Methylene chloride

Other names:	Aerothene MM
	CH2Cl2
	Chlorure de methylene
	DICHLOROMETHANE
	F 30
	F 30 (chlorocarbon)
	FREON 30
	HCC 30
	Khladon 30
	METHYLENE DICHLORIDE
	Metaclen
	Methane dichloride
	Methane, dichloro-
	Methoklone
	Methylene bichloride
	Metylenu chlorek
	NCI-C50102
	NSC 406122
	Narkotil
	R 30
	R-30
	Rcra waste number U080
	Salesthin
	Solaesthin
	Soleana VDA
	Solmethine
	UN 1593
Inchi:	InChI=1S/CH2Cl2/c2-1-3/h1H2
InchiKey:	YMWUJEATGCHHMB-UHFFFAOYSA-N
Formula:	CH2Cl2
SMILES:	ClCCl
Mol. weight [g/mol]:	84.93
CAS:	75-09-2

Physical Properties

Property code	Value	Unit	Source

af	0.1990		KDB
affp	628.00 ± 8.00	kJ/mol	NIST Webbook
basg	602.00 ± 8.00	kJ/mol	NIST Webbook
chl	-605.80 ± 8.40	kJ/mol	NIST Webbook
chl	-602.50	kJ/mol	NIST Webbook
dm	1.80	debye	KDB
gf	-68.91	kJ/mol	KDB
gyrad	2.4320		KDB
hf	-95.70 ± 1.30	kJ/mol	NIST Webbook
hf	-95.46	kJ/mol	KDB
hf	-95.10 ± 2.50	kJ/mol	NIST Webbook
hfl	-124.30	kJ/mol	NIST Webbook
hfl	-124.10 ± 2.50	kJ/mol	NIST Webbook
hfus	6.74	kJ/mol	Joback Method
hvap	29.00	kJ/mol	NIST Webbook
hvap	29.03 ± 0.08	kJ/mol	NIST Webbook
hvap	30.60 ± 0.10	kJ/mol	NIST Webbook
hvap	28.80	kJ/mol	NIST Webbook
hvap	28.50 ± 0.42	kJ/mol	NIST Webbook
ie	11.33 ± 0.04	eV	NIST Webbook
ie	11.35 ± 0.02	eV	NIST Webbook
ie	11.36	eV	NIST Webbook
ie	11.33	eV	NIST Webbook
ie	11.28	eV	NIST Webbook
ie	11.40	eV	NIST Webbook
ie	11.40	eV	NIST Webbook
ie	11.32	eV	NIST Webbook
ie	11.32 ± 0.01	eV	NIST Webbook
log10ws	-0.63		Aqueous Solubility Prediction Method
log10ws	-0.63		Estimated Solubility Method
logp	1.421		Crippen Method
mcvol	49.430	ml/mol	McGowan Method
nfpah	%!d(float64=2)		KDB
pc	6100.00	kPa	KDB
pc	6355.00 ± 15.00	kPa	NIST Webbook
rinpol	548.40		NIST Webbook
rinpol	520.00		NIST Webbook
rinpol	528.00		NIST Webbook
rinpol	531.00		NIST Webbook
rinpol	531.00		NIST Webbook
rinpol	488.00		NIST Webbook
rinpol	480.00		NIST Webbook
rinpol	520.00		NIST Webbook

