1-Butanol

Other names:	1-Butyl alcohol
	1-Hydroxybutane
	Alcool butylique
	Butan-1-ol
	Butanol
	Butanol-1
	Butanolen
	Butanolo
	Butyl alcohol
	Butyl hydroxide
	Butylowy alkohol
	Butyric alcohol
	CCS 203
	Hemostyp
	METHYLOLPROPANE
	N-BUTYL ALCOHOL
	NSC 62782
	Propylcarbinol
	Propylmethanol
	Rcra waste number U031
	n-Butan-1-ol
	n-Butanol
	n-C4H9OH
Inchi:	InChI=1S/C4H10O/c1-2-3-4-5/h5H,2-4H2,1H3
InchiKey:	LRHPLDYGYMQRHN-UHFFFAOYSA-N
Formula:	C4H10O
SMILES:	CCCCO
Mol. weight [g/mol]:	74.12
CAS:	71-36-3

Physical Properties

Property code	Value	Unit	Source
af	0.5930		KDB
affp	789.20	kJ/mol	NIST Webbook
aigt	616.48	K	KDB
basg	758.90	kJ/mol	NIST Webbook

chl	-2674.90 ± 0.84	kJ/mol	NIST Webbook
chl	-2677.40 ± 0.63	kJ/mol	NIST Webbook
chl	-2676.01 ± 0.24	kJ/mol	NIST Webbook
chl	-2675.61 ± 0.45	kJ/mol	NIST Webbook
chl	-2670.44	kJ/mol	NIST Webbook
chl	-2676.00	kJ/mol	NIST Webbook
chl	-2640.00 ± 120.00	kJ/mol	NIST Webbook
chl	-2676.18 ± 0.24	kJ/mol	NIST Webbook
chl	-2670.77	kJ/mol	NIST Webbook
chl	-2670.40 ± 0.42	kJ/mol	NIST Webbook
dm	1.80	debye	KDB
dvisc	0.0025383	Pa×s	Densities and Viscosities of Binary Liquid Mixtures of Trichloroethylene and Tetrachloroethylene with Some Polar and Nonpolar Solvents
dvisc	0.0025650	Pa×s	Thermodynamic properties of binary liquid mixtures of diethylenetriamine with alcohols at different temperatures
dvisc	0.0025730	Pa×s	Study of densities, viscosities, and speeds of sound of binary liquid mixtures of butan-1-ol with n-alkanes (C6, C8, and C10) at T = (298.15, 303.15, and 308.15) K
fll	1.40	% in Air	KDB
flu	11.20	% in Air	KDB
fpc	309.26	K	KDB
fpo	302.04	K	KDB
gf	-150.90	kJ/mol	KDB
gyrad	3.2250		KDB
hf	-275.28 ± 0.53	kJ/mol	NIST Webbook
hf	-280.00 ± 0.92	kJ/mol	NIST Webbook
hf	-273.50 ± 0.70	kJ/mol	NIST Webbook
hf	-277.00 ± 1.50	kJ/mol	NIST Webbook
hf	-281.40	kJ/mol	NIST Webbook
hf	-278.20	kJ/mol	NIST Webbook
hf	-274.90	kJ/mol	KDB
hf	-274.40	kJ/mol	NIST Webbook
hf	-275.60	kJ/mol	NIST Webbook
hf	-277.30 ± 3.00	kJ/mol	NIST Webbook
hf	-280.50 ± 0.40	kJ/mol	NIST Webbook
hf	-274.90 ± 0.30	kJ/mol	NIST Webbook
hf	-275.80	kJ/mol	NIST Webbook
hf	-274.70 ± 0.30	kJ/mol	NIST Webbook

hfl	-329.60	kJ/mol	NIST Webbook
hfl	-327.18 ± 0.24	kJ/mol	NIST Webbook
hfl	-325.80 ± 0.67	kJ/mol	NIST Webbook
hfl	-328.40 ± 0.84	kJ/mol	NIST Webbook
hfl	-332.80 ± 0.42	kJ/mol	NIST Webbook
hfl	-327.58 ± 0.52	kJ/mol	NIST Webbook
hfl	-327.01 ± 0.24	kJ/mol	NIST Webbook
hfus	10.20	kJ/mol	Joback Method
hvap	41.18	kJ/mol	Joback Method
ie	10.64 ± 0.07	eV	NIST Webbook
ie	10.10 ± 0.05	eV	NIST Webbook
ie	9.99 ± 0.05	eV	NIST Webbook
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ie	10.04	eV	NIST Webbook
ie	10.37	eV	NIST Webbook
ie	10.37	eV	NIST Webbook
ie	10.09 ± 0.02	eV	NIST Webbook
log10ws	0.00		Aqueous Solubility Prediction Method
log10ws	0.00		Estimated Solubility Method
logp	0.779		Crippen Method
mcvol	73.090	ml/mol	McGowan Method
nfpaf	%!d(float64=3)		KDB
nfpah	%!d(float64=1)		KDB
pc	4413.00	kPa	NIST Webbook
pc	4423.00	kPa	NIST Webbook
pc	4415.00	kPa	NIST Webbook
pc	4418.00 ± 20.00	kPa	NIST Webbook
pc	4410.00	kPa	NIST Webbook
pc	4410.00 ± 20.00	kPa	NIST Webbook
pc	4414.00	kPa	KDB
pc	4924.00 ± 30.00	kPa	NIST Webbook
pc	4418.00 ± 20.00	kPa	NIST Webbook
pc	4904.13 ± 60.79	kPa	NIST Webbook
pc	4413.00	kPa	NIST Webbook
rhoc	271.29 ± 1.48	kg/m3	NIST Webbook
rhoc	266.84	kg/m3	NIST Webbook
rhoc	269.88	kg/m3	NIST Webbook
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rhoc	269.80	kg/m3	NIST Webbook

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ripol	1125.00	NIST Webbook
ripol	1157.00	NIST Webbook
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ripol	1113.00	NIST Webbook
ripol	1160.00	NIST Webbook
ripol	1175.00	NIST Webbook
ripol	1164.00	NIST Webbook
ripol	1150.00	NIST Webbook
ripol	1150.00	NIST Webbook
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ripol	1160.00	NIST Webbook
ripol	1135.00	NIST Webbook
ripol	1143.00	NIST Webbook
ripol	1152.00	NIST Webbook
ripol	1138.00	NIST Webbook
ripol	1136.00	NIST Webbook
ripol	1115.00	NIST Webbook
ripol	1137.00	NIST Webbook
ripol	1152.00	NIST Webbook
ripol	1118.00	NIST Webbook
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ripol	1178.00	NIST Webbook
ripol	1135.00	NIST Webbook
ripol	1135.08	NIST Webbook
ripol	1129.04	NIST Webbook
ripol	1130.99	NIST Webbook
ripol	1133.33	NIST Webbook
ripol	1131.00	NIST Webbook
ripol	1153.00	NIST Webbook
ripol	1124.00	NIST Webbook
ripol	1148.00	NIST Webbook
ripol	1104.00	NIST Webbook
ripol	1126.00	NIST Webbook
ripol	1178.00	NIST Webbook
ripol	1146.00	NIST Webbook
ripol	1147.50	NIST Webbook
ripol	1120.00	NIST Webbook

ripol	1128.00		NIST Webbook
ripol	1133.00		NIST Webbook
sg	361.98	J/mol×K	NIST Webbook
sl	228.00	J/mol×K	NIST Webbook
sl	225.73	J/mol×K	NIST Webbook
sl	251.90	J/mol×K	NIST Webbook
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tb	390.65 ± 0.50	K	NIST Webbook
tb	390.17 ± 0.20	K	NIST Webbook
tb	389.15 ± 2.00	K	NIST Webbook
tb	389.65 ± 0.50	K	NIST Webbook
tb	389.65 ± 2.50	K	NIST Webbook
tb	390.65 ± 0.30	K	NIST Webbook
tb	390.90 ± 0.50	K	NIST Webbook
tb	389.65 ± 1.00	K	NIST Webbook
tb	390.25 ± 0.20	K	NIST Webbook
tb	390.75 ± 0.50	K	NIST Webbook
tb	390.86 ± 0.05	K	NIST Webbook
tb	389.55 ± 0.70	K	NIST Webbook
tb	390.75 ± 0.30	K	NIST Webbook
tb	390.65 ± 0.30	K	NIST Webbook
tb	390.25 ± 0.20	K	NIST Webbook
tb	390.80 ± 0.50	K	NIST Webbook
tb	390.85 ± 0.50	K	NIST Webbook
tb	390.65 ± 0.50	K	NIST Webbook
tb	390.00 ± 1.50	K	NIST Webbook
tb	390.60 ± 0.50	K	NIST Webbook
tb	390.48	K	Phase Equilibrium (VLE, LLE, and VLLE) Data of the Ternary System: Ionic Liquid [OMIM][PF6] + Butan-1-ol + Butyl Acetate
tb	391.00 ± 0.30	K	NIST Webbook
tb	390.70 ± 0.70	K	NIST Webbook
tb	389.75 ± 0.50	K	NIST Webbook
tb	390.95 ± 0.20	K	NIST Webbook
tb	390.90 ± 0.20	K	NIST Webbook
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tb	390.88	K	KDB

tb	390.83	K	Experimental Determination of Densities and Isobaric Vapor Liquid Equilibria of Methyl Acetate and Ethyl Acetate with Alcohols (C3 and C4) at 0.3 MPa
tb	390.75	K	Excess molar volumes of ternary mixtures of 1,3-dichlorobenzene and methyl ethyl ketone with 1-alkanols at 303.15K
tb	390.12	K	Vapor liquid equilibria for the binary mixtures of 2,3-butanediol with n-butanol, n-butyl acetate, and ethyl acetate at 101.3 kPa
tb	390.85	K	Measurement and correlation of binary vapor liquid equilibria of isomeric butanols with 1,4-dioxane
tb	390.12	K	Liquid liquid equilibria of water + 2,3-butanediol + 1-butanol at T = 298.15 K, T = 308.15K and T = 318.15K
tb	390.65	K	Isobaric VLE data for the system of butan-1-ol + butylethanoate + 1-butyl-3-methylimidazoliumbis[(trifluoromethyl)sulf
tb	390.82	K	Isobaric vapor-liquid equilibrium for the binary systems of 1-butanol +Poly methoxy methylal (n=2), and 1-butanol +Poly methoxy methylal (n=3) at 101.3 kPa
tb	390.80	K	Solubility and tie-line data for ternary aqueous mixtures of cyclopentanol with organic solvents at T = 298.2 K: Experiments and NRTL model
tb	390.84	K	Isobaric vapor liquid equilibrium for the binary systems of 1-propanol + 1-(methoxymethoxy)-butane and 1-butanol + 1-(methoxymethoxy)-butane at 101.3 kPa
tb	390.40	K	Isothermal and isobaric (vapour + liquid) equilibria of (alpha-pinene + n-butanol + n-octane)
tb	390.82	K	Isobaric Vapor Liquid Equilibrium for Two Binary Systems (n-Butanol + 1,4-Butanediol and .gamma.-Butyrolactone + 1,4-Butanediol) at p = (30.0, 50.0, and 70.0) kPa

tb	390.79	K	Isobaric Vapor Liquid Equilibrium for the Binary Systems of sec-Butyl Acetate + n-Butyl Alcohol, Isobutyl Alcohol, or tert-Butyl Alcohol at 101.3 kPa
tb	390.78	K	Isobaric Vapor Liquid Equilibrium for Three Binary Systems of Acetaldehyde + Ethanol, Ethyl Acetate, 1-Butanol at 101.3 kPa
tb	390.83	K	Measurement and Modelization of VLE for Butyl Acetate with Methanol, Ethanol, 1-Propanol, and 1-Butanol. Experimental Data at 0.15 MPa
tb	390.75	K	Vapor-Liquid Equilibrium for Binary of 1-Butanol + N,N-Dimethylacetamide and Methyl Isobutyl Ketone + N,N-Dimethylacetamide at 101.3 kPa
tb	390.91	K	Isobaric Phase Equilibria of Diethyl Carbonate with Five Alcohols at 101.3 kPa
tb	390.86	K	Measurement and Correlation of Excess Molar Enthalpies for Ethylene Glycol + Alkanol Systems at the Temperatures (298.15, 308.15, and 323.15) K
tb	390.65	K	Vapor Liquid Equilibrium at p/kPa = 101.3 of the Binary Mixtures of Ethenyl Acetate with Methanol and Butan-1-ol
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tb	390.85	K	Vapor-Liquid Equilibrium Behaviors of Vanillin in 1-Butanol, 2-Butanol, and 2-Methyl-1-propanol
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tb	390.77 ± 0.07	K	NIST Webbook
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tb	390.80 ± 0.25	K	NIST Webbook
tb	390.75 ± 0.40	K	NIST Webbook
tb	390.40	K	NIST Webbook
tb	391.00	K	NIST Webbook
tb	390.85 ± 0.25	K	NIST Webbook
tb	390.92 ± 0.10	K	NIST Webbook
tb	390.75 ± 0.30	K	NIST Webbook
tb	390.90	K	NIST Webbook
tb	390.85 ± 0.40	K	NIST Webbook
tb	390.85 ± 0.30	K	NIST Webbook
tb	450.74 ± 0.10	K	NIST Webbook
tb	390.15 ± 0.40	K	NIST Webbook
tb	389.85 ± 0.40	K	NIST Webbook
tb	390.80 ± 0.40	K	NIST Webbook
tb	390.10 ± 0.50	K	NIST Webbook
tb	390.65 ± 0.30	K	NIST Webbook
tb	390.55 ± 0.08	K	NIST Webbook
tb	390.80 ± 0.40	K	NIST Webbook
tb	390.80 ± 0.30	K	NIST Webbook
tb	390.80 ± 0.30	K	NIST Webbook
tb	391.20 ± 0.30	K	NIST Webbook
tb	390.51 ± 0.20	K	NIST Webbook
tb	390.90 ± 0.30	K	NIST Webbook
tb	390.75 ± 0.30	K	NIST Webbook
tb	391.00 ± 0.40	K	NIST Webbook
tb	390.35 ± 0.30	K	NIST Webbook
tb	390.90 ± 0.30	K	NIST Webbook
tb	390.85	K	NIST Webbook
tb	390.75 ± 0.30	K	NIST Webbook
tb	391.00 ± 0.15	K	NIST Webbook
tb	390.80 ± 0.50	K	NIST Webbook
tb	390.20 ± 2.00	K	NIST Webbook
tb	390.87 ± 0.50	K	NIST Webbook
tb	390.85 ± 0.30	K	NIST Webbook
tb	390.85 ± 0.40	K	NIST Webbook

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tb	390.85 ± 0.40	K	NIST Webbook
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tb	390.85	K	NIST Webbook
tb	390.15 ± 0.50	K	NIST Webbook
tb	390.85 ± 0.50	K	NIST Webbook
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tb	380.15 ± 2.00	K	NIST Webbook
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tb	390.55 ± 0.05	K	NIST Webbook
tb	390.85 ± 0.30	K	NIST Webbook
tb	391.25 ± 0.50	K	NIST Webbook
tb	390.95 ± 0.30	K	NIST Webbook
tb	391.00 ± 0.40	K	NIST Webbook
tb	390.95 ± 0.50	K	NIST Webbook
tb	390.80 ± 0.50	K	NIST Webbook
tb	390.00 ± 1.00	K	NIST Webbook

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tb	390.75 ± 0.30	K	NIST Webbook
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tb	390.65 ± 1.00	K	NIST Webbook
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tb	390.15 ± 1.00	K	NIST Webbook
tb	390.85 ± 0.50	K	NIST Webbook
tb	390.75 ± 0.50	K	NIST Webbook
tb	390.85 ± 0.50	K	NIST Webbook
tb	388.90 ± 2.00	K	NIST Webbook
tb	390.15 ± 1.00	K	NIST Webbook
tb	390.87 ± 0.05	K	NIST Webbook
tb	390.85 ± 0.50	K	NIST Webbook
tb	391.15 ± 0.50	K	NIST Webbook
tb	390.00 ± 3.00	K	NIST Webbook
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tb	390.84 ± 0.10	K	NIST Webbook
tb	390.50 ± 0.50	K	NIST Webbook
tb	390.55 ± 0.50	K	NIST Webbook
tb	390.57 ± 0.20	K	NIST Webbook
tb	390.90 ± 0.30	K	NIST Webbook
tb	390.90 ± 0.25	K	NIST Webbook
tb	390.30 ± 0.30	K	NIST Webbook
tb	390.45 ± 0.50	K	NIST Webbook
tb	390.40 ± 0.50	K	NIST Webbook
tb	390.90 ± 0.40	K	NIST Webbook
tb	391.15 ± 0.20	K	NIST Webbook
tb	389.30 ± 1.00	K	NIST Webbook
tb	355.70 ± 0.50	K	NIST Webbook
tc	563.00	K	KDB
tc	560.30	K	NIST Webbook
tc	561.20	K	NIST Webbook
tc	560.10 ± 0.20	K	NIST Webbook
tc	562.92	K	NIST Webbook
tc	560.20	K	NIST Webbook
tc	564.20	K	NIST Webbook
tc	560.20	K	NIST Webbook
tc	563.05	K	NIST Webbook
tc	561.00 ± 2.00	K	NIST Webbook

tc	563.01	K	NIST Webbook
tc	562.20 ± 0.40	K	NIST Webbook
tc	562.20 ± 0.30	K	NIST Webbook
tc	562.20 ± 0.30	K	NIST Webbook
tc	562.20 ± 0.20	K	NIST Webbook
tc	562.90	K	NIST Webbook
tc	562.40	K	NIST Webbook
tc	562.40 ± 0.60	K	NIST Webbook
tc	562.40 ± 0.60	K	NIST Webbook
tc	563.60	K	NIST Webbook
tc	563.00 ± 0.30	K	NIST Webbook
tc	563.00	K	Measurement of Critical Properties for Binary and Ternary Mixtures Containing n-Butanol and n-Alkane
tc	562.20 ± 0.40	K	NIST Webbook
tf	193.15 ± 5.00	K	NIST Webbook
tf	183.30	K	KDB
tf	183.63	K	Aqueous Solubility Prediction Method
tf	183.85	K	NIST Webbook
tf	184.00 ± 2.00	K	NIST Webbook
tf	193.65 ± 1.00	K	NIST Webbook
tf	183.35 ± 0.20	K	NIST Webbook
tf	183.85	K	NIST Webbook
tf	192.55 ± 0.50	K	NIST Webbook
tf	183.70	K	Phase equilibria of didecyldimethylammonium nitrate ionic liquid with water and organic solvents
tt	184.51	K	KDB
tt	184.54 ± 0.02	K	NIST Webbook
tt	184.51 ± 0.05	K	NIST Webbook
tt	183.90 ± 0.20	K	NIST Webbook
vc	0.274	m3/kmol	KDB
vc	0.274	m3/kmol	NIST Webbook
zc	0.2583670		KDB
zra	0.25		KDB

Temperature Dependent Properties

Property code	Value	Unit	Temperature [K]	Source
cpg	149.26 ± 0.79	J/mol×K	431.05	NIST Webbook

cpg	151.60 ± 0.79	J/mol×K	441.15	NIST Webbook
cpg	140.93 ± 0.79	J/mol×K	395.25	NIST Webbook
cpg	137.88	J/mol×K	398.15	NIST Webbook
cpg	143.00 ± 0.79	J/mol×K	404.15	NIST Webbook
cpg	144.16 ± 0.79	J/mol×K	409.15	NIST Webbook
cpg	142.06	J/mol×K	413.15	NIST Webbook
cpg	189.31 ± 0.79	J/mol×K	603.35	NIST Webbook
cpg	181.20 ± 0.79	J/mol×K	568.45	NIST Webbook
cpg	175.97 ± 0.79	J/mol×K	545.95	NIST Webbook
cpg	169.95 ± 0.79	J/mol×K	520.05	NIST Webbook
cpg	162.55 ± 0.79	J/mol×K	488.25	NIST Webbook
cpg	155.88 ± 0.79	J/mol×K	459.55	NIST Webbook
cpg	152.66	J/mol×K	453.15	NIST Webbook
cpg	146.58 ± 0.79	J/mol×K	419.55	NIST Webbook
cpg	147.42	J/mol×K	433.15	NIST Webbook
cpl	175.97	J/mol×K	298.15	NIST Webbook
cpl	173.80	J/mol×K	293.15	Molar heat capacities for (1-butanol + 1,4-butanediol, 2,3-butanediol, 1,2-butanediol, and 2-methyl-2,4-pentanediol) as function of temperature
cpl	189.10	J/mol×K	323.00	NIST Webbook
cpl	215.50	J/mol×K	302.60	NIST Webbook
cpl	183.30	J/mol×K	298.00	NIST Webbook
cpl	175.30	J/mol×K	294.00	NIST Webbook
cpl	180.30	J/mol×K	303.00	NIST Webbook
cpl	177.03	J/mol×K	298.15	NIST Webbook
cpl	179.50	J/mol×K	301.20	NIST Webbook
cpl	181.60	J/mol×K	303.50	NIST Webbook
cpl	174.30	J/mol×K	293.15	NIST Webbook
cpl	177.08	J/mol×K	298.15	NIST Webbook
cpl	176.69	J/mol×K	298.15	NIST Webbook
cpl	174.50	J/mol×K	298.00	NIST Webbook
cpl	177.18	J/mol×K	298.15	NIST Webbook
cpl	192.20	J/mol×K	321.05	NIST Webbook
cpl	177.70	J/mol×K	298.00	NIST Webbook
cpl	176.67	J/mol×K	298.15	NIST Webbook
cpl	176.86	J/mol×K	298.15	NIST Webbook
cpl	208.40	J/mol×K	338.15	Isobaric heat capacity of nanostructured liquids with potential use as lubricants

cpl	204.10	J/mol×K	333.15	Isobaric heat capacity of nanostructured liquids with potential use as lubricants
cpl	200.10	J/mol×K	328.15	Isobaric heat capacity of nanostructured liquids with potential use as lubricants
cpl	196.00	J/mol×K	323.15	Isobaric heat capacity of nanostructured liquids with potential use as lubricants
cpl	191.40	J/mol×K	318.15	Isobaric heat capacity of nanostructured liquids with potential use as lubricants
cpl	187.80	J/mol×K	313.15	Isobaric heat capacity of nanostructured liquids with potential use as lubricants
cpl	184.50	J/mol×K	308.15	Isobaric heat capacity of nanostructured liquids with potential use as lubricants
cpl	181.10	J/mol×K	303.15	Isobaric heat capacity of nanostructured liquids with potential use as lubricants
cpl	177.90	J/mol×K	298.15	Isobaric heat capacity of nanostructured liquids with potential use as lubricants
cpl	174.10	J/mol×K	293.15	Isobaric heat capacity of nanostructured liquids with potential use as lubricants
cpl	171.00	J/mol×K	288.15	Isobaric heat capacity of nanostructured liquids with potential use as lubricants

cpl	168.00	J/mol×K	283.15	Isobaric heat capacity of nanostructured liquids with potential use as lubricants
cpl	222.54	J/mol×K	353.15	Molar heat capacities for (1-butanol + 1,4-butanediol, 2,3-butanediol, 1,2-butanediol, and 2-methyl-2,4-pentanediol) as function of temperature
cpl	218.10	J/mol×K	348.15	Molar heat capacities for (1-butanol + 1,4-butanediol, 2,3-butanediol, 1,2-butanediol, and 2-methyl-2,4-pentanediol) as function of temperature
cpl	213.57	J/mol×K	343.15	Molar heat capacities for (1-butanol + 1,4-butanediol, 2,3-butanediol, 1,2-butanediol, and 2-methyl-2,4-pentanediol) as function of temperature
cpl	209.16	J/mol×K	338.15	Molar heat capacities for (1-butanol + 1,4-butanediol, 2,3-butanediol, 1,2-butanediol, and 2-methyl-2,4-pentanediol) as function of temperature
cpl	204.72	J/mol×K	333.15	Molar heat capacities for (1-butanol + 1,4-butanediol, 2,3-butanediol, 1,2-butanediol, and 2-methyl-2,4-pentanediol) as function of temperature

cpl	200.40	J/mol×K	328.15	Molar heat capacities for (1-butanol + 1,4-butanediol, 2,3-butanediol, 1,2-butanediol, and 2-methyl-2,4-pentanediol) as function of temperature
cpl	196.13	J/mol×K	323.15	Molar heat capacities for (1-butanol + 1,4-butanediol, 2,3-butanediol, 1,2-butanediol, and 2-methyl-2,4-pentanediol) as function of temperature
cpl	192.10	J/mol×K	318.15	Molar heat capacities for (1-butanol + 1,4-butanediol, 2,3-butanediol, 1,2-butanediol, and 2-methyl-2,4-pentanediol) as function of temperature
cpl	188.02	J/mol×K	313.15	Molar heat capacities for (1-butanol + 1,4-butanediol, 2,3-butanediol, 1,2-butanediol, and 2-methyl-2,4-pentanediol) as function of temperature
cpl	184.28	J/mol×K	308.15	Molar heat capacities for (1-butanol + 1,4-butanediol, 2,3-butanediol, 1,2-butanediol, and 2-methyl-2,4-pentanediol) as function of temperature
cpl	180.61	J/mol×K	303.15	Molar heat capacities for (1-butanol + 1,4-butanediol, 2,3-butanediol, 1,2-butanediol, and 2-methyl-2,4-pentanediol) as function of temperature

cpl	177.10	J/mol×K	298.15	Molar heat capacities for (1-butanol + 1,4-butanediol, 2,3-butanediol, 1,2-butanediol, and 2-methyl-2,4-pentanediol) as function of temperature
dvisc	0.0022962	Pa×s	303.15	Densities and Viscosities of Binary Mixtures of n-Butanol with 2-Butanol, Isobutanol, and tert-Butanol from (303.15 to 343.15) K
dvisc	0.0014140	Pa×s	323.15	Densities and Viscosities of Naphthalen-1-ol, Naphthalen-2-ol, and 1-Aminonaphthalene in the Solvents of Ethanol, Propan-1-ol, and Butan-1-ol
dvisc	0.0017690	Pa×s	313.15	Densities and Viscosities of Naphthalen-1-ol, Naphthalen-2-ol, and 1-Aminonaphthalene in the Solvents of Ethanol, Propan-1-ol, and Butan-1-ol
dvisc	0.0022600	Pa×s	303.15	Densities and Viscosities of Naphthalen-1-ol, Naphthalen-2-ol, and 1-Aminonaphthalene in the Solvents of Ethanol, Propan-1-ol, and Butan-1-ol
dvisc	0.0029290	Pa×s	293.15	Densities and Viscosities of Naphthalen-1-ol, Naphthalen-2-ol, and 1-Aminonaphthalene in the Solvents of Ethanol, Propan-1-ol, and Butan-1-ol

dvisc	0.0011430	Pa×s	333.15	Densities and Viscosities of Naphthalen-1-ol, Naphthalen-2-ol, and 1-Aminonaphthalene in the Solvents of Ethanol, Propan-1-ol, and Butan-1-ol
dvisc	0.0029530	Pa×s	293.15	Density, Viscosity, and Isobaric Heat Capacity of the Mixture (1-Butanol + 1-Hexene)
dvisc	0.0017442	Pa×s	313.15	Molecular interactions in binary mixtures of formamide with 1-butanol, 2-butanol, 1,3-butanediol and 1,4-butanediol at different temperatures: An ultrasonic and viscometric study
dvisc	0.0022630	Pa×s	303.15	Density, Viscosity, and Isobaric Heat Capacity of the Mixture (1-Butanol + 1-Hexene)
dvisc	0.0017690	Pa×s	313.15	Density, Viscosity, and Isobaric Heat Capacity of the Mixture (1-Butanol + 1-Hexene)
dvisc	0.0038220	Pa×s	283.15	Density and Viscosity of the Binary Systems Ethanol + Butan-1-ol, + Pentan-1-ol, + Heptan-1-ol, + Octan-1-ol, Nonan-1-ol, + Decan-1-ol at 0.1 MPa and Temperatures from 283.15 K to 313.15 K

dvisc	0.0028590	Pa×s	293.15	Density and Viscosity of the Binary Systems Ethanol + Butan-1-ol, + Pentan-1-ol, + Heptan-1-ol, + Octan-1-ol, Nonan-1-ol, + Decan-1-ol at 0.1 MPa and Temperatures from 283.15 K to 313.15 K	
dvisc	0.0025580	Pa×s	298.15	Density and Viscosity of the Binary Systems Ethanol + Butan-1-ol, + Pentan-1-ol, + Heptan-1-ol, + Octan-1-ol, Nonan-1-ol, + Decan-1-ol at 0.1 MPa and Temperatures from 283.15 K to 313.15 K	
dvisc	0.0022360	Pa×s	303.15	Density and Viscosity of the Binary Systems Ethanol + Butan-1-ol, + Pentan-1-ol, + Heptan-1-ol, + Octan-1-ol, Nonan-1-ol, + Decan-1-ol at 0.1 MPa and Temperatures from 283.15 K to 313.15 K	
dvisc	0.0017360	Pa×s	313.15	Density and Viscosity of the Binary Systems Ethanol + Butan-1-ol, + Pentan-1-ol, + Heptan-1-ol, + Octan-1-ol, Nonan-1-ol, + Decan-1-ol at 0.1 MPa and Temperatures from 283.15 K to 313.15 K	

dvisc	0.0033748	Pa×s	288.15	Densities, Viscosities and Refractive Indices of the Ternary Mixture Dimethyladipate + 2-Butanone + 1-Butanol at T = (288.15 to 323.15) K
dvisc	0.0029446	Pa×s	293.15	Densities, Viscosities and Refractive Indices of the Ternary Mixture Dimethyladipate + 2-Butanone + 1-Butanol at T = (288.15 to 323.15) K
dvisc	0.0025809	Pa×s	298.15	Densities, Viscosities and Refractive Indices of the Ternary Mixture Dimethyladipate + 2-Butanone + 1-Butanol at T = (288.15 to 323.15) K
dvisc	0.0022719	Pa×s	303.15	Densities, Viscosities and Refractive Indices of the Ternary Mixture Dimethyladipate + 2-Butanone + 1-Butanol at T = (288.15 to 323.15) K
dvisc	0.0020081	Pa×s	308.15	Densities, Viscosities and Refractive Indices of the Ternary Mixture Dimethyladipate + 2-Butanone + 1-Butanol at T = (288.15 to 323.15) K
dvisc	0.0017820	Pa×s	313.15	Densities, Viscosities and Refractive Indices of the Ternary Mixture Dimethyladipate + 2-Butanone + 1-Butanol at T = (288.15 to 323.15) K

dvisc	0.0015872	Pa×s	318.15	Densities, Viscosities and Refractive Indices of the Ternary Mixture Dimethyladipate + 2-Butanone + 1-Butanol at T = (288.15 to 323.15) K
dvisc	0.0014188	Pa×s	323.15	Densities, Viscosities and Refractive Indices of the Ternary Mixture Dimethyladipate + 2-Butanone + 1-Butanol at T = (288.15 to 323.15) K
dvisc	0.0025730	Pa×s	298.15	Temperature and Composition Dependence of the Densities, Viscosities, and Speeds of Sound of Binary Liquid Mixtures of 1-Butanol with Hexadecane and Squalane
dvisc	0.0020253	Pa×s	308.15	Densities and Viscosities of Binary Mixtures of n-Butanol with 2-Butanol, Isobutanol, and tert-Butanol from (303.15 to 343.15) K
dvisc	0.0019810	Pa×s	308.15	Temperature and Composition Dependence of the Densities, Viscosities, and Speeds of Sound of Binary Liquid Mixtures of 1-Butanol with Hexadecane and Squalane
dvisc	0.0029420	Pa×s	293.15	Viscosities and Densities of Binary Mixtures of (N-Acetylmorpholine + Alkanols) from (293.15 to 323.15) K

dvisc	0.0022750	Pa×s	303.15	Viscosities and Densities of Binary Mixtures of (N-Acetylmorpholine + Alkanols) from (293.15 to 323.15) K
dvisc	0.0016840	Pa×s	313.15	Viscosities and Densities of Binary Mixtures of (N-Acetylmorpholine + Alkanols) from (293.15 to 323.15) K
dvisc	0.0010980	Pa×s	323.15	Viscosities and Densities of Binary Mixtures of (N-Acetylmorpholine + Alkanols) from (293.15 to 323.15) K
dvisc	0.0025880	Pa×s	298.15	Densities and Viscosities of Binary Mixtures of Cyclopropanecarboxylic Acid with Methanol, Ethanol, Propan-1-ol, and Butan-1-ol at Different Temperatures
dvisc	0.0019910	Pa×s	308.15	Densities and Viscosities of Binary Mixtures of Cyclopropanecarboxylic Acid with Methanol, Ethanol, Propan-1-ol, and Butan-1-ol at Different Temperatures
dvisc	0.0015953	Pa×s	318.15	Densities and Viscosities of Binary Mixtures of n-Butanol with 2-Butanol, Isobutanol, and tert-Butanol from (303.15 to 343.15) K

dvisc	0.0014170	Pa×s	323.15	Densities and Viscosities of Binary Mixtures of n-Butanol with 2-Butanol, Isobutanol, and tert-Butanol from (303.15 to 343.15) K	
dvisc	0.0012714	Pa×s	328.15	Densities and Viscosities of Binary Mixtures of n-Butanol with 2-Butanol, Isobutanol, and tert-Butanol from (303.15 to 343.15) K	
dvisc	0.0011381	Pa×s	333.15	Densities and Viscosities of Binary Mixtures of n-Butanol with 2-Butanol, Isobutanol, and tert-Butanol from (303.15 to 343.15) K	
dvisc	0.0010275	Pa×s	338.15	Densities and Viscosities of Binary Mixtures of n-Butanol with 2-Butanol, Isobutanol, and tert-Butanol from (303.15 to 343.15) K	
dvisc	0.0009306	Pa×s	343.15	Densities and Viscosities of Binary Mixtures of n-Butanol with 2-Butanol, Isobutanol, and tert-Butanol from (303.15 to 343.15) K	
dvisc	0.0015550	Pa×s	318.15	Densities and Viscosities of Binary Mixtures of Cyclopropanecarboxylic Acid with Methanol, Ethanol, Propan-1-ol, and Butan-1-ol at Different Temperatures	

dvisc	0.0012280	Pa×s	328.15	Densities and Viscosities of Binary Mixtures of Cyclopropanecarboxylic Acid with Methanol, Ethanol, Propan-1-ol, and Butan-1-ol at Different Temperatures
dvisc	0.0009810	Pa×s	338.15	Densities and Viscosities of Binary Mixtures of Cyclopropanecarboxylic Acid with Methanol, Ethanol, Propan-1-ol, and Butan-1-ol at Different Temperatures
dvisc	0.0029410	Pa×s	293.15	Densities, Viscosities, and Excess Molar Enthalpies of 2-Pyrrolidone + Butanol Isomers at T) (293.15, 298.15, and 303.15) K
dvisc	0.0022750	Pa×s	303.15	Temperature and Composition Dependence of the Densities, Viscosities, and Speeds of Sound of Binary Liquid Mixtures of 1-Butanol with Hexadecane and Squalane
dvisc	0.0025720	Pa×s	298.15	Densities, Viscosities, and Excess Molar Enthalpies of 2-Pyrrolidone + Butanol Isomers at T) (293.15, 298.15, and 303.15) K
dvisc	0.0022750	Pa×s	303.15	Densities, Viscosities, and Excess Molar Enthalpies of 2-Pyrrolidone + Butanol Isomers at T) (293.15, 298.15, and 303.15) K

dvisc	0.0029401	Pa×s	293.15	Densities and Viscosities of exo-Tetrahydrodicyclopentadiene + n-Butanol and exo-Tetrahydrodicyclopentadiene + n-Pentanol at Temperatures of (293.15 to 313.15) K
dvisc	0.0025612	Pa×s	298.15	Densities and Viscosities of exo-Tetrahydrodicyclopentadiene + n-Butanol and exo-Tetrahydrodicyclopentadiene + n-Pentanol at Temperatures of (293.15 to 313.15) K
dvisc	0.0022866	Pa×s	303.15	Densities and Viscosities of exo-Tetrahydrodicyclopentadiene + n-Butanol and exo-Tetrahydrodicyclopentadiene + n-Pentanol at Temperatures of (293.15 to 313.15) K
dvisc	0.0029380	Pa×s	293.15	Viscosities and densities for binary mixtures of N-methylpiperazine with methanol, ethanol, n-propanol, iso-propanol, n-butanol and iso-butanol at 293.15, 298.15 and 303.15K
dvisc	0.0025720	Pa×s	298.15	Viscosities and densities for binary mixtures of N-methylpiperazine with methanol, ethanol, n-propanol, iso-propanol, n-butanol and iso-butanol at 293.15, 298.15 and 303.15K

dvisc	0.0022710	Pa×s	303.15	Viscosities and densities for binary mixtures of N-methylpiperazine with methanol, ethanol, n-propanol, iso-propanol, n-butanol and iso-butanol at 293.15, 298.15 and 303.15K
dvisc	0.0029356	Pa×s	293.15	Molecular interactions in binary mixtures of formamide with 1-butanol, 2-butanol, 1,3-butanediol and 1,4-butanediol at different temperatures: An ultrasonic and viscometric study
dvisc	0.0025416	Pa×s	298.15	Molecular interactions in binary mixtures of formamide with 1-butanol, 2-butanol, 1,3-butanediol and 1,4-butanediol at different temperatures: An ultrasonic and viscometric study
dvisc	0.0022751	Pa×s	303.15	Molecular interactions in binary mixtures of formamide with 1-butanol, 2-butanol, 1,3-butanediol and 1,4-butanediol at different temperatures: An ultrasonic and viscometric study
dvisc	0.0019978	Pa×s	308.15	Molecular interactions in binary mixtures of formamide with 1-butanol, 2-butanol, 1,3-butanediol and 1,4-butanediol at different temperatures: An ultrasonic and viscometric study

dvisc	0.0017954	Pa×s	313.15	Densities and Viscosities of Binary Mixtures of n-Butanol with 2-Butanol, Isobutanol, and tert-Butanol from (303.15 to 343.15) K	
dvisc	0.0015712	Pa×s	318.15	Molecular interactions in binary mixtures of formamide with 1-butanol, 2-butanol, 1,3-butanediol and 1,4-butanediol at different temperatures: An ultrasonic and viscometric study	
dvisc	0.0020104	Pa×s	308.15	Densities and Viscosities for Binary Mixtures of the Ionic Liquid N-Ethyl Piperazinium Propionate with n-Alcohols at Several Temperatures	
dvisc	0.0025710	Pa×s	298.15	Viscosity, density, and speed of sound of methylcyclopentane with primary and secondary alcohols at T = (293.15, 298.15, and 303.15) K	
dvisc	0.0022550	Pa×s	303.15	Viscosity, density, and speed of sound of methylcyclopentane with primary and secondary alcohols at T = (293.15, 298.15, and 303.15) K	

dvisc	0.0025850	Pa×s	298.15	Densities, viscosities, and ultrasonic velocity studies of binary mixtures of trichloromethane with methanol, ethanol, propan-1-ol, and butan-1-ol at T=(298.15 and 308.15) K	
dvisc	0.0020120	Pa×s	308.15	Densities, viscosities, and ultrasonic velocity studies of binary mixtures of trichloromethane with methanol, ethanol, propan-1-ol, and butan-1-ol at T=(298.15 and 308.15) K	
dvisc	0.0029410	Pa×s	293.15	Dynamic viscosities of binary mixtures of cycloalkanes with primary alcohols at T = (293.15, 298.15, and 303.15) K: New UNIFAC-VISCO interaction parameters	
dvisc	0.0025710	Pa×s	298.15	Dynamic viscosities of binary mixtures of cycloalkanes with primary alcohols at T = (293.15, 298.15, and 303.15) K: New UNIFAC-VISCO interaction parameters	
dvisc	0.0022550	Pa×s	303.15	Dynamic viscosities of binary mixtures of cycloalkanes with primary alcohols at T = (293.15, 298.15, and 303.15) K: New UNIFAC-VISCO interaction parameters	

dvisc	0.0038794	Pa×s	283.15	Density and viscosity of 1-butyl-3-methylimidazolium nitrate with ethanol, 1-propanol, or 1-butanol at several temperatures
dvisc	0.0033911	Pa×s	288.15	Density and viscosity of 1-butyl-3-methylimidazolium nitrate with ethanol, 1-propanol, or 1-butanol at several temperatures
dvisc	0.0030100	Pa×s	293.15	Density and viscosity of 1-butyl-3-methylimidazolium nitrate with ethanol, 1-propanol, or 1-butanol at several temperatures
dvisc	0.0026062	Pa×s	298.15	Density and viscosity of 1-butyl-3-methylimidazolium nitrate with ethanol, 1-propanol, or 1-butanol at several temperatures
dvisc	0.0023147	Pa×s	303.15	Density and viscosity of 1-butyl-3-methylimidazolium nitrate with ethanol, 1-propanol, or 1-butanol at several temperatures
dvisc	0.0020655	Pa×s	308.15	Density and viscosity of 1-butyl-3-methylimidazolium nitrate with ethanol, 1-propanol, or 1-butanol at several temperatures
dvisc	0.0018476	Pa×s	313.15	Density and viscosity of 1-butyl-3-methylimidazolium nitrate with ethanol, 1-propanol, or 1-butanol at several temperatures

dvisc	0.0016597	Pa×s	318.15	Density and viscosity of 1-butyl-3-methylimidazolium nitrate with ethanol, 1-propanol, or 1-butanol at several temperatures
dvisc	0.0014952	Pa×s	323.15	Density and viscosity of 1-butyl-3-methylimidazolium nitrate with ethanol, 1-propanol, or 1-butanol at several temperatures
dvisc	0.0013567	Pa×s	328.15	Density and viscosity of 1-butyl-3-methylimidazolium nitrate with ethanol, 1-propanol, or 1-butanol at several temperatures
dvisc	0.0012435	Pa×s	333.15	Density and viscosity of 1-butyl-3-methylimidazolium nitrate with ethanol, 1-propanol, or 1-butanol at several temperatures
dvisc	0.0029410	Pa×s	293.15	Properties of pure 1-methylimidazolium acetate ionic liquid and its binary mixtures with alcohols
dvisc	0.0025730	Pa×s	298.15	Properties of pure 1-methylimidazolium acetate ionic liquid and its binary mixtures with alcohols
dvisc	0.0022640	Pa×s	303.15	Properties of pure 1-methylimidazolium acetate ionic liquid and its binary mixtures with alcohols

dvisc	0.0020020	Pa×s	308.15	Properties of pure 1-methylimidazolium acetate ionic liquid and its binary mixtures with alcohols
dvisc	0.0017880	Pa×s	313.15	Properties of pure 1-methylimidazolium acetate ionic liquid and its binary mixtures with alcohols
dvisc	0.0029590	Pa×s	293.15	Viscosities of binary mixtures containing 1-butanol + 2,2,4-trimethylpentane or + 1,2,4-trimethylbenzene at high pressures for the thermophysical characterization of biofuels
dvisc	0.0017940	Pa×s	313.15	Viscosities of binary mixtures containing 1-butanol + 2,2,4-trimethylpentane or + 1,2,4-trimethylbenzene at high pressures for the thermophysical characterization of biofuels
dvisc	0.0011543	Pa×s	333.15	Viscosities of binary mixtures containing 1-butanol + 2,2,4-trimethylpentane or + 1,2,4-trimethylbenzene at high pressures for the thermophysical characterization of biofuels
dvisc	0.0007837	Pa×s	353.15	Viscosities of binary mixtures containing 1-butanol + 2,2,4-trimethylpentane or + 1,2,4-trimethylbenzene at high pressures for the thermophysical characterization of biofuels

dvisc	0.0014060	Pa×s	323.15	Thermophysical properties of 1,2,4-trimethylbenzene in admixtures with 1-butanol or 2-butanol at high pressures
dvisc	0.0011340	Pa×s	333.15	Thermophysical properties of 1,2,4-trimethylbenzene in admixtures with 1-butanol or 2-butanol at high pressures
dvisc	0.0022820	Pa×s	303.15	Viscosities and Densities of Binary Mixtures of 1,4-Dioxane, Carbon Tetrachloride, and Butanol at 303.15 K, 308.15 K, and 313.15 K
dvisc	0.0019073	Pa×s	308.15	Viscosities and Densities of Binary Mixtures of 1,4-Dioxane, Carbon Tetrachloride, and Butanol at 303.15 K, 308.15 K, and 313.15 K
dvisc	0.0018947	Pa×s	308.15	Viscosities and Densities of Binary Mixtures of 1,4-Dioxane, Carbon Tetrachloride, and Butanol at 303.15 K, 308.15 K, and 313.15 K
dvisc	0.0016813	Pa×s	313.15	Viscosities and Densities of Binary Mixtures of 1,4-Dioxane, Carbon Tetrachloride, and Butanol at 303.15 K, 308.15 K, and 313.15 K
dvisc	0.0026180	Pa×s	298.15	Excess Molar Volumes and Viscosity Deviations of Binary Mixtures of 2,4,6-Trimethyl-1,3,5-trioxane + Ethanol, 1-Propanol, and 1-Butanol at (298.15, 303.15, and 308.15) K

dvisc	0.0022970	Pa×s	303.15	Excess Molar Volumes and Viscosity Deviations of Binary Mixtures of 2,4,6-Trimethyl-1,3,5-trioxane + Ethanol, 1-Propanol, and 1-Butanol at (298.15, 303.15, and 308.15) K
dvisc	0.0020320	Pa×s	308.15	Excess Molar Volumes and Viscosity Deviations of Binary Mixtures of 2,4,6-Trimethyl-1,3,5-trioxane + Ethanol, 1-Propanol, and 1-Butanol at (298.15, 303.15, and 308.15) K
dvisc	0.0033460	Pa×s	288.15	Densities and Viscosities of Binary Mixtures of 1-Chlorobutane with Butanol Isomers at Several Temperatures
dvisc	0.0025600	Pa×s	298.15	Densities and Viscosities of Binary Mixtures of 1-Chlorobutane with Butanol Isomers at Several Temperatures
dvisc	0.0019910	Pa×s	308.15	Densities and Viscosities of Binary Mixtures of 1-Chlorobutane with Butanol Isomers at Several Temperatures
dvisc	0.0015690	Pa×s	318.15	Densities and Viscosities of Binary Mixtures of 1-Chlorobutane with Butanol Isomers at Several Temperatures

dvisc	0.0038690	Pa×s	283.15	Excess Enthalpy, Density, Viscosity, and Speed of Sound for the Mixture Tetrahydropyran + 1-Butanol at (283.15, 298.15 and, 313.15) K	
dvisc	0.0025609	Pa×s	298.15	Excess Enthalpy, Density, Viscosity, and Speed of Sound for the Mixture Tetrahydropyran + 1-Butanol at (283.15, 298.15 and, 313.15) K	
dvisc	0.0017567	Pa×s	313.15	Excess Enthalpy, Density, Viscosity, and Speed of Sound for the Mixture Tetrahydropyran + 1-Butanol at (283.15, 298.15 and, 313.15) K	
dvisc	0.0025640	Pa×s	298.15	Viscosities, Densities, and Ultrasonic Velocities of Binary Mixtures of Ethylbenzene with Ethanol, 1-Propanol, and 1-Butanol at (298.15 and 308.15)K	
dvisc	0.0020000	Pa×s	308.15	Viscosities, Densities, and Ultrasonic Velocities of Binary Mixtures of Ethylbenzene with Ethanol, 1-Propanol, and 1-Butanol at (298.15 and 308.15)K	
dvisc	0.0033460	Pa×s	288.15	Densities and Viscosities of Binary Mixtures of 1-Bromobutane with Butanol Isomers at Several Temperatures	

dvisc	0.0025600	Pa×s	298.15	Densities and Viscosities of Binary Mixtures of 1-Bromobutane with Butanol Isomers at Several Temperatures
dvisc	0.0019910	Pa×s	308.15	Densities and Viscosities of Binary Mixtures of 1-Bromobutane with Butanol Isomers at Several Temperatures
dvisc	0.0015690	Pa×s	318.15	Densities and Viscosities of Binary Mixtures of 1-Bromobutane with Butanol Isomers at Several Temperatures
dvisc	0.0022740	Pa×s	303.15	Densities, Viscosities, and Ultrasonic Velocity Studies of Binary Mixtures of Chloroform with Propan-1-ol and Butan-1-ol at (303.15 and 313.15) K
dvisc	0.0017830	Pa×s	313.15	Densities, Viscosities, and Ultrasonic Velocity Studies of Binary Mixtures of Chloroform with Propan-1-ol and Butan-1-ol at (303.15 and 313.15) K
dvisc	0.0025660	Pa×s	298.15	Density and Viscosity for Ethyl 3-Ethoxypropionate + Methacrylic Acid, + Benzyl Methacrylate, and + 2-Hydroxyethyl Methacrylate

dvisc	0.0022650	Pa×s	303.15	Density and Viscosity for Ethyl 3-Ethoxypropionate + Methacrylic Acid, + Benzyl Methacrylate, and + 2-Hydroxyethyl Methacrylate
dvisc	0.0019980	Pa×s	308.15	Density and Viscosity for Ethyl 3-Ethoxypropionate + Methacrylic Acid, + Benzyl Methacrylate, and + 2-Hydroxyethyl Methacrylate
dvisc	0.0017780	Pa×s	313.15	Density and Viscosity for Ethyl 3-Ethoxypropionate + Methacrylic Acid, + Benzyl Methacrylate, and + 2-Hydroxyethyl Methacrylate
dvisc	0.0038690	Pa×s	283.15	Excess Enthalpy, Density, Speed of Sound, and Viscosity for 2-Methyltetrahydrofuran + 1-Butanol at (283.15, 298.15, and, 313.15) K
dvisc	0.0025609	Pa×s	298.15	Excess Enthalpy, Density, Speed of Sound, and Viscosity for 2-Methyltetrahydrofuran + 1-Butanol at (283.15, 298.15, and, 313.15) K
dvisc	0.0017567	Pa×s	313.15	Excess Enthalpy, Density, Speed of Sound, and Viscosity for 2-Methyltetrahydrofuran + 1-Butanol at (283.15, 298.15, and, 313.15) K

dvisc	0.0029630	Pa×s	293.15	Density and Viscosity of Binary Mixtures of Diethyl Carbonate with Alcohols at (293.15 to 363.15) K and Predictive Results by UNIFAC-VISCO Group Contribution Method
dvisc	0.0026190	Pa×s	298.15	Density and Viscosity of Binary Mixtures of Diethyl Carbonate with Alcohols at (293.15 to 363.15) K and Predictive Results by UNIFAC-VISCO Group Contribution Method
dvisc	0.0022980	Pa×s	303.15	Density and Viscosity of Binary Mixtures of Diethyl Carbonate with Alcohols at (293.15 to 363.15) K and Predictive Results by UNIFAC-VISCO Group Contribution Method
dvisc	0.0017860	Pa×s	313.15	Density and Viscosity of Binary Mixtures of Diethyl Carbonate with Alcohols at (293.15 to 363.15) K and Predictive Results by UNIFAC-VISCO Group Contribution Method

