1-Octanol

Other names:	1-Hydroxyoctane
	1-Octyl alcohol
	Alcohol C-8
	Alfol 8
	Caprylic alcohol
	Dytol M-83
	Emery 3322
	Emery 3324
	Epal 8
	HEPTYL CARBINOL
	Lorol 20
	Lorol C 8-98
	N-OCTYL ALCOHOL
	NSC 9823
	OCTYL ALCOHOL
	Octan-1-ol
	Octanol
	Octanol-(1)
	Octilin
	Octyl alcohol, normal-primary
	Prim-n-octyl alcohol
	Primary octyl alcohol
	Sipol L8
	capryl alcohol
	n-Heptyl carbinol
	n-Octan-1-ol
	n-Octanol
Inchi:	InChI=1S/C8H18O/c1-2-3-4-5-6-7-8-9/h9H,2-8H2,1H3
InchiKey:	KBPLFHHGFOOTCA-UHFFFAOYSA-N
Formula:	C8H18O
SMILES:	CCCCCCCCO
Mol. weight [g/mol]:	130.23
CAS:	111-87-5

Physical Properties

Property code	Value	Unit	Source

af	0.5870		KDB
affp	799.00	kJ/mol	NIST Webbook
chl	-5294.00 ± 0.60	kJ/mol	NIST Webbook
chl	-5295.50 ± 1.00	kJ/mol	NIST Webbook
chl	-5285.70 ± 2.70	kJ/mol	NIST Webbook
chl	-5285.60	kJ/mol	NIST Webbook
dm	2.00	debye	KDB
dvisc	0.0075080	Pa×s	Study of molecular interactions in binary liquid mixtures of 1-octanol with n-hexane, n-octane, and n-decane using volumetric, viscometric, and acoustic properties
dvisc	0.0071430	Pa×s	Volumetric and Viscometric Properties of Binary Liquid Mixtures of Ethylene Glycol Monomethyl Ether + 1-Hexanol, 1-Octanol, and 1-Decanol at Temperatures of T = (293.15, 298.15, 303.15, and 308.15) K
gf	-120.20	kJ/mol	KDB
hf	-360.10	kJ/mol	KDB
hfl	-435.00 ± 2.70	kJ/mol	NIST Webbook
hfl	-425.20 ± 1.10	kJ/mol	NIST Webbook
hfl	-428.00 ± 1.10	kJ/mol	NIST Webbook
hfl	-426.60 ± 0.60	kJ/mol	NIST Webbook
hfus	23.70	kJ/mol	Measurements, Correlations, and Mod. UNIFAC (Do) Prediction of (Solid-Liquid) Phase Equilibria Diagrams in Binary Systems (Aliphatic Ketone + an Alcohol)
hfus	25.14	kJ/mol	Heat Capacities and Derived Thermodynamic Functions of 1-Hexanol, 1-Heptanol, 1-Octanol, and 1-Decanol between 5 K and 390 K
hsub	100.40	kJ/mol	NIST Webbook
hvap	50.08	kJ/mol	Joback Method
log10ws	-2.39		Aqueous Solubility Prediction Method
log10ws	-2.39		Estimated Solubility Method
logp	2.339		Crippen Method
mcvol	129.450	ml/mol	McGowan Method
nfpaf	%!d(float64=2)		KDB
nfpah	%!d(float64=1)		KDB
pc	2860.00	kPa	NIST Webbook

pc	1777.00 ± 20.00	kPa	NIST Webbook
pc	2777.00 ± 20.00	kPa	NIST Webbook
pc	2860.00	kPa	NIST Webbook
pc	2777.00	kPa	KDB
pc	2780.00 ± 30.00	kPa	NIST Webbook
pc	2850.00 ± 50.00	kPa	NIST Webbook
rhoc	261.76 ± 3.91	kg/m3	NIST Webbook
rhoc	265.66	kg/m3	NIST Webbook
rhoc	256.55 ± 6.51	kg/m3	NIST Webbook
rhoc	265.66 ± 2.60	kg/m3	NIST Webbook
rinpol	1065.00		NIST Webbook
rinpol	1074.00		NIST Webbook
rinpol	1068.00		NIST Webbook
rinpol	1067.00		NIST Webbook
rinpol	1059.00		NIST Webbook
rinpol	1066.00		NIST Webbook
rinpol	1039.00		NIST Webbook
rinpol	1082.00		NIST Webbook
rinpol	1076.00		NIST Webbook
rinpol	1069.00		NIST Webbook
rinpol	1051.00		NIST Webbook
rinpol	1052.00		NIST Webbook
rinpol	1083.00		NIST Webbook
rinpol	1053.00		NIST Webbook
rinpol	1075.00		NIST Webbook
rinpol	1070.00		NIST Webbook
rinpol	1073.00		NIST Webbook
rinpol	1072.00		NIST Webbook
rinpol	1071.00		NIST Webbook
rinpol	1056.00		NIST Webbook
rinpol	1080.00		NIST Webbook
rinpol	1061.00		NIST Webbook
rinpol	176.31		NIST Webbook
rinpol	177.62		NIST Webbook
rinpol	175.63		NIST Webbook
rinpol	1083.00		NIST Webbook
rinpol	1059.00		NIST Webbook
rinpol	1058.00		NIST Webbook
rinpol	1050.00		NIST Webbook
rinpol	1056.00		NIST Webbook
rinpol	1061.00		NIST Webbook
rinpol	1061.00		NIST Webbook
rinpol	1053.00		NIST Webbook
rinpol	1053.00		NIST Webbook

rinpol	1084.00	NIST Webbook
rinpol	1056.00	NIST Webbook
rinpol	1056.00	NIST Webbook
rinpol	1056.00	NIST Webbook
rinpol	1052.00	NIST Webbook
rinpol	1079.00	NIST Webbook
rinpol	1053.00	NIST Webbook
rinpol	1080.00	NIST Webbook
rinpol	1065.00	NIST Webbook
rinpol	1087.00	NIST Webbook
rinpol	1077.00	NIST Webbook
rinpol	1075.00	NIST Webbook
rinpol	1069.70	NIST Webbook
rinpol	1061.00	NIST Webbook
rinpol	1057.00	NIST Webbook
rinpol	1084.00	NIST Webbook
rinpol	1075.00	NIST Webbook
rinpol	1071.00	NIST Webbook
rinpol	1061.00	NIST Webbook
rinpol	1066.00	NIST Webbook
rinpol	1070.00	NIST Webbook
rinpol	1071.00	NIST Webbook
rinpol	1071.00	NIST Webbook
rinpol	1064.00	NIST Webbook
rinpol	1055.00	NIST Webbook
rinpol	1070.00	NIST Webbook
rinpol	1072.00	NIST Webbook
rinpol	1064.00	NIST Webbook
rinpol	1070.00	NIST Webbook
rinpol	1069.00	NIST Webbook
rinpol	1063.00	NIST Webbook
rinpol	1070.00	NIST Webbook
rinpol	1038.00	NIST Webbook
rinpol	1072.00	NIST Webbook
rinpol	1072.00	NIST Webbook
rinpol	1064.00	NIST Webbook
rinpol	1073.00	NIST Webbook
rinpol	1067.00	NIST Webbook
rinpol	1069.00	NIST Webbook
rinpol	1068.34	NIST Webbook
rinpol	1073.00	NIST Webbook
rinpol	1071.00	NIST Webbook
rinpol	1052.00	NIST Webbook
rinpol	1070.00	NIST Webbook

rinpol	1044.00	NIST Webbook
rinpol	1064.00	NIST Webbook
rinpol	1072.00	NIST Webbook
rinpol	1049.00	NIST Webbook
rinpol	1075.00	NIST Webbook
rinpol	1072.00	NIST Webbook
rinpol	1074.00	NIST Webbook
rinpol	1071.00	NIST Webbook
rinpol	1074.00	NIST Webbook
rinpol	1074.00	NIST Webbook
rinpol	1068.00	NIST Webbook
rinpol	1076.00	NIST Webbook
rinpol	1080.00	NIST Webbook
rinpol	1082.00	NIST Webbook
rinpol	1066.00	NIST Webbook
rinpol	1065.00	NIST Webbook
rinpol	1070.00	NIST Webbook
rinpol	1056.00	NIST Webbook
rinpol	1068.00	NIST Webbook
rinpol	1060.00	NIST Webbook
rinpol	1069.00	NIST Webbook
rinpol	1060.00	NIST Webbook
rinpol	1080.00	NIST Webbook
rinpol	1077.00	NIST Webbook
rinpol	1074.00	NIST Webbook
rinpol	1063.00	NIST Webbook
rinpol	1052.00	NIST Webbook
rinpol	1061.00	NIST Webbook
rinpol	1057.00	NIST Webbook
rinpol	1083.00	NIST Webbook
rinpol	1061.00	NIST Webbook
rinpol	1057.00	NIST Webbook
rinpol	1052.00	NIST Webbook
rinpol	1056.00	NIST Webbook
rinpol	1057.00	NIST Webbook
rinpol	1058.00	NIST Webbook
rinpol	1075.00	NIST Webbook
rinpol	1069.00	NIST Webbook
rinpol	1054.00	NIST Webbook
rinpol	1053.00	NIST Webbook
rinpol	1054.00	NIST Webbook
rinpol	1054.00	NIST Webbook
rinpol	1054.00	NIST Webbook
rinpol	1054.00	NIST Webbook

rinpol	1061.00	NIST Webbook
rinpol	1051.00	NIST Webbook
rinpol	1073.00	NIST Webbook
rinpol	1079.00	NIST Webbook
rinpol	1060.00	NIST Webbook
rinpol	1053.00	NIST Webbook
rinpol	1052.00	NIST Webbook
rinpol	1072.00	NIST Webbook
rinpol	1072.00	NIST Webbook
rinpol	1095.00	NIST Webbook
rinpol	1068.34	NIST Webbook
rinpol	1061.00	NIST Webbook
rinpol	1059.00	NIST Webbook
rinpol	1068.00	NIST Webbook
rinpol	1074.00	NIST Webbook
rinpol	1074.00	NIST Webbook
rinpol	1050.00	NIST Webbook
rinpol	1048.00	NIST Webbook
rinpol	1074.00	NIST Webbook
rinpol	1075.00	NIST Webbook
rinpol	1068.00	NIST Webbook
rinpol	1053.00	NIST Webbook
rinpol	1069.00	NIST Webbook
rinpol	1073.00	NIST Webbook
rinpol	1084.00	NIST Webbook
rinpol	1072.00	NIST Webbook
rinpol	1072.00	NIST Webbook
rinpol	1074.00	NIST Webbook
rinpol	1056.00	NIST Webbook
rinpol	1056.80	NIST Webbook
rinpol	1053.00	NIST Webbook
rinpol	1053.00	NIST Webbook
rinpol	1076.00	NIST Webbook
rinpol	1050.00	NIST Webbook
rinpol	1052.00	NIST Webbook
rinpol	1052.00	NIST Webbook
rinpol	1073.00	NIST Webbook
rinpol	1048.00	NIST Webbook
rinpol	1072.00	NIST Webbook
rinpol	1069.00	NIST Webbook
rinpol	1070.00	NIST Webbook
rinpol	1052.00	NIST Webbook
rinpol	1061.00	NIST Webbook
rinpol	1070.00	NIST Webbook

rinpol	1067.00	NIST Webbook
rinpol	1067.00	NIST Webbook
rinpol	1070.00	NIST Webbook
rinpol	1070.00	NIST Webbook
rinpol	1091.00	NIST Webbook
rinpol	1077.00	NIST Webbook
rinpol	1066.00	NIST Webbook
rinpol	1054.00	NIST Webbook
rinpol	1052.00	NIST Webbook
rinpol	1039.00	NIST Webbook
rinpol	1075.00	NIST Webbook
rinpol	1059.00	NIST Webbook
rinpol	1069.00	NIST Webbook
rinpol	1076.00	NIST Webbook
rinpol	1070.00	NIST Webbook
rinpol	1069.00	NIST Webbook
rinpol	1066.00	NIST Webbook
rinpol	1048.00	NIST Webbook
rinpol	1061.00	NIST Webbook
rinpol	1065.00	NIST Webbook
rinpol	1065.00	NIST Webbook
rinpol	1039.00	NIST Webbook
rinpol	1089.00	NIST Webbook
rinpol	1068.00	NIST Webbook
rinpol	1069.00	NIST Webbook
rinpol	1051.00	NIST Webbook
rinpol	1070.00	NIST Webbook
rinpol	1068.00	NIST Webbook
rinpol	1060.00	NIST Webbook
rinpol	1069.00	NIST Webbook
rinpol	1069.00	NIST Webbook
rinpol	1070.00	NIST Webbook
rinpol	1073.00	NIST Webbook
rinpol	1048.00	NIST Webbook
rinpol	1075.00	NIST Webbook
rinpol	1054.00	NIST Webbook
rinpol	1070.00	NIST Webbook
rinpol	1070.00	NIST Webbook
rinpol	1053.00	NIST Webbook
rinpol	1057.00	NIST Webbook
rinpol	1069.00	NIST Webbook
rinpol	1076.00	NIST Webbook
rinpol	1077.00	NIST Webbook
rinpol	1074.00	NIST Webbook

rinpol	1076.00	NIST Webbook
rinpol	1046.00	NIST Webbook
rinpol	1069.00	NIST Webbook
rinpol	1091.00	NIST Webbook
rinpol	1052.00	NIST Webbook
rinpol	1068.00	NIST Webbook
rinpol	1071.00	NIST Webbook
rinpol	1082.00	NIST Webbook
rinpol	1067.00	NIST Webbook
rinpol	1070.00	NIST Webbook
rinpol	1069.00	NIST Webbook
rinpol	1069.00	NIST Webbook
rinpol	1061.00	NIST Webbook
rinpol	1072.00	NIST Webbook
rinpol	1074.00	NIST Webbook
rinpol	1055.00	NIST Webbook
rinpol	1076.00	NIST Webbook
rinpol	1076.00	NIST Webbook
rinpol	1073.00	NIST Webbook
rinpol	1071.00	NIST Webbook
rinpol	1068.00	NIST Webbook
rinpol	1063.00	NIST Webbook
rinpol	1069.00	NIST Webbook
rinpol	1095.00	NIST Webbook
rinpol	1078.00	NIST Webbook
rinpol	1072.00	NIST Webbook
rinpol	1071.00	NIST Webbook
rinpol	1057.00	NIST Webbook
rinpol	1055.00	NIST Webbook
rinpol	1071.00	NIST Webbook
rinpol	1053.00	NIST Webbook
rinpol	1051.00	NIST Webbook
rinpol	1051.00	NIST Webbook
rinpol	1052.00	NIST Webbook
rinpol	1054.00	NIST Webbook
rinpol	1054.00	NIST Webbook
rinpol	1052.00	NIST Webbook
rinpol	1059.00	NIST Webbook
rinpol	1070.00	NIST Webbook
rinpol	1081.00	NIST Webbook
rinpol	1058.00	NIST Webbook
rinpol	1052.00	NIST Webbook
rinpol	1074.00	NIST Webbook
rinpol	1074.00	NIST Webbook

rinpol	1070.00	NIST Webbook
rinpol	1070.00	NIST Webbook
rinpol	1051.00	NIST Webbook
rinpol	1072.00	NIST Webbook
rinpol	1068.00	NIST Webbook
rinpol	1080.00	NIST Webbook
rinpol	1068.00	NIST Webbook
rinpol	1068.00	NIST Webbook
rinpol	1080.10	NIST Webbook
rinpol	1070.00	NIST Webbook
rinpol	1084.00	NIST Webbook
rinpol	1057.00	NIST Webbook
rinpol	1060.00	NIST Webbook
rinpol	1058.00	NIST Webbook
rinpol	1076.00	NIST Webbook
rinpol	1074.00	NIST Webbook
rinpol	1072.00	NIST Webbook
rinpol	1072.00	NIST Webbook
rinpol	1072.00	NIST Webbook
rinpol	1072.00	NIST Webbook
rinpol	1072.00	NIST Webbook
rinpol	1071.00	NIST Webbook
rinpol	1071.00	NIST Webbook
rinpol	1071.00	NIST Webbook
rinpol	1069.00	NIST Webbook
rinpol	1068.00	NIST Webbook
rinpol	1052.00	NIST Webbook
rinpol	1053.00	NIST Webbook
rinpol	1052.00	NIST Webbook
rinpol	1071.00	NIST Webbook
rinpol	1072.00	NIST Webbook
rinpol	1072.00	NIST Webbook
rinpol	1054.00	NIST Webbook
rinpol	1081.00	NIST Webbook
rinpol	1083.00	NIST Webbook
rinpol	1069.00	NIST Webbook
rinpol	1051.00	NIST Webbook
rinpol	1060.00	NIST Webbook
rinpol	1071.00	NIST Webbook
rinpol	1057.00	NIST Webbook
rinpol	1071.00	NIST Webbook
rinpol	1069.00	NIST Webbook
rinpol	1069.00	NIST Webbook
rinpol	1061.70	NIST Webbook

rinpol	1061.70	NIST Webbook
rinpol	1072.00	NIST Webbook
rinpol	1070.00	NIST Webbook
rinpol	1061.00	NIST Webbook
rinpol	1068.00	NIST Webbook
rinpol	1080.00	NIST Webbook
rinpol	1052.00	NIST Webbook
rinpol	1054.00	NIST Webbook
rinpol	1067.00	NIST Webbook
rinpol	1055.40	NIST Webbook
rinpol	1056.00	NIST Webbook
rinpol	1067.00	NIST Webbook
rinpol	1065.00	NIST Webbook
rinpol	1064.00	NIST Webbook
rinpol	1065.00	NIST Webbook
rinpol	1065.00	NIST Webbook
rinpol	1070.00	NIST Webbook
rinpol	1069.00	NIST Webbook
rinpol	1054.00	NIST Webbook
rinpol	1054.00	NIST Webbook
rinpol	1070.00	NIST Webbook
rinpol	1072.00	NIST Webbook
rinpol	1070.00	NIST Webbook
rinpol	1073.00	NIST Webbook
rinpol	1064.00	NIST Webbook
rinpol	1051.00	NIST Webbook
rinpol	1068.70	NIST Webbook
rinpol	1071.00	NIST Webbook
rinpol	1080.00	NIST Webbook
rinpol	1053.20	NIST Webbook
rinpol	1068.00	NIST Webbook
rinpol	1054.00	NIST Webbook
rinpol	1073.00	NIST Webbook
rinpol	1074.00	NIST Webbook
rinpol	1057.00	NIST Webbook
rinpol	1072.00	NIST Webbook
rinpol	1068.00	NIST Webbook
rinpol	1071.00	NIST Webbook
rinpol	1070.00	NIST Webbook
rinpol	1070.00	NIST Webbook
rinpol	1075.00	NIST Webbook
rinpol	1066.50	NIST Webbook
rinpol	1066.90	NIST Webbook
rinpol	1057.00	NIST Webbook

rinpol	1084.00	NIST Webbook
rinpol	1081.00	NIST Webbook
rinpol	1066.90	NIST Webbook
rinpol	1063.00	NIST Webbook
rinpol	1064.00	NIST Webbook
rinpol	1070.00	NIST Webbook
rinpol	1065.00	NIST Webbook
rinpol	1068.00	NIST Webbook
rinpol	1050.00	NIST Webbook
rinpol	1084.00	NIST Webbook
rinpol	1074.00	NIST Webbook
rinpol	1068.00	NIST Webbook
rinpol	1072.00	NIST Webbook
rinpol	1070.60	NIST Webbook
rinpol	1076.00	NIST Webbook
rinpol	1051.43	NIST Webbook
rinpol	1051.79	NIST Webbook
rinpol	1052.54	NIST Webbook
rinpol	1052.35	NIST Webbook
rinpol	1060.90	NIST Webbook
rinpol	1060.50	NIST Webbook
rinpol	1059.90	NIST Webbook
rinpol	1059.70	NIST Webbook
rinpol	1060.90	NIST Webbook
rinpol	1060.50	NIST Webbook
rinpol	1059.90	NIST Webbook
rinpol	1059.70	NIST Webbook
rinpol	1052.99	NIST Webbook
rinpol	1055.41	NIST Webbook
rinpol	1053.90	NIST Webbook
rinpol	1034.80	NIST Webbook
rinpol	1042.30	NIST Webbook
rinpol	1051.70	NIST Webbook
rinpol	1082.00	NIST Webbook
rinpol	1051.00	NIST Webbook
rinpol	1052.00	NIST Webbook
rinpol	1053.00	NIST Webbook
rinpol	1053.00	NIST Webbook
rinpol	1053.00	NIST Webbook
rinpol	1054.00	NIST Webbook
rinpol	1054.00	NIST Webbook
rinpol	1054.00	NIST Webbook
rinpol	1055.00	NIST Webbook
rinpol	1057.00	NIST Webbook