rinpol	524.00	NIST Webbook
rinpol	531.60	NIST Webbook
rinpol	512.70	NIST Webbook
rinpol	510.00	NIST Webbook
rinpol	514.00	NIST Webbook
rinpol	531.00	NIST Webbook
rinpol	513.00	NIST Webbook
rinpol	511.00	NIST Webbook
rinpol	511.00	NIST Webbook
rinpol	527.00	NIST Webbook
rinpol	528.00	NIST Webbook
rinpol	529.00	NIST Webbook
rinpol	515.00	NIST Webbook
rinpol	519.00	NIST Webbook
rinpol	515.00	NIST Webbook
rinpol	540.00	NIST Webbook
rinpol	515.00	NIST Webbook
rinpol	531.00	NIST Webbook
rinpol	515.00	NIST Webbook
rinpol	512.00	NIST Webbook
rinpol	515.00	NIST Webbook
rinpol	515.00	NIST Webbook
rinpol	504.00	NIST Webbook
rinpol	510.00	NIST Webbook
rinpol	527.00	NIST Webbook
rinpol	477.00	NIST Webbook
rinpol	497.90	NIST Webbook
rinpol	530.00	NIST Webbook
rinpol	553.70	NIST Webbook
rinpol	524.00	NIST Webbook
rinpol	524.00	NIST Webbook
rinpol	524.00	NIST Webbook
rinpol	486.00	NIST Webbook
rinpol	518.00	NIST Webbook
rinpol	506.30	NIST Webbook
rinpol	520.00	NIST Webbook
rinpol	519.00	NIST Webbook
rinpol	537.90	NIST Webbook
rinpol	537.80	NIST Webbook
rinpol	515.00	NIST Webbook
rinpol	516.90	NIST Webbook
rinpol	508.00	NIST Webbook
rinpol	504.90	NIST Webbook
rinpol	518.00	NIST Webbook

rinpol	542.20		NIST Webbook
rinpol	531.00		NIST Webbook
rinpol	511.00		NIST Webbook
rinpol	512.70		NIST Webbook
rinpol	553.50		NIST Webbook
rinpol	555.90		NIST Webbook
rinpol	516.50		NIST Webbook
rinpol	540.00		NIST Webbook
ripol	919.00		NIST Webbook
ripol	948.00		NIST Webbook
ripol	946.00		NIST Webbook
ripol	935.70		NIST Webbook
ripol	932.62		NIST Webbook
ripol	926.65		NIST Webbook
ripol	933.00		NIST Webbook
ripol	933.00		NIST Webbook
ripol	914.00		NIST Webbook
ripol	931.00		NIST Webbook
ripol	948.00		NIST Webbook
ripol	948.00		NIST Webbook
ripol	905.00		NIST Webbook
ripol	937.20		NIST Webbook
ripol	936.00		NIST Webbook
ripol	927.00		NIST Webbook
ripol	933.00		NIST Webbook
ripol	953.00		NIST Webbook
ripol	944.00		NIST Webbook
ripol	932.00		NIST Webbook
ripol	944.00		NIST Webbook
ripol	933.00		NIST Webbook
ripol	927.00		NIST Webbook
ripol	933.00		NIST Webbook
ripol	948.00		NIST Webbook
ripol	937.20		NIST Webbook
ripol	917.00		NIST Webbook
ripol	914.00		NIST Webbook
ripol	933.00		NIST Webbook
ripol	937.00		NIST Webbook
ripol	912.00		NIST Webbook
ripol	928.00		NIST Webbook
ripol	925.00		NIST Webbook
ripol	925.00		NIST Webbook
sl	174.50	J/mol×K	NIST Webbook
tb	313.30 ± 0.30	K	NIST Webbook

tb	314.70 ± 0.50	K	NIST Webbook
tb	312.95 ± 0.30	K	NIST Webbook
tb	313.35 ± 0.20	K	NIST Webbook
tb	313.20 ± 1.00	K	NIST Webbook
tb	313.20 ± 0.50	K	NIST Webbook
tb	313.00	K	NIST Webbook
tb	314.95 ± 0.50	K	NIST Webbook
tb	312.95 ± 0.50	K	NIST Webbook
tb	313.00	K	KDB
tb	312.93 ± 0.20	K	NIST Webbook
tb	312.92 ± 0.07	K	NIST Webbook
tb	313.15 ± 1.00	K	NIST Webbook
tc	510.00	K	NIST Webbook
tc	510.00	K	KDB
tc	508.00 ± 0.20	K	NIST Webbook
tf	178.01	K	KDB
tf	176.00 ± 1.50	K	NIST Webbook
tf	198.06 ± 0.40	K	NIST Webbook
tf	176.65 ± 0.40	K	NIST Webbook
tf	177.00 ± 2.00	K	NIST Webbook
tf	177.62	K	Aqueous Solubility Prediction Method
vc	0.190	m3/kmol	Joback Method
zra	0.26		KDB