dvisc	0.0014210	Pa×s	323.15	Density and Viscosity of Binary Mixtures of Diethyl Carbonate with Alcohols at (293.15 to 363.15) K and Predictive Results by UNIFAC-VISCO Group Contribution Method
dvisc	0.0011370	Pa×s	333.15	Density and Viscosity of Binary Mixtures of Diethyl Carbonate with Alcohols at (293.15 to 363.15) K and Predictive Results by UNIFAC-VISCO Group Contribution Method
dvisc	0.0009200	Pa×s	343.15	Density and Viscosity of Binary Mixtures of Diethyl Carbonate with Alcohols at (293.15 to 363.15) K and Predictive Results by UNIFAC-VISCO Group Contribution Method
dvisc	0.0007590	Pa×s	353.15	Density and Viscosity of Binary Mixtures of Diethyl Carbonate with Alcohols at (293.15 to 363.15) K and Predictive Results by UNIFAC-VISCO Group Contribution Method

dvisc	0.0006300	Pa×s	363.15	Density and Viscosity of Binary Mixtures of Diethyl Carbonate with Alcohols at (293.15 to 363.15) K and Predictive Results by UNIFAC-VISCO Group Contribution Method	
dvisc	0.0032860	Pa×s	288.15	Densities and Viscosities of Binary Mixtures of 1-Butanol with Methacrylic Acid, Benzyl Methacrylate, and 2-Hydroxyethyl Methacrylate between 288.15 K and 318.15 K	
dvisc	0.0025660	Pa×s	298.15	Densities and Viscosities of Binary Mixtures of 1-Butanol with Methacrylic Acid, Benzyl Methacrylate, and 2-Hydroxyethyl Methacrylate between 288.15 K and 318.15 K	
dvisc	0.0019980	Pa×s	308.15	Densities and Viscosities of Binary Mixtures of 1-Butanol with Methacrylic Acid, Benzyl Methacrylate, and 2-Hydroxyethyl Methacrylate between 288.15 K and 318.15 K	
dvisc	0.0015500	Pa×s	318.15	Densities and Viscosities of Binary Mixtures of 1-Butanol with Methacrylic Acid, Benzyl Methacrylate, and 2-Hydroxyethyl Methacrylate between 288.15 K and 318.15 K	

dvisc	0.0025768	Pa×s	298.15	Densities and Viscosities for Binary Mixtures of the Ionic Liquid N-Ethyl Piperazinium Propionate with n-Alcohols at Several Temperatures
dvisc	0.0022719	Pa×s	303.15	Densities and Viscosities for Binary Mixtures of the Ionic Liquid N-Ethyl Piperazinium Propionate with n-Alcohols at Several Temperatures
dvisc	0.0029410	Pa×s	293.15	Viscosity, density, and speed of sound of methylcyclopentane with primary and secondary alcohols at T = (293.15, 298.15, and 303.15) K
dvisc	0.0017856	Pa×s	313.15	Densities and Viscosities for Binary Mixtures of the Ionic Liquid N-Ethyl Piperazinium Propionate with n-Alcohols at Several Temperatures
dvisc	0.0017673	Pa×s	313.15	Densities and Viscosities of exo-Tetrahydrodicyclopentadiene + n-Butanol and exo-Tetrahydrodicyclopentadiene + n-Pentanol at Temperatures of (293.15 to 313.15) K
econd	8.00e-05	S/m	318.15	Electrical Conductivity of Caprolactam Tetrabutylammonium Bromide Ionic Liquids in Aqueous and Alcohol Binary Systems

econd	6.90e-05	S/m	313.15	Electrical Conductivity of Caprolactam Tetrabutylammonium Bromide Ionic Liquids in Aqueous and Alcohol Binary Systems
econd	6.58e-05	S/m	308.15	Electrical Conductivity of Caprolactam Tetrabutylammonium Bromide Ionic Liquids in Aqueous and Alcohol Binary Systems
econd	5.55e-05	S/m	303.15	Electrical Conductivity of Caprolactam Tetrabutylammonium Bromide Ionic Liquids in Aqueous and Alcohol Binary Systems
econd	4.40e-05	S/m	298.15	Electrical Conductivity of Caprolactam Tetrabutylammonium Bromide Ionic Liquids in Aqueous and Alcohol Binary Systems
hfust	9.37	kJ/mol	184.50	NIST Webbook
hfust	9.28	kJ/mol	183.90	NIST Webbook
hfust	9.28	kJ/mol	183.90	NIST Webbook
hfust	9.28	kJ/mol	183.90	NIST Webbook
hvapt	48.30	kJ/mol	363.50	NIST Webbook
hvapt	49.50 ± 0.10	kJ/mol	333.00	NIST Webbook
hvapt	49.00	kJ/mol	360.00	NIST Webbook
hvapt	47.20	kJ/mol	374.00	NIST Webbook
hvapt	37.40	kJ/mol	530.00	NIST Webbook
hvapt	41.90	kJ/mol	458.00	NIST Webbook
hvapt	43.80	kJ/mol	410.00	NIST Webbook
hvapt	45.40	kJ/mol	386.50	NIST Webbook
hvapt	51.60	kJ/mol	230.00	NIST Webbook
hvapt	41.90	kJ/mol	481.50	NIST Webbook
hvapt	50.10	kJ/mol	368.00	NIST Webbook
hvapt	45.30	kJ/mol	387.50	NIST Webbook
hvapt	45.30	kJ/mol	383.50	NIST Webbook
hvapt	49.90	kJ/mol	352.50	NIST Webbook
hvapt	20.80	kJ/mol	523.00	NIST Webbook

hvapt	29.60	kJ/mol	473.00	NIST Webbook
hvapt	38.20	kJ/mol	423.00	NIST Webbook
hvapt	46.00	kJ/mol	373.00	NIST Webbook
hvapt	43.29	kJ/mol	390.90	NIST Webbook
hvapt	43.09	kJ/mol	390.60	KDB
hvapt	48.60 ± 0.10	kJ/mol	343.00	NIST Webbook
hvapt	47.50 ± 0.10	kJ/mol	353.00	NIST Webbook
hvapt	46.40 ± 0.10	kJ/mol	363.00	NIST Webbook
hvapt	55.00	kJ/mol	346.00	NIST Webbook
hvapt	53.00	kJ/mol	343.00	NIST Webbook
hvapt	45.40 ± 0.10	kJ/mol	381.00	NIST Webbook
hvapt	43.10 ± 0.10	kJ/mol	391.00	NIST Webbook
hvapt	42.10	kJ/mol	491.00	NIST Webbook
hvapt	46.60	kJ/mol	380.00	NIST Webbook
hvapt	48.30	kJ/mol	352.00	NIST Webbook
hvapt	47.20 ± 0.10	kJ/mol	356.00	NIST Webbook
kvisc	0.0000032	m2/s	298.15	Viscosities of Binary Mixtures of Isomeric Butanols or Isomeric Chlorobutanes with 2-Methyltetrahydrofuran
kvisc	0.0000022	m2/s	313.15	Viscosities of Binary Mixtures of Isomeric Butanols or Isomeric Chlorobutanes with 2-Methyltetrahydrofuran
pvap	53.33	kPa	373.74	Isobaric Vapor-Liquid Equilibria for Tetrahydropyran and Alcohol Systems
pvap	3.60	kPa	318.40	Vapor-Liquid Equilibrium for 1-Butanol + 1-Butene at (318.4 and 364.5) K and Vapor-Liquid Equilibrium of 1-Butanol + 2-Methylpropane, + n-Butane and 1-Butene + 2-Methylpropane at 318.4 K

pvap	3.60	kPa	318.35	Vapor-Liquid Equilibrium for 1-Butanol + 1-Butene at (318.4 and 364.5) K and Vapor-Liquid Equilibrium of 1-Butanol + 2-Methylpropane, + n-Butane and 1-Butene + 2-Methylpropane at 318.4 K
pvap	3.90	kPa	318.35	Vapor-Liquid Equilibrium for 1-Butanol + 1-Butene at (318.4 and 364.5) K and Vapor-Liquid Equilibrium of 1-Butanol + 2-Methylpropane, + n-Butane and 1-Butene + 2-Methylpropane at 318.4 K
pvap	98.66	kPa	390.04	Isobaric Vapor-Liquid Equilibria for Tetrahydropyran and Alcohol Systems
pvap	93.32	kPa	388.49	Isobaric Vapor-Liquid Equilibria for Tetrahydropyran and Alcohol Systems
pvap	36.30	kPa	364.52	Vapor-Liquid Equilibrium for 1-Butanol + 1-Butene at (318.4 and 364.5) K and Vapor-Liquid Equilibrium of 1-Butanol + 2-Methylpropane, + n-Butane and 1-Butene + 2-Methylpropane at 318.4 K
pvap	1.83	kPa	308.15	Isothermal Vapor-Liquid Equilibrium of (1-Butanol + 1,8-Cineole) at 10 Temperatures between (278.15 and 323.15) K

pvap	36.90	kPa	364.50	Vapour liquid equilibrium for the systems butane + methanol, +2-propanol, +1-butanol, +2-butanol, +2-methyl-2-propanol at 364.5K
pvap	101.30	kPa	390.65	Isobaric VLE data for the system of butan-1-ol + butylethanoate + 1-butyl-3-methylimidazoliumbis[(trifluoromethyl)sulfonyl]imide
pvap	2.51	kPa	313.15	Isothermal vapor liquid equilibrium and molar excess Gibbs energies of two ternary systems containing either 1-butanol or 2-butanol + 1-hexene + methylbenzene at 313.15 K
pvap	1.79	kPa	308.15	(Vapour + liquid) equilibria of (1-butanol + benzene, or toluene, or o-, or m-, or p-xylene) at T = 308.15 K
pvap	0.91	kPa	298.15	(Vapour + liquid) equilibrium and excess Gibbs functions of ternary mixtures containing 1-butanol or 2-butanol, n-hexane, and 1-chlorobutane at T = 298.15 K
pvap	25.30	kPa	356.35	Isobaric (vapour + liquid) equilibria for the (1-propanol + 1-butanol) binary mixture at (53.3 and 91.3) kPa
pvap	29.30	kPa	359.45	Isobaric (vapour + liquid) equilibria for the (1-propanol + 1-butanol) binary mixture at (53.3 and 91.3) kPa

pvap	33.50	kPa	362.95	Isobaric (vapour + liquid) equilibria for the (1-propanol + 1-butanol) binary mixture at (53.3 and 91.3) kPa
pvap	42.80	kPa	368.60	Isobaric (vapour + liquid) equilibria for the (1-propanol + 1-butanol) binary mixture at (53.3 and 91.3) kPa
pvap	52.50	kPa	373.45	Isobaric (vapour + liquid) equilibria for the (1-propanol + 1-butanol) binary mixture at (53.3 and 91.3) kPa
pvap	61.20	kPa	377.30	Isobaric (vapour + liquid) equilibria for the (1-propanol + 1-butanol) binary mixture at (53.3 and 91.3) kPa
pvap	72.80	kPa	381.90	Isobaric (vapour + liquid) equilibria for the (1-propanol + 1-butanol) binary mixture at (53.3 and 91.3) kPa
pvap	84.90	kPa	386.05	Isobaric (vapour + liquid) equilibria for the (1-propanol + 1-butanol) binary mixture at (53.3 and 91.3) kPa
pvap	105.30	kPa	391.95	Isobaric (vapour + liquid) equilibria for the (1-propanol + 1-butanol) binary mixture at (53.3 and 91.3) kPa
pvap	120.10	kPa	395.80	Isobaric (vapour + liquid) equilibria for the (1-propanol + 1-butanol) binary mixture at (53.3 and 91.3) kPa

pvap	2.48	kPa	313.15	Vapour-liquid equilibria of binary and ternary mixtures containing 1-butanol, 2,2,4-trimethylpentane and 1-hexene at T = 313.15 K
pvap	34.47	kPa	363.20	(Vapour + liquid) equilibria, (VLE) excess molar enthalpies and infinite dilution activity coefficients of selected binary systems involving n-hexyl pyridinium bis(trifluoromethylsulphonyl)imide ionic liquid: Experimental and predictions using modified UNIFAC (Dortmund)
pvap	14.44	kPa	343.39	Thermogravimetric measurement of evaporation: Data analysis based on the Stefan tube
pvap	13.71	kPa	343.36	Thermogravimetric measurement of evaporation: Data analysis based on the Stefan tube
pvap	13.52	kPa	343.20	Thermogravimetric measurement of evaporation: Data analysis based on the Stefan tube
pvap	21.03	kPa	352.70	Thermogravimetric measurement of evaporation: Data analysis based on the Stefan tube
pvap	32.09	kPa	360.58	Thermogravimetric measurement of evaporation: Data analysis based on the Stefan tube

pvap	4.74	kPa	323.90	Thermogravimetric measurement of evaporation: Data analysis based on the Stefan tube
pvap	4.95	kPa	324.80	Thermogravimetric measurement of evaporation: Data analysis based on the Stefan tube
pvap	8.38	kPa	334.32	Thermogravimetric measurement of evaporation: Data analysis based on the Stefan tube
pvap	8.09	kPa	333.51	Thermogravimetric measurement of evaporation: Data analysis based on the Stefan tube
pvap	4.98	kPa	324.91	Thermogravimetric measurement of evaporation: Data analysis based on the Stefan tube
pvap	2.61	kPa	314.81	Thermogravimetric measurement of evaporation: Data analysis based on the Stefan tube
pvap	2.60	kPa	314.76	Thermogravimetric measurement of evaporation: Data analysis based on the Stefan tube
pvap	2.67	kPa	314.75	Thermogravimetric measurement of evaporation: Data analysis based on the Stefan tube
pvap	1.96	kPa	310.26	Thermogravimetric measurement of evaporation: Data analysis based on the Stefan tube

pvap	1.90	kPa	310.29	Thermogravimetric measurement of evaporation: Data analysis based on the Stefan tube
pvap	1.46	kPa	305.59	Thermogravimetric measurement of evaporation: Data analysis based on the Stefan tube
pvap	1.92	kPa	310.16	Thermogravimetric measurement of evaporation: Data analysis based on the Stefan tube
pvap	1.09	kPa	301.02	Thermogravimetric measurement of evaporation: Data analysis based on the Stefan tube
pvap	1.91	kPa	310.29	Thermogravimetric measurement of evaporation: Data analysis based on the Stefan tube
pvap	4.62	kPa	323.50	Thermogravimetric measurement of evaporation: Data analysis based on the Stefan tube
pvap	4.91	kPa	324.57	Thermogravimetric measurement of evaporation: Data analysis based on the Stefan tube
pvap	4.87	kPa	324.59	Thermogravimetric measurement of evaporation: Data analysis based on the Stefan tube

pvap	12.79	kPa	341.78	Thermogravimetric measurement of evaporation: Data analysis based on the Stefan tube
pvap	4.48	kPa	322.58	Thermogravimetric measurement of evaporation: Data analysis based on the Stefan tube
pvap	4.85	kPa	324.28	Thermogravimetric measurement of evaporation: Data analysis based on the Stefan tube
pvap	4.52	kPa	322.69	Thermogravimetric measurement of evaporation: Data analysis based on the Stefan tube
pvap	4.33	kPa	322.07	Thermogravimetric measurement of evaporation: Data analysis based on the Stefan tube
pvap	4.69	kPa	323.67	Thermogravimetric measurement of evaporation: Data analysis based on the Stefan tube
pvap	4.75	kPa	323.78	Thermogravimetric measurement of evaporation: Data analysis based on the Stefan tube
pvap	79.99	kPa	384.26	Isobaric Vapor-Liquid Equilibria for Tetrahydropyran and Alcohol Systems
pvap	4.15	kPa	321.34	Thermogravimetric measurement of evaporation: Data analysis based on the Stefan tube

pvap	4.19	kPa	321.34	Thermogravimetric measurement of evaporation: Data analysis based on the Stefan tube
pvap	4.54	kPa	323.09	Thermogravimetric measurement of evaporation: Data analysis based on the Stefan tube
pvap	2.50	kPa	314.76	Thermogravimetric measurement of evaporation: Data analysis based on the Stefan tube
pvap	4.79	kPa	324.18	Thermogravimetric measurement of evaporation: Data analysis based on the Stefan tube
pvap	101.47	kPa	390.69	Vapor-Liquid Equilibrium of Binary Mixtures Containing Isopropyl Acetate and Alkanols at 101.32 kPa
pvap	127.02	kPa	397.19	Vapor-Liquid Equilibrium of Binary Mixtures Containing Isopropyl Acetate and Alkanols at 101.32 kPa
pvap	152.42	kPa	402.79	Vapor-Liquid Equilibrium of Binary Mixtures Containing Isopropyl Acetate and Alkanols at 101.32 kPa
pvap	177.83	kPa	407.58	Vapor-Liquid Equilibrium of Binary Mixtures Containing Isopropyl Acetate and Alkanols at 101.32 kPa
pvap	203.23	kPa	411.88	Vapor-Liquid Equilibrium of Binary Mixtures Containing Isopropyl Acetate and Alkanols at 101.32 kPa

pvap	228.64	kPa	415.88	Vapor-Liquid Equilibrium of Binary Mixtures Containing Isopropyl Acetate and Alkanols at 101.32 kPa
pvap	85.00	kPa	385.75	Experimental Vapor-Liquid Equilibria and Thermodynamic Modeling of the Methanol + n-Heptane and 1-Butanol + Aniline Binary Systems
pvap	0.56	kPa	293.15	Dibutyl Citrate Synthesis, Physicochemical Characterization, and Px Data in Mixtures with Butanol
pvap	0.84	kPa	298.15	Dibutyl Citrate Synthesis, Physicochemical Characterization, and Px Data in Mixtures with Butanol
pvap	1.23	kPa	303.15	Dibutyl Citrate Synthesis, Physicochemical Characterization, and Px Data in Mixtures with Butanol
pvap	1.76	kPa	308.15	Dibutyl Citrate Synthesis, Physicochemical Characterization, and Px Data in Mixtures with Butanol
pvap	2.44	kPa	313.15	Dibutyl Citrate Synthesis, Physicochemical Characterization, and Px Data in Mixtures with Butanol
pvap	3.33	kPa	318.15	Dibutyl Citrate Synthesis, Physicochemical Characterization, and Px Data in Mixtures with Butanol

pvap	4.35	kPa	323.15	Dibutyl Citrate Synthesis, Physicochemical Characterization, and Px Data in Mixtures with Butanol
pvap	5.33	kPa	328.15	Dibutyl Citrate Synthesis, Physicochemical Characterization, and Px Data in Mixtures with Butanol
pvap	101.30	kPa	390.75	Vapor-Liquid Equilibrium for Binary of 1-Butanol + N,N-Dimethylacetamide and Methyl Isobutyl Ketone + N,N-Dimethylacetamide at 101.3 kPa
pvap	2.70	kPa	314.19	Vapor-Liquid Equilibrium Studies for Systems Containing n-Butylisocyanate at Temperatures between 323.15 K and 371.15 K
pvap	4.49	kPa	323.20	Vapor-Liquid Equilibrium Studies for Systems Containing n-Butylisocyanate at Temperatures between 323.15 K and 371.15 K
pvap	10.38	kPa	338.21	Vapor-Liquid Equilibrium Studies for Systems Containing n-Butylisocyanate at Temperatures between 323.15 K and 371.15 K
pvap	12.89	kPa	342.17	Vapor-Liquid Equilibrium Studies for Systems Containing n-Butylisocyanate at Temperatures between 323.15 K and 371.15 K

pvap	21.44	kPa	352.81	Vapor-Liquid Equilibrium Studies for Systems Containing n-Butylisocyanate at Temperatures between 323.15 K and 371.15 K
pvap	31.52	kPa	361.24	Vapor-Liquid Equilibrium Studies for Systems Containing n-Butylisocyanate at Temperatures between 323.15 K and 371.15 K
pvap	49.26	kPa	371.94	Vapor-Liquid Equilibrium Studies for Systems Containing n-Butylisocyanate at Temperatures between 323.15 K and 371.15 K
pvap	1.00	kPa	298.15	Vapor-Pressure Measurements of Liquid Solutions at Different Temperatures: Apparatus for Use over an Extended Temperature Range and Some New Data
pvap	1.32	kPa	303.15	Vapor-Pressure Measurements of Liquid Solutions at Different Temperatures: Apparatus for Use over an Extended Temperature Range and Some New Data
pvap	1.84	kPa	308.15	Vapor-Pressure Measurements of Liquid Solutions at Different Temperatures: Apparatus for Use over an Extended Temperature Range and Some New Data

pvap	2.54	kPa	313.15	Vapor-Pressure Measurements of Liquid Solutions at Different Temperatures: Apparatus for Use over an Extended Temperature Range and Some New Data	
pvap	3.45	kPa	318.15	Vapor-Pressure Measurements of Liquid Solutions at Different Temperatures: Apparatus for Use over an Extended Temperature Range and Some New Data	
pvap	4.65	kPa	323.15	Vapor-Pressure Measurements of Liquid Solutions at Different Temperatures: Apparatus for Use over an Extended Temperature Range and Some New Data	
pvap	6.17	kPa	328.15	Vapor-Pressure Measurements of Liquid Solutions at Different Temperatures: Apparatus for Use over an Extended Temperature Range and Some New Data	
pvap	8.10	kPa	333.15	Vapor-Pressure Measurements of Liquid Solutions at Different Temperatures: Apparatus for Use over an Extended Temperature Range and Some New Data	

pvap	10.56	kPa	338.15	Vapor-Pressure Measurements of Liquid Solutions at Different Temperatures: Apparatus for Use over an Extended Temperature Range and Some New Data	
pvap	13.60	kPa	343.15	Vapor-Pressure Measurements of Liquid Solutions at Different Temperatures: Apparatus for Use over an Extended Temperature Range and Some New Data	
pvap	17.37	kPa	348.15	Vapor-Pressure Measurements of Liquid Solutions at Different Temperatures: Apparatus for Use over an Extended Temperature Range and Some New Data	
pvap	21.98	kPa	353.15	Vapor-Pressure Measurements of Liquid Solutions at Different Temperatures: Apparatus for Use over an Extended Temperature Range and Some New Data	
pvap	27.65	kPa	358.15	Vapor-Pressure Measurements of Liquid Solutions at Different Temperatures: Apparatus for Use over an Extended Temperature Range and Some New Data	

pvap	34.41	kPa	363.15	Vapor-Pressure Measurements of Liquid Solutions at Different Temperatures: Apparatus for Use over an Extended Temperature Range and Some New Data	
pvap	833.00	kPa	468.15	Isothermal Vapor-Liquid Equilibria for the n-Pentane + 1-Butanol and n-Pentane + 2-Butanol Systems near the Critical Region of the Mixtures	
pvap	1182.00	kPa	483.15	Isothermal Vapor-Liquid Equilibria for the n-Pentane + 1-Butanol and n-Pentane + 2-Butanol Systems near the Critical Region of the Mixtures	
pvap	1558.00	kPa	498.15	Isothermal Vapor-Liquid Equilibria for the n-Pentane + 1-Butanol and n-Pentane + 2-Butanol Systems near the Critical Region of the Mixtures	
pvap	2027.00	kPa	513.15	Isothermal Vapor-Liquid Equilibria for the n-Pentane + 1-Butanol and n-Pentane + 2-Butanol Systems near the Critical Region of the Mixtures	

pvap	4.61	kPa	323.15	Vapor-Liquid Equilibria and HE for Binary Systems of Dimethyl Ether (DME) with C1-C4 Alkan-1-ols at 323.15 K and Liquid-Liquid Equilibria for Ternary System of DME + Methanol + Water at 313.15 K
pvap	0.20	kPa	278.15	Isothermal Vapor-Liquid Equilibrium of (1-Butanol + 1,8-Cineole) at 10 Temperatures between (278.15 and 323.15) K
pvap	0.31	kPa	283.15	Isothermal Vapor-Liquid Equilibrium of (1-Butanol + 1,8-Cineole) at 10 Temperatures between (278.15 and 323.15) K
pvap	0.45	kPa	288.15	Isothermal Vapor-Liquid Equilibrium of (1-Butanol + 1,8-Cineole) at 10 Temperatures between (278.15 and 323.15) K
pvap	0.66	kPa	293.15	Isothermal Vapor-Liquid Equilibrium of (1-Butanol + 1,8-Cineole) at 10 Temperatures between (278.15 and 323.15) K
pvap	0.94	kPa	298.15	Isothermal Vapor-Liquid Equilibrium of (1-Butanol + 1,8-Cineole) at 10 Temperatures between (278.15 and 323.15) K
pvap	1.32	kPa	303.15	Isothermal Vapor-Liquid Equilibrium of (1-Butanol + 1,8-Cineole) at 10 Temperatures between (278.15 and 323.15) K

pvap	66.66	kPa	379.43	Isobaric Vapor-Liquid Equilibria for Tetrahydropyran and Alcohol Systems
pvap	2.53	kPa	313.15	Isothermal Vapor-Liquid Equilibrium of (1-Butanol + 1,8-Cineole) at 10 Temperatures between (278.15 and 323.15) K
pvap	3.44	kPa	318.15	Isothermal Vapor-Liquid Equilibrium of (1-Butanol + 1,8-Cineole) at 10 Temperatures between (278.15 and 323.15) K
pvap	4.54	kPa	323.15	Isothermal Vapor-Liquid Equilibrium of (1-Butanol + 1,8-Cineole) at 10 Temperatures between (278.15 and 323.15) K
pvap	60.00	kPa	376.78	Isobaric Vapor Liquid Equilibria for the 1-Propanol + Ethylene Glycol Monopropyl Ether and 1-Butanol + Ethylene Glycol Monopropyl Ether Systems
pvap	80.00	kPa	384.28	Isobaric Vapor Liquid Equilibria for the 1-Propanol + Ethylene Glycol Monopropyl Ether and 1-Butanol + Ethylene Glycol Monopropyl Ether Systems
pvap	100.00	kPa	390.28	Isobaric Vapor Liquid Equilibria for the 1-Propanol + Ethylene Glycol Monopropyl Ether and 1-Butanol + Ethylene Glycol Monopropyl Ether Systems

pvap	101.30	kPa	390.65	Vapor Liquid Equilibrium at p/kPa = 101.3 of the Binary Mixtures of Ethenyl Acetate with Methanol and Butan-1-ol
pvap	8.39	kPa	333.15	Vapor Liquid Equilibrium for Several Compounds Relevant to the Biofuels Industry Modeled with the Wilson Equation
pvap	110.26	kPa	393.15	Vapor Liquid Equilibrium for Several Compounds Relevant to the Biofuels Industry Modeled with the Wilson Equation
pvap	101.30	kPa	390.48	Phase Equilibrium (VLE, LLE, and VLLE) Data of the Ternary System: Ionic Liquid [OMIM][PF6] + Butan-1-ol + Butyl Acetate
pvap	101.30	kPa	390.49	Vapor Liquid Equilibria Measurements for the Nine n-Alkane/Ketone Pairs Comprising 2-, 3-, and 4-Heptanone with n-Octane, n-Nonane, and n-Decane
pvap	40.00	kPa	366.76	Isobaric Vapor-Liquid Equilibria for Tetrahydropyran and Alcohol Systems
pvap	4.54	kPa	323.14	Thermogravimetric measurement of evaporation: Data analysis based on the Stefan tube

rfi	1.39523	303.15	Experimental Determination and Modeling of Densities and Refractive Indices of the Binary Systems Alcohol + Dicyclohexylamine at T = (288.15 to 323.15) K	
rfi	1.39686	298.15	Liquid-Liquid Equilibrium in Binary System [bmim][PF6] + 1-Butanol	
rfi	1.39300	308.15	Densities, Excess Molar Volumes, and Refractive Properties of the Binary Mixtures of the Amino Acid Ionic Liquid [bmim][Gly] with 1-Butanol or Isopropanol at T = (298.15 to 313.15) K	
rfi	1.39500	303.15	Densities, Excess Molar Volumes, and Refractive Properties of the Binary Mixtures of the Amino Acid Ionic Liquid [bmim][Gly] with 1-Butanol or Isopropanol at T = (298.15 to 313.15) K	
rfi	1.39700	298.15	Densities, Excess Molar Volumes, and Refractive Properties of the Binary Mixtures of the Amino Acid Ionic Liquid [bmim][Gly] with 1-Butanol or Isopropanol at T = (298.15 to 313.15) K	
rfi	1.39730	298.15	Excess Molar Enthalpies of 2-Methyl-2-butanol (1) + 1-Alkanols (C1-C5) (2) at 298.15 K	

rfi	1.39730	298.15	Volumetric Behavior of Carbon Dioxide + Butan-1-ol Mixtures at 313.15 K
rfi	1.39740	298.15	Excess Molar Enthalpies of Benzyl Alcohol + Alkanols (C1-C6) and Their Correlations at 298.15 K and Ambient Pressure
rfi	1.39100	313.15	Densities, Excess Molar Volumes, and Refractive Properties of the Binary Mixtures of the Amino Acid Ionic Liquid [bmim][Gly] with 1-Butanol or Isopropanol at T = (298.15 to 313.15) K
rfi	1.39930	293.15	Liquid-Liquid Equilibria of Water + 3-Hydroxy-2-butanone + 1-Butanol
rfi	1.39730	298.15	Isothermal Vapor Liquid Equilibrium Data of Propan-1-ol + 2,2,4-Trimethylpentane and Butan-1-ol + 2,2,4-Trimethylpentane at 318.15 K
rfi	1.39730	298.15	Density, Viscosity, and Refractive Index Properties for the Binary Mixtures of n-Butylammonium Acetate Ionic Liquid + Alkanols at Several Temperatures
rfi	1.39930	293.20	Vapor Liquid Equilibrium Data for Binary Systems of 1-Methyl-4-(1-methylethenyl)-cyclohexene + {Ethanol, Propan-1-ol, Propan-2-ol, Butan-1-ol, Pentan-1-ol, or Hexan-1-ol} at 40 kPa

rfi	1.39930	293.20	Vapor Liquid Equilibrium Data for Binary Systems of 1H-Pyrrole with Butan-1-ol, Propan-1-ol, or Pentan-1-ol
rfi	1.39720	298.15	Vapor Liquid Equilibria, Excess Enthalpy, and Excess Volume of Binary Mixtures Containing an Alcohol (1-Butanol, 2-Butanol, or 2-Methyl-2-butanol) and 2-Ethoxy-2-methylbutane
rfi	1.39920	293.15	Vapor Liquid Equilibrium Data for 1-Methyl-2-Pyrrolidone + (1-Butanol or 1-Hexene or Water) Binary Mixtures
rfi	1.39920	293.15	Vapor Liquid Equilibrium Data for Binary Systems of n-Dodecane + {Propan-1-ol, Butan-1-ol, 2-Methylpropan-1-ol} at 40 kPa
rfi	1.39928	293.15	Ternary Liquid Liquid Equilibrium Data for the Water + Acetonitrile + {Butan-1-ol or 2-Methylpropan-1-ol} Systems at (303.2, 323.2, 343.2) K and 1 atm
rfi	1.39726	298.15	Experimental Liquid-Liquid Equilibria of 1-Alkyl-3-methylimidazolium Hexafluorophosphate with 1-Alcohols
rfi	1.39600	298.15	Phase Equilibrium for the Esterification Reaction of Acetic Acid + Butan-1-ol at 101.3 kPa

rfi	1.39970	293.15	Densities, Excess Molar Volumes, and Isobaric Thermal Expansibilities for 1,2-Ethanediol + 1-Butanol, or 1-Hexanol, or 1-Octanol in the Temperature Range from (293.15 to 313.15) K
rfi	1.39780	298.15	Densities, Excess Molar Volumes, and Isobaric Thermal Expansibilities for 1,2-Ethanediol + 1-Butanol, or 1-Hexanol, or 1-Octanol in the Temperature Range from (293.15 to 313.15) K
rfi	1.39940	293.15	Refractive Indices and Deviations in Refractive Indices for Binary Mixtures of Formamide + 1-Butanol, + 2-Butanol, + 1,3-Butanediol, and + 1,4-Butanediol at Temperatures from (293.15 to 318.15) K
rfi	1.39740	298.15	Refractive Indices and Deviations in Refractive Indices for Binary Mixtures of Formamide + 1-Butanol, + 2-Butanol, + 1,3-Butanediol, and + 1,4-Butanediol at Temperatures from (293.15 to 318.15) K

rfi	1.39530	303.15	Refractive Indices and Deviations in Refractive Indices for Binary Mixtures of Formamide + 1-Butanol, + 2-Butanol, + 1,3-Butanediol, and + 1,4-Butanediol at Temperatures from (293.15 to 318.15) K	
rfi	1.39330	308.15	Refractive Indices and Deviations in Refractive Indices for Binary Mixtures of Formamide + 1-Butanol, + 2-Butanol, + 1,3-Butanediol, and + 1,4-Butanediol at Temperatures from (293.15 to 318.15) K	
rfi	1.39130	313.15	Refractive Indices and Deviations in Refractive Indices for Binary Mixtures of Formamide + 1-Butanol, + 2-Butanol, + 1,3-Butanediol, and + 1,4-Butanediol at Temperatures from (293.15 to 318.15) K	
rfi	1.38930	318.15	Refractive Indices and Deviations in Refractive Indices for Binary Mixtures of Formamide + 1-Butanol, + 2-Butanol, + 1,3-Butanediol, and + 1,4-Butanediol at Temperatures from (293.15 to 318.15) K	

rfi	1.39720	298.15	Effect of Pressure on the Static Relative Permittivities of Alkan-1-ols at 298.15 K
rfi	1.39740	298.15	Excess Molar Enthalpies of 1,2-Propanediol + Alkan-1-ols (C1-C6) and Their Correlations at 298.15 K and Ambient Pressure (81.5 kPa)
rfi	1.39370	308.15	Density, Viscosity, Refractive Index, and Speed of Sound for Binary Mixtures of 1,4-Dioxane with Different Organic Liquids at (298.15, 303.15, and 308.15) K
rfi	1.39550	303.15	Density, Viscosity, Refractive Index, and Speed of Sound for Binary Mixtures of 1,4-Dioxane with Different Organic Liquids at (298.15, 303.15, and 308.15) K
rfi	1.39830	298.15	Density, Viscosity, Refractive Index, and Speed of Sound for Binary Mixtures of 1,4-Dioxane with Different Organic Liquids at (298.15, 303.15, and 308.15) K
rfi	1.39730	298.15	Density, Surface Tension, and Refractive Index of Octane + 1-Alkanol Mixtures at T) 298.15 K.

rfi	1.39920	293.15	Limiting Activity Coefficients by Comparative Tensimetry: 1-Propanol and 1-Butanol in Heptane and in Octane
rfi	1.39320	308.15	Excess Molar Volumes, Viscosities, and Refractive Indexes for Binary Mixtures of 1-Chlorobutane with Four Alcohols at T = (288.15, 298.15 and 308.15) K
rfi	1.39720	298.15	Excess Molar Volumes, Viscosities, and Refractive Indexes for Binary Mixtures of 1-Chlorobutane with Four Alcohols at T = (288.15, 298.15 and 308.15) K
rfi	1.40100	288.15	Excess Molar Volumes, Viscosities, and Refractive Indexes for Binary Mixtures of 1-Chlorobutane with Four Alcohols at T = (288.15, 298.15 and 308.15) K
rfi	1.39740	298.15	Bubble-Temperature Measurements on Some Binary Mixtures Formed by Tetrahydrofuran or Amyl Alcohol with Hydrocarbons, Chlorohydrocarbons, or Butanols at (94.6 or 95.8) kPa

rfi	1.39800	298.15	Bubble Temperatures of the Binary Mixtures of Dimethylcarbonate with Some Alcohols at 95.8 kPa
rfi	1.39890	293.15	Solubility Data for Roflumilast and Maraviroc in Various Solvents between T = (278.2-323.2) K
rfi	1.39860	293.15	Solid-Liquid Equilibrium Measurements for Posaconazole and Voriconazole in Several Solvents between T = 278.2 and 323.2 K Using Differential Thermal Analysis/Thermal Gravimetric Analysis
rfi	1.39890	293.15	Isothermal Vapor-Liquid Equilibrium Measurements for the Pentan-2-one + Propan-1-ol/Butan-1-ol System within 342-363 K
rfi	1.39318	308.15	Experimental Determination and Modeling of Densities and Refractive Indices of the Binary Systems Alcohol + Dicyclohexylamine at T = (288.15 to 323.15) K
rfi	1.39940	293.15	Liquid Liquid Equilibrium Data of Water + Butyric Acid + {Butanal or n-Butanol} Ternary Systems at 293.15, 308.15, and 323.15 K

rfi	1.38688	323.15	Experimental Determination and Modeling of Densities, Refractive Indices of the Binary Mixtures of Dimethylphthalate (or Dimethyladipate) + 1-Butanol, or + 2-Butanol, or + 2-Butanone at T = (288.15 to 323.15) K
rfi	1.38901	318.15	Experimental Determination and Modeling of Densities, Refractive Indices of the Binary Mixtures of Dimethylphthalate (or Dimethyladipate) + 1-Butanol, or + 2-Butanol, or + 2-Butanone at T = (288.15 to 323.15) K
rfi	1.39110	313.15	Experimental Determination and Modeling of Densities, Refractive Indices of the Binary Mixtures of Dimethylphthalate (or Dimethyladipate) + 1-Butanol, or + 2-Butanol, or + 2-Butanone at T = (288.15 to 323.15) K
rfi	1.39318	308.15	Experimental Determination and Modeling of Densities, Refractive Indices of the Binary Mixtures of Dimethylphthalate (or Dimethyladipate) + 1-Butanol, or + 2-Butanol, or + 2-Butanone at T = (288.15 to 323.15) K

rfi	1.39523	303.15	Experimental Determination and Modeling of Densities, Refractive Indices of the Binary Mixtures of Dimethylphthalate (or Dimethyladipate) + 1-Butanol, or + 2-Butanol, or + 2-Butanone at T = (288.15 to 323.15) K
rfi	1.39730	298.15	Experimental Determination and Modeling of Densities, Refractive Indices of the Binary Mixtures of Dimethylphthalate (or Dimethyladipate) + 1-Butanol, or + 2-Butanol, or + 2-Butanone at T = (288.15 to 323.15) K
rfi	1.39932	293.15	Experimental Determination and Modeling of Densities, Refractive Indices of the Binary Mixtures of Dimethylphthalate (or Dimethyladipate) + 1-Butanol, or + 2-Butanol, or + 2-Butanone at T = (288.15 to 323.15) K
rfi	1.40135	288.15	Experimental Determination and Modeling of Densities, Refractive Indices of the Binary Mixtures of Dimethylphthalate (or Dimethyladipate) + 1-Butanol, or + 2-Butanol, or + 2-Butanone at T = (288.15 to 323.15) K

rfi	1.38688	323.15	Experimental Determination and Modeling of Densities and Refractive Indices of the Binary Systems Alcohol + Dicyclohexylamine at T = (288.15 to 323.15) K
rfi	1.38901	318.15	Experimental Determination and Modeling of Densities and Refractive Indices of the Binary Systems Alcohol + Dicyclohexylamine at T = (288.15 to 323.15) K
rfi	1.39110	313.15	Experimental Determination and Modeling of Densities and Refractive Indices of the Binary Systems Alcohol + Dicyclohexylamine at T = (288.15 to 323.15) K
rfi	1.39460	303.15	Measurements of the Properties of Binary Mixtures of Dimethylsulphoxide (DMSO) with 1-Alkanols (C4, C6, C7) at 303.15K
rfi	1.39730	298.15	Isobaric Vapor Liquid Equilibria of 1-Butanol p-Xylene System1
rfi	1.39930	293.15	Vapor liquid equilibria and excess volumes of the binary systems ethanol + ethyl lactate, isopropanol + isopropyl lactate and n-butanol + n-butyl lactate at 101.325 kPa

rfi	1.39740	298.15	Activity coefficients of the binary mixtures of a-cresol or p-cresol with C l-C4 aliphatic alcohols near ambient pressure
rfi	1.39760	298.15	Isobaric vapor-liquid equilibria for the binary systems 1-propyl alcohol + dipropyl ether and 1-butyl alcohol + dibutyl ether at 20 and 101.3 kPa
rfi	1.39732	298.15	Isobaric vapour liquid equilibria for the binary systems 4-methyl-2-pentanone + 1-butanol and + 2-butanol at 20 and 101.3 kPa
rfi	1.39740	298.15	Excess molar volumes, excess molar enthalpies and refractive index deviations for binary mixtures of propan-1-ol, butan-1-ol and pentan-1-ol with 2,2,4-trimethylpentane at 298.15 K
rfi	1.40130	288.20	Experimental and calculated liquid-liquid equilibrium data for water + furfural + solvents
rfi	1.39520	303.20	Experimental and calculated liquid-liquid equilibrium data for water + furfural + solvents
rfi	1.38920	318.20	Experimental and calculated liquid-liquid equilibrium data for water + furfural + solvents

rfi	1.39950	298.15	Investigation on vapor liquid equilibrium for 2-propanol + 1-butanol + 1-pentanol at 101.3 kPa
rfi	1.39830	293.20	Liquid-liquid equilibrium data for systems containing of formic acid, water, and primary normal alcohols at T = 298.2 K
rfi	1.39730	298.15	Isobaric (vapor+liquid) equilibrium for n-propyl acetate with 1-butanol or 2-butanol. Binary mixtures at 0.15 and 0.6 MPa
rfi	1.39886	298.15	Atmospheric densities and interfacial tensions for 1-alkanol (1-butanol to 1-octanol) + water and ether (MTBE, ETBE, DIPE, TAME and THP) + water demixed mixtures.
rfi	1.39700	298.15	Excess molar volumes and partial molar volumes for (propionitrile + an alkanol) at T = 298.15 K and p = 0.1 MPa
rfi	1.39970	293.15	Measurement and Correlation of the Solubilities of m-Phthalic Acid in Monobasic Alcohols

rfi	1.39590	303.15	Physico-chemical and excess properties of the binary mixtures of methylcyclohexane + ethanol, + propan-1-ol, + propan-2-ol, + butan-1-ol, + 2-methyl-1-propanol, or 3-methyl-1-butanol at T = (298.15, 303.15, and 308.15) K
rfi	1.39400	308.15	Physico-chemical and excess properties of the binary mixtures of methylcyclohexane + ethanol, + propan-1-ol, + propan-2-ol, + butan-1-ol, + 2-methyl-1-propanol, or 3-methyl-1-butanol at T = (298.15, 303.15, and 308.15) K
rfi	1.39680	298.15	Thermodynamic properties of (tetradecane + benzene, + toluene, + chlorobenzene, + bromobenzene, + anisole) binary mixtures at T = (298.15, 303.15, and 308.15) K
rfi	1.39440	303.15	Thermodynamic properties of (tetradecane + benzene, + toluene, + chlorobenzene, + bromobenzene, + anisole) binary mixtures at T = (298.15, 303.15, and 308.15) K
rfi	1.39230	308.15	Thermodynamic properties of (tetradecane + benzene, + toluene, + chlorobenzene, + bromobenzene, + anisole) binary mixtures at T = (298.15, 303.15, and 308.15) K

rfi	1.39740	298.15	(Vapor + liquid) equilibria of the binary mixtures of m-cresol with C1 C4 aliphatic alcohols at 95.5 kPa	
rfi	1.39680	298.15	Thermodynamic interactions in binary mixtures of anisole with ethanol, propan-1-ol, propan-2-ol, butan-1-ol, pentan-1-ol, and 3-methylbutan-1-ol at T = (298.15, 303.15, and 308.15) K	
rfi	1.39440	303.15	Thermodynamic interactions in binary mixtures of anisole with ethanol, propan-1-ol, propan-2-ol, butan-1-ol, pentan-1-ol, and 3-methylbutan-1-ol at T = (298.15, 303.15, and 308.15) K	
rfi	1.39230	308.15	Thermodynamic interactions in binary mixtures of anisole with ethanol, propan-1-ol, propan-2-ol, butan-1-ol, pentan-1-ol, and 3-methylbutan-1-ol at T = (298.15, 303.15, and 308.15) K	
rfi	1.39740	298.15	Excess Gibbs free energies of the binary mixtures of acetonitrile with butanols at 94.83 kPa	

rfi	1.39479	303.15	Density, viscosity, refractive index, excess molar enthalpy, viscosity, and refractive index deviations for the (1-butanol + 2-butanone) binary system at T = 303 K. A new adiabatic calorimeter for heat of mixing	
rfi	1.39729	298.15	Thermophysical properties of the binary mixtures (1,8-cineole + 1-alkanol) at T = (298.15 and 313.15) K and at atmospheric pressure	
rfi	1.39114	313.15	Thermophysical properties of the binary mixtures (1,8-cineole + 1-alkanol) at T = (298.15 and 313.15) K and at atmospheric pressure	
rfi	1.39640	298.15	Application of the ERAS model to volumetric properties of binary mixtures of banana oil with primary and secondary alcohols (C1- C4) at different temperatures	
rfi	1.39620	303.15	Experimental study on the calorimetric data of cyclohexanol with alkanols (C1-C4) and correlation with Wilson, NRTL and UNIQUAC models at 300 K	

rfi	1.39747	298.15	Densities, speeds of sound, and refractive indices for binary mixtures of 1-butyl-3-methylimidazolium methyl sulphate ionic liquid with alcohols at T = (298.15, 303.15, 308.15, and 313.15) K
rfi	1.39549	303.15	Densities, speeds of sound, and refractive indices for binary mixtures of 1-butyl-3-methylimidazolium methyl sulphate ionic liquid with alcohols at T = (298.15, 303.15, 308.15, and 313.15) K
rfi	1.39342	308.15	Densities, speeds of sound, and refractive indices for binary mixtures of 1-butyl-3-methylimidazolium methyl sulphate ionic liquid with alcohols at T = (298.15, 303.15, 308.15, and 313.15) K
rfi	1.39138	313.15	Densities, speeds of sound, and refractive indices for binary mixtures of 1-butyl-3-methylimidazolium methyl sulphate ionic liquid with alcohols at T = (298.15, 303.15, 308.15, and 313.15) K
rfi	1.39720	298.15	Properties of pure n-butylammonium nitrate ionic liquid and its binary mixtures of with alcohols at T = (293.15 to 313.15) K

rfi	1.39733	298.15	Volumetric behaviour of the (2,2,4-trimethylpentane + methylbenzene + butan-1-ol) ternary system and its binary sub-systems within the temperature range (298.15-328.15) K	
rfi	1.39720	298.15	Thermodynamic and acoustic properties of binary mixtures of 1-butanol with 1,2-butanediol. The comparison with the results for 1,3-, and 1,4-butanediol	
rfi	1.39680	298.20	Tie line data for the (water + butyric acid + n-butyl alcohol or amyl alcohol) at T = (298.2, 308.2, and 318.2) K and (water + butyric acid + isoamyl alcohol) at T = 298.2 K	
rfi	1.39733	298.15	Volumetric behavior of the ternary system (methyl tert-butyl ether + methylbenzene + butan-1-ol) and its binary sub-system (methyl tert-butyl ether + butan-1-ol) within the temperature range (298.15 to 328.15) K	
rfi	1.39930	293.15	Separation of water/butan-1-ol with ionic liquids in ternary liquid-liquid phase equilibrium	

rfi	1.40135	288.15	Experimental Determination and Modeling of Densities and Refractive Indices of the Binary Systems Alcohol + Dicyclohexylamine at T = (288.15 to 323.15) K
rfi	1.39932	293.15	Experimental Determination and Modeling of Densities and Refractive Indices of the Binary Systems Alcohol + Dicyclohexylamine at T = (288.15 to 323.15) K
rfi	1.39730	298.15	Experimental Determination and Modeling of Densities and Refractive Indices of the Binary Systems Alcohol + Dicyclohexylamine at T = (288.15 to 323.15) K
rfi	1.39700	298.15	Experimental Binary VLE Data of Morpholine with 1-Butanol and 3-Methyl-1-butanol Systems
rfi	1.39830	298.15	Physico-chemical and excess properties of the binary mixtures of methylcyclohexane + ethanol, + propan-1-ol, + propan-2-ol, + butan-1-ol, + 2-methyl-1-propanol, or 3-methyl-1-butanol at T = (298.15, 303.15, and 308.15) K

rhol	805.80	kg/m3	298.15	Apparent Molal Volumes and Viscosity B-Coefficients of Acetyl Salicylic Acid (2-Acetoxy Benzoic Acid) Solutions in Higher Alcohols at Different Temperatures	
rhol	805.87	kg/m3	298.15	Binary Liquid-Liquid Equilibrium (LLE) for Methyl tert-Amyl Ether (TAME) + Water from (288.15 to 313.15) K and Ternary LLE for Systems of TAME + C1-C4 Alcohols + Water at 298.15 K	
rhol	805.85	kg/m3	298.15	Binary Liquid-Liquid Equilibrium (LLE) for Dibutyl Ether (DBE) + Water from (288.15 to 318.15) K and Ternary LLE for Systems of DBE + C1 !less thanless than C4 Alcohols + Water at 298.15 K	
rhol	805.76	kg/m3	298.15	Liquid Liquid Equilibrium for Ternary Systems of Propyl Vinyl Ether + C3 or C4 Alcohols + Water at 298.15 K and Excess Molar Enthalpies for Ternary and Constituent Binary Systems of Propyl Vinyl Ether + Ethanol + Isooctane at 303.15 K	
rhol	805.92	kg/m3	298.15	Densities and Excess Molar Volumes of N-Methylmorpholine + 1-Alkanol Systems at 298.15 K	

rhol	790.27	kg/m3	318.15	Thermodynamic Properties of 1-Butyl-3-methylpyridinium Tetrafluoroborate and Its Mixtures with Water and Alkanols
rhol	805.95	kg/m3	298.15	Thermodynamic Properties of 1-Butyl-3-methylpyridinium Tetrafluoroborate and Its Mixtures with Water and Alkanols
rhol	786.25	kg/m3	323.15	Speed of sound, density and related thermodynamic excess properties of binary mixtures of butan-2-one with C1-C4 nalkanols and chloroform
rhol	790.24	kg/m3	318.15	Speed of sound, density and related thermodynamic excess properties of binary mixtures of butan-2-one with C1-C4 nalkanols and chloroform
rhol	794.19	kg/m3	313.15	Speed of sound, density and related thermodynamic excess properties of binary mixtures of butan-2-one with C1-C4 nalkanols and chloroform
rhol	798.10	kg/m3	308.15	Speed of sound, density and related thermodynamic excess properties of binary mixtures of butan-2-one with C1-C4 nalkanols and chloroform
rhol	801.97	kg/m3	303.15	Speed of sound, density and related thermodynamic excess properties of binary mixtures of butan-2-one with C1-C4 nalkanols and chloroform