rinpol	1060.00	NIST Webbook
rinpol	1061.00	NIST Webbook
rinpol	1066.00	NIST Webbook
rinpol	1049.00	NIST Webbook
rinpol	1050.00	NIST Webbook
rinpol	1049.00	NIST Webbook
rinpol	1042.00	NIST Webbook
rinpol	1057.00	NIST Webbook
rinpol	1060.00	NIST Webbook
rinpol	1058.00	NIST Webbook
rinpol	1044.00	NIST Webbook
rinpol	1064.00	NIST Webbook
rinpol	1058.00	NIST Webbook
rinpol	1060.00	NIST Webbook
rinpol	1042.00	NIST Webbook
rinpol	1061.00	NIST Webbook
rinpol	1071.00	NIST Webbook
rinpol	1073.00	NIST Webbook
rinpol	1059.00	NIST Webbook
rinpol	1059.00	NIST Webbook
rinpol	1068.00	NIST Webbook
rinpol	1071.00	NIST Webbook
rinpol	1062.00	NIST Webbook
rinpol	1053.00	NIST Webbook
rinpol	1060.00	NIST Webbook
rinpol	1078.00	NIST Webbook
rinpol	1078.00	NIST Webbook
rinpol	1060.00	NIST Webbook
rinpol	1053.00	NIST Webbook
rinpol	1060.00	NIST Webbook
rinpol	1059.00	NIST Webbook
rinpol	1061.00	NIST Webbook
rinpol	1063.00	NIST Webbook
rinpol	1061.00	NIST Webbook
rinpol	1062.00	NIST Webbook
rinpol	1067.00	NIST Webbook
rinpol	1070.00	NIST Webbook
rinpol	1061.00	NIST Webbook
rinpol	1063.00	NIST Webbook
rinpol	1055.00	NIST Webbook
rinpol	1060.00	NIST Webbook
rinpol	1065.00	NIST Webbook
rinpol	1061.00	NIST Webbook
rinpol	1059.00	NIST Webbook

rinpol	1061.00	NIST Webbook
rinpol	1070.00	NIST Webbook
rinpol	1061.00	NIST Webbook
rinpol	1093.00	NIST Webbook
rinpol	1093.00	NIST Webbook
rinpol	1071.00	NIST Webbook
rinpol	1074.00	NIST Webbook
rinpol	1064.00	NIST Webbook
rinpol	1082.00	NIST Webbook
rinpol	1054.00	NIST Webbook
rinpol	1070.00	NIST Webbook
rinpol	1076.00	NIST Webbook
rinpol	1069.00	NIST Webbook
rinpol	1079.00	NIST Webbook
rinpol	1074.00	NIST Webbook
rinpol	1070.00	NIST Webbook
rinpol	1068.00	NIST Webbook
rinpol	1052.00	NIST Webbook
rinpol	1073.00	NIST Webbook
rinpol	1045.00	NIST Webbook
ripol	1540.00	NIST Webbook
ripol	1559.00	NIST Webbook
ripol	1517.00	NIST Webbook
ripol	1554.00	NIST Webbook
ripol	1528.00	NIST Webbook
ripol	1551.00	NIST Webbook
ripol	1528.00	NIST Webbook
ripol	1557.00	NIST Webbook
ripol	1555.00	NIST Webbook
ripol	1550.00	NIST Webbook
ripol	1556.00	NIST Webbook
ripol	1556.00	NIST Webbook
ripol	1555.68	NIST Webbook
ripol	1573.00	NIST Webbook
ripol	1538.00	NIST Webbook
ripol	1544.00	NIST Webbook
ripol	1559.00	NIST Webbook
ripol	1520.00	NIST Webbook
ripol	1557.00	NIST Webbook
ripol	1546.00	NIST Webbook
ripol	1530.00	NIST Webbook
ripol	1516.00	NIST Webbook
ripol	1568.00	NIST Webbook
ripol	1534.00	NIST Webbook

ripol	1507.00	NIST Webbook
ripol	1540.00	NIST Webbook
ripol	1564.00	NIST Webbook
ripol	1562.00	NIST Webbook
ripol	1543.00	NIST Webbook
ripol	1555.00	NIST Webbook
ripol	1558.00	NIST Webbook
ripol	1567.00	NIST Webbook
ripol	1594.00	NIST Webbook
ripol	1572.00	NIST Webbook
ripol	1556.00	NIST Webbook
ripol	1519.00	NIST Webbook
ripol	1522.00	NIST Webbook
ripol	1561.00	NIST Webbook
ripol	1526.00	NIST Webbook
ripol	1550.00	NIST Webbook
ripol	1553.00	NIST Webbook
ripol	1549.00	NIST Webbook
ripol	1566.00	NIST Webbook
ripol	1565.00	NIST Webbook
ripol	1513.00	NIST Webbook
ripol	1519.00	NIST Webbook
ripol	1557.00	NIST Webbook
ripol	1557.00	NIST Webbook
ripol	1573.00	NIST Webbook
ripol	1573.00	NIST Webbook
ripol	1554.00	NIST Webbook
ripol	1507.00	NIST Webbook
ripol	1553.00	NIST Webbook
ripol	1519.00	NIST Webbook
ripol	1568.00	NIST Webbook
ripol	1548.00	NIST Webbook
ripol	1562.00	NIST Webbook
ripol	1562.00	NIST Webbook
ripol	1562.00	NIST Webbook
ripol	1562.00	NIST Webbook
ripol	1562.00	NIST Webbook
ripol	1562.00	NIST Webbook
ripol	1519.00	NIST Webbook
ripol	1549.00	NIST Webbook
ripol	1562.00	NIST Webbook
ripol	1562.00	NIST Webbook
ripol	1562.00	NIST Webbook
ripol	1562.00	NIST Webbook

ripol	1562.00	NIST Webbook
ripol	1556.00	NIST Webbook
ripol	1519.00	NIST Webbook
ripol	1567.00	NIST Webbook
ripol	1562.00	NIST Webbook
ripol	1562.00	NIST Webbook
ripol	1549.00	NIST Webbook
ripol	1551.00	NIST Webbook
ripol	1562.00	NIST Webbook
ripol	1562.00	NIST Webbook
ripol	1562.00	NIST Webbook
ripol	1562.00	NIST Webbook
ripol	1539.00	NIST Webbook
ripol	1538.00	NIST Webbook
ripol	1536.00	NIST Webbook
ripol	1562.00	NIST Webbook
ripol	1550.00	NIST Webbook
ripol	1562.00	NIST Webbook
ripol	1562.00	NIST Webbook
ripol	1562.00	NIST Webbook
ripol	1563.00	NIST Webbook
ripol	1557.00	NIST Webbook
ripol	1562.00	NIST Webbook
ripol	1543.00	NIST Webbook
ripol	1574.00	NIST Webbook
ripol	1561.00	NIST Webbook
ripol	1605.00	NIST Webbook
ripol	1548.00	NIST Webbook
ripol	1554.38	NIST Webbook
ripol	1559.00	NIST Webbook
ripol	1568.00	NIST Webbook
ripol	1565.00	NIST Webbook
ripol	1566.00	NIST Webbook
ripol	1562.00	NIST Webbook
ripol	1564.00	NIST Webbook
ripol	1561.00	NIST Webbook
ripol	1566.00	NIST Webbook
ripol	1561.00	NIST Webbook
ripol	1568.00	NIST Webbook
ripol	1558.00	NIST Webbook
ripol	1557.00	NIST Webbook
ripol	1570.00	NIST Webbook
ripol	1563.00	NIST Webbook
ripol	1555.00	NIST Webbook

ripol	1561.00	NIST Webbook
ripol	1530.00	NIST Webbook
ripol	1530.00	NIST Webbook
ripol	1585.00	NIST Webbook
ripol	1530.00	NIST Webbook
ripol	1519.00	NIST Webbook
ripol	1516.00	NIST Webbook
ripol	1562.00	NIST Webbook
ripol	1562.00	NIST Webbook
ripol	1519.00	NIST Webbook
ripol	1516.00	NIST Webbook
ripol	1568.00	NIST Webbook
ripol	1553.00	NIST Webbook
ripol	1554.00	NIST Webbook
ripol	1519.00	NIST Webbook
ripol	1534.00	NIST Webbook
ripol	1534.00	NIST Webbook
ripol	1519.00	NIST Webbook
ripol	1575.00	NIST Webbook
ripol	1534.00	NIST Webbook
ripol	1531.00	NIST Webbook
ripol	1507.00	NIST Webbook
ripol	1551.00	NIST Webbook
ripol	1546.00	NIST Webbook
ripol	1534.00	NIST Webbook
ripol	1558.00	NIST Webbook
ripol	1565.00	NIST Webbook
ripol	1564.00	NIST Webbook
ripol	1566.00	NIST Webbook
ripol	1569.00	NIST Webbook
ripol	1560.00	NIST Webbook
ripol	1564.00	NIST Webbook
ripol	1560.00	NIST Webbook
ripol	1552.00	NIST Webbook
ripol	1510.00	NIST Webbook
ripol	1562.00	NIST Webbook
ripol	1562.00	NIST Webbook
ripol	1562.00	NIST Webbook
ripol	1563.00	NIST Webbook
ripol	1519.00	NIST Webbook
ripol	1555.00	NIST Webbook
ripol	1540.00	NIST Webbook
ripol	1565.00	NIST Webbook
ripol	1562.00	NIST Webbook

ripol	1560.00	NIST Webbook
ripol	1555.00	NIST Webbook
ripol	1562.00	NIST Webbook
ripol	1556.00	NIST Webbook
ripol	1556.00	NIST Webbook
ripol	1560.00	NIST Webbook
ripol	1550.00	NIST Webbook
ripol	1566.00	NIST Webbook
ripol	1555.00	NIST Webbook
ripol	1560.00	NIST Webbook
ripol	1554.00	NIST Webbook
ripol	1547.00	NIST Webbook
ripol	1564.00	NIST Webbook
ripol	1541.00	NIST Webbook
ripol	1558.00	NIST Webbook
ripol	1553.35	NIST Webbook
ripol	1562.00	NIST Webbook
ripol	1562.00	NIST Webbook
ripol	1558.00	NIST Webbook
ripol	1530.00	NIST Webbook
ripol	1572.00	NIST Webbook
ripol	1573.00	NIST Webbook
ripol	1562.00	NIST Webbook
ripol	1543.00	NIST Webbook
ripol	1557.00	NIST Webbook
ripol	1556.00	NIST Webbook
ripol	1546.00	NIST Webbook
ripol	1543.00	NIST Webbook
ripol	1573.00	NIST Webbook
ripol	1563.00	NIST Webbook
ripol	1559.00	NIST Webbook
ripol	1567.00	NIST Webbook
ripol	1555.00	NIST Webbook
ripol	1557.00	NIST Webbook
ripol	1528.00	NIST Webbook
ripol	1531.00	NIST Webbook
ripol	1557.00	NIST Webbook
ripol	1558.00	NIST Webbook
ripol	1558.00	NIST Webbook
ripol	1537.00	NIST Webbook
ripol	1566.00	NIST Webbook
ripol	1539.00	NIST Webbook
ripol	1567.00	NIST Webbook
ripol	1551.00	NIST Webbook

ripol	1562.00	NIST Webbook
ripol	1515.00	NIST Webbook
ripol	1528.00	NIST Webbook
ripol	1594.00	NIST Webbook
ripol	1542.00	NIST Webbook
ripol	1560.00	NIST Webbook
ripol	1555.00	NIST Webbook
ripol	1568.00	NIST Webbook
ripol	1572.00	NIST Webbook
ripol	1558.00	NIST Webbook
ripol	1580.00	NIST Webbook
ripol	1548.00	NIST Webbook
ripol	1552.00	NIST Webbook
ripol	1556.00	NIST Webbook
ripol	1552.00	NIST Webbook
ripol	1584.40	NIST Webbook
ripol	1584.40	NIST Webbook
ripol	1550.00	NIST Webbook
ripol	1540.00	NIST Webbook
ripol	1568.00	NIST Webbook
ripol	1563.00	NIST Webbook
ripol	1556.00	NIST Webbook
ripol	1553.00	NIST Webbook
ripol	1560.00	NIST Webbook
ripol	1553.00	NIST Webbook
ripol	1576.00	NIST Webbook
ripol	1574.00	NIST Webbook
ripol	1549.00	NIST Webbook
ripol	1576.00	NIST Webbook
ripol	1547.00	NIST Webbook
ripol	1540.00	NIST Webbook
ripol	1563.00	NIST Webbook
ripol	1566.00	NIST Webbook
ripol	1565.00	NIST Webbook
ripol	1552.00	NIST Webbook
ripol	1566.00	NIST Webbook
ripol	1561.00	NIST Webbook
ripol	1565.00	NIST Webbook
ripol	1533.00	NIST Webbook
ripol	1526.00	NIST Webbook
ripol	1544.00	NIST Webbook
ripol	1565.00	NIST Webbook
ripol	1563.00	NIST Webbook
ripol	1564.00	NIST Webbook

ripol	1560.00	NIST Webbook
ripol	1562.00	NIST Webbook
ripol	1562.00	NIST Webbook
ripol	1550.00	NIST Webbook
ripol	1550.00	NIST Webbook
ripol	1550.00	NIST Webbook
ripol	1538.00	NIST Webbook
ripol	1550.00	NIST Webbook
ripol	1567.00	NIST Webbook
ripol	1530.00	NIST Webbook
ripol	1565.00	NIST Webbook
ripol	1565.00	NIST Webbook
ripol	1554.00	NIST Webbook
ripol	1554.00	NIST Webbook
ripol	1541.00	NIST Webbook
ripol	1562.00	NIST Webbook
ripol	1573.00	NIST Webbook
ripol	1562.00	NIST Webbook
ripol	1579.00	NIST Webbook
ripol	1579.00	NIST Webbook
ripol	1579.00	NIST Webbook
ripol	1555.00	NIST Webbook
ripol	1537.00	NIST Webbook
ripol	1520.00	NIST Webbook
ripol	1551.00	NIST Webbook
ripol	1560.00	NIST Webbook
ripol	1549.00	NIST Webbook
ripol	1559.00	NIST Webbook
ripol	1546.00	NIST Webbook
ripol	1578.00	NIST Webbook
ripol	1566.00	NIST Webbook
ripol	1558.00	NIST Webbook
ripol	1540.00	NIST Webbook
ripol	1562.00	NIST Webbook
ripol	1555.00	NIST Webbook
ripol	1589.00	NIST Webbook
ripol	1538.00	NIST Webbook
ripol	1544.00	NIST Webbook
ripol	1551.00	NIST Webbook
ripol	1558.00	NIST Webbook
ripol	1559.00	NIST Webbook
ripol	1544.00	NIST Webbook
ripol	1584.00	NIST Webbook
ripol	1559.00	NIST Webbook

ripol	1546.00	NIST Webbook
ripol	1546.00	NIST Webbook
ripol	1557.00	NIST Webbook
ripol	1558.00	NIST Webbook
ripol	1566.00	NIST Webbook
ripol	1544.00	NIST Webbook
ripol	1565.00	NIST Webbook
ripol	1566.00	NIST Webbook
ripol	1569.00	NIST Webbook
ripol	1534.00	NIST Webbook
ripol	1545.00	NIST Webbook
ripol	1566.00	NIST Webbook
ripol	1559.00	NIST Webbook
ripol	1555.00	NIST Webbook
ripol	1538.00	NIST Webbook
ripol	1507.00	NIST Webbook
ripol	1553.00	NIST Webbook
ripol	1559.00	NIST Webbook
ripol	1565.00	NIST Webbook
ripol	1550.00	NIST Webbook
ripol	1553.00	NIST Webbook
ripol	1566.00	NIST Webbook
ripol	1520.00	NIST Webbook
ripol	1537.00	NIST Webbook
ripol	1540.00	NIST Webbook
ripol	1552.00	NIST Webbook
ripol	1563.00	NIST Webbook
ripol	1558.00	NIST Webbook
ripol	1546.00	NIST Webbook
ripol	1566.00	NIST Webbook
ripol	1561.00	NIST Webbook
ripol	1561.00	NIST Webbook
ripol	1557.00	NIST Webbook
ripol	1556.00	NIST Webbook
ripol	1573.00	NIST Webbook
ripol	1569.00	NIST Webbook
ripol	1560.00	NIST Webbook
ripol	1546.00	NIST Webbook
ripol	1567.00	NIST Webbook
ripol	1562.00	NIST Webbook
ripol	1567.00	NIST Webbook
ripol	1521.00	NIST Webbook
ripol	1519.00	NIST Webbook
ripol	1530.00	NIST Webbook

ripol	1530.00		NIST Webbook
ripol	1530.00		NIST Webbook
ripol	1519.00		NIST Webbook
ripol	1522.00		NIST Webbook
ripol	1565.00		NIST Webbook
ripol	1523.00		NIST Webbook
ripol	1564.00		NIST Webbook
ripol	1564.00		NIST Webbook
ripol	1561.00		NIST Webbook
ripol	1547.00		NIST Webbook
ripol	1546.00		NIST Webbook
ripol	1519.00		NIST Webbook
ripol	1547.00		NIST Webbook
ripol	1526.00		NIST Webbook
ripol	1552.00		NIST Webbook
ripol	1555.00		NIST Webbook
ripol	1522.00		NIST Webbook
ripol	1518.00		NIST Webbook
ripol	1540.00		NIST Webbook
ripol	1523.00		NIST Webbook
ripol	1547.00		NIST Webbook
ripol	1560.00		NIST Webbook
ripol	1530.00		NIST Webbook
ripol	1539.00		NIST Webbook
ripol	1556.05		NIST Webbook
ripol	1554.00		NIST Webbook
tb	466.00	K	The role of organic diluents in the aspects of equilibrium, kinetics and thermodynamic model for silver ion extraction using an extractant D2EHPA
tb	468.00	K	(Liquid + liquid) equilibria of (water + lactic acid + alcohol) ternary systems
tb	468.20	K	Ternary phase diagrams for aqueous mixtures of butyric acid with several solvents: experimental and correlated data
tb	468.31	K	KDB
tb	468.00	K	(Liquid + liquid) equilibria of (water + propionic acid + alcohol) ternary systems
tc	652.60	K	NIST Webbook
tc	650.56 ± 0.30	K	NIST Webbook
tc	652.40 ± 0.60	K	NIST Webbook
tc	668.00 ± 2.00	K	NIST Webbook
tc	657.80	K	NIST Webbook

tc	658.60	K	NIST Webbook
tc	652.40 ± 0.60	K	NIST Webbook
tc	652.20 ± 0.80	K	NIST Webbook
tc	652.50 ± 0.50	K	NIST Webbook
tc	652.50	K	NIST Webbook
tc	652.50	K	KDB
tf	256.40 ± 0.50	K	NIST Webbook
tf	256.25 ± 0.40	K	NIST Webbook
tf	257.60	K	KDB
tf	256.20 ± 0.50	K	NIST Webbook
tf	258.03 ± 0.08	K	NIST Webbook
tf	258.35 ± 0.05	K	NIST Webbook
tf	257.65	K	Aqueous Solubility Prediction Method
tf	257.70	K	SLE and LLE for tri-butylphosphate or tri-octylamine contained systems; extractive solvents of Molybdenum
tf	257.85	K	Solid-Liquid Equilibria, Excess Molar Volumes, and Molar Refractivity Deviations for Extractive Solvents of Molybdenum
tf	259.15	K	Phase Equilibria Study of the Binary Systems (N-Butyl-4-methylpyridinium Tosylate Ionic Liquid + Organic Solvent, or Water)
tf	258.03	K	Experimental (Solid + Liquid) and (Liquid + Liquid) Equilibria and Excess Molar Volume of Alkanol + Acetonitrile, Propanenitrile, and Butanenitrile Mixtures
tf	258.20	K	Phase equilibria of didecyldimethylammonium nitrate ionic liquid with water and organic solvents
tf	256.90 ± 0.40	K	NIST Webbook
vc	0.497	m3/kmol	NIST Webbook
vc	0.497	m3/kmol	KDB
zc	0.2543990		KDB

