

# Methylene chloride

Other names:	Aerothene MM
	CH <sub>2</sub> Cl <sub>2</sub>
	Chlorure de methylene
	DICHLOROMETHANE
	F 30
	F 30 (chlorocarbon)
	FREON 30
	HCC 30
	Khladon 30
	METHYLENE DICHLORIDE
	Metaclen
	Methane dichloride
	Methane, dichloro-
	Methoklone
	Methylene bichloride
	Metylenu chlorek
	NCI-C50102
	NSC 406122
	Narkotil
	R 30
	R-30
	Rcra waste number U080
	Salesthin
	Solaesthin
	Soleana VDA
	Solmethine
	UN 1593
Inchi:	InChI=1S/CH <sub>2</sub> Cl <sub>2</sub> /c2-1-3/h1H <sub>2</sub>
InchiKey:	YMWUJEATGCHHMB-UHFFFAOYSA-N
Formula:	CH <sub>2</sub> Cl <sub>2</sub>
SMILES:	CICCl
Mol. weight [g/mol]:	84.93
CAS:	75-09-2

## Physical Properties

Property code	Value	Unit	Source
---------------	-------	------	--------

af	0.1990		KDB
affp	628.00 ± 8.00	kJ/mol	NIST Webbook
basg	602.00 ± 8.00	kJ/mol	NIST Webbook
chl	-605.80 ± 8.40	kJ/mol	NIST Webbook
chl	-602.50	kJ/mol	NIST Webbook
dm	1.80	debye	KDB
gf	-68.91	kJ/mol	KDB
gyrad	2.4320		KDB
hf	-95.70 ± 1.30	kJ/mol	NIST Webbook
hf	-95.10 ± 2.50	kJ/mol	NIST Webbook
hf	-95.46	kJ/mol	KDB
hfl	-124.10 ± 2.50	kJ/mol	NIST Webbook
hfl	-124.30	kJ/mol	NIST Webbook
hfus	6.74	kJ/mol	Joback Method
hvap	29.00	kJ/mol	NIST Webbook
hvap	28.80	kJ/mol	NIST Webbook
hvap	30.60 ± 0.10	kJ/mol	NIST Webbook
hvap	29.03 ± 0.08	kJ/mol	NIST Webbook
hvap	28.50 ± 0.42	kJ/mol	NIST Webbook
ie	11.33 ± 0.04	eV	NIST Webbook
ie	11.32	eV	NIST Webbook
ie	11.32 ± 0.01	eV	NIST Webbook
ie	11.33	eV	NIST Webbook
ie	11.36	eV	NIST Webbook
ie	11.35 ± 0.02	eV	NIST Webbook
ie	11.40	eV	NIST Webbook
ie	11.40	eV	NIST Webbook
ie	11.28	eV	NIST Webbook
log10ws	-0.63		Estimated Solubility Method
log10ws	-0.63		Aqueous Solubility Prediction Method
logp	1.421		Crippen Method
mcvol	49.430	ml/mol	McGowan Method
nfpah	%!d(float64=2)		KDB
pc	6355.00 ± 15.00	kPa	NIST Webbook
pc	6100.00	kPa	KDB
rinpol	515.00		NIST Webbook
rinpol	515.00		NIST Webbook
rinpol	524.00		NIST Webbook
rinpol	524.00		NIST Webbook
rinpol	530.00		NIST Webbook
rinpol	527.00		NIST Webbook
rinpol	510.00		NIST Webbook
rinpol	504.00		NIST Webbook

