Methylene chloride

Other names:	Aerothene MM
	CH2Cl2
	Chlorure de methylene
	DICHLOROMETHANE
	F 30
	F 30 (chlorocarbon)
	FREON 30
	HCC 30
	Khladon 30
	METHYLENE DICHLORIDE
	Metaclen
	Methane dichloride
	Methane, dichloro-
	Methoklone
	Methylene bichloride
	Metylenu chlorek
	NCI-C50102
	NSC 406122
	Narkotil
	R 30
	R-30
	Rcra waste number U080
	Salesthin
	Solaesthin
	Soleana VDA
	Solmethine
	UN 1593
Inchi:	InChI=1S/CH2Cl2/c2-1-3/h1H2
InchiKey:	YMWUJEATGCHHMB-UHFFFAOYSA-N
Formula:	CH2Cl2
SMILES:	ClCCl
Mol. weight [g/mol]:	84.93
CAS:	75-09-2

Physical Properties

Property code	Value	Unit	Source

af	0.1990		KDB
affp	628.00 ± 8.00	kJ/mol	NIST Webbook
basg	602.00 ± 8.00	kJ/mol	NIST Webbook
chl	-602.50	kJ/mol	NIST Webbook
chl	-605.80 ± 8.40	kJ/mol	NIST Webbook
dm	1.80	debye	KDB
gf	-68.91	kJ/mol	KDB
gyrad	2.4320		KDB
hf	-95.10 ± 2.50	kJ/mol	NIST Webbook
hf	-95.70 ± 1.30	kJ/mol	NIST Webbook
hf	-95.46	kJ/mol	KDB
hfl	-124.30	kJ/mol	NIST Webbook
hfl	-124.10 ± 2.50	kJ/mol	NIST Webbook
hfus	6.74	kJ/mol	Joback Method
hvap	29.00	kJ/mol	NIST Webbook
hvap	29.03 ± 0.08	kJ/mol	NIST Webbook
hvap	30.60 ± 0.10	kJ/mol	NIST Webbook
hvap	28.80	kJ/mol	NIST Webbook
hvap	28.50 ± 0.42	kJ/mol	NIST Webbook
ie	11.33	eV	NIST Webbook
ie	11.32 ± 0.01	eV	NIST Webbook
ie	11.36	eV	NIST Webbook
ie	11.35 ± 0.02	eV	NIST Webbook
ie	11.40	eV	NIST Webbook
ie	11.28	eV	NIST Webbook
ie	11.33 ± 0.04	eV	NIST Webbook
ie	11.32	eV	NIST Webbook
ie	11.40	eV	NIST Webbook
log10ws	-0.63		Estimated Solubility Method
log10ws	-0.63		Aqueous Solubility Prediction Method
logp	1.421		Crippen Method
mcvol	49.430	ml/mol	McGowan Method
nfpah	%!d(float64=2)		KDB
pc	6100.00	kPa	KDB
pc	6355.00 ± 15.00	kPa	NIST Webbook
rinpol	515.00		NIST Webbook
rinpol	524.00		NIST Webbook
rinpol	524.00		NIST Webbook
rinpol	530.00		NIST Webbook
rinpol	527.00		NIST Webbook
rinpol	515.00		NIST Webbook
rinpol	504.00		NIST Webbook
rinpol	515.00		NIST Webbook