Temperature Dependent Properties

Property code	Value	Unit	Temperature [K]	Source

cpg	343.66	J/mol×K	607.75	Joback Method
cpg	323.45	J/mol×K	554.50	Joback Method
cpg	353.18	J/mol×K	634.38	Joback Method
cpg	290.08	J/mol×K	474.62	Joback Method
cpg	301.62	J/mol×K	501.25	Joback Method
cpg	312.74	J/mol×K	527.87	Joback Method
cpg	333.75	J/mol×K	581.13	Joback Method
cpl	284.50	J/mol×K	286.00	NIST Webbook
cpl	312.10	J/mol×K	298.00	NIST Webbook
cpl	318.30	J/mol×K	310.67	NIST Webbook
cpl	324.30	J/mol×K	298.00	NIST Webbook
cpl	304.00	J/mol×K	298.15	Volumetric Properties, Viscosities, and Isobaric Heat Capacities of Imidazolium Octanoate Protic Ionic Liquid in Molecular Solvents
cpl	329.50	J/mol×K	303.20	NIST Webbook
cpl	304.00	J/mol×K	298.15	NIST Webbook
cpl	379.00	J/mol×K	313.15	Volumetric Properties, Viscosities, and Isobaric Heat Capacities of Imidazolium Octanoate Protic Ionic Liquid in Molecular Solvents
cpl	354.00	J/mol×K	308.15	Volumetric Properties, Viscosities, and Isobaric Heat Capacities of Imidazolium Octanoate Protic Ionic Liquid in Molecular Solvents
cpl	329.00	J/mol×K	303.15	Volumetric Properties, Viscosities, and Isobaric Heat Capacities of Imidazolium Octanoate Protic Ionic Liquid in Molecular Solvents
cpl	305.55	J/mol×K	298.15	NIST Webbook

dvisc	0.0038650	Pa×s	318.15	Density, Viscosity, and Speed of Sound of (1-Octanol + 2-Methoxyethanol), (1-Octanol + N,N-Dimethylacetamide), and (1-Octanol + Acetophenone) at Temperatures of (298.15, 308.15, and 318.15) K
dvisc	0.0076610	Pa×s	298.15	Acoustical and Excess Properties of {Chlorobenzene + 1-Hexanol, or 1-Heptanol, or 1-Octanol, or 1-Nonanol, or 1-Decanol} at (298.15, 303.15, 308.15, and 313.15) K
dvisc	0.0064020	Pa×s	303.15	Acoustical and Excess Properties of {Chlorobenzene + 1-Hexanol, or 1-Heptanol, or 1-Octanol, or 1-Nonanol, or 1-Decanol} at (298.15, 303.15, 308.15, and 313.15) K
dvisc	0.0054250	Pa×s	308.15	Acoustical and Excess Properties of {Chlorobenzene + 1-Hexanol, or 1-Heptanol, or 1-Octanol, or 1-Nonanol, or 1-Decanol} at (298.15, 303.15, 308.15, and 313.15) K
dvisc	0.0046280	Pa×s	313.15	Acoustical and Excess Properties of {Chlorobenzene + 1-Hexanol, or 1-Heptanol, or 1-Octanol, or 1-Nonanol, or 1-Decanol} at (298.15, 303.15, 308.15, and 313.15) K

dvisc	0.0073650	Pa×s	298.15	Density, Viscosity, and Speed of Sound of (1-Octanol + 2-Methoxyethanol), (1-Octanol + N,N-Dimethylacetamide), and (1-Octanol + Acetophenone) at Temperatures of (298.15, 308.15, and 318.15) K
dvisc	0.0052510	Pa×s	308.15	Density, Viscosity, and Speed of Sound of (1-Octanol + 2-Methoxyethanol), (1-Octanol + N,N-Dimethylacetamide), and (1-Octanol + Acetophenone) at Temperatures of (298.15, 308.15, and 318.15) K
dvisc	0.0046030	Pa×s	313.15	Densities, Viscosities, and Speed of Sound Studies of Binary Mixtures of Methylbenzene with Heptan-1-ol, Octan-1-ol, and Decan-1-ol at (303.15 and 313.15) K
dvisc	0.0110700	Pa×s	288.15	Densities and Viscosities of Four Binary Diethyl Carbonate + 1-Alcohol Systems from (288.15 to 313.15) K
dvisc	0.0108720	Pa×s	288.15	Measurement of Density and Viscosity of Binary 1-Alkanol Systems (C8-C11) at 101 kPa and Temperatures from (283.15 to 313.15) K

dvisc	0.0091610	Pa×s	293.15	Densities and Viscosities of Four Binary Diethyl Carbonate + 1-Alcohol Systems from (288.15 to 313.15) K
dvisc	0.0076880	Pa×s	298.15	Densities and Viscosities of Four Binary Diethyl Carbonate + 1-Alcohol Systems from (288.15 to 313.15) K
dvisc	0.0064040	Pa×s	303.15	Densities and Viscosities of Four Binary Diethyl Carbonate + 1-Alcohol Systems from (288.15 to 313.15) K
dvisc	0.0046200	Pa×s	313.15	Densities and Viscosities of Four Binary Diethyl Carbonate + 1-Alcohol Systems from (288.15 to 313.15) K
dvisc	0.0134930	Pa×s	283.15	Density and Viscosity of the Binary Systems Ethanol + Butan-1-ol, + Pentan-1-ol, + Heptan-1-ol, + Octan-1-ol, Nonan-1-ol, + Decan-1-ol at 0.1 MPa and Temperatures from 283.15 K to 313.15 K
dvisc	0.0064330	Pa×s	303.15	Densities, Viscosities, and Speed of Sound Studies of Binary Mixtures of Methylbenzene with Heptan-1-ol, Octan-1-ol, and Decan-1-ol at (303.15 and 313.15) K

dvisc	0.0090160	Pa×s	293.15	Density and Viscosity of the Binary Systems Ethanol + Butan-1-ol, + Pentan-1-ol, + Heptan-1-ol, + Octan-1-ol, Nonan-1-ol, + Decan-1-ol at 0.1 MPa and Temperatures from 283.15 K to 313.15 K	
dvisc	0.0073900	Pa×s	298.15	Density and Viscosity of the Binary Systems Ethanol + Butan-1-ol, + Pentan-1-ol, + Heptan-1-ol, + Octan-1-ol, Nonan-1-ol, + Decan-1-ol at 0.1 MPa and Temperatures from 283.15 K to 313.15 K	
dvisc	0.0063060	Pa×s	303.15	Density and Viscosity of the Binary Systems Ethanol + Butan-1-ol, + Pentan-1-ol, + Heptan-1-ol, + Octan-1-ol, Nonan-1-ol, + Decan-1-ol at 0.1 MPa and Temperatures from 283.15 K to 313.15 K	
dvisc	0.0045870	Pa×s	313.15	Density and Viscosity of the Binary Systems Ethanol + Butan-1-ol, + Pentan-1-ol, + Heptan-1-ol, + Octan-1-ol, Nonan-1-ol, + Decan-1-ol at 0.1 MPa and Temperatures from 283.15 K to 313.15 K	

dvisc	0.0073980	Pa×s	298.15	Density and Viscosity of Binary Mixtures of {1-Butyl-3-methylimidazolium Thiocyanate + 1-Heptanol, 1-Octanol, 1-Nonanol, or 1-Decanol}
dvisc	0.0052850	Pa×s	308.15	Density and Viscosity of Binary Mixtures of {1-Butyl-3-methylimidazolium Thiocyanate + 1-Heptanol, 1-Octanol, 1-Nonanol, or 1-Decanol}
dvisc	0.0038870	Pa×s	318.15	Density and Viscosity of Binary Mixtures of {1-Butyl-3-methylimidazolium Thiocyanate + 1-Heptanol, 1-Octanol, 1-Nonanol, or 1-Decanol}
dvisc	0.0046040	Pa×s	313.15	Densities, Viscosities, and Ultrasonic Velocity Studies of Binary Mixtures of Chloroform with Octan-1-ol and Decan-1-ol at (303.15 and 313.15) K
dvisc	0.0022620	Pa×s	338.15	Density and Viscosity of Binary Mixtures of {1-Butyl-3-methylimidazolium Thiocyanate + 1-Heptanol, 1-Octanol, 1-Nonanol, or 1-Decanol}
dvisc	0.0064460	Pa×s	303.15	Densities, Viscosities, and Ultrasonic Velocity Studies of Binary Mixtures of Chloroform with Octan-1-ol and Decan-1-ol at (303.15 and 313.15) K

dvisc	0.0134930	Pa×s	283.15	Measurement of Density and Viscosity of Binary 1-Alkanol Systems (C8-C11) at 101 kPa and Temperatures from (283.15 to 313.15) K	
dvisc	0.0045400	Pa×s	313.15	Measurement of Density and Viscosity of Binary 1-Alkanol Systems (C8-C11) at 101 kPa and Temperatures from (283.15 to 313.15) K	
dvisc	0.0053420	Pa×s	308.15	Measurement of Density and Viscosity of Binary 1-Alkanol Systems (C8-C11) at 101 kPa and Temperatures from (283.15 to 313.15) K	
dvisc	0.0062730	Pa×s	303.15	Measurement of Density and Viscosity of Binary 1-Alkanol Systems (C8-C11) at 101 kPa and Temperatures from (283.15 to 313.15) K	
dvisc	0.0074980	Pa×s	298.15	Measurement of Density and Viscosity of Binary 1-Alkanol Systems (C8-C11) at 101 kPa and Temperatures from (283.15 to 313.15) K	
dvisc	0.0017820	Pa×s	348.15	Density and Viscosity of Binary Mixtures of {1-Butyl-3-methylimidazolium Thiocyanate + 1-Heptanol, 1-Octanol, 1-Nonanol, or 1-Decanol}	

dvisc	0.0089440	Pa×s	293.15	Measurement of Density and Viscosity of Binary 1-Alkanol Systems (C8-C11) at 101 kPa and Temperatures from (283.15 to 313.15) K
dvisc	0.0029310	Pa×s	328.15	Density and Viscosity of Binary Mixtures of {1-Butyl-3-methylimidazolium Thiocyanate + 1-Heptanol, 1-Octanol, 1-Nonanol, or 1-Decanol}
hfust	25.24	kJ/mol	258.40	NIST Webbook
hvapt	61.60	kJ/mol	396.00	NIST Webbook
hvapt	69.60	kJ/mol	301.50	NIST Webbook
hvapt	68.70	kJ/mol	318.00	NIST Webbook
hvapt	67.30	kJ/mol	364.00	NIST Webbook
hvapt	50.63	kJ/mol	468.30	KDB
hvapt	56.60	kJ/mol	438.00	NIST Webbook
hvapt	47.80	kJ/mol	515.00	NIST Webbook
hvapt	64.00	kJ/mol	274.50	NIST Webbook
hvapt	67.50	kJ/mol	405.50	NIST Webbook
hvapt	65.00	kJ/mol	410.00	NIST Webbook
hvapt	70.40	kJ/mol	323.00	NIST Webbook
hvapt	64.00 ± 1.00	kJ/mol	267.00	NIST Webbook
hvapt	52.50	kJ/mol	452.00	NIST Webbook
pvap	9.99	kPa	400.95	Vapor Pressure and Isobaric Vapor-Liquid Equilibrium for Binary Systems of Furfural, 2-Acetylfuran, and 5-Methylfurfural at 3.60 and 5.18 kPa
pvap	40.53	kPa	438.31	Vapor pressure of 1-octanol below 5 kPa using DSC
pvap	0.22	kPa	333.69	Vapor pressure of 1-octanol below 5 kPa using DSC
pvap	0.39	kPa	342.10	Vapor pressure of 1-octanol below 5 kPa using DSC

pvap	0.60	kPa	349.10	Vapor pressure of 1-octanol below 5 kPa using DSC
pvap	0.81	kPa	353.96	Vapor pressure of 1-octanol below 5 kPa using DSC
pvap	1.07	kPa	357.90	Vapor pressure of 1-octanol below 5 kPa using DSC
pvap	1.45	kPa	363.30	Vapor pressure of 1-octanol below 5 kPa using DSC
pvap	2.13	kPa	369.90	Vapor pressure of 1-octanol below 5 kPa using DSC
pvap	3.49	kPa	379.49	Vapor pressure of 1-octanol below 5 kPa using DSC
pvap	5.17	kPa	387.20	Vapor pressure of 1-octanol below 5 kPa using DSC
pvap	13.58	kPa	408.44	Vapor pressure of 1-octanol below 5 kPa using DSC
pvap	40.53	kPa	438.58	Vapor pressure of 1-octanol below 5 kPa using DSC
pvap	0.24	kPa	334.46	Vapor pressure of 1-octanol below 5 kPa using DSC
pvap	0.39	kPa	341.76	Vapor pressure of 1-octanol below 5 kPa using DSC
pvap	0.57	kPa	348.20	Vapor pressure of 1-octanol below 5 kPa using DSC
pvap	0.75	kPa	352.09	Vapor pressure of 1-octanol below 5 kPa using DSC
pvap	1.06	kPa	357.94	Vapor pressure of 1-octanol below 5 kPa using DSC
pvap	1.46	kPa	363.51	Vapor pressure of 1-octanol below 5 kPa using DSC

pvap	2.09	kPa	369.76	Vapor pressure of 1-octanol below 5 kPa using DSC
pvap	3.44	kPa	379.04	Vapor pressure of 1-octanol below 5 kPa using DSC
pvap	5.16	kPa	387.28	Vapor pressure of 1-octanol below 5 kPa using DSC
pvap	13.55	kPa	408.90	Vapor pressure of 1-octanol below 5 kPa using DSC
pvap	40.47	kPa	435.48	Vapor pressure of 1-octanol below 5 kPa using DSC
pvap	0.06	kPa	319.82	Thermogravimetric measurement of evaporation: Data analysis based on the Stefan tube
pvap	0.40	kPa	343.39	Thermogravimetric measurement of evaporation: Data analysis based on the Stefan tube
pvap	0.21	kPa	333.93	Thermogravimetric measurement of evaporation: Data analysis based on the Stefan tube
pvap	0.09	kPa	324.46	Thermogravimetric measurement of evaporation: Data analysis based on the Stefan tube
pvap	1.36	kPa	362.32	Thermogravimetric measurement of evaporation: Data analysis based on the Stefan tube

pvap	6.14	kPa	390.95	Thermogravimetric measurement of evaporation: Data analysis based on the Stefan tube
pvap	0.54	kPa	348.15	Thermogravimetric measurement of evaporation: Data analysis based on the Stefan tube
pvap	0.38	kPa	343.23	Thermogravimetric measurement of evaporation: Data analysis based on the Stefan tube
pvap	0.85	kPa	354.85	Vapor Pressure and Isobaric Vapor-Liquid Equilibrium for Binary Systems of Furfural, 2-Acetylfuran, and 5-Methylfurfural at 3.60 and 5.18 kPa
pvap	1.07	kPa	358.15	Vapor Pressure and Isobaric Vapor-Liquid Equilibrium for Binary Systems of Furfural, 2-Acetylfuran, and 5-Methylfurfural at 3.60 and 5.18 kPa
pvap	1.51	kPa	363.55	Vapor Pressure and Isobaric Vapor-Liquid Equilibrium for Binary Systems of Furfural, 2-Acetylfuran, and 5-Methylfurfural at 3.60 and 5.18 kPa

pvap	2.01	kPa	368.25	Vapor Pressure and Isobaric Vapor-Liquid Equilibrium for Binary Systems of Furfural, 2-Acetylfuran, and 5-Methylfurfural at 3.60 and 5.18 kPa	
pvap	2.99	kPa	376.05	Vapor Pressure and Isobaric Vapor-Liquid Equilibrium for Binary Systems of Furfural, 2-Acetylfuran, and 5-Methylfurfural at 3.60 and 5.18 kPa	
pvap	4.01	kPa	381.75	Vapor Pressure and Isobaric Vapor-Liquid Equilibrium for Binary Systems of Furfural, 2-Acetylfuran, and 5-Methylfurfural at 3.60 and 5.18 kPa	
pvap	5.01	kPa	386.05	Vapor Pressure and Isobaric Vapor-Liquid Equilibrium for Binary Systems of Furfural, 2-Acetylfuran, and 5-Methylfurfural at 3.60 and 5.18 kPa	
pvap	6.01	kPa	389.95	Vapor Pressure and Isobaric Vapor-Liquid Equilibrium for Binary Systems of Furfural, 2-Acetylfuran, and 5-Methylfurfural at 3.60 and 5.18 kPa	

pvap	7.01	kPa	393.15	Vapor Pressure and Isobaric Vapor-Liquid Equilibrium for Binary Systems of Furfural, 2-Acetylfuran, and 5-Methylfurfural at 3.60 and 5.18 kPa
pvap	8.02	kPa	396.05	Vapor Pressure and Isobaric Vapor-Liquid Equilibrium for Binary Systems of Furfural, 2-Acetylfuran, and 5-Methylfurfural at 3.60 and 5.18 kPa
pvap	9.02	kPa	398.55	Vapor Pressure and Isobaric Vapor-Liquid Equilibrium for Binary Systems of Furfural, 2-Acetylfuran, and 5-Methylfurfural at 3.60 and 5.18 kPa
pvap	13.57	kPa	407.24	Vapor pressure of 1-octanol below 5 kPa using DSC
pvap	14.97	kPa	410.15	Vapor Pressure and Isobaric Vapor-Liquid Equilibrium for Binary Systems of Furfural, 2-Acetylfuran, and 5-Methylfurfural at 3.60 and 5.18 kPa
pvap	19.91	kPa	417.35	Vapor Pressure and Isobaric Vapor-Liquid Equilibrium for Binary Systems of Furfural, 2-Acetylfuran, and 5-Methylfurfural at 3.60 and 5.18 kPa

pvap	97.64	kPa	466.80	Low-Pressure VLE Measurements and Thermodynamic Modeling, with PSRK and NRTL, of Binary 1-Alcohol + n-Alkane Systems
pvap	89.16	kPa	463.50	Low-Pressure VLE Measurements and Thermodynamic Modeling, with PSRK and NRTL, of Binary 1-Alcohol + n-Alkane Systems
pvap	81.52	kPa	460.31	Low-Pressure VLE Measurements and Thermodynamic Modeling, with PSRK and NRTL, of Binary 1-Alcohol + n-Alkane Systems
pvap	74.25	kPa	457.08	Low-Pressure VLE Measurements and Thermodynamic Modeling, with PSRK and NRTL, of Binary 1-Alcohol + n-Alkane Systems
pvap	67.60	kPa	454.35	Low-Pressure VLE Measurements and Thermodynamic Modeling, with PSRK and NRTL, of Binary 1-Alcohol + n-Alkane Systems

pvap	60.03	kPa	450.45	Low-Pressure VLE Measurements and Thermodynamic Modeling, with PSRK and NRTL, of Binary 1-Alcohol + n-Alkane Systems	
pvap	52.15	kPa	445.86	Low-Pressure VLE Measurements and Thermodynamic Modeling, with PSRK and NRTL, of Binary 1-Alcohol + n-Alkane Systems	
pvap	48.62	kPa	443.61	Low-Pressure VLE Measurements and Thermodynamic Modeling, with PSRK and NRTL, of Binary 1-Alcohol + n-Alkane Systems	
pvap	40.08	kPa	437.67	Low-Pressure VLE Measurements and Thermodynamic Modeling, with PSRK and NRTL, of Binary 1-Alcohol + n-Alkane Systems	
pvap	40.00	kPa	436.67	Low-Pressure VLE Measurements and Thermodynamic Modeling, with PSRK and NRTL, of Binary 1-Alcohol + n-Alkane Systems	