rhol	805.81	kg/m3	298.15	Speed of sound, density and related thermodynamic excess properties of binary mixtures of butan-2-one with C1-C4 nalkanols and chloroform	
rhol	809.61	kg/m3	293.15	Speed of sound, density and related thermodynamic excess properties of binary mixtures of butan-2-one with C1-C4 nalkanols and chloroform	
rhol	786.29	kg/m3	323.15	Volumetric and Viscometric Study of Binary Systems of Ethyl Butyrate with Alcohols	
rhol	790.28	kg/m3	318.15	Volumetric and Viscometric Study of Binary Systems of Ethyl Butyrate with Alcohols	
rhol	794.22	kg/m3	313.15	Volumetric and Viscometric Study of Binary Systems of Ethyl Butyrate with Alcohols	
rhol	798.12	kg/m3	308.15	Volumetric and Viscometric Study of Binary Systems of Ethyl Butyrate with Alcohols	
rhol	801.99	kg/m3	303.15	Volumetric and Viscometric Study of Binary Systems of Ethyl Butyrate with Alcohols	
rhol	805.84	kg/m3	298.15	Volumetric and Viscometric Study of Binary Systems of Ethyl Butyrate with Alcohols	
rhol	809.66	kg/m3	293.15	Volumetric and Viscometric Study of Binary Systems of Ethyl Butyrate with Alcohols	

rhol	813.46	kg/m3	288.15	Volumetric and Viscometric Study of Binary Systems of Ethyl Butyrate with Alcohols
rhol	778.01	kg/m3	333.15	Densities, Refractive Indices, Viscosities, and Spectroscopic Study of 1-Amino-2-propanol + 1-Butanol and + 2-Butanol Solutions at (288.15 to 333.15) K
rhol	782.12	kg/m3	328.15	Densities, Refractive Indices, Viscosities, and Spectroscopic Study of 1-Amino-2-propanol + 1-Butanol and + 2-Butanol Solutions at (288.15 to 333.15) K
rhol	786.18	kg/m3	323.15	Densities, Refractive Indices, Viscosities, and Spectroscopic Study of 1-Amino-2-propanol + 1-Butanol and + 2-Butanol Solutions at (288.15 to 333.15) K
rhol	790.18	kg/m3	318.15	Densities, Refractive Indices, Viscosities, and Spectroscopic Study of 1-Amino-2-propanol + 1-Butanol and + 2-Butanol Solutions at (288.15 to 333.15) K

rhol	794.13	kg/m3	313.15	Densities, Refractive Indices, Viscosities, and Spectroscopic Study of 1-Amino-2-propanol + 1-Butanol and + 2-Butanol Solutions at (288.15 to 333.15) K
rhol	798.04	kg/m3	308.15	Densities, Refractive Indices, Viscosities, and Spectroscopic Study of 1-Amino-2-propanol + 1-Butanol and + 2-Butanol Solutions at (288.15 to 333.15) K
rhol	801.91	kg/m3	303.15	Densities, Refractive Indices, Viscosities, and Spectroscopic Study of 1-Amino-2-propanol + 1-Butanol and + 2-Butanol Solutions at (288.15 to 333.15) K
rhol	805.75	kg/m3	298.15	Densities, Refractive Indices, Viscosities, and Spectroscopic Study of 1-Amino-2-propanol + 1-Butanol and + 2-Butanol Solutions at (288.15 to 333.15) K
rhol	809.56	kg/m3	293.15	Densities, Refractive Indices, Viscosities, and Spectroscopic Study of 1-Amino-2-propanol + 1-Butanol and + 2-Butanol Solutions at (288.15 to 333.15) K

rhol	813.35	kg/m3	288.15	Densities, Refractive Indices, Viscosities, and Spectroscopic Study of 1-Amino-2-propanol + 1-Butanol and + 2-Butanol Solutions at (288.15 to 333.15) K
rhol	778.16	kg/m3	333.15	Densities and Viscosities of Diaminotoluene with Water, Ethanol, Propan-1-ol, and Butan-1-ol from (293.15 to 333.15) K
rhol	786.32	kg/m3	323.15	Densities and Viscosities of Diaminotoluene with Water, Ethanol, Propan-1-ol, and Butan-1-ol from (293.15 to 333.15) K
rhol	794.27	kg/m3	313.15	Densities and Viscosities of Diaminotoluene with Water, Ethanol, Propan-1-ol, and Butan-1-ol from (293.15 to 333.15) K
rhol	802.04	kg/m3	303.15	Densities and Viscosities of Diaminotoluene with Water, Ethanol, Propan-1-ol, and Butan-1-ol from (293.15 to 333.15) K
rhol	809.69	kg/m3	293.15	Densities and Viscosities of Diaminotoluene with Water, Ethanol, Propan-1-ol, and Butan-1-ol from (293.15 to 333.15) K

rhol	786.21	kg/m3	323.15	Volumetric and Transport Properties of Binary Mixtures of n-Octane + Ethanol, + 1-Propanol, + 1-Butanol, and + 1-Pentanol from (293.15 to 323.15) K at Atmospheric Pressure	
rhol	790.18	kg/m3	318.15	Volumetric and Transport Properties of Binary Mixtures of n-Octane + Ethanol, + 1-Propanol, + 1-Butanol, and + 1-Pentanol from (293.15 to 323.15) K at Atmospheric Pressure	
rhol	794.10	kg/m3	313.15	Volumetric and Transport Properties of Binary Mixtures of n-Octane + Ethanol, + 1-Propanol, + 1-Butanol, and + 1-Pentanol from (293.15 to 323.15) K at Atmospheric Pressure	
rhol	797.99	kg/m3	308.15	Volumetric and Transport Properties of Binary Mixtures of n-Octane + Ethanol, + 1-Propanol, + 1-Butanol, and + 1-Pentanol from (293.15 to 323.15) K at Atmospheric Pressure	
rhol	801.86	kg/m3	303.15	Volumetric and Transport Properties of Binary Mixtures of n-Octane + Ethanol, + 1-Propanol, + 1-Butanol, and + 1-Pentanol from (293.15 to 323.15) K at Atmospheric Pressure	

rhol	805.70	kg/m3	298.15	Volumetric and Transport Properties of Binary Mixtures of n-Octane + Ethanol, + 1-Propanol, + 1-Butanol, and + 1-Pentanol from (293.15 to 323.15) K at Atmospheric Pressure
rhol	809.53	kg/m3	293.15	Volumetric and Transport Properties of Binary Mixtures of n-Octane + Ethanol, + 1-Propanol, + 1-Butanol, and + 1-Pentanol from (293.15 to 323.15) K at Atmospheric Pressure
rhol	805.71	kg/m3	298.15	Thermophysical Properties of the Pure Ionic Liquid 1-Butyl-1-methylpyrrolidinium Dicyanamide and Its Binary Mixtures with Alcohols
rhol	809.50	kg/m3	293.15	Isobaric vapor liquid equilibium for the binary systerms of 1-butanol with o-xylene, m-xylene, p-xylene, and ethylbenzene at 101.33 kPa
rhol	805.71	kg/m3	298.15	ACSExcess Molar Enthalpies of Mixtures of (+-)-Linalool with Several Alkanols
rhol	805.95	kg/m3	298.15	Densities, Viscosities, and Speeds of Sound of Binary Liquid Mixtures of Ethylenediamine with Alcohols at T = (293.15 to 313.15) K

rhol	805.80	kg/m3	298.15	Surface Tensions of the Ternary Mixtures Containing an Isomeric Butanol + n-Hexane + 1-Chlorobutane at 298.15 K
rhol	778.03	kg/m3	333.15	Volumetric properties of the boldine + alcohol mixtures at atmospheric pressure from 283.15 to 333.15K A new method for the determination of the density of pure boldine
rhol	782.48	kg/m3	328.15	Excess Molar Volumes of 1,3-Diethyl Propanedioate with Methanol, Ethanol, Propan-1-ol, Propan-2-ol, Butan-2-ol, 2-Methyl-propan-1-ol, and Pentan-1-ol at T = (288.15, 298.15, 313.15, and 328.15) K
rhol	794.49	kg/m3	313.15	Excess Molar Volumes of 1,3-Diethyl Propanedioate with Methanol, Ethanol, Propan-1-ol, Propan-2-ol, Butan-2-ol, 2-Methyl-propan-1-ol, and Pentan-1-ol at T = (288.15, 298.15, 313.15, and 328.15) K
rhol	806.14	kg/m3	298.15	Excess Molar Volumes of 1,3-Diethyl Propanedioate with Methanol, Ethanol, Propan-1-ol, Propan-2-ol, Butan-2-ol, 2-Methyl-propan-1-ol, and Pentan-1-ol at T = (288.15, 298.15, 313.15, and 328.15) K

rhol	813.77	kg/m3	288.15	Excess Molar Volumes of 1,3-Diethyl Propanedioate with Methanol, Ethanol, Propan-1-ol, Propan-2-ol, Butan-2-ol, 2-Methyl-propan-1-ol, and Pentan-1-ol at T = (288.15, 298.15, 313.15, and 328.15) K
rhol	801.90	kg/m3	303.00	Fluid-Phase Behavior of {1-Hexyl-3-methylimidazolium Bis(trifluoromethylsulfonyl) Imide, [C6mim][NTf2], + C2-C8 n-Alcohol} Mixtures: Liquid-Liquid Equilibrium and Excess Volumes
rhol	801.90	kg/m3	303.00	Fluid-Phase Behavior of {1-Hexyl-3-methylimidazolium Bis(trifluoromethylsulfonyl) Imide, [C6mim][NTf2], + C2-C8 n-Alcohol} Mixtures: Liquid-Liquid Equilibrium and Excess Volumes
rhol	805.70	kg/m3	298.00	Fluid-Phase Behavior of {1-Hexyl-3-methylimidazolium Bis(trifluoromethylsulfonyl) Imide, [C6mim][NTf2], + C2-C8 n-Alcohol} Mixtures: Liquid-Liquid Equilibrium and Excess Volumes
rhol	805.70	kg/m3	298.00	Fluid-Phase Behavior of {1-Hexyl-3-methylimidazolium Bis(trifluoromethylsulfonyl) Imide, [C6mim][NTf2], + C2-C8 n-Alcohol} Mixtures: Liquid-Liquid Equilibrium and Excess Volumes

rhol	809.60	kg/m3	293.00	Fluid-Phase Behavior of {1-Hexyl-3-methylimidazolium Bis(trifluoromethylsulfonyl) Imide, [C6mim][NTf2], + C2-C8 n-Alcohol} Mixtures: Liquid-Liquid Equilibrium and Excess Volumes
rhol	809.60	kg/m3	293.00	Fluid-Phase Behavior of {1-Hexyl-3-methylimidazolium Bis(trifluoromethylsulfonyl) Imide, [C6mim][NTf2], + C2-C8 n-Alcohol} Mixtures: Liquid-Liquid Equilibrium and Excess Volumes
rhol	806.04	kg/m3	298.15	Measurement and Correlation of Vapor-Liquid Equilibria at T) 333.15 K and Excess Molar Volumes at T) 298.15 K for Ethanol + Dimethyl Carbonate (DMC), DMC + 1-Propanol, and DMC + 1-Butanol
rhol	805.80	kg/m3	298.15	Excess Molar Volumes and Surface Tensions of 1,2,4-Trimethylbenzene and 1,3,5-Trimethylbenzene with 1-Butanol, 2-Methyl-1-propanol, 2-Butanol, and 2-Methyl-2-propanol at 298.15 K
rhol	805.90	kg/m3	298.15	Dynamic Viscosities of Diethyl Carbonate with Linear and Secondary Alcohols at Several Temperatures
rhol	810.00	kg/m3	293.00	KDB

rhol	806.14	kg/m3	298.15	Speeds of Sound and Isentropic Compressibilities in Binary Mixtures of 2-Propanol with Several 1-Alkanols at 298.15K	
rhol	813.37	kg/m3	288.15	Temperature Dependence of Densities and Excess Molar Volumes of the Ternary Mixture (1-Butanol + Chloroform + Benzene) and its Binary Constituents (1-Butanol + Chloroform and 1-Butanol + Benzene)	
rhol	809.57	kg/m3	293.15	Temperature Dependence of Densities and Excess Molar Volumes of the Ternary Mixture (1-Butanol + Chloroform + Benzene) and its Binary Constituents (1-Butanol + Chloroform and 1-Butanol + Benzene)	
rhol	805.76	kg/m3	298.15	Temperature Dependence of Densities and Excess Molar Volumes of the Ternary Mixture (1-Butanol + Chloroform + Benzene) and its Binary Constituents (1-Butanol + Chloroform and 1-Butanol + Benzene)	

rhol	801.92	kg/m3	303.15	Temperature Dependence of Densities and Excess Molar Volumes of the Ternary Mixture (1-Butanol + Chloroform + Benzene) and its Binary Constituents (1-Butanol + Chloroform and 1-Butanol + Benzene)	
rhol	798.05	kg/m3	308.15	Temperature Dependence of Densities and Excess Molar Volumes of the Ternary Mixture (1-Butanol + Chloroform + Benzene) and its Binary Constituents (1-Butanol + Chloroform and 1-Butanol + Benzene)	
rhol	794.15	kg/m3	313.15	Temperature Dependence of Densities and Excess Molar Volumes of the Ternary Mixture (1-Butanol + Chloroform + Benzene) and its Binary Constituents (1-Butanol + Chloroform and 1-Butanol + Benzene)	
rhol	813.41	kg/m3	288.15	Volumetric Properties, Viscosities, and Refractive Indices of the Binary Systems 1-Butanol + PEG 200, + PEG 400, and + TEGDME	
rhol	809.62	kg/m3	293.15	Volumetric Properties, Viscosities, and Refractive Indices of the Binary Systems 1-Butanol + PEG 200, + PEG 400, and + TEGDME	

rhol	805.80	kg/m3	298.15	Volumetric Properties, Viscosities, and Refractive Indices of the Binary Systems 1-Butanol + PEG 200, + PEG 400, and + TEGDME
rhol	801.96	kg/m3	303.15	Volumetric Properties, Viscosities, and Refractive Indices of the Binary Systems 1-Butanol + PEG 200, + PEG 400, and + TEGDME
rhol	798.09	kg/m3	308.15	Volumetric Properties, Viscosities, and Refractive Indices of the Binary Systems 1-Butanol + PEG 200, + PEG 400, and + TEGDME
rhol	794.18	kg/m3	313.15	Volumetric Properties, Viscosities, and Refractive Indices of the Binary Systems 1-Butanol + PEG 200, + PEG 400, and + TEGDME
rhol	790.23	kg/m3	318.15	Volumetric Properties, Viscosities, and Refractive Indices of the Binary Systems 1-Butanol + PEG 200, + PEG 400, and + TEGDME
rhol	786.23	kg/m3	323.15	Volumetric Properties, Viscosities, and Refractive Indices of the Binary Systems 1-Butanol + PEG 200, + PEG 400, and + TEGDME
rhol	782.17	kg/m3	328.15	Volumetric Properties, Viscosities, and Refractive Indices of the Binary Systems 1-Butanol + PEG 200, + PEG 400, and + TEGDME

rhol	778.06	kg/m3	333.15	Volumetric Properties, Viscosities, and Refractive Indices of the Binary Systems 1-Butanol + PEG 200, + PEG 400, and + TEGDME
rhol	805.60	kg/m3	298.15	Densities, Ultrasonic Speeds, and Excess Properties of Binary Mixtures of Diethylene Glycol with 1-Butanol, 2-Butanol, and 1,4-Butanediol at Different Temperatures
rhol	801.90	kg/m3	303.15	Densities, Ultrasonic Speeds, and Excess Properties of Binary Mixtures of Diethylene Glycol with 1-Butanol, 2-Butanol, and 1,4-Butanediol at Different Temperatures
rhol	798.20	kg/m3	308.15	Densities, Ultrasonic Speeds, and Excess Properties of Binary Mixtures of Diethylene Glycol with 1-Butanol, 2-Butanol, and 1,4-Butanediol at Different Temperatures
rhol	794.50	kg/m3	313.15	Densities, Ultrasonic Speeds, and Excess Properties of Binary Mixtures of Diethylene Glycol with 1-Butanol, 2-Butanol, and 1,4-Butanediol at Different Temperatures

rhol	805.80	kg/m3	298.15	Revision of the volumetric method for measurements of liquid liquid equilibria in binary systems
rhol	805.74	kg/m3	298.15	Excess molar volumes and excess molar enthalpies of binary and ternary mixtures of (ethanol or 1-butanol), triethylamine and n-hexane
rhol	802.10	kg/m3	303.15	Viscous synergy and antagonism and isentropic compressibility of ternary mixtures containing 1,3-dioxolane, water and monoalkanols at 303.15K
rhol	817.15	kg/m3	283.15	Volumetric properties of the boldine + alcohol mixtures at atmospheric pressure from 283.15 to 333.15K A new method for the determination of the density of pure boldine
rhol	813.38	kg/m3	288.15	Volumetric properties of the boldine + alcohol mixtures at atmospheric pressure from 283.15 to 333.15K A new method for the determination of the density of pure boldine
rhol	809.59	kg/m3	293.15	Volumetric properties of the boldine + alcohol mixtures at atmospheric pressure from 283.15 to 333.15K A new method for the determination of the density of pure boldine

rhol	805.77	kg/m3	298.15	Volumetric properties of the boldine + alcohol mixtures at atmospheric pressure from 283.15 to 333.15K A new method for the determination of the density of pure boldine
rhol	801.92	kg/m3	303.15	Volumetric properties of the boldine + alcohol mixtures at atmospheric pressure from 283.15 to 333.15K A new method for the determination of the density of pure boldine
rhol	798.05	kg/m3	308.15	Volumetric properties of the boldine + alcohol mixtures at atmospheric pressure from 283.15 to 333.15K A new method for the determination of the density of pure boldine
rhol	794.14	kg/m3	313.15	Volumetric properties of the boldine + alcohol mixtures at atmospheric pressure from 283.15 to 333.15K A new method for the determination of the density of pure boldine
rhol	790.19	kg/m3	318.15	Volumetric properties of the boldine + alcohol mixtures at atmospheric pressure from 283.15 to 333.15K A new method for the determination of the density of pure boldine

rhol	786.19	kg/m3	323.15	Volumetric properties of the boldine + alcohol mixtures at atmospheric pressure from 283.15 to 333.15K A new method for the determination of the density of pure boldine	
rhol	782.14	kg/m3	328.15	Volumetric properties of the boldine + alcohol mixtures at atmospheric pressure from 283.15 to 333.15K A new method for the determination of the density of pure boldine	
rhol	805.67	kg/m3	298.15	Mixing Properties for the Ternary Mixture Methyl tert-Butyl Ether + 1-Butanol + Decane at 298.15 K	
rhol	805.71	kg/m3	298.15	Excess molar enthalpy of 1-alkanol + 1-octene mixtures at 298.15K Experimental results and theoretical description by means of the ERAS and TB models	
rhol	813.37	kg/m3	288.15	Influence of chain length and degree of branching of alcohol + chlorobenzene mixtures on determination and modelling of VE by CEOS and CEOS/GE mixing rules	

rhol	809.57	kg/m3	293.15	Influence of chain length and degree of branching of alcohol + chlorobenzene mixtures on determination and modelling of VE by CEOS and CEOS/GE mixing rules
rhol	805.76	kg/m3	298.15	Influence of chain length and degree of branching of alcohol + chlorobenzene mixtures on determination and modelling of VE by CEOS and CEOS/GE mixing rules
rhol	801.92	kg/m3	303.15	Influence of chain length and degree of branching of alcohol + chlorobenzene mixtures on determination and modelling of VE by CEOS and CEOS/GE mixing rules
rhol	798.05	kg/m3	308.15	Influence of chain length and degree of branching of alcohol + chlorobenzene mixtures on determination and modelling of VE by CEOS and CEOS/GE mixing rules
rhol	794.15	kg/m3	313.15	Influence of chain length and degree of branching of alcohol + chlorobenzene mixtures on determination and modelling of VE by CEOS and CEOS/GE mixing rules

rhol	805.78	kg/m3	298.15	Liquid liquid equilibria for the binary system of di-isopropyl ether (DIPE) +water in between 288.15 and 323.15K and the ternary systems of DIPE +water + C1 C4 alcohols at 298.15K
rhol	813.60	kg/m3	288.15	Excess molar enthalpies of diethyl malonate+ (1-butanol, 2-methyl-1-propanol, 1-pentanol, n-heptane, and ethyl acetate) at T= (288.2, 298.2, 313.2, 328.2, 338.2, and 348.2 K) and p = 101.3 kPa
rhol	805.98	kg/m3	298.15	Excess molar enthalpies of diethyl malonate+ (1-butanol, 2-methyl-1-propanol, 1-pentanol, n-heptane, and ethyl acetate) at T= (288.2, 298.2, 313.2, 328.2, 338.2, and 348.2 K) and p = 101.3 kPa
rhol	794.34	kg/m3	313.15	Excess molar enthalpies of diethyl malonate+ (1-butanol, 2-methyl-1-propanol, 1-pentanol, n-heptane, and ethyl acetate) at T= (288.2, 298.2, 313.2, 328.2, 338.2, and 348.2 K) and p = 101.3 kPa
rhol	782.32	kg/m3	328.15	Excess molar enthalpies of diethyl malonate+ (1-butanol, 2-methyl-1-propanol, 1-pentanol, n-heptane, and ethyl acetate) at T= (288.2, 298.2, 313.2, 328.2, 338.2, and 348.2 K) and p = 101.3 kPa

rhol	774.03	kg/m3	338.15	Excess molar enthalpies of diethyl malonate+ (1-butanol, 2-methyl-1-propanol, 1-pentanol, n-heptane, and ethyl acetate) at T= (288.2, 298.2, 313.2, 328.2, 338.2, and 348.2 K) and p = 101.3 kPa
rhol	765.48	kg/m3	348.15	Excess molar enthalpies of diethyl malonate+ (1-butanol, 2-methyl-1-propanol, 1-pentanol, n-heptane, and ethyl acetate) at T= (288.2, 298.2, 313.2, 328.2, 338.2, and 348.2 K) and p = 101.3 kPa
rhol	805.78	kg/m3	298.15	Phase behaviour of ionic liquid 1-butyl-1-methylpyrrolidinium tris(pentafluoroethyl)trifluorophosphate with alcohols, water and aromatic hydrocarbons
rhol	805.88	kg/m3	298.20	Liquid-liquid equilibrium data for ternary systems of water + lactic acid + C4-C7 alcohols at 298.2 K and atmospheric pressure
rhol	805.99	kg/m3	298.15	Thermodynamic and spectral investigations of binary liquid mixturesof 2-butoxy ethanol with alcohols at temperature range of 293.15-313.15 K
rhol	805.50	kg/m3	298.15	Solubility of androstenedione in lower alcohols
rhol	810.00	kg/m3	293.15	Liquid-liquid equilibrium of 1-butanol + water +tri-n-butyl phosphate + ammonium chloride system

rhol	807.96	kg/m3	298.20	Phase equilibria of (water + propionic acid or butyric acid + 2-methoxy-2-methylpropane) ternary systems at 298.2 K and 323.2 K
rhol	805.20	kg/m3	298.15	Liquid-liquid equilibria for the ternary system containing 1-Butanol + methoxy(methoxymethoxy)methane + water at temperatures of 303.15, 323.15 and 343.15 K
rhol	806.01	kg/m3	298.20	Isothermal vapor-liquid equilibria, excess molar volume and the deviation of refractive indices for binary mixtures of 1-butanol, 1-hexanol, 3-methyl-1-butanol and butyl acetate
rhol	806.27	kg/m3	298.15	Effect of composition and temperature variations on thermophysical properties of binary and ternary mixtures of 1-ethyl-3-methylimidazolium ethylsulfate with 1-butanol and /or methanol
rhol	802.42	kg/m3	303.15	Effect of composition and temperature variations on thermophysical properties of binary and ternary mixtures of 1-ethyl-3-methylimidazolium ethylsulfate with 1-butanol and /or methanol

rhol	798.54	kg/m3	308.15	Effect of composition and temperature variations on thermophysical properties of binary and ternary mixtures of 1-ethyl-3-methylimidazolium ethylsulfate with 1-butanol and /or methanol
rhol	794.62	kg/m3	313.15	Effect of composition and temperature variations on thermophysical properties of binary and ternary mixtures of 1-ethyl-3-methylimidazolium ethylsulfate with 1-butanol and /or methanol
rhol	790.65	kg/m3	318.15	Effect of composition and temperature variations on thermophysical properties of binary and ternary mixtures of 1-ethyl-3-methylimidazolium ethylsulfate with 1-butanol and /or methanol
rhol	786.64	kg/m3	323.15	Effect of composition and temperature variations on thermophysical properties of binary and ternary mixtures of 1-ethyl-3-methylimidazolium ethylsulfate with 1-butanol and /or methanol

rhol	805.98	kg/m3	298.15	A systematic study on volumetric and transport properties of binary systems 1-propanol + n-hexadecane, 1-butanol + n-hexadecane and 1-propanol + ethyl oleate at different temperatures: Experimental and modeling
rhol	802.13	kg/m3	303.15	A systematic study on volumetric and transport properties of binary systems 1-propanol + n-hexadecane, 1-butanol + n-hexadecane and 1-propanol + ethyl oleate at different temperatures: Experimental and modeling
rhol	798.25	kg/m3	308.15	A systematic study on volumetric and transport properties of binary systems 1-propanol + n-hexadecane, 1-butanol + n-hexadecane and 1-propanol + ethyl oleate at different temperatures: Experimental and modeling
rhol	805.77	kg/m3	298.15	Isobaric vapor-liquid equilibrium, density and speed of sound of binary mixtures 2,2,4-trimethylpentane + 1-butanol or dibutyl ether (DBE) at 101.3 kPa

rhol	809.38	kg/m3	293.20	Liquid-liquid equilibrium data and thermophysical properties for ternary systems composed of water, acetic acid and different solvents
rhol	793.78	kg/m3	313.20	Liquid-liquid equilibrium data and thermophysical properties for ternary systems composed of water, acetic acid and different solvents
rhol	805.82	kg/m3	298.15	Excess molar heat capacities for (1-butanol+1,3-butanediol) at temperatures from (285 to 353) K
rhol	805.80	kg/m3	298.15	Bubble point temperatures of the binary mixtures of nitrobenzene with C1 C4 aliphatic alcohols at 94.95 kPa
rhol	817.18	kg/m3	283.15	Excess molar volumes of binary mixtures of 1,3-dimethylimidazolidin-2-one with an alkan-1-ol at the temperatures 283.15 K, 298.15 K, and 313.15 K
rhol	805.79	kg/m3	298.15	Excess molar volumes of binary mixtures of 1,3-dimethylimidazolidin-2-one with an alkan-1-ol at the temperatures 283.15 K, 298.15 K, and 313.15 K
rhol	794.18	kg/m3	313.15	Excess molar volumes of binary mixtures of 1,3-dimethylimidazolidin-2-one with an alkan-1-ol at the temperatures 283.15 K, 298.15 K, and 313.15 K

rhol	809.52	kg/m3	293.15	Densities and volumetric properties of (N-(2-hydroxyethyl)morpholine + ethanol, + 1-propanol, + 2-propanol, + 1-butanol, and + 2-butanol) at (293.15, 298.15, 303.15, 313.15, and 323.15) K
rhol	805.88	kg/m3	298.15	Densities and volumetric properties of (N-(2-hydroxyethyl)morpholine + ethanol, + 1-propanol, + 2-propanol, + 1-butanol, and + 2-butanol) at (293.15, 298.15, 303.15, 313.15, and 323.15) K
rhol	801.99	kg/m3	303.15	Densities and volumetric properties of (N-(2-hydroxyethyl)morpholine + ethanol, + 1-propanol, + 2-propanol, + 1-butanol, and + 2-butanol) at (293.15, 298.15, 303.15, 313.15, and 323.15) K
rhol	794.34	kg/m3	313.15	Densities and volumetric properties of (N-(2-hydroxyethyl)morpholine + ethanol, + 1-propanol, + 2-propanol, + 1-butanol, and + 2-butanol) at (293.15, 298.15, 303.15, 313.15, and 323.15) K
rhol	786.67	kg/m3	323.15	Densities and volumetric properties of (N-(2-hydroxyethyl)morpholine + ethanol, + 1-propanol, + 2-propanol, + 1-butanol, and + 2-butanol) at (293.15, 298.15, 303.15, 313.15, and 323.15) K

rhol	805.64	kg/m3	298.15	Apparent molar volumes and compressibilities of tetrabutyl-ammonium bromide in organic solvents
rhol	805.60	kg/m3	298.15	Volumetric and surface properties of pure ionic liquid n-octyl-pyridinium nitrate and its binary mixture with alcohol
rhol	805.77	kg/m3	298.15	Effect of temperature and composition on the density, viscosity, surface tension, and thermodynamic properties of binary mixtures of N-octylisoquinolinium bis{(trifluoromethyl)sulfonyl}imide with alcohols
rhol	798.06	kg/m3	308.15	Effect of temperature and composition on the density, viscosity, surface tension, and thermodynamic properties of binary mixtures of N-octylisoquinolinium bis{(trifluoromethyl)sulfonyl}imide with alcohols
rhol	790.19	kg/m3	318.15	Effect of temperature and composition on the density, viscosity, surface tension, and thermodynamic properties of binary mixtures of N-octylisoquinolinium bis{(trifluoromethyl)sulfonyl}imide with alcohols

rhol	782.15	kg/m3	328.15	Effect of temperature and composition on the density, viscosity, surface tension, and thermodynamic properties of binary mixtures of N-octylisoquinolinium bis{(trifluoromethyl)sulfonyl}imide with alcohols
rhol	773.89	kg/m3	338.15	Effect of temperature and composition on the density, viscosity, surface tension, and thermodynamic properties of binary mixtures of N-octylisoquinolinium bis{(trifluoromethyl)sulfonyl}imide with alcohols
rhol	798.06	kg/m3	298.15	Extraction of butan-1-ol from water with ionic liquids at T = 308.15 K
rhol	805.85	kg/m3	298.15	Isothermal (vapour + liquid) equilibria for binary mixtures of diisopropyl ether with (methanol, or ethanol, or 1-butanol): Experimental data, correlations, and predictions
rhol	809.89	kg/m3	293.15	Thermodynamics of 1-alkanol + linear polyether mixtures
rhol	806.47	kg/m3	298.15	Thermodynamics of 1-alkanol + linear polyether mixtures
rhol	802.22	kg/m3	303.15	Thermodynamics of 1-alkanol + linear polyether mixtures

rhol	809.54	kg/m3	293.15	Osmotic and apparent molar properties of binary mixtures alcohol + 1-butyl-3-methylimidazolium trifluoromethanesulfonate ionic liquid
rhol	798.01	kg/m3	308.15	Osmotic and apparent molar properties of binary mixtures alcohol + 1-butyl-3-methylimidazolium trifluoromethanesulfonate ionic liquid
rhol	786.15	kg/m3	323.15	Osmotic and apparent molar properties of binary mixtures alcohol + 1-butyl-3-methylimidazolium trifluoromethanesulfonate ionic liquid
rhol	805.70	kg/m3	298.15	Density and surface tension of pure ionic liquid 1-butyl-3-methylimidazolium L-lactate and its binary mixture with alcohol and water
rhol	798.00	kg/m3	308.15	Density and surface tension of pure ionic liquid 1-butyl-3-methylimidazolium L-lactate and its binary mixture with alcohol and water
rhol	790.10	kg/m3	318.15	Density and surface tension of pure ionic liquid 1-butyl-3-methylimidazolium L-lactate and its binary mixture with alcohol and water

rhol	805.95	kg/m3	298.15	FT-IR studies on excess thermodynamic properties of binary liquid mixtures o-chlorotoluene with 1-propanol, 1-butanol, 1-pentanol, 1-hexanol and 1-heptanol at different temperatures
rhol	802.11	kg/m3	303.15	FT-IR studies on excess thermodynamic properties of binary liquid mixtures o-chlorotoluene with 1-propanol, 1-butanol, 1-pentanol, 1-hexanol and 1-heptanol at different temperatures
rhol	798.24	kg/m3	308.15	FT-IR studies on excess thermodynamic properties of binary liquid mixtures o-chlorotoluene with 1-propanol, 1-butanol, 1-pentanol, 1-hexanol and 1-heptanol at different temperatures
rhol	804.00	kg/m3	303.15	Study of molecular interactions in the mixtures of some primary alcohols with equimolar mixture of 1-propanol and alkylbenzoates at T = 303.15 K
rhol	786.32	kg/m3	323.15	Osmotic coefficients and apparent molar volumes of 1-hexyl-3-methylimidazolium trifluoromethanesulfonate ionic liquid in alcohols

rhol	809.81	kg/m3	293.15	Molecular interactions in binary mixtures of 1-butoxy-2-propanol with alcohols at different temperatures: A thermophysical and spectroscopic approach
rhol	805.99	kg/m3	298.15	Molecular interactions in binary mixtures of 1-butoxy-2-propanol with alcohols at different temperatures: A thermophysical and spectroscopic approach
rhol	802.14	kg/m3	303.15	Molecular interactions in binary mixtures of 1-butoxy-2-propanol with alcohols at different temperatures: A thermophysical and spectroscopic approach
rhol	798.26	kg/m3	308.15	Molecular interactions in binary mixtures of 1-butoxy-2-propanol with alcohols at different temperatures: A thermophysical and spectroscopic approach
rhol	794.38	kg/m3	313.15	Molecular interactions in binary mixtures of 1-butoxy-2-propanol with alcohols at different temperatures: A thermophysical and spectroscopic approach

rhol	805.79	kg/m3	298.15	A combined experimental and computational investigation of excess molar enthalpies of (nitrobenzene + alkanol) mixtures
rhol	805.50	kg/m3	298.15	Solubility and solution thermodynamics of sorbic acid in eight pure organic solvents
rhol	809.80	kg/m3	293.15	Isothermal (vapour + liquid) equilibrium and thermophysical properties for (1-butyl-3-methylimidazolium iodide + 1-butanol) binary system
rhol	805.90	kg/m3	298.15	Isothermal (vapour + liquid) equilibrium and thermophysical properties for (1-butyl-3-methylimidazolium iodide + 1-butanol) binary system
rhol	798.20	kg/m3	308.15	Isothermal (vapour + liquid) equilibrium and thermophysical properties for (1-butyl-3-methylimidazolium iodide + 1-butanol) binary system
rhol	809.72	kg/m3	293.15	Properties of pure 1,1,3,3-tetramethylguanidine imidazole ionic liquid and its binary mixtures with alcohols at T = (293.15 to 313.15) K
rhol	805.92	kg/m3	298.15	Properties of pure 1,1,3,3-tetramethylguanidine imidazole ionic liquid and its binary mixtures with alcohols at T = (293.15 to 313.15) K

rhol	802.08	kg/m3	303.15	Properties of pure 1,1,3,3-tetramethylguanidine imidazole ionic liquid and its binary mixtures with alcohols at T = (293.15 to 313.15) K
rhol	798.21	kg/m3	308.15	Properties of pure 1,1,3,3-tetramethylguanidine imidazole ionic liquid and its binary mixtures with alcohols at T = (293.15 to 313.15) K
rhol	794.30	kg/m3	313.15	Properties of pure 1,1,3,3-tetramethylguanidine imidazole ionic liquid and its binary mixtures with alcohols at T = (293.15 to 313.15) K
rhol	805.79	kg/m3	298.15	Measurements and equation-of-state modelling of thermodynamic properties of binary mixtures of 1-butyl-1-methylpyrrolidinium tetracyanoborate ionic liquid with molecular compounds
rhol	806.10	kg/m3	298.15	Solubility and solution thermodynamics of thymol in six pure organic solvents
rhol	817.10	kg/m3	283.15	Ionic association and conductance of [emim][BF4] and [bmim][BF4] in 1-butanol in a wide range of temperature
rhol	813.34	kg/m3	288.15	Ionic association and conductance of [emim][BF4] and [bmim][BF4] in 1-butanol in a wide range of temperature

rhol	809.55	kg/m3	293.15	Ionic association and conductance of [emim][BF4] and [bmim][BF4] in 1-butanol in a wide range of temperature
rhol	805.74	kg/m3	298.15	Ionic association and conductance of [emim][BF4] and [bmim][BF4] in 1-butanol in a wide range of temperature
rhol	801.90	kg/m3	303.15	Ionic association and conductance of [emim][BF4] and [bmim][BF4] in 1-butanol in a wide range of temperature
rhol	798.03	kg/m3	308.15	Ionic association and conductance of [emim][BF4] and [bmim][BF4] in 1-butanol in a wide range of temperature
rhol	794.13	kg/m3	313.15	Ionic association and conductance of [emim][BF4] and [bmim][BF4] in 1-butanol in a wide range of temperature
rhol	790.18	kg/m3	318.15	Ionic association and conductance of [emim][BF4] and [bmim][BF4] in 1-butanol in a wide range of temperature
rhol	809.75	kg/m3	293.15	(Liquid + liquid) equilibria of four alcohol-water systems containing 1,8-cineole at T = 298.15 K
rhol	809.82	kg/m3	293.15	Thermodynamic and spectroscopic interpretation of molecular interactions of nicotine + alcohol binary mixtures

rhol	806.00	kg/m3	298.15	Thermodynamic and spectroscopic interpretation of molecular interactions of nicotine + alcohol binary mixtures	
rhol	802.15	kg/m3	303.15	Thermodynamic and spectroscopic interpretation of molecular interactions of nicotine + alcohol binary mixtures	
rhol	798.28	kg/m3	308.15	Thermodynamic and spectroscopic interpretation of molecular interactions of nicotine + alcohol binary mixtures	
rhol	794.37	kg/m3	313.15	Thermodynamic and spectroscopic interpretation of molecular interactions of nicotine + alcohol binary mixtures	
rhol	790.43	kg/m3	318.15	Thermodynamic and spectroscopic interpretation of molecular interactions of nicotine + alcohol binary mixtures	
rhol	786.44	kg/m3	323.15	Thermodynamic and spectroscopic interpretation of molecular interactions of nicotine + alcohol binary mixtures	
rhol	778.01	kg/m3	333.15	Densities and Viscosities of Ternary System n-Dodecane (1) + Bicyclohexyl (2) + n-Butanol (3) and Corresponding Binaries at T = (293.15 to 333.15) K	

rhol	809.52	kg/m3	293.15	Thermodynamic and spectroscopic properties of binary mixtures of n-butylammonium butanoate ionic liquid with alcohols at T = (293.15-313.15) K	
rhol	805.71	kg/m3	298.15	Thermodynamic and spectroscopic properties of binary mixtures of n-butylammonium butanoate ionic liquid with alcohols at T = (293.15-313.15) K	
rhol	801.87	kg/m3	303.15	Thermodynamic and spectroscopic properties of binary mixtures of n-butylammonium butanoate ionic liquid with alcohols at T = (293.15-313.15) K	
rhol	798.00	kg/m3	308.15	Thermodynamic and spectroscopic properties of binary mixtures of n-butylammonium butanoate ionic liquid with alcohols at T = (293.15-313.15) K	
rhol	794.10	kg/m3	313.15	Thermodynamic and spectroscopic properties of binary mixtures of n-butylammonium butanoate ionic liquid with alcohols at T = (293.15-313.15) K	

rhol	809.49	kg/m3	293.15	Investigation on molecular interactions of antibiotics in alcohols using volumetric and acoustic studies at different temperatures	
rhol	805.65	kg/m3	298.15	Investigation on molecular interactions of antibiotics in alcohols using volumetric and acoustic studies at different temperatures	
rhol	803.68	kg/m3	303.15	Investigation on molecular interactions of antibiotics in alcohols using volumetric and acoustic studies at different temperatures	
rhol	800.50	kg/m3	308.15	Investigation on molecular interactions of antibiotics in alcohols using volumetric and acoustic studies at different temperatures	
rhol	794.11	kg/m3	313.15	Investigation on molecular interactions of antibiotics in alcohols using volumetric and acoustic studies at different temperatures	
rhol	806.00	kg/m3	298.00	Comparative study of physical properties of binary mixtures of halogen free ionic liquids with alcohols	
rhol	802.00	kg/m3	303.00	Comparative study of physical properties of binary mixtures of halogen free ionic liquids with alcohols	

rhol	798.00	kg/m3	308.00	Comparative study of physical properties of binary mixtures of halogen free ionic liquids with alcohols	
rhol	794.00	kg/m3	313.00	Comparative study of physical properties of binary mixtures of halogen free ionic liquids with alcohols	
rhol	805.81	kg/m3	298.20	Liquid-liquid equilibrium in systems used for the production of 5-hydroxymethylfurfural from biomass using alcohols as solvents	
rhol	809.50	kg/m3	293.15	Measurement and correlation of isobaric vapour-liquid equilibria for binary systems of 1-(methoxymethoxy)-2-methyl-propane with 1-butanol and isobutanol at 101.33 kPa	
rhol	817.11	kg/m3	283.15	Excess molar enthalpies of R-fenchone + butan-1-ol or + pentan-1-ol. Modeling with COSMO-RS and UNIFAC	
rhol	805.72	kg/m3	298.15	Excess molar enthalpies of R-fenchone + butan-1-ol or + pentan-1-ol. Modeling with COSMO-RS and UNIFAC	
rhol	794.40	kg/m3	313.15	Excess molar enthalpies of R-fenchone + butan-1-ol or + pentan-1-ol. Modeling with COSMO-RS and UNIFAC	

rhol	782.11	kg/m3	328.15	Excess molar enthalpies of R-fenchone + butan-1-ol or + pentan-1-ol. Modeling with COSMO-RS and UNIFAC
rhol	809.53	kg/m3	293.15	Intermolecular interactions in binary mixtures of 2-diethylethanolamine with 1-propanol and 1-butanol at different temperatures
rhol	802.00	kg/m3	303.15	Intermolecular interactions in binary mixtures of 2-diethylethanolamine with 1-propanol and 1-butanol at different temperatures
rhol	794.70	kg/m3	313.15	Intermolecular interactions in binary mixtures of 2-diethylethanolamine with 1-propanol and 1-butanol at different temperatures
rhol	810.00	kg/m3	293.15	Density, speed of sound and refractive index of mixtures containing 2-phenoxyethanol with propanol or butanol at various temperatures
rhol	802.00	kg/m3	303.15	Density, speed of sound and refractive index of mixtures containing 2-phenoxyethanol with propanol or butanol at various temperatures

rhol	794.00	kg/m3	313.15	Density, speed of sound and refractive index of mixtures containing 2-phenoxyethanol with propanol or butanol at various temperatures
rhol	787.00	kg/m3	323.15	Density, speed of sound and refractive index of mixtures containing 2-phenoxyethanol with propanol or butanol at various temperatures
rhol	806.00	kg/m3	298.15	Density, speed of sound and refractive index of mixtures containing 2-phenoxyethanol with propanol or butanol at various temperatures
rhol	809.77	kg/m3	293.15	Volumetric properties of monoethanolamine and alcohol binary mixtures at different temperatures and 0.1 MPa
rhol	805.92	kg/m3	298.15	Volumetric properties of monoethanolamine and alcohol binary mixtures at different temperatures and 0.1 MPa
rhol	802.00	kg/m3	303.15	Volumetric properties of monoethanolamine and alcohol binary mixtures at different temperatures and 0.1 MPa
rhol	798.14	kg/m3	308.15	Volumetric properties of monoethanolamine and alcohol binary mixtures at different temperatures and 0.1 MPa

rhol	794.24	kg/m3	313.15	Volumetric properties of monoethanolamine and alcohol binary mixtures at different temperatures and 0.1 MPa
rhol	790.20	kg/m3	318.15	Volumetric properties of monoethanolamine and alcohol binary mixtures at different temperatures and 0.1 MPa
rhol	786.21	kg/m3	323.15	Volumetric properties of monoethanolamine and alcohol binary mixtures at different temperatures and 0.1 MPa
rhol	817.18	kg/m3	283.15	Volumetric and sound speed study of aqueous 1-butanol liquid mixtures at different temperatures
rhol	813.40	kg/m3	288.15	Volumetric and sound speed study of aqueous 1-butanol liquid mixtures at different temperatures
rhol	809.60	kg/m3	293.15	Volumetric and sound speed study of aqueous 1-butanol liquid mixtures at different temperatures
rhol	805.78	kg/m3	298.15	Volumetric and sound speed study of aqueous 1-butanol liquid mixtures at different temperatures
rhol	801.94	kg/m3	303.15	Volumetric and sound speed study of aqueous 1-butanol liquid mixtures at different temperatures

rhol	801.91	kg/m3	303.15	Vapor-liquid equilibrium and excess properties of the binary mixtures formed by ethyl isobutyrate and n-alkanols	
rhol	786.62	kg/m3	323.15	Vapor-liquid equilibrium and excess properties of the binary mixtures formed by ethyl isobutyrate and n-alkanols	
rhol	805.75	kg/m3	298.15	Excess molar enthalpies of methyl isobutyl ketone (MIBK) with alkan-1-ols (C1-C6) and their correlations at 298.15 K	
rhol	801.98	kg/m3	303.15	Excess molar enthalpies and heat capacities of dimethyl sulfoxide + seven normal alkanols at 303.15K and atmospheric pressure	
rhol	801.92	kg/m3	303.15	Studies of viscosities of dilute solutions of alkylamines in non-electrolyte solvents III. Alkylamines in butanols 303.15K	
rhol	801.90	kg/m3	303.15	Volumetric and transport properties of ternary mixtures containing 1-alkanol + ethyl ethanoate + cyclohexane at 303.15 K: Experimental data, correlation and prediction by ERAS model	
rhol	809.50	kg/m3	298.15	Topological investigations of the molecular species and molecular interactions that characterize pyrrolidin-2-one + lower alkanol mixtures	

rhol	813.37	kg/m3	288.15	Experimental determination and modelling of densities and excess molar volumes of ternary system (1-butanol + cyclohexylamine + n-heptane) and corresponding binaries from 288.15 to 323.15K	
rhol	809.57	kg/m3	293.15	Experimental determination and modelling of densities and excess molar volumes of ternary system (1-butanol + cyclohexylamine + n-heptane) and corresponding binaries from 288.15 to 323.15K	
rhol	805.76	kg/m3	298.15	Experimental determination and modelling of densities and excess molar volumes of ternary system (1-butanol + cyclohexylamine + n-heptane) and corresponding binaries from 288.15 to 323.15K	
rhol	801.92	kg/m3	303.15	Experimental determination and modelling of densities and excess molar volumes of ternary system (1-butanol + cyclohexylamine + n-heptane) and corresponding binaries from 288.15 to 323.15K	

rhol	798.05	kg/m3	308.15	Experimental determination and modelling of densities and excess molar volumes of ternary system (1-butanol + cyclohexylamine + n-heptane) and corresponding binaries from 288.15 to 323.15K	
rhol	794.15	kg/m3	313.15	Experimental determination and modelling of densities and excess molar volumes of ternary system (1-butanol + cyclohexylamine + n-heptane) and corresponding binaries from 288.15 to 323.15K	
rhol	790.23	kg/m3	318.15	Experimental determination and modelling of densities and excess molar volumes of ternary system (1-butanol + cyclohexylamine + n-heptane) and corresponding binaries from 288.15 to 323.15K	
rhol	786.24	kg/m3	323.15	Experimental determination and modelling of densities and excess molar volumes of ternary system (1-butanol + cyclohexylamine + n-heptane) and corresponding binaries from 288.15 to 323.15K	
rhol	809.78	kg/m3	293.15	Excess molar volumes of Diisopropylamine + (C1-C5) Alkan-1-ols: application of the ERAS model and cubic EOS	

rhol	806.05	kg/m3	298.15	Excess molar volumes of Diisopropylamine + (C1-C5) Alkan-1-ols: application of the ERAS model and cubic EOS	
rhol	802.03	kg/m3	303.15	Excess molar volumes of Diisopropylamine + (C1-C5) Alkan-1-ols: application of the ERAS model and cubic EOS	
rhol	794.12	kg/m3	313.15	Excess molar volumes of Diisopropylamine + (C1-C5) Alkan-1-ols: application of the ERAS model and cubic EOS	
rhol	805.90	kg/m3	298.15	Thermodynamics of mixtures containing amines. XVI. CE pm of 1-butanol, 1-octanol or 1-decanol + benzylamine systems at (298.15, 308.15, 318.15 and 333.15) K	
rhol	817.21	kg/m3	283.15	Effects of alkyl group and temperature on the interactions between furfural and alcohol: Insight from density and sound velocity studies	
rhol	809.65	kg/m3	293.15	Effects of alkyl group and temperature on the interactions between furfural and alcohol: Insight from density and sound velocity studies	

rhol	802.00	kg/m3	303.15	Effects of alkyl group and temperature on the interactions between furfural and alcohol: Insight from density and sound velocity studies	
rhol	794.21	kg/m3	313.15	Effects of alkyl group and temperature on the interactions between furfural and alcohol: Insight from density and sound velocity studies	
rhol	805.76	kg/m3	298.15	Experimental and theoretical excess molar properties of imidazolium based ionic liquids with isomers of butanol	
rhol	798.05	kg/m3	308.15	Experimental and theoretical excess molar properties of imidazolium based ionic liquids with isomers of butanol	
rhol	790.21	kg/m3	318.15	Experimental and theoretical excess molar properties of imidazolium based ionic liquids with isomers of butanol	
rhol	809.78	kg/m3	293.15	Density, Speed of Sound, Viscosity, Excess Properties, and Prigogine Flory Patterson (PFP) Theory of Binary Mixtures of Amine and Alcohols	

rhol	805.95	kg/m3	298.15	Density, Speed of Sound, Viscosity, Excess Properties, and Prigogine Flory Patterson (PFP) Theory of Binary Mixtures of Amine and Alcohols	
rhol	802.10	kg/m3	303.15	Density, Speed of Sound, Viscosity, Excess Properties, and Prigogine Flory Patterson (PFP) Theory of Binary Mixtures of Amine and Alcohols	
rhol	798.23	kg/m3	308.15	Density, Speed of Sound, Viscosity, Excess Properties, and Prigogine Flory Patterson (PFP) Theory of Binary Mixtures of Amine and Alcohols	
rhol	794.31	kg/m3	313.15	Density, Speed of Sound, Viscosity, Excess Properties, and Prigogine Flory Patterson (PFP) Theory of Binary Mixtures of Amine and Alcohols	
rhol	802.01	kg/m3	303.15	Density, Refractive Index, and Sound Velocity for the Binary Mixtures of Tri-n-Butyl Phosphate and n-Butanol between 303.15 K and 323.15 K	
rhol	798.14	kg/m3	308.15	Density, Refractive Index, and Sound Velocity for the Binary Mixtures of Tri-n-Butyl Phosphate and n-Butanol between 303.15 K and 323.15 K	