rinpol	531.00	NIST Webbook
rinpol	515.00	NIST Webbook
rinpol	515.00	NIST Webbook
rinpol	512.00	NIST Webbook
rinpol	511.00	NIST Webbook
rinpol	531.00	NIST Webbook
rinpol	515.00	NIST Webbook
rinpol	515.00	NIST Webbook
rinpol	512.70	NIST Webbook
rinpol	553.50	NIST Webbook
rinpol	555.90	NIST Webbook
rinpol	553.70	NIST Webbook
rinpol	548.40	NIST Webbook
rinpol	542.20	NIST Webbook
rinpol	506.30	NIST Webbook
rinpol	518.00	NIST Webbook
rinpol	504.90	NIST Webbook
rinpol	508.00	NIST Webbook
rinpol	516.90	NIST Webbook
rinpol	516.50	NIST Webbook
rinpol	537.80	NIST Webbook
rinpol	537.90	NIST Webbook
rinpol	519.00	NIST Webbook
rinpol	520.00	NIST Webbook
rinpol	518.00	NIST Webbook
rinpol	486.00	NIST Webbook
rinpol	524.00	NIST Webbook
rinpol	497.90	NIST Webbook
rinpol	477.00	NIST Webbook
rinpol	511.00	NIST Webbook
rinpol	513.00	NIST Webbook
rinpol	531.00	NIST Webbook
rinpol	540.00	NIST Webbook
rinpol	540.00	NIST Webbook
rinpol	519.00	NIST Webbook
rinpol	512.70	NIST Webbook
rinpol	531.60	NIST Webbook
rinpol	524.00	NIST Webbook
rinpol	520.00	NIST Webbook
rinpol	480.00	NIST Webbook
rinpol	488.00	NIST Webbook
rinpol	531.00	NIST Webbook
rinpol	531.00	NIST Webbook
rinpol	528.00	NIST Webbook

rinpol	514.00		NIST Webbook
rinpol	520.00		NIST Webbook
rinpol	511.00		NIST Webbook
rinpol	527.00		NIST Webbook
rinpol	528.00		NIST Webbook
rinpol	529.00		NIST Webbook
rinpol	515.00		NIST Webbook
rinpol	510.00		NIST Webbook
ripol	914.00		NIST Webbook
ripol	948.00		NIST Webbook
ripol	948.00		NIST Webbook
ripol	946.00		NIST Webbook
ripol	935.70		NIST Webbook
ripol	932.62		NIST Webbook
ripol	926.65		NIST Webbook
ripol	933.00		NIST Webbook
ripol	948.00		NIST Webbook
ripol	937.20		NIST Webbook
ripol	917.00		NIST Webbook
ripol	914.00		NIST Webbook
ripol	933.00		NIST Webbook
ripol	927.00		NIST Webbook
ripol	933.00		NIST Webbook
ripol	928.00		NIST Webbook
ripol	925.00		NIST Webbook
ripol	919.00		NIST Webbook
ripol	933.00		NIST Webbook
ripol	944.00		NIST Webbook
ripol	933.00		NIST Webbook
ripol	937.00		NIST Webbook
ripol	932.00		NIST Webbook
ripol	944.00		NIST Webbook
ripol	953.00		NIST Webbook
ripol	933.00		NIST Webbook
ripol	927.00		NIST Webbook
ripol	936.00		NIST Webbook
ripol	937.20		NIST Webbook
ripol	905.00		NIST Webbook
ripol	948.00		NIST Webbook
ripol	931.00		NIST Webbook
ripol	925.00		NIST Webbook
ripol	912.00		NIST Webbook
sl	174.50	J/molxK	NIST Webbook
tb	313.00	K	KDB

tb	312.95 ± 0.30	K	NIST Webbook
tb	313.35 ± 0.20	K	NIST Webbook
tb	313.20 ± 1.00	K	NIST Webbook
tb	313.20 ± 0.50	K	NIST Webbook
tb	313.15 ± 1.00	K	NIST Webbook
tb	314.70 ± 0.50	K	NIST Webbook
tb	313.30 ± 0.30	K	NIST Webbook
tb	312.92 ± 0.07	K	NIST Webbook
tb	313.00	K	NIST Webbook
tb	312.93 ± 0.20	K	NIST Webbook
tb	314.95 ± 0.50	K	NIST Webbook
tb	312.95 ± 0.50	K	NIST Webbook
tc	510.00	K	NIST Webbook
tc	508.00 ± 0.20	K	NIST Webbook
tc	510.00	K	KDB
tf	178.01	K	KDB
tf	176.65 ± 0.40	K	NIST Webbook
tf	176.00 ± 1.50	K	NIST Webbook
tf	198.06 ± 0.40	K	NIST Webbook
tf	177.62	K	Aqueous Solubility Prediction Method
tf	177.00 ± 2.00	K	NIST Webbook
vc	0.190	m3/kmol	Joback Method
zra	0.26		KDB