Temperature Dependent Properties

Property code	Value	Unit	Temperature [K]	Source
cpg	55.35	J/mol×K	297.14	Joback Method
cpg	65.49	J/mol×K	448.32	Joback Method
cpg	63.64	J/mol×K	418.08	Joback Method
cpg	61.71	J/mol×K	387.85	Joback Method
cpg	59.68	J/mol×K	357.61	Joback Method
cpg	57.56	J/mol×K	327.38	Joback Method
cpg	67.25	J/mol×K	478.55	Joback Method
cpl	100.50	J/mol×K	292.50	NIST Webbook
cpl	100.80	J/mol×K	292.50	NIST Webbook
cpl	102.30	J/mol×K	298.15	NIST Webbook
cpl	105.50	J/mol×K	303.20	NIST Webbook
cpl	100.00	J/mol×K	298.00	NIST Webbook
cpl	100.00	J/mol×K	298.10	NIST Webbook

cpl	129.30	J/mol×K	298.00	NIST Webbook
dvisc	0.0014784	Pa×s	183.58	Joback Method
dvisc	0.0026264	Pa×s	160.87	Joback Method
dvisc	0.0006594	Pa×s	229.00	Joback Method
dvisc	0.0004912	Pa×s	251.72	Joback Method
dvisc	0.0003842	Pa×s	274.43	Joback Method
dvisc	0.0003120	Pa×s	297.14	Joback Method
dvisc	0.0009444	Pa×s	206.29	Joback Method
hfust	6.16	kJ/mol	178.20	NIST Webbook
hfust	6.16	kJ/mol	178.22	NIST Webbook
hfust	6.16	kJ/mol	178.20	NIST Webbook
hvapt	28.06	kJ/mol	313.00	NIST Webbook
hvapt	30.20	kJ/mol	273.00	NIST Webbook
hvapt	29.00	kJ/mol	347.00	NIST Webbook
hvapt	30.30	kJ/mol	287.50	NIST Webbook
hvapt	29.20	kJ/mol	308.00	NIST Webbook
hvapt	29.40	kJ/mol	249.00	NIST Webbook
pvap	45.70	kPa	291.90	Effect of Dissolved Poly(lactic acid) on the Solubility of CO2, N2, and He Gases in Dichloromethane
pvap	69.33	kPa	302.67	Limiting activity coefficient measurements in binary mixtures of dichloromethane and 1-alkanols (C1 - C4)
pvap	197.94	kPa	333.15	Vapor Liquid Equilibrium and Excess Enthalpy Data for Systems Containing N,N-Dimethylacetamide
pvap	70.49	kPa	303.15	Total Vapor Pressure Measurements for 2-Ethoxyethanol with Carbon Tetrachloride, Chloroform, and Dichloromethane at 303.15 K
pvap	79.99	kPa	306.43	Limiting activity coefficient measurements in binary mixtures of dichloromethane and 1-alkanols (C1 - C4)

pvap	53.33	kPa	296.05	Limiting activity coefficient measurements in binary mixtures of dichloromethane and 1-alkanols (C1 - C4)
pvap	58.66	kPa	298.41	Limiting activity coefficient measurements in binary mixtures of dichloromethane and 1-alkanols (C1 - C4)
pvap	85.33	kPa	308.17	Limiting activity coefficient measurements in binary mixtures of dichloromethane and 1-alkanols (C1 - C4)
pvap	90.66	kPa	309.82	Limiting activity coefficient measurements in binary mixtures of dichloromethane and 1-alkanols (C1 - C4)
pvap	95.99	kPa	311.39	Limiting activity coefficient measurements in binary mixtures of dichloromethane and 1-alkanols (C1 - C4)
pvap	37.10	kPa	287.00	Effect of Dissolved Poly(lactic acid) on the Solubility of CO2, N2, and He Gases in Dichloromethane
pvap	67.40	kPa	301.80	Effect of Dissolved Poly(lactic acid) on the Solubility of CO2, N2, and He Gases in Dichloromethane
pvap	63.99	kPa	300.61	Limiting activity coefficient measurements in binary mixtures of dichloromethane and 1-alkanols (C1 - C4)