pvap	1.00e-03	kPa	288.15	Vapor Liquid Equilibrium between (278.15 and 323.15) K and Excess Functions at T) 298.15 K for 1-Bromobutane with 1-Octanol or 1-Decanol
pvap	3.00e-03	kPa	293.15	Vapor Liquid Equilibrium between (278.15 and 323.15) K and Excess Functions at T) 298.15 K for 1-Bromobutane with 1-Octanol or 1-Decanol
pvap	7.00e-03	kPa	298.15	Vapor Liquid Equilibrium between (278.15 and 323.15) K and Excess Functions at T) 298.15 K for 1-Bromobutane with 1-Octanol or 1-Decanol
pvap	0.01	kPa	303.15	Vapor Liquid Equilibrium between (278.15 and 323.15) K and Excess Functions at T) 298.15 K for 1-Bromobutane with 1-Octanol or 1-Decanol
pvap	0.02	kPa	308.15	Vapor Liquid Equilibrium between (278.15 and 323.15) K and Excess Functions at T) 298.15 K for 1-Bromobutane with 1-Octanol or 1-Decanol
pvap	0.03	kPa	313.15	Vapor Liquid Equilibrium between (278.15 and 323.15) K and Excess Functions at T) 298.15 K for 1-Bromobutane with 1-Octanol or 1-Decanol

pvap	0.05	kPa	318.15	Vapor Liquid Equilibrium between (278.15 and 323.15) K and Excess Functions at T) 298.15 K for 1-Bromobutane with 1-Octanol or 1-Decanol
pvap	0.08	kPa	323.15	Vapor Liquid Equilibrium between (278.15 and 323.15) K and Excess Functions at T) 298.15 K for 1-Bromobutane with 1-Octanol or 1-Decanol
pvap	0.01	kPa	298.15	Vapor Pressure Determination of the Aliphatic C5 to C8 1-Alcohols
pvap	0.02	kPa	303.15	Vapor Pressure Determination of the Aliphatic C5 to C8 1-Alcohols
pvap	0.03	kPa	308.15	Vapor Pressure Determination of the Aliphatic C5 to C8 1-Alcohols
pvap	0.04	kPa	313.15	Vapor Pressure Determination of the Aliphatic C5 to C8 1-Alcohols
pvap	0.06	kPa	318.15	Vapor Pressure Determination of the Aliphatic C5 to C8 1-Alcohols
pvap	0.09	kPa	323.15	Vapor Pressure Determination of the Aliphatic C5 to C8 1-Alcohols
pvap	0.14	kPa	328.15	Vapor Pressure Determination of the Aliphatic C5 to C8 1-Alcohols
pvap	0.20	kPa	333.15	Vapor Pressure Determination of the Aliphatic C5 to C8 1-Alcohols
pvap	0.29	kPa	338.15	Vapor Pressure Determination of the Aliphatic C5 to C8 1-Alcohols
pvap	0.41	kPa	343.15	Vapor Pressure Determination of the Aliphatic C5 to C8 1-Alcohols

pvap	0.57	kPa	348.15	Vapor Pressure Determination of the Aliphatic C5 to C8 1-Alcohols
pvap	0.79	kPa	353.15	Vapor Pressure Determination of the Aliphatic C5 to C8 1-Alcohols
pvap	1.07	kPa	358.15	Vapor Pressure Determination of the Aliphatic C5 to C8 1-Alcohols
pvap	1.45	kPa	363.15	Vapor Pressure Determination of the Aliphatic C5 to C8 1-Alcohols
pvap	1.93	kPa	368.15	Vapor Pressure Determination of the Aliphatic C5 to C8 1-Alcohols
pvap	2.54	kPa	373.15	Vapor Pressure Determination of the Aliphatic C5 to C8 1-Alcohols
pvap	3.32	kPa	378.15	Vapor Pressure Determination of the Aliphatic C5 to C8 1-Alcohols
pvap	4.29	kPa	383.15	Vapor Pressure Determination of the Aliphatic C5 to C8 1-Alcohols
pvap	5.48	kPa	388.15	Vapor Pressure Determination of the Aliphatic C5 to C8 1-Alcohols
pvap	6.94	kPa	393.15	Vapor Pressure Determination of the Aliphatic C5 to C8 1-Alcohols
pvap	8.71	kPa	398.15	Vapor Pressure Determination of the Aliphatic C5 to C8 1-Alcohols
pvap	10.83	kPa	403.15	Vapor Pressure Determination of the Aliphatic C5 to C8 1-Alcohols
pvap	13.36	kPa	408.15	Vapor Pressure Determination of the Aliphatic C5 to C8 1-Alcohols
pvap	16.36	kPa	413.15	Vapor Pressure Determination of the Aliphatic C5 to C8 1-Alcohols
pvap	19.88	kPa	418.15	Vapor Pressure Determination of the Aliphatic C5 to C8 1-Alcohols

pvap	23.99	kPa	423.15	Vapor Pressure Determination of the Aliphatic C5 to C8 1-Alcohols
pvap	28.76	kPa	428.15	Vapor Pressure Determination of the Aliphatic C5 to C8 1-Alcohols
pvap	34.27	kPa	433.15	Vapor Pressure Determination of the Aliphatic C5 to C8 1-Alcohols
pvap	40.61	kPa	438.15	Vapor Pressure Determination of the Aliphatic C5 to C8 1-Alcohols
pvap	47.88	kPa	443.15	Vapor Pressure Determination of the Aliphatic C5 to C8 1-Alcohols
pvap	56.17	kPa	448.15	Vapor Pressure Determination of the Aliphatic C5 to C8 1-Alcohols
pvap	65.45	kPa	453.15	Vapor Pressure Determination of the Aliphatic C5 to C8 1-Alcohols
pvap	75.95	kPa	458.15	Vapor Pressure Determination of the Aliphatic C5 to C8 1-Alcohols
pvap	87.73	kPa	463.15	Vapor Pressure Determination of the Aliphatic C5 to C8 1-Alcohols
pvap	5.17	kPa	387.92	Vapor pressure of 1-octanol below 5 kPa using DSC
pvap	3.49	kPa	379.45	Vapor pressure of 1-octanol below 5 kPa using DSC
pvap	2.10	kPa	369.69	Vapor pressure of 1-octanol below 5 kPa using DSC
pvap	1.43	kPa	362.88	Vapor pressure of 1-octanol below 5 kPa using DSC
pvap	1.06	kPa	357.99	Vapor pressure of 1-octanol below 5 kPa using DSC

pvap	0.80	kPa	352.98	Vapor pressure of 1-octanol below 5 kPa using DSC
pvap	0.59	kPa	348.72	Vapor pressure of 1-octanol below 5 kPa using DSC
pvap	0.37	kPa	343.33	Vapor pressure of 1-octanol below 5 kPa using DSC
pvap	0.22	kPa	333.97	Vapor pressure of 1-octanol below 5 kPa using DSC
pvap	101.40	kPa	468.47	Vapor pressure of 1-octanol below 5 kPa using DSC
pvap	101.40	kPa	468.15	Vapor pressure of 1-octanol below 5 kPa using DSC
pvap	40.53	kPa	438.80	Vapor pressure of 1-octanol below 5 kPa using DSC
pvap	13.55	kPa	409.30	Vapor pressure of 1-octanol below 5 kPa using DSC
pvap	5.18	kPa	387.87	Vapor pressure of 1-octanol below 5 kPa using DSC
pvap	3.49	kPa	379.35	Vapor pressure of 1-octanol below 5 kPa using DSC
pvap	2.13	kPa	370.00	Vapor pressure of 1-octanol below 5 kPa using DSC
pvap	1.47	kPa	364.35	Vapor pressure of 1-octanol below 5 kPa using DSC
pvap	0.79	kPa	354.96	Vapor pressure of 1-octanol below 5 kPa using DSC
pvap	0.59	kPa	349.38	Vapor pressure of 1-octanol below 5 kPa using DSC
pvap	0.39	kPa	344.50	Vapor pressure of 1-octanol below 5 kPa using DSC

pvap	0.20	kPa	338.55	Vapor pressure of 1-octanol below 5 kPa using DSC	
pvap	0.08	kPa	323.15	Isothermal (vapour + liquid) equilibrium at several temperatures of (1-chlorobutane + 1-octanol, or 1-decanol)	
pvap	0.05	kPa	318.15	Isothermal (vapour + liquid) equilibrium at several temperatures of (1-chlorobutane + 1-octanol, or 1-decanol)	
pvap	0.03	kPa	313.15	Isothermal (vapour + liquid) equilibrium at several temperatures of (1-chlorobutane + 1-octanol, or 1-decanol)	
pvap	0.02	kPa	308.15	Isothermal (vapour + liquid) equilibrium at several temperatures of (1-chlorobutane + 1-octanol, or 1-decanol)	
pvap	0.01	kPa	303.15	Isothermal (vapour + liquid) equilibrium at several temperatures of (1-chlorobutane + 1-octanol, or 1-decanol)	
pvap	7.00e-03	kPa	298.15	Isothermal (vapour + liquid) equilibrium at several temperatures of (1-chlorobutane + 1-octanol, or 1-decanol)	
pvap	3.00e-03	kPa	293.15	Isothermal (vapour + liquid) equilibrium at several temperatures of (1-chlorobutane + 1-octanol, or 1-decanol)	

pvap	100.30	kPa	466.61	Isobaric vapor-liquid equilibrium for binary systems containing benzothiophene
pvap	91.10	kPa	463.21	Isobaric vapor-liquid equilibrium for binary systems containing benzothiophene
pvap	81.00	kPa	459.23	Isobaric vapor-liquid equilibrium for binary systems containing benzothiophene
pvap	71.00	kPa	454.52	Isobaric vapor-liquid equilibrium for binary systems containing benzothiophene
pvap	60.70	kPa	449.55	Isobaric vapor-liquid equilibrium for binary systems containing benzothiophene
pvap	50.50	kPa	443.90	Isobaric vapor-liquid equilibrium for binary systems containing benzothiophene
pvap	41.50	kPa	437.35	Isobaric vapor-liquid equilibrium for binary systems containing benzothiophene
pvap	30.40	kPa	428.60	Isobaric vapor-liquid equilibrium for binary systems containing benzothiophene
pvap	20.50	kPa	417.83	Isobaric vapor-liquid equilibrium for binary systems containing benzothiophene
pvap	93.42	kPa	465.25	Vapor pressure of selected aliphatic alcohols by ebulliometry. Part 1

pvap	74.88	kPa	457.52	Vapor pressure of selected aliphatic alcohols by ebulliometry. Part 1
pvap	60.37	kPa	450.34	Vapor pressure of selected aliphatic alcohols by ebulliometry. Part 1
pvap	47.73	kPa	442.88	Vapor pressure of selected aliphatic alcohols by ebulliometry. Part 1
pvap	40.83	kPa	438.14	Vapor pressure of selected aliphatic alcohols by ebulliometry. Part 1
pvap	40.83	kPa	438.12	Vapor pressure of selected aliphatic alcohols by ebulliometry. Part 1
pvap	34.89	kPa	433.49	Vapor pressure of selected aliphatic alcohols by ebulliometry. Part 1
pvap	29.45	kPa	428.65	Vapor pressure of selected aliphatic alcohols by ebulliometry. Part 1
pvap	24.62	kPa	423.70	Vapor pressure of selected aliphatic alcohols by ebulliometry. Part 1
pvap	20.33	kPa	418.63	Vapor pressure of selected aliphatic alcohols by ebulliometry. Part 1
pvap	20.33	kPa	418.62	Vapor pressure of selected aliphatic alcohols by ebulliometry. Part 1
pvap	20.33	kPa	418.61	Vapor pressure of selected aliphatic alcohols by ebulliometry. Part 1
pvap	16.83	kPa	413.79	Vapor pressure of selected aliphatic alcohols by ebulliometry. Part 1

pvap	13.95	kPa	409.14	Vapor pressure of selected aliphatic alcohols by ebulliometry. Part 1
pvap	11.10	kPa	403.71	Vapor pressure of selected aliphatic alcohols by ebulliometry. Part 1
pvap	8.36	kPa	397.32	Vapor pressure of selected aliphatic alcohols by ebulliometry. Part 1
pvap	93.54	kPa	465.29	Vapor pressure of selected aliphatic alcohols by ebulliometry. Part 1
pvap	1.07	kPa	358.41	Vapor pressure of 1-octanol below 5 kPa using DSC
rfi	1.42900		293.15	Effect of temperature and chain length on the viscosity and surface tension of binary systems of N,N-dimethylformamide with 1-octanol, 1-nonanol and 1-decanol
rfi	1.42980		293.15	Densities, Excess Molar Volumes, and Isobaric Thermal Expansibilities for 1,2-Ethanediol + 1-Butanol, or 1-Hexanol, or 1-Octanol in the Temperature Range from (293.15 to 313.15) K
rfi	1.42695		298.15	Liquid Liquid Phase Equilibria of 1-Propanol + Water + n-Alcohol Ternary Systems at 298.15 K and Atmospheric Pressure
rfi	1.42750		298.15	Effect of Pressure on the Static Relative Permittivities of Alkan-1-ols at 298.15 K

rfi	1.42740	298.15	Density, Surface Tension, and Refractive Index of Octane + 1-Alkanol Mixtures at T) 298.15 K.	
rfi	1.42910	293.15	Experimental solubility for betulin and estrone in various solvents within the temperature range T = (293.2 to 328.2) K	
rfi	1.42700	298.20	Tie-line data for the aqueous solutions of phenol with organic solvents at T = 298.2 K	
rfi	1.41850	318.20	A thermodynamic study of solute solvent interactions through dielectric properties of the mixtures consisting of 1,4-butanediol, 1-octanol, and 1,4-dioxane at different temperatures	
rfi	1.41800	318.20	A thermodynamic study of solute solvent interactions through dielectric properties of the mixtures consisting of 1,4-butanediol, 1-octanol, and 1,4-dioxane at different temperatures	
rfi	1.42280	308.20	A thermodynamic study of solute solvent interactions through dielectric properties of the mixtures consisting of 1,4-butanediol, 1-octanol, and 1,4-dioxane at different temperatures	

rfi	1.42760	298.20	A thermodynamic study of solute solvent interactions through dielectric properties of the mixtures consisting of 1,4-butanediol, 1-octanol, and 1,4-dioxane at different temperatures
rfi	1.42700	298.20	Ternary equilibrium data of mixtures consisting of 2-butanol, water, and heavy alcohols at T = 298.2 K
rfi	1.42930	293.15	Experimental study and ERAS modeling of the excess molar enthalpy of (acetonitrile + 1-heptanol or 1-octanol) mixtures at (298.15, 313.15, and 323.15) K and atmospheric pressure
rfi	1.42562	303.15	Densities, excess molar volumes, and refractive indices of 1,1,2,2-tetrabromoethane and 1-alkanols binary mixtures
rfi	1.42958	293.15	Densities, excess molar volumes, and refractive indices of 1,1,2,2-tetrabromoethane and 1-alkanols binary mixtures
rfi	1.42563	303.15	Densities, excess molar volumes, and refractive indices of 1,1,2,2-tetrachloroethane and 1-alkanols binary mixtures
rfi	1.42941	293.15	Densities, excess molar volumes, and refractive indices of 1,1,2,2-tetrachloroethane and 1-alkanols binary mixtures

rfi	1.42958		293.15	Excess molar volumes and viscosities of (1,1,2,2-tetrabromoethane + 1-alkanols) at T = (293.15 and 303.15) K
rfi	1.42884		298.15	Atmospheric densities and interfacial tensions for 1-alkanol (1-butanol to 1-octanol) + water and ether (MTBE, ETBE, DIPE, TAME and THP) + water demixed mixtures.
rfi	1.42760		298.20	Solubility and tie line data for the aqueous solutions of butyric acid with 1-octanol and 2-ethyl-1-hexanol at various temperatures
rfi	1.42800		298.15	Densities, Excess Molar Volumes, and Isobaric Thermal Expansibilities for 1,2-Ethanediol + 1-Butanol, or 1-Hexanol, or 1-Octanol in the Temperature Range from (293.15 to 313.15) K
rhol	818.24	kg/m3	303.15	Excess molar enthalpies and heat capacities of dimethyl sulfoxide + seven normal alkanols at 303.15K and atmospheric pressure
rhol	821.60	kg/m3	298.15	Thermodynamics of 1,3-dimethylurea in eight alcohols

rhol	823.60	kg/m3	293.20	Ternary and Quaternary Liquid-Liquid Equilibria for Systems of Methyl Butyl Ketone + Water + Hydroquinone + Phenol at 313.2 K and Atmospheric Pressure
rhol	832.07	kg/m3	283.15	Densities and Viscosities for Binary Liquid Mixtures of n-Undecane + 1-Heptanol, 1-Octanol, 1-Nonanol, and 1-Decanol from 283.15 to 363.15 K at 0.1 MPa
rhol	828.66	kg/m3	288.15	Densities and Viscosities for Binary Liquid Mixtures of n-Undecane + 1-Heptanol, 1-Octanol, 1-Nonanol, and 1-Decanol from 283.15 to 363.15 K at 0.1 MPa
rhol	825.24	kg/m3	293.15	Densities and Viscosities for Binary Liquid Mixtures of n-Undecane + 1-Heptanol, 1-Octanol, 1-Nonanol, and 1-Decanol from 283.15 to 363.15 K at 0.1 MPa
rhol	821.80	kg/m3	298.15	Densities and Viscosities for Binary Liquid Mixtures of n-Undecane + 1-Heptanol, 1-Octanol, 1-Nonanol, and 1-Decanol from 283.15 to 363.15 K at 0.1 MPa

rhol	818.35	kg/m3	303.15	Densities and Viscosities for Binary Liquid Mixtures of n-Undecane + 1-Heptanol, 1-Octanol, 1-Nonanol, and 1-Decanol from 283.15 to 363.15 K at 0.1 MPa
rhol	814.87	kg/m3	308.15	Densities and Viscosities for Binary Liquid Mixtures of n-Undecane + 1-Heptanol, 1-Octanol, 1-Nonanol, and 1-Decanol from 283.15 to 363.15 K at 0.1 MPa
rhol	811.38	kg/m3	313.15	Densities and Viscosities for Binary Liquid Mixtures of n-Undecane + 1-Heptanol, 1-Octanol, 1-Nonanol, and 1-Decanol from 283.15 to 363.15 K at 0.1 MPa
rhol	807.85	kg/m3	318.15	Densities and Viscosities for Binary Liquid Mixtures of n-Undecane + 1-Heptanol, 1-Octanol, 1-Nonanol, and 1-Decanol from 283.15 to 363.15 K at 0.1 MPa
rhol	804.29	kg/m3	323.15	Densities and Viscosities for Binary Liquid Mixtures of n-Undecane + 1-Heptanol, 1-Octanol, 1-Nonanol, and 1-Decanol from 283.15 to 363.15 K at 0.1 MPa

rhol	800.70	kg/m3	328.15	Densities and Viscosities for Binary Liquid Mixtures of n-Undecane + 1-Heptanol, 1-Octanol, 1-Nonanol, and 1-Decanol from 283.15 to 363.15 K at 0.1 MPa
rhol	797.08	kg/m3	333.15	Densities and Viscosities for Binary Liquid Mixtures of n-Undecane + 1-Heptanol, 1-Octanol, 1-Nonanol, and 1-Decanol from 283.15 to 363.15 K at 0.1 MPa
rhol	793.41	kg/m3	338.15	Densities and Viscosities for Binary Liquid Mixtures of n-Undecane + 1-Heptanol, 1-Octanol, 1-Nonanol, and 1-Decanol from 283.15 to 363.15 K at 0.1 MPa
rhol	789.70	kg/m3	343.15	Densities and Viscosities for Binary Liquid Mixtures of n-Undecane + 1-Heptanol, 1-Octanol, 1-Nonanol, and 1-Decanol from 283.15 to 363.15 K at 0.1 MPa
rhol	785.95	kg/m3	348.15	Densities and Viscosities for Binary Liquid Mixtures of n-Undecane + 1-Heptanol, 1-Octanol, 1-Nonanol, and 1-Decanol from 283.15 to 363.15 K at 0.1 MPa