## Temperature Dependent Properties

Property code	Value	Unit	Temperature [K]	Source
cpg	65.49	J/molxK	448.32	Joback Method
cpg	61.71	J/molxK	387.85	Joback Method
cpg	59.68	J/molxK	357.61	Joback Method
cpg	57.56	J/molxK	327.38	Joback Method
cpg	67.25	J/molxK	478.55	Joback Method
cpg	55.35	J/molxK	297.14	Joback Method
cpg	63.64	J/molxK	418.08	Joback Method
cpl	100.00	J/molxK	298.00	NIST Webbook
cpl	100.50	J/molxK	292.50	NIST Webbook
cpl	100.80	J/molxK	292.50	NIST Webbook
cpl	102.30	J/molxK	298.15	NIST Webbook
cpl	105.50	J/molxK	303.20	NIST Webbook
cpl	129.30	J/molxK	298.00	NIST Webbook

cpl	100.00	J/molxK	298.10	NIST Webbook
dvisc	0.0006594	Paxs	229.00	Joback Method
dvisc	0.0004912	Paxs	251.72	Joback Method
dvisc	0.0003842	Paxs	274.43	Joback Method
dvisc	0.0026264	Paxs	160.87	Joback Method
dvisc	0.0014784	Paxs	183.58	Joback Method
dvisc	0.0009444	Paxs	206.29	Joback Method
dvisc	0.0003120	Paxs	297.14	Joback Method
hfust	6.16	kJ/mol	178.22	NIST Webbook
hfust	6.16	kJ/mol	178.20	NIST Webbook
hfust	6.16	kJ/mol	178.20	NIST Webbook
hvapt	29.00	kJ/mol	347.00	NIST Webbook
hvapt	28.06	kJ/mol	313.00	NIST Webbook
hvapt	29.40	kJ/mol	249.00	NIST Webbook
hvapt	29.20	kJ/mol	308.00	NIST Webbook
hvapt	30.30	kJ/mol	287.50	NIST Webbook
hvapt	30.20	kJ/mol	273.00	NIST Webbook
pvap	197.94	kPa	333.15	Vapor Liquid Equilibrium and Excess Enthalpy Data for Systems Containing N,N-Dimethylacetamide
pvap	34.66	kPa	285.89	Limiting activity coefficient measurements in binary mixtures of dichloromethane and 1-alkanols (C1 - C4)
pvap	37.33	kPa	287.58	Limiting activity coefficient measurements in binary mixtures of dichloromethane and 1-alkanols (C1 - C4)
pvap	41.33	kPa	289.94	Limiting activity coefficient measurements in binary mixtures of dichloromethane and 1-alkanols (C1 - C4)

pvap	70.49	kPa	303.15	Total Vapor Pressure Measurements for 2-Ethoxyethanol with Carbon Tetrachloride, Chloroform, and Dichloromethane at 303.15 K
pvap	49.33	kPa	294.15	Limiting activity coefficient measurements in binary mixtures of dichloromethane and 1-alkanols (C1 - C4)
pvap	53.33	kPa	296.05	Limiting activity coefficient measurements in binary mixtures of dichloromethane and 1-alkanols (C1 - C4)
pvap	58.66	kPa	298.41	Limiting activity coefficient measurements in binary mixtures of dichloromethane and 1-alkanols (C1 - C4)
pvap	63.99	kPa	300.61	Limiting activity coefficient measurements in binary mixtures of dichloromethane and 1-alkanols (C1 - C4)
pvap	69.33	kPa	302.67	Limiting activity coefficient measurements in binary mixtures of dichloromethane and 1-alkanols (C1 - C4)
pvap	74.66	kPa	304.61	Limiting activity coefficient measurements in binary mixtures of dichloromethane and 1-alkanols (C1 - C4)

