

# N,N,2,2-Tetramethyl-1,3-propanediamine

<b>Other names:</b>	1,3-Propanediamine, N,N,2,2-tetramethyl-N,N,2,2-tetramethylpropane-1,3-diamine
<b>Inchi:</b>	InChI=1S/C7H18N2/c1-7(2,5-8)6-9(3)4/h5-6,8H2,1-4H3
<b>InchiKey:</b>	ULDIVZQLPBUHAG-UHFFFAOYSA-N
<b>Formula:</b>	C7H18N2
<b>SMILES:</b>	CN(C)CC(C)(C)CN
<b>Mol. weight [g/mol]:</b>	130.23
<b>CAS:</b>	53369-71-4

## Physical Properties

Property code	Value	Unit	Source
gf	188.13	kJ/mol	Joback Method
hf	-95.24	kJ/mol	Joback Method
hfus	14.69	kJ/mol	Joback Method
hvap	42.56	kJ/mol	Joback Method
log10ws	-0.52		Crippen Method
logp	0.533		Crippen Method
mvol	129.450	ml/mol	McGowan Method
pc	3032.27	kPa	Joback Method
tb	429.70	K	NIST Webbook
tc	629.59	K	Joback Method
tf	286.80	K	Joback Method
vc	0.464	m <sup>3</sup> /kmol	Joback Method

## Temperature Dependent Properties

Property code	Value	Unit	Temperature [K]	Source
cpg	283.71	J/molxK	441.30	Joback Method
cpg	298.62	J/molxK	472.68	Joback Method
cpg	312.72	J/molxK	504.06	Joback Method
cpg	326.05	J/molxK	535.44	Joback Method
cpg	338.63	J/molxK	566.83	Joback Method
cpg	350.51	J/molxK	598.21	Joback Method
cpg	361.72	J/molxK	629.59	Joback Method

# Correlations

Information	Value
Property code	pvap
Equation	$\ln(P_{vp}) = A + B/(T + C)$
Coeff. A	1.56608e+01
Coeff. B	-4.07585e+03
Coeff. C	-6.05930e+01
Temperature range (K), min.	302.05
Temperature range (K), max.	454.42

## Sources

<b>NIST Webbook:</b>	<a href="http://webbook.nist.gov/cgi/cbook.cgi?ID=C53369714&amp;Units=SI">http://webbook.nist.gov/cgi/cbook.cgi?ID=C53369714&amp;Units=SI</a>
<b>The Yaws Handbook of Vapor Pressure:</b>	<a href="https://www.sciencedirect.com/book/9780128029992/the-yaws-handbook-of-vapor-pressure">https://www.sciencedirect.com/book/9780128029992/the-yaws-handbook-of-vapor-pressure</a>
<b>Crippen Method:</b>	<a href="http://pubs.acs.org/doi/abs/10.1021/ci9903071">http://pubs.acs.org/doi/abs/10.1021/ci9903071</a>
<b>Crippen Method:</b>	<a href="https://www.chemeo.com/doc/models/crippen_log10ws">https://www.chemeo.com/doc/models/crippen_log10ws</a>
<b>Joback Method:</b>	<a href="https://en.wikipedia.org/wiki/Joback_method">https://en.wikipedia.org/wiki/Joback_method</a>
<b>McGowan Method:</b>	<a href="http://link.springer.com/article/10.1007/BF02311772">http://link.springer.com/article/10.1007/BF02311772</a>

## Legend

<b>cp<sub>g</sub>:</b>	Ideal gas heat capacity
<b>gf:</b>	Standard Gibbs free energy of formation
<b>hf:</b>	Enthalpy of formation at standard conditions
<b>hfus:</b>	Enthalpy of fusion at standard conditions
<b>hvap:</b>	Enthalpy of vaporization at standard conditions
<b>log<sub>10</sub>ws:</b>	Log <sub>10</sub> of Water solubility in mol/l
<b>logp:</b>	Octanol/Water partition coefficient
<b>mcvol:</b>	McGowan's characteristic volume
<b>pc:</b>	Critical Pressure
<b>pvap:</b>	Vapor pressure
<b>tb:</b>	Normal Boiling Point Temperature
<b>tc:</b>	Critical Temperature

**tf:** Normal melting (fusion) point

**vc:** Critical Volume

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