rinpol	515.00	NIST Webbook
rinpol	512.00	NIST Webbook
rinpol	515.00	NIST Webbook
rinpol	531.00	NIST Webbook
rinpol	531.00	NIST Webbook
rinpol	515.00	NIST Webbook
rinpol	519.00	NIST Webbook
rinpol	515.00	NIST Webbook
rinpol	529.00	NIST Webbook
rinpol	528.00	NIST Webbook
rinpol	527.00	NIST Webbook
rinpol	511.00	NIST Webbook
rinpol	520.00	NIST Webbook
rinpol	514.00	NIST Webbook
rinpol	528.00	NIST Webbook
rinpol	519.00	NIST Webbook
rinpol	531.00	NIST Webbook
rinpol	531.00	NIST Webbook
rinpol	488.00	NIST Webbook
rinpol	480.00	NIST Webbook
rinpol	520.00	NIST Webbook
rinpol	524.00	NIST Webbook
rinpol	531.60	NIST Webbook
rinpol	512.70	NIST Webbook
rinpol	510.00	NIST Webbook
rinpol	540.00	NIST Webbook
rinpol	540.00	NIST Webbook
rinpol	531.00	NIST Webbook
rinpol	513.00	NIST Webbook
rinpol	511.00	NIST Webbook
rinpol	477.00	NIST Webbook
rinpol	497.90	NIST Webbook
rinpol	524.00	NIST Webbook
rinpol	486.00	NIST Webbook
rinpol	511.00	NIST Webbook
rinpol	518.00	NIST Webbook
rinpol	520.00	NIST Webbook
rinpol	537.90	NIST Webbook
rinpol	512.70	NIST Webbook
rinpol	537.80	NIST Webbook
rinpol	553.50	NIST Webbook
rinpol	555.90	NIST Webbook
rinpol	553.70	NIST Webbook
rinpol	548.40	NIST Webbook

rinpol	542.20		NIST Webbook
rinpol	506.30		NIST Webbook
rinpol	518.00		NIST Webbook
rinpol	504.90		NIST Webbook
rinpol	508.00		NIST Webbook
rinpol	516.90		NIST Webbook
rinpol	516.50		NIST Webbook
rinpol	510.00		NIST Webbook
ripol	927.00		NIST Webbook
ripol	927.00		NIST Webbook
ripol	948.00		NIST Webbook
ripol	946.00		NIST Webbook
ripol	935.70		NIST Webbook
ripol	932.62		NIST Webbook
ripol	933.00		NIST Webbook
ripol	933.00		NIST Webbook
ripol	914.00		NIST Webbook
ripol	925.00		NIST Webbook
ripol	931.00		NIST Webbook
ripol	948.00		NIST Webbook
ripol	948.00		NIST Webbook
ripol	905.00		NIST Webbook
ripol	937.20		NIST Webbook
ripol	936.00		NIST Webbook
ripol	933.00		NIST Webbook
ripol	948.00		NIST Webbook
ripol	937.20		NIST Webbook
ripol	917.00		NIST Webbook
ripol	914.00		NIST Webbook
ripol	933.00		NIST Webbook
ripol	912.00		NIST Webbook
ripol	928.00		NIST Webbook
ripol	925.00		NIST Webbook
ripol	919.00		NIST Webbook
ripol	933.00		NIST Webbook
ripol	944.00		NIST Webbook
ripol	937.00		NIST Webbook
ripol	932.00		NIST Webbook
ripol	944.00		NIST Webbook
ripol	953.00		NIST Webbook
ripol	933.00		NIST Webbook
ripol	926.65		NIST Webbook
sl	174.50	J/mol×K	NIST Webbook
tb	314.70 ± 0.50	K	NIST Webbook

tb	313.00	K	NIST Webbook
tb	312.92 ± 0.07	K	NIST Webbook
tb	313.30 ± 0.30	K	NIST Webbook
tb	312.95 ± 0.50	K	NIST Webbook
tb	313.15 ± 1.00	K	NIST Webbook
tb	313.20 ± 0.50	K	NIST Webbook
tb	313.20 ± 1.00	K	NIST Webbook
tb	313.35 ± 0.20	K	NIST Webbook
tb	314.95 ± 0.50	K	NIST Webbook
tb	312.95 ± 0.30	K	NIST Webbook
tb	313.00	K	KDB
tb	312.93 ± 0.20	K	NIST Webbook
tc	510.00	K	KDB
tc	508.00 ± 0.20	K	NIST Webbook
tc	510.00	K	NIST Webbook
tf	177.62	K	Aqueous Solubility Prediction Method
tf	177.00 ± 2.00	K	NIST Webbook
tf	198.06 ± 0.40	K	NIST Webbook
tf	178.01	K	KDB
tf	176.65 ± 0.40	K	NIST Webbook
tf	176.00 ± 1.50	K	NIST Webbook
vc	0.190	m3/kmol	Joback Method
zra	0.26		KDB

