b>	Critical Volume

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