

# 2,3-Pentanedione

<b>Other names:</b>	2,3-Pentadione 2,3-Pentandione Acetylpropionyl CH <sub>3</sub> C(O)C(O)C <sub>2</sub> H <sub>5</sub> Pentan-2,3-dione Pentane-2,3-dione
<b>Inchi:</b>	InChI=1S/C5H8O2/c1-3-5(7)4(2)6/h3H2,1-2H3
<b>InchiKey:</b>	TZMFJUDUGYTVRY-UHFFFAOYSA-N
<b>Formula:</b>	C <sub>5</sub> H <sub>8</sub> O <sub>2</sub>
<b>SMILES:</b>	CCC(=O)C(C)=O
<b>Mol. weight [g/mol]:</b>	100.12
<b>CAS:</b>	600-14-6

## Physical Properties

Property code	Value	Unit	Source
gf	-266.62	kJ/mol	Joback Method
hf	-371.69	kJ/mol	Joback Method
hfus	7.84	kJ/mol	Phase diagrams of binary systems containing n-alkanes, or cyclohexane, or 1-alkanols and 2,3-pentanedione at atmospheric and high pressure
hvap	40.22	kJ/mol	Joback Method
log10ws	-0.48		Crippen Method
logp	0.554		Crippen Method
mcvol	84.450	ml/mol	McGowan Method
pc	4046.64	kPa	Joback Method
rinpol	681.00		NIST Webbook
rinpol	700.00		NIST Webbook
rinpol	700.00		NIST Webbook
rinpol	702.00		NIST Webbook
rinpol	696.00		NIST Webbook
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rinpol	674.00	NIST Webbook
rinpol	676.00	NIST Webbook
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ripol	1082.00		NIST Webbook
tb	384.20	K	NIST Webbook
tb	381.20	K	NIST Webbook
tc	613.03	K	Joback Method
tf	245.97	K	Joback Method
vc	0.328	m <sup>3</sup> /kmol	Joback Method

## Temperature Dependent Properties

Property code	Value	Unit	Temperature [K]	Source
cpg	200.02	J/molxK	613.03	Joback Method
cpg	164.27	J/molxK	453.45	Joback Method
cpg	172.09	J/molxK	485.37	Joback Method
cpg	179.57	J/molxK	517.28	Joback Method
cpg	186.71	J/molxK	549.20	Joback Method
cpg	193.52	J/molxK	581.11	Joback Method
cpg	156.11	J/molxK	421.54	Joback Method
dvisc	0.0004592	Paxs	392.28	Joback Method
dvisc	0.0005966	Paxs	363.02	Joback Method
dvisc	0.0008115	Paxs	333.75	Joback Method
dvisc	0.0011710	Paxs	304.49	Joback Method
dvisc	0.0018268	Paxs	275.23	Joback Method
dvisc	0.0003666	Paxs	421.54	Joback Method

dvisc	0.0031678	Paxs	245.97	Joback Method
hfust	7.84	kJ/mol	221.20	NIST Webbook
pvap	45.32	kPa	358.85	Vapor Liquid Equilibrium for Binary Systems of 2,3-Pentanedione with Diacetyl and Acetone
pvap	55.21	kPa	364.69	Vapor Liquid Equilibrium for Binary Systems of 2,3-Pentanedione with Diacetyl and Acetone
pvap	60.20	kPa	367.09	Vapor Liquid Equilibrium for Binary Systems of 2,3-Pentanedione with Diacetyl and Acetone
pvap	65.19	kPa	369.44	Vapor Liquid Equilibrium for Binary Systems of 2,3-Pentanedione with Diacetyl and Acetone
pvap	70.18	kPa	371.61	Vapor Liquid Equilibrium for Binary Systems of 2,3-Pentanedione with Diacetyl and Acetone
pvap	75.27	kPa	373.80	Vapor Liquid Equilibrium for Binary Systems of 2,3-Pentanedione with Diacetyl and Acetone
pvap	40.33	kPa	355.74	Vapor Liquid Equilibrium for Binary Systems of 2,3-Pentanedione with Diacetyl and Acetone
pvap	35.34	kPa	352.15	Vapor Liquid Equilibrium for Binary Systems of 2,3-Pentanedione with Diacetyl and Acetone