rhol	794.22	kg/m3	313.15	Density, Refractive Index, and Sound Velocity for the Binary Mixtures of Tri-n-Butyl Phosphate and n-Butanol between 303.15 K and 323.15 K	
rhol	790.28	kg/m3	318.15	Density, Refractive Index, and Sound Velocity for the Binary Mixtures of Tri-n-Butyl Phosphate and n-Butanol between 303.15 K and 323.15 K	
rhol	786.29	kg/m3	323.15	Density, Refractive Index, and Sound Velocity for the Binary Mixtures of Tri-n-Butyl Phosphate and n-Butanol between 303.15 K and 323.15 K	
rhol	809.84	kg/m3	293.15	Density and Excess Molar Volumes of 1-Butanol + Methanol + Electrolyte Systems in the Temperature Range (293.15 308.15) K	
rhol	806.03	kg/m3	298.15	Density and Excess Molar Volumes of 1-Butanol + Methanol + Electrolyte Systems in the Temperature Range (293.15 308.15) K	
rhol	802.19	kg/m3	303.15	Density and Excess Molar Volumes of 1-Butanol + Methanol + Electrolyte Systems in the Temperature Range (293.15 308.15) K	

rhol	798.38	kg/m3	308.15	Density and Excess Molar Volumes of 1-Butanol + Methanol + Electrolyte Systems in the Temperature Range (293.15 308.15) K	
rhol	807.96	kg/m3	298.15	Tie-Line Data for Aqueous Mixtures of Butyric Acid with Diisopropyl Ether at Various Temperatures	
rhol	806.00	kg/m3	298.15	Thermodynamics of 1,3-dimethylurea in eight alcohols	
rhol	817.07	kg/m3	283.15	Densities and Viscosities for Binary Liquid Mixtures of n-Undecane + 1-Propanol, + 1-Butanol, + 1-Pentanol, and + 1-Hexanol from 283.15 to 363.15 K at 0.1 MPa	
rhol	813.32	kg/m3	288.15	Densities and Viscosities for Binary Liquid Mixtures of n-Undecane + 1-Propanol, + 1-Butanol, + 1-Pentanol, and + 1-Hexanol from 283.15 to 363.15 K at 0.1 MPa	
rhol	809.55	kg/m3	293.15	Densities and Viscosities for Binary Liquid Mixtures of n-Undecane + 1-Propanol, + 1-Butanol, + 1-Pentanol, and + 1-Hexanol from 283.15 to 363.15 K at 0.1 MPa	

rhol	805.76	kg/m3	298.15	Densities and Viscosities for Binary Liquid Mixtures of n-Undecane + 1-Propanol, + 1-Butanol, + 1-Pentanol, and + 1-Hexanol from 283.15 to 363.15 K at 0.1 MPa
rhol	801.94	kg/m3	303.15	Densities and Viscosities for Binary Liquid Mixtures of n-Undecane + 1-Propanol, + 1-Butanol, + 1-Pentanol, and + 1-Hexanol from 283.15 to 363.15 K at 0.1 MPa
rhol	798.09	kg/m3	308.15	Densities and Viscosities for Binary Liquid Mixtures of n-Undecane + 1-Propanol, + 1-Butanol, + 1-Pentanol, and + 1-Hexanol from 283.15 to 363.15 K at 0.1 MPa
rhol	794.20	kg/m3	313.15	Densities and Viscosities for Binary Liquid Mixtures of n-Undecane + 1-Propanol, + 1-Butanol, + 1-Pentanol, and + 1-Hexanol from 283.15 to 363.15 K at 0.1 MPa
rhol	790.27	kg/m3	318.15	Densities and Viscosities for Binary Liquid Mixtures of n-Undecane + 1-Propanol, + 1-Butanol, + 1-Pentanol, and + 1-Hexanol from 283.15 to 363.15 K at 0.1 MPa

rhol	786.30	kg/m3	323.15	Densities and Viscosities for Binary Liquid Mixtures of n-Undecane + 1-Propanol, + 1-Butanol, + 1-Pentanol, and + 1-Hexanol from 283.15 to 363.15 K at 0.1 MPa
rhol	782.27	kg/m3	328.15	Densities and Viscosities for Binary Liquid Mixtures of n-Undecane + 1-Propanol, + 1-Butanol, + 1-Pentanol, and + 1-Hexanol from 283.15 to 363.15 K at 0.1 MPa
rhol	778.18	kg/m3	333.15	Densities and Viscosities for Binary Liquid Mixtures of n-Undecane + 1-Propanol, + 1-Butanol, + 1-Pentanol, and + 1-Hexanol from 283.15 to 363.15 K at 0.1 MPa
rhol	774.03	kg/m3	338.15	Densities and Viscosities for Binary Liquid Mixtures of n-Undecane + 1-Propanol, + 1-Butanol, + 1-Pentanol, and + 1-Hexanol from 283.15 to 363.15 K at 0.1 MPa
rhol	769.80	kg/m3	343.15	Densities and Viscosities for Binary Liquid Mixtures of n-Undecane + 1-Propanol, + 1-Butanol, + 1-Pentanol, and + 1-Hexanol from 283.15 to 363.15 K at 0.1 MPa

rhol	765.49	kg/m3	348.15	Densities and Viscosities for Binary Liquid Mixtures of n-Undecane + 1-Propanol, + 1-Butanol, + 1-Pentanol, and + 1-Hexanol from 283.15 to 363.15 K at 0.1 MPa
rhol	761.12	kg/m3	353.15	Densities and Viscosities for Binary Liquid Mixtures of n-Undecane + 1-Propanol, + 1-Butanol, + 1-Pentanol, and + 1-Hexanol from 283.15 to 363.15 K at 0.1 MPa
rhol	756.64	kg/m3	358.15	Densities and Viscosities for Binary Liquid Mixtures of n-Undecane + 1-Propanol, + 1-Butanol, + 1-Pentanol, and + 1-Hexanol from 283.15 to 363.15 K at 0.1 MPa
rhol	752.08	kg/m3	363.15	Densities and Viscosities for Binary Liquid Mixtures of n-Undecane + 1-Propanol, + 1-Butanol, + 1-Pentanol, and + 1-Hexanol from 283.15 to 363.15 K at 0.1 MPa
rhol	805.94	kg/m3	298.15	Isobaric Vapor Liquid Equilibrium for Two Binary Systems (Methanol + Dibutyl Carbonate) and (n-Butanol + Dibutyl Carbonate) at p = 40.0, 70.0, and 100.0 kPa

rhol	805.75	kg/m3	298.15	Physical Properties of the Pure Deep Eutectic Solvent, [ChCl]:[Lev] (1:2) DES, and Its Binary Mixtures with Alcohols
rhol	798.06	kg/m3	308.15	Physical Properties of the Pure Deep Eutectic Solvent, [ChCl]:[Lev] (1:2) DES, and Its Binary Mixtures with Alcohols
rhol	790.22	kg/m3	318.15	Physical Properties of the Pure Deep Eutectic Solvent, [ChCl]:[Lev] (1:2) DES, and Its Binary Mixtures with Alcohols
rhol	801.92	kg/m3	303.15	Isobaric Vapor Liquid Equilibrium for the Binary Systems Dimethyl Disulfide + C1 C4 n-Alkanol at 40.000 and 101.325 kPa
rhol	809.50	kg/m3	298.15	Measurement and Correlation of Isobaric Vapor Liquid Equilibrium for Binary Systems of 1-(Ethoxymethoxy)-2-methyl-propane with Isobutyl Alcohol or 1-Butanol at 101.33 kPa
rhol	809.01	kg/m3	293.15	Densities and Viscosities for Binary Liquid Mixtures of Biodiesel + 1-Butanol, + Isobutyl Alcohol, or + 2-Butanol from 293.15 to 333.15 K at 0.1 MPa

rhol	805.87	kg/m3	298.15	Densities and Viscosities for Binary Liquid Mixtures of Biodiesel + 1-Butanol, + Isobutyl Alcohol, or + 2-Butanol from 293.15 to 333.15 K at 0.1 MPa
rhol	802.06	kg/m3	303.15	Densities and Viscosities for Binary Liquid Mixtures of Biodiesel + 1-Butanol, + Isobutyl Alcohol, or + 2-Butanol from 293.15 to 333.15 K at 0.1 MPa
rhol	798.18	kg/m3	308.15	Densities and Viscosities for Binary Liquid Mixtures of Biodiesel + 1-Butanol, + Isobutyl Alcohol, or + 2-Butanol from 293.15 to 333.15 K at 0.1 MPa
rhol	794.26	kg/m3	313.15	Densities and Viscosities for Binary Liquid Mixtures of Biodiesel + 1-Butanol, + Isobutyl Alcohol, or + 2-Butanol from 293.15 to 333.15 K at 0.1 MPa
rhol	790.30	kg/m3	318.15	Densities and Viscosities for Binary Liquid Mixtures of Biodiesel + 1-Butanol, + Isobutyl Alcohol, or + 2-Butanol from 293.15 to 333.15 K at 0.1 MPa

rhol	786.29	kg/m3	323.15	Densities and Viscosities for Binary Liquid Mixtures of Biodiesel + 1-Butanol, + Isobutyl Alcohol, or + 2-Butanol from 293.15 to 333.15 K at 0.1 MPa
rhol	782.02	kg/m3	328.15	Densities and Viscosities for Binary Liquid Mixtures of Biodiesel + 1-Butanol, + Isobutyl Alcohol, or + 2-Butanol from 293.15 to 333.15 K at 0.1 MPa
rhol	777.81	kg/m3	333.15	Densities and Viscosities for Binary Liquid Mixtures of Biodiesel + 1-Butanol, + Isobutyl Alcohol, or + 2-Butanol from 293.15 to 333.15 K at 0.1 MPa
rhol	805.98	kg/m3	298.15	Densities and Viscosities for Binary Liquid Mixtures of Butan-1-ol + Propane-1,2-diol, + Butane-1,2-diol and 2-Methylpropan-1-ol + Propane-1,2-diol, + Butane-1,2-diol from 298.15 to 333.15 K at 0.1 MPa
rhol	802.16	kg/m3	303.15	Densities and Viscosities for Binary Liquid Mixtures of Butan-1-ol + Propane-1,2-diol, + Butane-1,2-diol and 2-Methylpropan-1-ol + Propane-1,2-diol, + Butane-1,2-diol from 298.15 to 333.15 K at 0.1 MPa

rhol	798.28	kg/m3	308.15	Densities and Viscosities for Binary Liquid Mixtures of Butan-1-ol + Propane-1,2-diol, + Butane-1,2-diol and 2-Methylpropan-1-ol + Propane-1,2-diol, + Butane-1,2-diol from 298.15 to 333.15 K at 0.1 MPa
rhol	794.37	kg/m3	313.15	Densities and Viscosities for Binary Liquid Mixtures of Butan-1-ol + Propane-1,2-diol, + Butane-1,2-diol and 2-Methylpropan-1-ol + Propane-1,2-diol, + Butane-1,2-diol from 298.15 to 333.15 K at 0.1 MPa
rhol	790.42	kg/m3	318.15	Densities and Viscosities for Binary Liquid Mixtures of Butan-1-ol + Propane-1,2-diol, + Butane-1,2-diol and 2-Methylpropan-1-ol + Propane-1,2-diol, + Butane-1,2-diol from 298.15 to 333.15 K at 0.1 MPa
rhol	786.41	kg/m3	323.15	Densities and Viscosities for Binary Liquid Mixtures of Butan-1-ol + Propane-1,2-diol, + Butane-1,2-diol and 2-Methylpropan-1-ol + Propane-1,2-diol, + Butane-1,2-diol from 298.15 to 333.15 K at 0.1 MPa

rhol	782.25	kg/m3	328.15	Densities and Viscosities for Binary Liquid Mixtures of Butan-1-ol + Propane-1,2-diol, + Butane-1,2-diol and 2-Methylpropan-1-ol + Propane-1,2-diol, + Butane-1,2-diol from 298.15 to 333.15 K at 0.1 MPa
rhol	778.09	kg/m3	333.15	Densities and Viscosities for Binary Liquid Mixtures of Butan-1-ol + Propane-1,2-diol, + Butane-1,2-diol and 2-Methylpropan-1-ol + Propane-1,2-diol, + Butane-1,2-diol from 298.15 to 333.15 K at 0.1 MPa
rhol	809.50	kg/m3	293.15	Three Binary Azeotropic Systems for 1-(Methoxymethoxy)-propane, 1-(Ethoxymethoxy)-propane, and Methoxy(methoxymethoxy)methane with Three Alcohols at 101.33 kPa: Experimental Data, Correlation, and Purification
rhol	805.88	kg/m3	298.15	Liquid-Liquid Equilibrium for Ternary Systems, Water + 5-Hydroxymethylfurfural + (1-Butanol, Isobutanol, Methyl Isobutyl Ketone), at 313.15, 323.15, and 333.15 K
rhol	809.50	kg/m3	293.15	Densities and Viscosities of Ternary System n-Dodecane (1) + Bicyclohexyl (2) + n-Butanol (3) and Corresponding Binaries at T = (293.15 to 333.15) K

rhol	805.70	kg/m3	298.15	Densities and Viscosities of Ternary System n-Dodecane (1) + Bicyclohexyl (2) + n-Butanol (3) and Corresponding Binaries at T = (293.15 to 333.15) K
rhol	801.87	kg/m3	303.15	Densities and Viscosities of Ternary System n-Dodecane (1) + Bicyclohexyl (2) + n-Butanol (3) and Corresponding Binaries at T = (293.15 to 333.15) K
rhol	798.00	kg/m3	308.15	Densities and Viscosities of Ternary System n-Dodecane (1) + Bicyclohexyl (2) + n-Butanol (3) and Corresponding Binaries at T = (293.15 to 333.15) K
rhol	794.10	kg/m3	313.15	Densities and Viscosities of Ternary System n-Dodecane (1) + Bicyclohexyl (2) + n-Butanol (3) and Corresponding Binaries at T = (293.15 to 333.15) K
rhol	790.15	kg/m3	318.15	Densities and Viscosities of Ternary System n-Dodecane (1) + Bicyclohexyl (2) + n-Butanol (3) and Corresponding Binaries at T = (293.15 to 333.15) K
rhol	786.16	kg/m3	323.15	Densities and Viscosities of Ternary System n-Dodecane (1) + Bicyclohexyl (2) + n-Butanol (3) and Corresponding Binaries at T = (293.15 to 333.15) K

rhol	782.12	kg/m3	328.15	Densities and Viscosities of Ternary System n-Dodecane (1) + Bicyclohexyl (2) + n-Butanol (3) and Corresponding Binaries at T = (293.15 to 333.15) K
rhol	805.80	kg/m3	298.15	Experimental study on the calorimetric data of 2-butoxyethanol with aliphatic alcohols (C1-C4) and correlation with the Wilson, NRTL and UNIQUAC models at T = 298 K
rhol	806.00	kg/m3	298.15	Isobaric Vapor-Liquid Equilibrium for the Binary Systems of 2-Butanol + 2-(Methoxymethoxy)-butane and 1-Butanol + 2-(Methoxymethoxy)-butane at 101.3 kPa
sfust	50.46	J/mol×K	183.90	NIST Webbook
sfust	50.79	J/mol×K	184.50	NIST Webbook
speedsl	1212.00	m/s	308.15	Densities, Speeds of Sound, Excess Molar Volumes, and Excess Isentropic Compressibilities at T = (298.15 and 308.15) K for Methyl Methacrylate + 1-Alkanols (1-Butanol, 1-Pentanol, and 1-Heptanol) + Cyclohexane, + Benzene, + Toluene, + p-Xylene, and + Ethylbenzene
speedsl	1191.40	m/s	313.15	Excess Enthalpy, Density, and Speed of Sound for the Mixtures a-Pinene + 1-Butanol or 2-Butanol at (283.15, 298.15, and 313.15) K

speedsl	1240.70	m/s	298.15	Excess Enthalpy, Density, and Speed of Sound for the Mixtures a-Pinene + 1-Butanol or 2-Butanol at (283.15, 298.15, and 313.15) K	
speedsl	1292.20	m/s	283.15	Excess Enthalpy, Density, and Speed of Sound for the Mixtures a-Pinene + 1-Butanol or 2-Butanol at (283.15, 298.15, and 313.15) K	
speedsl	1241.60	m/s	298.15	Compressibility Studies of Binary Solutions Involving Water as a Solute in Nonaqueous Solvents at T) 298.15 K	
speedsl	1244.00	m/s	298.15	Densities, Speeds of Sound, Excess Molar Volumes, and Excess Isentropic Compressibilities at T = (298.15 and 308.15) K for Methyl Methacrylate + 1-Alkanols (1-Butanol, 1-Pentanol, and 1-Heptanol) + Cyclohexane, + Benzene, + Toluene, + p-Xylene, and + Ethylbenzene	
speedsl	1172.12	m/s	318.15	Thermodynamic Properties of dodecane + 1-butanol and + 2-butanol systems	
speedsl	1205.50	m/s	308.15	Thermodynamic Properties of dodecane + 1-butanol and + 2-butanol systems	

speedsl	1239.29	m/s	298.15	Thermodynamic Properties of dodecane + 1-butanol and + 2-butanol systems	
speedsl	1273.51	m/s	288.15	Thermodynamic Properties of dodecane + 1-butanol and + 2-butanol systems	
speedsl	1308.19	m/s	278.15	Thermodynamic Properties of dodecane + 1-butanol and + 2-butanol systems	
speedsl	1170.00	m/s	318.15	Ultrasonic study of molecular interactions in binary mixtures of methyl acrylate with 1-alkanols (C4 to C10) at temperatures from (288.15 to 318.15) K	
speedsl	1189.20	m/s	313.15	Ultrasonic study of molecular interactions in binary mixtures of methyl acrylate with 1-alkanols (C4 to C10) at temperatures from (288.15 to 318.15) K	
speedsl	1207.60	m/s	308.15	Ultrasonic study of molecular interactions in binary mixtures of methyl acrylate with 1-alkanols (C4 to C10) at temperatures from (288.15 to 318.15) K	
speedsl	1225.80	m/s	303.15	Ultrasonic study of molecular interactions in binary mixtures of methyl acrylate with 1-alkanols (C4 to C10) at temperatures from (288.15 to 318.15) K	

speedsl	1242.60	m/s	298.15	Ultrasonic study of molecular interactions in binary mixtures of methyl acrylate with 1-alkanols (C4 to C10) at temperatures from (288.15 to 318.15) K
speedsl	1262.40	m/s	293.15	Ultrasonic study of molecular interactions in binary mixtures of methyl acrylate with 1-alkanols (C4 to C10) at temperatures from (288.15 to 318.15) K
speedsl	1282.20	m/s	288.15	Ultrasonic study of molecular interactions in binary mixtures of methyl acrylate with 1-alkanols (C4 to C10) at temperatures from (288.15 to 318.15) K
speedsl	1156.90	m/s	323.15	Acoustic, volumetric and osmotic properties of binary mixtures containing the ionic liquid 1-butyl-3-methylimidazolim dicyanamide mixed with primary and secondary alcohols
speedsl	1206.50	m/s	308.15	Acoustic, volumetric and osmotic properties of binary mixtures containing the ionic liquid 1-butyl-3-methylimidazolim dicyanamide mixed with primary and secondary alcohols

speedsl	1256.70	m/s	293.15	Acoustic, volumetric and osmotic properties of binary mixtures containing the ionic liquid 1-butyl-3-methylimidazolim dicyanamide mixed with primary and secondary alcohols
speedsl	1241.30	m/s	298.15	Ultrasonic speeds and isentropic compressibilities of {difurylmethane + (C1 C6) n-alkanol} binary mixtures at T = 298.15 K
speedsl	1239.80	m/s	298.15	Excess enthalpy, density, and speed of sound determination for the ternary mixture (methyl tert-butyl ether + 1-butanol + n-hexane)
speedsl	1205.90	m/s	308.15	Excess enthalpy, density, and speed of sound determination for the ternary mixture (methyl tert-butyl ether + 1-butanol + n-hexane)
speedsl	1222.70	m/s	303.15	Excess enthalpy, density, and speed of sound determination for the ternary mixture (methyl tert-butyl ether + 1-butanol + n-hexane)
speedsl	1257.00	m/s	293.15	Excess enthalpy, density, and speed of sound determination for the ternary mixture (methyl tert-butyl ether + 1-butanol + n-hexane)

speedsl	1273.70	m/s	288.15	Excess enthalpy, density, and speed of sound determination for the ternary mixture (methyl tert-butyl ether + 1-butanol + n-hexane)
speedsl	1256.30	m/s	293.15	Acoustic and Thermophysical Properties of Binary Liquid Mixtures of Primary Butanols with Hexane and Cyclohexane at 293.15 K
srf	0.02	N/m	303.15	Density and Surface Tension Variation with Temperature for Heptane + 1-Alkanol
srf	0.02	N/m	303.15	Study of the Temperature Dependence of Surface Tensions of Some Alkanol + Hexane Mixtures
srf	0.02	N/m	313.15	Study of the Temperature Dependence of Surface Tensions of Some Alkanol + Hexane Mixtures
srf	0.02	N/m	298.00	Surface Tension and Density of Pure Ionic Liquids and Some Binary Mixtures with 1-Propanol and 1-Butanol
srf	0.03	N/m	293.20	KDB
srf	0.02	N/m	298.15	Study of the Temperature Dependence of Surface Tensions of Some Alkanol + Hexane Mixtures
srf	0.02	N/m	293.15	Study of the Temperature Dependence of Surface Tensions of Some Alkanol + Hexane Mixtures

srf	0.03	N/m	288.15	Study of the Temperature Dependence of Surface Tensions of Some Alkanol + Hexane Mixtures
srf	0.02	N/m	308.15	Study of the Temperature Dependence of Surface Tensions of Some Alkanol + Hexane Mixtures
srf	0.03	N/m	283.15	Study of the Temperature Dependence of Surface Tensions of Some Alkanol + Hexane Mixtures
srf	0.02	N/m	308.15	Density and Surface Tension Variation with Temperature for Heptane + 1-Alkanol
srf	0.02	N/m	298.15	Density and Surface Tension Variation with Temperature for Heptane + 1-Alkanol
srf	0.02	N/m	293.15	Density and Surface Tension Variation with Temperature for Heptane + 1-Alkanol
srf	0.02	N/m	288.15	Density and Surface Tension Variation with Temperature for Heptane + 1-Alkanol
srf	0.02	N/m	293.15	Density and Surface Tension of Binary Mixtures of Acetonitrile + 1-Alkanol at 293.15 K
srf	0.02	N/m	323.15	Surface Tension of Dilute Solutions of Linear Alcohols in Benzyl Alcohol
srf	0.02	N/m	318.15	Surface Tension of Dilute Solutions of Linear Alcohols in Benzyl Alcohol

srf	0.02	N/m	313.15	Surface Tension of Dilute Solutions of Linear Alcohols in Benzyl Alcohol	
srf	0.02	N/m	308.15	Surface Tension of Dilute Solutions of Linear Alcohols in Benzyl Alcohol	
srf	0.02	N/m	303.15	Surface Tension of Dilute Solutions of Linear Alcohols in Benzyl Alcohol	
srf	0.02	N/m	298.15	Surface Tension of Dilute Solutions of Linear Alcohols in Benzyl Alcohol	
srf	0.02	N/m	298.15	Concentration Dependence of Surface Tension for Very Dilute Aqueous Solutions of Organic Non-Electrolytes	
srf	0.02	N/m	333.20	Surface tension and interfacial compositions of binary glycerol/alcohol mixtures	
srf	0.02	N/m	323.20	Surface tension and interfacial compositions of binary glycerol/alcohol mixtures	
srf	0.02	N/m	313.20	Surface tension and interfacial compositions of binary glycerol/alcohol mixtures	
srf	0.02	N/m	303.20	Surface tension and interfacial compositions of binary glycerol/alcohol mixtures	
srf	0.02	N/m	293.20	Surface tension and interfacial compositions of binary glycerol/alcohol mixtures	

srf	0.02	N/m	328.15	Study of surface tension and surface properties of binary systems of DMSO with long chain alcohols at various temperatures	
srf	0.02	N/m	318.15	Study of surface tension and surface properties of binary systems of DMSO with long chain alcohols at various temperatures	
srf	0.02	N/m	308.15	Study of surface tension and surface properties of binary systems of DMSO with long chain alcohols at various temperatures	
srf	0.02	N/m	298.15	Study of surface tension and surface properties of binary systems of DMSO with long chain alcohols at various temperatures	
srf	0.02	N/m	288.15	Study of surface tension and surface properties of binary systems of DMSO with long chain alcohols at various temperatures	
srf	0.02	N/m	293.15	Surface Tension of Dilute Solutions of Linear Alcohols in Benzyl Alcohol	

Pressure Dependent Properties

Property code	Value	Unit	Pressure [kPa]	Source

tbp	390.65	K	101.40	Isobaric Vapor Liquid Equilibrium Data for the 1,2-Dichloroethane-1-Butanol System
tbp	386.45	K	86.71	Isobaric Vapor Liquid Equilibrium Data for the 1,2-Dichloroethane-1-Butanol System
tbp	384.35	K	80.04	Isobaric Vapor Liquid Equilibrium Data for the 1,2-Dichloroethane-1-Butanol System
tbp	382.05	K	73.37	Isobaric Vapor Liquid Equilibrium Data for the 1,2-Dichloroethane-1-Butanol System
tbp	379.65	K	66.66	Isobaric Vapor Liquid Equilibrium Data for the 1,2-Dichloroethane-1-Butanol System
tbp	376.95	K	59.97	Isobaric Vapor Liquid Equilibrium Data for the 1,2-Dichloroethane-1-Butanol System

Correlations

Information	Value
Property code	pvap
Equation	ln(Pvp) = A + B/(T + C)
Coeff. A	1.47969e+01
Coeff. B	-2.95887e+03
Coeff. C	-9.99030e+01
Temperature range (K), min.	303.83
Temperature range (K), max.	411.84

Information	Value
Property code	pvap

Equation	ln(Pvp) = A + B/T + C*ln(T) + D*T^2
Coeff. A	8.67221e+01
Coeff. B	-9.12750e+03
Coeff. C	-9.88261e+00
Coeff. D	1.42848e-06
Temperature range (K), min.	183.85
Temperature range (K), max.	563.00

Datasets

Molar heat capacity at constant pressure, J/K/mol

Temperature, K - Liquid	Pressure, kPa - Liquid	Molar heat capacity at constant pressure, J/K/mol - Liquid
273.15	100.00	159.50
273.15	5000.00	158.70
273.15	10000.00	158.40
273.15	15000.00	158.20
273.15	20000.00	157.90
273.15	25000.00	157.20
293.15	100.00	175.20
293.15	5000.00	173.90
293.15	10000.00	172.80
293.15	15000.00	173.50
293.15	20000.00	172.10
293.15	25000.00	171.70
313.15	100.00	190.40
313.15	5000.00	189.80
313.15	10000.00	189.00
313.15	15000.00	187.60
313.15	20000.00	186.70
313.15	25000.00	186.00
333.15	100.00	206.60
333.15	5000.00	206.50
333.15	10000.00	204.90
333.15	15000.00	202.00
333.15	20000.00	200.70
333.15	25000.00	199.90

Reference *https://www.doi.org/10.1016/j.jct.2011.08.012*

Viscosity, Pa*s

Temperature, K - Liquid	Pressure, kPa - Liquid	Viscosity, Pa*s - Liquid
298.15	81.50	0.0025720

Reference *https://www.doi.org/10.1016/j.jct.2016.12.036*

Refractive index (Na D-line)

Pressure, kPa - Liquid	Temperature, K - Liquid	Refractive index (Na D-line) - Liquid
100.00	298.15	1.3975

Reference *https://www.doi.org/10.1016/j.jct.2016.01.012*

Temperature, K	Pressure, kPa	Refractive index (Na D-line)
294.15	100.00	1.39782

Reference *https://www.doi.org/10.1021/acs.jced.5b01024*

Mass density, kg/m3

Temperature, K - Gas	Pressure, kPa - Gas	Mass density, kg/m3 - Gas
528.31	2642.00	113.76
539.69	3127.00	113.76
543.30	3315.00	113.76
546.79	3456.00	113.76
548.15	3531.00	113.76
551.66	3610.00	113.76
558.64	3788.00	113.76
574.48	4178.00	113.76
590.87	4586.00	113.76

605.17	4925.00	113.76
616.28	5197.00	113.76
531.38	2775.00	167.7
543.89	3315.00	167.7
550.45	3647.00	167.7
555.53	3893.00	167.7
557.65	4020.00	167.7
562.14	4275.00	167.7
569.15	4559.00	167.7
580.80	5000.00	167.7
593.16	5452.00	167.7
604.17	5852.00	167.7
616.27	6275.00	167.7
533.03	2834.00	203.9
548.49	3579.00	203.9
555.70	3941.00	203.9
559.33	4118.00	203.9
563.04	4340.00	203.9
568.44	4598.00	203.9
577.73	5059.00	203.9
588.79	5613.00	203.9
600.42	6162.00	203.9
614.97	6834.00	203.9

Reference *https://www.doi.org/10.1016/j.fluid.2012.09.026*

Mass density, kg/m3

Temperature, K - Liquid	Pressure, kPa - Liquid	Mass density, kg/m3 - Liquid
293.15	100.00	809.5
293.15	10000.00	816.4
293.15	20000.00	822.9
293.15	30000.00	828.9
293.15	40000.00	834.5
293.15	50000.00	839.8
293.15	60000.00	844.7
293.15	70000.00	849.6
293.15	80000.00	854.2
293.15	90000.00	858.5
293.15	100000.00	862.9
293.15	110000.00	866.7

293.15	120000.00	870.8
293.15	130000.00	874.4
293.15	140000.00	878.1
303.15	100.00	802.0
303.15	10000.00	809.4
303.15	20000.00	816.0
303.15	30000.00	822.3
303.15	40000.00	828.0
303.15	50000.00	833.6
303.15	60000.00	838.7
303.15	70000.00	843.6
303.15	80000.00	848.3
303.15	90000.00	852.7
303.15	100000.00	857.0
303.15	110000.00	861.2
303.15	120000.00	865.0
303.15	130000.00	868.9
303.15	140000.00	872.6
313.15	100.00	794.1
313.15	10000.00	801.6
313.15	20000.00	809.1
313.15	30000.00	815.4
313.15	40000.00	821.4
313.15	50000.00	827.2
313.15	60000.00	832.6
313.15	70000.00	837.8
313.15	80000.00	842.6
313.15	90000.00	847.3
313.15	100000.00	851.6
313.15	110000.00	856.0
313.15	120000.00	860.1
313.15	130000.00	864.0
313.15	140000.00	867.9
323.15	100.00	786.5
323.15	10000.00	794.5
323.15	20000.00	801.7
323.15	30000.00	808.5
323.15	40000.00	815.0
323.15	50000.00	821.0
323.15	60000.00	826.5
323.15	70000.00	831.5
323.15	80000.00	836.3
323.15	90000.00	841.1
323.15	100000.00	845.8

323.15	110000.00	850.1
323.15	120000.00	854.3
323.15	130000.00	858.3
323.15	140000.00	862.1
333.15	100.00	778.3
333.15	10000.00	786.8
333.15	20000.00	794.6
333.15	30000.00	801.7
333.15	40000.00	808.4
333.15	50000.00	814.6
333.15	60000.00	820.3
333.15	70000.00	825.6
333.15	80000.00	830.8
333.15	90000.00	836.0
333.15	100000.00	840.5
333.15	110000.00	844.9
333.15	120000.00	849.0
333.15	130000.00	853.2
333.15	140000.00	857.0
343.15	100.00	769.9
343.15	10000.00	779.0
343.15	20000.00	787.1
343.15	30000.00	794.4
343.15	40000.00	801.3
343.15	50000.00	807.8
343.15	60000.00	813.6
343.15	70000.00	819.3
343.15	80000.00	824.6
343.15	90000.00	829.5
343.15	100000.00	834.4
343.15	110000.00	838.9
343.15	120000.00	843.3
343.15	130000.00	847.5
343.15	140000.00	851.6
353.15	100.00	760.9
353.15	10000.00	770.6
353.15	20000.00	779.2
353.15	30000.00	787.1
353.15	40000.00	794.3
353.15	50000.00	800.9
353.15	60000.00	807.7
353.15	70000.00	813.2
353.15	80000.00	818.9
353.15	90000.00	823.9

353.15	100000.00	828.8
353.15	110000.00	833.5
353.15	120000.00	838.1
353.15	130000.00	842.1
353.15	140000.00	846.5
363.15	100.00	751.9
363.15	10000.00	762.6
363.15	20000.00	771.7
363.15	30000.00	780.1
363.15	40000.00	787.4
363.15	50000.00	794.3
363.15	60000.00	800.9
363.15	70000.00	807.0
363.15	80000.00	812.5
363.15	90000.00	817.9
363.15	100000.00	822.9
363.15	110000.00	827.7
363.15	120000.00	832.4
363.15	130000.00	836.8
363.15	140000.00	841.0
373.15	100.00	742.4
373.15	10000.00	753.4
373.15	20000.00	763.3
373.15	30000.00	772.2
373.15	40000.00	780.0
373.15	50000.00	787.2
373.15	60000.00	793.8
373.15	70000.00	800.2
373.15	80000.00	805.8
373.15	90000.00	811.5
373.15	100000.00	817.0
373.15	110000.00	821.9
373.15	120000.00	826.3
373.15	130000.00	831.0
373.15	140000.00	835.3
383.15	100.00	732.2
383.15	10000.00	744.7
383.15	20000.00	754.6
383.15	30000.00	764.3
383.15	40000.00	772.2
383.15	50000.00	779.6
383.15	60000.00	786.6
383.15	70000.00	793.2
383.15	80000.00	799.1

383.15	90000.00	804.9
383.15	100000.00	810.3
383.15	110000.00	815.4
383.15	120000.00	820.4
383.15	130000.00	825.0
383.15	140000.00	829.6
393.15	10000.00	735.3
393.15	20000.00	746.4
393.15	30000.00	755.9
393.15	40000.00	764.9
393.15	50000.00	772.3
393.15	60000.00	779.6
393.15	70000.00	786.3
393.15	80000.00	792.8
393.15	90000.00	798.8
393.15	100000.00	804.3
393.15	110000.00	809.7
393.15	120000.00	814.7
393.15	130000.00	819.3
393.15	140000.00	823.9
403.15	10000.00	725.6
403.15	20000.00	737.3
403.15	30000.00	747.7
403.15	40000.00	756.9
403.15	50000.00	765.4
403.15	60000.00	772.8
403.15	70000.00	779.9
403.15	80000.00	786.5
403.15	90000.00	792.6
403.15	100000.00	798.5
403.15	110000.00	803.9
403.15	120000.00	809.1
403.15	130000.00	814.1
403.15	140000.00	819.3

Reference *https://www.doi.org/10.1016/j.fluid.2010.11.020*

Temperature, K	Pressure, kPa	Mass density, kg/m3
313.10	1013.00	794.97
313.10	2008.00	795.76
313.10	3007.00	796.55
313.10	3990.00	797.33
313.10	5000.00	798.11

313.10	6014.00	798.9
313.10	7006.00	799.69
313.10	8016.00	800.47
313.10	9008.00	801.23
313.10	9996.00	801.97
313.10	11012.00	802.72
313.10	12008.00	803.48
313.10	13036.00	804.23
313.10	14009.00	804.93
313.10	15034.00	805.67
313.10	16014.00	806.36
313.10	17013.00	807.07
313.10	18011.00	807.76
313.10	19027.00	808.46
313.10	20015.00	809.14
313.10	21034.00	809.84
313.10	22060.00	810.53
313.10	23039.00	811.19
313.10	24049.00	811.9
313.10	25048.00	812.56
323.08	1040.00	787.23
323.08	2043.00	788.06
323.08	3005.00	788.88
323.08	4009.00	789.7
323.08	5014.00	790.53
323.08	6024.00	791.38
323.08	7062.00	792.21
323.08	8033.00	792.99
323.08	9025.00	793.79
323.08	10008.00	794.58
323.08	11020.00	795.36
323.08	12027.00	796.15
323.08	13044.00	796.95
323.08	14035.00	797.69
323.08	15055.00	798.46
323.08	16063.00	799.21
323.08	17039.00	799.93
323.08	18057.00	800.67
323.08	18995.00	801.37
323.08	20045.00	802.11
323.08	21090.00	802.85
323.08	22039.00	803.52
323.08	23078.00	804.27
323.08	24010.00	804.94

323.08	25055.00	805.68
332.95	1095.00	779.28
332.95	2064.00	780.15
332.95	3035.00	781.02
332.95	4039.00	781.89
332.95	5007.00	782.76
332.95	6035.00	783.65
332.95	7048.00	784.53
332.95	8021.00	785.37
332.95	9035.00	786.23
332.95	10030.00	787.06
332.95	11057.00	787.9
332.95	12026.00	788.71
332.95	13041.00	789.54
332.95	14034.00	790.33
332.95	15074.00	791.16
332.95	16022.00	791.89
332.95	16957.00	792.61
332.95	18103.00	793.49
332.95	19053.00	794.22
332.95	20044.00	794.97
332.95	21090.00	795.75
332.95	21974.00	796.4
332.95	22960.00	797.13
332.95	24014.00	797.92
332.95	25045.00	798.67
342.89	1021.00	771.26
342.89	2060.00	772.26
342.89	3051.00	773.2
342.89	4044.00	774.13
342.89	5014.00	775.03
342.89	6024.00	775.96
342.89	7050.00	776.88
342.89	8037.00	777.76
342.89	9052.00	778.65
342.89	10057.00	779.53
342.89	11053.00	780.38
342.89	12027.00	781.22
342.89	13077.00	782.09
342.89	14026.00	782.86
342.89	15080.00	783.73
342.89	16028.00	784.48
342.89	17031.00	785.27
342.89	18020.00	786.04

342.89	19009.00	786.8
342.89	20030.00	787.58
342.89	21066.00	788.35
342.89	21978.00	789.03
342.89	23061.00	789.83
342.89	24031.00	790.55
342.89	25045.00	791.3
352.76	1033.00	762.72
352.76	2075.00	763.8
352.76	3038.00	764.75
352.76	4004.00	765.74
352.76	5017.00	766.75
352.76	6042.00	767.75
352.76	7041.00	768.72
352.76	8024.00	769.64
352.76	9042.00	770.61
352.76	10044.00	771.56
352.76	11026.00	772.44
352.76	12073.00	773.4
352.76	13002.00	774.25
352.76	14050.00	775.17
352.76	15000.00	776.03
352.76	15999.00	776.89
352.76	17008.00	777.76
352.76	18066.00	778.65
352.76	19002.00	779.43
352.76	20014.00	780.28
352.76	21043.00	781.12
352.76	22024.00	781.94
352.76	23025.00	782.75
352.76	24015.00	783.56
352.76	25100.00	784.44
362.68	1031.00	753.17
362.68	2065.00	754.33
362.68	3038.00	755.4
362.68	4026.00	756.45
362.68	5050.00	757.54
362.68	6018.00	758.55
362.68	6995.00	759.56
362.68	8100.00	760.7
362.68	9002.00	761.62
362.68	10021.00	762.62
362.68	10995.00	763.58
362.68	12059.00	764.62

362.68	13056.00	765.56
362.68	14029.00	766.47
362.68	15024.00	767.42
362.68	16010.00	768.33
362.68	17020.00	769.25
362.68	18041.00	770.16
362.68	19031.00	771.05
362.68	20045.00	771.93
362.68	21016.00	772.76
362.68	22022.00	773.64

Reference *https://www.doi.org/10.1016/j.jct.2006.07.009*

Temperature, K	Pressure, kPa	Mass density, kg/m3
278.15	100.00	820.9
278.15	1000.00	821.6
278.15	5000.00	824.4
278.15	10000.00	827.7
278.15	15000.00	830.9
278.15	20000.00	834.0
278.15	25000.00	837.0
278.15	30000.00	839.9
278.15	35000.00	842.7
278.15	40000.00	845.4
278.15	45000.00	848.1
278.15	50000.00	850.7
278.15	55000.00	853.2
278.15	60000.00	855.6
288.15	100.00	813.5
288.15	1000.00	814.1
288.15	5000.00	817.0
288.15	10000.00	820.5
288.15	15000.00	823.9
288.15	20000.00	827.2
288.15	25000.00	830.3
288.15	30000.00	833.2
288.15	35000.00	836.2
288.15	40000.00	839.0
288.15	45000.00	841.7
288.15	50000.00	844.4
288.15	55000.00	846.8
288.15	60000.00	849.4
298.15	100.00	805.8

298.15	1000.00	806.6
298.15	5000.00	809.6
298.15	10000.00	813.3
298.15	15000.00	816.8
298.15	20000.00	820.2
298.15	25000.00	823.5
298.15	30000.00	826.6
298.15	35000.00	829.6
298.15	40000.00	832.6
298.15	45000.00	835.4
298.15	50000.00	838.2
298.15	55000.00	840.9
298.15	60000.00	843.3
308.15	100.00	798.0
308.15	1000.00	798.8
308.15	5000.00	802.0
308.15	10000.00	805.9
308.15	15000.00	809.5
308.15	20000.00	813.0
308.15	25000.00	816.4
308.15	30000.00	819.7
308.15	35000.00	822.8
308.15	40000.00	825.9
308.15	45000.00	828.8
308.15	50000.00	831.7
308.15	55000.00	834.5
308.15	60000.00	837.0
318.15	100.00	790.3
318.15	1000.00	791.1
318.15	5000.00	794.5
318.15	10000.00	798.5
318.15	15000.00	802.4
318.15	20000.00	805.6
318.15	25000.00	809.5
318.15	30000.00	812.9
318.15	35000.00	816.1
318.15	40000.00	819.3
318.15	45000.00	822.4
318.15	50000.00	825.3
318.15	55000.00	828.2
318.15	60000.00	830.9
328.15	100.00	782.3
328.15	1000.00	783.1
328.15	5000.00	786.7

328.15	10000.00	790.9
328.15	15000.00	795.0
328.15	20000.00	798.8
328.15	25000.00	802.4
328.15	30000.00	806.0
328.15	35000.00	809.4
328.15	40000.00	812.7
328.15	45000.00	815.8
328.15	50000.00	818.9
328.15	55000.00	821.8
328.15	60000.00	824.6
338.15	100.00	774.0
338.15	1000.00	774.9
338.15	5000.00	778.7
338.15	10000.00	783.2
338.15	15000.00	787.4
338.15	20000.00	791.4
338.15	25000.00	795.3
338.15	30000.00	799.0
338.15	35000.00	802.5
338.15	40000.00	805.9
338.15	45000.00	809.2
338.15	50000.00	812.4
338.15	55000.00	815.4
338.15	60000.00	818.4
348.15	100.00	765.6
348.15	1000.00	766.3
348.15	5000.00	770.5
348.15	10000.00	775.3
348.15	15000.00	779.8
348.15	20000.00	784.0
348.15	25000.00	788.0
348.15	30000.00	791.8
348.15	35000.00	795.6
348.15	40000.00	799.1
348.15	45000.00	802.5
348.15	50000.00	805.8
348.15	55000.00	809.0
348.15	60000.00	811.9
358.15	100.00	756.6
358.15	1000.00	757.7
358.15	5000.00	761.7
358.15	10000.00	767.0
358.15	15000.00	771.7

358.15	20000.00	776.2
358.15	25000.00	780.4
358.15	30000.00	784.4
358.15	35000.00	788.3
358.15	40000.00	792.0
358.15	45000.00	795.5
358.15	50000.00	799.0
358.15	55000.00	802.3
358.15	60000.00	805.3

Reference *https://www.doi.org/10.1016/j.jct.2011.10.023*

Pressure, kPa	Temperature, K	Mass density, kg/m3
100.00	302.90	802.0
1020.00	302.90	802.7
2010.00	302.90	803.4
3010.00	302.90	804.1
4000.00	302.90	804.9
5010.00	302.90	805.7
6000.00	302.90	806.4
7020.00	302.90	807.2
8010.00	302.90	807.9
9020.00	302.90	808.7
10020.00	302.90	809.4
100.00	323.10	785.6
1010.00	323.10	786.6
2020.00	323.10	787.5
3010.00	323.10	788.3
4000.00	323.10	789.2
5020.00	323.10	790.0
6020.00	323.10	790.9
7000.00	323.10	791.7
8010.00	323.10	792.5
9020.00	323.10	793.3
10020.00	323.10	794.2

Reference *https://www.doi.org/10.1016/j.jct.2013.08.007*

Temperature, K	Pressure, kPa	Mass density, kg/m3
298.15	100.00	805.95

Reference *https://www.doi.org/10.1016/j.jct.2018.07.013*

Temperature, K	Pressure, kPa	Mass density, kg/m3
298.15	81.50	806.07

Reference *https://www.doi.org/10.1021/acs.jced.5b00162*

Temperature, K	Pressure, kPa	Mass density, kg/m3
283.63	80.00	817.19
283.64	1000.00	817.88
283.63	5000.00	820.7
283.66	10000.00	824.01
283.66	20000.00	830.39
283.65	30000.00	836.43
283.64	40000.00	842.03
283.64	60000.00	852.39
283.65	80000.00	861.72
283.64	100000.00	870.33
303.45	100.00	801.81
303.45	1000.00	802.54
303.46	5000.00	805.65
303.46	10000.00	809.35
303.47	20000.00	816.29
303.47	30000.00	822.84
303.47	40000.00	828.87
303.47	60000.00	839.95
303.48	80010.00	849.82
303.48	100000.00	858.94
323.26	100.00	786.36
323.25	1000.00	787.19
323.26	5000.00	790.63
323.27	10000.00	794.68
323.27	20000.00	802.33
323.27	30000.00	809.41
323.28	40000.00	816.0
323.28	60000.00	827.77
323.28	80000.00	838.31
323.28	100000.00	847.65
343.03	100.00	769.54
343.03	1000.00	770.45
343.04	5000.00	774.32
343.04	10000.00	778.88
343.05	20000.00	787.37

343.05	30000.00	795.06
343.05	40000.00	802.13
343.05	60000.00	814.83
343.05	80000.00	826.08
343.05	100000.00	836.2
362.68	100.00	752.67
362.68	1000.00	753.64
362.69	5000.00	758.03
362.69	10000.00	763.16
362.69	20000.00	772.56
362.68	30000.00	781.04
362.70	40000.00	788.67
362.70	60000.00	802.36
362.71	80000.00	814.44
362.71	100000.00	825.18

Reference *https://www.doi.org/10.1021/acs.jced.7b00176*

Temperature, K	Pressure, kPa	Mass density, kg/m3
298.12	2010.00	806.8
298.12	3910.00	808.2
298.12	6020.00	809.8
298.12	8030.00	811.2
298.12	9910.00	812.5
298.13	13010.00	814.6
298.12	16020.00	816.6
298.12	19030.00	818.5
298.12	22000.00	820.4
298.12	25020.00	822.3
298.12	28000.00	824.0
298.12	30010.00	825.1
303.15	2010.00	803.3
303.15	4000.00	804.8
303.15	5980.00	806.3
303.15	8010.00	807.8
303.15	9990.00	809.2
303.15	12990.00	811.3
303.15	15990.00	813.3
303.16	19000.00	815.3
303.15	22010.00	817.3
303.15	25010.00	819.1
303.15	27970.00	820.9
303.15	29990.00	822.1

313.16	1990.00	795.5
313.16	3980.00	797.1
313.16	5990.00	798.7
313.16	7980.00	800.2
313.15	10020.00	801.8
313.15	12980.00	803.9
313.16	15980.00	806.1
313.16	19010.00	808.2
313.15	22020.00	810.2
313.15	25010.00	812.2
313.16	27990.00	814.1
313.16	29990.00	815.3
323.14	2000.00	787.9
323.14	3920.00	789.6
323.14	6030.00	791.3
323.14	8000.00	792.9
323.14	10020.00	794.5
323.14	13040.00	796.9
323.14	16050.00	799.2
323.14	19000.00	801.3
323.14	21990.00	803.4
323.14	24970.00	805.5
323.14	28000.00	807.5
323.14	29970.00	808.8

Reference *https://www.doi.org/10.1021/acs.jced.8b00812*

Temperature, K	Pressure, kPa	Mass density, kg/m3
293.15	100.00	809.28
293.15	5000.00	813.04
293.15	10000.00	816.54
293.15	15000.00	819.83
293.15	20000.00	822.97
293.15	25000.00	826.0
293.15	30000.00	828.99
293.15	35000.00	831.99
303.15	100.00	801.9
303.15	5000.00	806.0
303.15	10000.00	809.77
303.15	15000.00	813.26
303.15	20000.00	816.52
303.15	25000.00	819.61
303.15	30000.00	822.58

303.15	35000.00	825.48
313.15	100.00	794.32
313.15	5000.00	798.42
313.15	10000.00	802.22
313.15	15000.00	805.8
313.15	20000.00	809.2
313.15	25000.00	812.49
313.15	30000.00	815.73
313.15	35000.00	818.97

Reference *https://www.doi.org/10.1021/acs.jced.8b00975*

Temperature, K	Pressure, kPa	Mass density, kg/m3
283.15	550.80	818.14
283.15	11959.00	825.705
283.15	18047.00	829.455
283.15	32971.00	838.117
283.15	62025.00	852.74
283.15	267.60	817.937
283.15	2020.20	819.134
283.15	7034.80	822.538
283.15	9991.90	824.485
283.15	15011.00	827.609
283.15	25044.00	833.615
283.15	28046.00	835.368
283.15	32970.00	838.092
283.15	34960.00	839.141
283.15	37943.00	840.737
283.15	44940.00	844.325
283.15	49996.00	846.902
283.15	51982.00	847.942
283.15	57949.00	850.826
283.15	991.30	818.415
283.15	2914.80	819.722
283.15	4946.50	821.145
283.15	9021.60	823.807
283.15	34017.00	838.557
283.15	40004.00	841.811
283.15	42058.00	842.896
283.15	54965.00	849.358
283.15	64998.00	854.215
283.15	878.60	818.388
283.15	20045.00	830.642

