2-Methyl-1-butanol

Other names: 1-butanol, 2-methyl-
Inchi: InChI=1S/C5H12O/c1-3-5(2)4-6/h5-6H,3-4H2,1-2H3
InchiKey: QPRQEDXDYOZYLA-UHFFFAOYSA-N
Formula: C5H12O
SMILES: CCC(C)CO
Mol. weight [g/mol]: 88.15
CAS: 137-32-6

Physical Properties

Property code	Value	Unit	Source
gf	-148.04	kJ/mol	Joback Method
hf	-304.04	kJ/mol	Joback Method
hfus	9.27	kJ/mol	Joback Method
hvap	43.02	kJ/mol	Joback Method
log10ws	-0.47		Aqueous Solubility Prediction Method
logp	1.025		Crippen Method
mcvol	87.180	ml/mol	McGowan Method
pc	3916.03	kPa	Joback Method
tb	402.45	K	Vapor-Liquid Equilibrium of the Ethanol + 2-Methyl-1-butanol System
tb	401.93	K	Vapor-Liquid Equilibrium of Binary Mixtures Containing Ethyl Acetate + 2-Methyl-1-propanol and Ethyl Acetate + 2-Methyl-1-butanol at 101.3 kPa
tb	401.93	K	Vapor-Liquid Equilibria at 101.3 kPa for Binary Mixtures Containing 2-Methyl-1-propanol + 2-Methyl-1-butanol, 2-Methyl-1-propanol + 3-Methyl-1-butanol, and 2-Methyl-1-propanol + 1-Pentanol
tb	401.27	K	Isobaric Vapor Liquid Equilibria for the Binary and Ternary Systems of 2-Methyl-1-butanol, 2-Methyl-butanol Acetate, and Dimethylformamide (DMF) at 101.3 kPa

tc	570.76	K	Joback Method
tf	191.93	K	Joback Method
tt	118.00	K	Heat Capacities in the Solid and in the Liquid Phase of Isomeric Pentanols
vc	0.329	m3/kmol	Joback Method

Temperature Dependent Properties

Property code	Value	Unit	Temperature [K]	Source
cpg	216.94	J/mol×K	570.76	Joback Method
cpg	202.42	J/mol×K	515.68	Joback Method
cpg	194.72	J/mol×K	488.15	Joback Method
cpg	186.73	J/mol×K	460.61	Joback Method
cpg	178.43	J/mol×K	433.08	Joback Method
cpg	169.82	J/mol×K	405.54	Joback Method
cpg	209.82	J/mol×K	543.22	Joback Method
dvisc	0.0379452	Pa×s	227.53	Joback Method
dvisc	0.0026559	Pa×s	298.74	Joback Method
dvisc	0.0010745	Pa×s	334.34	Joback Method
dvisc	0.0083860	Pa×s	263.13	Joback Method
dvisc	0.0005174	Pa×s	369.94	Joback Method
dvisc	0.3005917	Pa×s	191.93	Joback Method
dvisc	0.0002833	Pa×s	405.54	Joback Method
kvisc	0.0000047	m2/s	303.15	Density, Viscosity, Excess Molar Volume, and Viscosity Deviation of Three Amyl Alcohols + Ethanol Binary Mixtures from 293.15 to 323.15 K
kvisc	0.0000041	m2/s	308.15	Density, Viscosity, Excess Molar Volume, and Viscosity Deviation of Three Amyl Alcohols + Ethanol Binary Mixtures from 293.15 to 323.15 K

kvisc	0.0000053	m2/s	298.15	Density, Viscosity, Excess Molar Volume, and Viscosity Deviation of Three Amyl Alcohols + Ethanol Binary Mixtures from 293.15 to 323.15 K	
kvisc	0.0000035	m2/s	313.15	Density, Viscosity, Excess Molar Volume, and Viscosity Deviation of Three Amyl Alcohols + Ethanol Binary Mixtures from 293.15 to 323.15 K	
kvisc	0.0000027	m2/s	323.15	Density, Viscosity, Excess Molar Volume, and Viscosity Deviation of Three Amyl Alcohols + Ethanol Binary Mixtures from 293.15 to 323.15 K	
kvisc	0.0000031	m2/s	318.15	Density, Viscosity, Excess Molar Volume, and Viscosity Deviation of Three Amyl Alcohols + Ethanol Binary Mixtures from 293.15 to 323.15 K	
kvisc	0.0000061	m2/s	293.15	Density, Viscosity, Excess Molar Volume, and Viscosity Deviation of Three Amyl Alcohols + Ethanol Binary Mixtures from 293.15 to 323.15 K	
pvap	32.00	kPa	371.74	Vapor pressure of selected aliphatic alcohols by ebulliometry. Part 2	

