

# Propanamide, N,N-bis(2-ethylhexyl)-2-methyl-

<b>Other names:</b>	N,N-di-2-ethylhexylisobutyramide
<b>Inchi:</b>	InChI=1S/C20H41NO/c1-7-11-13-18(9-3)15-21(20(22)17(5)6)16-19(10-4)14-12-8-2/h17-
<b>InchiKey:</b>	TYDFWJHDNHXBRF-UHFFFAOYSA-N
<b>Formula:</b>	C20H41NO
<b>SMILES:</b>	CCCCC(CC)CN(CC(CC)CCCC)C(=O)C(C)C
<b>Mol. weight [g/mol]:</b>	311.55

## Physical Properties

Property code	Value	Unit	Source
gf	92.06	kJ/mol	Joback Method
hf	-517.02	kJ/mol	Joback Method
hfus	79.08	kJ/mol	Enthalpies of vaporization of N,N-dialkyl monamides at 298.15K
hvap	67.74	kJ/mol	Joback Method
log10ws	-5.81		Crippen Method
logp	5.904		Crippen Method
mcvol	304.210	ml/mol	McGowan Method
pc	1072.87	kPa	Joback Method
rinpol	1949.00		NIST Webbook
rinpol	1949.00		NIST Webbook
tb	721.99	K	Joback Method
tc	895.97	K	Joback Method
tf	352.56	K	Joback Method
vc	1.161	m <sup>3</sup> /kmol	Joback Method

## Temperature Dependent Properties

Property code	Value	Unit	Temperature [K]	Source
cpg	919.74	J/mol×K	721.99	Joback Method
cpg	940.34	J/mol×K	750.99	Joback Method
cpg	959.93	J/mol×K	779.98	Joback Method
cpg	978.56	J/mol×K	808.98	Joback Method
cpg	996.26	J/mol×K	837.97	Joback Method

cpg	1013.07	J/mol×K	866.97	Joback Method
cpg	1029.02	J/mol×K	895.97	Joback Method

## Sources

<b>Crippen Method:</b>	<a href="https://www.chemeo.com/doc/models/crippen_log10ws">https://www.chemeo.com/doc/models/crippen_log10ws</a>
<b>Enthalpies of vaporization of N,N-dialkyl monamides at 298.15K:</b>	<a href="https://www.doi.org/10.1016/j.tca.2009.05.007">https://www.doi.org/10.1016/j.tca.2009.05.007</a>
<b>Determination of Activity Coefficients at Infinite Dilution of Solutes in Non-polar (2-ethylhexyl)isobutyramide</b>	<a href="https://www.doi.org/10.1021/acs.jced.8b00635">https://www.doi.org/10.1021/acs.jced.8b00635</a>
<b>Joback Method</b>	<a href="https://en.wikipedia.org/wiki/Joback_method">https://en.wikipedia.org/wiki/Joback_method</a>
<b>Using Inverse Gas-Liquid Chromatography:</b>	<a href="http://link.springer.com/article/10.1007/BF02311772">http://link.springer.com/article/10.1007/BF02311772</a>
<b>NIST Webbook:</b>	<a href="http://webbook.nist.gov/cgi/cbook.cgi?ID=U308084&amp;Units=SI">http://webbook.nist.gov/cgi/cbook.cgi?ID=U308084&amp;Units=SI</a>
<b>Crippen Method:</b>	<a href="http://pubs.acs.org/doi/abs/10.1021/ci9903071">http://pubs.acs.org/doi/abs/10.1021/ci9903071</a>

## Legend

<b>cpg:</b>	Ideal gas heat capacity
<b>gf:</b>	Standard Gibbs free energy of formation
<b>hf:</b>	Enthalpy of formation at standard conditions
<b>hfus:</b>	Enthalpy of fusion at standard conditions
<b>hvap:</b>	Enthalpy of vaporization at standard conditions
<b>log10ws:</b>	Log10 of Water solubility in mol/l
<b>logp:</b>	Octanol/Water partition coefficient
<b>mcvol:</b>	McGowan's characteristic volume
<b>pc:</b>	Critical Pressure
<b>rinpol:</b>	Non-polar retention indices
<b>tb:</b>	Normal Boiling Point Temperature
<b>tc:</b>	Critical Temperature
<b>tf:</b>	Normal melting (fusion) point
<b>vc:</b>	Critical Volume

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