283.15	30052.00	836.501
283.15	48022.00	845.954
283.15	59996.00	851.847
283.15	22057.00	831.928
283.16	65971.00	854.69
283.16	101.20	817.825
288.15	882.70	814.204
288.15	2917.80	815.651
288.15	9018.60	819.847
288.15	10040.00	820.583
288.15	20693.00	827.38
288.15	32084.00	834.074
288.15	37993.00	837.311
288.15	40041.00	838.449
288.15	44979.00	841.064
288.15	52032.00	844.696
288.15	55077.00	846.251
288.15	61974.00	849.542
288.15	64942.00	850.93
288.15	391.30	813.84
288.15	66025.00	851.488
288.15	7088.80	818.541
288.15	12072.00	821.91
288.15	17968.00	825.72
288.15	78.50	813.612
288.15	25002.00	829.984
288.15	28026.00	831.832
288.15	30035.00	832.896
288.15	33050.00	834.714
288.15	34036.00	835.152
288.15	50068.00	843.734
288.15	554.40	813.953
288.15	2057.50	815.033
288.15	5041.40	817.16
288.15	15001.00	823.877
288.15	20068.00	826.976
288.15	22020.00	828.211
288.15	35077.00	835.674
288.15	42023.00	839.459
288.15	989.10	814.282
288.15	48045.00	842.655
288.15	58004.00	847.666
288.15	59988.00	848.618
293.15	33151.00	832.036

293.15	28066.00	829.163
293.15	58072.00	845.282
293.15	825.10	811.053
293.15	66045.00	849.224
293.15	21987.00	825.417
293.15	27986.00	829.06
293.15	35000.00	833.078
293.15	45068.00	838.658
293.15	1971.70	811.916
293.15	3060.60	812.72
293.15	5011.00	814.074
293.15	9042.10	816.992
293.15	91.10	810.528
293.15	10036.00	817.626
293.15	11947.00	818.963
293.15	14941.00	820.924
293.15	34099.00	832.526
293.15	40081.00	835.959
293.15	42099.00	837.008
293.15	54966.00	843.81
293.16	1045.20	811.221
293.16	7096.80	815.616
293.16	17889.00	822.832
293.16	19992.00	824.147
293.16	24983.00	827.251
293.16	29976.00	830.11
293.16	32232.00	831.521
293.16	476.40	810.81
293.16	38027.00	834.783
293.16	48022.00	840.192
293.16	50085.00	841.31
293.16	52032.00	842.308
293.16	60018.00	846.257
293.16	64991.00	848.649
293.16	305.20	810.686
293.16	61962.00	847.209
298.15	33113.00	828.365
298.15	45021.00	835.052
298.15	55016.00	840.375
298.15	65018.00	845.412
298.15	10107.00	813.67
298.15	38097.00	831.213
298.15	40161.00	832.357
298.15	57931.00	841.819

298.15	59914.00	842.826
298.15	22004.00	821.641
298.15	50068.00	837.797
298.15	62036.00	843.934
298.15	5067.60	810.04
298.15	305.40	806.45
298.15	42029.00	833.462
298.15	1980.90	807.747
298.15	48107.00	836.745
298.15	3167.80	808.637
298.15	7065.30	811.55
298.15	854.50	806.874
298.15	20128.00	820.445
298.15	25081.00	823.55
298.15	33924.00	828.845
298.15	98.10	806.282
298.15	42359.00	833.614
298.16	9120.90	813.012
298.16	11970.00	814.957
298.16	15220.00	817.234
298.16	17984.00	819.076
298.16	28040.00	825.42
298.16	30089.00	826.686
298.16	32030.00	827.814
298.16	35217.00	829.575
298.16	52020.00	838.835
298.16	493.70	806.599
298.16	1012.60	806.992
298.16	66118.00	845.919
303.15	498.90	802.435
303.15	32862.00	824.267
303.15	34081.00	824.9
303.15	39904.00	828.27
303.15	1977.60	803.554
303.15	52073.00	834.942
303.15	55075.00	836.525
303.15	57989.00	838.099
303.15	64988.00	841.599
303.15	33112.00	824.372
303.15	304.10	802.278
303.15	62052.00	840.165
303.15	4930.40	805.792
303.15	42014.00	829.398
303.15	44896.00	831.035

303.15	50065.00	833.864
303.15	979.10	802.787
303.15	19974.00	816.207
303.15	21920.00	817.555
303.15	31958.00	823.711
303.15	34980.00	825.355
303.15	59979.00	839.13
303.15	66030.00	842.119
303.15	11975.00	810.901
303.15	900.60	802.733
303.15	15044.00	812.956
303.15	24992.00	819.44
303.15	37951.00	827.164
303.15	47951.00	832.781
303.16	10014.00	809.449
303.16	18021.00	814.973
303.16	27933.00	821.256
303.16	29990.00	822.537
303.16	8858.20	808.638
303.16	2964.90	804.3
303.16	100.00	802.125
303.16	7045.60	807.335
308.15	12009.00	807.267
308.15	803.30	798.75
308.15	17850.00	811.315
308.15	21988.00	814.131
308.15	27955.00	817.999
308.15	32886.00	820.911
308.15	1916.70	799.642
308.15	4990.20	802.048
308.15	2980.70	800.489
308.15	19930.00	812.759
308.15	25012.00	816.144
308.15	31929.00	820.41
308.15	32941.00	821.059
308.15	37971.00	823.987
308.15	54982.00	833.542
308.15	9863.10	805.692
308.15	40011.00	825.154
308.15	8864.30	804.962
308.15	494.10	798.504
308.15	58009.00	835.036
308.15	1039.50	798.954
308.15	29906.00	819.191

308.15	34011.00	821.636
308.15	45017.00	828.081
308.15	50044.00	830.856
308.15	61985.00	837.136
308.15	64981.00	838.691
308.15	41998.00	826.392
308.15	48033.00	829.748
308.15	51977.00	831.933
308.15	6964.90	803.548
308.15	274.30	798.317
308.15	60030.00	836.098
308.16	15069.00	809.4
308.16	34949.00	822.224
308.16	85.20	798.155
308.16	66088.00	839.237
313.15	17921.00	807.72
313.15	24862.00	812.41
313.15	15059.00	805.716
313.15	19932.00	809.082
313.15	10209.00	802.164
313.15	6921.80	799.704
313.15	12020.00	803.525
313.15	5111.80	798.292
313.15	27869.00	814.438
313.15	29957.00	815.731
313.15	54924.00	830.321
313.15	32980.00	817.674
313.15	33981.00	818.214
313.15	2109.70	795.902
313.15	41978.00	822.973
313.15	59932.00	832.96
313.15	1088.80	795.074
313.15	32107.00	817.091
313.15	33002.00	817.783
313.15	38127.00	820.76
313.15	47880.00	826.39
313.15	103.10	794.254
313.15	52097.00	828.677
313.15	62004.00	834.048
313.15	66113.00	836.222
313.15	22058.00	810.55
313.15	316.00	794.433
313.15	9210.70	801.445
313.15	35045.00	818.819

313.15	44946.00	824.676
313.15	3137.00	796.707
313.15	58082.00	831.994
313.15	569.60	794.643
313.15	64954.00	835.652
313.16	860.70	794.888
313.16	39997.00	821.817
313.16	50021.00	827.556
318.15	11896.00	799.896
318.15	66028.00	833.327
318.15	33016.00	814.458
318.15	37838.00	817.415
318.15	59923.00	830.12
318.15	2933.20	792.866
318.15	8989.00	797.678
318.15	44884.00	821.648
318.15	72.70	790.456
318.15	18102.00	804.461
318.15	47944.00	823.402
318.15	150.10	790.516
318.15	57972.00	829.016
318.15	336.30	790.675
318.15	22094.00	807.257
318.15	29998.00	812.541
318.15	19985.00	805.78
318.15	40093.00	818.817
318.15	1105.70	791.334
318.15	54944.00	827.433
318.15	10075.00	798.523
318.15	1929.10	792.026
318.15	65022.00	832.785
318.16	27909.00	811.139
318.16	31957.00	813.824
318.16	891.00	791.16
318.16	34110.00	815.009
318.16	35048.00	815.63
318.16	6926.70	796.098
318.16	41996.00	819.89
318.16	5020.80	794.556
318.16	49950.00	824.565
318.16	14800.00	802.072
318.16	52037.00	825.741
318.16	61930.00	831.241
318.16	24873.00	809.191

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323.15	58045.00	825.293
323.15	17981.00	800.333
323.15	2006.90	787.814
323.15	54964.00	823.594
323.15	60050.00	826.448
323.15	64999.00	829.093
323.15	35113.00	811.773
323.15	9029.90	793.564
323.15	2966.60	788.632
323.15	48025.00	819.642
323.15	95.00	786.145
323.15	52114.00	821.994
323.15	20038.00	801.815
323.15	4996.20	790.328
323.15	34081.00	811.155
323.15	24993.00	805.264
323.15	7042.10	792.001
323.15	15111.00	798.229
323.15	62131.00	827.565
323.15	32079.00	809.955
323.15	37953.00	813.623
323.15	528.40	786.535
323.15	10090.00	794.392
323.15	42011.00	816.075
323.15	21995.00	803.198
323.15	994.20	786.944
323.15	11914.00	795.8
323.16	30055.00	808.674
323.16	27958.00	807.28
323.16	33067.00	810.61
323.16	40047.00	814.861
323.16	44937.00	817.856
323.16	920.70	786.888
323.16	66131.00	829.687
328.15	33066.00	807.772
328.15	35030.00	808.989
328.15	38010.00	810.915
328.15	44953.00	815.144
328.15	47978.00	816.959
328.15	49927.00	818.128
328.15	54947.00	821.045

328.15	59991.00	823.77
328.15	27979.00	804.324
328.15	61982.00	824.89
328.15	65032.00	826.439
328.15	66046.00	826.965
328.15	5013.40	786.446
328.15	2996.70	784.66
328.15	24924.00	802.164
328.15	57939.00	822.658
328.15	994.10	782.825
328.15	39969.00	812.181
328.15	34013.00	808.285
328.15	41959.00	813.311
328.15	289.90	782.179
328.15	51997.00	819.32
328.15	32060.00	807.104
328.15	29958.00	805.68
328.15	33107.00	807.703
328.15	9017.00	789.881
328.15	97.90	781.987
328.15	10040.00	790.727
328.15	897.50	782.733
328.15	6974.90	788.15
328.15	1988.80	783.741
328.16	20279.00	798.799
328.16	12210.00	792.522
328.16	18061.00	797.122
328.16	22013.00	800.057
328.16	15055.00	794.783
328.16	501.00	782.371
333.15	34086.00	805.827
333.15	37894.00	808.332
333.15	4997.40	783.921
333.15	47936.00	814.509
333.15	51953.00	816.927
333.15	58057.00	820.5
333.15	32067.00	804.524
333.15	62049.00	822.689
333.15	35075.00	806.492
333.15	40111.00	809.759
333.15	30016.00	803.156
333.15	22034.00	797.567
333.15	66005.00	824.853
333.15	44974.00	812.823

333.15	45490.00	813.093
333.15	49920.00	815.727
333.15	54876.00	818.691
333.15	7002.40	785.686
333.15	59993.00	821.547
333.15	64970.00	824.293
333.15	10210.00	788.367
333.15	41947.00	810.938
333.15	9032.00	787.401
333.15	532.60	779.903
333.15	28040.00	801.782
333.15	1987.20	781.249
333.15	897.80	780.245
333.15	3023.40	782.177
333.15	15008.00	792.251
333.16	303.50	779.691
333.16	25027.00	799.667
333.16	12011.00	789.847
333.16	20036.00	796.123
333.16	1003.00	780.346
333.16	17992.00	794.568
333.16	87.70	779.482
333.16	33073.00	805.134
338.15	25026.00	796.362
338.15	22041.00	794.077
338.15	14885.00	788.349
338.15	35012.00	803.345
338.15	52127.00	814.375
338.15	60014.00	819.058
338.15	65012.00	821.872
338.15	66057.00	822.451
338.15	47971.00	811.881
338.15	57940.00	817.865
338.15	32882.00	801.907
338.15	10034.00	784.276
338.15	44978.00	809.926
338.15	19927.00	792.473
338.15	49999.00	813.099
338.15	898.50	775.911
338.15	41990.00	808.025
338.15	5018.60	779.821
338.15	309.80	775.333
338.15	2936.10	777.859
338.15	29994.00	799.911

338.15	39987.00	806.717
338.15	17956.00	790.925
338.15	80.90	775.106
338.15	33937.00	802.603
338.15	561.60	775.578
338.15	1014.10	776.019
338.16	2000.20	776.987
338.16	6917.30	781.55
338.16	11959.00	785.981
338.16	27976.00	798.446
338.16	54992.00	816.085
338.16	31949.00	801.318
338.16	62074.00	820.252
338.16	37962.00	805.345
338.16	9046.90	783.457
338.16	67534.00	823.221
343.15	1017.70	770.563
343.15	52026.00	808.534
343.15	4994.80	774.313
343.15	30082.00	794.441
343.15	297.30	769.853
343.15	911.70	770.459
343.15	65070.00	815.911
343.15	44927.00	804.186
343.15	33097.00	796.546
343.15	520.30	770.07
343.15	2332.50	771.83
343.15	2968.30	772.434
343.15	61967.00	814.212
343.15	10060.00	778.847
343.15	6976.60	776.117
343.15	47908.00	806.033
343.15	1994.60	771.497
343.15	73.70	769.631
343.15	12025.00	780.506
343.15	59944.00	813.181
343.15	19985.00	786.983
343.15	66059.00	816.518
343.15	17975.00	785.398
343.15	9002.60	777.941
343.15	28045.00	793.012
343.15	40012.00	801.058
343.15	25091.00	790.858
343.15	14928.00	782.929

343.15	35030.00	797.771
343.15	50059.00	807.349
343.15	22033.00	788.56
343.15	32170.00	795.891
343.15	41972.00	802.307
343.15	37943.00	799.701
343.16	34034.00	797.101
343.16	54919.00	810.24
343.16	57986.00	812.021
348.15	49950.00	804.611
348.15	88.80	765.578
348.15	3051.10	768.513
348.15	859.80	766.355
348.15	64942.00	813.608
348.15	5094.00	770.529
348.15	40026.00	798.192
348.15	47915.00	803.304
348.15	22163.00	785.27
348.15	66153.00	814.319
348.15	1974.90	767.461
348.15	35045.00	794.751
348.15	57945.00	809.531
348.15	52167.00	805.991
348.15	20111.00	783.649
348.15	55022.00	807.759
348.15	33959.00	793.985
348.15	30149.00	791.333
348.15	32161.00	792.797
348.15	462.20	765.945
348.15	14792.00	779.258
348.15	974.20	766.465
348.15	37937.00	796.729
348.15	10025.00	775.087
348.15	60030.00	810.761
348.15	42066.00	799.57
348.15	28080.00	789.8
348.16	7050.20	772.352
348.16	45051.00	801.465
348.16	32942.00	793.364
348.16	9003.20	774.145
348.16	291.70	765.77
348.16	62063.00	811.961
348.16	17899.00	781.843
348.16	25100.00	787.55

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353.15	21894.00	781.405
353.15	4963.90	766.051
353.15	17977.00	778.173
353.15	336.00	761.273
353.15	40143.00	795.172
353.15	19937.00	779.792
353.15	29936.00	787.797
353.15	12128.00	772.98
353.15	52082.00	803.184
353.15	10005.00	771.015
353.15	45002.00	798.55
353.15	48029.00	800.525
353.15	926.00	761.914
353.15	7000.90	768.126
353.15	32961.00	790.048
353.15	37913.00	793.571
353.15	1016.40	762.008
353.15	50145.00	801.944
353.15	34945.00	791.423
353.15	60098.00	808.146
353.15	14954.00	775.544
353.15	42067.00	796.559
353.15	65063.00	811.071
353.15	2992.10	764.106
353.15	31876.00	789.263
353.15	86.20	760.995
353.15	24908.00	783.869
353.15	27850.00	786.199
353.15	9062.30	770.136
353.15	57980.00	806.879
353.15	55132.00	805.077
353.15	572.60	761.533
353.16	2113.40	763.181
353.16	62043.00	809.275
353.16	33964.00	790.701
353.16	66024.00	811.636
358.15	60043.00	804.453
358.15	17041.00	772.871
358.15	32793.00	786.01
358.15	50038.00	798.154
358.15	926.80	756.64
358.15	31978.00	785.37
358.15	4930.00	760.98

358.15	52081.00	799.474
358.15	37872.00	789.731
358.15	55073.00	801.393
358.15	62067.00	805.674
358.15	28035.00	782.278
358.15	33850.00	786.716
358.15	17965.00	773.738
358.15	66109.00	808.067
358.15	47958.00	796.761
358.15	40037.00	791.31
358.15	35027.00	787.554
358.15	64986.00	807.407
358.15	22175.00	777.415
358.15	14905.00	770.903
358.15	1013.90	756.732
358.15	12137.00	768.259
358.15	42100.00	792.749
358.15	25059.00	779.849
358.15	3021.20	758.947
358.15	508.40	756.19
358.15	57918.00	803.161
358.15	10073.00	766.29
358.15	264.10	755.935
358.15	30002.00	783.844
358.16	2083.90	757.886
358.16	95.10	755.754
358.16	45054.00	794.81
358.16	9030.20	765.202
358.16	20162.00	775.665
358.16	6993.50	763.142
363.15	50050.00	794.458
363.15	34000.00	783.178
363.15	54940.00	797.589
363.15	37919.00	786.081
363.15	31961.00	781.745
363.15	4979.20	757.84
363.15	11971.00	764.806
363.15	47947.00	793.03
363.15	27993.00	778.673
363.15	94.00	752.456
363.15	35028.00	783.945
363.15	42016.00	788.986
363.15	59980.00	800.724
363.15	30062.00	780.285

363.15	17973.00	770.255
363.15	33033.00	782.563
363.15	57917.00	799.479
363.15	601.70	753.061
363.15	892.90	753.404
363.15	19940.00	771.991
363.15	39916.00	787.488
363.15	52023.00	795.718
363.15	334.00	752.741
363.15	2940.40	755.648
363.15	961.60	753.487
363.15	1986.00	754.594
363.15	64924.00	803.734
363.15	66088.00	804.418
363.15	1034.70	753.567
363.15	9959.50	762.855
363.15	62043.00	802.008
363.15	14936.00	767.534
363.15	24872.00	776.146
363.16	8964.80	761.886
363.16	21933.00	773.696
363.16	44987.00	791.029
363.16	6920.20	759.821

Reference *https://www.doi.org/10.1021/je5010214*

Temperature, K	Pressure, kPa	Mass density, kg/m3
313.08	1006.00	795.22
313.08	2017.00	796.08
313.08	3007.00	796.85
313.08	4010.00	797.64
313.08	5016.00	798.44
313.08	6017.00	799.22
313.08	7011.00	800.0
313.08	8006.00	800.77
313.08	9019.00	801.53
313.08	10008.00	802.27
313.08	11014.00	803.01
313.08	12016.00	803.73
313.08	13016.00	804.46
313.08	14013.00	805.18
313.08	15010.00	805.91
313.08	16003.00	806.58

313.08	16999.00	807.29
313.08	17999.00	807.98
313.08	19024.00	808.67
313.08	20005.00	809.33
313.08	21013.00	810.09
313.08	22003.00	810.75
313.08	23024.00	811.41
313.08	24027.00	812.09
313.08	24234.00	812.21
323.05	999.00	787.38
323.05	2026.00	788.22
323.05	3007.00	789.03
323.05	4024.00	789.91
323.05	5012.00	790.75
323.05	6016.00	791.57
323.05	7023.00	792.4
323.05	8004.00	793.16
323.05	9009.00	793.98
323.05	10007.00	794.78
323.05	11004.00	795.57
323.05	11995.00	796.32
323.05	13007.00	797.09
323.05	14022.00	797.85
323.05	14989.00	798.63
323.05	16006.00	799.33
323.05	16998.00	800.1
323.05	18001.00	800.84
323.05	19025.00	801.56
323.05	19993.00	802.28
323.05	21032.00	803.06
323.05	22006.00	803.74
323.05	22993.00	804.41
323.05	24012.00	805.16
323.05	24418.00	805.43
332.93	1013.00	779.33
332.93	2026.00	780.27
332.93	3014.00	781.21
332.93	4006.00	782.04
332.93	5000.00	782.91
332.93	6013.00	783.78
332.93	6997.00	784.6
332.93	8003.00	785.5
332.93	9005.00	786.32
332.93	10001.00	787.17

332.93	10996.00	787.97
332.93	12003.00	788.8
332.93	12993.00	789.62
332.93	14001.00	790.38
332.93	15016.00	791.22
332.93	15998.00	791.96
332.93	16999.00	792.72
332.93	18009.00	793.53
332.93	18986.00	794.29
332.93	20043.00	795.08
332.93	21056.00	795.87
332.93	22040.00	796.59
332.93	23009.00	797.3
332.93	24001.00	798.04
332.93	24515.00	798.46
342.86	1026.00	770.92
342.86	2007.00	771.87
342.86	3003.00	772.85
342.86	4051.00	773.84
342.86	5008.00	774.72
342.86	6023.00	775.59
342.86	7010.00	776.55
342.86	8000.00	777.42
342.86	9001.00	778.28
342.86	10005.00	779.2
342.86	11006.00	780.06
342.86	12010.00	780.9
342.86	13033.00	781.77
342.86	14017.00	782.62
342.86	15001.00	783.45
342.86	16009.00	784.29
342.86	17016.00	785.14
342.86	17991.00	785.93
342.86	19023.00	786.79
342.86	19999.00	787.57
342.86	21000.00	788.37
342.86	22003.00	789.15
342.86	22989.00	789.91
342.86	24004.00	790.69
342.86	24621.00	791.2
352.79	1020.00	762.66
352.79	2014.00	763.68
352.79	3007.00	764.68
352.79	3998.00	765.67

352.79	4998.00	766.69
352.79	5997.00	767.68
352.79	7006.00	768.64
352.79	8005.00	769.61
352.79	9007.00	770.53
352.79	9997.00	771.44
352.79	11008.00	772.38
352.79	12000.00	773.26
352.79	13043.00	774.22
352.79	14002.00	775.11
352.79	15000.00	775.98
352.79	16000.00	776.85
352.79	16992.00	777.72
352.79	17988.00	778.56
352.79	18994.00	779.41
352.79	20012.00	780.22
352.79	21003.00	781.04
352.79	22010.00	781.87
352.79	22971.00	782.68
352.79	24021.00	783.49
352.79	24573.00	783.95
362.67	1043.00	753.5
362.67	2116.00	754.77
362.67	3026.00	755.74
362.67	4013.00	756.82
362.67	5024.00	757.87
362.67	6047.00	758.93
362.67	7014.00	759.93
362.67	8012.00	760.98
362.67	9025.00	762.03
362.67	10015.00	762.94
362.67	11002.00	763.92
362.67	12002.00	764.89
362.67	13011.00	765.86
362.67	14011.00	766.84
362.67	15014.00	767.7
362.67	16006.00	768.67
362.67	17009.00	769.58
362.67	18008.00	770.48
362.67	18993.00	771.37
362.67	19990.00	772.21
362.67	21007.00	773.1
362.67	21997.00	773.96
362.67	23014.00	774.89

362.67	24000.00	775.58
362.67	24563.00	776.12

Reference *https://www.doi.org/10.1021/je700261b*

Speed of sound, m/s

Temperature, K - Liquid	Pressure, kPa - Liquid	Speed of sound, m/s - Liquid
253.15	108.00	1399.203
253.15	5065.00	1423.053
253.15	10086.00	1446.233
253.15	15035.00	1468.319
253.15	20115.00	1490.177
253.15	25062.00	1510.75
253.15	30526.00	1532.696
273.15	103.00	1326.117
273.15	5008.00	1351.87
273.15	10094.00	1377.167
273.15	15185.00	1401.443
273.15	20115.00	1424.124
273.15	25088.00	1446.132
273.15	30170.00	1467.891
293.15	104.00	1256.423
293.15	5067.00	1284.566
293.15	10144.00	1311.663
293.15	15081.00	1336.897
293.15	20094.00	1361.485
293.15	25087.00	1384.998
293.15	30176.00	1408.068
313.15	100.00	1189.161
313.15	5050.00	1219.204
313.15	10082.00	1248.164
313.15	15167.00	1275.967
313.15	20187.00	1302.165
313.15	25473.00	1328.5
313.15	30162.00	1351.047
333.15	99.00	1122.893
333.15	5055.00	1155.491
333.15	10099.00	1186.761
333.15	15058.00	1215.841
333.15	20149.00	1244.214

333.15	25167.00	1270.825
333.15	30276.00	1296.753
353.15	100.00	1056.353
353.15	5013.00	1091.582
353.15	10032.00	1125.304
353.15	15066.00	1157.065
353.15	20085.00	1186.968
353.15	25075.00	1215.14
353.15	30204.00	1242.721

Reference *https://www.doi.org/10.1016/j.jct.2015.10.006*

Temperature, K	Pressure, kPa	Speed of sound, m/s
250.00	67.00	1411.8
218.17	100.00	1536.6
249.99	100.00	1412.0
300.04	101.00	1234.1
350.11	104.00	1067.1
400.14	145.00	891.2
449.96	576.00	696.4
350.11	911.00	1073.1
400.14	933.00	898.8
300.03	1100.00	1240.0
218.25	1111.00	1540.5
449.96	1176.00	704.8
499.96	1636.00	472.8
400.14	1839.00	907.3
449.96	2010.00	716.0
249.99	2060.00	1421.4
350.11	2067.00	1081.4
218.26	2129.00	1544.8
300.03	2196.00	1246.4
499.96	2336.00	490.0
350.11	4569.00	1099.0
400.14	4864.00	934.6
499.96	4923.00	544.0
249.99	5075.00	1435.7
300.03	5369.00	1264.5
218.31	5399.00	1558.2
499.96	7243.00	586.7
249.98	9057.00	1453.9
350.11	9077.00	1129.2
400.12	9483.00	973.3

449.96	9527.00	803.9
218.28	9570.00	1575.2
300.03	9681.00	1288.1
499.96	9870.00	626.9
499.95	19314.00	741.4
218.22	19695.00	1614.9
300.03	20060.00	1341.4
350.11	20660.00	1199.6
400.14	20722.00	1055.9
449.96	21165.00	910.4
249.98	22437.00	1511.7
499.91	25968.00	806.2
400.14	36029.00	1150.7
217.92	39005.00	1685.4
449.96	39912.00	1043.3
300.03	40032.00	1432.5
499.95	40048.00	918.1
250.01	40124.00	1581.1
350.10	40651.00	1303.9
249.98	58985.00	1648.4
400.14	59027.00	1269.1
300.02	59057.00	1509.0
350.10	59555.00	1388.7
218.12	59881.00	1753.2
449.96	59899.00	1156.8
499.96	60162.00	1045.7
400.14	79223.00	1357.8
249.98	79282.00	1714.5
300.02	79463.00	1582.8
449.96	79678.00	1251.8
218.18	79839.00	1813.2
499.96	80004.00	1149.3
350.09	80590.00	1472.2

Reference *https://www.doi.org/10.1021/acs.jced.8b00938*

Sources

Isobaric VLE data for the system of butan-1-ol + butylethanoate + 1-butyl-3-methylimidazoliumbis[(trifluoromethyl)sulfonyl]amide: https://www.doi.org/10.1016/j.fluid.2013.05.020

Solubility and Solution Thermodynamics of Betaine in Different Pure Solvents and Binary Mixtures: https://www.doi.org/10.1021/je2011659

Solubilities of Betulinic Acid in Thirteen Organic Solvents at Different Temperatures: https://www.doi.org/10.1021/je200531k

Topological investigations of the molecular species and molecular interactions that characterize pyrrolidin-2-one + lower alkanol mixtures:	https://www.doi.org/10.1016/j.tca.2009.03.014
Solubility Modeling and Mixing Thermodynamics of Thiamphenicol in Water and Twelve Neat Organic Solvents from T = (278.15 to 318.15) K:	https://www.doi.org/10.1021/acs.jced.7b00542
Solubility Determination and Model Correlation of Pymetrozine and Evaluation of the Effect of Solvent Properties and Composition:	https://www.doi.org/10.1021/acs.jced.9b00229
Activity coefficients at infinite dilution for different alcohols and ketones in [EMpy][ESO4]: Experimental data and modeling with PC-SAFT:	https://www.doi.org/10.1016/j.fluid.2015.09.003
Experimental Measurements and Correlations of Liquid−Liquid Equilibria for the Ternary and Quaternary Mixtures Formed by Dialkyl Carbonates at T = 298.15 K:	https://www.doi.org/10.1021/je100878k
Ammonium ionic liquids in separation of water/butan-1-ol using liquid-liquid equilibrium diagrams in ternary systems:	https://www.doi.org/10.1016/j.fluid.2018.12.009
Solubility of N-phenylanthranilic acid in nine organic solvents from T = (283.15 to 318.15) K: Determination and modelling:	https://www.doi.org/10.1016/j.jct.2016.08.017
Thermochemistry of Benzyl Alcohol: Reaction Equilibria Involving Benzyl Alcohol and tert-Alkyl Ethers:	https://www.doi.org/10.1021/je049823k
Density and Viscosity of the Binary Systems Ethanol + Butan-1-ol, + Pentan-1-ol, + Heptan-1-ol, + Octan-1-ol, Nonan-1-ol, + Decan-1-ol at 0.1 MPa and Temperatures from 283.15 K to 313.15 K:	https://www.doi.org/10.1021/je400630f
Excess molar volumes and excess molar enthalpies of binary and ternary mixtures of (1-butanol + tri-n-butylamine or tri-n-octylamine) and n-hexane: Experimental results and ERAS-model calculations:	https://www.doi.org/10.1016/j.jct.2005.12.006
Study of the LLE, VLE, and VLLE of the ternary System Water + 1-Butanol + Isoamyl Alcohol at 101.3 kPa:	https://www.doi.org/10.1021/acs.jced.8b00308
Measurement and Correlation of the Solubility of Azoxystrobin in Seven Monosolvents and Two Different Binary Mixed Solvents:	https://www.doi.org/10.1021/acs.jced.7b00669
Vapor-liquid equilibria and phase densities at saturation of carbon dioxide + 1-butanol and carbon dioxide + 2-butanol from 313 to 363 K:	https://www.doi.org/10.1016/j.fluid.2010.03.042
(Vapour + liquid) equilibria (VLE), excess molar enthalpies and infinite dilution activity coefficients of selected binary systems involving n-hexyl pyridinium bis(trifluoromethylsulfonyl)imide ionic liquid: Experimental and predictions using modified UNIFAC (Dortmund):	https://www.doi.org/10.1016/j.jct.2015.06.023
Thermodynamic characterization of the mixture (1-butanol + iso-octane): Densities, viscosities, and isobaric heat capacities at high pressures:	https://www.doi.org/10.1016/j.jct.2011.08.012
Solubility Study and Mixing Property of 3,5-Dinitro-2-methylbenzoic Acid in 13 Pure Solvents from 288.15 to 333.15 K:	https://www.doi.org/10.1021/acs.jced.9b00497
Measurement and Correlation of the Solubility for 4,4′-Diaminodiphenylmethane in Different Solvents:	https://www.doi.org/10.1021/je501111d
Measurement and correlation of solubility of ciclesonide in seven pure organic solvents:	https://www.doi.org/10.1016/j.jct.2016.10.014
Determination of Solubility and Thermodynamic Properties of Benzophenone in Different Pure Solvents:	https://www.doi.org/10.1021/acs.jced.8b00196
(Liquid + liquid) equilibrium of {water + phenol + (1-butanol, or 2-butanol, or tert-butanol)} systems:	https://www.doi.org/10.1016/j.jct.2010.06.007
Liquid-liquid equilibrium data and thermophysical properties for ternary systems composed of water, acetic acid and different solvents:	https://www.doi.org/10.1016/j.fluid.2018.10.021
Measurement and correlation of solubilities of apigenin and apigenin 7-O-rhamnosylglucoside in seven solvents at different temperatures:	https://www.doi.org/10.1016/j.jct.2010.09.013
Solid-liquid equilibrium and thermodynamic of 2,5-thiophenedicarboxylic acid in different organic solvents:	https://www.doi.org/10.1016/j.fluid.2014.05.003
Solubility of Hesperetin in Various Solvents from (288.2 to 323.2) K:	https://www.doi.org/10.1021/je800078j
Solubility of lansoprazole in different solvents:	https://www.doi.org/10.1016/j.fluid.2012.06.011
Phase Equilibria Study of the Binary Systems (N-Butyl-4-methylpyridinium Tosylate Ionic Liquid + Organic Solvent, or Water):	https://www.doi.org/10.1021/je8006309
Volumetric Properties, Viscosities, and Refractive Indices of the Binary Systems 1-Butanol + PEG 200, + PEG 400, and + TEGDME:	https://www.doi.org/10.1007/s10765-013-1469-0
Densities and Viscosities of Pure 1-Methyl-3-octylimidazolium Nitrate and its Binary Mixtures with Alcohols at Several Temperatures:	https://www.doi.org/10.1021/je901058s
Liquid-Liquid Extraction of Butanol from Heptane + Butanol Mixture by Ionic Liquids:	https://www.doi.org/10.1021/acs.jced.7b00635
Thermodynamic properties of enantiotropic polymorphs of glycolide:	https://www.doi.org/10.1016/j.jct.2017.03.011
Volumetric and viscometric study with FT-IR analysis of binary systems with diethyl succinate and alcohols:	https://www.doi.org/10.1016/j.jct.2016.06.016
Viscosities and Interfacial Properties of I-Methyl-3-butylimidazolium Hexafluorophosphate and 1-Isobutenyl-3-methylimidazolium Tetrafluoroborate Ionic Liquids:	https://www.doi.org/10.1021/je0604277
Solubility of Pravastatin Sodium in Water, Methanol, Ethanol, 2-Propanol, 1-Propanol, and 1-Butanol from (278 to 333) K:	https://www.doi.org/10.1021/je800196k
(Solid + liquid) phase equilibria of tetraphenyl piperazine-1,4-diyldiphosphonate in pure solvents:	https://www.doi.org/10.1016/j.jct.2014.06.021
Viscous synergy and antagonism and isentropic compressibility of ternary mixtures containing 1,3-dioxolane, water and monoalkanols at 303.15K:	https://www.doi.org/10.1016/j.fluid.2006.02.023
Thermophysical properties of pure 1-ethyl-3-methylimidazolium methylsulfate and its binary mixtures with alcohols:	https://www.doi.org/10.1016/j.fluid.2011.03.008
Densities and Viscosities of Binary Mixtures of 1-Bromobutane with Butanol Isomers at Several Temperatures:	https://www.doi.org/10.1021/je0501235
Solubility Modeling and Mixing Properties for Benzoin in Different Monosolvents and Solvent Mixtures at the Temperature Range from 273.15 to 313.15 K:	https://www.doi.org/10.1021/acs.jced.7b00743
Thermodynamic and spectroscopic interpretation of molecular interactions of nicotine + alcohol binary mixtures:	https://www.doi.org/10.1016/j.jct.2016.07.005
Solubilities of rutin in eight solvents at T = 283.15, 298.15, 313.15, 323.15, and 333.15 K:	https://www.doi.org/10.1016/j.fluid.2007.07.030
Thermodynamic equilibrium of hydrocortisone form I in different solvent systems:	https://www.doi.org/10.1016/j.fluid.2017.09.007
Measurement and Correlation of the Solubility of Penicillin V potassium in Ethanol + Water and 1-Butyl alcohol + Water Systems:	https://www.doi.org/10.1021/je5008422
Liquid-liquid equilibrium and critical states for the system acetic acid + n-butanol + n-butyl acetate + water at 308.15 K:	https://www.doi.org/10.1016/j.fluid.2014.11.004

Solubility Modeling and Solvent Effect of 2-Amino-6-chloropurine in Twelve Neat Solvents:	https://www.doi.org/10.1021/acs.jced.8b01014
Determination of the solubility, dissolution enthalpy and entropy of 3-pentadecylphenyl acrylate in different organic solvents:	https://www.doi.org/10.1016/j.tca.2012.06.025
Solubility of Clopidogrel Hydrogen Sulfate (Form II) in Different Solvents:	https://www.doi.org/10.1021/je100022w
Thermodynamic equilibrium of hydroxyacetic acid in pure and binary solvent systems:	https://www.doi.org/10.1016/j.jct.2017.01.004
Density and viscosity of 1-butyl-3-methylimidazolium nitrate with ethanol, 1-propanol, or 1-butanol at several temperatures:	https://www.doi.org/10.1016/j.jct.2009.06.023
Solubility Measurement and Correlation of 2-Oxindole in 12 Pure Organic Solvents:	https://www.doi.org/10.1021/acs.jced.9b00308
Vapor–Liquid Equilibrium for Several Compounds Relevant to the Biofuels Industry Modeled with the Wilson Equation:	https://www.doi.org/10.1021/je400885z
Solubility Determination and Thermodynamic Modeling of 2-Mercaptobenzimidazole in 12 Solvents from T = 278.15 K to T = 318.15 K:	https://www.doi.org/10.1021/acs.jced.9b00190
Ultrasonic speeds and isentropic compressibilities of {difurylmethane + (C1-C6) n-alkanol} binary mixtures at T = 298.15 K:	https://www.doi.org/10.1016/j.jct.2010.05.016
Binary, ternary and quaternary liquid–liquid equilibria in 1-butanol, oleic acid, water and n-heptane mixtures:	https://www.doi.org/10.1016/j.fluid.2009.06.013
Thermodynamic Models for Determination of Solid-Liquid Equilibrium of the Buprofezin in Pure and Binary Organic Solvents:	https://www.doi.org/10.1021/acs.jced.9b00172
Measurement and Correlation of the Solubility of Pyrimethanil in Seven Monosolvents and Two Different Binary Mixed Solvents:	https://www.doi.org/10.1021/acs.jced.8b00124
Isobaric Vapor–Liquid Equilibria for Tetrahydropyran and Alcohol Systems:	https://www.doi.org/10.1021/je800145h
Properties of pure 1-methylimidazolium acetate ionic liquid and its binary mixtures with alcohols:	https://www.doi.org/10.1016/j.jct.2012.01.013
Solubility Determination and Thermodynamic Mixing Properties of 5-Methyl-2-pyrazinecarboxylic Acid in Different Solvents:	https://www.doi.org/10.1021/acs.jced.9b00406
Determination and Modeling of d-Histidine Solubility in Several Pure Solvents from 293.15 to 333.15 K:	https://www.doi.org/10.1021/acs.jced.9b00659
Speeds of Sound and Isentropic Compressibilities in Binary Mixtures of 2-Propanol with Several 1-Alkanols at 298.15 K:	https://www.doi.org/10.1007/s10765-005-8101-x
Determination and Correlation of Solubility and Thermodynamic Properties of trans-Cinnamyl Alcohol in Pure and Binary Solvents from 253.15 K to 293.15 K:	https://www.doi.org/10.1021/acs.jced.7b00665
Solubility of Phenacetinum in Methanol, Ethanol, 1-Propanol, 1-Butanol, 1-Pentanol, Tetrahydrofuran, Ethyl Acetate, and Benzene between 282.83 K and 333.70 K:	https://www.doi.org/10.1021/je700209v
Measurements and equation-of-state modelling of thermodynamic properties of binary mixtures of 1-butyl-1-methylpyrrolidinium tetracyanoborate ionic liquid with molecular compounds:	https://www.doi.org/10.1016/j.jct.2015.06.004
Measurements of activity coefficients at infinite dilution of aliphatic and aromatic hydrocarbons, alcohols, thiophene, tetrahydrofuran, MTBE, and water in ionic liquid [BMIM][SCN] using GLC:	https://www.doi.org/10.1016/j.jct.2008.12.018
Solubility of Ethyl p-Aminobenzoate in Six Alcohols within (283.15 to 327.15) K:	https://www.doi.org/10.1021/acs.jced.6b00037
Paramagnetic Ionic Liquid 1-Butyl-3-methylimidazolium Tetrabromidocobaltate(II): Activity Coefficients at Infinite Dilution of Organic Solutes and Crystal Structure:	https://www.doi.org/10.1021/je800846j
Densities and Excess Molar Volumes of N-Methylmorpholine + 1-Alkanol Systems at 298.15 K:	https://www.doi.org/10.1021/je7005055
Study of molecular interactions in the mixtures of some primary alcohols with equimolar mixture of 1-propanol and alkylbenzoates at T = 303.15 K:	https://www.doi.org/10.1016/j.jct.2013.09.005
Thermodynamic modelling for solubility of 4-nitrobenzaldehyde in different solvents at temperature range from (273.15 to 313.15) K and mixing properties of solutions:	https://www.doi.org/10.1016/j.jct.2016.09.012
VLE measurements using a static cell vapor phase manual sampling method accompanied with an empirical data consistency test:	https://www.doi.org/10.1016/j.jct.2014.08.024
Measurement and Correlation of Solubilities of Decanedioic Acid in C4-C6 Alcohol Solvents:	https://www.doi.org/10.1021/je900587b
Excess Molar Enthalpies of 2-Methyl-2-butanol (1) + 1-Alkanols (C1-C5) (2) at 298.15 K:	https://www.doi.org/10.1021/je101031j
Salt effect on vapor–liquid equilibria for binary systems of propanol/CaCl2 and butanol/CaCl2:	https://www.doi.org/10.1016/j.fluid.2005.07.023
Activity coefficients at infinite dilution measurements for organic solutes and water in the ionic liquid triethylsulphonium bis(trifluoromethylsulfonyl)imide:	https://www.doi.org/10.1016/j.jct.2008.12.005
Solubility Measurement and Thermodynamic Model Correlation of [illegible] in Different Solvents:	https://www.doi.org/10.1021/acs.jced.9b00232
Measurement and Correlation of Isobaric Vapor Liquid Equilibrium for Binary Systems of 1-(Ethoxymethoxy)-2-methyl-propane with Isobutyl Alcohol or 1-Butanol at 101.3 kPa:	https://www.doi.org/10.1021/acs.jced.7b00377
Solubility measurement and thermodynamic functions of dehydroepiandrosterone acetate in different solvents at evaluated temperatures:	https://www.doi.org/10.1016/j.jct.2016.01.023
Determination of partial molar excess enthalpies at infinite dilution for the systems of four alcohols + [bmim]PF6 at different temperatures by isothermal titration calorimeter:	https://www.doi.org/10.1016/j.fluid.2007.06.007
Solubility and Thermodynamic Analysis of Tenoxicam in Different Pure Solvents at Different Temperatures:	https://www.doi.org/10.1021/acs.jced.5b00382
Measurement and correlation of the solubility of (1-benzyl-1H-1,2,3-triazole-4-yl)methanol in water and alcohols at temperatures from 292.15 K to 310.15 K:	https://www.doi.org/10.1016/j.tca.2016.01.009
A 1-alkylcyanopyridinium-based ionic liquid in the separation processes:	https://www.doi.org/10.1016/j.jct.2016.01.017
Measurement and correlation of solubility of Tetracycline hydrochloride in six organic solvents:	https://www.doi.org/10.1016/j.jct.2012.08.007
Solubility and Thermodynamic Modeling of Sulfanilamide in 12 Mono Solvents and 4 Binary Solvent Mixtures from 278.15 to 318.15 K:	https://www.doi.org/10.1021/acs.jced.9b00411
Measurement and correlation of solubility of carbendazim in lower alcohols:	https://www.doi.org/10.1016/j.tca.2017.12.007
Excess molar enthalpies of ethylformate and (1-propanol, 2-propanol, 1-butanol, 2-butanol, and 1-pentanol) at T = (298.15 and 308.15) K and p = (5 and 10) MPa:	https://www.doi.org/10.1016/j.jct.2005.04.015
Osmotic and apparent molar properties of binary mixtures alcohol + 1-butyl-3-methylimidazolium trifluoromethanesulfonate ionic liquid:	https://www.doi.org/10.1016/j.jct.2013.01.032

Solubility determination and thermodynamic modelling of 2,4-dihydro-5-methyl-2-(4-methylphenyl)-3H-pyrazol-3-one in twelve organic solvents from T = (278.15 to 313.15) K and mixing properties of solutions.	https://www.doi.org/10.1016/j.jct.2016.08.007
Vapor-Liquid Equilibrium of Binary Mixtures. 1. Ethanol + 1-Butanol, Ethanol + Octane, 1-Butanol + Octane:	https://www.doi.org/10.1021/je0600045
Liquid-Liquid Equilibrium Data of Water + Butyric Acid + {Butanal or n-Butanol} Ternary Systems at 293.15, 308.15, and 323.15 K:	https://www.doi.org/10.1021/acs.jced.6b00941
Isothermal (vapour + liquid) equilibria for binary mixtures of diisopropyl ether with (methanol, or ethanol, or 1-butanol): Experimental data, correlations, and predictions:	https://www.doi.org/10.1016/j.jct.2012.11.005
Solubility determination and thermodynamic modelling of 3,5-dimethylpyrazole in nine organic solvents from T = (283.15 to 313.15) K and mixing properties of solutions:	https://www.doi.org/10.1016/j.jct.2017.02.011
Enthalpies of dilution of aqueous solutions of n-butanol, butanediols, 1,2,4-butanetriol and 1,2,3,4-butanetetrol at 298.15K:	https://www.doi.org/10.1016/j.tca.2005.06.014
Physico-chemical and excess properties of the binary mixtures of methylcyclohexane + ethanol, + propan-1-ol, + propan-2-ol, + butan-1-ol, + 2-methyl-1-propanol, or 3-methyl-1-butanol at T = (298.15, 303.15, and 308.15) K:	https://www.doi.org/10.1016/j.jct.2005.06.011
Solubility of carbon dioxide, methane, and ethane in 1-butanol and saturated liquid densities and viscosities:	https://www.doi.org/10.1016/j.jct.2013.08.007
Surface Tension and Density of Pure Ionic Liquids and Some Binary Mixtures with 1-Propanol and 1-Butanol:	https://www.doi.org/10.1021/je700621d
Liquid-liquid equilibria of aqueous systems containing alcohol and ammonium sulfate:	https://www.doi.org/10.1016/j.fluid.2011.12.023
Thermodynamic Analysis of the Experimental Equilibria for the Liquid-Phase Etherification of Isobutene with C1 to C4 Linear Primary Alcohols:	https://www.doi.org/10.1021/acs.jced.5b00557
Equilibrium Solubility, Model Correlation, and Solvent Effect of Indole-3-acetic Acid in Twelve Pure Solvents:	https://www.doi.org/10.1021/acs.jced.8b01265
Solubility of Triethylenediamine in Methanol, Ethanol, Isopropanol, 1-Butanol, and Ethyl Acetate between 276.15 K and 363.15 K:	https://www.doi.org/10.1021/je700186v
Experimental measurement and correlation of the solubility of aloe-emodin in seven pure solvents:	https://www.doi.org/10.1016/j.jct.2016.02.020
Solubility of CO2 in Alcohols, Glycols, Ethers, and Ketones at High Pressures from (288.15 to 318.15) K:	https://www.doi.org/10.1021/je101344v
Solubility of Trimethoprim (TMP) in Different Organic Solvents from (278 to 333) K:	https://www.doi.org/10.1021/je700497h
Density, Viscosity, and Isobaric Heat Capacity of the Mixture (1-Butanol + 1-Hexene):	https://www.doi.org/10.1021/je301301j
Activity Coefficients at Infinite Dilution Measurements for Organic Solutes and Water in the Ionic Liquid 1-Hexyl-3-methylimidazolium Thiocyanate:	https://www.doi.org/10.1021/je900890u
Determination of the thermodynamic parameters of ionic liquid 1-propyl-3-methylimidazolium bromide by gas-liquid chromatography:	https://www.doi.org/10.1016/j.jct.2018.09.023
Activity coefficients at infinite dilution and physicochemical properties for organic solutes and water in the ionic liquid 1-(3-hydroxypropyl)pyridinium bis(trifluoromethylsulfonyl)-amide:	https://www.doi.org/10.1016/j.jct.2011.04.018
Vapor Pressure Measurements of Liquid Solutions at Different Temperatures: Apparatus for Use over an Extended Temperature Range and Some New Data:	https://www.doi.org/10.1021/je034197x
Liquid-Liquid Equilibrium Measurements and Correlation for Ternary Systems (Butyl Acetate + 1-Butanol + Ethylene Glycol/1,3-Propanediol/Ethanolamine)	https://www.doi.org/10.1021/acs.jced.8b01219
Surface tension and interfacial compositions of binary glycerol/alcohol mixtures:	https://www.doi.org/10.1016/j.jct.2019.03.014
Solubility of Berberine Chloride in Various Solvents:	https://www.doi.org/10.1021/je0504360
FT-IR studies on excess thermodynamic properties of binary liquid mixtures o-chlorotoluene with 1-propanol, 1-butanol, 1-pentanol, 1-hexanol and 1-heptanol at different temperatures:	https://www.doi.org/10.1016/j.jct.2013.08.013
Densities, viscosities, and isobaric heat capacities of the system (1-butanol + cyclohexane) at high pressures:	https://www.doi.org/10.1016/j.jct.2014.01.020
Thermodynamics of 1,3-dimethylurea in eight alcohols:	https://www.doi.org/10.1021/acs.jced.5b00882
Isobaric Vapor Liquid Equilibria for the 1-Propanol + Ethylene Glycol Monopropyl Ether and 1-Butanol + Ethylene Glycol Monopropyl Ether Systems:	https://www.doi.org/10.1021/je2008236
Measurement and Correlation of Solubility of Amorphous Cefmetazole Sodium in Pure Solvents and Binary Solvent Mixtures:	https://www.doi.org/10.1021/acs.jced.6b00741
Solubility evaluation and thermodynamic modeling of beta-lapachone in water and ten organic solvents at different temperatures:	https://www.doi.org/10.1016/j.fluid.2018.05.005
Activity coefficients at infinite dilution and physicochemical properties for organic solutes and water in the ionic liquid trihexyl-tetradecyl-phosphonium tricyanomethanide:	https://www.doi.org/10.1016/j.jct.2018.01.003
Vapor-liquid equilibria and excess volumes of the binary systems ethanol + ethyl lactate, isopropanol + isopropyl lactate and n-butanol + n-butyl lactate at 101.325 kPa:	https://www.doi.org/10.1016/j.fluid.2005.02.015
Liquid-Liquid Equilibria of Room Temperature Ionic Liquids and Butan-1-ol:	https://www.doi.org/10.1021/je0340093
Experimental study of the PVT and CVVT properties of n-butanol in the critical region:	https://www.doi.org/10.1016/j.fluid.2012.09.026
Densities and Viscosities of Binary Liquid Mixtures of Trichloroethylene and Tetrachloroethylene with Some Polar and Nonpolar Solvents:	https://www.doi.org/10.1021/je034203p
Phase Equilibrium for the Esterification Reaction of Acetic Acid + Butan-1-ol at 101.3 kPa:	https://www.doi.org/10.1021/je700411p
Solubility of d-Aspartic Acid in Several Neat Solvents: Determination, Modeling, and Solvent Effect Analysis:	https://www.doi.org/10.1021/acs.jced.9b00320
Solubility and Solvent Effect of 1-(2-Bromo-phenyl)-pyrrole-2,5-dione in 14 Pure Solvents from 278.15 to 323.15 K:	https://www.doi.org/10.1021/acs.jced.9b00560
Solubility of Bioactive Compound Hesperidin in Six Pure Solvents at (298.15 to 338.15) K:	https://www.doi.org/10.1021/je500206w
Solubility of Acetoguanamine in Twelve Neat Solvents from 283.15 to 323.15 K:	https://www.doi.org/10.1021/acs.jced.9b00593
Thermodynamic Properties of Mixtures Containing Ionic Liquids. 5. Activity Coefficients at Infinite Dilution of Hydrocarbons, Alcohols, Esters, and Aldehydes in 1-Methyl-3-butyl-imidazolium Bis(trifluoromethyl-sulfonyl) Imide Using Gas-Liquid Chromatography:	https://www.doi.org/10.1021/je050125p
Volumetric behavior of the ternary system (methyl tert-butyl ether + chlorobenzene + butan-1-ol) and its binary sub-system (methyl tert-butyl ether + butan-1-ol) within the temperature range (298.15 to 328.15) K:	https://www.doi.org/10.1016/j.jct.2015.06.009
McGowan Method:	http://link.springer.com/article/10.1007/BF02311772