pvap	50.00	kPa	382.75	Effect of addition of calcium chloride on vapor pressure of 2-methyl-1-butanol
pvap	63.30	kPa	388.95	Effect of addition of calcium chloride on vapor pressure of 2-methyl-1-butanol
pvap	70.00	kPa	391.65	Effect of addition of calcium chloride on vapor pressure of 2-methyl-1-butanol
pvap	76.60	kPa	394.15	Effect of addition of calcium chloride on vapor pressure of 2-methyl-1-butanol
pvap	56.60	kPa	385.65	Effect of addition of calcium chloride on vapor pressure of 2-methyl-1-butanol
pvap	83.30	kPa	396.65	Effect of addition of calcium chloride on vapor pressure of 2-methyl-1-butanol
pvap	9.33	kPa	345.51	Vapor pressure of selected aliphatic alcohols by ebulliometry. Part 2
pvap	12.00	kPa	350.44	Vapor pressure of selected aliphatic alcohols by ebulliometry. Part 2
pvap	14.66	kPa	354.51	Vapor pressure of selected aliphatic alcohols by ebulliometry. Part 2
pvap	17.33	kPa	358.02	Vapor pressure of selected aliphatic alcohols by ebulliometry. Part 2
pvap	20.66	kPa	361.80	Vapor pressure of selected aliphatic alcohols by ebulliometry. Part 2
pvap	24.00	kPa	365.12	Vapor pressure of selected aliphatic alcohols by ebulliometry. Part 2

pvap	28.00	kPa	368.62	Vapor pressure of selected aliphatic alcohols by ebulliometry. Part 2
pvap	90.00	kPa	398.85	Effect of addition of calcium chloride on vapor pressure of 2-methyl-1-butanol
pvap	36.00	kPa	374.56	Vapor pressure of selected aliphatic alcohols by ebulliometry. Part 2
pvap	40.00	kPa	377.13	Vapor pressure of selected aliphatic alcohols by ebulliometry. Part 2
pvap	44.00	kPa	379.50	Vapor pressure of selected aliphatic alcohols by ebulliometry. Part 2
pvap	48.66	kPa	382.05	Vapor pressure of selected aliphatic alcohols by ebulliometry. Part 2
pvap	53.33	kPa	384.41	Vapor pressure of selected aliphatic alcohols by ebulliometry. Part 2
pvap	58.66	kPa	386.92	Vapor pressure of selected aliphatic alcohols by ebulliometry. Part 2
pvap	63.99	kPa	389.23	Vapor pressure of selected aliphatic alcohols by ebulliometry. Part 2
pvap	69.33	kPa	391.40	Vapor pressure of selected aliphatic alcohols by ebulliometry. Part 2
pvap	74.66	kPa	393.43	Vapor pressure of selected aliphatic alcohols by ebulliometry. Part 2
pvap	79.99	kPa	395.35	Vapor pressure of selected aliphatic alcohols by ebulliometry. Part 2

