

D-Valine

Other names:	(R)-Valine (S)-(+)-valine (S)-3-amino-2-methylbutanoic acid (S)-Valine 2-Amino-3-methylbutanoic acid(D) L-2-amino-3-methylbutanoic acid L-2-amino-3-methylbutyric acid L-2-aminoisovaleric acid L-valine NSC 20654 Valine, D- butanoic acid, 2-amino-3-methyl-, (S)-
Inchi:	InChI=1S/C5H11NO2/c1-3(2)4(6)5(7)8/h3-4H,6H2,1-2H3,(H,7,8)/t4-/m0/s1
InchiKey:	KZSNJWFQEVHDMF-BYPYZUCNSA-N
Formula:	C5H11NO2
SMILES:	CC(C)C(N)C(=O)O
Mol. weight [g/mol]:	117.15
CAS:	640-68-6

Physical Properties

Property code	Value	Unit	Source
gf	-212.95	kJ/mol	Joback Method
hf	-388.11	kJ/mol	Joback Method
hfus	12.54	kJ/mol	Joback Method
hvap	60.01	kJ/mol	Joback Method
log10ws	-0.32		Crippen Method
logp	0.054		Crippen Method
mcvol	98.730	ml/mol	McGowan Method
pc	4627.70	kPa	Joback Method
tb	531.50	K	Joback Method
tc	722.61	K	Joback Method
tf	310.12	K	Joback Method
vc	0.357	m3/kmol	Joback Method

Temperature Dependent Properties

Property code	Value	Unit	Temperature [K]	Source
cpg	232.36	J/mol×K	531.50	Joback Method
cpg	241.05	J/mol×K	563.35	Joback Method
cpg	249.30	J/mol×K	595.20	Joback Method
cpg	257.14	J/mol×K	627.06	Joback Method
cpg	264.58	J/mol×K	658.91	Joback Method
cpg	271.63	J/mol×K	690.76	Joback Method
cpg	278.30	J/mol×K	722.61	Joback Method
cps	158.00	J/mol×K	298.00	NIST Webbook
cps	158.20	J/mol×K	298.00	NIST Webbook

Sources

[illegible]

<https://www.doi.org/10.1016/j.jct.2015.10.002>

<https://www.doi.org/10.1021/je100857s>

<https://www.doi.org/10.1016/j.jct.2013.11.015>

<https://www.doi.org/10.1021/je060149b>

<https://www.doi.org/10.1016/j.tca.2014.06.028>

https://en.wikipedia.org/wiki/Joback_method

<https://www.doi.org/10.1021/acs.jced.7b00257>

<https://www.doi.org/10.1016/j.jct.2016.06.026>

<https://www.doi.org/10.1016/j.jct.2015.08.009>

<https://www.doi.org/10.1016/j.tca.2005.04.031>

<https://www.doi.org/10.1021/je8001464>

<https://www.doi.org/10.1021/je500255d>

<https://www.doi.org/10.1016/j.ijct.2013.11.002>

<http://webbook.nist.gov/cgi/cbook.cgi?ID=C640686&Units=SI>

<https://www.doi.org/10.1016/j.ijct.2013.08.010>

<https://www.doi.org/10.1016/j.tca.2008.10.023>

<https://www.doi.org/10.1016/j.ijct.2008.07.019>

<https://www.doi.org/10.1007/s10765-015-2006-0>

<https://www.doi.org/10.1016/j.ijct.2017.12.010>

<https://www.doi.org/10.1021/acs.jced.9b00026>

<https://www.doi.org/10.1016/j.ijct.2011.01.004>

<https://www.doi.org/10.1016/j.ijct.2004.12.015>

<https://www.doi.org/10.1016/j.ijct.2019.06.002>

Volumetric, ultrasonic, and viscometric behaviour of glycine, DL-alanine, and L-valine in interactions between some amino acids and carbon tetrachloride aqueous solutions at 298.15 K. Studies of Amino Acids in Mannitol Aqueous Solutions at 298.15 and 310.15 K: Solubility of L-valine in Water, Ethanol, Nitroform, acetonitrile and acetone, and solvent-solute interactions at 313.15 and 318.15 K. Thermodynamic Aspects Concerning Studies of Some Amino Acids in Aqueous Solution at Different Temperatures: Determination of Solubility and Thermodynamic Properties of Glycine and L-Alanine in Aqueous Solutions as a Function of Temperature; Solubility of L-Alanine, L-Valine and L-Isoleucine in Aqueous Solutions at Different Temperatures; Thermodynamic Properties of Glycine, L-Alanine, and L-Valine in Aqueous Solutions at Different Temperatures.