Solubility of 4-Methylsulfonylacetophenone in Nine Pure Solvents and (Ethyl Acetate + n-Hexane) Binary Solvent Mixtures	https://www.doi.org/10.1021/acs.jced.6b00617
Densities and Viscosities for Binary Mixtures of the Ionic Liquid N-Ethyl Piperazinium Propionate with n-Alcohols at Several Temperatures:	https://www.doi.org/10.1021/je201116t
Study of interaction between organic compounds and some dicationic oxygenated ionic liquids using gas chromatography:	https://www.doi.org/10.1016/j.fluid.2014.11.020
Cloud point measurements of 1-butyl-2,3-dimethylimidazolium tetrafluoroborate with alcohols:	https://www.doi.org/10.1016/j.jct.2010.07.005
Apparent Molar Volumes and Viscosity B-Coefficients of Acetyl Salicylic Acid (2-Acetoxy Benzoic Acid) Solutions in Higher Alcohols at Different Temperatures:	https://www.doi.org/10.1021/je9004646
Activity coefficients of the binary mixtures of m-cresol or p-cresol with C alcohols near ambient pressure:	https://www.doi.org/10.1016/j.fluid.2006.03.018
Measurements of activity coefficients at infinite dilution for organic solutes and water in the ionic liquid	https://www.doi.org/10.1016/j.jct.2011.11.025
Vapour–liquid equilibria of binary and ternary mixtures containing 1-butanol, 2,2,4-trimethylpentane and 1-hexene at T = 313.15 K:	https://www.doi.org/10.1016/j.jct.2013.04.003
Refractive Indices and Deviations in Refractive Indices for Binary Mixtures of Formamide + 1-Butanol, + 2-Butanol, + 1,3-Butanediol, and + 1,4-Butanediol at Temperatures from (293.15 to 318.15) K:	https://www.doi.org/10.1021/je7007632
Ternary Liquid–Liquid Equilibrium Data for the Water + Acetonitrile + ... Systems at (303.2, 323.2, 343.2) K and 1 atm:	https://www.doi.org/10.1021/je5007218
Excess Enthalpy, Density, Speed of Sound, and Viscosity for 2-Methyltetrahydrofuran + 1-Butanol at (283.15, 298.15, and 313.15) K:	https://www.doi.org/10.1021/je060017i
Solubility Determination of Nicotinamide and Its Application for the Cocrystallization with Benzoic Acid:	https://www.doi.org/10.1021/acs.jced.8b00560
Determination and modelling of troxerutin solubility in eleven mono-solvents and (1,4-dioxane + 2-propanol) binary solvents at temperatures from 288.15 K to 323.15 K:	https://www.doi.org/10.1016/j.jct.2016.09.015
Measurement and Correlation of Solubility on Reactive Crystallization of Methyl R-(-)-4-Hydroxy-phenylglycinate:	https://www.doi.org/10.1021/acs.jced.6b00415
The solubility of maleic acid in various inorganic salts and organic solvents: Thermodynamics model and parameters:	https://www.doi.org/10.1016/j.fluid.2017.07.012
Measurements of Compressed Liquid Densities for CO2 (1) + Butan-1-ol (2) via a Vibrating Tube Densimeter at Temperatures from (313 to 363) K and Pressures up to 25 MPa:	https://www.doi.org/10.1021/je700261b
Determination and Modeling of the Solubility of 2,4-Dimethoxybenzoic Acid in (Isopropanol + Ethyl Acetate Mixed Organic Solvents at Temperatures from (288.15 to ...) K:	https://www.doi.org/10.1021/je501045s
Thermophysical properties of the binary mixtures (1,8-cineole + 1-alkanol) at T = (298.15 and 313.15) K and at atmospheric pressure:	https://www.doi.org/10.1016/j.jct.2009.08.016
Solubility Determination, Modeling, and Thermodynamic Dissolution Properties of Benzenesulfonamide in 16 Neat Solvents from 273.15 to 324.45 K:	https://www.doi.org/10.1021/acs.jced.9b00360
Measurement of Diffusion Coefficients in Thermodynamically Nonideal Systems:	https://www.doi.org/10.1021/je0497303
Separation of (water/butan-1-ol) binary systems based on activity coefficients at infinite dilution with phosphonium ionic liquid:	https://www.doi.org/10.1016/j.jct.2019.05.011
Solubility measurement and modelling of ethyl 5-amino-4-cyano-3-(2-ethoxy-2-oxoethyl)-2-thiophenecarboxylate in four groups mixed solvents:	https://www.doi.org/10.1016/j.jct.2016.09.001
Liquid density of 1-butanol at pressures up to 140MPa and from 293.15K to 403.15K:	https://www.doi.org/10.1016/j.fluid.2010.11.020
Liquid–Liquid Equilibria Data of Protic Ionic Liquid (Ethyl-2-hydroxyethylammonium Propionate, Diethylammonium Propionate, or Butylammonium Propionate) with 1-Butanol + Water at 298.15 K:	https://www.doi.org/10.1021/acs.jced.8b01009
Biocompatible choline based ionic salts: Solubility in short-chain alcohols:	https://www.doi.org/10.1016/j.jct.2013.07.025
Excess molar heat capacities for (1-butanol+1,3-butanediol) at temperatures from (285 to 353) K:	https://www.doi.org/10.1016/j.jct.2004.09.007
Thermodynamic Speed of Sound Data for Liquid and Supercritical Alcohols:	https://www.doi.org/10.1021/acs.jced.8b00938
Experimental and Computational	https://www.doi.org/10.1021/acs.jced.5b00162
Thermodynamic Properties of (Benzyl alcohol + Alkanols) Mixtures:	https://www.doi.org/10.1021/je301029j
Determination and Correlation of Solubility of Tris(3-hydroxypropyl)phosphine Oxide in Selected Solvents:	https://www.doi.org/10.1021/acs.jced.7b00301
Measurement and Correlation of Solubility of Two Isomers of Cyanopyridine in Eight Pure Solvents from 288.15 to 318.15 K:	https://www.doi.org/10.1016/j.jct.2015.02.004
Solubility and solution thermodynamics of sorbic acid in eight pure organic solvents:	https://www.doi.org/10.1021/je020228x
Solubility of Benzimidazoles in Alcohols:	https://www.doi.org/10.1016/j.fluid.2011.01.003
Solubility of HFCs in lower alcohols:	https://www.doi.org/10.1016/j.jct.2004.01.001
(Vapour + liquid) equilibria of (1-butanol + benzene, or toluene, or o-, or m-, or p-xylene) at T = 308.15 K:	https://www.doi.org/10.1016/j.jct.2013.09.007
The study of activity coefficients at infinite dilution for organic solutes and water in 1-butyl-4-methylpyridinium dicyanamide, [B4MPy][DCA] using GLC:	https://www.doi.org/10.1016/j.jct.2019.03.006
Volumetric and sound speed study of aqueous 1-butanol liquid mixtures at different temperatures:	https://www.doi.org/10.1016/j.tca.2006.10.025
Excess molar enthalpies and heat capacities of dimethyl sulfoxide + seven normal alkanols at 303.15K and atmospheric pressure:	https://www.doi.org/10.1016/j.jct.2011.12.003
Effect of temperature and composition on the density, viscosity, surface tension, and thermodynamic properties of binary mixtures of N-octylisoquinolinium bis{(trifluoromethyl)sulfonyl}imide with alcohols:	https://www.doi.org/10.1021/je300529j
Biomass-Derived Platform Chemicals: Thermodynamic Studies on the Extraction of 5-Hydroxymethylfurfural from	https://www.doi.org/10.1021/acs.jced.6b00630
Solubility, Correlation, and Thermodynamic Analysis of Sorafenib Free Base and Sorafenib Tosylate in Monosolvents and Binary Solvent Mixtures:	https://www.doi.org/10.1016/j.fluid.2014.03.005
Unusual behavior of isotope effect on solubility of 1-hexyl-3-methylimidazolium bis(trifluorosulfonyl)imide with 1-alkanols:	https://www.doi.org/10.1021/acs.jced.6b00361
Measurement and Correlation of the Solubility of p-Coumaric Acid in Nine Pure and Water + Ethanol Mixed Solvents at Temperatures from 293.15 to 333.15 K:	https://www.doi.org/10.1016/j.jct.2005.07.022
(Liquid + liquid) equilibrium of (water + 2-propanol + 1-butanol + salt) systems at T = 313.15 K and T = 353.15 K: Experimental data and correlation:	

The use of ionic liquids for separation of binary hydrocarbons mixtures based on gamma infinity data measurements: https://www.doi.org/10.1016/j.jct.2018.07.024

KDB Vapor Pressure Data: https://www.cheric.org/research/kdb/hcprop/showprop.php?cmpid=821

Estimated Solubility Method: http://pubs.acs.org/doi/suppl/10.1021/ci034243x/suppl_file/ci034243xsi20040112_053635.txt

Thermodynamic and acoustic properties of binary mixtures of 1-butanol with 1,2-butanediol. The comparison with the results for 1,3-, and 1,4-butanediol: https://www.doi.org/10.1016/j.jct.2013.08.033

Solid-liquid equilibrium of sulbactam in pure solvents and binary solvent mixtures: https://www.doi.org/10.1016/j.fluid.2014.09.008

Excess enthalpies of binary and ternary mixtures containing dibutyl ether (DBE), 1-butanol, and heptane at T = 298.15 K and 313.15 K: https://www.doi.org/10.1016/j.jct.2009.07.014

Measurement and correlation of solubility of methyleneaminoacetonitrile in pure and binary solvents and thermodynamic properties of solution: https://www.doi.org/10.1016/j.jct.2019.01.015

Partition Coefficients of Organic Compounds in New Imidazolium and Tetralkylammonium Based Ionic Liquids Using Inverse Gas Chromatography: https://www.doi.org/10.1021/je9003178

Density and Derived Properties of Binary Mixtures Containing 2-(Dimethylamino)ethyl Methacrylate + Alcohols at Temperatures from T = (293.15 to 313.15) K and Pressures of up to 35 MPa: https://www.doi.org/10.1021/acs.jced.8b00975

Densities and Viscosities for Binary Liquid Mixtures of n-Undecane + 1-Propanol, 1-Butanol, 1-Pentanol, and 1-Hexanol from 283.15 to 363.15 K at 0.1 MPa: https://www.doi.org/10.1021/acs.jced.6b00121

Viscosities of Binary Mixtures of Isomeric Butanols or Isomeric Chloropropanes with 2-Methyltetrahydrofuran: https://www.doi.org/10.1021/je030167i

Fluid Phase Behavior of {1-Hexyl-3-methylimidazolium Bis(trifluoromethylsulfonyl) Imide, [C6mim][NTf2], + C2-C8 n-Alcohol} Mixtures: Liquid-Liquid Equilibrium and Excess Volumes: https://www.doi.org/10.1021/je060307z

Concentration Dependence of Surface Tension for Very Dilute Aqueous Solutions of Organic Nonelectrolytes: https://www.doi.org/10.1021/je049955d

Density and Viscosity of Binary Mixtures of Diethyl Carbonate with Alcohols at (293.15 to 363.15) K and Predictive Results by UNIFAC-VISCO Group Contribution Method: https://www.doi.org/10.1021/je0600808

Thermodynamics of risperidone and solubility in pure organic solvents: https://www.doi.org/10.1016/j.fluid.2014.04.028

Excess Molar Enthalpies of Binary Mixtures Containing 1-Methyl-3-octyl-imidazolium Tetrafluoroborate and Alcohols at T = 298.15 K: https://www.doi.org/10.1021/je400709f

High selective water/butan-1-ol separation on investigation of limiting activity coefficients with [P8,8,8,8][NTf2] ionic liquid: https://www.doi.org/10.1016/j.fluid.2017.06.001

Ultrasonic study of molecular interactions in binary mixtures of methyl acrylate with 1-alkanols (C4 to C10) at temperatures from (288.15 to 318.15) K: https://www.doi.org/10.1016/j.jct.2012.11.027

Activity coefficients at infinite dilution and physicochemical properties for organic solutes and water in the ionic liquid 4-(3-hydroxypropyl)-4-methylmorpholinium bis(trifluoromethylsulfonyl)-amide: https://www.doi.org/10.1016/j.jct.2015.02.024

Experimental Vapor-Liquid Equilibria and Thermodynamic Modeling of the Methanol + [illegible] and 1-Butanol [illegible] Azeotropic Binary Systems [illegible]: https://www.doi.org/10.1021/acs.jced.7b00766

Vapor-Liquid Equilibrium for Binary of 1-Butanol + N,N-Dimethylacetamide and Methyl Isobutyl Ketone + N,N-Dimethylacetamide at 101.3 kPa: https://www.doi.org/10.1021/acs.jced.9b00535

Liquid-Liquid Equilibrium in Systems with an Ionic Liquid: Experimental Data for an Example of the Biphasic Acid Scavenging Utilizing Ionic Liquids Process: https://www.doi.org/10.1021/je1004966

Evaluation of the Performance of Trigeminal Tricationic Ionic Liquids for Separation Problems: https://www.doi.org/10.1021/je201129y

Solubility and Data Correlation of b-Arbutin in Different Monosolvents from 283.15 to 323.15 K: https://www.doi.org/10.1021/acs.jced.9b00703

Excess molar enthalpies of diethyl malonate+ (1-butanol, 2-methyl-1-propanol, 1-pentanol, n-heptane, and ethyl acetate) at T = (288.2, 298.2, 313.2, 328.2, 338.2, and 348.2 K) and p = 101.3 kPa: https://www.doi.org/10.1016/j.fluid.2009.11.030

Isobaric heat capacity of nanostructured liquids with potential use as lubricants: https://www.doi.org/10.1016/j.jct.2018.03.031

Volumetric properties and enthalpies of solution of alcohols CkH2k+1OH (k = 1, 2, 6) in 1-methyl-3-alkylimidazolium bis(trifluoromethylsulfonyl)imide {[C1Cnim][NTf2] n = 2, 4, 6, 8, 10} ionic liquids: https://www.doi.org/10.1016/j.jct.2011.05.033

Activity Coefficients at Infinite Dilution by GLC in Alkanediamines as Stationary Phases: https://www.doi.org/10.1021/je200628n

Measurement and correlation of the solubility of 2,3,4,5-tetrabromothiophene in different solvents: https://www.doi.org/10.1016/j.jct.2012.06.005

Densities and Viscosities of Naphthalen-1-ol, Naphthalen-2-ol, and 1-Aminonaphthalene in the Solvents of Ethanol, Propan-1-ol, and Butan-1-ol: https://www.doi.org/10.1021/je201222n

Solid-Liquid Phase Equilibrium of trans-Cinnamic Acid in Several Alcohols: Measurements and Thermodynamic Modeling: https://www.doi.org/10.1021/acs.jced.5b00814

A systematic study on volumetric and transport properties of binary systems 1-propanol + n-hexadecane, 1-butanol + n-hexadecane and 1-propanol + ethyl oleate at different temperatures: https://www.doi.org/10.1016/j.fluid.2018.05.028

Activity Coefficients at Infinite Dilution for Organic Compounds Dissolved in 1-Alkyl-1-methylpyrrolidinium Bis(trifluoromethylsulfonyl)imide Ionic Liquids Having Six-, Eight-, and Ten-Carbon Alkyl Chains: https://www.doi.org/10.1021/je300692s

Thermodynamic Models for Determination of the Solubility of [illegible] in Different Solvents at Temperatures from (278.15 to 328.15) K: https://www.doi.org/10.1021/je500437m

Densities, Speeds of Sound, Excess Molar Volumes, and Excess Isentropic Compressibilities at T = (298.15 and 308.15) K for Methyl Methacrylate + 1-Alkanols (1-Butanol, 1-Pentanol, and 1-Hexanol) Systems: https://www.doi.org/10.1021/je100652b

Activity coefficients at infinite dilution measurements for organic solutes and water in the ionic liquid 1-hexyloxymethyl-3-methyl-imidazolium and 1,3-dihexyloxymethyl-imidazolium bis(trifluoromethylsulfonyl)-imide ionic liquids. The cation influence: https://www.doi.org/10.1016/j.fluid.2009.08.017

Isothermal Vapor-Liquid Equilibrium of (1-Butanol + 1,8-Cineole) at 10 Temperatures between (323.15 and 368.15) K: https://www.doi.org/10.1021/je200011e

Measurement and Correlation of the Solubility of 2,6-Dihydroxybenzoic Acid in Alcohols and Binary Solvents: https://www.doi.org/10.1021/acs.jced.6b00957

High pressure vapor-liquid equilibria for CO2 + alkanol systems and densities of n-dodecane and n-tridecane: https://www.doi.org/10.1016/j.fluid.2007.04.023

Solubility determination and thermodynamic modeling of paclobutrazol in nine organic solvents from T = (278.15 to 318.15) K and mixing properties of solutions: https://www.doi.org/10.1016/j.jct.2016.09.038

Activity coefficients at infinite dilution and physicochemical properties for organic solutes and water in the ionic liquid 1-(2-methoxyethyl)-1-methylpyrrolidinium trifluorotris(perfluoroethyl)phosphate: https://www.doi.org/10.1016/j.jct.2013.01.007

Separation of binary mixtures hexane/hex-1-ene, cyclohexane/cyclohexene and ethylbenzene/styrene based on gamma infinity data measurements: https://www.doi.org/10.1016/j.jct.2017.11.017

Densities and Viscosities for Binary Liquid Mixtures of Biodiesel + 1-Butanol, + Isobutyl Alcohol, or + 2-Butanol from 293.15 to 333.15 K at 0.1 MPa: https://www.doi.org/10.1021/acs.jced.7b00440

Solubility Modeling and Solvent Effect for Flubendazole in 12 Neat Solvents:	https://www.doi.org/10.1021/acs.jced.8b01126
Measurement and Correlation of Solubility of Azithromycin Monohydrate in Five Pure Solvents:	https://www.doi.org/10.1021/je400899e
Measurement and modelling of econazole nitrate in twelve pure organic solvents at temperatures from 273.15 K to 318.15 K:	https://www.doi.org/10.1016/j.jct.2016.07.043
Solubility of Anthracene in Binary Alcohol + Acetonitrile Solvent Mixtures at 298.2 K:	https://www.doi.org/10.1021/je0201727
Vapour pressures and osmotic coefficients of binary mixtures containing alcohol and pyrrolidinium-based ionic liquids:	https://www.doi.org/10.1016/j.jct.2013.05.027
Solubility of androstenedione in lower alcohols:	https://www.doi.org/10.1016/j.fluid.2013.11.008
Solubility and Density of 2,6-Dimethylnaphthalene in C1-C7 1-Alkanols:	https://www.doi.org/10.1021/je049967z
Infinite Dilution Activity Coefficients of Solutes Dissolved in Two Trihexyl(tetradecyl)phosphonium Ionic Liquids:	https://www.doi.org/10.1021/je500050p
ACSExcess Molar Enthalpies of Mixtures of (+-)-Linalool with Several Alkanols:	https://www.doi.org/10.1021/je200457q
Solubility behaviour and thermodynamic analysis of methyleneaminoacetonitrile in binary (ethanol + water, ethanol + 2-propanol, ethanol + n-butanol) solvents:	https://www.doi.org/10.1016/j.jct.2019.06.009
Application of the extended real associated solution (ERAS) theory to excess molar enthalpies of benzaldehyde + 1-alkanols (C1 to C5) at T = 298.15 K:	https://www.doi.org/10.1016/j.jct.2011.05.014
o-Nitrophenylacetonitrile Solubility in Several Pure Solvents: Measurement, Correlation, and Solvent Effect Analysis:	https://www.doi.org/10.1021/acs.jced.9b00243
Study of the Temperature Dependence of Surface Tensions of Some Alkanol + Hexane Mixtures:	https://www.doi.org/10.1021/je700215z
Measurement and modeling of osmotic coefficients of binary mixtures (alcohol + 1,3-dimethylpyridinium methylsulfate) at T = 323.15 K:	https://www.doi.org/10.1016/j.jct.2011.01.010
Bubble-Temperature Measurements on Some Binary Mixtures Formed by Tetrahydrofuran or Amyl Alcohol with Hydrocarbons, Chlorohydrocarbons, or Butanols at (94.6 or 95.8) kPa:	https://www.doi.org/10.1021/je030141r
NIST Webbook:	http://webbook.nist.gov/cgi/cbook.cgi?ID=C71363&Units=SI
Deduction of Physicochemical Properties from Solubilities: 2,4-Dihydroxybenzophenone, Biotin, and Caprolactam as Examples:	https://www.doi.org/10.1021/je501140p
Measurement and correlation of the solubility of maleic anhydride in different organic solvents:	https://www.doi.org/10.1016/j.fluid.2014.01.008
Measurements of activity coefficients at infinite dilution of aromatic and aliphatic hydrocarbons, alcohols, and water in the new ionic liquid [EMIM][SCN] using GLC:	https://www.doi.org/10.1016/j.jct.2008.01.004
Solubility of Phenanthrene in Binary Mixtures of C1-C4 Alcohols + 2-Propanol and Ethanol + Methanol at 298.15 K:	https://www.doi.org/10.1021/je900027h
Solubility of Anthracene in Binary Diisopropyl Ether + Alcohol Solvent Mixtures at 298.15 K:	https://www.doi.org/10.1021/je600525e
Solubility determination and thermodynamic modelling for 2-amino-4-chlorobenzoic acid in eleven organic solvents from T = (278.15 to 318.15) K and mixing properties of solutions:	https://www.doi.org/10.1016/j.jct.2016.11.019
Binary Solid-Liquid Solubility Determination and Model Correlation of Ou[illegible] Solvents:	https://www.doi.org/10.1021/acs.jced.8b01181
Vapor Liquid Equilibrium Data for Binary Systems of 1-Methyl-4-(1-methylethenyl)-cyclohexene + {Ethanol, Propan-1-ol, Propan-2-ol, Butan-1-ol, Pentan-1-ol, or Hexan-1-ol} at 40 kPa:	https://www.doi.org/10.1021/je300347z
Study of densities, viscosities, and speeds of sound of binary liquid mixtures of butan-1-ol with n-alkanes (C6, C8, and C10) at T = (298.15, 303.15, and 308.15) K:	https://www.doi.org/10.1016/j.jct.2007.05.016
Excess molar enthalpies of dimethyl carbonate and (butanol isomers and n-heptane) at T = (298.15, 313.15, and 328.15) K and p = (0.1, 1.0, and 10.0) MPa:	https://www.doi.org/10.1016/j.jct.2005.07.011
Solubility of Alcohols and Aromatic Compounds in Imidazolium-Based Ionic Liquids:	https://www.doi.org/10.1021/je800331n
Separation of Butanol Using Tetradecyl(trihexyl)phosphonium Bis(2,4,4-trimethylpentyl)phosphinate, Oleyl Alcohol, and Castor Oil:	https://www.doi.org/10.1021/acs.jced.9b00211
Volumetric and viscometric properties of binary mixtures of {methyl tert-butyl ether (MTBE) + alcohol} at several temperatures and p = 0.1 MPa: Experimental results and application of the ERAS model:	https://www.doi.org/10.1016/j.jct.2011.02.018
Thermodynamic models for determination of the solubility of [illegible] in organic solvents from T = (278.15 to 333.15) K:	https://www.doi.org/10.1016/j.jct.2015.11.005
Viscosity, density, and speed of sound of methylcyclopentane with primary and secondary alcohols at T = (293.15, 298.15 and 303.15) K:	https://www.doi.org/10.1016/j.jct.2005.12.010
Activity coefficients at infinite dilution, physicochemical and thermodynamic properties for organic solutes and water in the ionic liquid ethyl-dimethyl-(2-methoxyethyl)ammonium tris(pentafluoroethyl)trifluorophosphate:	https://www.doi.org/10.1016/j.jct.2015.05.022
Solubility and Mixing Thermodynamics Properties of p-Toluenesulfonamide and o-Toluenesulfonamide in Seven Monosolvents at Different Temperatures:	https://www.doi.org/10.1021/acs.jced.7b00714
Thermodynamics and activity coefficients at infinite dilution for organic solutes and water in the ionic liquid 1-butyl-1-methylmorpholinium tricyanomethanide:	https://www.doi.org/10.1016/j.jct.2013.08.030
Solubility and Molecular Interactions of Trimetazidine Hydrochloride in 12 Monosolvents and Binary Solvent Mixtures of Methanol + (Ethanol, N,N-Dimethylformamide or Ethyl Acetate):	https://www.doi.org/10.1021/acs.jced.8b00235
Isothermal Vapor-Liquid Equilibrium Measurements for the Pentan-2-one + Propan-1-ol System within [illegible] K:	https://www.doi.org/10.1021/acs.jced.9b00100
Activity coefficients at infinite dilution of organic solvents and water in 1-butyl-3-methylimidazolium dicyanamide. A literature review of hexane/hex-1-ene separation:	https://www.doi.org/10.1016/j.fluid.2016.02.004
Solubility Determination, Correlation, and Solute-Solvent Molecular Interactions of 5-Aminotetrazole in Various Pure Solvents:	https://www.doi.org/10.1021/acs.jced.9b00385
Solubility of oxcarbazepine in eight solvents within the temperature range T = (288.15-308.15) K:	https://www.doi.org/10.1016/j.jct.2016.09.011
Thermochemistry of Branched Ethers: Experimental Study of Chemical Equilibrium in the Reacting System of tert-Amyl Alkyl Ether Synthesis:	https://www.doi.org/10.1021/je034172y
Phase equilibria of the water + 1-butanol + toluene ternary system at 101.3 kPa:	https://www.doi.org/10.1016/j.fluid.2014.10.038
Renewable platform chemicals: Thermochemical study of levulinic acid esters:	https://www.doi.org/10.1016/j.tca.2017.12.006
Mixing Properties for the Ternary Mixture Methyl tert-Butyl Ether + 1-Butanol + Decane at 298.15 K:	https://www.doi.org/10.1021/je025579q

Solubility determination and thermodynamic dissolution functions of 1,3-diphenylguanidine in nine organic solvents at evaluated temperatures:	https://www.doi.org/10.1016/j.jct.2016.03.011
Solubility of Indinavir Sulfate in Different Solvents from (278.35 to [illegible]	https://www.doi.org/10.1021/je800689a
Solubility of N-Ethylcarbazole in different organic solvent at 279.15 - 319.15 K:	https://www.doi.org/10.1016/j.fluid.2014.06.030
Solubility and solution thermodynamics of thymol in six pure organic solvents:	https://www.doi.org/10.1016/j.jct.2015.09.010
Solubility Measurement and Correlation of Two Erlotinib Hydrochloride Polymorphs in Alcohols and Ketones from 273.15 to 323.15 K:	https://www.doi.org/10.1021/acs.jced.6b00796
Thermodynamic Properties of Polymorphs of 2,2'-Thiodiethylene Bis[3-(3,5-di-tert-butyl-4-hydroxyphenyl)propionate]:	https://www.doi.org/10.1021/je500869p
Liquid phase equilibria of (water + phosphoric acid + 1-butanol or butyl acetate) ternary systems at T = 308.2 K:	https://www.doi.org/10.1016/j.jct.2008.07.001
Quaternary liquid-liquid equilibrium, solubility and critical states: Acetic acid - n-butanol - n-butyl acetate - water at 313.15 K and atmospheric pressure:	https://www.doi.org/10.1016/j.fluid.2019.04.020
Solubility in Different Solvents, Correlation, and Solvent Effect in the Solvent Crystallization Process of Iohexol:	https://www.doi.org/10.1021/acs.jced.8b01101
Experimental and calculated liquid-liquid equilibrium data for water + furfural + solvents:	https://www.doi.org/10.1016/j.fluid.2012.07.032
Measurement of Critical Properties for Binary and Ternary Mixtures Containing n-Butanol and n-Alkane:	https://www.doi.org/10.1021/acs.jced.8b00585
Isobaric Vapor-Liquid Equilibrium Data for the 1,2-Dichloroethane-1-Butanol System:	https://www.doi.org/10.1021/acs.jced.9b00484
Liquid-Liquid Equilibrium for Ternary Systems Containing Water, Oleic Acid, and Alcohols at 313.15 K: Effect of Alcohol Chain Length:	https://www.doi.org/10.1021/je501174y
Thermodynamic Properties of Mixtures Containing Ionic Liquids. 9. Activity Coefficients at Infinite Dilution of Hydrocarbons, Alcohols, Esters, and Aldehydes in Trimethyl-butylammonium Bis(trifluoromethylsulfonyl) Imide Using Gas-Liquid Chromatography and Static Method:	https://www.doi.org/10.1021/je050440b
VLE for CO2 + glycerol + (ethanol or 1-propanol or 1-butanol):	https://www.doi.org/10.1016/j.fluid.2011.01.019
Isobaric (vapour + liquid) equilibria for the (1-propanol + 1-butanol) binary mixture at (53.3 and 91.3) kPa:	https://www.doi.org/10.1016/j.jct.2010.02.001
Effect of Solvent Properties and Composition on the Solubility of Ganciclovir Form I:	https://www.doi.org/10.1021/acs.jced.8b01080
Experimental and theoretical study of interaction between organic compounds and 1-(4-sulfobutyl)-3-methylimidazolium based ionic liquids:	https://www.doi.org/10.1016/j.fluid.2014.06.021
Solubility and solution thermodynamics of 2-methyl-4-nitroaniline in eleven organic solvents at elevated temperatures:	https://www.doi.org/10.1016/j.jct.2016.10.037
Volumetric and Viscometric Study of Binary Systems of Ethyl Butyrate with Alcohols:	https://www.doi.org/10.1021/je5005752
Volumetric properties of monoethanolamine and alcohol binary mixtures at different temperatures and 0.1 MPa:	https://www.doi.org/10.1016/j.jct.2019.01.025
Mixing thermodynamic properties of 1-butyl-4-methylpyridinium tetrafluoroborate [b4mpy][BF4] with water and with an alkan-1-ol (methanol to pentanol):	https://www.doi.org/10.1016/j.jct.2008.02.019
Thermodynamic Properties of Mixtures Containing Ionic Liquids 4. LLE of Binary Mixtures of [C2MIM][NTf2] with Propan-1-ol, Butan-1-ol, and Pentan-1-ol and [C4MIM][NTf2] with Cyclohexanol and 1,2-Hexanediol Including Studies of the Influence of [illegible]	https://www.doi.org/10.1021/je049564z
Design of a Salt-Free Aqueous Two-Phase System Containing Butanol, [illegible]	https://www.doi.org/10.1021/acs.jced.9b00305
Physicochemical properties of binary mixtures of [illegible]	https://www.doi.org/10.1016/j.jct.2016.01.012
Densities, Ultrasonic Speeds, and Excess Properties of Binary Mixtures of Diethylene Glycol with 1-Butanol, 2-Butanol, and 1,4-Butanediol at Different Temperatures:	https://www.doi.org/10.1007/s10765-015-1980-6
Isobaric Vapor-Liquid Equilibrium for the Binary Systems of 2-Butanol + 2-(Methoxymethoxy)butane and 1-Butanol + 2-(Methoxymethoxy)-butane at 101.3 kPa:	https://www.doi.org/10.1021/acs.jced.8b00499
Solubilities of Ibuprofen in Different Pure Solvents:	https://www.doi.org/10.1021/je100255z
Thermodynamics and selectivity of separation based on activity coefficients at infinite dilution of various solutes in ionic liquid [HMMIM][BF4]:	https://www.doi.org/10.1016/j.jct.2018.05.017
Measurement and correlation for solubility of levofloxacin in six solvents at temperatures from 288.15 to 328.15 K:	https://www.doi.org/10.1016/j.fluid.2012.05.027
Volumetric Behavior of Carbon Dioxide + Butan-1-ol Mixtures at 313.15 K:	https://www.doi.org/10.1021/je100788g
The liquid-liquid coexistence curves of {x 1-octyl-3-methylimidazolium hexafluorophosphate + (1-x) 1-butanol} and {x 1-octyl-3-methylimidazolium hexafluorophosphate + (1 - x) 2-butanol} in the critical region:	https://www.doi.org/10.1016/j.fluid.2013.03.020
Activity Coefficients at Infinite Dilution Measurements for Organic Solutes and Water in the Ionic Liquid 1-Butyl-1-methylpyrrolidinium Trifluoromethanesulfonate:	https://www.doi.org/10.1021/je1000582
Osmotic coefficients of alcoholic mixtures containing BMpyrDCA: Experimental determination and correlation:	https://www.doi.org/10.1016/j.jct.2013.12.002
Vapor-liquid equilibria for the binary mixtures of 2,3-butanediol with n-butanol, n-butyl acetate, and ethyl acetate at 101.3 kPa:	https://www.doi.org/10.1016/j.fluid.2007.09.001
Measurement, correlation and dissolution thermodynamics of biological active chalcone in organic solvents at different temperatures:	https://www.doi.org/10.1016/j.jct.2014.03.026
Equilibrium study and diversified models of drug Norfloxacin in eight pure organic and binary solvents at T = (278.15 to 328.15) K:	https://www.doi.org/10.1016/j.fluid.2016.12.012
Solubility and dissolution characteristics of capecitabine in pure lower alcohols and water with methanol mixture solvents at atmospheric pressure and different temperatures:	https://www.doi.org/10.1016/j.fluid.2017.12.024
Experimental and theoretically study of interaction between organic compounds and tricyanomethanide based ionic liquids:	https://www.doi.org/10.1016/j.jct.2014.12.027
Liquid-liquid equilibria for water + 1-propanol (or 1-butanol) + potassium chloride + ammonium chloride quaternary systems at 298.15 K:	https://www.doi.org/10.1016/j.fluid.2015.03.050
Solubility and Thermodynamic Properties of (2S)-Pyrrolidine-2-carboxylic Acid in Water, Alcohols, and Water-Alcohol Mixtures:	https://www.doi.org/10.1021/je5007949
Thermodynamic properties of (tetradecane + benzene, + toluene, + chlorobenzene, + bromobenzene, + anisole) binary mixtures at T = (298.15, 303.15, and 308.15) K:	https://www.doi.org/10.1016/j.jct.2005.10.022
Solubility and Solution Thermodynamics of 2,6-Dichloro-4-nitroaniline in 12 Pure Solvents at Temperatures from 278.15 to 323.15 K:	https://www.doi.org/10.1021/acs.jced.9b00778

Thermodynamics and limiting activity coefficients measurements for organic solutes and water in the ionic liquid 1-dodecyl-3-methylimidazolium bis(trifluoromethylsulfonyl) imide:	https://www.doi.org/10.1016/j.jct.2016.08.008
Vapor–Liquid Equilibrium of Binary Mixtures Containing Isopropyl Acetate and Alkanols at 101.30 kPa:	https://www.doi.org/10.1021/acs.jced.5b00360
Ternary (liquid + liquid) equilibria of {bis(trifluoromethylsulfonyl)-amide based ionic liquids + butan-1-ol + water}:	https://www.doi.org/10.1016/j.jct.2015.11.002
Solubility of 1-methyl-4-nitropyrazole in seventeen pure solvents at temperatures from 283.15 K to 323.15 K:	https://www.doi.org/10.1016/j.fluid.2018.05.023
Isobaric Vapor–Liquid–Liquid Phase Equilibria Measurements of Three Ternary Water + 2-Butanone + Aliphatic Alcohol (Ethanol, 1-Propanol, 2-Propanol) Systems at 101.3 kPa:	https://www.doi.org/10.1021/acs.jced.6b00725
Solubility, Solubility Modeling, and Antisolvent Precipitation of 1,3-Bis(9-carbazolyl)benzene in Organic Solvents:	https://www.doi.org/10.1021/acs.jced.9b00432
Vapour pressures, osmotic and activity coefficients for binary mixtures containing (1-ethylpyridinium ethylsulfate + several alcohols) at T = 323.15 K:	https://www.doi.org/10.1016/j.jct.2009.12.001
Solubility measurement and thermodynamic model correlation of baricitinib in twelve pure solvents:	https://www.doi.org/10.1021/acs.jced.9b00693
Thermophysical properties of 1,2,4-trimethylbenzene in admixtures with 1-butanol or 2-butanol at high pressures:	https://www.doi.org/10.1016/j.jct.2017.03.006
Liquid–Liquid Phase Equilibria of the Binary Ionic Liquid Systems [Cn MIM][NTf2] + n-Butanol, + n-Pentanol, + H2O Using UV Spectroscopic and Densimetric Analytical Methods:	https://www.doi.org/10.1021/je300672q
Experimental and Predicted Results of Anomeric Equilibrium of Glucose in Alcohols:	https://www.doi.org/10.1021/je050193b
Measurement of activity coefficients at infinite dilution in 1-hexadecyl-3-methylimidazolium tetrafluoroborate ionic liquid:	https://www.doi.org/10.1016/j.jct.2007.01.004
Partition Coefficients of Organic Compounds in Four New Tetraalkylammonium Bis(trifluoromethylsulfonyl)imide Ionic Liquids Using Inverse Gas Chromatography:	https://www.doi.org/10.1021/je200637v
Liquid–Liquid Equilibrium in Mixtures of the Ionic Liquid 1-n-Butyl-3-methylimidazolium Hexafluorophosphate and an Alkanol:	https://www.doi.org/10.1021/je050474j
Isobaric vapour–liquid equilibria for the binary systems 4-methyl-2-pentanone + 1-butanol and + 2-butanol at 20 and 101.3 kPa:	https://www.doi.org/10.1016/j.fluid.2008.11.012
Thermodynamics and activity coefficients at infinite dilution for organic solutes, water and diols in the ionic liquid choline bis(trifluoromethylsulfonyl)imide:	https://www.doi.org/10.1016/j.jct.2014.04.024
Liquid–Liquid Phase Behavior of Solutions of 1,3-Dimethylimidazolium- and 1-Methyl-3-ethylimidazolium Bis(trifluoromethyl)sulfonyl)amide (C1,3mimNTf2) in n-Alkyl Alcohols:	https://www.doi.org/10.1021/je400253e
Isobaric Vapor Liquid Equilibria of 1-Butanol p-Xylene System:	https://www.doi.org/10.1007/s10765-006-0018-5
Measurement and Correlation of Solubility of Clopidogrel Hydrogen Sulfate (Metastable Form) in Lower Alcohols:	https://www.doi.org/10.1021/je200067e
The co-solubility of 2-ethylhexanoic acid and some liquid alcohols in supercritical carbon dioxide:	https://www.doi.org/10.1016/j.fluid.2005.09.023
Thermodynamics of Phase and Chemical Equilibrium in a Strongly Nonideal Esterification System:	https://www.doi.org/10.1021/je0498199
Determination and correlation of the solubility of four Bronsted-acidic ionic liquids based on benzothiazolium cations in six alcohols:	https://www.doi.org/10.1016/j.jct.2013.12.031
Solubility of Podophyllotoxin in Six Organic Solvents from (283.2 to 323.2) K:	https://www.doi.org/10.1021/je8007028
Experimental solid-liquid phase equilibria of {cholesterol + binary solvent mixture: 1-Alcohol (C4-C10) + cyclohexane}:	https://www.doi.org/10.1016/j.fluid.2009.10.009
Excess molar enthalpies of methyl isobutyl ketone (MIBK) with alkan-1-ols (C1-C6) and their correlations at 298.15 K:	https://www.doi.org/10.1016/j.tca.2004.09.012
Solubility Determination and Thermodynamic Models for 2-Methylnaphthalene in Different Solvents from T = (278.15 to 303.15) K:	https://www.doi.org/10.1021/je5010627
Solubility Modeling of 4-(Methylsulfonyl)benzaldehyde in Nine Organic Solvents at Elevated Temperatures:	https://www.doi.org/10.1021/acs.jced.5b01053
Ether + alcohol + hydrocarbon mixtures in fuels and bio-fuels: Excess enthalpies of binary mixtures containing dibutyl ether (DBE) or 1-butanol and 1-hexene or methylcyclohexane or toluene or cyclohexane or 2,2,4-trimethylpentane at 298.15 K and 313.15 K:	https://www.doi.org/10.1016/j.fluid.2011.11.005
Densities and Viscosities of Ternary System n-Dodecane (1) + Bicyclohexyl (2) + Butanol (3) and Corresponding Binaries at T = (293.15 to 343.15) K:	https://www.doi.org/10.1021/acs.jced.8b00438
Isothermal (liquid + liquid) equilibrium data at T = 313.15 K and isobaric (vapor + liquid + liquid) equilibrium data at 101.3 kPa for the ternary system (water + 1-butanol + p-xylene):	https://www.doi.org/10.1016/j.jct.2014.07.021
Experimental and theoretical study on liquid-liquid equilibrium of 1-butanol + water + NMP at 25 and 35 deg C:	https://www.doi.org/10.1016/j.fluid.2010.09.013
Investigation of the vapor-liquid equilibrium for 2-propanol + 1-butanol + 1-pentanol at 101.3 kPa:	https://www.doi.org/10.1016/j.fluid.2012.12.020
Phase behaviour of ionic liquid 1-butyl-1-methylpyrrolidinium tris(pentafluoroethyl)trifluorophosphate with alcohols, water and aromatic hydrocarbons:	https://www.doi.org/10.1016/j.fluid.2013.01.027
Solubility Measurement and Correlation of Probenecid in 12 Pure Organic Solvents and Thermodynamic Properties of Mixing of Solutions:	https://www.doi.org/10.1021/acs.jced.8b00863
Biobutanol and n-propanol recovery using a low density phosphonium based ionic liquid at T = 298.15 K and p = 1 atm:	https://www.doi.org/10.1016/j.fluid.2013.06.047
Thermogravimetric measurement of evaporation: Data analysis based on the Stefan tube:	https://www.doi.org/10.1016/j.tca.2014.09.011
Assessment of Pyrrolidinium-Based Ionic Liquid for the Separation of Binary Mixtures Based on Activity Coefficients at Infinite Dilution:	https://www.doi.org/10.1021/acs.jced.9b00341
Separation of aliphatic from aromatic hydrocarbons and sulphur compounds from fuel based on measurements of activity coefficients at infinite dilution for organic solutes and water in the ionic liquid N,N-diethyl-N-methyl-N-(2-methoxy-ethyl)ammonium bis(trifluoromethylsulfonyl)imide:	https://www.doi.org/10.1016/j.jct.2016.07.017
Excess Enthalpy, Density, Viscosity, and Speed of Sound for the Mixture Tetrahydropyran + 1-Butanol at (283.15, 298.15, and 313.15) K:	https://www.doi.org/10.1021/je0498457
Thermodynamic Functions for the Solubility of 3-Nitrobenzonitrile in 12 Organic Solvents from T/K = (278.15 to 318.15):	https://www.doi.org/10.1021/acs.jced.7b00615
Thermodynamic Study of Solubility for Imatinib Mesylate in Nine Monosolvents and Two Binary Solvent Mixtures from 278.15 to 318.15 K:	https://www.doi.org/10.1021/acs.jced.8b00551
Solubility of Salicylamide in Alcohols from (283.15 to 308.15) K:	https://www.doi.org/10.1021/je800668j
Study of Ether-, Alcohol-, or Cyano-Functionalized Ionic Liquids Using Inverse Gas Chromatography:	https://www.doi.org/10.1021/je900838a

Solubility of d-Tryptophan and l-Tyrosine in Several Organic Solvents: Determination and Solvent Effect:	https://www.doi.org/10.1021/acs.jced.9b00258
Physicochemical properties and activity coefficients at infinite dilution for organic solutes and water in a novel bicyclic guanidinium superbase-derived protic ionic liquid:	https://www.doi.org/10.1016/j.jct.2012.09.033
Solubilities of Adipic Acid and Succinic Acid in Glutaric Acid + Acetone or	https://www.doi.org/10.1021/je500682v
Investigation on molecular interactions of antibiotics in alcohols using volumetric and acoustic studies at different temperatures:	https://www.doi.org/10.1016/j.jct.2016.09.037
Revision of the volumetric method for measurements of liquid liquid equilibria in binary systems:	https://www.doi.org/10.1016/j.fluid.2004.12.004
Measurement and Correlation of the Solubility of e-CL-20 in 12 Organic Solvents at Temperatures Ranging from 283.15 to 318.15 K:	https://www.doi.org/10.1021/acs.jced.6b00761
Speed of sound, density and related thermodynamic excess properties of binary mixtures of butan-2-one with C1–C4 alkanols and chloroform:	https://www.doi.org/10.1021/je5007604
Liquid liquid equilibria of water + 1-butanol + methyl methacrylate or butyl methacrylate or isobutyl methacrylate at 288.15K and 318.15K:	https://www.doi.org/10.1016/j.fluid.2007.07.011
Solubility and solution thermodynamics of ethyl 5-amino-4-cyano-3-(2-ethoxy-2-oxoethyl)-2-thiophenecarboxylate in nine organic solvents at evaluated temperatures:	https://www.doi.org/10.1016/j.jct.2015.11.025
Determination and correlation of solubility of N-methyl-3,4,5-trinitropyrazole (MTNP) in ten pure solvents from 283.15 K to 323.15 K:	https://www.doi.org/10.1016/j.fluid.2017.04.008
Thermodynamic Properties of dodecane + 1-butanol and 2-butanol systems:	https://www.doi.org/10.1021/je0497810
Measurement and correlation of the solubility of N-(phosphonomethyl)iminodiacetic acid in different organic solvents:	https://www.doi.org/10.1016/j.fluid.2014.08.036
Activity Coefficients at Infinite Dilution for Hydrocarbons in Fatty Alcohols Determined by Gas-Liquid Chromatography:	https://www.doi.org/10.1021/je1005517
Liquid-liquid equilibrium for the ternary system of 1-butanol + 3-methyl-1-butanol + water at different temperatures:	https://www.doi.org/10.1016/j.fluid.2012.08.012
Binary Equilibrium Solubility of Amidinothiourea in Monosolvents: Experimental Determination, Model Correlation, and Solvent Effect Analysis:	https://www.doi.org/10.1021/acs.jced.9b00458
Solubility of Bis(benzoxazolyl-2-methyl) Sulfide in Different Pure Solvents and Ethanol + Water Binary Mixtures between (273.25 and 325.20) K:	https://www.doi.org/10.1021/je8005419
Phase Behavior of Binary Mixtures Containing Succinic Acid or Its Esters:	https://www.doi.org/10.1021/acs.jced.7b00005
Activity coefficients at infinite dilution for organic solutes and water in 1-ethyl-1-methylpyrrolidinium lactate:	https://www.doi.org/10.1016/j.jct.2015.05.014
Densities and Viscosities of Binary Solutions of Benzene-1,3-diol + Water, Ethanol, Propan-1-ol, and Butan-1-ol at T = (293.15 to 333.15) K:	https://www.doi.org/10.1021/je101173w
Solubility of Phenanthrene in Binary Mixtures of C1-C4 Alcohols at 298.2 K:	https://www.doi.org/10.1021/je900599q
Measurement and correlation of isobaric vapour-liquid equilibria for binary systems of 1-(methoxymethoxy)-2-methylpropane with 1-butanol and isobutanol at 101.33 kPa:	https://www.doi.org/10.1016/j.jct.2017.08.006
Properties of pure 1,1,3,3-tetramethylguanidine imidazole ionic liquid and its binary mixtures with alcohols at T = (293.15 to 318.15) K:	https://www.doi.org/10.1016/j.jct.2015.04.027
Measurement and Correlation of the Solubility of 5-Fluorouracil in Pure and Binary Solvents:	https://www.doi.org/10.1021/acs.jced.8b00425
Solubility and thermodynamic function of a bioactive compound bergenin in various pharmaceutically acceptable neat solvents at different temperatures:	https://www.doi.org/10.1016/j.jct.2016.05.003
Solubility of Baicalein in Different Solvents from (287 to 323) K:	https://www.doi.org/10.1007/s10765-014-1740-z
Isobaric vapor-liquid equilibrium, density and speed of sound of binary mixtures 2,2,4-trimethylpentane + 1-butanol or dibutyl ether (DBE) at 101.3 kPa:	https://www.doi.org/10.1016/j.fluid.2018.07.027
Determination and Correlation of Excess Molar Enthalpies of Eight Binary Systems Containing Acetophenone at Different Temperatures:	https://www.doi.org/10.1021/je800193a
Molecular dynamic simulation, molecular interactions and structural properties of 1-butyl-3-methylimidazolium bis(trifluoromethylsulfonyl)imide + 2-methyl-1-propanol mixtures at (298.15-323.15) K and 0.1 MPa:	https://www.doi.org/10.1016/j.fluid.2018.04.025
Molar heat capacities for (1-butanol + 1,4-butanediol, 2,3-butanediol, 1,2-butanediol, and 2-methyl-2,4-pentanediol) as function of temperature:	https://www.doi.org/10.1016/j.jct.2007.04.011
Activity Coefficients at Infinite Dilution of Organic Compounds in Trihexyl(tetradecyl)phosphonium Bis(trifluoromethylsulfonyl)imide Using Inverse Gas Chromatography:	https://www.doi.org/10.1021/je800754w
Determination and correlation of solubility of 4'-bromomethyl-2-cyanobiphenyl in acetone + (ethanol, n-propanol, n-butanol) mixtures:	https://www.doi.org/10.1016/j.jct.2016.06.032
Solubility Determination and Thermodynamic Modeling of … in Different Solvents and Mixing Properties of Solutions:	https://www.doi.org/10.1021/acs.jced.8b01099
Solubilities of Rutaecarpine in Twelve Organic Solvents from (283.2 to 323.2) K:	https://www.doi.org/10.1021/je4001334
Separation of binary mixtures based on gamma infinity data using [OMMIM][BF4] ionic liquid and modelling of thermodynamic functions:	https://www.doi.org/10.1016/j.jct.2018.09.003
Application of trihexyltetradecylphosphonium dicyanamide ionic liquid for various types of separations problems: Activity coefficients at infinite dilution measurements using GLC method:	https://www.doi.org/10.1016/j.fluid.2019.03.023
Measurements of activity coefficients at infinite dilution of organic solutes and water in 1-propyl-1-methylpiperidinium bis(trifluoromethylsulfonyl)imide ionic liquid using g.l.c.:	https://www.doi.org/10.1016/j.jct.2010.05.017
Solubility, Model Correlation, and Solvent Effect of …	https://www.doi.org/10.1021/acs.jced.8b01226
Volumetric properties of the boldine + alcohol mixtures at atmospheric pressure from 288.15 to 333.15K A new method for the determination of the density of pure boldine:	https://www.doi.org/10.1016/j.fluid.2007.04.030
Ebulliometric determination and prediction of (vapor + liquid) equilibria for binary and ternary mixtures containing alcohols (C1–C4) and dimethyl carbonate:	https://www.doi.org/10.1016/j.jct.2011.08.006
Thermodynamics and selectivity of separation based on activity coefficients at infinite dilution of various solutes in 1-allyl-3-methylimidazolium bis(trifluoromethylsulfonyl)imide ionic liquid:	https://www.doi.org/10.1016/j.jct.2016.06.028
Compressed Liquid Densities of Binary Mixtures of 1-Butanol and Diethylene Glycol Dimethyl Ether from (283 to 363) K at Pressures up to 100 MPa:	https://www.doi.org/10.1021/acs.jced.7b00176
Solubility of Andrographolide in Various Solvents from (288.2 to 323.2) K:	https://www.doi.org/10.1021/je100344z
Activity Coefficients at Infinite Dilution of Organic Compounds in 1-Butyl-3-methylimidazolium Tetrafluoroborate Using Inverse Gas Chromatography:	https://www.doi.org/10.1021/je800658v