pvap	30.35	kPa	348.13	Vapor Liquid Equilibrium for Binary Systems of 2,3-Pentanedione with Diacetyl and Acetone
pvap	25.35	kPa	343.73	Vapor Liquid Equilibrium for Binary Systems of 2,3-Pentanedione with Diacetyl and Acetone
pvap	20.36	kPa	338.38	Vapor Liquid Equilibrium for Binary Systems of 2,3-Pentanedione with Diacetyl and Acetone
pvap	15.37	kPa	331.84	Vapor Liquid Equilibrium for Binary Systems of 2,3-Pentanedione with Diacetyl and Acetone
pvap	50.21	kPa	361.89	Vapor Liquid Equilibrium for Binary Systems of 2,3-Pentanedione with Diacetyl and Acetone
rho1	982.00	kg/m3	292.15	Vapor Liquid Equilibrium Data for 2,3-Pentanedione + (Acetaldehyde or Acetone) at (100, 150, and 200) kPa

## Correlations

Information	Value
Property code	pvap
Equation	$\ln(P_{vp}) = A + B/(T + C)$
Coeff. A	1.45607e+01
Coeff. B	-3.35146e+03
Coeff. C	-4.71100e+01
Temperature range (K), min.	281.92
Temperature range (K), max.	409.46

# Sources

Phase diagrams of binary systems containing n-alkanes, or cyclohexane, or 2,3-pentanedione at atmospheric and high pressure: Joback Method:

<https://www.doi.org/10.1016/j.fluid.2006.02.001>

Crippen Method:

<http://link.springer.com/article/10.1007/BF02311772>

[https://en.wikipedia.org/wiki/Joback\\_method](https://en.wikipedia.org/wiki/Joback_method)

Vapor Liquid Equilibrium Data for 2,3-Pentanedione + (Acetaldehyde or Acetone) at Infinite Dilution: Activity Coefficients and Partition of Volatile Compounds in Water: Effect of Temperature and Salt Concentration: NIST Webbook

[https://www.chemeo.com/doc/models/crippen\\_log10ws](https://www.chemeo.com/doc/models/crippen_log10ws)

<https://www.doi.org/10.1021/acs.jced.8b01196>

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The Yaws Handbook of Vapor Pressure: Crippen Method:

<https://www.sciencedirect.com/book/9780128029992/the-yaws-handbook-of-vapor-pressure>

<http://pubs.acs.org/doi/abs/10.1021/ci9903071>

Vapor Liquid Equilibrium for Binary Systems of 2,3-Pentanedione with Diacetyl and Acetone:

<https://www.doi.org/10.1021/je7005924>

# Legend

<b>cpg:</b>	Ideal gas heat capacity
<b>dvisc:</b>	Dynamic viscosity
<b>gf:</b>	Standard Gibbs free energy of formation
<b>hf:</b>	Enthalpy of formation at standard conditions
<b>hfus:</b>	Enthalpy of fusion at standard conditions
<b>hfust:</b>	Enthalpy of fusion at a given temperature
<b>hvap:</b>	Enthalpy of vaporization at standard conditions
<b>log10ws:</b>	Log10 of Water solubility in mol/l
<b>logp:</b>	Octanol/Water partition coefficient
<b>mccvol:</b>	McGowan's characteristic volume
<b>pc:</b>	Critical Pressure
<b>pvap:</b>	Vapor pressure
<b>rho:</b>	Liquid Density
<b>rinpol:</b>	Non-polar retention indices
<b>ripol:</b>	Polar retention indices
<b>tb:</b>	Normal Boiling Point Temperature
<b>tc:</b>	Critical Temperature
<b>tf:</b>	Normal melting (fusion) point
<b>vc:</b>	Critical Volume

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