pvap	37.33	kPa	287.58	Limiting activity coefficient measurements in binary mixtures of dichloromethane and 1-alkanols (C1 - C4)
pvap	41.33	kPa	289.94	Limiting activity coefficient measurements in binary mixtures of dichloromethane and 1-alkanols (C1 - C4)
pvap	45.33	kPa	292.12	Limiting activity coefficient measurements in binary mixtures of dichloromethane and 1-alkanols (C1 - C4)
pvap	49.33	kPa	294.15	Limiting activity coefficient measurements in binary mixtures of dichloromethane and 1-alkanols (C1 - C4)
pvap	55.90	kPa	296.80	Effect of Dissolved Poly(lactic acid) on the Solubility of CO2, N2, and He Gases in Dichloromethane
pvap	34.66	kPa	285.89	Limiting activity coefficient measurements in binary mixtures of dichloromethane and 1-alkanols (C1 - C4)
pvap	74.66	kPa	304.61	Limiting activity coefficient measurements in binary mixtures of dichloromethane and 1-alkanols (C1 - C4)
rfi	1.42760		288.15	Partial Molar Volumes of N,N'-1,2-Ethyl-bis(salicyladimine) Schiff Base (Salen) in Organic Solvents at T = (283.15 to 318.15) K

rfi	1.42190		298.15	Partial Molar Volumes of N,N'-1,2-Ethyl-bis(salicyladimine) Schiff Base (Salen) in Organic Solvents at T = (283.15 to 318.15) K
rfi	1.43050		283.15	Partial Molar Volumes of N,N'-1,2-Ethyl-bis(salicyladimine) Schiff Base (Salen) in Organic Solvents at T = (283.15 to 318.15) K
rfi	1.42370		293.15	Solubilities of Phosphorus-Containing Compounds in Selected Solvents
rfi	1.42370		293.15	Solubilities of Some Phosphaspirocyclic Compounds in Selected Solvents
rfi	1.42510		293.15	Partial Molar Volumes of N,N'-1,2-Ethyl-bis(salicyladimine) Schiff Base (Salen) in Organic Solvents at T = (283.15 to 318.15) K
rhol	1325.67	kg/m3	293.15	Volumetric properties of dichloromethane with aniline or nitrobenzene at different temperatures: A theoretical and experimental study
rhol	1316.58	kg/m3	298.15	Volumetric properties of dichloromethane with aniline or nitrobenzene at different temperatures: A theoretical and experimental study

rhol	1307.90	kg/m3	303.15	Volumetric properties of dichloromethane with aniline or nitrobenzene at different temperatures: A theoretical and experimental study
rhol	1334.81	kg/m3	288.15	Volumetric properties of dichloromethane with aniline or nitrobenzene at different temperatures: A theoretical and experimental study
rhol	1317.00	kg/m3	298.00	KDB
rhol	1325.79	kg/m3	293.15	Volumetric properties of binary liquid mixtures: Application of the Prigogine Flory Patterson theory to excess molar volumes of dichloromethane with benzene or toluene
rhol	1316.75	kg/m3	298.15	Volumetric properties of binary liquid mixtures: Application of the Prigogine Flory Patterson theory to excess molar volumes of dichloromethane with benzene or toluene
rhol	1307.51	kg/m3	303.15	Volumetric properties of binary liquid mixtures: Application of the Prigogine Flory Patterson theory to excess molar volumes of dichloromethane with benzene or toluene

rhol	1307.60	kg/m3	303.15	Viscosity and Density for Binary Mixtures of Carbon Tetrachloride + Chloroform, Carbon Tetrachloride + Dichloromethane, and Chloroform + Dichloromethane and One Ternary Mixture of Chloroform + 1:1 (Carbon Tetrachloride + Dichloromethane) at 303.15 K
rhol	1327.00	kg/m3	293.15	Interfacial Properties, Densities, and Contact Angles of Task Specific Ionic Liquids
rhol	1334.08	kg/m3	288.15	Volumetric properties of binary liquid mixtures: Application of the Prigogine Flory Patterson theory to excess molar volumes of dichloromethane with benzene or toluene
srf	0.03	N/m	293.20	KDB