Mutual diffusion behavior of short chain alcohols + n-octane mixtures:	https://www.doi.org/10.1016/j.tca.2015.11.024
Determination and Correlation of Solubility of Phenylbutazone in Monosolvents and Binary Solvent Mixtures:	https://www.doi.org/10.1021/acs.jced.6b00911
Solubilities of crotonic acid in different organic solvents at several temperatures from (278.15 to 383.15) K:	https://www.doi.org/10.1016/j.fluid.2013.06.038
Thermophysical Properties of the Pure Ionic Liquid 1-Butyl-1-methylpyrrolidinium Dicyanamide and Its Binary Mixtures with Alcohols:	https://www.doi.org/10.1021/je300384g
Solubility of Artemisinin in Different Single and Binary Solvent Mixtures Between (284.15 and 323.15) K and NRTL Interaction Parameters:	https://www.doi.org/10.1021/je100125x
Research on the application of 7-chloro quinaldine adducts in 7-chloroquinaldine separation process:	https://www.doi.org/10.1016/j.jct.2019.02.002
Solubility of Acenaphthylene in Different Solvents between (278 and 323) K:	https://www.doi.org/10.1021/je700347h
Study of the influence of the structure of the alcohol on vapor pressures and osmotic coefficients of binary mixtures alcohol + 1-hexyl-3-methylimidazolium bis(trifluoromethylsulfonyl)imide at T = 323.15 K:	https://www.doi.org/10.1016/j.fluid.2011.09.037
Excess enthalpies of ether + alcohol + hydrocarbon mixtures: Binary and ternary mixtures containing dibutyl ether (DBE), 1-butanol and benzene at 298.15K and 313.15K:	https://www.doi.org/10.1016/j.fluid.2009.06.015
Densities, Viscosities, and Excess Molar Enthalpies of 2-Pyrrolidone + Butanol Isomers at (293.15, 298.15, and 303.15) K:	https://www.doi.org/10.1021/je801006q
Isobaric Vapor Liquid Equilibria for Tertiary Amine and n-Alkane/Alkanol Binary Mixtures: Experimental Measurements and Modeling with GC-PPC-SAFT:	https://www.doi.org/10.1021/je300568h
Viscosities and Densities of Binary Mixtures of 1,4-Dioxane, Carbon Tetrachloride, and Butanol at 303.15 K, 308.15 K, and 313.15 K:	https://www.doi.org/10.1021/je030133a
Excess Molar Volumes and Viscosity Deviations of Binary Mixtures of 2,4,6-Trimethyl-1,3,5-trioxane + Ethanol, 1-Propanol, and 1-Butanol at (298.15, 303.15, and 308.15) K:	https://www.doi.org/10.1021/je049556i
P,T-Data for 1-Butanol and Isobutyl Alcohol from (283.15 to 363.15) K at Pressures up to 30 MPa:	https://www.doi.org/10.1021/je5010214
(Vapor + liquid) equilibria of the binary mixtures of m-cresol with C1 C4 aliphatic alcohols at 95.5 kPa:	https://www.doi.org/10.1016/j.jct.2006.03.006
Experimental Determination and Thermodynamic Models for Solid-Liquid Equilibrium of 4-(4-Aminophenoxy)-N-methylpyridine-2-carboxamide in Pure and Binary Solvent Mixtures for T = (278.15 to 318.15) K:	https://www.doi.org/10.1021/acs.jced.8b00165
Activity coefficients at infinite dilution measurements for organic solutes and water in the ionic liquid 1-butyl-1-methylpyrrolidinium trifluoromethanesulfonate using GLC:	https://www.doi.org/10.1016/j.fluid.2009.01.011
Tie-line data for the (water + butyric acid + n-butyl alcohol or amyl alcohol) and (water + butyric acid + isoamyl alcohol) at T = 298.2 K:	https://www.doi.org/10.1016/j.jct.2013.11.028
Solubility and Thermodynamic Properties of Cinnarizine:	https://www.doi.org/10.1021/acs.jced.5b00075
Solubility of Cefazolin Sodium Pentahydrate in Different Solvents between 278 K and 310 K:	https://www.doi.org/10.1021/je050242k
Isobaric vapor-liquid-liquid-solid equilibrium of the water + NaCl + 1-butanol system at 101.3 kPa:	https://www.doi.org/10.1016/j.jct.2016.04.013
Vapor Liquid Equilibrium Data for 1-Methyl-2-Pyrrolidone + (1-Butanol or 1-Hexene or Water) Binary Mixtures:	https://www.doi.org/10.1021/je500092v
Densities and Viscosities of Binary Mixtures of n-Butanol with 2-Butanol, Isobutanol, and tert-Butanol from (303.15 to 343.15) K:	https://www.doi.org/10.1021/je900722m
Experimental and theoretical excess molar properties of imidazolium based ionic liquids with isomers of butanol:	https://www.doi.org/10.1016/j.tca.2016.03.026
Measurements of activity coefficients at infinite dilution for organic solutes and water in N-hexylisoquinolinium thiocyanate, [HiQuin][SCN] using GLC:	https://www.doi.org/10.1016/j.jct.2013.02.004
Combined effects of salinity and temperature on the solubility of organic compounds:	https://www.doi.org/10.1016/j.jct.2011.11.020
Thermodynamic equilibrium of hexamethylenediammonium glutarate in mono-solvents and binary solvent systems:	https://www.doi.org/10.1016/j.fluid.2018.12.008
Dissolution behavior and thermodynamic properties of lapatinib ditosylate in pure and mixed organic solvents from T = (283.15 to 323.15) K:	https://www.doi.org/10.1016/j.fluid.2019.01.005
Activity coefficients at infinite dilution and physicochemical properties for organic solutes and water in the ionic liquid 4-(2-methoxyethyl)-4-methylmorpholinium trifluorotris(perfluoroethyl)phosphate:	https://www.doi.org/10.1016/j.jct.2012.05.017
Ternary Liquid-Liquid Equilibrium of Azeotropes (Ester + Alcohol) with Different Ionic Liquids at T = 298.15 K:	https://www.doi.org/10.1021/acs.jced.6b00811
Activity coefficient at infinite dilution measurements for organic solutes (polar and non-polar) in fatty compounds: Saturated fatty acids:	https://www.doi.org/10.1016/j.jct.2012.06.009
Electrical Conductivity of Caprolactam Tetrabutylammonium Bromide Ionic Liquids in Aqueous and Alcohol Binary Systems:	https://www.doi.org/10.1021/je100361s
Crippen Method:	http://pubs.acs.org/doi/abs/10.1021/ci990307l
Research on Dissolution Capability of Several Antofloxacin Salts:	https://www.doi.org/10.1021/acs.jced.8b00288
Measurement and Correlation of the Solubility of Vanillic Acid in Eight Pure and Water + Ethanol Mixed Solvents at Temperatures from (293.15 to 323.15) K:	https://www.doi.org/10.1021/acs.jced.5b00619
Solubility of CF4 in lower alcohols:	https://www.doi.org/10.1016/j.fluid.2010.02.007
Isobaric Vapor Liquid Equilibrium for Three Binary Systems of Acetaldehyde + Ethanol, Ethyl Acetate, 1-Butanol at 101.3 kPa:	https://www.doi.org/10.1021/acs.jced.7b00231
Liquid Liquid, Vapor Liquid and Vapor Liquid Liquid Equilibrium Data for the Water n-Butanol Cyclohexane System at Atmospheric Pressure: Experimental Determination and Correlation:	https://www.doi.org/10.1021/je400501z
Density and Surface Tension of Binary Mixtures of Acetonitrile + 1-Alkanol at 293.15 K:	https://www.doi.org/10.1021/je050519g
Phase equilibrium properties of binary and ternary systems containing di-isopropyl ether + 1-butanol + benzene at 313.15 K:	https://www.doi.org/10.1016/j.jct.2005.07.004
Solubility modelling and thermodynamic dissolution functions of phthalimide in ten organic solvents:	https://www.doi.org/10.1016/j.jct.2015.10.024
Isobaric vapor-liquid equilibrium for the binary systems of 1-butanol +Poly methoxy methylal (n=2), and 1-butanol +Poly methoxy methylal (n=3) at 101.3 kPa:	https://www.doi.org/10.1016/j.fluid.2014.06.019
Liquid-Liquid Equilibria Determination and Prediction for Ternary Systems of Water + 1,6-Hexanediamine + (1-Butanol or 1-Pentanol) at 298.2, 308.2, and 318.2 K:	https://www.doi.org/10.1021/acs.jced.7b00678

Solubility of Genistein in Water, Methanol, Ethanol, Propan-2-ol, 1-Butanol, and Ethyl Acetate from (280 to 333) K	https://www.doi.org/10.1021/je100261w
Measurements of activity coefficients at infinite dilution for organic solutes and water in the ionic liquid 1-butyl-1-methylpyrrolidinium tricyanomethanide	https://www.doi.org/10.1016/j.jct.2013.07.004
Thermodynamic interactions in binary mixtures of anisole with ethanol, propan-1-ol, propan-2-ol, butan-1-ol, butan-2-ol, and 2-methylpropan-1-ol at T = (293.15, 303.15, and 313.15) K	https://www.doi.org/10.1016/j.jct.2006.03.018
Thermodynamic and Spectroscopic Study of the Ternary System [illegible] furan + 1-Butanol at T = (288.15 to 323.15) K	https://www.doi.org/10.1021/je4003916
Excess molar enthalpies of ethane-1,2-diamine plus primary and secondary alkanols (C1-C4) and correlation with Redlich-Kister, Wilson, NRTL and UNIQUAC models at T = 298.15 K	https://www.doi.org/10.1016/j.jct.2016.12.036
(Liquid + liquid) equilibria of four alcohol-water systems containing [illegible]	https://www.doi.org/10.1016/j.jct.2016.06.010
Solubility of Veratric Acid in Eight Monosolvents and Ethanol + 1-Butanol at Various Temperatures	https://www.doi.org/10.1021/je400029t
Estimation and confirmation of the thermodynamic stability relationships of the enantiotropic polymorphs of glycolide	https://www.doi.org/10.1016/j.jct.2017.10.011
Determination and Thermodynamic Analysis of the Solubility of Limonin in Eight Organic Solvents and Ethyl Acetate + 2-Propanol Binary Solvents from 282.2 to 323.2 K	https://www.doi.org/10.1021/acs.jced.7b00851
Solubility determination and thermodynamic functions of [illegible] in nine organic solvents from T = 283.15 K to T = 323.15 K [illegible]	https://www.doi.org/10.1016/j.jct.2016.11.029
Apparent Molar Volume and Compressibility of [illegible] in Various Solvents	https://www.doi.org/10.1021/je900888e
Thermodynamic functions for solubility of 2-mercaptobenzothiazole in eleven pure organic solvents at temperatures from 273.15 K to 318.15 K and mixing properties of solutions	https://www.doi.org/10.1016/j.jct.2017.05.013
Density and Viscosity for Ethyl 3-Ethoxypropionate + Methacrylic Acid, [illegible]	https://www.doi.org/10.1021/je050223a
Solubility and Thermodynamic Modeling of Dimethyl Terephthalate in Pure Solvents and the Evaluation of the Mixing Properties of the Solutions	https://www.doi.org/10.1021/acs.jced.9b00658
Binary Liquid-Liquid Equilibrium (LLE) for Methyl tert-Amyl Ether (TAME) + Water from (288.15 to 318.15) K and Ternary LLE for Systems of TAME + C1-C4 Alcohols + Water at 298.15 K	https://www.doi.org/10.1021/je8006698
Measurement of infinite dilution activity coefficients of 1-alkanols (C1-C5) in [illegible] (273 [illegible]) K	https://www.doi.org/10.1016/j.fluid.2005.08.018
Excess molar volumes of ternary mixtures of 1,3-dichlorobenzene and methyl ethyl ketone with 1-alkanols at 303.15 K	https://www.doi.org/10.1016/j.fluid.2006.04.010
Liquid-liquid equilibrium in systems used for the production of 5-hydroxymethylfurfural from biomass using alcohols as solvents	https://www.doi.org/10.1016/j.jct.2017.03.019
Extraction of n-Alcohols from n-Heptane Using Ionic Liquids	https://www.doi.org/10.1021/je200481u
Solubility and Thermodynamic Behavior of Syringic Acid in Eight Pure and Water + Methanol Mixed Solvents	https://www.doi.org/10.1021/acs.jced.7b00333
Pressurized Synthetic Methodology for Solubility Determination at Elevated Temperatures with Application to Paracetamol in Pure Solvents	https://www.doi.org/10.1021/acs.jced.7b00118
Thermodynamic and spectroscopic properties of binary mixtures of n-butylammonium butanoate ionic liquid with alcohols at T = (293.15–313.15) K	https://www.doi.org/10.1016/j.jct.2016.09.007
Densities, speeds of sound, and refractive indices for binary mixtures of 1-butyl-3-methylimidazolium methyl sulphate ionic liquid with alcohols at T = (298.15, 303.15, 308.15, and 313.15) K	https://www.doi.org/10.1016/j.jct.2012.08.030
Excess molar enthalpies of methyl formate + (1-propanol, [illegible] 1-pentanol) at T = 298.15 K, p = (5.0, 10.0) MPa, and methyl formate [illegible]	https://www.doi.org/10.1016/j.tca.2005.11.041
Activity Coefficients at Infinite Dilution of Polar Solutes in 1-Butyl-3-methylimidazolium Tetrafluoroborate Using Gas-Liquid Chromatography	https://www.doi.org/10.1021/je900704b
Density, Viscosity, Refractive Index, and Speed of Sound for the Binary Mixtures of 1,4-Dioxane with Different Organic Liquids at (298.15, 303.15, and 308.15) K	https://www.doi.org/10.1021/je049609w
Vapor-Liquid Equilibrium [illegible] 1-Butanol [illegible] at (318.4 and 364.5) K and [illegible]	https://www.doi.org/10.1021/je8004672
Vapor-Liquid Equilibrium of 1-Butanol + 2-Methylpropane, + n-Butane and [illegible] at 318.4 K; Solubilities [illegible] in Acetone, Toluene, Xylene, Ethanol, and [illegible]	https://www.doi.org/10.1021/je800405b
Solubility of Ivermectin B1a in Some Pure and Mixed Solvents from (278.2 to 313.2) K	https://www.doi.org/10.1021/je900772h
Molecular interactions in binary mixtures of formamide with 1-butanol, 2-butanol, 1,3-butanediol and 1,4-butanediol at different temperatures: An ultrasonic and viscometric study	https://www.doi.org/10.1016/j.fluid.2007.12.007
Densities, Excess Molar Volumes, and Refractive Properties of the Binary [illegible] Isopropanol at T = (298.15 to 313.15) K	https://www.doi.org/10.1021/je200030k
Solubility and phase behavior of binary [illegible] Liquid [illegible]	https://www.doi.org/10.1016/j.jct.2013.03.014
Separation of binary mixtures based on limiting activity coefficients data using specific ammonium-based ionic liquid and modelling of thermodynamic functions	https://www.doi.org/10.1016/j.fluid.2017.12.029
Thermodynamics of 4-bromomethyl-2-cyanobiphenyl in different solvents	https://www.doi.org/10.1016/j.jct.2014.12.002
[illegible] hexylamine + n-heptane at temperature range (288.15 323.15 K): Experimental density data, excess [illegible]	https://www.doi.org/10.1016/j.fluid.2010.07.011
Isobaric Vapor-Liquid Equilibria of Heptane + 1-Butanol and Heptane + [illegible] (91.8) kPa	https://www.doi.org/10.1021/je9006629
Isobaric Vapor-Liquid Equilibrium for the Binary Systems of sec-Butyl Acetate + n-Butyl Alcohol, Isobutyl Alcohol, or tert-Butyl Alcohol at 101.3 kPa	https://www.doi.org/10.1021/acs.jced.6b00471
Compact Analysis of a Homogeneous Ternary Mixture	https://www.doi.org/10.1021/acs.jced.5b01024
Acoustic, volumetric and osmotic properties of binary mixtures containing the ionic liquid 1-butyl-3-methylimidazolium [illegible]	https://www.doi.org/10.1016/j.jct.2012.02.007
Phase behaviour of 1-hexyloxymethyl-3-methyl-imidazolium and 1,3-dihexyloxymethyl-imidazolium based ionic liquids with alcohols, water, ketones and hydrocarbons: The effect of cation and anion on solubility	https://www.doi.org/10.1016/j.fluid.2006.07.005
Determination and Correlation of the Solubility of Musk Ketone in Pure and Binary Solvents [illegible]	https://www.doi.org/10.1021/acs.jced.9b00310
Binary Liquid-Liquid Equilibrium (LLE) for Dibutyl Ether (DBE) + Water from (288.15 to 318.15) K and Ternary LLE for Systems of DBE + C1 less than C4 Alcohols + Water at 298.15 K	https://www.doi.org/10.1021/je800446c
Excess molar enthalpies of binary and ternary mixtures of (ethanol or 1-butanol) + triethylamine and n-hexane:	https://www.doi.org/10.1016/j.fluid.2005.08.001

Measurement and correlation of the solubility of 3,4-bis(3-nitrofurazan-4-yl)furoxan (DNTF) in different solvents:	https://www.doi.org/10.1016/j.jct.2015.06.003
Thermodynamic behaviour of alkyl lactate-alkanol systems:	https://www.doi.org/10.1016/j.jct.2018.07.013
Physical properties of systems of interest to the edible oil industry: Viscosities and densities of model systems formed by (triacylglycerol + fatty acid + solvent):	https://www.doi.org/10.1016/j.jct.2017.06.012
Acoustic and Thermophysical Properties of Binary Liquid Mixtures of [illegible] with Hexane and Cyclohexane at 298.15 K:	https://www.doi.org/10.1007/s10765-009-0665-4
Activity coefficients at infinite dilution and physicochemical properties for organic solutes and water in the ionic liquid 1-(2-methoxyethyl)-1-methylpyrrolidinium bis(trifluoromethylsulfonyl)-amide:	https://www.doi.org/10.1016/j.jct.2012.03.015
Effects of alkyl group and temperature on the interactions between furfural and [illegible]	https://www.doi.org/10.1016/j.tca.2014.11.009
Experimental Study of Thermodynamic Properties of Mixtures Containing Ionic Liquid 1-Ethyl-3-methylimidazolium Ethyl Sulfate Using Gas-Liquid Chromatography and Transpiration Method:	https://www.doi.org/10.1021/je0602723
Determination and Correlation of Solubilities of Lauric Acid in Eight Alcohols:	https://www.doi.org/10.1021/je500222s
Comparison of Extractability of Oxalic Acid from Dilute Aqueous Solutions Using Dioctylamine and Trioctylphosphine Oxide:	https://www.doi.org/10.1021/acs.jced.8b01155
Solubility Modeling and Solvent Effects of Allopurinol in 15 Neat Solvents:	https://www.doi.org/10.1021/acs.jced.8b00430
Density and surface tension of pure ionic liquid 1-butyl-3-methylimidazolium L-lactate and its binary mixture with alcohol and water:	https://www.doi.org/10.1016/j.jct.2013.04.015
Solubility of 3-chloro-N-phenylphthalimide in ten organic solvents from T = (283.15 to 323.15) K: Determination and modelling:	https://www.doi.org/10.1016/j.jct.2016.01.003
Measurement of the Solubility of Diclazuril and Discussion of the Effect of Mixing Solvent:	https://www.doi.org/10.1021/acs.jced.9b00353
Measurement and Correlation of Solubilities of Asiaticoside in Water, Methanol, Ethanol, n-Propanol, n-Butanol, and a Methanol + Water Mixture from (278.15 to 343.15) K:	https://www.doi.org/10.1021/je101208p
Activity coefficients at infinite dilution and physicochemical properties for organic solutes and water in the ionic liquid 4-(2-methoxyethyl)-4-methylmorpholinium bis(trifluoromethylsulfonyl)-amide:	https://www.doi.org/10.1016/j.jct.2011.11.021
Liquid-Liquid Equilibrium for Ternary Systems of Propyl Vinyl Ether + C3 or C4 Alcohol + Water at 298.15 K and Excess Molar [illegible] for Ternary and Constituent Binary Systems of Propyl Vinyl Ether + Ethanol + Isooctane at 303.15 K:	https://www.doi.org/10.1021/je7005743
Measurement and Correlation of Solubility of 3-Hydroxy-2-naphthoic Acid in Pure and Binary Mixed Organic Solvents from T = (293.15 to [illegible]) K:	https://www.doi.org/10.1021/acs.jced.8b00578
Density, Surface Tension, and Refractive Index of Octane + 1-Alkanol Mixtures at T = 298.15 K:	https://www.doi.org/10.1021/je034053i
Solubilities of succinimide in different pure solvents and binary methanol+ethyl acetate solvent mixtures:	https://www.doi.org/10.1016/j.tca.2012.03.007
Determination and Correlation of Dipyridamole p-Toluene Sulfonate Solubility in Seven Alcohol Solvents and Three Binary Solvents:	https://www.doi.org/10.1021/acs.jced.7b00825
Solid-Liquid Equilibrium of Azacyclotridecan-2-one in 15 Pure Solvents from T = 273.15 to 323.15 K: Experimental Determination and Thermodynamic Modeling:	https://www.doi.org/10.1021/acs.jced.8b01193
Experimental Determination of Densities and Isobaric Vapor-Liquid Equilibria of Methyl Acetate and Ethyl Acetate with Alcohols (C3 and C4) at 0.3 MPa:	https://www.doi.org/10.1007/s10765-013-1483-2
Solubility Determination and Correlation of Glibenclamide in 11 Monosolvents and (Acetone + Acetonitrile) Binary Solvents from 283.15 K to 323.15 K:	https://www.doi.org/10.1021/acs.jced.8b00717
Properties of pure n-butylammonium nitrate ionic liquid and its binary mixtures of with alcohols at T = (293.15 to 313.15) K:	https://www.doi.org/10.1016/j.jct.2012.10.008
Bubble point temperatures of the binary mixtures of nitrobenzene with C1–C4 aliphatic alcohols at 94.95 kPa:	https://www.doi.org/10.1016/j.jct.2005.08.011
Phase equilibria study of the binary systems (N-butyl-4-methylpyridinium tosylate ionic liquid + an alcohol):	https://www.doi.org/10.1016/j.jct.2009.03.008
Dynamic Viscosities of Diethyl Carbonate with Linear and Secondary Alcohols at Several Temperatures:	https://www.doi.org/10.1021/je0341413
Solubility Data for Roflumilast and Maraviroc in Various Solvents between T = (273.2–323.2) K:	https://www.doi.org/10.1021/acs.jced.9b00710
Ternary Liquid-Liquid Equilibria for Systems Containing Alcohol or Acetic Acid + Ester + 1-Ethyl-3-methylimidazolium Hydrogen Sulfate at 313.2 K Using Headspace Gas Chromatography:	https://www.doi.org/10.1021/je800547k
Densities and Viscosities for Binary Liquid Mixtures of Butan-1-ol + Propane-1,2-diol, Butan-1-ol + Butane-1,2-diol and 2-Methoxyethanol + Propane-1,2-diol, + Butane-1,2-diol from 293.15 to 333.15 K:	https://www.doi.org/10.1021/acs.jced.7b00621
Solubility of Benzoin in Six Monosolvents and in Some Binary Solvent Mixtures at Various Temperatures:	https://www.doi.org/10.1021/acs.jced.7b00238
A [illegible] model for the concentration-dependent surface tension of binary systems:	https://www.doi.org/10.1016/j.fluid.2016.04.007
Viscosities of binary mixtures containing 1-butanol + 2,2,4-trimethylpentane or + 1,2,4-trimethylbenzene at high pressures for the thermophysical characterization of biofuels:	https://www.doi.org/10.1016/j.jct.2016.07.008
Densities, Viscosities and Refractive Indices of the Ternary Mixture Dimethyl Carbonate + 1-Butanol + 2-Butanol at T = (288.15 to 323.15) K:	https://www.doi.org/10.1021/je5007696
Solubility determination and correlation of cyromazine in sixteen pure solvents and mixing properties of solutions:	https://www.doi.org/10.1016/j.fluid.2018.07.024
Solubility of 2-Chlorophenylacetic Acid in 12 Pure Solvents from T = (273.15/283.15 to 318.15) K: Determination and Modeling:	https://www.doi.org/10.1021/acs.jced.8b00139
Comparative study of physical properties of binary mixtures of halogen-free ionic liquids with alcohols:	https://www.doi.org/10.1016/j.jct.2017.02.022
Solubility of Diphenyl Carbonate in Pure Alcohols from (283 to 333) K:	https://www.doi.org/10.1021/je800569v
Thermodynamic study of the solubility of 2,4'-dihydroxydiphenyl sulfone in nine organic solvents from T = (278.15 to 313.15) K and thermodynamic properties of dissolution:	https://www.doi.org/10.1016/j.jct.2016.07.001
Solubility of Dichloronitrobenzene in Eight Organic Solvents from T = (278.15 to 303.15) K: Measurement and Thermodynamic Modeling:	https://www.doi.org/10.1021/je401044h
Thermodynamic Properties of Mixtures Containing Ionic Liquids. 8. Activity Coefficients at Infinite Dilution of Hydrocarbons, Alcohols, Esters, and Aldehydes in 1-Hexyl-3-methylimidazolium Bis(trifluoromethylsulfonyl) imide Using Gas-Liquid Chromatography:	https://www.doi.org/10.1021/je0503554
Bubble Temperatures of the Binary Mixtures of Dimethylcarbonate with [illegible] Alcohols at 95.8 kPa:	https://www.doi.org/10.1021/je030131q
Measurement and correlation of solubility of trans-resveratrol in 11 solvents at T = (278.2, 288.2, 298.2, 308.2, and 318.2) K:	https://www.doi.org/10.1016/j.jct.2007.10.006

Ammonium ionic liquids in extraction of bio-butan-1-ol from water phase using activity coefficients at infinite dilution:	https://www.doi.org/10.1016/j.fluid.2018.09.024
Temperature and Composition Dependence of the Densities, Viscosities, and Speeds of Sound of Binary Liquid Mixtures of 1-Butanol with Hexadecane and Squalane:	https://www.doi.org/10.1021/je7007654
Osmotic coefficients and apparent molar volumes of 1-hexyl-3-methylimidazolium trifluoromethanesulfonate ionic liquid in alcohols:	https://www.doi.org/10.1016/j.jct.2013.09.041
Density and Excess Molar Volumes of 1-Butanol + Methanol + Electrolyte Systems in the Temperature Range (293.15 - 308.15) K:	https://www.doi.org/10.1021/acs.jced.5b00366
Measurement of the thermal conductivity of biofuel mixtures: Methyl caprate components of biodiesel and alcohols:	https://www.doi.org/10.1016/j.fluid.2019.112263
Vapor-Liquid Equilibrium of Binary Mixtures. 2. Ethanol + 2,2,4-Trimethylpentane, 1-Butanol + 2,2,4-Trimethylpentane, and Ethanol + o-Xylene:	https://www.doi.org/10.1021/je060005x
Screening of environmental friendly ionic liquid as a solvent for the different types of separations problem: Insight from activity coefficients at infinite dilution measurement using (gas + liquid) chromatography technique:	https://www.doi.org/10.1016/j.jct.2015.08.017
Solubility and Solution Thermodynamics of 3-Nitrophthalonitrile in 12 Neat Solvents at Temperatures from 278.15 to 318.15 K:	https://www.doi.org/10.1021/acs.jced.8b01250
LLE, VLE and VLLE data for the water n-butanol n-hexane system at atmospheric pressure:	https://www.doi.org/10.1016/j.fluid.2011.11.025
Effect of composition and temperature variations on thermophysical properties of binary and ternary mixtures of 1-ethyl-3-methylimidazolium ethylsulfate with 1-butanol and methanol:	https://www.doi.org/10.1016/j.fluid.2017.01.019
Isobaric Vapor Liquid Equilibrium for Two Binary Systems (Methanol + Dibutyl Carbonate and n-Butanol + Dibutyl Carbonate) at 101.3 kPa:	https://www.doi.org/10.1021/acs.jced.6b00517
Limiting activity coefficient measurements in binary mixtures of dichloromethane and 1-alkanols (C1 - C4):	https://www.doi.org/10.1016/j.fluid.2015.11.037
Phase equilibria and excess molar enthalpies study of the binary systems (pyrrole + hydrocarbon, or an alcohol) and modeling:	https://www.doi.org/10.1016/j.fluid.2013.10.029
Joback Method:	https://en.wikipedia.org/wiki/Joback_method
Isobaric Vapor Liquid Equilibrium for Two Binary Systems (n-Butanol + 1,4-Butanediol and gamma-Butyrolactone + 1,4-Butanediol) at p = (30.0, 50.0, and 70.0) kPa:	https://www.doi.org/10.1021/acs.jced.6b00088
Volumetric properties, viscosity and refractive index of the protic ionic liquid, pyrrolidinium octanoate, in molecular solvents:	https://www.doi.org/10.1016/j.jct.2010.01.013
Solubility and thermodynamic properties of vanillyl alcohol in some pure solvents:	https://www.doi.org/10.1016/j.jct.2016.11.030
Densities and Viscosities of Binary Mixtures of Cyclopropanecarboxylic Acid with Methanol, Ethanol, Propan-1-ol, and Butan-1-ol at Different Temperatures:	https://www.doi.org/10.1021/je800930v
Activity coefficients at infinite dilution measurements for organic solutes and water in the ionic liquid 1-ethyl-3-methylimidazolium tetracyanoborate:	https://www.doi.org/10.1016/j.jct.2011.02.012
Excess molar enthalpy of 1-alkanol + 1-octene mixtures at 298.15K experimental results and theoretical description by means of the ERAS and TB models:	https://www.doi.org/10.1016/j.fluid.2007.07.063
Thermodynamics of 1-alkanol + linear polyether mixtures:	https://www.doi.org/10.1016/j.jct.2012.12.007
Intermolecular interactions in binary mixtures of 2-diethylethanolamine with 1-propanol and 1-butanol at different temperatures:	https://www.doi.org/10.1016/j.jct.2018.06.007
Experimental study on the calorimetric data of 2-butoxyethanol with aliphatic alcohols (C1-C4) and correlation with the Wilson, NRTL and UNIQUAC models at T = 298 K:	https://www.doi.org/10.1016/j.jct.2016.08.003
Phase equilibria of (water + propionic acid or butyric acid + 2-methoxy-2-methylpropane) ternary systems at 298.2 K and 323.2 K:	https://www.doi.org/10.1016/j.fluid.2015.06.003
Thermodynamic properties of paracetamol impurities 4-nitrophenol and 4'-chloroacetanilide and the impact of such impurities on the crystallisation of paracetamol from solution:	https://www.doi.org/10.1016/j.jct.2019.02.004
Solubilities of 3-Carboxy-3-hydroxypentanedioic Acid in Ethanol, Butan-1-ol, Water, Acetone, and Methylbenzene:	https://www.doi.org/10.1021/je101167z
Measurement and Modelization of VLE for Butyl Acetate with Methanol, Ethanol, 1-Propanol, and 1-Butanol. Experimental Data at 0.15 MPa:	https://www.doi.org/10.1021/acs.jced.8b00182
Influence of electrolytes on liquid-liquid equilibria of water/1-butanol and on the partitioning of 5-hydroxymethylfurfural in water/1-butanol:	https://www.doi.org/10.1016/j.fluid.2016.05.001
Phase Equilibrium (VLE, LLE, and VLLE) Data of the Ternary System: Ionic Liquid [OMIM][PF_6] + Butan-1-ol + Butyl Acetate:	https://www.doi.org/10.1021/je5000296
Experimental measurement and modeling of solubility of LiBr and $LiNO_3$ in methanol, ethanol, 1-propanol, 2-propanol and 1-butanol:	https://www.doi.org/10.1016/j.fluid.2011.03.017
Isobaric Vapor-Liquid Equilibrium for the Binary Systems Dimethyl Disulfide + C1-C4 n-Alkanol at 40.000 and 101.325 kPa:	https://www.doi.org/10.1021/acs.jced.7b00078
Dynamic viscosities of binary mixtures of cycloalkanes with primary alcohols at T = (293.15, 298.15, and 303.15) K: New UNIFAC-VISCO interaction parameters:	https://www.doi.org/10.1016/j.jct.2006.06.008
Speed of sound measurements of liquid C1-C4 alkanols:	https://www.doi.org/10.1016/j.jct.2015.10.006
Excess molar enthalpies for binary mixtures of trimethyl phosphate with alkanols {$CH_3(CH_2)_nOH$, n = 0 3} at 298.15K:	https://www.doi.org/10.1016/j.tca.2008.02.015
Solubility of Potassium Clavulanate in Ethanol, 1-Propanol, 1-Butanol, 2-Propanol, and 2-Methyl-1-propanol between 273 K and 305 K:	https://www.doi.org/10.1021/je050145r
Solubilities of 4',5,7-Triacetoxyflavanone in Fourteen Organic Solvents at Different Temperatures:	https://www.doi.org/10.1021/acs.jced.6b00929
Solubility in Binary Solvent Mixtures: Anthracene Dissolved in Alcohol + 2-Methyl-1-butanol Mixtures at 298.2 K:	https://www.doi.org/10.1021/je030175z
Solubility Determination and Modeling of EGCG Peracetate in 12 Pure Solvents at Temperatures from 278.15 to 318.15 K:	https://www.doi.org/10.1021/acs.jced.9b00428
Determination and correlation of solubility and mixing properties of isonicotinamide (form II) in some pure solvents:	https://www.doi.org/10.1016/j.tca.2016.01.012
Molecular interactions in binary mixtures of 1-butoxy-2-propanol with alcohols at different temperatures: A thermophysical and spectroscopic approach:	https://www.doi.org/10.1016/j.jct.2014.07.019
Solubility of Gastrodin in Pure and Mixed Solvents at 273.15-313.15 K and Its Correlation with Different Thermodynamic Models:	https://www.doi.org/10.1021/acs.jced.9b00286
Solid-Liquid Equilibrium Measurements for Posaconazole and Voriconazole in Several Solvents between T = 278.2 and 323.2 K Using Differential Thermal Analysis/Thermal Gravimetric Analysis:	https://www.doi.org/10.1021/acs.jced.9b00179
Excess Molar Volumes and Surface Tensions of 1,2,4-Trimethylbenzene and 1,3,5-Trimethylbenzene with 1-Butanol, 2-Methyl-1-propanol, 2-Butanol, and 2-Methyl-2-propanol at 298.15 K:	https://www.doi.org/10.1021/je049807n

Saturated Solubility and Thermodynamic Evaluation of L-Tryptophan in Eight Pure Solvents and Three Groups of Binary Mixed Solvents by the Gravimetric Method at T = 278.15–333.15 K:	https://www.doi.org/10.1021/acs.jced.9b00562
Surface Tensions of the Ternary Mixtures Containing an Isomeric Butanol + n-Hexane + 1-Chlorobutane at 298.15 K:	https://www.doi.org/10.1021/je100177u
Phase equilibria of the water + 1-butanol + 2-pentanol ternary system at 101.3 kPa:	https://www.doi.org/10.1016/j.jct.2018.03.024
Determination and correlation of solubility and solution thermodynamics of musk xylene in different pure solvents:	https://www.doi.org/10.1016/j.jct.2019.03.029
Liquid-Liquid Equilibrium for Ternary System n-Butanol + n-Hexane + Ethylene Glycol at 298.15-323.15 K:	https://www.doi.org/10.1021/acs.jced.7b00577
Solubility of Chlorpheniramine Maleate in Ethanol, 1-Propanol, 1-Butanol, Benzene, Ethyl Acetate, Ethyl Formate, and Butyl Acetate between 283 K and 333 K:	https://www.doi.org/10.1021/je7001682
Temperature Dependence of Densities and Excess Molar Volumes of the Ternary Mixture (1-Butanol + Chloroform + Benzene) and its Binary Constituents (1-Butanol + Chloroform and 1-Butanol + Benzene):	https://www.doi.org/10.1007/s10765-008-0390-4
Determination and correlation of solubility and solution thermodynamics of [illegible] in different pure solvents:	https://www.doi.org/10.1016/j.fluid.2014.03.022
High-Pressure Vapor-Liquid Equilibria in the System Carbon Dioxide + 1-Butanol at Temperatures from (293.15 to 324.15) K:	https://www.doi.org/10.1021/je0499521
Solubility measurement and thermodynamic functions of 3-nitrobenzaldehyde in different solvents at elevated temperatures:	https://www.doi.org/10.1016/j.jct.2016.10.040
Separation of water/butan-1-ol based on activity coefficients at infinite dilution in 1,3-didecyl-2-methylimidazolium dicyanamide ionic liquid:	https://www.doi.org/10.1016/j.jct.2017.10.003
Excess molar volumes of Diisopropylamine + (C1-C5) Alkanols: application of the ERAS model and cubic EOS:	https://www.doi.org/10.1016/j.tca.2011.05.006
Excess Molar Enthalpies of Benzyl Alcohol + Alkanols (C1-C6) and Their Correlations at 298.15 K and Ambient Pressure:	https://www.doi.org/10.1021/je100095p
KDB Pure (Korean Thermophysical Properties Databank):	https://www.cheric.org/research/kdb/hcprop/showprop.php?cmpid=821
Mutual diffusion coefficients of n-butanol + n-heptane and n-pentanol + n-heptane from 288.15 K to 318.15 K:	https://www.doi.org/10.1016/j.jct.2017.05.028
Isobaric vapor-liquid phase equilibria for the binary systems of tert-butanol + 2-ethyl-1-hexanol and n-butanol + 2-ethyl-1-hexanol:	https://www.doi.org/10.1016/j.fluid.2005.03.034
Determination and correlation thermodynamic models for solid liquid equilibrium of the Nifedipine in pure and mixture organic solvents:	https://www.doi.org/10.1016/j.jct.2016.07.035
Solubilities of Betulin in Fourteen Organic Solvents at Different Temperatures:	https://www.doi.org/10.1021/je700069g
Solubility and limiting activity coefficient of simvastatin in different organic solvents:	https://www.doi.org/10.1016/j.fluid.2009.03.006
Vapor-Liquid Equilibrium Behaviors of 5-Methyl-2-(1-methylethyl)phenol in Alcohol:	https://www.doi.org/10.1021/je7007582
Measurement and Correlation of Solubilities of Luteolin in Organic Solvents at Different Temperatures:	https://www.doi.org/10.1021/je060133l
Excess Molar Volumes of 1,3-Diethyl Propanedioate with Methanol, Ethanol, Propan-1-ol, Propan-2-ol, Butan-2-ol, 2-Methyl-propan-1-ol, and Pentan-1-ol at T = (288.15, 298.15, 313.15, and 328.15) K:	https://www.doi.org/10.1021/je100100n
Measurement and correlation solubility and mixing properties of L-malic acid in pure and mixture organic solvents:	https://www.doi.org/10.1016/j.fluid.2013.10.011
Thermophysical properties and phase equilibria study of the binary systems {N-hexylquinolinium bis(trifluoromethylsulfonyl)imide + aromatic hydrocarbons, or an alcohol}:	https://www.doi.org/10.1016/j.jct.2010.12.021
Thermodynamic and transport properties of (1-Butanol + 1,4-Butanediol) at temperatures from (298.15 to 318.15) K:	https://www.doi.org/10.1016/j.jct.2009.09.013
Solubility of daidzin in different organic solvents and (ethyl alcohol + water) mixed solvents:	https://www.doi.org/10.1016/j.jct.2015.04.025
Separation of water/butan-1-ol with ionic liquids in ternary liquid-liquid phase equilibrium:	https://www.doi.org/10.1016/j.jct.2019.03.001
Thermodynamics and activity coefficients at infinite dilution for organic solutes in the ionic liquid 1-butyl-1-methylpyrrolidinium dicyanamide:	https://www.doi.org/10.1016/j.fluid.2018.06.013
Excess Enthalpy, Density, and Speed of Sound for the Mixtures α-Pinene + 1-Butanol or 2-Butanol at (283.15, 298.15, and 313.15) K:	https://www.doi.org/10.1021/je050294+
New ionic liquid [P4,4,4,4][NTf2] in bio-butanol extraction on investigation of limiting activity coefficients:	https://www.doi.org/10.1016/j.fluid.2018.07.028
Solubility of Carbamazepine (Form III) in Different Solvents from (275 to 343) K:	https://www.doi.org/10.1021/je8002157
Isobaric vapor-liquid equilibria for water + n-propanol + n-butanol ternary system at atmospheric pressure:	https://www.doi.org/10.1016/j.fluid.2010.06.004
Experimental and theoretical study on infinite dilution activity coefficients of various solutes in piperidinium ionic liquids:	https://www.doi.org/10.1016/j.jct.2013.01.005
Liquid-liquid equilibria of water + 2,3-butanediol + 1-butanol at T = 298.15 K, T = 308.15K and T = 318.15K:	https://www.doi.org/10.1016/j.fluid.2007.12.003
Solubility of caprolactam in different organic solvents:	https://www.doi.org/10.1016/j.fluid.2011.12.004
Solubility of 4-methyl-2-nitroaniline in fourteen organic solvents from T = (278.15 to 318.15) K and mixing properties of solutions:	https://www.doi.org/10.1016/j.jct.2017.02.008
Gas liquid chromatography measurements of activity coefficients at infinite dilution of various organic solutes and water in tri-iso-butylmethylphosphonium tosylate ionic liquid:	https://www.doi.org/10.1016/j.jct.2010.01.004
Compressibility Studies of Binary Solutions Involving Water as a Solute in Nonaqueous Solvents at T = 298.15 K:	https://www.doi.org/10.1021/je0601098
Experimental and theoretical study on liquid-liquid equilibrium of 1-butanol + water + NH_4Cl at 298.15, 308.15 and 318.15 K:	https://www.doi.org/10.1016/j.fluid.2012.03.026
Volumetric and compressibility behaviour of poly(propylene glycol) - Amino acid aqueous solutions at different temperatures:	https://www.doi.org/10.1016/j.jct.2015.06.024
Limiting activity coefficients in binary mixtures of 1-alkanols and ethylbenzene:	https://www.doi.org/10.1016/j.fluid.2007.09.021
Vapor-Liquid Equilibria of Binary Mixtures Containing 1-Butanol and Hydrocarbons at 313.15 K:	https://www.doi.org/10.1021/je200840e
A thermodynamic study on solubility and liquid-liquid phase behaviour for 2,2,2-trifluoroethanol + cyclohexane + 1-butanol at three temperatures and pressure 101.3 kPa:	https://www.doi.org/10.1016/j.jct.2010.05.010

Solubility of Sulfachloropyridazine in Pure and Binary Solvent Mixtures and Investigation of Intermolecular Interactions:	https://www.doi.org/10.1021/acs.jced.7b01134
Phase equilibria of glycerol containing systems and their description with the Cubic-Plus-Association (CPA) Equation of State:	https://www.doi.org/10.1016/j.fluid.2009.03.011
Excess enthalpy, density, and speed of sound determination for the ternary mixture (methyl tert-butyl ether + 1-butanol + n-hexane):	https://www.doi.org/10.1016/j.jct.2007.02.002
High pressure phase equilibrium for ethylene + 1-butanol at 283.65 K and 290.80 K:	https://www.doi.org/10.1016/j.fluid.2007.07.017
Solubility measurement and correlation of the form A of ibrutinib in organic solvents from 278.15 to 323.15 K:	https://www.doi.org/10.1016/j.jct.2016.08.034
Measurement and correlation of the solubility of genistin in eleven organic solvents from T = (283.2 to 323.2) K:	https://www.doi.org/10.1016/j.jct.2015.05.019
Liquid-liquid equilibrium of 1-butanol + water +tri-n-butyl phosphate + ammonium chloride system:	https://www.doi.org/10.1016/j.fluid.2014.06.023
Solubility and Solvent Effect of Acetamiprid in Thirteen Pure Solvents and Aqueous Solutions of Ethanol:	https://www.doi.org/10.1021/acs.jced.9b00294
Solid-Liquid Phase Equilibrium and Solubility of Dibenzo[b,d]furan and 9H-Fluoren-9-one in Organic Solvents:	https://www.doi.org/10.1021/je201282d
Excess molar volumes of binary mixtures of 1,3-dimethylimidazolidin-2-one with an alkan-1-ol at the temperatures 283.15 K, 298.15 K, and 313.15 K:	https://www.doi.org/10.1016/j.jct.2006.09.017
Thermodynamic functions of 1-methyl-4-(methylsulfonyl)benzene solubility in nine organic solvents from T = (278.15 to 318.15) K:	https://www.doi.org/10.1016/j.jct.2016.08.021
Molar excess enthalpies at T = 298.15 K for (1-alkanol + dibutylether) systems:	https://www.doi.org/10.1016/j.jct.2009.07.001
Thermodynamics and activity coefficients at infinite dilution measurements for organic solutes and water in the ionic liquid 1-butyl-1-methylpyrrolidinium tetracyanoborate:	https://www.doi.org/10.1016/j.jct.2011.06.007
Measurements of activity coefficients at infinite dilution of organic compounds and water in isoquinolinium-based ionic liquid [C8iQuin][NTf2] using GLC:	https://www.doi.org/10.1016/j.jct.2010.10.026
Activity Coefficients at Infinite Dilution of Polar Solutes in 1-Butyl-3-methylimidazolium Trifluoromethanesulfonate Using Gas-Liquid Chromatography:	https://www.doi.org/10.1021/je700560s
Solubility of Benzoic Acid and Aspirin in Pure Solvents Using Focused Beam Reflective Measurement:	https://www.doi.org/10.1021/acs.jced.5b00746
Measurement and Correlation of Solubility of 2-Chloro-4-ethylamino-6-isopropylamino-1,3,5-triazine in Different Organic Solvents:	https://www.doi.org/10.1021/je400639m
Solubility of Minoxidil in Methanol, Ethanol, 1-Propanol, 2-Propanol, 1-Butanol, and Water from (278.15 to 333.15) K:	https://www.doi.org/10.1021/je1012839
Activity coefficient at infinite dilution measurements for organic solutes (polar and non-polar) in fatty compounds – Part II: C18 fatty acids:	https://www.doi.org/10.1016/j.jct.2012.12.009
Solubility of Fluorene in Different Solvents from 278.98 K to 338.35 K:	https://www.doi.org/10.1021/je700179q
Solubility of antioxidant 1010 in pure alkanols:	https://www.doi.org/10.1016/j.fluid.2008.11.005
Excess Molar Enthalpies of Five Binary Systems Containing Ethyl Acetoacetate at Different Temperatures:	https://www.doi.org/10.1021/je800889y
Solubility and Dissolution Thermodynamics for 1,3,5-Trichlorobenzene in Organic Solvents:	https://www.doi.org/10.1021/acs.jced.5b00607
Measurement and Correlation of the Solubility of 2-Cyanoacetamide in 14 Pure Solvents and the Mixing Properties of Solutions:	https://www.doi.org/10.1021/acs.jced.8b01205
On the solubility of nicotinic acid and isonicotinic acid in water and organic solvents:	https://www.doi.org/10.1016/j.jct.2013.01.024
Activity Coefficients at Infinite Dilution for Organic Solutes Dissolved in Three 1-Alkyl-1-methylpyrrolidinium Bis(trifluoromethylsulfonyl)imide Ionic Liquids Bearing Short Linear Alkyl Side Chains of Three to Five Carbons:	https://www.doi.org/10.1021/je4001894
Densities and Viscosities of Binary Mixtures of 1-Butanol with Methacrylic Acid, Butyl Methacrylate, and 2-Hydroxyethyl Methacrylate between 288.15 K and 318.15 K:	https://www.doi.org/10.1021/je0601255
Solubilities of 3-Pentadecylphenol in Ethanol, 1-Butanol, Toluene, Acetone, Tetrachloromethane, and Ethyl Acetate:	https://www.doi.org/10.1021/je900346t
Evaluation of the solvent structural effect upon the vapor-liquid equilibrium of [C4C1im][Cl] + Alcohols:	https://www.doi.org/10.1016/j.fluid.2017.02.016
Liquid-liquid equilibria for the binary system of di-isopropyl ether (DIPE) +water in between 288.15 and 323.15K and the ternary systems of DIPE +water +C1-C4 alcohols at 298.15K:	https://www.doi.org/10.1016/j.fluid.2008.04.010
Measurement and Correlation of the Solubilities of m-Phthalic Acid in Monobasic Alcohols:	https://www.doi.org/10.1021/je800520w
Densities, Excess Molar Volumes, and Isobaric Thermal Expansibilities for 1,2-Ethanediol + 1-Butanol, or 1-Hexanol, or 1-Octanol in the Temperature Range from (293.15 to 353.15) K:	https://www.doi.org/10.1021/je700580g
Solubility Measurement and Modeling of 4,4'-Dihydroxydiphenyl Sulfone in Nine Organic Solvents from (278.15 to 313.15) K and Thermodynamic Property of Dissolution:	https://www.doi.org/10.1021/acs.jced.5b00714
Vapor-Liquid Equilibrium Behaviors of 3-Ethoxy-4-hydroxybenzaldehyde in Alcohols:	https://www.doi.org/10.1021/je700316s
Vapor-Liquid Equilibria and HE for Binary Systems of Dimethyl Ether (DME) with C1-C4 Alkan-1-ols at 323.15 K and Liquid-Liquid Equilibria for Ternary System of DME + Methanol + Water at 313.15 K:	https://www.doi.org/10.1021/je0603574
Densities, Viscosities, and Ultrasonic Velocity Studies of Binary Mixtures of Chloroform with Pentan-1-ol, Hexan-1-ol, and Heptan-1-ol at (303.15 and 313.15) K:	https://www.doi.org/10.1021/je050169y
Measurement and Correlation of Solubility and Thermodynamic Properties of Vinpocetine in Nine Pure Solvents and (Ethanol + Water) Binary Solvent:	https://www.doi.org/10.1021/acs.jced.8b00663
Determination and Correlation of Solubility of Quetiapine Fumarate in Nine Pure Solvents and Two Aqueous Binary Solvents:	https://www.doi.org/10.1021/acs.jced.7b00518
Prediction of different thermodynamic properties for systems of alcohols and sulfate-based anion Ionic Liquids using modified UNIFAC:	https://www.doi.org/10.1016/j.fluid.2012.11.003
Solubility and Solution Thermodynamics of Gibberellin A4 in Different Organic Solvents from 278.15 K to 333.15 K:	https://www.doi.org/10.1021/acs.jced.5b00190
(Liquid + liquid) equilibria in {water + acrylic acid + (1-butanol, or 2-butanol, or 1-pentanol)} systems at T = 293.2 K, T = 303.2 K, and T = 313.2 K and atmospheric pressure:	https://www.doi.org/10.1016/j.jct.2011.04.009
Solubility and tie-line data for ternary aqueous mixtures of cyclopentanol and organic solvents at T = 298.2 K: Experiments and NRTL model:	https://www.doi.org/10.1016/j.fluid.2014.07.011
Henry's Constants of 1-Alkanols and 2-Ketones in Salt Solutions:	https://www.doi.org/10.1021/je0600956