Temperature Dependent Properties

Property code	Value	Unit	Temperature [K]	Source
cpg	55.35	J/mol×K	297.14	Joback Method
cpg	57.56	J/mol×K	327.38	Joback Method
cpg	59.68	J/mol×K	357.61	Joback Method
cpg	61.71	J/mol×K	387.85	Joback Method
cpg	63.64	J/mol×K	418.08	Joback Method
cpg	65.49	J/mol×K	448.32	Joback Method
cpg	67.25	J/mol×K	478.55	Joback Method
cpl	100.80	J/mol×K	292.50	NIST Webbook
cpl	100.00	J/mol×K	298.00	NIST Webbook
cpl	102.30	J/mol×K	298.15	NIST Webbook
cpl	100.00	J/mol×K	298.10	NIST Webbook
cpl	100.50	J/mol×K	292.50	NIST Webbook
cpl	129.30	J/mol×K	298.00	NIST Webbook

cpl	105.50	J/mol×K	303.20	NIST Webbook
dvisc	0.0009444	Pa×s	206.29	Joback Method
dvisc	0.0003120	Pa×s	297.14	Joback Method
dvisc	0.0014784	Pa×s	183.58	Joback Method
dvisc	0.0006594	Pa×s	229.00	Joback Method
dvisc	0.0026264	Pa×s	160.87	Joback Method
dvisc	0.0003842	Pa×s	274.43	Joback Method
dvisc	0.0004912	Pa×s	251.72	Joback Method
hfust	6.16	kJ/mol	178.20	NIST Webbook
hfust	6.16	kJ/mol	178.22	NIST Webbook
hfust	6.16	kJ/mol	178.20	NIST Webbook
hvapt	28.06	kJ/mol	313.00	NIST Webbook
hvapt	29.00	kJ/mol	347.00	NIST Webbook
hvapt	30.30	kJ/mol	287.50	NIST Webbook
hvapt	29.20	kJ/mol	308.00	NIST Webbook
hvapt	29.40	kJ/mol	249.00	NIST Webbook
hvapt	30.20	kJ/mol	273.00	NIST Webbook
pvap	74.66	kPa	304.61	Limiting activity coefficient measurements in binary mixtures of dichloromethane and 1-alkanols (C1 - C4)
pvap	45.33	kPa	292.12	Limiting activity coefficient measurements in binary mixtures of dichloromethane and 1-alkanols (C1 - C4)
pvap	53.33	kPa	296.05	Limiting activity coefficient measurements in binary mixtures of dichloromethane and 1-alkanols (C1 - C4)
pvap	58.66	kPa	298.41	Limiting activity coefficient measurements in binary mixtures of dichloromethane and 1-alkanols (C1 - C4)

pvap	63.99	kPa	300.61	Limiting activity coefficient measurements in binary mixtures of dichloromethane and 1-alkanols (C1 - C4)
pvap	69.33	kPa	302.67	Limiting activity coefficient measurements in binary mixtures of dichloromethane and 1-alkanols (C1 - C4)
pvap	37.33	kPa	287.58	Limiting activity coefficient measurements in binary mixtures of dichloromethane and 1-alkanols (C1 - C4)
pvap	79.99	kPa	306.43	Limiting activity coefficient measurements in binary mixtures of dichloromethane and 1-alkanols (C1 - C4)
pvap	85.33	kPa	308.17	Limiting activity coefficient measurements in binary mixtures of dichloromethane and 1-alkanols (C1 - C4)
pvap	90.66	kPa	309.82	Limiting activity coefficient measurements in binary mixtures of dichloromethane and 1-alkanols (C1 - C4)
pvap	95.99	kPa	311.39	Limiting activity coefficient measurements in binary mixtures of dichloromethane and 1-alkanols (C1 - C4)
pvap	67.40	kPa	301.80	Effect of Dissolved Poly(lactic acid) on the Solubility of CO2, N2, and He Gases in Dichloromethane