Pressure Dependent Properties

Property code	Value	Unit	Pressure [kPa]	Source
tbp	313.25	K	96.60	Low cost apparatus for rapid boiling point determination of small air sensitive samples under inert atmosphere

Correlations

Information	Value
Property code	pvap
Equation	ln(Pvp) = A + B/(T + C)
Coeff. A	1.43555e+01
Coeff. B	-2.65134e+03
Coeff. C	-4.07080e+01
Temperature range (K), min.	229.18
Temperature range (K), max.	510.00

Information	Value
Property code	pvap
Equation	ln(Pvp) = A + B/T + C*ln(T) + D*T^2
Coeff. A	8.08779e+01
Coeff. B	-6.03061e+03
Coeff. C	-1.00863e+01
Coeff. D	9.81251e-06
Temperature range (K), min.	178.01
Temperature range (K), max.	510.00

Sources

Study of Ether-, Alcohol-, or Cyano-Functionalized Ionic Liquids Using Inverse Gas Chromatography:	https://www.doi.org/10.1021/je900838a
Measurement and Correlation of the Dissolution Equilibria of o-Iodoaniline and p-Iodoaniline in Pure Solvents:	https://www.doi.org/10.1021/acs.jced.7b00840
Activity coefficients at infinite dilution of organic solutes in the ionic liquid 1-ethyl-3-methylimidazolium ethanesulfonate:	https://www.doi.org/10.1016/j.fluid.2010.10.008
Activity Coefficients at Infinite Dilution of Organic Compounds in Four New Imidazolium-Based Ionic Liquids:	https://www.doi.org/10.1021/je200195q
Separation of azeotropic mixtures based on gamma infinity data using [OMIM][BF4] ionic liquid and [OMIM][DCA] ionic liquid:	https://www.doi.org/10.1016/j.jct.2018.09.003
Solubility Measurement and Correlation of (+)-Biotin Intermediate Lactone in Different Organic Solvents from (287.15 to 328.15) K:	https://www.doi.org/10.1021/acs.jced.5b00857
Solubility Measurement and Correlation of 4-Amino-3,6-dichloropyridine in Seven Pure Organic Solvents from (278.15 to 338.15) K:	https://www.doi.org/10.1021/je500286x
Thermodynamics and activity coefficients at infinite dilution for organic compounds in the ionic liquid 1-hexyl-3-methylimidazolium chloride:	https://www.doi.org/10.1016/j.jct.2018.08.028
Activity Coefficients at Infinite Dilution of Organic Compounds in the Ionic Liquid 1-Hexyl-3-methylimidazolium Chloride:	https://www.doi.org/10.1021/je200628n
Solubility Measurement and Correlation of 2-Oxindole in 12 Pure Organic Solvents:	https://www.doi.org/10.1021/acs.jced.9b00308
Liquid-vapor equilibrium data of CO2+ dichloromethane + medroxyprogesterone system:	https://www.doi.org/10.1016/j.fluid.2013.10.037
Aqueous Solubility Prediction Method:	http://onschallenge.wikispaces.com/file/view/AqueousDataset002.xlsx/351826032/AqueousDa
Conductivities of Binary Mixtures of Ionic Liquids with Polar Solvents:	https://www.doi.org/10.1021/je800468h
Solubilities of Rutaecarpine in Twelve Organic Solvents from (283.2 to 323.2) K:	https://www.doi.org/10.1021/je4001334
Determination of Infinite Dilution Activity Coefficients of Several Organic Solutes in 1-Butylpyridinium Nitrate by Gas-Liquid Chromatography:	https://www.doi.org/10.1021/acs.jced.7b00244
Solubilities of Some Pharmaceutical Compounds in Selected Pure Solvents:	https://www.doi.org/10.1021/je060138i
Solubility of Diosgenin in twelve organic solvents from T = (283.2 to 323.2) K:	https://www.doi.org/10.1016/j.jct.2012.11.021
Interactions of Volatile Organic Compounds with the Ionic Liquid 1-Butyl-1-methylpyrrolidinium Dicyanamide:	https://www.doi.org/10.1021/je200822w
Solubilities of (5Z)-2-Amino-a-(methoxyimino)-4-thiazoleethanethioic Acid S-2-Benzothiazolyl Ester in Different Pure Solvents and Binary Mixtures of Tetrahydrofuran + Dichloromethane or 1,2-Dichloroethane:	https://www.doi.org/10.1021/je100890f