Liquid liquid equilibrium of ternary systems containing nicotine:	https://www.doi.org/10.1016/j.fluid.2007.02.016
Solubility of rosuvastatin calcium in different neat solvents at different temperatures:	https://www.doi.org/10.1016/j.jct.2015.11.019
Determination and correlation of molar excess enthalpies of binary systems 2,4-pentanedione + (1-butanol, + 2-methyl-1-propanol, + 1-pentanol, + 1-heptane, + ethyl acetate, and + water):	https://www.doi.org/10.1016/j.fluid.2007.12.008
Densities and volumetric properties of (N-(2-hydroxyethyl)morpholine + ethanol, + 1-propanol, + 2-propanol, + 1-butanol, and + 2-butanol) at (293.15, 298.15, 303.15, 313.15, and 323.15) K:	https://www.doi.org/10.1016/j.jct.2007.11.010
Thermodynamics and activity coefficients at infinite dilution for organic solutes and water in the ionic liquid 1-hexyl-3-methylimidazolium chloride:	https://www.doi.org/10.1016/j.jct.2018.08.028
Solubility Measurement and Phase Equilibrium Modeling of 2-Aminobenzamide in 15 Pure Solvents:	https://www.doi.org/10.1021/acs.jced.9b00350
Thermodynamics of solubility of thiomalic acid in different organic solvents from 278.15 K to 333.15 K:	https://www.doi.org/10.1016/j.fluid.2013.05.002
Heat Capacities of alpha,omega-Dichloroalkanes at Temperatures from 284.15 K to 353.15 K and a Group Additivity Analysis:	https://www.doi.org/10.1021/je020042y
Viscosity and Density of n-Alcohols at Temperatures between (298.15 and 323.15) K and Pressures up to 30 MPa:	https://www.doi.org/10.1021/acs.jced.8b00812
(Solid + liquid) and (liquid + liquid) phase equilibria study and correlation of the binary systems {N-butyl-3-methylpyridinium tosylate + water, or + an alcohol, or + a hydrocarbon}:	https://www.doi.org/10.1016/j.fluid.2010.01.025
Volumetric and transport properties of ternary mixtures containing 1-alkanol + ethyl ethanoate + cyclohexane at 298.15 K: Experimental data, correlation and prediction by ERAS model:	https://www.doi.org/10.1016/j.tca.2009.02.022
Experimental study of the LL, VL and VLL equilibria of water + 1-butanol + [illegible] at 101.3 kPa:	https://www.doi.org/10.1016/j.fluid.2018.07.037
Thermodynamics and activity coefficients at infinite dilution for organic compounds and water in the ionic liquid 1-butyl-3-methylimidazolium perchlorate:	https://www.doi.org/10.1016/j.jct.2017.07.012
Solubilities of Naringin and Naringenin in Different Solvents and Dissociation Constants of Naringenin:	https://www.doi.org/10.1021/je501004g
Activity coefficients at infinite dilution and physicochemical properties for organic solutes and water in the ionic liquid 1-(2-methoxyethyl)-1-methylpiperidinium bis(trifluoromethylsulfonyl)-amide:	https://www.doi.org/10.1016/j.jct.2012.01.019
Isothermal and isobaric (vapour + liquid) equilibria of (alpha-pinene + [illegible]):	https://www.doi.org/10.1016/j.jct.2004.08.012
Solubilities of Phloretin in 12 Solvents at Different Temperatures:	https://www.doi.org/10.1021/je101168w
Liquid-liquid equilibrium data for ternary systems of water + lactic acid + C4-C7 alcohols at 298.2 K and atmospheric pressure:	https://www.doi.org/10.1016/j.fluid.2013.06.007
Determination and Modeling of Solubility of 4-Aminobenzamide in Different Pure Solvents:	https://www.doi.org/10.1021/acs.jced.8b01144
Viscosities and Densities of Binary Mixtures of (N-Acetylmorpholine + Alkanols) from (293.15 to 323.15) K:	https://www.doi.org/10.1021/je8003194
Excess molar volumes, excess molar enthalpies and refractive index deviations for binary mixtures of propan-1-ol, butan-1-ol and pentan-1-ol with 2,2,4-trimethylpentane at 298.15 K:	https://www.doi.org/10.1016/j.fluid.2012.06.031
Solubility and mixing thermodynamic properties of levamisole hydrochloride in twelve pure solvents at various temperatures:	https://www.doi.org/10.1016/j.jct.2019.105882
Effect of the Alkyl Chain on the Physical Properties of Imidazolium-Based Ionic Liquids with 1-Propanol, 1-Butanol, and 1-Pentanol Binary Mixtures:	https://www.doi.org/10.1021/acs.jced.8b00948
Determination and thermodynamic modeling of solid-liquid phase equilibrium for 2,5-dichloroaniline in pure solvents and ternary 3,5-dichloroaniline + [illegible] system:	https://www.doi.org/10.1016/j.jct.2016.03.007
Determination and Correlation of Pyridazin-3-amine Solubility in Eight Organic Solvents at Temperatures Ranging from (288.15 to 328.15) K:	https://www.doi.org/10.1021/je300517q
Measurement of activity coefficient at infinite dilution for some bio-oil components in water and mass transfer study of bubbles in the dilutor:	https://www.doi.org/10.1016/j.fluid.2015.01.010
Solubility of [illegible] in Ethanol, 1-Propanol, 2-Propanol, 1-Butanol, Acetonitrile, Acetone, Ethyl Acetate, and Water from 293 K to 343 K:	https://www.doi.org/10.1021/je900177h
The effect of temperature on the liquid-liquid equilibria of some aliphatic alcohols + di-sodium hydrogen citrate + water systems: Experimental and correlation:	https://www.doi.org/10.1016/j.fluid.2013.05.031
Isobaric vapor liquid equilibrium for the binary systems of 1-butanol with o-xylene, m-xylene, p-xylene, and ethylbenzene at 101.33 kPa:	https://www.doi.org/10.1021/je201053v
Liquid Phase Behavior of 1-Butyl-1-methylpyrrolidinium Bis(trifluoromethylsulfonyl)imide with 1-Butanol or Hexan-1-ol: Measurements of Coexistence Curves, Density, and Interfacial Tension:	https://www.doi.org/10.1021/je201391j
Experimental Liquid-Liquid Equilibria of 1-Alkyl-3-methylimidazolium Hexafluorophosphate with 1-Alcohols:	https://www.doi.org/10.1021/je700099w
Thermodynamic study of solubility for pyrazinamide in ten solvents from T = (283.15 to 323.15) K:	https://www.doi.org/10.1016/j.jct.2017.04.014
Influence of chain length and degree of branching of alcohol + chlorobenzene mixtures on determination and modelling of VE by CEOS and CEOS/GE mixing rules:	https://www.doi.org/10.1016/j.fluid.2007.09.028
Isothermal vapor liquid equilibrium and molar excess Gibbs energies of two ternary systems containing either 1-butanol or 2-butanol + 1- hexene + methylbenzene at 313.15 K:	https://www.doi.org/10.1016/j.fluid.2014.11.010
Activity Coefficients at Infinite Dilution of Polar Solutes in 1-Butyl-3-methylimidazolium Tetrafluoroborate Using Gas-Liquid Chromatography:	https://www.doi.org/10.1021/je060305e
Experimental Determination and Modeling of Densities and Refractive Indices of the Binary Systems Alcohol + Dicyclohexylamine at T = (288.15 to 323.15) K:	https://www.doi.org/10.1016/j.tca.2011.08.002
Solubilities of Proxetine Hydrochloride Hemihydrate between 286 K and 363 K:	https://www.doi.org/10.1021/je0498762
Measurement and Study on the Liquid Liquid Equilibrium of Ionic Liquids with Aliphatic Alcohols:	https://www.doi.org/10.1021/je101261z
Liquid-Liquid Equilibria of the System 1-(2-Hydroxyethyl)-3-methylimidozolium Tetrafluoroborate or 1-(2-Hydroxyethyl)-2,3-dimethylimidozolium Tetrafluoroborate + Water + 1-Butanol at 293.15 K:	https://www.doi.org/10.1021/je050455+
Activity coefficients at infinite dilution of organic solutes in [illegible]phosphonium-based ionic liquids using gas-liquid chromatography:	https://www.doi.org/10.1016/j.jct.2015.02.023
Solid-Liquid Equilibrium Solubility, Thermodynamic Properties, and Molecular Simulation of Phenylphosphonic Acid in 15 Pure Solvents:	https://www.doi.org/10.1021/acs.jced.9b00362
Solubility of Amorphous Clopidogrel Hydrogen Sulfate in Different Pure Solvents at Different Temperatures:	https://www.doi.org/10.1021/acs.jced.5b00306
Thermodynamic Properties of Mixtures Containing Ionic Liquids. 6. Activity Coefficients at Infinite Dilution of Hydrocarbons, Alcohols, Esters, and Aldehydes in 1-Methyl-3-octyl-imidazolium Tetrafluoroborate Using Gas-Liquid Chromatography:	https://www.doi.org/10.1021/je050134y

Measurement and Correlation of Solubility of Gatifloxacin in 12 Pure Solvents from 273.15 K to 318.15 K:	https://www.doi.org/10.1021/acs.jced.8b00902
Determination and correlation of solubility and thermodynamic properties of pyraclostrobin in pure and binary solvents:	https://www.doi.org/10.1016/j.jct.2016.05.005
Determination and correlation of the solubility for diosgenin in alcohols solvents:	https://www.doi.org/10.1016/j.jct.2012.02.002
Ionic association and conductance of [emim][BF4] and [bmim][BF4] in 1-butanol in a wide range of temperature:	https://www.doi.org/10.1016/j.jct.2015.12.029
Liquid-Liquid Equilibrium for Ternary Systems, Water + 5-Hydroxymethylfurfural + (1-Butanol, Isobutanol, Methyl Isobutyl Ketone), at 313.15, 323.15, and 333.15 K:	https://www.doi.org/10.1021/acs.jced.8b00120
Viscosities, Densities, and Ultrasonic Velocities of Binary Mixtures of Ethylbenzene with 1-Propanol, and 1-Butanol at (298.15 and 308.15) K:	https://www.doi.org/10.1021/je049912x
Determination and Correlation of Solubility of Cefradine Form I in Five Pure Solvents from (283.15 to 308.15) K:	https://www.doi.org/10.1021/je400218d
Equilibrium solubility of dimetomorph in several neat solvents and binary aqueous co-solvent mixtures: Experimental determination and thermodynamic analysis:	https://www.doi.org/10.1016/j.jct.2019.01.013
Vapor Liquid Equilibrium Data for Binary Systems of n-Dodecane + (Propanol, n-Butanol, 2-Methylpropan-1-ol) at 40 kPa:	https://www.doi.org/10.1021/je500237g
Thermodynamic Models for Determination of the Solubility of Boc-(R)-3-Amino-4-(2,4,5-trifluorophenyl)butanoic acid in Different Pure Solvents and (Tetrahydrofuran + n-Butanol) Binary Mixtures with Temperatures from 280.15 to 330.15 K:	https://www.doi.org/10.1021/acs.jced.5b00758
Excess molar volumes and partial molar volumes for (propionitrile + an alkan-1-ol) at 298.15 K and p = 0.1 MPa:	https://www.doi.org/10.1016/j.jct.2005.05.007
Solubility of salicylic acid in pure alcohols at different temperatures:	https://www.doi.org/10.1016/j.jct.2012.09.006
Volumetric and Transport Properties of Binary Mixtures of n-Octane + Ethanol, + 1-Propanol, + 1-Butanol, and + 1-Pentanol from (293.15 to 323.15) K at Atmospheric Pressure:	https://www.doi.org/10.1021/je4004806
Density and Surface Tension Variation with Temperature for Heptane + 1-Alkanol:	https://www.doi.org/10.1021/je060179e
Solubility and mixing thermodynamic properties of (2,4,6-trimethylbenzoyl) diphenylphosphine oxide in pure and binary solvents:	https://www.doi.org/10.1016/j.fluid.2018.01.015
Solubility of Organic Systems Containing 1,4-Dioxan-2-one:	https://www.doi.org/10.1021/je050406x
The Solid-Liquid Equilibrium and Crystal Habit of l-Carnitine Fumarate:	https://www.doi.org/10.1021/acs.jced.7b00768
Excess Enthalpy, Density, and Speed of Sound for the Ternary Mixture Methyl tert-Butyl Ether (1) + Butan-1-ol (2) + Octane (3):	https://www.doi.org/10.1021/je8004613
Compressed liquid densities of 1-butanol and 2-butanol at temperatures from 313 K to 363 K and pressures up to 25 MPa:	https://www.doi.org/10.1016/j.jct.2006.07.009
Solubility determination and thermodynamic modelling of terephthaldialdehyde in ten organic solvents from T = (273.15 to 318.15) K and mixing properties of solutions:	https://www.doi.org/10.1016/j.jct.2016.07.013
Activity Coefficients at Infinite Dilution of Organic Compounds in Four New Imidazolium-Based Ionic Liquids:	https://www.doi.org/10.1021/je200195q
Densities, Refractive Indices, Viscosities, and Spectroscopic Study of 1-Amino-2-propanol + 1-Butanol and + 2-Butanol Solutions at (288.15 to 333.15) K:	https://www.doi.org/10.1021/je401036f
Three Binary Azeotropic Systems for 1-(Methoxymethoxy)-propane, 2-(methoxymethoxy)-propane, and Methoxy(methoxymethoxy)methane with Three Alcohols at 101.33 kPa:	https://www.doi.org/10.1021/acs.jced.7b00740
Measurement and Correlation of the Solubility of Tetramethylpyrazine in Nine Monosolvents and Two Binary Solvent Systems: Experimental Data, Correlation, and Purification:	https://www.doi.org/10.1021/acs.jced.8b00888
Vapor Liquid Equilibrium Data for Binary Systems of 1H-Pyrrole with Butan-1-ol, Propan-1-ol, or Pentan-1-ol	https://www.doi.org/10.1021/je3004194
Solubility of D(-)-p-Hydroxyphenylglycine Dane Salt in Eight Alcohols between (293 and 343) K:	https://www.doi.org/10.1021/je800506q
Solubility of p-Aminobenzenesulfonamide in Different Solvents from (283.15 to 323.15) K:	https://www.doi.org/10.1021/je900038m
(Vapor + liquid) equilibria of alcohol + 1-methyl-1-propylpiperidinium triflate ionic liquid: VPO measurements and modeling:	https://www.doi.org/10.1016/j.jct.2016.01.026
Liquid density of biofuel mixtures: (Dibutyl ether + 1-butanol) system at pressures up to 140 MPa and temperatures from (293.15 to 393.15) K:	https://www.doi.org/10.1016/j.jct.2011.06.006
Solubilities of 2,4-Dinitro-L-phenylalanine in Monosolvents at (273.15 to 368.15) K:	https://www.doi.org/10.1021/je100352r
Liquid-Liquid Equilibrium in Binary System [bmim][PF6] + 1-Butanol:	https://www.doi.org/10.1021/je060260h
Experimental study on the calorimetric data of cyclohexanol with alkanols (C1-C4) and correlation with Wilson, NRTL and UNIQUAC models at 300 K:	https://www.doi.org/10.1016/j.jct.2012.02.037
Density, Speed of Sound, Viscosity, Excess Properties, and Prigogine-Flory-Patterson (PFP) Theory of Binary Mixtures of Amine and Alcohols:	https://www.doi.org/10.1021/acs.jced.5b00216
Solubility of 1-(2-Dimethylaminoethyl)-5-mercapto-1H-tetrazole in Water, DMF, Methanol, Ethanol, 1-Propanol, 2-Propanol, 1-Butanol, and 2-Butanol between (293 and 343) K:	https://www.doi.org/10.1021/je8003244
Excess Enthalpies of Binary and Ternary Mixtures Containing Dibutyl Ether, Cyclohexane, and 1-Butanol at 298.15 K:	https://www.doi.org/10.1021/je800751m
Activity coefficients at infinite dilution and physicochemical properties for organic solutes and water in the ionic liquid 1-(2-hydroxyethyl)-3-methylimidazolium tris(pentafluoroethyl)trifluorophosphate:	https://www.doi.org/10.1016/j.jct.2013.05.008
Thermodynamics of mixtures containing amines. XVI. CE pm of benzylamine + 1-propanol, 1-butanol, 1-pentanol, 1-hexanol, 1-octanol or 1-decanol systems at (293.15, 308.15, 318.15 and 333.15) K:	https://www.doi.org/10.1016/j.tca.2014.10.018
Measurement and Correlation of the Solubility of Mifepristone in Eight Pure and Water + Methanol Mixed Solvents at Temperatures from 273.15 to 318.15 K:	https://www.doi.org/10.1021/acs.jced.8b01062
Thermodynamic Properties of Mixtures Containing Ionic Liquids: Liquid-Liquid Equilibria of Binary Mixtures of 1-Ethyl-3-methylimidazolium Bis(trifluoromethylsulfonyl)imide with Propan-1-ol, Butan-1-ol, and Pentan-1-ol:	https://www.doi.org/10.1021/je0201931
Excess Molar Enthalpies of Methyl Acetate and (1-Propanol, 2-Propanol, 1-Butanol, 2-Butanol, and 1-Pentanol) at 298.15 K and (5.0 MPa):	https://www.doi.org/10.1021/je050175u
Volumetric and surface properties of pure ionic liquid n-octyl-pyridinium bis(trifluoromethylsulfonyl)imide and binary mixture with alcohol:	https://www.doi.org/10.1016/j.jct.2011.10.013
Feasibility of bio-based lactate esters as extractant for biobutanol recovery: (Liquid + liquid) equilibria:	https://www.doi.org/10.1016/j.jct.2015.10.003
Activity Coefficients at Infinite Dilution of Polar Solutes in 1-Propyl-2,3-dimethylimidazolium Tetrafluoroborate Using Gas Liquid Chromatography:	https://www.doi.org/10.1021/je700640r

Solubility of Xylitol in Ethanol, Acetone, N,N-Dimethylformamide, 1-Butanol, 1-Pentanol, Toluene, 2-Propanol, and Water:	https://www.doi.org/10.1021/je060348v
Density, viscosity, refractive index excess molar enthalpy, viscosity, and refractive index deviations for the (1-butanol + 2-butanone) binary system at T = 303 K. A new adiabatic calorimeter for heat of mixing:	https://www.doi.org/10.1016/j.jct.2007.06.009
Effect of Pressure on the Static Relative Permittivities of Alkan-1-ols at 298.15 K:	https://www.doi.org/10.1021/je060248p
Measurement and prediction of multicomponent diffusion coefficients in four ternary liquid systems:	https://www.doi.org/10.1016/j.fluid.2009.12.009
Measurements of Infinite Dilution Activity Coefficients of Alkanols in the Ionic Liquid Tributylmethylammonium Methyl Sulfate Using HS-SPME/GC/FID:	https://www.doi.org/10.1021/acs.jced.5b00919
Solubility Determination and Modeling of p-Nitrobenzamide Dissolved in Twelve Neat Solvents from 283.15 to 328.15 K:	https://www.doi.org/10.1021/acs.jced.9b00065
Isobaric vapor–liquid equilibria for the binary systems 1-propyl alcohol + dipropyl ether and 1-butyl alcohol + dibutyl ether at 20 and 101.3 kPa:	https://www.doi.org/10.1016/j.fluid.2006.06.017
Activity coefficients at infinite dilution measurements for organic solutes and water in the ionic liquid 4-methyl-N-butyl-pyridinium bis(trifluoromethylsulfonyl)-imide:	https://www.doi.org/10.1016/j.jct.2009.06.011
Physical Properties of the Pure Deep Eutectic Solvent, [ChCl]:[Lev] (1:2) DES, and Its Binary Mixtures with Alcohols:	https://www.doi.org/10.1021/acs.jced.6b00563
Vapor–Liquid–Liquid Equilibria Measurements for the Dehydration of Ethanol, Isopropanol, and n-Propanol via Azeotropic Distillation Using DIPE and Isooctane as Entrainers:	https://www.doi.org/10.1021/je300847v
Liquid–liquid equilibrium and critical states for the quaternary system propionic acid–n-butanol–n-butyl propionate–water at 303.15 K:	https://www.doi.org/10.1016/j.fluid.2017.12.023
Experimental determination and correlation of the solubility of 4-hydroxy-2,5-dimethyl-3(2H)-furanone (DMHF) in six different solvents:	https://www.doi.org/10.1016/j.jct.2015.08.028
Solubility of Oleanolic Acid in Various Solvents from (288.3 to 328.3) K:	https://www.doi.org/10.1021/je700312r
Experimental and predicted excess molar enthalpies of binary mixtures containing 2,2'-oxybis[propane], benzene, butan-1-ol, 2-methylpropan-1-ol, 2-methyl-2-ene-1-propanol at 303.15 K:	https://www.doi.org/10.1016/j.fluid.2010.08.024
Densities and Viscosities of Binary Mixtures of 1-Chlorobutane with Butanol Isomers at Several Temperatures:	https://www.doi.org/10.1021/je0496140
Salt effect on (liquid + liquid) equilibrium of (water + tertbutanol + 1-butanol) system: Experimental data and correlation:	https://www.doi.org/10.1016/j.jct.2005.03.025
Experimental Binary VLE Data of Morpholine with 1-Butanol and 3-Methyl-1-butanol Systems:	https://www.doi.org/10.1021/acs.jced.7b01128
Experimental determination and modelling of densities and excess molar volumes of ternary system (1-butanol + cyclohexylamine + n-heptane) and corresponding binaries from 298.15 to 323.15 K:	https://www.doi.org/10.1016/j.tca.2009.07.002
Thermodynamic and spectral investigations of binary liquid mixtures of 2-butoxyethanol with alcohols at the temperature range of 293.15-318.15 K:	https://www.doi.org/10.1016/j.fluid.2013.06.021
Isothermal Vapor-Liquid Equilibria for the n-Pentane + 1-Butanol and n-Pentane + 2-Butanol Systems near the Critical Region of the Mixtures:	https://www.doi.org/10.1021/je049894j
Apparent molar volumes and compressibilities of tetrabutyl ammonium bromide in organic solvents:	https://www.doi.org/10.1016/j.jct.2011.01.005
Separation of binary mixtures hexane/hex-1-ene, cyclohexane/cyclohexene and ethylbenzene/styrene based on limiting activity coefficients:	https://www.doi.org/10.1016/j.jct.2017.03.004
Density, Viscosity, and Refractive Index Properties for the Binary Mixtures of n-Butylammonium Acetate Ionic Liquid + Alkanols at Several Temperatures:	https://www.doi.org/10.1021/je200707b
PρT measurements and derived properties of liquid 1-alkanols:	https://www.doi.org/10.1016/j.jct.2011.10.023
Vapor Liquid Equilibrium at p/kPa = 101.3 of the Binary Mixtures of Ethenyl Acetate with Methanol and Butan-1-ol:	https://www.doi.org/10.1021/je300835h
Thermodynamics of Fluconazole Solubility in Various Solvents at Different Temperatures:	https://www.doi.org/10.1021/je4010257
Molar excess enthalpies of binary mixtures of [tributyl phosphate + {CH3(CH2)nOH (n = 0 to 3)}] at 298.15 K:	https://www.doi.org/10.1016/j.jct.2006.08.001
Measurements and correlation of liquid-liquid equilibria of 4-methyl-2-pentanone + ethanol + water and 4-methyl-2-pentanone + n-butanol + water ternary systems between 283.2 and 323.2 K:	https://www.doi.org/10.1016/j.fluid.2012.01.009
Solubility determination and thermodynamic modeling of 2-methyl-4-nitrobenzoic acid in twelve organic solvents from T = (283.15–318.15) K and mixing properties of solutions:	https://www.doi.org/10.1016/j.jct.2016.07.023
Determination and correlation of solubility of spironolactone form II in pure solvents and binary solvent mixtures:	https://www.doi.org/10.1016/j.jct.2014.07.011
Measurement and Correlation of Liquid–Liquid Equilibria for the Ternary Systems of Water + d-Fructose + 1-Butanol, Water + d-Glucose + 1-Butanol, and Water + d-Galactose + 1-Butanol at (288.2, 303.2 and 318.2) K:	https://www.doi.org/10.1021/acs.jced.7b00302
Experimental studies on the thermal conductivity of methyl laurate + components of biodiesel with the alcohols:	https://www.doi.org/10.1016/j.jct.2019.105881
High-pressure phase behavior of ternary systems (carbon dioxide + alkanol + hydrophobic ionic liquid):	https://www.doi.org/10.1016/j.fluid.2010.02.015
Densities and Viscosities for Binary Liquid Mixtures of Ethanol + 1-Propanol, 1-Butanol, and 1-Pentanol from (293.15 to 328.15) K at 0.1 MPa:	https://www.doi.org/10.1021/je300632p
Solubility and Thermodynamic Behavior of Veramoss in Different Pure Solvents and {Ethanol + Water} Mixtures from T = 283.15 to 333.15 K:	https://www.doi.org/10.1021/acs.jced.6b00646
Experimental Determination and Modeling of Densities, Refractive Indices of the Binary Mixtures of Dimethylphthalate (or Dimethyladipate) + 1-Butanol, or + 2-Butanol, or + 2-Butanone at T = (288.15 to 343.15) K:	https://www.doi.org/10.1016/j.tca.2012.01.013
Densities and Viscosities of exo-Tetrahydrodicyclopentadiene + n-Butanol and exo-Tetrahydrodicyclopentadiene + n-Pentanol at Temperatures of (293.15 to 313.15) K:	https://www.doi.org/10.1021/je9005129
Excess Molar Enthalpies of 1,2-Propanediol + Alkan-1-ols (C1-C6) and Their Correlation at 298.15 K and Atmospheric Pressure (81.5 kPa):	https://www.doi.org/10.1021/je060086x
Thermophysical properties and thermodynamic phase behavior of ionic liquids:	https://www.doi.org/10.1016/j.tca.2006.06.018
Measurement and Correlation of the Solubility of Florfenicol Form A in Several Pure and Binary Solvents:	https://www.doi.org/10.1021/acs.jced.8b00043
Isobaric (vapor+liquid) equilibrium for n-propyl acetate with 1-butanol or 2-butanol. Binary mixtures at 0.15 and 0.6 MPa:	https://www.doi.org/10.1016/j.fluid.2014.10.035
Measurement and Correlation of Solubility of Loratadine in Different Pure Solvents and Binary Mixtures:	https://www.doi.org/10.1021/acs.jced.6b00721
Extraction of Citric Acid and Maleic Acid from Their Aqueous Solutions Using a Phosphorus-Bonded Extractant, Tri-n-octylphosphineoxide, and a Secondary Amine, Dioctylamine:	https://www.doi.org/10.1021/acs.jced.7b00562

Aqueous Solubility Prediction Method:	http://onschallenge.wikispaces.com/file/view/AqueousDataset002.xlsx/351826032/AqueousDa
Solubility determination and thermodynamic modeling of 2,4-dinitroaniline in nine organic solvents from T = (278.15 to 318.15) K and mixing properties of solutions:	https://www.doi.org/10.1016/j.jct.2016.07.009
Excess enthalpies of oxygenated compounds + hydrocarbon mixtures: Binary and ternary mixtures containing dibutyl ether (DBE), 1-butanol and 2,2,4-trimethylpentane at 298.15 K:	https://www.doi.org/10.1016/j.fluid.2009.11.010
Solubility Measurement and Correlation for 2-Naphthaldehyde in Pure Organic Solvents and Methanol + Ethanol Mixtures:	https://www.doi.org/10.1021/acs.jced.5b00377
Measurement and Correlation for Solubility of Dexibuprofen in Different Solvents from (283.15 to 293.15) K:	https://www.doi.org/10.1016/j.tca.2012.04.016
Solubility determination and thermodynamic modeling of 5-nitro-8-hydroxyquinoline in ten organic solvents from T = (278.15 to 313.15) K and mixing properties of solutions:	https://www.doi.org/10.1016/j.jct.2016.04.018
Application of the ERAS model to volumetric properties of binary mixtures of benzene with primary and secondary alcohols (C1- C4) at different temperatures:	https://www.doi.org/10.1016/j.jct.2010.12.009
Solubility Measurement and Thermodynamic Modeling for [illegible]onamide in 16 Solvents from T = 273.15 to 323.85 K:	https://www.doi.org/10.1021/acs.jced.9b00445
Isothermal (vapour + liquid) equilibrium and thermophysical properties for (1-butyl-3-methylimidazolium iodide + 1-butanol) binary system:	https://www.doi.org/10.1016/j.jct.2015.03.023
Liquid-liquid equilibria for the ternary system containing 1-Butanol + methoxy(methoxymethoxy)methane + water at temperatures of 303.15, 323.15 and 343.15 K:	https://www.doi.org/10.1016/j.fluid.2015.11.009
Isobaric Phase Equilibria of Diethyl Carbonate with Five Alcohols at 101.3 kPa:	https://www.doi.org/10.1021/je0201018
Tie-Line Data for Aqueous Mixtures of Butyric Acid with Diisopropyl Ether at Various Temperatures:	https://www.doi.org/10.1021/acs.jced.5b00527
Thermodynamic Properties of 1-Butyl-3-methylpyridinium Tetrafluoroborate and Its Mixtures with Water and Alkanols:	https://www.doi.org/10.1021/je700294p
Solubility of 6-Chloropyridazin-3-amine in Different Solvents:	https://www.doi.org/10.1021/je3001188
Solubility of N-tert-Butylbenzothiazole-2-sulfenamide in Several Pure and Binary Solvents:	https://www.doi.org/10.1021/acs.jced.8b00954
Liquid Liquid Equilibria of Water + 1-Butanol + 2-Ethyl-1-hexanol System:	https://www.doi.org/10.1021/je030202l
Vapor Liquid Equilibria Measurements for the Nine n-Alkane/Ketone Pairs Comprising 2-, 3-, and 4-Heptanone with n-Octane, n-Nonane, and n-Decane:	https://www.doi.org/10.1021/je500731x
Vapour liquid equilibria, azeotropic data, excess enthalpies, activity coefficients at infinite dilution and solid liquid equilibria for binary alcohol ketone systems:	https://www.doi.org/10.1016/j.fluid.2008.02.021
Naphthalene Solubility in Binary Solvent Mixtures of 2,2,4-Trimethylpentane + Alcohols at 298.15 K:	https://www.doi.org/10.1021/je7005049
Solubility Determination and Modeling of Sodium Mononitrobenzenesulphonate in Different Solvents at Temperatures Ranging from (283.15 to 323.15) K:	https://www.doi.org/10.1021/acs.jced.5b00783
Solubilities of (2 Z)-2-Amino-2-(methoxyimino)-4-thiazole-ethanethioic Acid S-2-Benzothiazolyl Ester in Pure Solvents and a Mixture of Acetonitrile and Dichloromethane:	https://www.doi.org/10.1021/je900784v
Solubility measurement and modelling of 1,8-dinitronaphthalene in nine organic solvents from T = (273.15 to 308.15) K and mixing properties of solutions:	https://www.doi.org/10.1016/j.jct.2015.07.010
Phase equilibria of didecyldimethylammonium nitrate ionic liquid with water and organic solvents:	https://www.doi.org/10.1016/j.jct.2006.10.009
Atmospheric densities and interfacial tensions for 1- alkanol (1-butanol to 1-octanol) + water and ether (MTBE, ETBE, DIPE, TAME and THF) + water demixed mixtures:	https://www.doi.org/10.1016/j.fluid.2015.03.040
Volumetric properties of ternary (IL + 2-propanol or 1-butanol or 2-butanol + ethyl acetate) systems and binary (IL + 2-propanol or 1-butanol or 2-butanol) and (1-butanol or 2-butanol + ethyl acetate) systems:	https://www.doi.org/10.1016/j.jct.2012.01.002
Solid-Liquid Phase Equilibrium and Thermodynamic Analysis of [illegible] in Different Solvent Systems:	https://www.doi.org/10.1021/acs.jced.9b00854
Density, speed of sound and refractive index of mixtures containing 2-phenoxyethanol with propanol or butanol at various temperatures:	https://www.doi.org/10.1016/j.jct.2018.08.029
Limiting Activity Coefficients by Comparative Tensimetry: 1-Propanol and 1-Butanol in Heptane and in Octane:	https://www.doi.org/10.1021/je0302295
Vapor-Liquid Equilibrium Behaviors of Vanillin in 1-Butanol, 2-Butanol, and 2-Methyl-1-propanol:	https://www.doi.org/10.1021/je700063c
Solubility of Acetylene in Alcohols and Ketones:	https://www.doi.org/10.1021/acs.jced.8b00126
Measurement and Correlation of Excess Molar Enthalpies for Ethylene Glycol + Alkanol Systems at the Temperatures (298.15, 308.15, and 323.15) K:	https://www.doi.org/10.1021/je030207i
Henry's Constants and Infinite Dilution Activity Coefficients of Propane, Propene, Butane, Isobutane, 1-Butene, Isobutene, trans-2-Butene and 1,3-Butadiene in 1-Butanol and 2-Butanol at T = (250 to 330) K:	https://www.doi.org/10.1021/je034265f
Study of surface tension and surface properties of binary systems of DMSO with long chain alcohols at various temperatures:	https://www.doi.org/10.1016/j.jct.2013.04.009
Excess Molar Volumes, Viscosities, and Refractive Indexes for Binary Mixtures of 1-Chlorobutane with Four Alcohols at T = (288.15, 298.15 and 308.15) K:	https://www.doi.org/10.1021/je030226s
The Yaws Handbook of Vapor Pressure:	https://www.sciencedirect.com/book/9780128029992/the-yaws-handbook-of-vapor-pressure
KDB:	https://www.cheric.org/files/research/kdb/mol/mol821.mol
Volumetric behaviour of the (2,2,4-trimethylpentane + methylbenzene + butan-1-ol) ternary system and its binary sub-systems within the temperature range (298.15–328.15) K:	https://www.doi.org/10.1016/j.jct.2013.05.009
Present status of the modified UNIFAC Model for the prediction of phase equilibria and excess enthalpies for systems with ionic liquids:	https://www.doi.org/10.1016/j.fluid.2014.03.006
Extraction of butan-1-ol from water with ionic liquids at T = 308.15 K:	https://www.doi.org/10.1016/j.jct.2012.04.017
Thermodynamic models for determination of solid liquid equilibrium of the 6-benzyladenine in pure and binary organic solvents:	https://www.doi.org/10.1016/j.jct.2016.12.002
A combined experimental and computational investigation of excess molar enthalpies of (nitrobenzene + alkanol) mixtures:	https://www.doi.org/10.1016/j.jct.2014.09.001
Densities, Viscosities, and Speeds of Sound of Binary Liquid Mixtures of Ethylenediamine with Alcohols at T = (293.15 to 313.15) K:	https://www.doi.org/10.1021/je1012857
Solubilities of 4-Amino-4?-nitrodiphenyl Sulfide in Six Organic Solvents from 287.45 K to 365.55 K:	https://www.doi.org/10.1021/je4000982

Surface Tension of Dilute Solutions of Linear Alcohols in Benzyl Alcohol:	https://www.doi.org/10.1021/je0500431
Thermodynamic analysis for determination of solid - liquid equilibrium of dibenzothiophene in a binary solvent of formic acid and butan-1-ol between (278.15 and 383.15)	https://www.doi.org/10.1016/j.fluid.2015.08.023
Thermodynamic behavior of binary mixtures CnMpyNTf2 ionic liquids with	https://www.doi.org/10.1016/j.tca.2012.09.009
Solubility of Anthracene in Binary Propyl Acetate + Alcohol Solvent Mixtures at 298.15 K:	https://www.doi.org/10.1021/je700502x
Separation of binary mixtures based on gamma infinity data using [OMIM][NTf2] ionic liquid and modelling of thermodynamic functions:	https://www.doi.org/10.1016/j.jct.2017.12.012
Liquid-liquid equilibrium data for systems containing of formic acid, water, and primary normal alcohols at T = 298.2 K:	https://www.doi.org/10.1016/j.fluid.2013.06.006
Solid Forms, Crystal Habits, and Solubility of Danthron:	https://www.doi.org/10.1021/acs.jced.5b00192
Vapor-Liquid Equilibrium Studies for Systems Containing n-Butylisocyanate at Temperatures between 323.15 K and 371.15 K:	https://www.doi.org/10.1021/je030222n
Solubility of triclosan and iodopropynyl butylcarbamate in pure alkanols at several temperatures:	https://www.doi.org/10.1016/j.fluid.2012.05.020
Solubility of 2,5-Furandicarboxylic Acid in Eight Pure Solvents and Two Binary Solvent Systems at 313.15-363.15 K:	https://www.doi.org/10.1021/acs.jced.7b00927
Measurement and Correlation of Solubility of Marbofloxacin in 12 Pure Solvents from 283.15 to 328.15 K:	https://www.doi.org/10.1021/acs.jced.9b00490
Measurement and Correlation for Solubility of 11alpha-Hydroxy-16alpha,17alpha-epoxyprogesterone and 16alpha,17alpha-Epoxyprogesterone in Solvents	https://www.doi.org/10.1021/je101072d
Excess molar enthalpies of R-fenchone + butan-1-ol or + pentan-1-ol. Modeling	https://www.doi.org/10.1016/j.jct.2018.01.005
Liquid-Liquid Extraction of Alcohols from Their Azeotropic Mixtures with Hexane Using Ionic Liquid:	https://www.doi.org/10.1021/acs.jced.8b00362
Thermodynamic Behaviour of Binary Mixtures of 1-Butyl-1-methylpyrrolidinium iodide and Alcohols:	https://www.doi.org/10.1021/acs.jced.5b00184
Isothermal vapor-liquid equilibria, excess molar volume and the deviation of refractive indices for binary mixtures of 1-butanol, 1-hexanol, 3-methyl-1-butanol and butyl acetate:	https://www.doi.org/10.1016/j.fluid.2016.12.025
Determination and Correlation of Solubility of 3,4-Dinitro-1H-pyrazole in Different Pure Solvents from 298.15 K to 338.15 K:	https://www.doi.org/10.1021/acs.jced.6b00219
Densities and Viscosities of Diaminotoluene with Water, Ethanol, Propan-1-ol, and Butan-1-ol from (293.15 to 333.15) K:	https://www.doi.org/10.1021/je400968v
Solubility in Binary Solvent Mixtures: Anthracene Dissolved in Alcohol + Carbon Tetrachloride Mixtures at 298.2 K:	https://www.doi.org/10.1021/je0301904
Thermodynamic Solubility and Mixing Properties of Phenformin in 14 Pure Solvents at Temperatures Ranging from 278.15 to 323.15 K:	https://www.doi.org/10.1021/acs.jced.9b00844
Studies of viscosities of dilute solutions of alkylamines in non-electrolyte solvents III. Alkylamines in butanols 303.15K:	https://www.doi.org/10.1016/j.tca.2009.02.005
Solubility in Binary Solvent Mixtures: Pyrene Dissolved in Alcohol + Acetonitrile Mixtures at 299.2 K:	https://www.doi.org/10.1021/je020226c
(Vapour + liquid) equilibrium and excess Gibbs functions of ternary mixtures containing 1-butanol or 2-butanol, n-hexane, and 1-chlorobutane at T = 298.15 K:	https://www.doi.org/10.1016/j.jct.2009.04.005
Thermodynamic study of molecular interaction selectivity in separation processes based on limiting activity coefficients:	https://www.doi.org/10.1016/j.jct.2018.02.014
Measurement and Correlation of Vapor-Liquid Equilibria at T) 333.15 K and Excess Molar Volumes at T) 298.15 K for Ethanol + Dimethyl Carbonate (DMC), DMC + 1-Propanol, and DMC + 1-Butanol:	https://www.doi.org/10.1021/je060217u
Solubility Behavior of 2-Chloro-3-(trifluoromethyl)pyridine in (Ethyl Acetate + n-Butanol, DMF + n-Butanol, DMF + Ethanol, and Methanol) Solvent Mixtures:	https://www.doi.org/10.1021/acs.jced.8b00989
Activity coefficients at infinite dilution of hydrocarbons, alkylbenzenes, and alcohols in the paramagnetic ionic liquid 1-butyl-3-methyl-imidazolium tetrachloridoferrate(III) using gas liquid chromatography:	https://www.doi.org/10.1016/j.jct.2008.09.004
Measurements of the Properties of Binary Mixtures of Dimethylsulphoxide (DMSO) with 1-Alkanols (C4, C6, C7) at 298.15 K:	https://www.doi.org/10.1007/s10765-005-8102-9
Isothermal Vapor-Liquid Equilibrium Data for Propan-1-ol + 2,2,4-Trimethylpentane and Butan-1-ol + 2,2,4-Trimethylpentane at 318.15 K:	https://www.doi.org/10.1021/je200279q
Densities, viscosities, and ultrasonic velocity studies of binary mixtures of trichloromethane with methanol, ethanol, propan-1-ol, and butan-1-ol at T = (298.15 and 308.15) K:	https://www.doi.org/10.1016/j.jct.2006.03.010
Thermodynamic Parameters of Anionic Surfactant-Additive Systems at the Cloud Point:	https://www.doi.org/10.1021/je100502b
Liquid-Liquid Equilibrium for the Quaternary System Water + Tetrahydrofuran + Toluene + 1-Butanol Mixture at 25 deg C and Atmospheric Pressure:	https://www.doi.org/10.1021/je700289p
Excess Gibbs free energies of the binary mixtures of acetonitrile with [illegible] at 94.83 kPa:	https://www.doi.org/10.1016/j.jct.2006.07.014
Solubility of climbazole in various alcohols at different temperatures:	https://www.doi.org/10.1016/j.jct.2014.04.011
Measurement of activity coefficients at infinite dilution of organic solutes in the ionic liquid 1-hexyl-1,4-diaza[2.2.2]bicyclooctanium bis(trifluoromethylsulfonyl)imide using gas-liquid chromatography:	https://www.doi.org/10.1016/j.jct.2013.10.026
Solubility of Agomelatine Crystal Form I and Form II in Pure Solvents and [illegible]	https://www.doi.org/10.1021/acs.jced.5b00586
Measurement of Infinite Dilution Activity Coefficients of Alcohols, Ketones, and Aromatic Hydrocarbons in 4-Methyl-N-butylpyridinium Tetrafluoroborate and 1-Butyl-3-methylimidazolium Hexafluorophosphate by Gas-Liquid Chromatography:	https://www.doi.org/10.1021/je800246v
Vapour-liquid equilibrium for the systems butane + methanol, +2-propanol, +1-butanol, +2-butanol, +2-methyl-2-propanol at 364.5 K:	https://www.doi.org/10.1016/j.fluid.2007.02.028
Vapor-liquid equilibrium and excess properties of the binary mixtures formed by ethyl isobutyrate and alkanols:	https://www.doi.org/10.1016/j.jct.2019.105884
Henry's law constants and infinite dilution activity coefficients of cis-2-butene, dimethylether, chloroethane, and 1,1-difluoroethane in methanol, 1-propanol, 2-propanol, 1-butanol, 2-butanol, isobutanol, tert-butanol, 1-pentanol, 2-pentanol, 3-pentanol, 2-methyl-1-butanol, 3-methyl-1-butanol, and 2-methyl-2-butanol:	https://www.doi.org/10.1016/j.jct.2005.08.004
Measurement and correlation of binary vapor liquid equilibria of isomeric [illegible]	https://www.doi.org/10.1016/j.fluid.2007.11.002
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Measurement and Correlation of Solubility of Indole-2-carboxylic Acid in Ten Different Pure Solvents from (278.15 to 360.15) K:	https://www.doi.org/10.1021/je400813d

Solubility Measurement and Thermodynamic Model Correlation and Evaluation of 2-Chloro-5-nitroaniline in 12 Pure Solvents	https://www.doi.org/10.1021/acs.jced.8b00931
Activity Coefficients at Infinite Dilution of Alkanols in the Ionic Liquids 1-Butyl-3-methylimidazolium Hexafluorophosphate, 1-Butyl-3-methylimidazolium Methyl Sulfate, and 1-Hexyl-3-methylimidazolium Bis(trifluoromethylsulfonyl) Amide Using the Dilutor Technique:	https://www.doi.org/10.1021/je800355e
Dortmund Data Bank Vapor-Liquid Equilibrium Data:	http://www.ddbst.com/en/EED/VLE/VLE%20Acetonitrile%3B1-Butanol.php
Determination and Correlation of Solubilities of (S)-Indoline-2-carboxylic Acid in Six Different Solvents from (283.15 to 358.15) K:	https://www.doi.org/10.1021/je400188h
Thermodynamic properties of binary mixtures of diethylenetriamine with alcohols at different temperatures:	https://www.doi.org/10.1016/j.tca.2011.06.003
Excess Volume of Binary Mixtures Containing an Alcohol (1-Butanol, 2-Butanol, 2-Methyl-2-butanol, and Butanol) and 2-Ethoxy-2-methylbutane:	https://www.doi.org/10.1021/je300670n
Dibutyl Citrate Synthesis, Physicochemical Characterization, and	https://www.doi.org/10.1021/acs.jced.7b01064
Isobaric vapor-liquid equilibrium for the binary systems of 1-propanol + 1-(methoxymethoxy)-butane and 1-butanol + 1-(methoxymethoxy)-butane at 101.3 kPa:	https://www.doi.org/10.1016/j.fluid.2014.12.029
Liquid-Liquid Equilibria for Systems of 1-Butanol + Water	https://www.doi.org/10.1021/je8007843
Liquid-Liquid Equilibria of Water + 1-Butanol + Water + 2-Butanone + 1-Butanol:	https://www.doi.org/10.1021/je2000759
Determination and Correlation of Solubility Data and Dissolution Thermodynamic Data of Cefixime Trihydrate in Seven Pure Solvents:	https://www.doi.org/10.1021/je5000605
Solubility determination and modelling of 4-nitro-1,2-phenylenediamine in eleven organic solvents from T = (283.15 to 318.15) K and	https://www.doi.org/10.1016/j.jct.2016.11.014
Viscosities and densities for binary mixtures of N-methylpiperazine with thermodynamic properties of solutions: 1-propanol, n-butanol, and iso-butanol at 293.15, 298.15 and 303.15K:	https://www.doi.org/10.1016/j.fluid.2005.05.012
Solubilities of evodiamine in twelve organic solvents from T = (283.2 to 323.2) K:	https://www.doi.org/10.1016/j.jct.2012.11.021
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Density, Refractive Index, and Sound Velocity for the Binary Mixtures of ... between 303.15 K and 323.15 K:	https://www.doi.org/10.1021/acs.jced.5b00343
Activity coefficients at infinite dilution and physicochemical properties for organic solutes and water in the ionic liquid 1-(2-methoxyethyl)-1-methylpiperidinium trifluorotris(perfluoroethyl)phosphate:	https://www.doi.org/10.1016/j.jct.2012.08.016

Legend

af:	Acentric Factor
affp:	Proton affinity
aigt:	Autoignition Temperature
basg:	Gas basicity
chl:	Standard liquid enthalpy of combustion
cpg:	Ideal gas heat capacity
cpl:	Liquid phase heat capacity
dm:	Dipole Moment
dvisc:	Dynamic viscosity
econd:	Electrical conductivity
fll:	Lower Flammability Limit
flu:	Upper Flammability Limit
fpc:	Flash Point (Closed Cup Method)
fpo:	Flash Point (Open Cup Method)
gf:	Standard Gibbs free energy of formation
gyrad:	Radius of Gyration
hf:	Enthalpy of formation at standard conditions
hfl:	Liquid phase enthalpy of formation at standard conditions
hfus:	Enthalpy of fusion at standard conditions
hfust:	Enthalpy of fusion at a given temperature
hvap:	Enthalpy of vaporization at standard conditions

hvapt:	Enthalpy of vaporization at a given temperature
ie:	Ionization energy
kvisc:	Kinematic viscosity
log10ws:	Log10 of Water solubility in mol/l
logp:	Octanol/Water partition coefficient
mcvol:	McGowan's characteristic volume
nfpaf:	NFPA Fire Rating
nfpah:	NFPA Health Rating
pc:	Critical Pressure
pvap:	Vapor pressure
rfi:	Refractive Index
rhoc:	Critical density
rhog:	Gas Density
rhol:	Liquid Density
rinpol:	Non-polar retention indices
ripol:	Polar retention indices
sfust:	Entropy of fusion at a given temperature
sg:	Molar entropy at standard conditions
sl:	Liquid phase molar entropy at standard conditions
speedsl:	Speed of sound in fluid
srf:	Surface Tension
tb:	Normal Boiling Point Temperature
tbp:	Boiling point at given pressure
tc:	Critical Temperature
tf:	Normal melting (fusion) point
tt:	Triple Point Temperature
vc:	Critical Volume
zc:	Critical Compressibility
zra:	Rackett Parameter

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