pvap	85.33	kPa	397.17	Vapor pressure of selected aliphatic alcohols by ebulliometry. Part 2
pvap	90.66	kPa	398.90	Vapor pressure of selected aliphatic alcohols by ebulliometry. Part 2
pvap	97.33	kPa	400.96	Vapor pressure of selected aliphatic alcohols by ebulliometry. Part 2
pvap	16.70	kPa	357.85	Vapor-Liquid Equilibrium of the Ethanol + 2-Methyl-1-butanol System
pvap	23.30	kPa	364.85	Vapor-Liquid Equilibrium of the Ethanol + 2-Methyl-1-butanol System
pvap	30.00	kPa	370.45	Vapor-Liquid Equilibrium of the Ethanol + 2-Methyl-1-butanol System
pvap	33.30	kPa	372.95	Vapor-Liquid Equilibrium of the Ethanol + 2-Methyl-1-butanol System
pvap	43.30	kPa	379.05	Effect of addition of calcium chloride on vapor pressure of 2-methyl-1-butanol
pvap	43.30	kPa	379.05	Vapor-Liquid Equilibrium of the Ethanol + 2-Methyl-1-butanol System
pvap	50.00	kPa	382.75	Vapor-Liquid Equilibrium of the Ethanol + 2-Methyl-1-butanol System
pvap	56.60	kPa	385.65	Vapor-Liquid Equilibrium of the Ethanol + 2-Methyl-1-butanol System
pvap	63.30	kPa	388.95	Vapor-Liquid Equilibrium of the Ethanol + 2-Methyl-1-butanol System

pvap	66.60	kPa	390.35	Vapor-Liquid Equilibrium of the Ethanol + 2-Methyl-1-butanol System
pvap	70.00	kPa	391.65	Vapor-Liquid Equilibrium of the Ethanol + 2-Methyl-1-butanol System
pvap	76.60	kPa	394.15	Vapor-Liquid Equilibrium of the Ethanol + 2-Methyl-1-butanol System
pvap	96.60	kPa	401.05	Effect of addition of calcium chloride on vapor pressure of 2-methyl-1-butanol
pvap	90.00	kPa	398.85	Vapor-Liquid Equilibrium of the Ethanol + 2-Methyl-1-butanol System
pvap	96.60	kPa	401.05	Vapor-Liquid Equilibrium of the Ethanol + 2-Methyl-1-butanol System
pvap	101.30	kPa	402.45	Vapor-Liquid Equilibrium of the Ethanol + 2-Methyl-1-butanol System
pvap	23.30	kPa	364.85	Effect of addition of calcium chloride on vapor pressure of 2-methyl-1-butanol
pvap	30.00	kPa	370.45	Effect of addition of calcium chloride on vapor pressure of 2-methyl-1-butanol
pvap	33.30	kPa	373.05	Effect of addition of calcium chloride on vapor pressure of 2-methyl-1-butanol
pvap	36.70	kPa	375.15	Effect of addition of calcium chloride on vapor pressure of 2-methyl-1-butanol
pvap	83.30	kPa	396.65	Vapor-Liquid Equilibrium of the Ethanol + 2-Methyl-1-butanol System

pvap	36.70	kPa	375.15	Vapor-Liquid Equilibrium of the Ethanol + 2-Methyl-1-butanol System
rfi	1.40860		298.15	Study of the Effects of Temperature and Pressure on the Thermodynamic and Acoustic Properties of 2-Methyl-1-butanol at Temperatures from 293K to 318K and Pressures up to 100MPa
rfi	1.40840		298.15	Bubble-Temperature Measurements on Some Binary Mixtures Formed by Tetrahydrofuran or Amyl Alcohol with Hydrocarbons, Chlorohydrocarbons, or Butanols at (94.6 or 95.8) kPa
rfi	1.40872		298.15	Density, Refractive Index, Speed of Sound at 298.15 K, and Vapor-Liquid Equilibria at 101.3 kPa for Binary Mixtures of Propanol + 2-Methyl-1-butanol and Propanol + 3-Methyl-1-butanol
rfi	1.40872		298.15	Density, Refractive Index, Speed of Sound at 298.15 K, and Vapor-Liquid Equilibria at 101.3 kPa for Binary Mixtures of Methanol + 2-Methyl-1-butanol and Ethanol + 2-Methyl-1-butanol