High-Pressure Vapor-Liquid Equilbria of Some Carbon Dioxide + Organic Binary Systems:	https://www.doi.org/10.1021/je0498560
Evaluation of the Performance of Trigeminal Tricationic Ionic Liquids for Separation Problems:	https://www.doi.org/10.1021/je201129y
Solubilities of Phenylphosphinic Acid, Methylphenylphosphinic Acid, Hexachlorocyclotriphosphazene, and Hexaphenoxycyclotriphosphazene in Selected Solvents:	https://www.doi.org/10.1021/je1009812
Thermodynamics and selectivity of separation based on activity coefficients at infinite dilution of various solutes in ionic liquid [HMMIM][BF4]:	https://www.doi.org/10.1016/j.jct.2018.05.017
Solubility of Ofloxacin in 1,2-Dichloromethane, Chloroform, Carbon Tetrachloride, and Water from (293.15 to 318.15) K:	https://www.doi.org/10.1021/je900540d
Activity Coefficients at Infinite Dilution of Organic Solutes in 1-Ethyl-3-methylimidazolium Tetrafluoroborate Using Gas-Liquid Chromatography:	https://www.doi.org/10.1021/je800218g
Solubility of Imidazoles, Benzimidazoles, and Phenylimidazoles in Dichloromethane, 1-Chlorobutane, Toluene, and 2-Nitrotoluene:	https://www.doi.org/10.1021/je049907t
Stability of Ion Pairs of Bis(trifluoromethanesulfonyl)amide-Based Ionic Liquids in Dichloromethane:	https://www.doi.org/10.1021/je900694m
NIST Webbook:	http://webbook.nist.gov/cgi/cbook.cgi?ID=C75092&Units=SI
Activity Coefficients at Infinite Dilution of Polar Solutes in 1-(2-Hydroxyethyl)-3-methylimidazolium Tetrafluoroborate Using Gas-Liquid Chromatography:	https://www.doi.org/10.1021/je900704b
Infinite Dilution Activity Coefficients of Solutes Dissolved in Two Trihexyl(tetradecyl)phosphonium Ionic Liquids:	https://www.doi.org/10.1021/je500050p
Measurement of Activity Coefficients of Solutes at Infinite Dilution in (Dimethyl Sulfoxide + Acetamide, or Formamide, or Urea) Using Gas-Liquid Chromatography at the Temperature 298.15 K:	https://www.doi.org/10.1021/je030122h
Solubilities of Tris(o-phenylenedioxy)cyclotriphosphazene	https://www.doi.org/10.1021/je200252c
Determination of Gas Concentrations in Liquids by Nuclear Magnetic Resonance: Hydrogen in Organic Solvents:	https://www.doi.org/10.1021/je2013582
Partial Molar Volumes of N,N'-1,2-Ethyl-bis(salicyladimine) Schiff Base (Salen) in Organic Solvents at T = (283.15 to 318.15) K:	https://www.doi.org/10.1021/je100369a
Ion Pair Formation of Alkylimidazolium Ionic Liquids in Dichloromethane:	https://www.doi.org/10.1021/je800152d
Phase Separation in Binary Mixtures Containing Linear Perfluoroalkanes:	https://www.doi.org/10.1021/je049620w
Solubilities of Phosphorus-Containing Compounds in Selected Solvents:	https://www.doi.org/10.1021/je100341q
Gas Solubilities (CO2, O2, Ar, N2, H2, and He) in Liquid Chlorinated Methanes:	https://www.doi.org/10.1021/je800200j
Thermodynamics and activity coefficients at infinite dilution for organic compounds and water in the ionic liquid 1-butyl-3-methylimidazolium tetracyanoborate:	https://www.doi.org/10.1016/j.jct.2017.07.012
Solubility Measurement and Correlation for ...2,4,6,8,10,12-hexaazaisowurtzitane in Organic Solvents at Temperatures between 283.15 and 333.15 K and Different Chloralkane + Ethyl Acetate Binary Solvents at Temperatures between 283.15 and 323.15 K:	https://www.doi.org/10.1021/acs.jced.6b00664
High-Pressure Phase Equilibria of the Related Substances in the Limonene Oxidation in Supercritical CO2:	https://www.doi.org/10.1021/je020150k
Solubility of Pyoluteorin in Water, Dichloromethane, Chloroform, and Carbon Tetrachloride from (278.2 to 323.2) K:	https://www.doi.org/10.1021/je800369k
Thermodynamic Properties of Mixtures Containing Ionic Liquids. Activity Coefficients at Infinite Dilution of Organic Compounds in 1-Propyl Boronic Acid-3-Alkylimidazolium Bromide and 1-Propenyl-3-alkylimidazolium Bromide Using Inverse Gas Chromatography:	https://www.doi.org/10.1021/je060033f
Excess properties of binary mixtures containing 1-hexyl-3-methylimidazolium bis(trifluoromethylsulfonyl)imide ionic liquid and organic compounds:	https://www.doi.org/10.1016/j.jct.2011.11.007
Solubility Determination, Modeling, and Thermodynamic Dissolution Properties of Benzenesulfonamide in 16 Neat Solvents from 273.15 to 324.45 K:	https://www.doi.org/10.1021/acs.jced.9b00360
Crippen Method:	http://pubs.acs.org/doi/abs/10.1021/ci990307l
Estimated Solubility Method:	http://pubs.acs.org/doi/suppl/10.1021/ci034243x/suppl_file/ci034243xsi20040112_053635.txt
Thermodynamic Studies of Solubility for Naphthalene in 12 Solvents from 279 to 330 K:	https://www.doi.org/10.1021/acs.jced.9b00064
Densities and Excess Molar Volumes for Binary Mixtures of Ionic Liquid 1-Ethyl-3-methylimidazolium Ethylsulfate with Solvents:	https://www.doi.org/10.1021/je1002237
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Legend

