

# 1-Ethenyl-7-methylnonyl Prop-2-enoate

<b>Inchi:</b>	InChI=1S/C15H26O2/c1-5-13(4)11-9-8-10-12-14(6-2)17-15(16)7-3/h6-7,13-14H,2-3,5,8-
<b>InchiKey:</b>	TWXJMPDHSKXVID-UHFFFAOYSA-N
<b>Formula:</b>	C15H26O2
<b>SMILES:</b>	C=CC(=O)OC(C=C)CCCCC(C)CC
<b>Mol. weight [g/mol]:</b>	238.37

## Physical Properties

Property code	Value	Unit	Source
gf	12.30	kJ/mol	Joback Method
hf	-357.43	kJ/mol	Joback Method
hfus	27.79	kJ/mol	Joback Method
hvap	56.02	kJ/mol	Joback Method
log10ws	-4.54		Crippen Method
logp	4.267		Crippen Method
mcvol	221.050	ml/mol	McGowan Method
pc	1584.75	kPa	Joback Method
rinsol	1552.00		NIST Webbook
tb	611.37	K	Joback Method
tc	789.05	K	Joback Method
tf	297.45	K	Joback Method
vc	0.850	m <sup>3</sup> /kmol	Joback Method

## Temperature Dependent Properties

Property code	Value	Unit	Temperature [K]	Source
cpg	577.55	J/molxK	611.37	Joback Method
cpg	594.60	J/molxK	640.98	Joback Method
cpg	610.86	J/molxK	670.60	Joback Method
cpg	626.36	J/molxK	700.21	Joback Method
cpg	641.10	J/molxK	729.82	Joback Method
cpg	655.12	J/molxK	759.43	Joback Method
cpg	668.44	J/molxK	789.05	Joback Method
dvisc	0.0043488	Paxs	297.45	Joback Method
dvisc	0.0015618	Paxs	349.77	Joback Method

dvisc	0.0007321	Paxs	402.09	Joback Method
dvisc	0.0004086	Paxs	454.41	Joback Method
dvisc	0.0002573	Paxs	506.73	Joback Method
dvisc	0.0001766	Paxs	559.05	Joback Method
dvisc	0.0001293	Paxs	611.37	Joback Method

## Sources

<b>NIST Webbook:</b>	<a href="http://webbook.nist.gov/cgi/cbook.cgi?ID=R412598&amp;Units=SI">http://webbook.nist.gov/cgi/cbook.cgi?ID=R412598&amp;Units=SI</a>
<b>Crippen Method:</b>	<a href="http://pubs.acs.org/doi/abs/10.1021/ci9903071">http://pubs.acs.org/doi/abs/10.1021/ci9903071</a>
<b>Crippen Method:</b>	<a href="https://www.chemeo.com/doc/models/crippen_log10ws">https://www.chemeo.com/doc/models/crippen_log10ws</a>
<b>Joback Method:</b>	<a href="https://en.wikipedia.org/wiki/Joback_method">https://en.wikipedia.org/wiki/Joback_method</a>
<b>McGowan Method:</b>	<a href="http://link.springer.com/article/10.1007/BF02311772">http://link.springer.com/article/10.1007/BF02311772</a>

## Legend

<b>cpg:</b>	Ideal gas heat capacity
<b>dvisc:</b>	Dynamic viscosity
<b>gf:</b>	Standard Gibbs free energy of formation
<b>hf:</b>	Enthalpy of formation at standard conditions
<b>hfus:</b>	Enthalpy of fusion at standard conditions
<b>hvap:</b>	Enthalpy of vaporization at standard conditions
<b>log10ws:</b>	Log10 of Water solubility in mol/l
<b>logp:</b>	Octanol/Water partition coefficient
<b>mcvol:</b>	McGowan's characteristic volume
<b>pc:</b>	Critical Pressure
<b>rinpol:</b>	Non-polar retention indices
<b>tb:</b>	Normal Boiling Point Temperature
<b>tc:</b>	Critical Temperature
<b>tf:</b>	Normal melting (fusion) point
<b>vc:</b>	Critical Volume

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