

# Propallylonal

<b>Other names:</b>	2,4,6(1H,3H,5H)-Pyrimidinetrione, 5-(2-bromo-2-propenyl)-5-(1-methylethyl)-Barbituric acid, 5-(2-bromoallyl)-5-isopropyl-5-(2'-Bromoallyl)-5-isopropylbarbituric acid Bromoaprobarbital 5-(2-Bromo-2-propenyl)-5-(1-methylethyl)-2,4,6(1H,3H,5H)-pyrimidinetrione lbomal 5-Isopropyl-5-bromoallylbarbituric acid 5-Isopropyl-5-(2-bromoallyl)barbituate Kwietal Noctal Noctenal Nostal Nostral Propaldon Quietal Quietalum 5-(2-Bromoallyl)-5-isopropylbarbituric acid 5-Isopropyl-5-bromoallylbarbituric acid 5-Isopropyl-5-(2-bromoallyl)barbiturate 5-(«beta»-Bromoallyl)-5-isopropylbarbituric acid
<b>Inchi:</b>	InChI=1S/C10H13BrN2O3/c1-5(2)10(4-6(3)11)7(14)12-9(16)13-8(10)15/h5H,3-4H2,1-2H
<b>InchiKey:</b>	KTGWBBOJAGDSHN-UHFFFAOYSA-N
<b>Formula:</b>	C10H13BrN2O3
<b>SMILES:</b>	<chem>C=C(Br)CC1(C(C)C)C(=O)NC(=O)NC1=O</chem>
<b>Mol. weight [g/mol]:</b>	289.13
<b>CAS:</b>	545-93-7

## Physical Properties

Property code	Value	Unit	Source
gf	-48.90	kJ/mol	Joback Method
hf	-380.96	kJ/mol	Joback Method
hfus	24.07	kJ/mol	Joback Method
hvap	68.85	kJ/mol	Joback Method
log10ws	-2.88		Crippen Method
logp	1.293		Crippen Method
mcvol	178.770	ml/mol	McGowan Method
pc	3677.55	kPa	Joback Method

rmpol	1890.00		NIST Webbook
rmpol	1866.00		NIST Webbook
rmpol	1875.00		NIST Webbook
rmpol	1883.00		NIST Webbook
rmpol	1875.00		NIST Webbook
rmpol	1883.00		NIST Webbook
rmpol	1866.00		NIST Webbook
tb	810.83	K	Joback Method
tc	1087.84	K	Joback Method
tf	677.54	K	Joback Method
vc	0.659	m <sup>3</sup> /kmol	Joback Method

## Temperature Dependent Properties

Property code	Value	Unit	Temperature [K]	Source
cpg	506.29	J/mol×K	810.83	Joback Method
cpg	522.82	J/mol×K	857.00	Joback Method
cpg	538.42	J/mol×K	903.17	Joback Method
cpg	553.14	J/mol×K	949.34	Joback Method
cpg	567.00	J/mol×K	995.50	Joback Method
cpg	580.04	J/mol×K	1041.67	Joback Method
cpg	592.31	J/mol×K	1087.84	Joback Method

## Sources

<b>Joback Method:</b>	<a href="https://en.wikipedia.org/wiki/Joback_method">https://en.wikipedia.org/wiki/Joback_method</a>
<b>McGowan Method:</b>	<a href="http://link.springer.com/article/10.1007/BF02311772">http://link.springer.com/article/10.1007/BF02311772</a>
<b>NIST Webbook:</b>	<a href="http://webbook.nist.gov/cgi/cbook.cgi?ID=C545937&amp;Units=SI">http://webbook.nist.gov/cgi/cbook.cgi?ID=C545937&amp;Units=SI</a>
<b>Crippen Method:</b>	<a href="http://pubs.acs.org/doi/abs/10.1021/ci990307l">http://pubs.acs.org/doi/abs/10.1021/ci990307l</a>
<b>Crippen Method:</b>	<a href="https://www.chemeo.com/doc/models/crippen_log10ws">https://www.chemeo.com/doc/models/crippen_log10ws</a>

## Legend

<b>cpg:</b>	Ideal gas heat capacity
<b>gf:</b>	Standard Gibbs free energy of formation

<b>hf:</b>	Enthalpy of formation at standard conditions
<b>hfus:</b>	Enthalpy of fusion at standard conditions
<b>hvac:</b>	Enthalpy of vaporization at standard conditions
<b>log10ws:</b>	Log10 of Water solubility in mol/l
<b>logp:</b>	Octanol/Water partition coefficient
<b>mcvol:</b>	McGowan's characteristic volume
<b>pc:</b>	Critical Pressure
<b>rinpol:</b>	Non-polar retention indices
<b>tb:</b>	Normal Boiling Point Temperature
<b>tc:</b>	Critical Temperature
<b>tf:</b>	Normal melting (fusion) point
<b>vc:</b>	Critical Volume

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