

(3-Nitrophenyl) methanol, isopropyl ether

Inchi:	InChI=1S/C10H13NO3/c1-8(2)14-7-9-4-3-5-10(6-9)11(12)13/h3-6,8H,7H2,1-2H3
InchiKey:	JYLYIZFXTADDBT-UHFFFAOYSA-N
Formula:	C10H13NO3
SMILES:	CC(C)OCc1cccc([N+](=O)[O-])c1
Mol. weight [g/mol]:	195.22

Physical Properties

Property code	Value	Unit	Source
gf	64.21	kJ/mol	Joback Method
hf	-172.93	kJ/mol	Joback Method
hfus	24.33	kJ/mol	Joback Method
hvap	59.41	kJ/mol	Joback Method
log10ws	-3.46		Crippen Method
logp	2.520		Crippen Method
mvol	151.290	ml/mol	McGowan Method
pc	2893.62	kPa	Joback Method
rinpol	1555.00		NIST Webbook
rinpol	1555.00		NIST Webbook
tb	633.68	K	Joback Method
tc	869.80	K	Joback Method
tf	392.24	K	Joback Method
vc	0.582	m ³ /kmol	Joback Method

Temperature Dependent Properties

Property code	Value	Unit	Temperature [K]	Source
cpg	382.74	J/mol×K	633.68	Joback Method
cpg	396.77	J/mol×K	673.03	Joback Method
cpg	409.84	J/mol×K	712.39	Joback Method
cpg	421.98	J/mol×K	751.74	Joback Method
cpg	433.23	J/mol×K	791.09	Joback Method
cpg	443.60	J/mol×K	830.44	Joback Method
cpg	453.12	J/mol×K	869.80	Joback Method

Sources

McGowan Method:	http://link.springer.com/article/10.1007/BF02311772
NIST Webbook:	http://webbook.nist.gov/cgi/cbook.cgi?ID=U374649&Units=SI
Crippen Method:	http://pubs.acs.org/doi/abs/10.1021/ci990307l
Crippen Method:	https://www.cheméo.com/doc/models/crippen_log10ws
Joback Method:	https://en.wikipedia.org/wiki/Joback_method

Legend

cpg:	Ideal gas heat capacity
gf:	Standard Gibbs free energy of formation
hf:	Enthalpy of formation at standard conditions
hfus:	Enthalpy of fusion at standard conditions
hvap:	Enthalpy of vaporization at standard conditions
log10ws:	Log10 of Water solubility in mol/l
logp:	Octanol/Water partition coefficient
mcvol:	McGowan's characteristic volume
pc:	Critical Pressure
rinpola:	Non-polar retention indices
tb:	Normal Boiling Point Temperature
tc:	Critical Temperature
tf:	Normal melting (fusion) point
vc:	Critical Volume

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