pvap	79.99	kPa	306.43	Limiting activity coefficient measurements in binary mixtures of dichloromethane and 1-alkanols (C1 - C4)
pvap	85.33	kPa	308.17	Limiting activity coefficient measurements in binary mixtures of dichloromethane and 1-alkanols (C1 - C4)
pvap	90.66	kPa	309.82	Limiting activity coefficient measurements in binary mixtures of dichloromethane and 1-alkanols (C1 - C4)
pvap	95.99	kPa	311.39	Limiting activity coefficient measurements in binary mixtures of dichloromethane and 1-alkanols (C1 - C4)
pvap	67.40	kPa	301.80	Effect of Dissolved Poly(lactic acid) on the Solubility of CO <sub>2</sub> , N <sub>2</sub> , and He Gases in Dichloromethane
pvap	55.90	kPa	296.80	Effect of Dissolved Poly(lactic acid) on the Solubility of CO <sub>2</sub> , N <sub>2</sub> , and He Gases in Dichloromethane
pvap	45.70	kPa	291.90	Effect of Dissolved Poly(lactic acid) on the Solubility of CO <sub>2</sub> , N <sub>2</sub> , and He Gases in Dichloromethane
pvap	37.10	kPa	287.00	Effect of Dissolved Poly(lactic acid) on the Solubility of CO <sub>2</sub> , N <sub>2</sub> , and He Gases in Dichloromethane



pvap	45.33	kPa	292.12	Limiting activity coefficient measurements in binary mixtures of dichloromethane and 1-alkanols (C1 - C4)
rfi	1.42190		298.15	Partial Molar Volumes of N,N'-1,2-Ethyl-bis(salicyladimine) Schiff Base (Salen) in Organic Solvents at T = (283.15 to 318.15) K
rfi	1.42370		293.15	Solubilities of Some Phosphaspirocyclic Compounds in Selected Solvents
rfi	1.42370		293.15	Solubilities of Phosphorus-Containing Compounds in Selected Solvents
rfi	1.43050		283.15	Partial Molar Volumes of N,N'-1,2-Ethyl-bis(salicyladimine) Schiff Base (Salen) in Organic Solvents at T = (283.15 to 318.15) K
rfi	1.42760		288.15	Partial Molar Volumes of N,N'-1,2-Ethyl-bis(salicyladimine) Schiff Base (Salen) in Organic Solvents at T = (283.15 to 318.15) K
rfi	1.42510		293.15	Partial Molar Volumes of N,N'-1,2-Ethyl-bis(salicyladimine) Schiff Base (Salen) in Organic Solvents at T = (283.15 to 318.15) K
rhoI	1325.67	kg/m3	293.15	Volumetric properties of dichloromethane with aniline or nitrobenzene at different temperatures: A theoretical and experimental study
rhoI	1317.00	kg/m3	298.00	KDB

rhoI	1334.81	kg/m3	288.15	Volumetric properties of dichloromethane with aniline or nitrobenzene at different temperatures: A theoretical and experimental study
rhoI	1327.00	kg/m3	293.15	Interfacial Properties, Densities, and Contact Angles of Task Specific Ionic Liquids
rhoI	1316.58	kg/m3	298.15	Volumetric properties of dichloromethane with aniline or nitrobenzene at different temperatures: A theoretical and experimental study
rhoI	1307.90	kg/m3	303.15	Volumetric properties of dichloromethane with aniline or nitrobenzene at different temperatures: A theoretical and experimental study
rhoI	1334.08	kg/m3	288.15	Volumetric properties of binary liquid mixtures: Application of the Prigogine Flory Patterson theory to excess molar volumes of dichloromethane with benzene or toluene
rhoI	1325.79	kg/m3	293.15	Volumetric properties of binary liquid mixtures: Application of the Prigogine Flory Patterson theory to excess molar volumes of dichloromethane with benzene or toluene