af:	Acentric Factor
affp:	Proton affinity
basg:	Gas basicity

chl:	Standard liquid enthalpy of combustion
cpg:	Ideal gas heat capacity
cpl:	Liquid phase heat capacity
dm:	Dipole Moment
dvisc:	Dynamic viscosity
gf:	Standard Gibbs free energy of formation
gyrad:	Radius of Gyration
hf:	Enthalpy of formation at standard conditions
hfl:	Liquid phase enthalpy of formation at standard conditions
hfus:	Enthalpy of fusion at standard conditions
hfust:	Enthalpy of fusion at a given temperature
hvap:	Enthalpy of vaporization at standard conditions
hvapt:	Enthalpy of vaporization at a given temperature
ie:	Ionization energy
log10ws:	Log10 of Water solubility in mol/l
logp:	Octanol/Water partition coefficient
mcvol:	McGowan's characteristic volume
nfpah:	NFPA Health Rating
pc:	Critical Pressure
pvap:	Vapor pressure
rfi:	Refractive Index
rhol:	Liquid Density
rinpol:	Non-polar retention indices
ripol:	Polar retention indices
sl:	Liquid phase molar entropy at standard conditions
srf:	Surface Tension
tb:	Normal Boiling Point Temperature
tbp:	Boiling point at given pressure
tc:	Critical Temperature
tf:	Normal melting (fusion) point
vc:	Critical Volume
zra:	Rackett Parameter