

# Acetic acid, trifluoro-, 2,2,2-trifluoroethyl ester

<b>Other names:</b>	2,2,2-Trifluoroethyl trifluoroacetate CF <sub>3</sub> C(O)OCH <sub>2</sub> CF <sub>3</sub> Trifluoroacetic acid trifluoroethyl ester
<b>Inchi:</b>	InChI=1S/C4H2F6O2/c5-3(6,7)1-12-2(11)4(8,9)10/h1H2
<b>InchiKey:</b>	ZKUJOCJJXCPCFS-UHFFFAOYSA-N
<b>Formula:</b>	C <sub>4</sub> H <sub>2</sub> F <sub>6</sub> O <sub>2</sub>
<b>SMILES:</b>	O=C(OCC(F)(F)F)C(F)(F)F
<b>Mol. weight [g/mol]:</b>	196.05
<b>CAS:</b>	407-38-5

## Physical Properties

Property code	Value	Unit	Source
gf	-1414.30	kJ/mol	Joback Method
hf	-1564.85	kJ/mol	Joback Method
hfus	12.55	kJ/mol	Joback Method
hvap	26.16	kJ/mol	Joback Method
log10ws	-1.68		Crippen Method
logp	1.654		Crippen Method
mcvol	85.280	ml/mol	McGowan Method
pc	3127.99	kPa	Joback Method
tb	328.00	K	NIST Webbook
tb	328.20	K	NIST Webbook
tc	502.85	K	Joback Method
tf	215.38	K	Joback Method
vc	0.369	m <sup>3</sup> /kmol	Joback Method

## Temperature Dependent Properties

Property code	Value	Unit	Temperature [K]	Source
cpg	181.39	J/mol×K	356.37	Joback Method
cpg	188.79	J/mol×K	380.78	Joback Method
cpg	195.79	J/mol×K	405.20	Joback Method
cpg	202.40	J/mol×K	429.61	Joback Method
cpg	208.64	J/mol×K	454.03	Joback Method

cpg	214.53	J/mol×K	478.44	Joback Method
cpg	220.06	J/mol×K	502.85	Joback Method
hvapt	31.80	kJ/mol	330.00	NIST Webbook

## Correlations

Information	Value
Property code	pvap
Equation	$\ln(P_{vp}) = A + B/(T + C)$
Coeff. A	1.62909e+01
Coeff. B	-3.82920e+03
Temperature range (K), min.	239.27
Temperature range (K), max.	348.76

## Sources

<b>NIST Webbook:</b>	<a href="http://webbook.nist.gov/cgi/cbook.cgi?ID=C407385&amp;Units=SI">http://webbook.nist.gov/cgi/cbook.cgi?ID=C407385&amp;Units=SI</a>
<b>The Yaws Handbook of Vapor Pressure:</b>	<a href="https://www.sciencedirect.com/book/9780128029992/the-yaws-handbook-of-vapor-pressure">https://www.sciencedirect.com/book/9780128029992/the-yaws-handbook-of-vapor-pressure</a>
<b>Crippen Method:</b>	<a href="http://pubs.acs.org/doi/abs/10.1021/ci9903071">http://pubs.acs.org/doi/abs/10.1021/ci9903071</a>
<b>Crippen Method:</b>	<a href="https://www.chemeo.com/doc/models/crippen_log10ws">https://www.chemeo.com/doc/models/crippen_log10ws</a>
<b>Joback Method:</b>	<a href="https://en.wikipedia.org/wiki/Joback_method">https://en.wikipedia.org/wiki/Joback_method</a>
<b>KDB:</b>	<a href="https://www.thermo.com/files/research/kdb/mol/mol1785.mol">https://www.thermo.com/files/research/kdb/mol/mol1785.mol</a>
<b>McGowan Method:</b>	<a href="http://link.springer.com/article/10.1007/BF02311772">http://link.springer.com/article/10.1007/BF02311772</a>

## Legend

<b>cpg:</b>	Ideal gas heat capacity
<b>gf:</b>	Standard Gibbs free energy of formation
<b>hf:</b>	Enthalpy of formation at standard conditions
<b>hfus:</b>	Enthalpy of fusion at standard conditions
<b>hvap:</b>	Enthalpy of vaporization at standard conditions
<b>hvapt:</b>	Enthalpy of vaporization at a given temperature
<b>log10ws:</b>	Log10 of Water solubility in mol/l
<b>logp:</b>	Octanol/Water partition coefficient
<b>mcvol:</b>	McGowan's characteristic volume

<b>pc:</b>	Critical Pressure
<b>pvap:</b>	Vapor pressure
<b>tb:</b>	Normal Boiling Point Temperature
<b>tc:</b>	Critical Temperature
<b>tf:</b>	Normal melting (fusion) point
<b>vc:</b>	Critical Volume

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