Tetraethylammonium bromide

Other names:	Ammonium, tetraethyl-, bromide
	Beparon
	Bromethyl
	Bromure de tetraethylammonium
	Etambro
	Etamon
	Ethanaminium, N,N,N-triethyl-, bromide
	Ethylon
	Etylon
	Sympatektoman
	TEA
	TEA bromide
	TEAB
	TMD 10
	Teamon
	Tetranium
	Tetrylammonium Bromide
	USAF do-32
Inchi:	InChI=1S/C8H20N.BrH/c1-5-9(6-2,7-3)8-4;/h5-8H2,1-4H3;1H/q+1;/p-1
InchiKey:	HWCKGOZZJDHMNC-UHFFFAOYSA-M
Formula:	C8H20BrN
SMILES:	CC[N+](CC)(CC)CC.[Br-]
Mol. weight [g/mol]:	210.16
CAS:	71-91-0

Physical Properties

Property code	Value	Unit	Source
tf	447.00 ± 2.00	K	NIST Webbook

Temperature Dependent Properties

Property code	Value	Unit	Temperature [K]	Source
cps	246.90	J/mol×K	298.00	NIST Webbook

hfust	20.30	kJ/mol	447.00	NIST Webbook
sfust	45.00	J/mol×K	447.00	NIST Webbook

Sources

Viscometric study of glycine, L-alanine, glycylglycine in aqueous tetra-n-alkylammonium bromide solutions at 298.15 K:	https://www.doi.org/10.1016/j.jct.2004.07.006
Electrical Conductivity Studies of Tetraalkylammonium Bromides in Aqueous Solutions at Several Temperatures:	https://www.doi.org/10.1021/je101043c
Electrical Conductance of Some Tetraalkylammonium Bromide Salts in 2-Butoxyethanol (1) + Water (2) Mixtures at 298.15, 303.15, 308.15, and 313.15 K:	https://www.doi.org/10.1021/je4004405
Vapor Pressure Osmometry Determination of Osmotic and Activity Coefficients of Dilute Aqueous Solutions of Symmetrical Tetraalkylammonium Halides at 308.15 K: NIST Webbook:	https://www.doi.org/10.1021/je400821q
Surface and Micellar Properties of Ionic Liquids 1-Dodecyl-3-methylimidazolium Bromide and Salting-out effects of organic and inorganic ammonium salts on the polymer solutions:	https://www.doi.org/10.1021/je5010005
Salting-in and salting-out effects of Organic Electrolytes:	https://www.doi.org/10.1016/j.jct.2018.03.030
NIST Webbook:	http://webbook.nist.gov/cgi/cbook.cgi?ID=C71910&Units=SI
Experimental study of the excess molar volume of ternary mixtures containing (water + 1,2-propanediol, or 1,3-butanediol, or 2,3-butanediol, or 1,4-butanediol, or 2-butoxyethanol) at a temperature of 298.15 K and atmospheric pressure:	https://www.doi.org/10.1016/j.jct.2009.11.018
Investigation of solubilities of carbon dioxide in five ionic liquids based systems for cases with both polymers and electrolytes at 298.15 K and atmospheric pressure:	https://www.doi.org/10.1016/j.jct.2016.08.015
The capability of tetra-alkyl ammonium bromides for aqueous biphasic systems formation at 298.15 K and atmospheric pressure:	https://www.doi.org/10.1016/j.fluid.2018.03.002
Interactions of some ex-amino acids with tetra-n-alkylammonium bromides in aqueous medium at different temperatures:	https://www.doi.org/10.1016/j.fluid.2011.09.016
Standard partial molar volumes of some electrolytes in ethylene carbonate based mixtures:	https://www.doi.org/10.1016/j.jct.2006.08.010
Temperature dependence of the pair interaction between hydrophobic and hydrophilic solutes: A calorimetric study:	https://www.doi.org/10.1016/j.jct.2003.09.005
Thermodynamic Parameters of Anionic Surfactant-Additive Systems at the Cloud Point:	https://www.doi.org/10.1016/j.tca.2005.05.012
Effect of tetra-n-alkylammonium bromides on the volumetric properties of glycine, L-alanine and glycylglycine at 298.15 K:	https://www.doi.org/10.1021/je100502b
Compressibilities of Tetraalkylammonium Bromides in Aqueous Propane-1,2-diol: An Attempt to Test Hydration Limits:	https://www.doi.org/10.1016/j.jct.2003.09.008
Densities and Viscosities of Polyhydroxy Solutes in Aqueous Tetraethylammonium Bromide Solutions at Different Temperatures:	https://www.doi.org/10.1021/acs.jced.5b00964
	https://www.doi.org/10.1021/acs.jced.5b00940

Legend

cps:	Solid phase heat capacity
hfust:	Enthalpy of fusion at a given temperature
sfust:	Entropy of fusion at a given temperature
tf:	Normal melting (fusion) point

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