rhol	782.16	kg/m3	353.15	Densities and Viscosities for Binary Liquid Mixtures of n-Undecane + 1-Heptanol, 1-Octanol, 1-Nonanol, and 1-Decanol from 283.15 to 363.15 K at 0.1 MPa
rhol	778.31	kg/m3	358.15	Densities and Viscosities for Binary Liquid Mixtures of n-Undecane + 1-Heptanol, 1-Octanol, 1-Nonanol, and 1-Decanol from 283.15 to 363.15 K at 0.1 MPa
rhol	774.43	kg/m3	363.15	Densities and Viscosities for Binary Liquid Mixtures of n-Undecane + 1-Heptanol, 1-Octanol, 1-Nonanol, and 1-Decanol from 283.15 to 363.15 K at 0.1 MPa
rhol	821.70	kg/m3	298.20	Physical Behavior of the Phases from the Liquid-Liquid Equilibrium of Citrus Essential Oils Systems at 298.2 K
rhol	824.90	kg/m3	293.15	Correlation Studies of Cyclohexanone/(C5-C10) Alkan-1-ol Binary Mixtures: PC-SAFT Model and Free Volume Theory
rhol	821.40	kg/m3	298.15	Correlation Studies of Cyclohexanone/(C5-C10) Alkan-1-ol Binary Mixtures: PC-SAFT Model and Free Volume Theory

rhol	817.90	kg/m3	303.15	Correlation Studies of Cyclohexanone/(C5-C10) Alkan-1-ol Binary Mixtures: PC-SAFT Model and Free Volume Theory
rhol	814.40	kg/m3	308.15	Correlation Studies of Cyclohexanone/(C5-C10) Alkan-1-ol Binary Mixtures: PC-SAFT Model and Free Volume Theory
rhol	810.90	kg/m3	313.15	Correlation Studies of Cyclohexanone/(C5-C10) Alkan-1-ol Binary Mixtures: PC-SAFT Model and Free Volume Theory
rhol	807.30	kg/m3	318.15	Correlation Studies of Cyclohexanone/(C5-C10) Alkan-1-ol Binary Mixtures: PC-SAFT Model and Free Volume Theory
rhol	803.70	kg/m3	323.15	Correlation Studies of Cyclohexanone/(C5-C10) Alkan-1-ol Binary Mixtures: PC-SAFT Model and Free Volume Theory
rhol	824.90	kg/m3	293.15	Investigation of Molecular Interactions in Binary Mixtures of n-Butyl Acetate and (C6 - C10) 1-Alkanol: PC-SAFT Model
rhol	821.40	kg/m3	298.15	Investigation of Molecular Interactions in Binary Mixtures of n-Butyl Acetate and (C6 - C10) 1-Alkanol: PC-SAFT Model

rhol	817.90	kg/m3	303.15	Investigation of Molecular Interactions in Binary Mixtures of n-Butyl Acetate and (C6 - C10) 1-Alkanol: PC-SAFT Model
rhol	814.40	kg/m3	308.15	Investigation of Molecular Interactions in Binary Mixtures of n-Butyl Acetate and (C6 - C10) 1-Alkanol: PC-SAFT Model
rhol	810.90	kg/m3	313.15	Investigation of Molecular Interactions in Binary Mixtures of n-Butyl Acetate and (C6 - C10) 1-Alkanol: PC-SAFT Model
rhol	807.30	kg/m3	318.15	Investigation of Molecular Interactions in Binary Mixtures of n-Butyl Acetate and (C6 - C10) 1-Alkanol: PC-SAFT Model
rhol	803.70	kg/m3	323.15	Investigation of Molecular Interactions in Binary Mixtures of n-Butyl Acetate and (C6 - C10) 1-Alkanol: PC-SAFT Model
rhol	821.70	kg/m3	298.15	Viscosities and Densities of Fatty Alcohol Mixtures from 298.15 to 338.15 K: Estimation by Kay's Rule and Prediction by the UNIFAC-VISCO and GC-UNIMOD Group Contribution Methods

rhol	814.60	kg/m3	308.15	Viscosities and Densities of Fatty Alcohol Mixtures from 298.15 to 338.15 K: Estimation by Kay's Rule and Prediction by the UNIFAC-VISCO and GC-UNIMOD Group Contribution Methods	
rhol	807.50	kg/m3	318.15	Viscosities and Densities of Fatty Alcohol Mixtures from 298.15 to 338.15 K: Estimation by Kay's Rule and Prediction by the UNIFAC-VISCO and GC-UNIMOD Group Contribution Methods	
rhol	800.40	kg/m3	328.15	Viscosities and Densities of Fatty Alcohol Mixtures from 298.15 to 338.15 K: Estimation by Kay's Rule and Prediction by the UNIFAC-VISCO and GC-UNIMOD Group Contribution Methods	
rhol	793.10	kg/m3	338.15	Viscosities and Densities of Fatty Alcohol Mixtures from 298.15 to 338.15 K: Estimation by Kay's Rule and Prediction by the UNIFAC-VISCO and GC-UNIMOD Group Contribution Methods	
rhol	822.21	kg/m3	298.15	Experimental Solid + Liquid Equilibria and Excess Molar Volume of Alkanol + Octylamine Mixtures. Analysis in Terms of ERAS, DISQUAC, and Modified UNIFAC	

rhol	825.05	kg/m3	293.15	Speeds of Sound, Densities, Isobaric Thermal Expansion, Compressibilities, and Internal Pressures of Heptan-1-ol, Octan-1-ol, Nonan-1-ol, and Decan-1-ol at Temperatures from (293 to 318) K and Pressures up to 100 MPa	
rhol	821.64	kg/m3	298.15	Speeds of Sound, Densities, Isobaric Thermal Expansion, Compressibilities, and Internal Pressures of Heptan-1-ol, Octan-1-ol, Nonan-1-ol, and Decan-1-ol at Temperatures from (293 to 318) K and Pressures up to 100 MPa	
rhol	818.17	kg/m3	303.15	Speeds of Sound, Densities, Isobaric Thermal Expansion, Compressibilities, and Internal Pressures of Heptan-1-ol, Octan-1-ol, Nonan-1-ol, and Decan-1-ol at Temperatures from (293 to 318) K and Pressures up to 100 MPa	
rhol	814.67	kg/m3	308.15	Speeds of Sound, Densities, Isobaric Thermal Expansion, Compressibilities, and Internal Pressures of Heptan-1-ol, Octan-1-ol, Nonan-1-ol, and Decan-1-ol at Temperatures from (293 to 318) K and Pressures up to 100 MPa	

rhol	811.17	kg/m3	313.15	Speeds of Sound, Densities, Isobaric Thermal Expansion, Compressibilities, and Internal Pressures of Heptan-1-ol, Octan-1-ol, Nonan-1-ol, and Decan-1-ol at Temperatures from (293 to 318) K and Pressures up to 100 MPa	
rhol	807.65	kg/m3	318.15	Speeds of Sound, Densities, Isobaric Thermal Expansion, Compressibilities, and Internal Pressures of Heptan-1-ol, Octan-1-ol, Nonan-1-ol, and Decan-1-ol at Temperatures from (293 to 318) K and Pressures up to 100 MPa	
rhol	832.11	kg/m3	283.15	Densities and Viscosities of Three Binary Monoglyme + 1-Alcohol Systems from (283.15 to 313.15) K	
rhol	825.13	kg/m3	293.15	Densities and Viscosities of Three Binary Monoglyme + 1-Alcohol Systems from (283.15 to 313.15) K	
rhol	821.63	kg/m3	298.15	Densities and Viscosities of Three Binary Monoglyme + 1-Alcohol Systems from (283.15 to 313.15) K	
rhol	818.14	kg/m3	303.15	Densities and Viscosities of Three Binary Monoglyme + 1-Alcohol Systems from (283.15 to 313.15) K	

rhol	811.15	kg/m3	313.15	Densities and Viscosities of Three Binary Monoglyme + 1-Alcohol Systems from (283.15 to 313.15) K
rhol	821.65	kg/m3	298.15	Densities and Excess Molar Volumes of N-Methylmorpholine + 1-Alkanol Systems at 298.15 K
rhol	831.00	kg/m3	293.15	Interfacial Properties, Densities, and Contact Angles of Task Specific Ionic Liquids
rhol	803.90	kg/m3	323.15	Densities, Viscosities, and Refractive Indices of Dimethyl Carbonate + 1-Hexanol/1-Octanol Binary Mixtures at Different Temperatures
rhol	811.40	kg/m3	313.15	A thermodynamic study of sublimation, dissolution and distribution processes of anti-inflammatory drug Clonixin
rhol	814.80	kg/m3	308.15	A thermodynamic study of sublimation, dissolution and distribution processes of anti-inflammatory drug Clonixin
rhol	818.30	kg/m3	303.15	A thermodynamic study of sublimation, dissolution and distribution processes of anti-inflammatory drug Clonixin
rhol	821.70	kg/m3	298.15	A thermodynamic study of sublimation, dissolution and distribution processes of anti-inflammatory drug Clonixin

rhol	825.10	kg/m3	293.15	A thermodynamic study of sublimation, dissolution and distribution processes of anti-inflammatory drug Clonixin	
rhol	804.58	kg/m3	323.15	Thermodynamics of mixing for binary mixtures of 1-octanol and 1-decanol with n-dodecane and ternary mixture of (TBP + 1-octanol + dodecane) at T = (298.15 to 323.15) K	
rhol	808.17	kg/m3	318.15	Thermodynamics of mixing for binary mixtures of 1-octanol and 1-decanol with n-dodecane and ternary mixture of (TBP + 1-octanol + dodecane) at T = (298.15 to 323.15) K	
rhol	811.73	kg/m3	313.15	Thermodynamics of mixing for binary mixtures of 1-octanol and 1-decanol with n-dodecane and ternary mixture of (TBP + 1-octanol + dodecane) at T = (298.15 to 323.15) K	
rhol	815.27	kg/m3	308.15	Thermodynamics of mixing for binary mixtures of 1-octanol and 1-decanol with n-dodecane and ternary mixture of (TBP + 1-octanol + dodecane) at T = (298.15 to 323.15) K	
rhol	818.79	kg/m3	303.15	Thermodynamics of mixing for binary mixtures of 1-octanol and 1-decanol with n-dodecane and ternary mixture of (TBP + 1-octanol + dodecane) at T = (298.15 to 323.15) K	

rhol	806.90	kg/m3	318.15	Densities, Viscosities, and Refractive Indices of Dimethyl Carbonate + 1-Hexanol/1-Octanol Binary Mixtures at Different Temperatures
rhol	822.99	kg/m3	298.15	The physicochemical properties and solubility of pharmaceuticals - Methyl xanthines
rhol	821.84	kg/m3	298.15	Densities, speeds of sound and viscosities of binary mixtures of tetrahydrofuran with 1-hexanol, 1-octanol and 1-decanol at T = (298.15 to 313.15) K
rhol	804.12	kg/m3	323.15	Volumetric and compressibility studies on tri-n-butyl phosphate (TBP)-phase modifier (1-octanol, 1-decanol and isodecanol) interactions from T = (298.15 to 323.15) K
rhol	807.69	kg/m3	318.15	Volumetric and compressibility studies on tri-n-butyl phosphate (TBP)-phase modifier (1-octanol, 1-decanol and isodecanol) interactions from T = (298.15 to 323.15) K

rhol	811.24	kg/m3	313.15	Volumetric and compressibility studies on tri-n-butyl phosphate (TBP)-phase modifier (1-octanol, 1-decanol and isodecanol) interactions from T = (298.15 to 323.15) K	
rhol	814.76	kg/m3	308.15	Volumetric and compressibility studies on tri-n-butyl phosphate (TBP)-phase modifier (1-octanol, 1-decanol and isodecanol) interactions from T = (298.15 to 323.15) K	
rhol	818.26	kg/m3	303.15	Volumetric and compressibility studies on tri-n-butyl phosphate (TBP)-phase modifier (1-octanol, 1-decanol and isodecanol) interactions from T = (298.15 to 323.15) K	
rhol	821.74	kg/m3	298.15	Volumetric and compressibility studies on tri-n-butyl phosphate (TBP)-phase modifier (1-octanol, 1-decanol and isodecanol) interactions from T = (298.15 to 323.15) K	
rhol	821.88	kg/m3	298.15	Thermodynamic, transport and excess properties of 2-butoxy ethanol +1-alkanol (C6,C8,C10) at different temperatures	

rhol	807.63	kg/m3	318.15	Volume effects for binary mixtures of propane-1,2-diol with methanol, propan-1-ol, hexan-1-ol, octan-1-ol, or nonan-1-ol at temperatures (293.15 to 318.15) K	
rhol	811.17	kg/m3	313.15	Volume effects for binary mixtures of propane-1,2-diol with methanol, propan-1-ol, hexan-1-ol, octan-1-ol, or nonan-1-ol at temperatures (293.15 to 318.15) K	
rhol	814.69	kg/m3	308.15	Volume effects for binary mixtures of propane-1,2-diol with methanol, propan-1-ol, hexan-1-ol, octan-1-ol, or nonan-1-ol at temperatures (293.15 to 318.15) K	
rhol	821.65	kg/m3	298.15	Volume effects for binary mixtures of propane-1,2-diol with methanol, propan-1-ol, hexan-1-ol, octan-1-ol, or nonan-1-ol at temperatures (293.15 to 318.15) K	
rhol	825.09	kg/m3	293.15	Volume effects for binary mixtures of propane-1,2-diol with methanol, propan-1-ol, hexan-1-ol, octan-1-ol, or nonan-1-ol at temperatures (293.15 to 318.15) K	

rhol	821.22	kg/m3	298.15	Physico-chemical properties of binary mixtures of N,N-dimethylformamide with 1-octanol, 1-nonanol and 1-decanol at different temperatures
rhol	821.79	kg/m3	298.15	Studies of thermophysical properties of binary liquid mixtures of amine and alcohols at various temperatures
rhol	821.67	kg/m3	298.15	Excess molar enthalpies of binary systems containing 2-octanone, hexanoic acid, or octanoic acid at T = 298.15 K
rhol	822.11	kg/m3	298.15	Partial molar volumes of some drug and pro-drug substances in 1-octanol at T = 298.15 K
rhol	811.21	kg/m3	313.15	Excess molar volumes of binary mixtures of 1,3-dimethylimidazolidin-2-one with an alkan-1-ol at the temperatures 283.15 K, 298.15 K, and 313.15 K
rhol	821.72	kg/m3	298.15	Excess molar volumes of binary mixtures of 1,3-dimethylimidazolidin-2-one with an alkan-1-ol at the temperatures 283.15 K, 298.15 K, and 313.15 K
rhol	832.06	kg/m3	283.15	Excess molar volumes of binary mixtures of 1,3-dimethylimidazolidin-2-one with an alkan-1-ol at the temperatures 283.15 K, 298.15 K, and 313.15 K

rhol	796.90	kg/m3	333.15	Effect of polarity and temperature on the binary interaction between D2EHPA extractant and organic solvents (kerosene, n-heptane, chlorobenzene and 1-octanol): Experimental and thermodynamics
rhol	810.60	kg/m3	313.15	Densities, Viscosities, and Refractive Indices of Dimethyl Carbonate + 1-Hexanol/1-Octanol Binary Mixtures at Different Temperatures
rhol	813.80	kg/m3	308.15	Densities, Viscosities, and Refractive Indices of Dimethyl Carbonate + 1-Hexanol/1-Octanol Binary Mixtures at Different Temperatures
rhol	817.20	kg/m3	303.15	Densities, Viscosities, and Refractive Indices of Dimethyl Carbonate + 1-Hexanol/1-Octanol Binary Mixtures at Different Temperatures
rhol	820.70	kg/m3	298.15	Densities, Viscosities, and Refractive Indices of Dimethyl Carbonate + 1-Hexanol/1-Octanol Binary Mixtures at Different Temperatures
rhol	824.60	kg/m3	293.15	Densities, Viscosities, and Refractive Indices of Dimethyl Carbonate + 1-Hexanol/1-Octanol Binary Mixtures at Different Temperatures

rhol	796.88	kg/m3	333.15	Isobaric phase equilibrium of morpholine + 1-decanol, volumetric properties and molar refractivity from 293.15 to 333.15 K of morpholine + 1-decanol and 1-octanol + toluene system with applications of Prigogine-Flory-Patterson theory
rhol	804.05	kg/m3	323.15	Isobaric phase equilibrium of morpholine + 1-decanol, volumetric properties and molar refractivity from 293.15 to 333.15 K of morpholine + 1-decanol and 1-octanol + toluene system with applications of Prigogine-Flory-Patterson theory
rhol	807.70	kg/m3	318.15	Isobaric phase equilibrium of morpholine + 1-decanol, volumetric properties and molar refractivity from 293.15 to 333.15 K of morpholine + 1-decanol and 1-octanol + toluene system with applications of Prigogine-Flory-Patterson theory

rhol	811.09	kg/m3	313.15	Isobaric phase equilibrium of morpholine + 1-decanol, volumetric properties and molar refractivity from 293.15 to 333.15 K of morpholine + 1-decanol and 1-octanol + toluene system with applications of Prigogine-Flory-Patterson theory
rhol	814.31	kg/m3	308.15	Isobaric phase equilibrium of morpholine + 1-decanol, volumetric properties and molar refractivity from 293.15 to 333.15 K of morpholine + 1-decanol and 1-octanol + toluene system with applications of Prigogine-Flory-Patterson theory
rhol	818.08	kg/m3	303.15	Isobaric phase equilibrium of morpholine + 1-decanol, volumetric properties and molar refractivity from 293.15 to 333.15 K of morpholine + 1-decanol and 1-octanol + toluene system with applications of Prigogine-Flory-Patterson theory

rhol	821.90	kg/m3	298.15	Isobaric phase equilibrium of morpholine + 1-decanol, volumetric properties and molar refractivity from 293.15 to 333.15 K of morpholine + 1-decanol and 1-octanol + toluene system with applications of Prigogine-Flory-Patterson theory
rhol	825.57	kg/m3	293.15	Isobaric phase equilibrium of morpholine + 1-decanol, volumetric properties and molar refractivity from 293.15 to 333.15 K of morpholine + 1-decanol and 1-octanol + toluene system with applications of Prigogine-Flory-Patterson theory
rhol	821.72	kg/m3	298.15	Thermodynamics of mixtures containing amines. XVI. CE pm of 1-butanol, 1-octanol or 1-decanol + benzylamine systems at (298.15, 308.15, 318.15 and 333.15) K
rhol	804.20	kg/m3	323.15	Effect of polarity and temperature on the binary interaction between D2EHPA extractant and organic solvents (kerosene, n-heptane, chlorobenzene and 1-octanol): Experimental and thermodynamics

rhol	811.04	kg/m3	313.15	Effect of chain length of alcohol on thermodynamic properties of their binary mixtures with benzylalcohol
rhol	814.68	kg/m3	308.15	Effect of chain length of alcohol on thermodynamic properties of their binary mixtures with benzylalcohol
rhol	818.27	kg/m3	303.15	Effect of chain length of alcohol on thermodynamic properties of their binary mixtures with benzylalcohol
rhol	821.87	kg/m3	298.15	Effect of chain length of alcohol on thermodynamic properties of their binary mixtures with benzylalcohol
rhol	817.52	kg/m3	303.15	Volumetric and transport properties of ternary mixtures containing 1-alkanol + ethyl ethanoate + cyclohexane at 303.15 K: Experimental data, correlation and prediction by ERAS model
rhol	826.00	kg/m3	293.00	KDB
rhol	811.20	kg/m3	313.15	Effect of polarity and temperature on the binary interaction between D2EHPA extractant and organic solvents (kerosene, n-heptane, chlorobenzene and 1-octanol): Experimental and thermodynamics

rhol	818.30	kg/m3	303.15	Effect of polarity and temperature on the binary interaction between D2EHPA extractant and organic solvents (kerosene, n-heptane, chlorobenzene and 1-octanol): Experimental and thermodynamics
rhol	821.60	kg/m3	298.15	Determination and prediction of solubilities of active pharmaceutical ingredients in selected organic solvents
rhol	823.27	kg/m3	298.15	Phase behaviour of tricyanomethanide-based ionic liquids with alcohols and hydrocarbons
rhol	825.95	kg/m3	293.20	Modeling extraction equilibria of butyric acid distributed between water and tri-n-butyl amine/diluent or tri-n-butyl phosphate/diluent system: Extension of the LSER approach
rhol	823.27	kg/m3	298.15	Phase behaviour of ionic liquid 1-butyl-1-methylpyrrolidinium tris(pentafluoroethyl)trifluorophosphate with alcohols, water and aromatic hydrocarbons
rhol	822.47	kg/m3	298.15	Excess molar enthalpy of 1-alkanol + 1-octene mixtures at 298.15K Experimental results and theoretical description by means of the ERAS and TB models