pvap	55.90	kPa	296.80	Effect of Dissolved Poly(lactic acid) on the Solubility of CO2, N2, and He Gases in Dichloromethane
pvap	45.70	kPa	291.90	Effect of Dissolved Poly(lactic acid) on the Solubility of CO2, N2, and He Gases in Dichloromethane
pvap	37.10	kPa	287.00	Effect of Dissolved Poly(lactic acid) on the Solubility of CO2, N2, and He Gases in Dichloromethane
pvap	197.94	kPa	333.15	Vapor Liquid Equilibrium and Excess Enthalpy Data for Systems Containing N,N-Dimethylacetamide
pvap	70.49	kPa	303.15	Total Vapor Pressure Measurements for 2-Ethoxyethanol with Carbon Tetrachloride, Chloroform, and Dichloromethane at 303.15 K
pvap	34.66	kPa	285.89	Limiting activity coefficient measurements in binary mixtures of dichloromethane and 1-alkanols (C1 - C4)
pvap	41.33	kPa	289.94	Limiting activity coefficient measurements in binary mixtures of dichloromethane and 1-alkanols (C1 - C4)
pvap	49.33	kPa	294.15	Limiting activity coefficient measurements in binary mixtures of dichloromethane and 1-alkanols (C1 - C4)

rfi	1.42370		293.15	Solubilities of Some Phosphaspirocyclic Compounds in Selected Solvents
rfi	1.42190		298.15	Partial Molar Volumes of N,N'-1,2-Ethyl-bis(salicyladimine) Schiff Base (Salen) in Organic Solvents at T = (283.15 to 318.15) K
rfi	1.42510		293.15	Partial Molar Volumes of N,N'-1,2-Ethyl-bis(salicyladimine) Schiff Base (Salen) in Organic Solvents at T = (283.15 to 318.15) K
rfi	1.42760		288.15	Partial Molar Volumes of N,N'-1,2-Ethyl-bis(salicyladimine) Schiff Base (Salen) in Organic Solvents at T = (283.15 to 318.15) K
rfi	1.43050		283.15	Partial Molar Volumes of N,N'-1,2-Ethyl-bis(salicyladimine) Schiff Base (Salen) in Organic Solvents at T = (283.15 to 318.15) K
rfi	1.42370		293.15	Solubilities of Phosphorus-Containing Compounds in Selected Solvents
rhol	1307.51	kg/m3	303.15	Volumetric properties of binary liquid mixtures: Application of the Prigogine Flory Patterson theory to excess molar volumes of dichloromethane with benzene or toluene
rhol	1327.00	kg/m3	293.15	Interfacial Properties, Densities, and Contact Angles of Task Specific Ionic Liquids

rhol	1307.60	kg/m3	303.15	Viscosity and Density for Binary Mixtures of Carbon Tetrachloride + Chloroform, Carbon Tetrachloride + Dichloromethane, and Chloroform + Dichloromethane and One Ternary Mixture of Chloroform + 1:1 (Carbon Tetrachloride + Dichloromethane) at 303.15 K
rhol	1317.00	kg/m3	298.00	KDB
rhol	1316.75	kg/m3	298.15	Volumetric properties of binary liquid mixtures: Application of the Prigogine Flory Patterson theory to excess molar volumes of dichloromethane with benzene or toluene
rhol	1325.79	kg/m3	293.15	Volumetric properties of binary liquid mixtures: Application of the Prigogine Flory Patterson theory to excess molar volumes of dichloromethane with benzene or toluene
rhol	1334.08	kg/m3	288.15	Volumetric properties of binary liquid mixtures: Application of the Prigogine Flory Patterson theory to excess molar volumes of dichloromethane with benzene or toluene

rhol	1307.90	kg/m3	303.15	Volumetric properties of dichloromethane with aniline or nitrobenzene at different temperatures: A theoretical and experimental study
rhol	1316.58	kg/m3	298.15	Volumetric properties of dichloromethane with aniline or nitrobenzene at different temperatures: A theoretical and experimental study
rhol	1325.67	kg/m3	293.15	Volumetric properties of dichloromethane with aniline or nitrobenzene at different temperatures: A theoretical and experimental study
rhol	1334.81	kg/m3	288.15	Volumetric properties of dichloromethane with aniline or nitrobenzene at different temperatures: A theoretical and experimental study
srf	0.03	N/m	293.20	KDB

