

# 2,6,10,14,18-Pentamethyl-7-(3-methyl-pentyl)-nonane

<b>Inchi:</b>	InChI=1S/C30H62/c1-10-26(6)20-22-30(29(9)19-12-15-25(4)5)23-21-28(8)18-13-17-27(7)
<b>InchiKey:</b>	ZYUQNRLZGGWDMML-UHFFFAOYSA-N
<b>Formula:</b>	C30H62
<b>SMILES:</b>	CCC(C)CCC(CCC(C)CCCC(C)CCCC(C)C)C(C)CCCC(C)C
<b>Mol. weight [g/mol]:</b>	422.81

## Physical Properties

Property code	Value	Unit	Source
gf	184.64	kJ/mol	Joback Method
hf	-699.49	kJ/mol	Joback Method
hfus	48.80	kJ/mol	Joback Method
hvap	79.66	kJ/mol	Joback Method
log10ws	-10.69		Crippen Method
logp	10.940		Crippen Method
mcvol	433.560	ml/mol	McGowan Method
pc	612.08	kPa	Joback Method
rinpol	2525.00		NIST Webbook
rinpol	2525.00		NIST Webbook
tb	882.72	K	Joback Method
tc	1081.45	K	Joback Method
tf	322.86	K	Joback Method
vc	1.673	m <sup>3</sup> /kmol	Joback Method

## Temperature Dependent Properties

Property code	Value	Unit	Temperature [K]	Source
cpg	1484.45	J/molxK	882.72	Joback Method
cpg	1601.40	J/molxK	1048.33	Joback Method
cpg	1580.74	J/molxK	1015.20	Joback Method
cpg	1558.79	J/molxK	982.08	Joback Method
cpg	1535.47	J/molxK	948.96	Joback Method
cpg	1510.72	J/molxK	915.84	Joback Method
cpg	1620.84	J/molxK	1081.45	Joback Method
dvisc	0.0000130	Paxs	882.72	Joback Method

dvisc	0.0000205	Paxs	789.41	Joback Method
dvisc	0.0000365	Paxs	696.10	Joback Method
dvisc	0.0000777	Paxs	602.79	Joback Method
dvisc	0.0002183	Paxs	509.48	Joback Method
dvisc	0.0009747	Paxs	416.17	Joback Method
dvisc	0.0103350	Paxs	322.86	Joback Method

## Sources

<b>NIST Webbook:</b>	<a href="http://webbook.nist.gov/cgi/cbook.cgi?ID=R495919&amp;Units=SI">http://webbook.nist.gov/cgi/cbook.cgi?ID=R495919&amp;Units=SI</a>
<b>Crippen Method:</b>	<a href="http://pubs.acs.org/doi/abs/10.1021/ci9903071">http://pubs.acs.org/doi/abs/10.1021/ci9903071</a>
<b>Crippen Method:</b>	<a href="https://www.chemeo.com/doc/models/crippen_log10ws">https://www.chemeo.com/doc/models/crippen_log10ws</a>
<b>Joback Method:</b>	<a href="https://en.wikipedia.org/wiki/Joback_method">https://en.wikipedia.org/wiki/Joback_method</a>
<b>McGowan Method:</b>	<a href="http://link.springer.com/article/10.1007/BF02311772">http://link.springer.com/article/10.1007/BF02311772</a>

## Legend

<b>cp<sub>g</sub>:</b>	Ideal gas heat capacity
<b>dvisc:</b>	Dynamic viscosity
<b>gf:</b>	Standard Gibbs free energy of formation
<b>hf:</b>	Enthalpy of formation at standard conditions
<b>hfus:</b>	Enthalpy of fusion at standard conditions
<b>hvap:</b>	Enthalpy of vaporization at standard conditions
<b>log<sub>10</sub>ws:</b>	Log <sub>10</sub> of Water solubility in mol/l
<b>log<sub>p</sub>:</b>	Octanol/Water partition coefficient
<b>m<sub>c</sub>vol:</b>	McGowan's characteristic volume
<b>pc:</b>	Critical Pressure
<b>rin<sub>pol</sub>:</b>	Non-polar retention indices
<b>tb:</b>	Normal Boiling Point Temperature
<b>tc:</b>	Critical Temperature
<b>tf:</b>	Normal melting (fusion) point
<b>vc:</b>	Critical Volume

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