rfi	1.40950		293.15	Isobaric Vapor - Liquid Equilibria for the Ternary System of 2-Methyl-1-butanol, 3-Methyl-1-butanol, and Ethylene Glycol at 101.3 kPa	
rhol	796.39	kg/m3	323.15	Densities and Viscosities for Binary Liquid Mixtures of Biodiesel + 1-Pentanol, 2-Pentanol, or 2-Methyl-1-Butanol from (288.15 to 338.15) K at 0.1 MPa	
rhol	812.12	kg/m3	303.15	Densities and Viscosities for Binary Liquid Mixtures of Pentanol Isomers from (288.15 to 328.15) K at 0.1 MPa	
rhol	808.29	kg/m3	308.15	Densities and Viscosities for Binary Liquid Mixtures of Pentanol Isomers from (288.15 to 328.15) K at 0.1 MPa	
rhol	814.83	kg/m3	298.15	Vapor Liquid Equilibrium for 2-Methyl-1-butanol + Ethylbenzene + Xylene Isomers at 101.33 kPa	
rhol	804.41	kg/m3	313.15	Densities and Viscosities for Binary Liquid Mixtures of Pentanol Isomers from (288.15 to 328.15) K at 0.1 MPa	
rhol	800.47	kg/m3	318.15	Densities and Viscosities for Binary Liquid Mixtures of Pentanol Isomers from (288.15 to 328.15) K at 0.1 MPa	

rhol	796.47	kg/m3	323.15	Densities and Viscosities for Binary Liquid Mixtures of Pentanol Isomers from (288.15 to 328.15) K at 0.1 MPa
rhol	814.98	kg/m3	298.15	Excess molar enthalpies of the binary systems: (Dibutyl ether + isomers of pentanol) at T = (298.15 and 308.15) K
rhol	807.63	kg/m3	308.15	Excess molar enthalpies of the binary systems: (Dibutyl ether + isomers of pentanol) at T = (298.15 and 308.15) K
rhol	823.02	kg/m3	288.15	Densities and Viscosities for Binary Liquid Mixtures of Biodiesel + 1-Pentanol, 2-Pentanol, or 2-Methyl-1-Butanol from (288.15 to 338.15) K at 0.1 MPa
rhol	819.56	kg/m3	293.15	Densities and Viscosities for Binary Liquid Mixtures of Biodiesel + 1-Pentanol, 2-Pentanol, or 2-Methyl-1-Butanol from (288.15 to 338.15) K at 0.1 MPa
rhol	815.82	kg/m3	298.15	Densities and Viscosities for Binary Liquid Mixtures of Biodiesel + 1-Pentanol, 2-Pentanol, or 2-Methyl-1-Butanol from (288.15 to 338.15) K at 0.1 MPa

rhol	819.64	kg/m3	293.15	Densities and Viscosities for Binary Liquid Mixtures of Pentanol Isomers from (288.15 to 328.15) K at 0.1 MPa
rhol	808.20	kg/m3	308.15	Densities and Viscosities for Binary Liquid Mixtures of Biodiesel + 1-Pentanol, 2-Pentanol, or 2-Methyl-1-Butanol from (288.15 to 338.15) K at 0.1 MPa
rhol	804.33	kg/m3	313.15	Densities and Viscosities for Binary Liquid Mixtures of Biodiesel + 1-Pentanol, 2-Pentanol, or 2-Methyl-1-Butanol from (288.15 to 338.15) K at 0.1 MPa
rhol	800.39	kg/m3	318.15	Densities and Viscosities for Binary Liquid Mixtures of Biodiesel + 1-Pentanol, 2-Pentanol, or 2-Methyl-1-Butanol from (288.15 to 338.15) K at 0.1 MPa
rhol	792.41	kg/m3	328.15	Densities and Viscosities for Binary Liquid Mixtures of Pentanol Isomers from (288.15 to 328.15) K at 0.1 MPa
rhol	792.34	kg/m3	328.15	Densities and Viscosities for Binary Liquid Mixtures of Biodiesel + 1-Pentanol, 2-Pentanol, or 2-Methyl-1-Butanol from (288.15 to 338.15) K at 0.1 MPa