Pressure Dependent Properties

Property code	Value	Unit	Pressure [kPa]	Source
tbp	313.25	K	96.60	Low cost apparatus for rapid boiling point determination of small air sensitive samples under inert atmosphere

Correlations

Information	Value
Property code	pvap
Equation	ln(Pvp) = A + B/(T + C)
Coeff. A	1.43555e+01
Coeff. B	-2.65134e+03
Coeff. C	-4.07080e+01
Temperature range (K), min.	229.18
Temperature range (K), max.	510.00

Information	Value
Property code	pvap
Equation	ln(Pvp) = A + B/T + C*ln(T) + D*T^2
Coeff. A	8.08779e+01
Coeff. B	-6.03061e+03
Coeff. C	-1.00863e+01
Coeff. D	9.81251e-06
Temperature range (K), min.	178.01
Temperature range (K), max.	510.00

Sources

High-pressure phase equilibrium data for carbon dioxide p dichloromethane p acetone and carbon dioxide p dichloromethane p acetone p N-acetylcysteine (NAC):	https://www.doi.org/10.1016/j.fluid.2018.06.008
Study of interaction between organic compounds and mono or dicationic oxygenated ionic liquids using gas chromatography:	https://www.doi.org/10.1016/j.fluid.2014.11.020
Conductivities of Binary Mixtures of Ionic Liquids with Polar Solvents:	https://www.doi.org/10.1021/je800468h
Solubility of Imidacloprid in Different Solvents:	https://www.doi.org/10.1021/je7004038
Interactions of volatile organic compounds with the ionic liquid 1-ethyl-3-methylimidazolium tetracyanoborate:	https://www.doi.org/10.1016/j.jct.2011.09.028
Crippen Method:	http://pubs.acs.org/doi/abs/10.1021/ci990307l
Solubilities of (a,Z)-2-Amino-a-(methoxyimino)-4-thiazoleethanethioic Acid S-2-Benzothiazolyl Ester in Different Pure Solvents and Binary Mixtures of Tetrahydrofuran between 278.15 K and 317.15 K:	https://www.doi.org/10.1021/je100890f
Synthesis and Solubility of 5,5-Dimethyl-2-(phenyl(phenylamino)methyl)-1,3,2-dioxaphosphinane 2-oxide in Selected Solvents:	https://www.doi.org/10.1021/acs.jced.7b00585
Effect of Dissolved Poly(lactic acid) on the Solubility of CO2, N2, and He Gases in Dichloromethane or 1,2-Dichloroethane:	https://www.doi.org/10.1021/acs.jced.5b00268
Solubilities of Phosphonic Acid, P,P'-(1,4-Piperazinediyl)bis-,P,P,P',P'-tetraphenyl Ester in Selected Solvents:	https://www.doi.org/10.1021/je400412z
Solubility of Toltrazuril and the Effect of a Cosolvent:	https://www.doi.org/10.1021/acs.jced.8b00581
A simple apparatus for data solubility determination:	https://www.doi.org/10.1016/j.fluid.2007.03.032
Joback Method:	https://en.wikipedia.org/wiki/Joback_method
Solubility of 3,4-Bis(3-nitrofurazan-4-yl)furoxan in Common Solvents at Temperatures between 293.15 K and 313.15 K:	https://www.doi.org/10.1021/je4005992
Liquid-vapor equilibrium data of CO2+ dichloromethane+ medroxyprogesterone system:	https://www.doi.org/10.1016/j.fluid.2013.10.037