rhol	821.72	kg/m3	298.15	Speeds of Sound and Isentropic Compressibilities in Binary Mixtures of 2-Propanol with Several 1-Alkanols at 298.15K
rhol	832.13	kg/m3	283.15	Volumetric properties of the boldine + alcohol mixtures at atmospheric pressure from 283.15 to 333.15K A new method for the determination of the density of pure boldine
rhol	828.71	kg/m3	288.15	Volumetric properties of the boldine + alcohol mixtures at atmospheric pressure from 283.15 to 333.15K A new method for the determination of the density of pure boldine
rhol	825.27	kg/m3	293.15	Volumetric properties of the boldine + alcohol mixtures at atmospheric pressure from 283.15 to 333.15K A new method for the determination of the density of pure boldine
rhol	821.81	kg/m3	298.15	Volumetric properties of the boldine + alcohol mixtures at atmospheric pressure from 283.15 to 333.15K A new method for the determination of the density of pure boldine

rhol	818.33	kg/m3	303.15	Volumetric properties of the boldine + alcohol mixtures at atmospheric pressure from 283.15 to 333.15K A new method for the determination of the density of pure boldine	
rhol	814.84	kg/m3	308.15	Volumetric properties of the boldine + alcohol mixtures at atmospheric pressure from 283.15 to 333.15K A new method for the determination of the density of pure boldine	
rhol	811.31	kg/m3	313.15	Volumetric properties of the boldine + alcohol mixtures at atmospheric pressure from 283.15 to 333.15K A new method for the determination of the density of pure boldine	
rhol	807.76	kg/m3	318.15	Volumetric properties of the boldine + alcohol mixtures at atmospheric pressure from 283.15 to 333.15K A new method for the determination of the density of pure boldine	
rhol	804.19	kg/m3	323.15	Volumetric properties of the boldine + alcohol mixtures at atmospheric pressure from 283.15 to 333.15K A new method for the determination of the density of pure boldine	

rhol	800.58	kg/m3	328.15	Volumetric properties of the boldine + alcohol mixtures at atmospheric pressure from 283.15 to 333.15K A new method for the determination of the density of pure boldine	
rhol	796.93	kg/m3	333.15	Volumetric properties of the boldine + alcohol mixtures at atmospheric pressure from 283.15 to 333.15K A new method for the determination of the density of pure boldine	
rhol	822.28	kg/m3	298.15	Thermodynamics of mixing for binary mixtures of 1-octanol and 1-decanol with n-dodecane and ternary mixture of (TBP + 1-octanol + dodecane) at T = (298.15 to 323.15) K	
rhol	818.18	kg/m3	303.15	Volume effects for binary mixtures of propane-1,2-diol with methanol, propan-1-ol, hexan-1-ol, octan-1-ol, or nonan-1-ol at temperatures (293.15 to 318.15) K	
speedsl	1330.58	m/s	303.15	Micellar Properties and Related Thermodynamic Parameters of Aqueous Anionic Surfactants in the Presence of Monohydric Alcohols	

speedsl	1297.72	m/s	313.15	Thermodynamic and spectral studies of molecular interactions in binary liquid mixtures of 1-butoxy-2-propanol with 1-alcohols	
speedsl	1314.28	m/s	308.15	Thermodynamic and spectral studies of molecular interactions in binary liquid mixtures of 1-butoxy-2-propanol with 1-alcohols	
speedsl	1330.99	m/s	303.15	Thermodynamic and spectral studies of molecular interactions in binary liquid mixtures of 1-butoxy-2-propanol with 1-alcohols	
speedsl	1347.32	m/s	298.15	Thermodynamic and spectral studies of molecular interactions in binary liquid mixtures of 1-butoxy-2-propanol with 1-alcohols	
speedsl	1364.08	m/s	293.15	Thermodynamic and spectral studies of molecular interactions in binary liquid mixtures of 1-butoxy-2-propanol with 1-alcohols	
speedsl	1282.90	m/s	318.15	Ultrasonic study of molecular interactions in binary mixtures of methyl acrylate with 1-alkanols (C4 to C10) at temperatures from (288.15 to 318.15) K	

speedsl	1298.90	m/s	313.15	Ultrasonic study of molecular interactions in binary mixtures of methyl acrylate with 1-alkanols (C4 to C10) at temperatures from (288.15 to 318.15) K
speedsl	1315.70	m/s	308.15	Ultrasonic study of molecular interactions in binary mixtures of methyl acrylate with 1-alkanols (C4 to C10) at temperatures from (288.15 to 318.15) K
speedsl	1333.10	m/s	303.15	Ultrasonic study of molecular interactions in binary mixtures of methyl acrylate with 1-alkanols (C4 to C10) at temperatures from (288.15 to 318.15) K
speedsl	1352.40	m/s	298.15	Ultrasonic study of molecular interactions in binary mixtures of methyl acrylate with 1-alkanols (C4 to C10) at temperatures from (288.15 to 318.15) K
speedsl	1371.50	m/s	293.15	Ultrasonic study of molecular interactions in binary mixtures of methyl acrylate with 1-alkanols (C4 to C10) at temperatures from (288.15 to 318.15) K
speedsl	1390.20	m/s	288.15	Ultrasonic study of molecular interactions in binary mixtures of methyl acrylate with 1-alkanols (C4 to C10) at temperatures from (288.15 to 318.15) K

speedsl	1313.92	m/s	308.15	Micellar Properties and Related Thermodynamic Parameters of Aqueous Anionic Surfactants in the Presence of Monohydric Alcohols
speedsl	1347.40	m/s	298.15	Micellar Properties and Related Thermodynamic Parameters of Aqueous Anionic Surfactants in the Presence of Monohydric Alcohols
srf	0.03	N/m	288.15	Density and Surface Tension Variation with Temperature for Heptane + 1-Alkanol
srf	0.03	N/m	293.15	Density and Surface Tension Variation with Temperature for Heptane + 1-Alkanol
srf	0.03	N/m	298.15	Density and Surface Tension Variation with Temperature for Heptane + 1-Alkanol
srf	0.03	N/m	303.15	Density and Surface Tension Variation with Temperature for Heptane + 1-Alkanol
srf	0.03	N/m	308.15	Density and Surface Tension Variation with Temperature for Heptane + 1-Alkanol
srf	0.03	N/m	293.20	KDB

Pressure Dependent Properties

Property code	Value	Unit	Pressure [kPa]	Source
tbrp	371.20	K	2.50	NIST Webbook

Correlations

Information	Value
Property code	pvap
Equation	ln(Pvp) = A + B/(T + C)
Coeff. A	1.45038e+01
Coeff. B	-3.54647e+03
Coeff. C	-1.09245e+02
Temperature range (K), min.	358.71
Temperature range (K), max.	495.05

Information	Value
Property code	pvap
Equation	ln(Pvp) = A + B/T + C*ln(T) + D*T^2
Coeff. A	2.04801e+02
Coeff. B	-1.63274e+04
Coeff. C	-2.73260e+01
Coeff. D	1.22481e-05
Temperature range (K), min.	257.65
Temperature range (K), max.	652.50

Datasets

Thermal conductivity, W/m/K

Temperature, K - Liquid	Pressure, kPa - Liquid	Thermal conductivity, W/m/K - Liquid
274.92	100.00	0.1580
274.85	5000.00	0.1598
274.82	10000.00	0.1611
274.77	20000.00	0.1639
274.63	30000.00	0.1669
294.93	100.00	0.1545
294.62	5000.00	0.1560

294.61	10100.00	0.1576
294.59	20100.00	0.1606
294.58	30000.00	0.1634
314.41	100.00	0.1511
314.95	5000.00	0.1530
314.91	10000.00	0.1545
314.86	20000.00	0.1578
314.86	30000.00	0.1609
334.98	100.00	0.1471
334.92	5000.00	0.1490
334.90	10000.00	0.1509
334.84	20000.00	0.1544
334.83	30000.00	0.1576
354.98	100.00	0.1427
354.41	5000.00	0.1445
354.36	10000.00	0.1466
354.50	20000.00	0.1505
354.35	30000.00	0.1539
374.96	100.00	0.1380
374.43	5100.00	0.1400
374.41	10000.00	0.1421
374.81	20000.00	0.1464
374.84	30000.00	0.1503
394.88	300.00	0.1325
395.13	5000.00	0.1349
395.05	10000.00	0.1377
394.82	20000.00	0.1420
394.85	30000.00	0.1461

Reference *https://www.doi.org/10.1021/acs.jced.9b00628*

Mass density, kg/m3

Temperature, K - Liquid	Pressure, kPa - Liquid	Mass density, kg/m3 - Liquid
298.15	100.00	821.4

Reference *https://www.doi.org/10.1016/j.fluid.2019.02.006*

Temperature, K	Pressure, kPa	Mass density, kg/m3
278.15	100.00	835.2

278.15	1000.00	835.7
278.15	5000.00	838.0
278.15	10000.00	841.0
278.15	15000.00	843.7
278.15	20000.00	846.4
278.15	25000.00	849.0
278.15	30000.00	851.5
278.15	35000.00	854.0
278.15	40000.00	856.4
278.15	45000.00	858.8
278.15	50000.00	861.1
278.15	55000.00	863.4
278.15	60000.00	865.5
288.15	100.00	828.4
288.15	1000.00	829.0
288.15	5000.00	831.5
288.15	10000.00	834.4
288.15	15000.00	837.3
288.15	20000.00	840.1
288.15	25000.00	842.8
288.15	30000.00	845.5
288.15	35000.00	848.0
288.15	40000.00	850.5
288.15	45000.00	852.9
288.15	50000.00	855.3
288.15	55000.00	857.6
288.15	60000.00	859.8
298.15	100.00	821.4
298.15	1000.00	822.0
298.15	5000.00	824.6
298.15	10000.00	827.7
298.15	15000.00	830.7
298.15	20000.00	833.6
298.15	25000.00	836.4
298.15	30000.00	839.1
298.15	35000.00	841.8
298.15	40000.00	844.4
298.15	45000.00	846.9
298.15	50000.00	849.4
298.15	55000.00	851.7
298.15	60000.00	853.9
308.15	100.00	814.5
308.15	1000.00	815.1
308.15	5000.00	817.8

308.15	10000.00	821.1
308.15	15000.00	824.2
308.15	20000.00	827.2
308.15	25000.00	830.1
308.15	30000.00	833.0
308.15	35000.00	835.7
308.15	40000.00	838.4
308.15	45000.00	841.0
308.15	50000.00	843.5
308.15	55000.00	846.0
308.15	60000.00	848.2
318.15	100.00	807.5
318.15	1000.00	808.1
318.15	5000.00	811.0
318.15	10000.00	814.4
318.15	15000.00	817.7
318.15	20000.00	820.8
318.15	25000.00	823.8
318.15	30000.00	826.8
318.15	35000.00	829.6
318.15	40000.00	832.4
318.15	45000.00	835.1
318.15	50000.00	837.7
318.15	55000.00	840.3
318.15	60000.00	842.6
328.15	100.00	800.4
328.15	1000.00	801.1
328.15	5000.00	804.1
328.15	10000.00	807.7
328.15	15000.00	811.1
328.15	20000.00	814.4
328.15	25000.00	817.5
328.15	30000.00	820.5
328.15	35000.00	823.5
328.15	40000.00	826.4
328.15	45000.00	829.1
328.15	50000.00	831.9
328.15	55000.00	834.5
328.15	60000.00	836.9
338.15	100.00	793.2
338.15	1000.00	793.9
338.15	5000.00	796.9
338.15	10000.00	800.8
338.15	15000.00	804.4

338.15	20000.00	807.8
338.15	25000.00	811.1
338.15	30000.00	814.2
338.15	35000.00	817.3
338.15	40000.00	820.3
338.15	45000.00	823.1
338.15	50000.00	825.9
338.15	55000.00	828.7
338.15	60000.00	831.2
348.15	100.00	785.6
348.15	1000.00	786.4
348.15	5000.00	789.8
348.15	10000.00	793.6
348.15	15000.00	797.4
348.15	20000.00	801.0
348.15	25000.00	804.5
348.15	30000.00	807.8
348.15	35000.00	810.9
348.15	40000.00	814.0
348.15	45000.00	817.0
348.15	50000.00	819.9
348.15	55000.00	822.7
348.15	60000.00	825.5
358.15	100.00	778.0
358.15	1000.00	778.8
358.15	5000.00	782.3
358.15	10000.00	786.4
358.15	15000.00	790.4
358.15	20000.00	794.1
358.15	25000.00	797.7
358.15	30000.00	801.1
358.15	35000.00	804.5
358.15	40000.00	807.7
358.15	45000.00	810.8
358.15	50000.00	813.6
358.15	55000.00	816.7
358.15	60000.00	819.6

Reference *https://www.doi.org/10.1016/j.jct.2011.10.023*

Temperature, K	Pressure, kPa	Mass density, kg/m3
298.15	2000.00	823.6
298.15	4010.00	824.8

298.15	6000.00	826.0
298.15	8000.00	827.2
298.15	10000.00	828.4
298.15	13010.00	830.1
298.15	16010.00	831.8
298.15	18990.00	833.4
298.15	22010.00	835.0
298.15	25010.00	836.6
298.15	28010.00	838.3
298.15	30010.00	839.3
303.15	1990.00	819.4
303.15	3990.00	820.6
303.15	5980.00	821.9
303.16	7980.00	823.1
303.15	10010.00	824.4
303.16	13000.00	826.2
303.15	16020.00	827.9
303.15	19030.00	829.7
303.16	22010.00	831.3
303.15	25000.00	833.0
303.16	28000.00	834.6
303.15	29980.00	835.6
313.15	1970.00	812.3
313.15	4000.00	813.6
313.15	6010.00	815.0
313.15	8000.00	816.3
313.15	10010.00	817.5
313.15	13030.00	819.4
313.15	15980.00	821.2
313.15	18990.00	823.0
313.15	21990.00	824.8
313.15	25000.00	826.5
313.15	27980.00	828.2
313.15	29990.00	829.3
323.15	2010.00	805.6
323.15	4010.00	807.0
323.15	6010.00	808.4
323.15	8000.00	809.7
323.15	10010.00	811.1
323.15	13020.00	813.0
323.15	16020.00	814.9
323.15	19010.00	816.8
323.15	22000.00	818.6
323.15	25010.00	820.4

323.15	28010.00	822.1
323.15	30020.00	823.3

Reference *https://www.doi.org/10.1021/acs.jced.8b00812*

Temperature, K	Pressure, kPa	Mass density, kg/m3
298.15	100.00	821.4

Reference *https://www.doi.org/10.1021/acs.jced.8b01189*

Temperature, K	Pressure, kPa	Mass density, kg/m3
313.14	1004.00	811.26
313.14	2019.00	811.95
313.14	3000.00	812.59
313.14	4020.00	813.22
313.14	4995.00	813.89
313.14	6000.00	814.53
313.14	7007.00	815.21
313.14	8012.00	815.83
313.14	8994.00	816.51
313.14	10014.00	817.19
313.14	11007.00	817.79
313.14	12000.00	818.42
313.14	12997.00	819.05
313.14	14001.00	819.67
313.14	14997.00	820.27
313.14	16002.00	820.89
313.14	17001.00	821.48
313.14	17997.00	822.1
313.14	19004.00	822.69
313.14	20007.00	823.28
313.14	21004.00	823.86
323.09	1007.00	804.26
323.09	2016.00	804.97
323.09	3019.00	805.7
323.09	4025.00	806.4
323.09	5012.00	807.07
323.09	6093.00	807.82
323.09	7143.00	808.56
323.09	8009.00	809.12
323.09	8989.00	809.77
323.09	10014.00	810.47

323.09	11031.00	811.14
323.09	12010.00	811.82
323.09	13005.00	812.44
323.09	14003.00	813.11
323.09	15012.00	813.76
323.09	16009.00	814.4
323.09	17005.00	815.04
323.09	18013.00	815.66
323.09	19003.00	816.26
323.09	20025.00	816.9
323.09	21081.00	817.56
323.09	22009.00	818.09
333.05	1038.00	797.31
333.05	2029.00	798.03
333.05	2999.00	798.76
333.05	4033.00	799.51
333.05	5020.00	800.25
333.05	6014.00	800.97
333.05	7029.00	801.69
333.05	8011.00	802.44
333.05	9005.00	803.11
333.05	10022.00	803.83
333.05	11011.00	804.55
333.05	12021.00	805.23
333.05	13022.00	805.92
333.05	14031.00	806.61
333.05	15010.00	807.28
333.05	16023.00	807.92
333.05	17012.00	808.59
333.05	18035.00	809.23
333.05	19009.00	809.87
333.05	20002.00	810.51
333.05	21011.00	811.15
333.05	22019.00	811.82
343.03	1049.00	790.6
343.03	2003.00	791.16
343.03	2981.00	791.93
343.03	4017.00	792.68
343.03	5000.00	793.37
343.03	6005.00	794.05
343.03	7018.00	794.88
343.03	8104.00	795.68
343.03	9014.00	796.44
343.03	10022.00	797.1

343.03	11028.00	797.87
343.03	11997.00	798.58
343.03	13030.00	799.3
343.03	14005.00	800.0
343.03	15002.00	800.66
343.03	16032.00	801.38
343.03	17016.00	801.99
343.03	18012.00	802.68
343.03	19005.00	803.29
343.03	20018.00	803.98
343.03	21014.00	804.61
343.03	22004.00	805.23
352.89	1023.00	782.71
352.89	2011.00	783.55
352.89	3032.00	784.42
352.89	4022.00	785.42
352.89	5005.00	786.18
352.89	6026.00	787.13
352.89	7003.00	787.9
352.89	8016.00	788.71
352.89	9007.00	789.47
352.89	10003.00	790.22
352.89	11022.00	791.0
352.89	12170.00	791.78
352.89	13015.00	792.51
352.89	14018.00	793.24
352.89	15009.00	794.0
352.89	16023.00	794.74
352.89	17010.00	795.48
352.89	18002.00	796.16
352.89	19029.00	796.9
352.89	20046.00	797.63
352.89	21023.00	798.3
362.77	1051.00	775.67
362.77	2025.00	776.41
362.77	3016.00	777.19
362.77	4009.00	778.05
362.77	4974.00	778.77
362.77	6005.00	779.51
362.77	7006.00	780.3
362.77	8011.00	780.95
362.77	9008.00	781.79
362.77	10034.00	782.62
362.77	11013.00	783.32

362.77	11987.00	784.03
362.77	13015.00	784.77
362.77	14006.00	785.51
362.77	14997.00	786.3
362.77	16001.00	787.03
362.77	17013.00	787.79
362.77	17997.00	788.55
362.77	19000.00	789.33
362.77	20023.00	790.04
362.77	21004.00	790.83
Reference		*https://www.doi.org/10.1021/je700145e*