rhol	788.20	kg/m3	333.15	Densities and Viscosities for Binary Liquid Mixtures of Biodiesel + 1-Pentanol, 2-Pentanol, or 2-Methyl-1-Butanol from (288.15 to 338.15) K at 0.1 MPa
rhol	783.84	kg/m3	338.15	Densities and Viscosities for Binary Liquid Mixtures of Biodiesel + 1-Pentanol, 2-Pentanol, or 2-Methyl-1-Butanol from (288.15 to 338.15) K at 0.1 MPa
rhol	823.35	kg/m3	288.15	Densities and Viscosities for Binary Liquid Mixtures of Pentanol Isomers from (288.15 to 328.15) K at 0.1 MPa
rhol	812.03	kg/m3	303.15	Densities and Viscosities for Binary Liquid Mixtures of Biodiesel + 1-Pentanol, 2-Pentanol, or 2-Methyl-1-Butanol from (288.15 to 338.15) K at 0.1 MPa
rhol	815.85	kg/m3	298.15	Densities and Viscosities for Binary Liquid Mixtures of Pentanol Isomers from (288.15 to 328.15) K at 0.1 MPa
srf	0.02	N/m	298.15	Surface Tensions of Three Amyl Alcohol + Ethanol Binary Mixtures from (293.15 to 323.15) K
srf	0.02	N/m	303.15	Surface Tensions of Three Amyl Alcohol + Ethanol Binary Mixtures from (293.15 to 323.15) K

srf	0.02	N/m	308.15	Surface Tensions of Three Amyl Alcohol + Ethanol Binary Mixtures from (293.15 to 323.15) K
srf	0.02	N/m	313.15	Surface Tensions of Three Amyl Alcohol + Ethanol Binary Mixtures from (293.15 to 323.15) K
srf	0.02	N/m	318.15	Surface Tensions of Three Amyl Alcohol + Ethanol Binary Mixtures from (293.15 to 323.15) K
srf	0.02	N/m	323.15	Surface Tensions of Three Amyl Alcohol + Ethanol Binary Mixtures from (293.15 to 323.15) K
srf	0.02	N/m	293.15	Surface Tensions of Three Amyl Alcohol + Ethanol Binary Mixtures from (293.15 to 323.15) K

Correlations

Information	Value
Property code	pvap
Equation	ln(Pvp) = A + B/(T + C)
Coeff. A	1.43135e+01
Coeff. B	-2.78286e+03
Coeff. C	-1.14962e+02
Temperature range (K), min.	313.37
Temperature range (K), max.	424.10

Information	Value
Property code	pvap
Equation	ln(Pvp) = A + B/T + C*ln(T) + D*T^2
Coeff. A	2.88019e+02
Coeff. B	-1.78736e+04
Coeff. C	-4.06446e+01
Coeff. D	2.96397e-05