rhoI	1316.75	kg/m3	298.15	Volumetric properties of binary liquid mixtures: Application of the Prigogine Flory Patterson theory to excess molar volumes of dichloromethane with benzene or toluene
rhoI	1307.51	kg/m3	303.15	Volumetric properties of binary liquid mixtures: Application of the Prigogine Flory Patterson theory to excess molar volumes of dichloromethane with benzene or toluene
rhoI	1307.60	kg/m3	303.15	Viscosity and Density for Binary Mixtures of Carbon Tetrachloride + Chloroform, Carbon Tetrachloride + Dichloromethane, and Chloroform + Dichloromethane and One Ternary Mixture of Chloroform + 1:1 (Carbon Tetrachloride + Dichloromethane) at 303.15 K
srf	0.03	N/m	293.20	KDB

## Pressure Dependent Properties

Property code	Value	Unit	Pressure [kPa]	Source
tbp	313.25	K	96.60	Low cost apparatus for rapid boiling point determination of small air sensitive samples under inert atmosphere

# Correlations

Information	Value
Property code	pvap
Equation	$\ln(P_{vp}) = A + B/(T + C)$
Coeff. A	1.43555e+01
Coeff. B	-2.65134e+03
Coeff. C	-4.07080e+01
Temperature range (K), min.	229.18
Temperature range (K), max.	510.00

Information	Value
Property code	pvap
Equation	$\ln(P_{vp}) = A + B/T + C \cdot \ln(T) + D \cdot T^2$
Coeff. A	8.08779e+01
Coeff. B	-6.03061e+03
Coeff. C	-1.00863e+01
Coeff. D	9.81251e-06
Temperature range (K), min.	178.01
Temperature range (K), max.	510.00

## Sources

Activity Coefficients at Infinite Dilution of Organic Solutes in Substitution of Phosphoric Acid, N-methylmethyl-Phosphamide, in Selected Solvents: activity coefficients of volatile organic compounds, in two Activity Coefficients at Infinite Dilution of Organic Compounds in Ethanol, 1,2-Dichloroethane, and a mixture of these with water, Using Inverse Gas Chromatography. *Journal of Chromatography*, 1980, 199, 1-11.

Equilibrium Vapor Pressure Using Gas Density Chromatography and Supercritical Carbon Dioxide at Infinite Dilution for Organic Solutes at infinite dilution of organic solutes in mixtures of methanol, 1,2-dichloroethane, and water, and parameters of enthalpy of sorption, using gas liquid chromatography, equilibrium coefficients of systems (water, phosphoric acid) at infinite dilution for hydrocarbons and for alcohols Determination and Correlation of Solubility of n-butylacrylamide in Several Different Solvents at Equilibrium Conditions in the Organic Solvent Systems for rapid boiling point determination of small air sensitive samples under inert atmosphere:

