

# 2-Propenoic acid, 1-methylpropyl ester

<b>Other names:</b>	Acrylic acid, sec-butyl ester s-Butyl acrylate sec-Butyl Acrylate
<b>Inchi:</b>	InChI=1S/C7H12O2/c1-4-6(3)9-7(8)5-2/h5-6H,2,4H2,1,3H3
<b>InchiKey:</b>	RNOOHTVUSNIPCJ-UHFFFAOYSA-N
<b>Formula:</b>	C7H12O2
<b>SMILES:</b>	C=CC(=O)OC(C)CC
<b>Mol. weight [g/mol]:</b>	128.17
<b>CAS:</b>	2998-08-5

## Physical Properties

Property code	Value	Unit	Source
gf	-140.46	kJ/mol	Joback Method
hf	-312.46	kJ/mol	Joback Method
hfus	11.87	kJ/mol	Joback Method
hvap	39.27	kJ/mol	Joback Method
log10ws	-1.58		Crippen Method
logp	1.514		Crippen Method
mcvol	112.630	ml/mol	McGowan Method
pc	3107.10	kPa	Joback Method
tb	432.09	K	Joback Method
tc	616.21	K	Joback Method
tf	224.05	K	Joback Method
vc	0.426	m3/kmol	Joback Method

## Temperature Dependent Properties

Property code	Value	Unit	Temperature [K]	Source
cpg	222.93	J/molxK	432.09	Joback Method
cpg	233.73	J/molxK	462.78	Joback Method
cpg	244.12	J/molxK	493.46	Joback Method
cpg	254.10	J/molxK	524.15	Joback Method
cpg	263.67	J/molxK	554.84	Joback Method
cpg	272.85	J/molxK	585.52	Joback Method

cpg	281.64	J/mol×K	616.21	Joback Method
dvisc	0.0042260	Paxs	224.05	Joback Method
dvisc	0.0019297	Paxs	258.72	Joback Method
dvisc	0.0010605	Paxs	293.40	Joback Method
dvisc	0.0006615	Paxs	328.07	Joback Method
dvisc	0.0004515	Paxs	362.74	Joback Method
dvisc	0.0003294	Paxs	397.42	Joback Method
dvisc	0.0002529	Paxs	432.09	Joback Method

## Correlations

Information	Value
Property code	pvap
Equation	$\ln(P_{vp}) = A + B/(T + C)$
Coeff. A	1.38311e+01
Coeff. B	-3.38339e+03
Coeff. C	-5.35040e+01
Temperature range (K), min.	303.32
Temperature range (K), max.	450.63

## Sources

<b>Crippen Method:</b>	<a href="http://pubs.acs.org/doi/abs/10.1021/ci9903071">http://pubs.acs.org/doi/abs/10.1021/ci9903071</a>
<b>Crippen Method:</b>	<a href="https://www.chemeo.com/doc/models/crippen_log10ws">https://www.chemeo.com/doc/models/crippen_log10ws</a>
<b>Joback Method:</b>	<a href="https://en.wikipedia.org/wiki/Joback_method">https://en.wikipedia.org/wiki/Joback_method</a>
<b>McGowan Method:</b>	<a href="http://link.springer.com/article/10.1007/BF02311772">http://link.springer.com/article/10.1007/BF02311772</a>
<b>NIST Webbook:</b>	<a href="http://webbook.nist.gov/cgi/cbook.cgi?ID=C2998085&amp;Units=SI">http://webbook.nist.gov/cgi/cbook.cgi?ID=C2998085&amp;Units=SI</a>
<b>The Yaws Handbook of Vapor Pressure:</b>	<a href="https://www.sciencedirect.com/book/9780128029992/the-yaws-handbook-of-vapor-pressure">https://www.sciencedirect.com/book/9780128029992/the-yaws-handbook-of-vapor-pressure</a>

## Legend

<b>cpg:</b>	Ideal gas heat capacity
<b>dvisc:</b>	Dynamic viscosity
<b>gf:</b>	Standard Gibbs free energy of formation
<b>hf:</b>	Enthalpy of formation at standard conditions

<b>hfus:</b>	Enthalpy of fusion at standard conditions
<b>hvap:</b>	Enthalpy of vaporization at standard conditions
<b>log10ws:</b>	Log10 of Water solubility in mol/l
<b>logp:</b>	Octanol/Water partition coefficient
<b>mcvol:</b>	McGowan's characteristic volume
<b>pc:</b>	Critical Pressure
<b>pvap:</b>	Vapor pressure
<b>tb:</b>	Normal Boiling Point Temperature
<b>tc:</b>	Critical Temperature
<b>tf:</b>	Normal melting (fusion) point
<b>vc:</b>	Critical Volume

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