Sources

Volume effects for binary mixtures of propane-1,2-diol with methanol, propan-1-ol, hexan-1-ol, octan-1-ol, or nonan-1-ol at temperatures (293.15 to 318.15) K:	https://www.doi.org/10.1016/j.jct.2012.12.003
On the photothermal characterization of liquid thermoelectrics - New methodology based on coupled pyroelectric-Seebeck effects, together with frequency and thickness scanning procedures:	https://www.doi.org/10.1016/j.tca.2017.04.013
Partition coefficient n-octanol/water of propranolol and atenolol at different temperatures: Experimental and theoretical studies:	https://www.doi.org/10.1016/j.jct.2012.05.021
Experimental solubility of clotrimazole and some thermodynamic aspects of dissolution in different solvents:	https://www.doi.org/10.1016/j.tca.2019.178431
Densities and Viscosities of Four Binary Diethyl Carbonate + 1-Alcohol Systems from (288.15 to 313.15) K:	https://www.doi.org/10.1021/je1013476
Excess Volumes, Densities, Speeds of Sound, and Viscosities for the Binary Systems of 1-Octanol with Hexadecane and Squalane at (298.15, 303.15 and 308.15) K:	https://www.doi.org/10.1007/s10765-008-0491-0
Henry's Law Constants of Organic Compounds in Water and n-Octane at T = 303.15 K:	https://www.doi.org/10.1021/je900711h
Excess molar volumes of binary mixtures of 1,3-dimethylimidazolidin-2-one with an alkan-1-ol at the temperatures 283.15 K, 298.15 K, and 313.15 K:	https://www.doi.org/10.1016/j.jct.2006.09.017
Densities, speeds of sound and viscosities of binary mixtures of tetrahydropyran with 1-hexanol, 1-octanol and 1-decanol at T = (298.15 to 313.15) K:	https://www.doi.org/10.1016/j.jct.2013.10.025
Densities, Viscosities, and Speed of Sound Studies of Binary Mixtures of Methylbenzene with Heptan-1-ol, Octan-1-ol, and Decan-1-ol at (303.15 and 313.15) K:	https://www.doi.org/10.1021/je0602519
Excess molar enthalpies and heat capacities of dimethyl sulfoxide + seven normal alkanols at 303.15K and atmospheric pressure:	https://www.doi.org/10.1016/j.tca.2006.10.025
Parametric Analysis of Mandelic Acid Separation from Aqueous Solutions by Using Secondary Amine Mixture (Amberlite LA-2) in Various Diluents:	https://www.doi.org/10.1021/acs.jced.9b00169
Partial molar volumes of some drug and pro-drug substances in 1-octanol at T = 298.15 K:	https://www.doi.org/10.1016/j.jct.2009.10.002
Liquid-liquid phase equilibrium of (piperidinium-based ionic liquid + an alcohol) binary systems and modelling with NRHB and PCP-SAFT:	https://www.doi.org/10.1016/j.fluid.2011.03.005
The physicochemical properties and solubility of pharmaceuticals - Methyl xanthines:	https://www.doi.org/10.1016/j.jct.2014.05.005
Solubility of the Antimicrobial Agent Triclosan in Organic Solvents of Different Hydrogen Bonding Capabilities at Several Temperatures:	https://www.doi.org/10.1021/je800426w
Determination of Solubility and Thermodynamic Properties of Benzophenone in Different Pure Solvents:	https://www.doi.org/10.1021/acs.jced.8b00196
Tie-line data for the aqueous solutions of phenol with organic solvents at T = 298.2 K:	https://www.doi.org/10.1016/j.jct.2012.10.028
Densities, Viscosities, and Refractive Indices of Dimethyl Carbonate + 1-Hexanol/1-Octanol Binary Mixtures at Different Temperatures:	https://www.doi.org/10.1021/acs.jced.5b00830
Solubility of Berberine Chloride in Various Solvents:	https://www.doi.org/10.1021/je0504360
Thermodynamics of 1,3-dimethylurea in eight alcohols:	https://www.doi.org/10.1021/acs.jced.5b00882
Speeds of Sound and Isentropic Compressibilities in Binary Mixtures of 2-Propanol with Several 1-Alkanols at 298.15K:	https://www.doi.org/10.1007/s10765-005-8101-x
KDB Pure (Korean Thermophysical Properties Databank):	https://www.cheric.org/research/kdb/hcprop/showprop.php?cmpid=841
Volumetric properties of the boldine + alcohol mixtures at atmospheric pressure from 283.15 to 333.15K A new method for the determination of the density of pure boldine:	https://www.doi.org/10.1016/j.fluid.2007.04.030
Excess molar enthalpy of 1-alkanol + 1-octene mixtures at 298.15K Experimental results and theoretical description by means of the ERAS and TB models:	https://www.doi.org/10.1016/j.fluid.2007.07.063
Density, Viscosity, and Speed of Sound of (1-Octanol + 2-Methoxyethanol), (1-Octanol + N,N-Dimethylacetamide), and (1-Octanol + Acetophenone) at Temperatures of (298.15, 308.15, and 318.15) K:	https://www.doi.org/10.1021/je100170v

Study of molecular interactions and preferential solvation in binary mixtures of cyclohexane and (C5-C10) 1-alkanol by Kirkwood-Buff integrals:	https://www.doi.org/10.1016/j.fluid.2019.02.006
Ternary phase diagrams for aqueous mixtures of butyric acid with several solvents: experimental and correlated data:	https://www.doi.org/10.1016/j.fluid.2014.03.008
Measurement and Modeling of Solubility of para-tert-Butylbenzoic Acid in Pure and Mixed Organic Solvents at Different Temperatures:	https://www.doi.org/10.1021/acs.jced.6b00965
Thermodynamic aspects of solubility and partitioning processes of some sulfonamides in the solvents modeling biological media:	https://www.doi.org/10.1016/j.jct.2013.09.027
Solution thermodynamics of pyrazinamide, isoniazid, and p-aminobenzoic acid in buffers and octanol:	https://www.doi.org/10.1016/j.jct.2015.08.022
Ternary and Quaternary Liquid-Liquid Equilibria for Systems of Methyl Butyl Ketone + Water + Hydroquinone + Phenol at 313.2 K and Atmospheric Pressure:	https://www.doi.org/10.1021/acs.jced.5b00918
Effect of polarity and temperature on the binary interaction between D2EHPA extractant and organic solvents (kerosene, n-heptane, chlorobenzene and 1-octanol): Experimental and thermodynamics:	https://www.doi.org/10.1016/j.fluid.2016.11.029
Ultrasonic study of molecular interactions in binary mixtures of methyl acetate with 1-alkanols (C4 to C8) at temperatures from (288.15 to 318.15) K:	https://www.doi.org/10.1016/j.jct.2012.11.027
Solubility of 1-Alkyl-3-ethylimidazolium-Based Ionic Liquids in Water and 1-Octanol:	https://www.doi.org/10.1021/je700693z
Studies of phase equilibria and the physicochemical properties of selected pharmaceuticals:	https://www.doi.org/10.1016/j.fluid.2015.07.037
Acoustical and Excess Properties of {Chlorobenzene + 1-Hexanol, or 1-Heptanol, or 1-Octanol, or 1-Nonanol, or 1-Decanol} at (298.15, 303.15, 308.15, and 313.15) K:	https://www.doi.org/10.1021/je060353z
Vapor pressure of selected aliphatic alcohols by ebulliometry. Part 1:	https://www.doi.org/10.1016/j.fluid.2010.06.019
Measurements of the Thermal Conductivity of n-Octane, Isooctane, 1-Octene, and 1-Octanol in the Temperature Range from 253 to 393 K at Pressures up to 30 MPa:	https://www.doi.org/10.1021/acs.jced.9b00628
Vapor pressure of 1-octanol below 5 kPa using DSC:	https://www.doi.org/10.1016/j.tca.2013.03.018
SLE and LLE for tri-butylphosphate or tri-octylamine contained systems; extractive solvents of Molybdenum:	https://www.doi.org/10.1016/j.fluid.2010.04.016
Effect of Pressure on the Static Relative Permittivities of Alkan-1-ols at 298.15 K:	https://www.doi.org/10.1021/je060248p
Fluid-Phase Behavior of {1-Hexyl-3-methylimidazolium Bis(trifluoromethylsulfonyl) imide, [C6mim][NTf2], + C2−C8 n-Alcohol} Mixtures: Liquid−Liquid Equilibrium and Excess Volumes:	https://www.doi.org/10.1021/je060307z
Measurement of Density and Viscosity of Binary 1-Alkanol Systems (C8-C11) at 101 kPa and Temperatures from (293.15 to 313.15) K:	https://www.doi.org/10.1021/je050200w
Solubility of Water in Octan-1-ol from (275 to 369) K:	https://www.doi.org/10.1021/je3001427
Physico-chemical properties of binary mixtures of N,N-dimethylformamide with 1-octanol, 1-nonanol and 1-decanol at different temperatures:	https://www.doi.org/10.1016/j.jct.2012.11.015
Liquid-Liquid Equilibria for the Ternary and Quaternary Systems of Water + Trimethylolpropane + (2-Ethylhexan-1-ol/Octan-1-ol) + Formaldehyde at 298.15 K:	https://www.doi.org/10.1021/acs.jced.9b00519
Physical Behavior of the Phases from the Liquid-Liquid Equilibrium of Citrus Essential Oil Systems at 298.2 K:	https://www.doi.org/10.1021/acs.jced.8b00086
Solubility of ionic liquids in water and octan-1-ol and octan-1-ol/water, or 2-phenylethanol/water partition coefficients:	https://www.doi.org/10.1016/j.jct.2012.06.003
Densities and Viscosities for Binary Liquid Mixtures of n-Undecane + 1-Heptanol, 1-Octanol, 1-Nonanol, and 1-Decanol from 283.15 to 363.15 K at 0.1 MPa:	https://www.doi.org/10.1021/acs.jced.6b00834
(Liquid + liquid) phase equilibria of 1-alkyl-3-methylimidazolium methylsulfate with alcohols, or ethers, or ketones:	https://www.doi.org/10.1016/j.jct.2005.07.024
Liquid-Liquid Phase Behavior of Solutions of 1,3-Dimethylimidazolium- and 1-Methyl-3-propylimidazolium Bis(trifluoromethylsulfonyl)amide (C1,3mimNTf2) in 1-Alkyl Alcohols:	https://www.doi.org/10.1021/je400253e
(Solid + liquid) and (liquid + liquid) phase equilibria study and correlation of the binary systems of {N-butyl-3-methylpyridinium tosylate + water, or + an alcohol, or + a hydrocarbon}:	https://www.doi.org/10.1016/j.fluid.2010.01.025
Densities, Excess Molar Volumes, and Isobaric Thermal Expansibilities for 1,2-Ethanediol + 1-Butanol, or 1-Hexanol, or 1-Octanol in the Temperature Range from (293.15 to 313.15) K:	https://www.doi.org/10.1021/je700580g
Effect of chain length of alcohol on thermodynamic properties of their binary mixtures with benzylalcohol:	https://www.doi.org/10.1016/j.tca.2014.01.027
Measurements, Correlations, and Mod. UNIFAC (Do) Prediction of (Solid-Liquid) Phase Equilibria Diagrams in Binary Systems (Aliphatic Ketone + an Alcohol):	https://www.doi.org/10.1021/je100725a
Heat Capacities and Derived Thermodynamic Functions of 1-Hexanol, 1-Heptanol, 1-Octanol, and 1-Decanol between 5 K and 390 K:	https://www.doi.org/10.1021/je0340856
Physico-chemical characterization of antituberculosis thioacetazone: Vapor pressure, solubility and lipophilicity:	https://www.doi.org/10.1016/j.jct.2016.12.034
Liquid-Liquid Phase Behavior of Solutions of 1-Octyl- and 1-Decyl-3-methylimidazolium Bis(trifluoromethylsulfonyl)imide (C8,10mimNTf2) in n-Alkyl Alcohols:	https://www.doi.org/10.1021/je900988a
Interfacial Properties, Densities, and Contact Angles of Task Specific Ionic Liquids:	https://www.doi.org/10.1021/je900502s
Thermodynamic Modeling of Triglycerides using PC-SAFT:	https://www.doi.org/10.1021/acs.jced.8b01046
Solubility Determination and Thermodynamic Modeling of Buprofezin in Different Solvents and Mixing Properties of Solutions:	https://www.doi.org/10.1021/acs.jced.8b01099
Solubility and tieline data for the aqueous solutions of butyric acid with 1-octanol and 2-ethyl-1-hexanol at various temperatures:	https://www.doi.org/10.1016/j.fluid.2013.10.022
Chemical Equilibrium of the Liquid-Phase Dehydration of 1-Octanol to 1-(Octyloxy)octane:	https://www.doi.org/10.1021/je301236k
Densities, Isothermal Compressibilities, and Isobaric Thermal Expansivities of Hexan-2-ol, Octan-1-ol and Decan-1-ol from (313 to 363) K and Pressures up to 22 MPa:	https://www.doi.org/10.1021/je700145e
Liquid Liquid Phase Equilibria of 1-Propanol + Water + n-Alcohol Ternary Systems at 298.15 K and Atmospheric Pressure:	https://www.doi.org/10.1021/je2002539
Benchmarks for the fifth industrial fluid properties simulation challenge:	https://www.doi.org/10.1016/j.fluid.2009.07.004
Densities, excess molar volumes, and refractive indices of 1,1,2,2-tetrachloroethane and 1-alkanols binary mixtures:	https://www.doi.org/10.1016/j.jct.2005.06.015
Modeling extraction equilibria of butyric acid distributed between water and tri-n-butyl amine/diluent or tri-n-butyl phosphate/diluent system: Extension of the LSER approach:	https://www.doi.org/10.1016/j.fluid.2014.10.043

Liquid Liquid Equilibrium for Ternary Systems Containing Water, Oleic Acid, and Alcohols at 313.15 K. Effect of Alcohol Chain Length:	https://www.doi.org/10.1021/je501174y
Liquid phase behavior of imidazolium-based ionic liquids with alcohols: effect of hydrogen bonding and non-polar interactions:	https://www.doi.org/10.1016/j.fluid.2004.07.014
Excess molar enthalpies for mixtures of carbon dioxide + a modifier (5 mol% methanol or 1-octanol) and hexane at 308.15K and 12.40MPa:	https://www.doi.org/10.1016/j.fluid.2005.12.029
Isothermal (vapour + liquid) equilibrium at several temperatures of (1-chlorobutane + 1-octanol, or 1-decanol):	https://www.doi.org/10.1016/j.jct.2003.09.009
Phase equilibria of didecyldimethylammonium nitrate ionic liquid with water and organic solvents:	https://www.doi.org/10.1016/j.jct.2006.10.009
Densities, excess molar volumes, and refractive indices of 1,1,2,2-tetrabromoethane and 1-alkanols binary mixtures:	https://www.doi.org/10.1016/j.jct.2005.11.012
Solubilities of Sulfamethazine, Sulfadimethoxine, Sulfamethoxydiazine, Sulfamonomethoxine, Sulfamethoxazole, and Sulfaquinoxaline in 1-Octanol from (298.15 to 333.15) K	https://www.doi.org/10.1021/je800867a
Density and Surface Tension Variation with Temperature for Heptane + 1-Alkanol	https://www.doi.org/10.1021/je060179e
Experimental (Solid + Liquid) and (Liquid + Liquid) Equilibria and Excess Molar Volume of Alkanol + Acetonitrile, Propanenitrile, and Butanenitrile Mixtures:	https://www.doi.org/10.1021/je050262m
Liquid-Liquid Equilibrium for Ternary Systems of Water + 2,2,3,3-Tetrafluoro-1-propanol + Anisole/1-Octanol at 298.2, 308.2, and 318.2 K.	https://www.doi.org/10.1021/acs.jced.8b00393
Experimental study and ERAS modeling of the excess molar enthalpy of (acetonitrile + 1-heptanol or 1-octanol) mixtures at (298.15, 313.15, and 323.15) K and atmospheric pressure	https://www.doi.org/10.1016/j.jct.2008.01.017
Solubility of naproxen in several organic solvents at different temperature:	https://www.doi.org/10.1016/j.fluid.2007.03.029
Towards automated characterisation of liquid-liquid equilibria:	https://www.doi.org/10.1016/j.fluid.2013.10.048
Thermodynamic aspects of solubility, solvation and partitioning processes of some sulfonamides:	https://www.doi.org/10.1016/j.jct.2010.12.007
Solubility and dissolution characteristics of capecitabine in pure lower alcohols and water with methanol mixture solvents at atmospheric pressure and different temperatures:	https://www.doi.org/10.1016/j.fluid.2017.12.024
Sulfasalazine: Dissolution and Distribution in Pharmaceutically Relevant Mediums:	https://www.doi.org/10.1021/acs.jced.6b00497
Isobaric phase equilibrium of morpholine + 1-decanol, volumetric properties and molar refractivity from 293.15 to 333.15 K of morpholine + 1-decanol and 1-octanol + toluene system with applications of Prigogine-Flory-Patterson theory:	https://www.doi.org/10.1016/j.tca.2016.12.010
Evaluation of (vapor + liquid) equilibria for the binary systems (1-octanol + cyclohexane) and (1-octanol + n-hexane) with [illegible]	https://www.doi.org/10.1016/j.jct.2008.06.005
Liquid Phase Behavior of Ionic Liquids with Water and 1-Octanol and Modeling of 1-Octanol/Water Partition Coefficients:	https://www.doi.org/10.1021/je7003935
Liquid-Liquid Phase Behavior of Solutions of 1-Hexyl-3-methylimidazolium Bis((trifluoromethyl)sulfonyl)amide (C6mimNTf2) in n-Alkyl Alcohols :	https://www.doi.org/10.1021/je1011357
Estimated Solubility Method:	http://pubs.acs.org/doi/suppl/10.1021/ci034243x/suppl_file/ci034243xsi20040112_053635.txt
Solubility Determination and Correlation of (2R,3S,4S,5S)-6-(Hydroxymethyl)tetrahydro-2H-pyran-2,3,4,5-tetraol in Fatty Alcohol:	https://www.doi.org/10.1021/je4011097
Study of molecular interactions in binary liquid mixtures of 1-octanol with n-hexane, n-octane, and n-decane using volumetric, viscometric, and acoustic properties:	https://www.doi.org/10.1016/j.jct.2008.02.005
A thermodynamic study of solute solvent interactions through dielectric properties of binary mixtures consisting of 1,4-butanediol, 1-octanol, and 1,4-dioxane at different temperatures:	https://www.doi.org/10.1016/j.jct.2012.05.032
Unusual behavior of the isotope effect on miscibility of 1-hexyl-3-methylimidazolium bis(trifluorosulfonyl)imide with 1-alkanols:	https://www.doi.org/10.1016/j.fluid.2014.03.005
Enhanced partition model of 4-nitrophenol in water - octanol system. Effects of association/dissociation processes:	https://www.doi.org/10.1016/j.fluid.2016.08.021
The Yaws Handbook of Vapor Pressure:	https://www.sciencedirect.com/book/9780128029992/the-yaws-handbook-of-vapor-pressure
Studies of thermophysical properties of binary liquid mixtures of amine and alcohols at various temperatures:	https://www.doi.org/10.1016/j.jct.2012.01.022
Solubility and distribution of salinazid and vanillin isoniazid in the solvents modeling biological media:	https://www.doi.org/10.1021/acs.jced.5b00352
Phase behaviour of ionic liquid 1-butyl-1-methylpyrrolidinium tris(pentafluoroethyl)trifluorophosphate with alcohols, water and aromatic hydrocarbons:	https://www.doi.org/10.1016/j.fluid.2013.01.027
Impact of structural modification of 1,2,4-thiadiazole derivatives on thermodynamics of solubility and solvation processes in 1-octanol and n-hexane:	https://www.doi.org/10.1016/j.jct.2015.12.016
Excess molar enthalpies of binary systems containing 2-octanone + hexanoic acid, or octanoic acid at T = 298.15 K:	https://www.doi.org/10.1016/j.jct.2011.07.018
Measurement and Study on the Liquid Liquid Equilibrium of Ionic Liquids with Aliphatic Alcohols:	https://www.doi.org/10.1021/je101261z
Molar excess enthalpies at T = 298.15 K for (1-alkanol + dibutylether) systems:	https://www.doi.org/10.1016/j.jct.2009.07.001
Volumetric and transport properties of ternary mixtures containing 1-alkanol + ethyl ethanoate + cyclohexane at 303.15 K: Experimental data, correlation and prediction by ERAS model:	https://www.doi.org/10.1016/j.tca.2009.02.022
Atmospheric densities and interfacial tensions for 1-alkanol (1-butanol to 1-decanol) + water and ethers (MTBE, ETBE, DIPE, TAME and THF) + water demixed mixtures:	https://www.doi.org/10.1016/j.fluid.2015.03.040
PρT measurements and derived properties of liquid 1-alkanols:	https://www.doi.org/10.1016/j.jct.2011.10.023
Volumetric, acoustic, and viscometric studies of molecular interactions in binary mixtures of diethylene glycol monomethyl ether with 1-alkanols at temperatures from (293.15 to 308.15) K:	https://www.doi.org/10.1016/j.jct.2009.08.009
Impact of Sulfonamide Structure on Solubility and Transfer Processes in Biologically Relevant Solvents:	https://www.doi.org/10.1021/je500918t
Kirkwood-Buff Integrals, Excess Volume, and Preferential Solvation in Pentanoic Acid/(C5-C10) 1-Alkanol Binary Mixtures:	https://www.doi.org/10.1021/acs.jced.8b01189
Solubility and Solution Thermodynamics of Novel Bicyclic Derivatives of 1,3-Selenazine in Biological Relevant Solvents:	https://www.doi.org/10.1021/je500363r
McGowan Method:	http://link.springer.com/article/10.1007/BF02311772
Solid-liquid equilibrium, excess molar volume, and deviations in the molar refractivity for the binary and ternary mixtures of Alamine 304-1 with 1-octanol, 2-octanol, and 1-decanol:	https://www.doi.org/10.1016/j.fluid.2012.03.025