Total Vapor Pressure Measurements for 2-Ethoxyethanol with Carbon Tetrachloride, Chloroform, and Dichloromethane at 303.15 K:	https://www.doi.org/10.1021/je034153o
Solubility of Simvastatin and Lovastatin in Mixtures of Dichloromethane and Supercritical Carbon Dioxide:	https://www.doi.org/10.1021/je700019h
KDB Vapor Pressure Data:	https://www.cheric.org/research/kdb/hcprop/showprop.php?cmpid=1526
Determination of the thermodynamic parameters of ionic liquid 1-propyl-3-methylimidazolium bromide by gas-liquid chromatography:	https://www.doi.org/10.1016/j.jct.2018.09.023
Solid-Liquid Equilibrium of Azacyclotridecan-2-one in 15 Pure Solvents from T = 273.15 to 323.15 K: Experimental Determination and Thermodynamic Modeling:	https://www.doi.org/10.1021/acs.jced.8b01193
Stability of Ion Pairs of Bis(trifluoromethanesulfonyl)amide-Based Ionic Liquids in Dichloromethane:	https://www.doi.org/10.1021/je900694m
Solubility Determination, Modeling, and Thermodynamic Dissolution Properties of Benzenesulfonamide in 16 Neat Solvents from 273.15 to 324.45 K:	https://www.doi.org/10.1021/acs.jced.9b00360
(Solid + liquid) phase equilibria of tetraphenyl piperazine-1,4-diyldiphosphonate in pure solvents:	https://www.doi.org/10.1016/j.jct.2014.06.021
Activity Coefficients at Infinite Dilution of Polar Solutes in 1-Butyl-3-methylimidazolium Trifluoromethanesulfonate Using Gas Liquid Chromatography:	https://www.doi.org/10.1021/je700560s
Activity coefficients at infinite dilution of organic solutes in the ionic liquid 1-butyl-3-methylimidazolium methanesulfonate:	https://www.doi.org/10.1016/j.fluid.2010.10.008
Solubility Measurement and Correlation of (+)-Biotin Intermediate Lactone in Different Organic Solvents from (287.15 to 323.75) K:	https://www.doi.org/10.1021/acs.jced.5b00857
Evaluation of the Performance of Trigeminal Tricationic Ionic Liquids for Separation Problems:	https://www.doi.org/10.1021/je201129y
Phase Separation in Binary Mixtures Containing Linear Perfluoroalkanes:	https://www.doi.org/10.1021/je049620w
Determination of Infinite Dilution Activity Coefficients of Several Organic Solutes in N-Butylpyridinium Nitrate/N-Octylpyridinium Nitrate by Blend Inverse Gas Chromatography:	https://www.doi.org/10.1021/acs.jced.7b00244
Volumetric properties of dichloromethane with aniline or pyridine at different temperatures: A theoretical and experimental study:	https://www.doi.org/10.1016/j.jct.2008.08.013
Determination and Correlation of Liquid Liquid Equilibrium Data for the Ternary Dichloromethane + Water + N,N-Dimethylacetamide System:	https://www.doi.org/10.1021/je401001z
Solubilities of Phenylphosphinic Acid, Methylphenylphosphinic Acid, Hexachlorocyclotriphosphazene, and Hexaphenoxycyclotriphosphazene in Selected Solvents:	https://www.doi.org/10.1021/je1009812
Salting-out effect on the (liquid + liquid) equilibrium for the ternary systems (water + phosphoric acid + dichloromethane or 1,2-dichloroethane):	https://www.doi.org/10.1016/j.jct.2013.11.034
Solubility of Valsartan in Different Organic Solvents and Ethanol + Water Binary Mixtures from (278.15 to 313.15) K:	https://www.doi.org/10.1021/je8007815
Solubility Measurement and Correlation for e-2,4,6,8,10,12-Hexanitro-2,4,6,8,10,12-hexaazaisowurtzitane in Five Organic Solvents at Temperatures between 283.15 and 323.15 K and Different Chloralkane + Ethyl Acetate Binary Solvents at Temperatures between 283.15 and 323.15 K:	https://www.doi.org/10.1021/acs.jced.6b00664
Solubilities of Phosphoramidic Acid, N-(phenylmethyl)-, Diphenyl Ester in Selected Solvents:	https://www.doi.org/10.1021/acs.jced.5b00007
Activity Coefficients at Infinite Dilution for Hydrocarbons in Fatty Alcohols Determined by Gas-Liquid Chromatography:	https://www.doi.org/10.1021/je1005517
Activity Coefficients at Infinite Dilution [illegible] in Alkanediamines as Stationary Phases:	https://www.doi.org/10.1021/je200628n
Solubilities of Rutaecarpine in Twelve Organic Solvents from (283.2 to 323.2) K:	https://www.doi.org/10.1021/je4001334
Solubility Measurement and Correlation of 2-Oxindole in 12 Pure Organic Solvents:	https://www.doi.org/10.1021/acs.jced.9b00308
Solubility of Enrofloxacin Sodium in Various Solvents at Various Temperatures:	https://www.doi.org/10.1021/je8005826
High Pressure Phase Equilibria of the Related Substances in the Limonene Oxidation in Supercritical CO2:	https://www.doi.org/10.1021/je020150k
Thermodynamic Properties of Mixtures Containing Ionic Liquids: Activity Coefficients at Infinite Dilution of Organic Compounds in 1-Propyl Boronic Acid-3-Alkylimidazolium Bromide and 1-Propenyl-3-alkylimidazolium Bromide Using Inverse Gas Chromatography:	https://www.doi.org/10.1021/je060033f
Activity coefficients at infinite dilution of organic solutes in 1-hexyl-3-methylimidazolium [illegible] acetate and [illegible] interfacial adsorption using gas-liquid chromatography:	https://www.doi.org/10.1016/j.jct.2013.10.038
Solvent Polarity Effect when Amberlite LA2 Is Used in the Extraction [illegible]	https://www.doi.org/10.1021/acs.jced.6b00970
Solubilities of Betulin in Fourteen Organic Solvents at Different Temperatures:	https://www.doi.org/10.1021/je700069g
Infinite dilution activity coefficients of volatile organic compounds in two ionic liquids composed of the tris(pentafluoroethyl)trifluorophosphate ([FAP]) anion and a [illegible]	https://www.doi.org/10.1016/j.jct.2013.05.035
Activity Coefficients at Infinite Dilution for Organic Solutes Dissolved in Three 1-Alkyl-1-methylpyrrolidinium Bis(trifluoromethylsulfonyl)imide Ionic Liquids Bearing Short Linear Alkyl Side Chains of Three to Five Carbons:	https://www.doi.org/10.1021/je4001894
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Aqueous Solubility Prediction Method:	http://onschallenge.wikispaces.com/file/view/AqueousDataset002.xlsx/351826032/AqueousDa
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Legend