<https://www.doi.org/10.1021/je200050q>

<https://www.doi.org/10.1021/acs.jced.5b00007>

<https://www.doi.org/10.1016/j.jct.2013.05.035>

<https://www.doi.org/10.1021/je800754w>

<https://www.doi.org/10.1016/j.jct.2012.09.017>

<https://www.doi.org/10.1021/je700560s>

<https://www.doi.org/10.1021/je700019h>

<https://www.doi.org/10.1016/j.ijct.2016.07.021>

<https://www.doi.org/10.1016/j.ijct.2013.10.038>

<https://www.doi.org/10.1016/j.ijct.2018.09.023>

<https://www.doi.org/10.1016/j.ijct.2013.11.034>

<https://www.doi.org/10.1021/je1005517>

<https://www.doi.org/10.1021/acs.iced.5b00135>

<https://www.doi.org/10.1021/je0498560>

<https://www.doi.org/10.1016/j.tca.2017.05.005>

<https://www.doi.org/10.1021/je500050p>  
<https://www.doi.org/10.1016/j.jct.2014.06.021>  
<https://www.doi.org/10.1021/acs.jced.9b00360>  
<https://www.doi.org/10.1021/je020150k>  
<https://www.doi.org/10.1016/j.jct.2015.08.016>  
<https://www.doi.org/10.1016/j.fluid.2015.02.022>  
<http://onschallenge.wikispaces.com/file/view/AqueousDatas>  
<https://www.doi.org/10.1016/j.tca.2012.06.025>  
<https://www.doi.org/10.1021/je100626x>  
<https://www.doi.org/10.1021/acs.jced.8b01205>  
<https://www.doi.org/10.1021/je060305e>  
<https://www.doi.org/10.1021/je400412z>  
<https://www.doi.org/10.1021/je4001334>  
<http://webbook.nist.gov/cgi/cbook.cgi?ID=C75092&Units=SI>

<https://www.doi.org/10.1021/je100369a>  
<http://link.springer.com/article/10.1007/BF02311772>  
<https://www.doi.org/10.1021/acs.jced.5b00857>  
<https://www.doi.org/10.1021/je8007815>  
<https://www.doi.org/10.1021/je900704b>  
<https://www.doi.org/10.1021/je060033f>  
<https://www.doi.org/10.1021/je7007389>  
<https://www.doi.org/10.1021/je700640r>  
<https://www.doi.org/10.1021/acs.jced.9b00445>  
<https://www.doi.org/10.1021/je700069g>  
<https://www.doi.org/10.1021/je050242k>  
<https://www.doi.org/10.1021/je200822w>  
<https://www.doi.org/10.1021/acs.jced.8b00080>  
<https://www.doi.org/10.1021/acs.jced.8b00581>  
<https://www.doi.org/10.1021/acs.jced.8b01017>  
<https://www.doi.org/10.1016/j.jct.2009.05.003>  
<https://www.doi.org/10.1021/je7007457>  
<https://www.doi.org/10.1016/j.jct.2007.01.004>  
<https://www.doi.org/10.1021/je800369k>  
<https://www.doi.org/10.1021/je200637v>  
<https://www.doi.org/10.1021/je900838a>  
<https://www.doi.org/10.1016/j.jct.2012.11.021>  
[http://pubs.acs.org/doi/suppl/10.1021/ci034243x/suppl\\_file/ci034243xsi20040112\\_053635.txt](http://pubs.acs.org/doi/suppl/10.1021/ci034243x/suppl_file/ci034243xsi20040112_053635.txt)

<https://www.doi.org/10.1021/acs.jced.9b00432>  
<https://www.doi.org/10.1021/je1009812>  
<https://www.doi.org/10.1021/acs.jced.6b00664>  
<https://www.doi.org/10.1016/j.jct.2018.09.003>  
<https://www.doi.org/10.1021/je800658v>  
<https://www.doi.org/10.1021/acs.jced.7b00119>

Activity Coefficients at Infinite Dilution by GLC in Alkanediamines as Solubility of Enrofloxacin Sodium in Various Solvents at Various Temperatures:

Solubilities of Adefovir Dipivoxil in Different Binary Solvents at 298.15 K: Solubilities of Phosphorus-Containing Compounds in Selected Solvents: Densities and Excess Molar Volumes for Binary Mixtures of Ionic Liquid Solubility of Acetophenone in Different Solvents from 293.15 to 327.60 K: Activity Coefficients at Infinite Dilution of Organic Solutes in

Measurements and Thermodynamic Models for Separation Gas-Liquid Equilibrium Data: Activity Coefficients at Infinite Dilution of Selected Organic Compounds in Binary Mixtures of CO<sub>2</sub> and Dichloromethane

Solubility of 1,2-Dichloroethane in the Ionic Liquid 2,2,2-Trifluoroethylmethylphosphorinane-5,5-dimethyl-2-

(2,4,6-Trichlorophenyl)-1,3,5-Triazine: Thermodynamic and Activity of Separation

Thermodynamic and Activity of Separation: Thermodynamic and Activity of Separation: Thermodynamic and Activity of Separation