Thermogravimetric measurement of evaporation: Data analysis based on the Stefan tube:	https://www.doi.org/10.1016/j.tca.2014.09.011
Thermodynamic Studies of Solubility for Naphthalene in 12 Solvents from 279 to 330 K:	https://www.doi.org/10.1021/acs.jced.9b00064
Densities, Viscosities, and Ultrasonic Velocity Studies of Binary Mixtures of Chloroform with Octan-1-ol and Decan-1-ol at (303.15 and 313.15) K:	https://www.doi.org/10.1021/je060192z
Viscosity and Density of n-Alcohols at Temperatures between (298.15 and 323.15) K and Pressures up to 30 MPa:	https://www.doi.org/10.1021/acs.jced.8b00812
The solubility of ethionamide and structural analogues in buffer solutions, octanol and hexane at several temperatures:	https://www.doi.org/10.1016/j.tca.2016.07.020
Ternary equilibrium data of mixtures consisting of 2-butanol, water, and heavy alcohols at T = 298.2 K:	https://www.doi.org/10.1016/j.jct.2012.01.003
Density, Surface Tension, and Refractive Index of Octane + 1-Alkanol Mixtures at T) 298.15 K:	https://www.doi.org/10.1021/je034053i
Solubility and vapor pressure data of bioactive 6-(acetylamino)-N-(5-ethyl-1,3,4-thiadiazol-2-yl) hexanamide:	https://www.doi.org/10.1016/j.jct.2019.03.015
A thermodynamic study of sublimation, dissolution and distribution processes of anti-inflammatory drug Clonixin:	https://www.doi.org/10.1016/j.jct.2019.01.006
Low Pressure VLE Measurements and Thermodynamic Modeling, with PSRK and NRTL, of Binary 1-Alcohol + n-Alkane Systems:	https://www.doi.org/10.1021/acs.jced.8b00680
Crippen Method:	http://pubs.acs.org/doi/abs/10.1021/ci990307l
Comparison of the Efficiencies of Amine Extractants on Lactic Acid with Different Organic Solvents:	https://www.doi.org/10.1021/je900446d
Extraction of Citric Acid and Maleic Acid from Their Aqueous Solutions Using a Phosphorus-Bonded Extractant, Tri-n-octylphosphineoxide, and a Secondary Amine, Dioctylamine:	https://www.doi.org/10.1021/acs.jced.7b00562
Vapor Pressure and Isobaric Vapor-Liquid Equilibrium for Binary Systems of Furfural + 2-Acetylfuran and 5-Methylfurfural at 3.60 and 5.18 kPa:	https://www.doi.org/10.1021/acs.jced.7b00574
Densities and Viscosities of Three Binary Monoglyme + 1-Alcohol Systems from (283.15 to 313.15) K:	https://www.doi.org/10.1021/je301095j
Viscosities and Densities of Fatty Alcohol Mixtures from 298.15 to 338.15 K: Estimation by Kay's Rule and Prediction by the UNIFAC-VISCO and GC-UNIMOD Group Contribution Methods:	https://www.doi.org/10.1021/acs.jced.8b00992
Volumetric Properties, Viscosities, and Isobaric Heat Capacities of Imidazolium Octanoate Protic Ionic Liquid in Molecular Solvents:	https://www.doi.org/10.1021/je100671v
Investigation of Molecular Interactions in Binary Mixtures of n-Butyl Acetate and (C6−C10) 1-Alkanol: PC-SAFT Model:	https://www.doi.org/10.1021/acs.jced.8b00524
Measurement and prediction of pyrene solubility in pure, binary, ternary and quaternary solvent systems:	https://www.doi.org/10.1016/j.fluid.2007.11.001
Phase behaviour of tricyanomethanide-based ionic liquids with alcohols and hydrocarbons:	https://www.doi.org/10.1016/j.fluid.2014.12.003
Effect of temperature and chain length on the viscosity and surface tension of binary systems of N,N-dimethylformamide with 1-octanol, 1-nonanol and 1-decanol:	https://www.doi.org/10.1016/j.jct.2014.03.022
Speeds of Sound, Densities, Isobaric Thermal Expansion, Compressibilities, and Internal Pressures of Heptan-1-ol, Octan-1-ol, Nonan-1-ol, and Decan-1-ol at Temperatures from (293 to 318) K and Pressures up to 100 MPa:	https://www.doi.org/10.1021/je0604378
NIST Webbook:	http://webbook.nist.gov/cgi/cbook.cgi?ID=C111875&Units=SI
Experimental Study on Reactive Extraction of Malonic Acid with Validation by Fourier Transform Infrared Spectroscopy:	https://www.doi.org/10.1021/acs.jced.8b00972
Thermodynamic, transport and excess properties of 2-butoxy ethanol +1-alkanol (C6,C8,C10) at different temperatures:	https://www.doi.org/10.1016/j.jct.2013.04.021
An experimental investigation on the solubility of methane in 1-octanol and n-dodecane at ambient temperatures:	https://www.doi.org/10.1016/j.fluid.2016.08.009
Joback Method:	https://en.wikipedia.org/wiki/Joback_method
Thermophysical properties and thermodynamic phase behavior of ionic liquids:	https://www.doi.org/10.1016/j.tca.2006.06.018
Determination and Correlation for Solubilities of Ofloxacin, Norfloxacin, Lomefloxacin, Ciprofloxacin, Pefloxacin, and Pipemidic Acid in 1-Octanol from (293.15 to 333.15) K:	https://www.doi.org/10.1021/je100116u
Thermodynamics of mixtures containing amines. XVI. CEpm of benzylamine systems at (298.15, 308.15, 318.15 and 333.15) K:	https://www.doi.org/10.1016/j.tca.2014.10.018
Solubility of pharmaceuticals in water and alcohols:	https://www.doi.org/10.1016/j.fluid.2015.02.018
Phase behaviour of l-hexyloxymethyl-3-methyl-imidazolium and 1,3-dihexyloxymethyl-imidazolium based ionic liquids with alcohols, water, ketones and hydrocarbons: The effect of cation and anion on solubility:	https://www.doi.org/10.1016/j.fluid.2006.07.005
Correlation Studies of Cyclohexanone/(C5-C10) Alkan-1-ol Binary Mixtures: PC-SAFT and the Volume Theory:	https://www.doi.org/10.1021/acs.jced.8b00201
Phase equilibrium for ternary liquid systems of water + di-(2-ethylhexyl)phosphoric acid + organic diluents: Thermodynamic study:	https://www.doi.org/10.1016/j.fluid.2015.05.017
Thermodynamic and spectral studies of molecular interactions in binary liquid mixtures of 1-butoxy-2-propanol with 1-alcohols:	https://www.doi.org/10.1021/acs.jced.5b00031
Solid-Liquid Equilibrium Solubility, Thermodynamic Properties, and Molecular Simulation of Phenylphosphonic Acid in 13 Pure Solvents at Different Temperatures:	https://www.doi.org/10.1021/acs.jced.9b00362
Comparison Study on Temperature Dependence of the Interfacial Tension of n-Alkane/Water and n-Alcohol/Water Two Binary Systems:	https://www.doi.org/10.1021/je500450w
Measurement and Correlation for Solubility of Dexibuprofen in Different Solvents from (283.15 to 293.15) K:	https://www.doi.org/10.1016/j.tca.2012.04.016
Micellar Properties and Related Thermodynamic Parameters of Aqueous Anionic Surfactants in the Presence of Monohydric Alcohols :	https://www.doi.org/10.1021/je101358p
Vapor Liquid Equilibrium between (278.15 and 323.15) K and Excess Functions at T) 298.15 K for 1-Bromobutane with 1-Octanol or 1-Decanol:	https://www.doi.org/10.1021/je030168a
Solubility Measurement and Phase Equilibrium Modeling of 2-Bromobenzamide in 15 Pure Solvents:	https://www.doi.org/10.1021/acs.jced.9b00350
Volumetric and Viscometric Properties of Binary Liquid Mixtures of Ethylene Glycol Monomethyl Ether + 1-Hexanol, 1-Octanol, and 1-Decanol at Temperatures of T = (293.15, 298.15, 303.15, and 308.15) K:	https://www.doi.org/10.1021/je9005344
The solid-liquid phase diagrams of binary mixtures of even saturated fatty alcohols:	https://www.doi.org/10.1016/j.fluid.2011.01.028

Experimental Solid + Liquid Equilibria and Excess Molar Volume of Alkanol + Octylamine Mixtures. Analysis in Terms of ERAS, DISQUAC, and Modified UNIFAC:	https://www.doi.org/10.1021/je0301895
Determination and prediction of solubilities of active pharmaceutical ingredients in selected organic solvents:	https://www.doi.org/10.1016/j.fluid.2015.07.032
Deduction of Physicochemical Properties from Solubilities: 2,4-Dihydroxybenzophenone, Biotin, and Caprolactam as Examples:	https://www.doi.org/10.1021/je501140p
Excess molar volumes and viscosities of (1,1,2,2-tetrabromoethane + 1-alkanols) at T = (293.15 and 303.15) K:	https://www.doi.org/10.1016/j.jct.2005.03.027
Solid-Liquid Equilibria, Excess Molar Volumes, and Molar Refractivity Deviations for Extractive Solvents of Molybdenum:	https://www.doi.org/10.1021/je900586f
Solubility of Rifapentine in the Binary System of Acetic Acid and n-Octanol Solvent Mixtures:	https://www.doi.org/10.1021/je8002578
Experimental solubility of diosgenin and estriol in various solvents between T = (293.2-328.2) K:	https://www.doi.org/10.1016/j.jct.2016.11.017
Solubility and Transfer Processes of Some Hydrazones in Biologically Relevant Solvents:	https://www.doi.org/10.1021/je400557z
(Liquid + liquid) equilibria of (water + propionic acid + alcohol) ternary systems:	https://www.doi.org/10.1016/j.jct.2006.05.001
Excess molar volumes and enthalpies of mixing for the ternary system (butyl butyrate + 1-octanol + dodecane) at the temperature 308.15 K.:	https://www.doi.org/10.1021/je020138s
KDB Vapor Pressure Data:	https://www.cheric.org/research/kdb/hcprop/showprop.php?cmpid=841
The role of organic diluents in the aspects of equilibrium, kinetics and thermodynamic model for silver ion extraction using an extractant D2EHPA:	https://www.doi.org/10.1016/j.fluid.2014.12.033
(Liquid + liquid) equilibria of (water + lactic acid + alcohol) ternary systems:	https://www.doi.org/10.1016/j.jct.2008.07.014
Experimental solid-liquid phase equilibria of {cholesterol + binary solvent mixture: 1-Alcohol (C4-C10) + cyclohexane}:	https://www.doi.org/10.1016/j.fluid.2009.10.009
Equilibrium Study on the Extraction of Levulinic Acid from Aqueous Solution with Aliquat 336 Dissolved in Different Diluents: Solvent's Polarity Effect and Column Design,	https://www.doi.org/10.1021/acs.jced.6b00164
Study of the physicochemical properties of protein kinase CK2 inhibitors: TBBi and 2-Me-TBBi:	https://www.doi.org/10.1016/j.fluid.2018.09.015
Solubility measurement and thermodynamic model correlation of sancycline in twelve pure solvents:	https://www.doi.org/10.1021/acs.jced.9b00693
Isobaric vapor-liquid equilibrium for binary systems containing benzothiophene:	https://www.doi.org/10.1016/j.fluid.2011.04.024
Thermodynamic Properties of 1-Hexyl-3-methylimidazolium Nitrate and 1-Alkanols Mixtures: PC-SAFT Model:	https://www.doi.org/10.1021/acs.jced.9b00507
Viscosities and Interfacial Properties of 1-Methyl-3-butylimidazolium Hexafluorophosphate and 1-Isobutenyl-3-methylimidazolium Tetrafluoroborate Ionic Liquids:	https://www.doi.org/10.1021/je0604277
Solubility and pKa of select pharmaceuticals in water, ethanol, and 1-octanol:	https://www.doi.org/10.1016/j.jct.2010.07.001
Equilibrium data and GC-PC SAFT predictions for furanic extraction:	https://www.doi.org/10.1016/j.fluid.2016.09.019
Densities and Excess Molar Volumes of N-Methylmorpholine + 1-Alkanol Systems at 298.15 K:	https://www.doi.org/10.1021/je7005055
Experimental solubility for betulin and estrone in various solvents within the temperature range T = (293.2 to 328.2) K:	https://www.doi.org/10.1016/j.jct.2016.02.006
Thermodynamics of Dissolution for Bis(triazine)-Bipyridine-Class Ligands in Different Diluents and Its Reflection on Extraction:	https://www.doi.org/10.1021/je1005246
Experimental Study of Chemical Equilibria in the Liquid-Phase Reaction between 1-Octanol and Ethanol to 1-Ethoxyoctane:	https://www.doi.org/10.1021/je400291y
Measurements and equation-of-state modelling of thermodynamic properties of binary mixtures of 1-butyl-1-methylpyrrolidinium tetracyanoborate ionic liquid with molecular compounds:	https://www.doi.org/10.1016/j.jct.2015.06.004
High-Pressure Phase Equilibria of Carbon Dioxide + 1-Octanol Binary System:	https://www.doi.org/10.1021/acs.jced.7b00865
Solubilities of Benzoic Acid, p-Methylbenzoic Acid, m-Methylbenzoic Acid, o-Methylbenzoic Acid, p-Hydroxybenzoic Acid, and o-Nitrobenzoic Acid in 1-Octanol:	https://www.doi.org/10.1021/je700677d
Thermodynamics of mixing for binary mixtures of 1-octanol and 1-decanol with n-dodecane and ternary mixture of {TBP + 1-octanol + dodecane} at T = (298.15 to 323.15) K:	https://www.doi.org/10.1016/j.jct.2015.06.029
Thermophysical properties and phase equilibria study of the binary systems {N-hexylquinolinium bis(trifluoromethylsulfonyl)imide + aromatic hydrocarbons, or an alcohol}:	https://www.doi.org/10.1016/j.jct.2010.12.021
Solubility and limiting activity coefficient of simvastatin in different organic solvents:	https://www.doi.org/10.1016/j.fluid.2009.03.006
Liquid-liquid equilibria for ternary systems of (water + carboxylic acid + 1-octanol) at 293.15 K: modeling phase equilibria using a solvatochromic approach:	https://www.doi.org/10.1016/j.fluid.2004.10.029
Solubility of d-Tryptophan and L-Tyrosine in Several Organic Solvents: Determination and Solvent Effect:	https://www.doi.org/10.1021/acs.jced.9b00258
Reactive Extraction of Gallic Acid Using Tributyl Phosphate in Different Classes of Diluents:	https://www.doi.org/10.1021/acs.jced.9b00192
Liquid phase equilibria of the system (water + phosphoric acid + 1-octanol) at T = (298.2, 308.2, and 318.2) K:	https://www.doi.org/10.1016/j.fluid.2011.12.016
Density and Viscosity of the Binary Systems Ethanol + Butan-1-ol, + Pentan-1-ol, + Heptan-1-ol, + Octan-1-ol, Nonan-1-ol, + Decan-1-ol at 0.1 MPa and Temperatures from 283.15 K to 313.15 K:	https://www.doi.org/10.1021/je400630f
KDB:	https://www.cheric.org/files/research/kdb/mol/mol841.mol
Density and Viscosity of Binary Mixtures of {1-Butyl-3-methylimidazolium Thiocyanate + 1-Heptanol, 1-Octanol, 1-Nonanol, or 1-Decanol}:	https://www.doi.org/10.1021/je901043q
FTIR studies and excess thermodynamic properties of binary liquid mixtures of 2-(2-butoxyethoxy) ethanol with 1-hexanol, 1-octanol and 1-decanol at different temperatures:	https://www.doi.org/10.1016/j.jct.2018.06.015
Aqueous Solubility Prediction Method:	http://onschallenge.wikispaces.com/file/view/AqueousDataset002.xlsx/351826032/AqueousDa
Studies on the importance of chain length of alkanols on the thermodynamic and transport properties of liquid mixtures at various temperatures:	https://www.doi.org/10.1016/j.jct.2016.12.019
Physico-chemical study of bioactive N-(5-ethyl-1,3,4-thiadiazole-2-yl)-4-nitrobenzamide: Sublimation, solubility and distribution:	https://www.doi.org/10.1016/j.tca.2017.09.010

Pressure-Volume-Temperature Properties for 1-Octanol + Acetophenone, Poly(Propylene Glycol) + 1-Octanol + Acetophenone, and Poly(Ethylene Glycol) + Poly(Ethylene Glycol) Methyl Ether + Anisole:	https://www.doi.org/10.1021/je049930e
Volumetric and compressibility studies on tri-n-butyl phosphate (TBP) phase modifier (1-octanol/1-decanol) association interaction (298.15 to 323.15) K:	https://www.doi.org/10.1016/j.jct.2013.10.018
Thermophysical Properties of an Insensitive Munitions Compound, 2,4-Dinitroanisole:	https://www.doi.org/10.1021/je7006764
Phase Equilibria Study of the Binary Systems (N-Butyl-4-methylpyridinium Tosylate Ionic Liquid + Organic Solvent, or Water):	https://www.doi.org/10.1021/je8006309
Solubility and Thermodynamic Modeling of Sulfanilamide in 12 Mono Solvents and 4 Binary Solvent Mixtures from 278.15 to 318.15 K:	https://www.doi.org/10.1021/acs.jced.9b00411
L(+)-tartaric Acid Separations Using Aliquat 336 in n-Heptane, Kerosene, and 1-Octanol at 300 +- 1 K:	https://www.doi.org/10.1021/acs.jced.6b01070
Experiment and Computation of Solubility and Dissolution Properties for Enalapril Maleate and its Intermediate in Pure Solvents:	https://www.doi.org/10.1021/acs.jced.8b00333
Thermodynamic study of aceclofenac solubility, distribution and sublimation:	https://www.doi.org/10.1016/j.jct.2019.05.014
Vapor Pressure Determination of the Aliphatic C5 to C8 1-Alcohols:	https://www.doi.org/10.1021/je049600u
Liquid Liquid Phase Behavior of Solutions of 1-Butyl-3-methylimidazolium Bis((trifluoromethyl)sulfonyl)amide (C4mimNTf2) in n-Alkyl Alcohols:	https://www.doi.org/10.1021/je200792v
Thermodynamic Aspects of Solubility and Solvation of Bioactive Bicyclic Derivatives in Organic Solvents:	https://www.doi.org/10.1021/acs.jced.7b00641

Legend

af:	Acentric Factor
affp:	Proton affinity
chl:	Standard liquid enthalpy of combustion
cpg:	Ideal gas heat capacity
cpl:	Liquid phase heat capacity
dm:	Dipole Moment
dvisc:	Dynamic viscosity
gf:	Standard Gibbs free energy of formation
hf:	Enthalpy of formation at standard conditions
hfl:	Liquid phase enthalpy of formation at standard conditions
hfus:	Enthalpy of fusion at standard conditions
hfust:	Enthalpy of fusion at a given temperature
hsub:	Enthalpy of sublimation at standard conditions
hvap:	Enthalpy of vaporization at standard conditions
hvapt:	Enthalpy of vaporization at a given temperature
log10ws:	Log10 of Water solubility in mol/l
logp:	Octanol/Water partition coefficient
mcvol:	McGowan's characteristic volume
nfpaf:	NFPA Fire Rating
nfpah:	NFPA Health Rating
pc:	Critical Pressure
pvap:	Vapor pressure
rfi:	Refractive Index
rhoc:	Critical density
rhol:	Liquid Density
rinpol:	Non-polar retention indices
ripol:	Polar retention indices

speedsl:	Speed of sound in fluid
srf:	Surface Tension
tb:	Normal Boiling Point Temperature
tbrp:	Boiling point at reduced pressure
tc:	Critical Temperature
tcondl:	Liquid thermal conductivity
tf:	Normal melting (fusion) point
vc:	Critical Volume
zc:	Critical Compressibility