af:	Acentric Factor
affp:	Proton affinity
basg:	Gas basicity
chl:	Standard liquid enthalpy of combustion
cpg:	Ideal gas heat capacity
cpl:	Liquid phase heat capacity
dm:	Dipole Moment
dvisc:	Dynamic viscosity
gf:	Standard Gibbs free energy of formation
gyrad:	Radius of Gyration
hf:	Enthalpy of formation at standard conditions
hfl:	Liquid phase enthalpy of formation at standard conditions
hfus:	Enthalpy of fusion at standard conditions
hfust:	Enthalpy of fusion at a given temperature
hvap:	Enthalpy of vaporization at standard conditions
hvapt:	Enthalpy of vaporization at a given temperature
ie:	Ionization energy
log10ws:	Log10 of Water solubility in mol/l
logp:	Octanol/Water partition coefficient
mcvol:	McGowan's characteristic volume
nfpah:	NFPA Health Rating
pc:	Critical Pressure
pvap:	Vapor pressure
rfi:	Refractive Index
rhol:	Liquid Density
rinpol:	Non-polar retention indices
ripol:	Polar retention indices
sl:	Liquid phase molar entropy at standard conditions
srf:	Surface Tension
tb:	Normal Boiling Point Temperature
tbp:	Boiling point at given pressure
tc:	Critical Temperature
tf:	Normal melting (fusion) point
vc:	Critical Volume
zra:	Rackett Parameter

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