

# Nonyl 2-bromobutanoate

<b>Inchi:</b>	InChI=1S/C13H25BrO2/c1-3-5-6-7-8-9-10-11-16-13(15)12(14)4-2/h12H,3-11H2,1-2H3
<b>InchiKey:</b>	QLHKAJYBOHBZQL-UHFFFAOYSA-N
<b>Formula:</b>	C13H25BrO2
<b>SMILES:</b>	CCCCCCCCCOC(=O)C(Br)CC
<b>Mol. weight [g/mol]:</b>	293.24

## Physical Properties

Property code	Value	Unit	Source
gf	-163.46	kJ/mol	Joback Method
hf	-535.40	kJ/mol	Joback Method
hfus	33.98	kJ/mol	Joback Method
hvap	59.73	kJ/mol	Joback Method
log10ws	-4.67		Crippen Method
logp	4.454		Crippen Method
mvol	218.970	ml/mol	McGowan Method
pc	1821.61	kPa	Joback Method
rinpol	1700.00		NIST Webbook
rinpol	1700.00		NIST Webbook
tb	638.85	K	Joback Method
tc	822.28	K	Joback Method
tf	353.23	K	Joback Method
vc	0.844	m <sup>3</sup> /kmol	Joback Method

## Temperature Dependent Properties

Property code	Value	Unit	Temperature [K]	Source
cpg	564.89	J/mol×K	638.85	Joback Method
cpg	580.63	J/mol×K	669.42	Joback Method
cpg	595.63	J/mol×K	699.99	Joback Method
cpg	609.91	J/mol×K	730.56	Joback Method
cpg	623.47	J/mol×K	761.14	Joback Method
cpg	636.33	J/mol×K	791.71	Joback Method
cpg	648.53	J/mol×K	822.28	Joback Method
dvisc	0.0023891	Paxs	353.23	Joback Method

dvisc	0.0011305	Paxs	400.83	Joback Method
dvisc	0.0006271	Paxs	448.44	Joback Method
dvisc	0.0003895	Paxs	496.04	Joback Method
dvisc	0.0002630	Paxs	543.64	Joback Method
dvisc	0.0001891	Paxs	591.25	Joback Method
dvisc	0.0001429	Paxs	638.85	Joback Method

## Sources

<b>NIST Webbook:</b>	<a href="http://webbook.nist.gov/cgi/cbook.cgi?ID=R23483&amp;Units=SI">http://webbook.nist.gov/cgi/cbook.cgi?ID=R23483&amp;Units=SI</a>
<b>Crippen Method:</b>	<a href="http://pubs.acs.org/doi/abs/10.1021/ci9903071">http://pubs.acs.org/doi/abs/10.1021/ci9903071</a>
<b>Crippen Method:</b>	<a href="https://www.chemeo.com/doc/models/crippen_log10ws">https://www.chemeo.com/doc/models/crippen_log10ws</a>
<b>Joback Method:</b>	<a href="https://en.wikipedia.org/wiki/Joback_method">https://en.wikipedia.org/wiki/Joback_method</a>
<b>McGowan Method:</b>	<a href="http://link.springer.com/article/10.1007/BF02311772">http://link.springer.com/article/10.1007/BF02311772</a>

## Legend

<b>cp<sub>g</sub>:</b>	Ideal gas heat capacity
<b>dvisc:</b>	Dynamic viscosity
<b>gf:</b>	Standard Gibbs free energy of formation
<b>hf:</b>	Enthalpy of formation at standard conditions
<b>hfus:</b>	Enthalpy of fusion at standard conditions
<b>hvap:</b>	Enthalpy of vaporization at standard conditions
<b>log<sub>10</sub>ws:</b>	Log <sub>10</sub> of Water solubility in mol/l
<b>logp:</b>	Octanol/Water partition coefficient
<b>m<sub>cvol</sub>:</b>	McGowan's characteristic volume
<b>pc:</b>	Critical Pressure
<b>rinpol:</b>	Non-polar retention indices
<b>tb:</b>	Normal Boiling Point Temperature
<b>tc:</b>	Critical Temperature
<b>tf:</b>	Normal melting (fusion) point
<b>vc:</b>	Critical Volume

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