Solubility of 1,2-Dichloroethane in the Ionic Liquid 2,2,2-Trifluoroethylmethylphosphorinane-5,5-dimethyl-2-

(2,4,6-Trichlorophenyl)-1,3,5-Triazine: Thermodynamic and Activity of Separation

Solubility of 1,2-Dichloroethane in the Ionic Liquid 2,2,2-Trifluoroethylmethylphosphorinane-5,5-dimethyl-2-

(2,4,6-Trichlorophenyl)-1,3,5-Triazine: Thermodynamic and Activity of Separation

Solubility of 1,2-Dichloroethane in the Ionic Liquid 2,2,2-Trifluoroethylmethylphosphorinane-5,5-dimethyl-2-

(2,4,6-Trichlorophenyl)-1,3,5-Triazine: Thermodynamic and Activity of Separation

Solubility of 1,2-Dichloroethane in the Ionic Liquid 2,2,2-Trifluoroethylmethylphosphorinane-5,5-dimethyl-2-

(2,4,6-Trichlorophenyl)-1,3,5-Triazine: Thermodynamic and Activity of Separation

Solubility of 1,2-Dichloroethane in the Ionic Liquid 2,2,2-Trifluoroethylmethylphosphorinane-5,5-dimethyl-2-

(2,4,6-Trichlorophenyl)-1,3,5-Triazine: Thermodynamic and Activity of Separation

Solubility of 1,2-Dichloroethane in the Ionic Liquid 2,2,2-Trifluoroethylmethylphosphorinane-5,5-dimethyl-2-

(2,4,6-Trichlorophenyl)-1,3,5-Triazine: Thermodynamic and Activity of Separation

Solubility of 1,2-Dichloroethane in the Ionic Liquid 2,2,2-Trifluoroethylmethylphosphorinane-5,5-dimethyl-2-

(2,4,6-Trichlorophenyl)-1,3,5-Triazine: Thermodynamic and Activity of Separation

Solubility of 1,2-Dichloroethane in the Ionic Liquid 2,2,2-Trifluoroethylmethylphosphorinane-5,5-dimethyl-2-

(2,4,6-Trichlorophenyl)-1,3,5-Triazine: Thermodynamic and Activity of Separation

Solubility of 1,2-Dichloroethane in the Ionic Liquid 2,2,2-Trifluoroethylmethylphosphorinane-5,5-dimethyl-2-

(2,4,6-Trichlorophenyl)-1,3,5-Triazine: Thermodynamic and Activity of Separation

Solubility of 1,2-Dichloroethane in the Ionic Liquid 2,2,2-Trifluoroethylmethylphosphorinane-5,5-dimethyl-2-

(2,4,6-Trichlorophenyl)-1,3,5-Triazine: Thermodynamic and Activity of Separation

<https://www.doi.org/10.1021/je200628n>

<https://www.doi.org/10.1021/je8005826>

<https://www.cheric.org/research/kdb/hcprop/showprop.php?cmpid=1526>

<https://www.doi.org/10.1021/je700664g>

<https://www.doi.org/10.1021/je100341q>

<https://www.doi.org/10.1021/je1002237>

<https://www.doi.org/10.1021/je0603630>

<https://www.doi.org/10.1021/je800218g>

<https://www.doi.org/10.1021/acs.jced.7b00160>

<https://www.doi.org/10.1021/acs.jced.8b00635>

<https://www.doi.org/10.1016/j.fluid.2013.10.037>

<https://www.doi.org/10.1016/j.jct.2014.04.020>

<https://www.doi.org/10.1016/j.fluid.2013.09.023>

<https://www.doi.org/10.1016/j.jct.2018.05.017>

<https://www.doi.org/10.1016/j.jct.2018.08.028>

<https://www.doi.org/10.1021/je400813d>

<https://www.doi.org/10.1016/j.jct.2008.08.013>

<https://www.doi.org/10.1021/acs.jced.7b00840>

<https://www.doi.org/10.1021/je800468h>

<https://www.doi.org/10.1016/j.fluid.2015.11.037>

[https://en.wikipedia.org/wiki/Joback\\_method](https://en.wikipedia.org/wiki/Joback_method)