<https://www.doi.org/10.1016/j.jct.2017.02.021>
<https://www.doi.org/10.1016/j.jct.2011.12.020>
<https://www.doi.org/10.1016/j.jct.2017.02.024>
<https://www.doi.org/10.1016/j.jct.2013.10.022>
<https://www.doi.org/10.1016/j.jct.2016.03.016>
<https://www.doi.org/10.1007/s10765-008-0514-x>
<https://www.doi.org/10.1016/j.jct.2017.03.025>
<https://www.doi.org/10.1016/j.jct.2019.03.011>
<https://www.doi.org/10.1021/je049582g>
<https://www.doi.org/10.1021/je500975a>
<https://www.doi.org/10.1016/j.jct.2016.09.027>
<http://link.springer.com/article/10.1007/BF02311772>
<https://www.doi.org/10.1016/j.jct.2018.03.022>
<https://www.doi.org/10.1016/j.fluid.2013.03.030>
<https://www.doi.org/10.1016/j.fluid.2011.10.015>
<https://www.doi.org/10.1016/j.jct.2015.11.015>
<https://www.doi.org/10.1016/j.jct.2016.06.018>
<https://www.doi.org/10.1016/j.jct.2010.08.004>
<https://www.doi.org/10.1016/j.jct.2013.11.001>
<https://www.doi.org/10.1016/j.fluid.2010.04.002>
<https://www.doi.org/10.1016/j.jct.2003.09.010>
<https://www.doi.org/10.1021/je7001418>
<https://www.doi.org/10.1016/j.jct.2015.07.038>
<https://www.doi.org/10.1016/j.jct.2016.03.012>
<https://www.doi.org/10.1016/j.jct.2017.08.010>
<http://pubs.acs.org/doi/abs/10.1021/ci990307I>
<https://www.doi.org/10.1021/je400077c>
<https://www.doi.org/10.1021/acs.jced.7b00452>
<https://www.doi.org/10.1021/acs.jced.5b01031>
<https://www.doi.org/10.1016/j.jct.2012.12.010>
<https://www.doi.org/10.1021/acs.jced.6b00766>
https://www.chemeo.com/doc/models/crippen_log10ws
<https://www.doi.org/10.1016/j.jct.2008.09.008>
<https://www.doi.org/10.1016/j.tca.2006.07.009>
<https://www.doi.org/10.1021/je501178z>
<https://www.doi.org/10.1021/je500035r>
<https://www.doi.org/10.1016/j.jct.2016.06.030>
<https://www.doi.org/10.1021/je100909b>
<https://www.doi.org/10.1016/j.jct.2018.09.026>
<https://www.doi.org/10.1021/je900882r>
<https://www.doi.org/10.1016/j.jct.2012.05.009>
<https://www.doi.org/10.1016/j.jct.2013.08.018>
<https://www.doi.org/10.1021/je400894j>

Interaction of some hydrophobic amino acids, peptides, and protein with <i>Bacillus subtilis</i> and <i>Staphylococcus aureus</i> and its effect on the growth of <i>Escherichia coli</i> and <i>Salmonella typhimurium</i> in aqueous solution	https://www.doi.org/10.1016/j.jct.2010.11.015
Densities and Speeds of Sound of Aqueous Solutions of Alanine and Valine in the Temperature Range of 25–40 °C	https://www.doi.org/10.1021/je500324a
Thermodynamic Properties of Amino Acids in Aqueous Solution	https://www.doi.org/10.1016/j.jct.2011.05.012
Thermal Stabilities of Different Amino Acids in Aqueous Solution	https://www.doi.org/10.1016/j.jct.2013.09.009
Temperature Dependence of the Solubility of Amino Acids in Aqueous Solution	https://www.doi.org/10.1016/j.jct.2014.03.015
Thermal Stabilities of Amino Acids in Aqueous Solution	https://www.doi.org/10.1016/j.tca.2013.06.017
Thermal Stabilities of Amino Acids in Aqueous Solution	https://www.doi.org/10.1021/acs.jced.5b00198

<https://www.doi.org/10.1021/je500324a>

<https://www.doi.org/10.1016/j.jct.2011.05.012>

<https://www.doi.org/10.1016/j.jct.2013.09.009>

<https://www.doi.org/10.1016/j.jct.2014.03.015>

<https://www.doi.org/10.1016/j.tca.2013.06.017>

<https://www.doi.org/10.1021/acs.iced.5b00198>

Legend

cpg:	Ideal gas heat capacity
cps:	Solid phase heat capacity
gf:	Standard Gibbs free energy of formation
hf:	Enthalpy of formation at standard conditions
hfus:	Enthalpy of fusion at standard conditions
hvap:	Enthalpy of vaporization at standard conditions
log10ws:	Log10 of Water solubility in mol/l
logp:	Octanol/Water partition coefficient
mcvol:	McGowan's characteristic volume
pc:	Critical Pressure
tb:	Normal Boiling Point Temperature
tc:	Critical Temperature
tf:	Normal melting (fusion) point
vc:	Critical Volume

Latest version available from:

<https://www.chemeo.com/cid/33-710-6/D-Valine.pdf>

Generated by Cheméo on 2025-12-05 23:39:37.876975323 +0000 UTC m=+4726175.407015976.

Cheméo (<https://www.chemeo.com>) is the biggest free database of chemical and physical data for the process industry.