Temperature range (K), min.	203.00
Temperature range (K), max.	565.00

Sources

Isothermal vapor liquid equilibria for different binary mixtures involved in the Kharchenko distillation:	https://www.doi.org/10.1016/j.fluid.2008.03.002
Vapor-Liquid Equilibrium for 2-Methyl-1-butanol + Ethylbenzene + Xylene Isomers at 101.33 kPa:	https://www.doi.org/10.1021/je301220s
Densities and Viscosities for Binary Liquid Mixtures of Pentanol Isomers from (288.15 to 328.15) K at 0.1 MPa:	https://www.doi.org/10.1021/acs.jced.8b00979
Excess molar enthalpies of the binary systems: (Dibutyl ether + isomers of pentanol) at T = (298.15 and 308.15) K:	https://www.doi.org/10.1016/j.jct.2015.03.021
Effect of addition of calcium chloride on vapor pressure of 2-methyl-1-butanol:	https://www.doi.org/10.1021/je0340147
The Yaws Handbook of Vapor Pressure:	https://www.sciencedirect.com/book/9780128029992/the-yaws-handbook-of-vapor-pressure
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Heat Capacities in the Solid and in the Liquid Phase of Isomeric Pentanols:	https://www.doi.org/10.1021/je060411g
Density, Refractive Index, Speed of Sound at 298.15 K, and Vapor-Liquid Equilibria at 101.3 kPa for Binary Mixtures of 1-Propanol + 2-Methyl-1-butanol and 1-Propanol + 2-Methyl-2-butanol:	https://www.doi.org/10.1021/je0502142
Mutual diffusion coefficients of 3-methyl-1-butanol + n-heptane and 2-methyl-1-butanol + n-heptane mixtures at 293.15 K:	https://www.doi.org/10.1016/j.jct.2018.10.023
Bubble-Temperature Measurements on Some Binary Mixtures:	https://www.doi.org/10.1021/je030141r
Enthalpy of Solution of Tetrahydrofuran in Amyl Alcohols at 298.15 K:	https://www.doi.org/10.1021/je901063s
Solubility of Anthracene in Binary Mixtures of Propanol + 2-Methyl-1-propanol + 2-Methyl-2-propanol:	http://onschallenge.wikispaces.com/file/view/AqueousDataset002.xlsx/351826032/AqueousDa
Isobaric Vapor Liquid Equilibria for the Binary and Ternary Systems of 2-Methyl-1-propanol, 2-Methyl-2-propanol and 2-Methyl-1-butanol at 101.3 kPa:	https://www.doi.org/10.1021/je0201727
Vapor pressure of selected aliphatic alcohols by ebulliometry. Part 2 : Crippen Method:	https://www.doi.org/10.1021/je3012925
Study of the Effects of Temperature and Pressure on the Thermodynamic and Acoustic Properties of 2-Methyl-1-butanol at Temperatures from (293.15 to 313.15) K and Pressures up to 100 MPa. Electrolyte Models:	https://www.doi.org/10.1021/je0500330
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Solubility in Binary Solvent Mixtures: 2-Methylbutan-1-ol), (Propyl Acetate + Hexan-1-ol, or Propyl Acetate + Heptan-1-ol) at 101.3 kPa:	https://www.doi.org/10.1021/je060399s
Antiaraenn Dissolved in Alcohol + Hexane Binary Mixtures at 298.2 K:	https://www.doi.org/10.1021/je050416y
Isobaric Vapor Liquid Equilibria for Binary Mixtures Containing 2-Methyl-1-propanol, 2-Methyl-1-butanol, and 2-Methyl-1-propanol + 1-Pentanol:	https://www.doi.org/10.1021/acs.jced.7b00996
The influence of cation, anion and temperature of the liquid-liquid equilibrium of 2-methyl-1-propanol + 3-methyl-1-butanol, and 2-Methyl-1-propanol + 1-Pentanol:	https://www.doi.org/10.1021/je030216r
McGowan Method:	http://link.springer.com/article/10.1007/BF02311772
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Isobaric Vapor Liquid Equilibria for Binary Mixtures Containing 2-Methyl-1-propanol, 2-Methyl-1-butanol, and 2-Methyl-1-propanol + 1-Pentanol:	https://www.doi.org/10.1016/j.fluid.2011.03.018
Joback Method:	https://en.wikipedia.org/wiki/Joback_method
The influence of cation, anion and temperature of the liquid-liquid equilibrium of 2-methyl-1-propanol + 3-methyl-1-butanol, and 2-Methyl-1-propanol + 1-Pentanol:	https://www.doi.org/10.1021/je030216r
Joback Method: Alkanes, pentanols-water systems:	https://en.wikipedia.org/wiki/Joback_method

Henry's law constants and infinite dilution activity coefficients of cis-2-butene, dimethylether, chloroethane, and 1,1-dichloroethane in methanol, 1-propanol, 2-propanol, 1-butanol, 2-butanol, isobutanol, tert-butanol, 1-pentanol, 2-pentanol, 3-pentanol, 3-methyl-1-butanol, and 2-methyl-2-butanol: https://www.doi.org/10.1016/j.jct.2005.08.004

Liquid-Liquid Equilibria for the Ternary Systems of Water + 2-Butanone + 2-Methylbutan-1-ol, 2-Ethylbutan-1-ol, and 4-Methylpentan-2-ol at 298.15 and 308.15 K https://www.doi.org/10.1021/acs.jced.6b00945

Legend

cpg:	Ideal gas heat capacity
dvisc:	Dynamic viscosity
gf:	Standard Gibbs free energy of formation
hf:	Enthalpy of formation at standard conditions
hfus:	Enthalpy of fusion at standard conditions
hvap:	Enthalpy of vaporization at standard conditions
kvisc:	Kinematic viscosity
log10ws:	Log10 of Water solubility in mol/l
logp:	Octanol/Water partition coefficient
mcvol:	McGowan's characteristic volume
pc:	Critical Pressure
pvap:	Vapor pressure
rfi:	Refractive Index
rhol:	Liquid Density
srf:	Surface Tension
tb:	Normal Boiling Point Temperature
tc:	Critical Temperature
tf:	Normal melting (fusion) point
tt:	Triple Point Temperature
vc:	Critical Volume