<https://www.doi.org/10.1021/je049616k>

<https://www.doi.org/10.1021/je020067x>

<https://www.doi.org/10.1021/je9003178>

<https://www.doi.org/10.1021/acs.jced.8b01193>

<https://www.doi.org/10.1021/acs.jced.5b00268>

<https://www.doi.org/10.1021/je800152d>

<https://www.doi.org/10.1016/j.jct.2017.12.012>

<https://www.doi.org/10.1021/je900502s>

<https://www.doi.org/10.1016/j.jct.2009.02.004>

<https://www.doi.org/10.1021/je201129y>

<https://www.doi.org/10.1021/je200252c>

<https://www.doi.org/10.1021/je800200j>

<https://www.sciencedirect.com/book/9780128029992/the-yaws-handbook-of-vapor-pressure>

<https://www.doi.org/10.1016/j.fluid.2007.03.032>

<https://www.doi.org/10.1016/j.fluid.2014.06.021>

<https://www.doi.org/10.1016/j.tca.2009.07.011>

<https://www.doi.org/10.1016/j.fluid.2012.05.006>

<https://www.doi.org/10.1016/j.jct.2015.02.023>

<https://www.doi.org/10.1021/acs.jced.7b00244>

<https://www.doi.org/10.1021/je2013582>

<https://www.doi.org/10.1021/je300692s>

<https://www.doi.org/10.1021/je049620w>



[illegible]

# Legend

<b>af:</b>	Acentric Factor
<b>affp:</b>	Proton affinity
<b>basg:</b>	Gas basicity
<b>chl:</b>	Standard liquid enthalpy of combustion
<b>cpg:</b>	Ideal gas heat capacity
<b>cpl:</b>	Liquid phase heat capacity
<b>dm:</b>	Dipole Moment
<b>dvisc:</b>	Dynamic viscosity
<b>gf:</b>	Standard Gibbs free energy of formation
<b>gyrad:</b>	Radius of Gyration
<b>hf:</b>	Enthalpy of formation at standard conditions
<b>hfl:</b>	Liquid phase enthalpy of formation at standard conditions
<b>hfus:</b>	Enthalpy of fusion at standard conditions
<b>hfust:</b>	Enthalpy of fusion at a given temperature
<b>hvap:</b>	Enthalpy of vaporization at standard conditions
<b>hvapt:</b>	Enthalpy of vaporization at a given temperature
<b>ie:</b>	Ionization energy
<b>log10ws:</b>	Log10 of Water solubility in mol/l
<b>logp:</b>	Octanol/Water partition coefficient
<b>mcvol:</b>	McGowan's characteristic volume
<b>nfpah:</b>	NFPA Health Rating
<b>pc:</b>	Critical Pressure
<b>pvap:</b>	Vapor pressure
<b>rfi:</b>	Refractive Index
<b>rhof:</b>	Liquid Density
<b>rinpol:</b>	Non-polar retention indices
<b>ripol:</b>	Polar retention indices
<b>sl:</b>	Liquid phase molar entropy at standard conditions
<b>srf:</b>	Surface Tension
<b>tb:</b>	Normal Boiling Point Temperature
<b>tbp:</b>	Boiling point at given pressure
<b>tc:</b>	Critical Temperature
<b>tf:</b>	Normal melting (fusion) point
<b>vc:</b>	Critical Volume
<b>zra:</b>	Rackett Parameter

Latest version available from:

<https://www.chemeo.com/cid/31-051-0/Methylene-chloride.pdf>

Generated by Cheméo on 2024-04-29 15:32:25.847406713 +0000 UTC m=+16693994.767984026.



Cheméo (<https://www.chemeo.com>) is the biggest free database of chemical and physical data for the process industry.