

sodium bromide

Inchi:	InChI=1S/BrH.Na/h1H;/q;+1/p-1
InchiKey:	JHJLBTNAGRQEKS-UHFFFAOYSA-M
Formula:	BrNa
SMILES:	[Br-].[Na+]
Mol. weight [g/mol]:	102.89
CAS:	7647-15-6

Physical Properties

Property code	Value	Unit	Source
ea	0.79 ± 0.01	eV	NIST Webbook
ea	0.94	eV	NIST Webbook
ie	8.30 ± 0.10	eV	NIST Webbook
ie	8.30 ± 0.10	eV	NIST Webbook
ie	8.30 ± 0.10	eV	NIST Webbook
ie	8.70	eV	NIST Webbook
tt	1114.00	K	Study of the NaBr DyBr ₃ phase diagram by differential thermal analysis

Correlations

Information	Value
Property code	pvap
Equation	$\ln(P_{vp}) = A + B/(T + C)$
Coeff. A	1.53489e+01
Coeff. B	-1.64472e+04
Coeff. C	-1.31070e+02
Temperature range (K), min.	1020.00
Temperature range (K), max.	1720.00

Solubility of Sodium Salts in Ammonium-Based Deep Eutectic Solvents: molar volumes, expansibilities, and isentropic compressibilities of Selected Alkali Bromides in Aqueous Binary Mixtures of γ -Butyrolactone and Butyrene Methyl Ether and Ternary Mixtures of KBr Solutions at 298.15, 303.15, 308.15, 313.15, 318.15, and 323.15 K: Hexanol
 Experimental Study of the Solubility of Selected Alkali Bromide Salts in the Systems $\text{Na}_2\text{B}_4\text{O}_7$, NaBr and NaCl in Methanol, Ethanol, and Their Mixtures: Studies of Tetrabutylammonium Bromide, Sodium Bromide, and Sodium Tetrakisphenylborate in 2-Butoxyethanol (1) + Water (2) Mixtures at (298.15, 303.15, 308.15, and 313.15) K:

<https://www.doi.org/10.1021/je400045d>
<https://www.doi.org/10.1016/j.jct.2008.04.008>
<https://www.doi.org/10.1021/je900709n>
<https://www.doi.org/10.1021/je049814b>
<https://www.doi.org/10.1021/je1007394>
<https://www.doi.org/10.1016/j.fluid.2013.10.047>
<https://www.doi.org/10.1021/je5009944>
<https://www.doi.org/10.1016/j.jct.2010.07.003>
<https://www.doi.org/10.1016/j.jct.2013.08.018>
<https://www.doi.org/10.1021/je049922y>
<https://www.doi.org/10.1021/je300737t>

Legend

ea: Electron affinity
 ie: Ionization energy
 pvap: Vapor pressure
 tt: Triple Point Temperature

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