Acetone

Other names:	(CH3)2CO
	2-PROPANONE
	Chevron acetone
	DIMETHYL KETONE
	Dimethylformaldehyde
	Dimethylketal
	KETONE PROPANE
	Ketone, dimethyl-
	Methyl ketone
	NSC 135802
	Propan-2-one
	Propanone
	Pyroacetic ether
	Rcra waste number U002
	Sasetone
	UN 1090
	«beta»-Ketopropane
	Â«betaÂ»-Ketopropane
Inchi:	InChI=1S/C3H6O/c1-3(2)4/h1-2H3
InchiKey:	CSCPPACGZOOCGX-UHFFFAOYSA-N
Formula:	C3H6O
SMILES:	CC(C)=O
Mol. weight [g/mol]:	58.08
CAS:	67-64-1

Physical Properties

Property code	Value	Unit	Source
af	0.3040		KDB
affp	812.00	kJ/mol	NIST Webbook
affp	814.30	kJ/mol	NIST Webbook
affp	812.60 ± 0.20	kJ/mol	NIST Webbook
affp	811.50 ± 3.40	kJ/mol	NIST Webbook
affp	811.50 ± 3.40	kJ/mol	NIST Webbook
affp	815.20	kJ/mol	NIST Webbook
aigt	738.15	K	KDB
basg	782.10 ± 1.50	kJ/mol	NIST Webbook
basg	782.10	kJ/mol	NIST Webbook

basg	784.70	kJ/mol	NIST Webbook
basg	782.20	kJ/mol	NIST Webbook
basg	782.00 ± 0.20	kJ/mol	NIST Webbook
basg	782.10 ± 1.50	kJ/mol	NIST Webbook
chg	-1821.40 ± 0.84	kJ/mol	NIST Webbook
chl	-1804.20	kJ/mol	NIST Webbook
chl	-1772.00	kJ/mol	NIST Webbook
dm	2.90	debye	KDB
dvisc	0.0003121	Pa×s	Densities and Viscosities of Binary Liquid Mixtures of Trichloroethylene and Tetrachloroethylene with Some Polar and Nonpolar Solvents
dvisc	0.0003000	Pa×s	Densities and Viscosities of 1-Butyl-3-methylimidazolium Tetrafluoroborate + Molecular Solvent Binary Mixtures
ea	0.00	eV	NIST Webbook
fll	2.60	% in Air	KDB
flu	12.80	% in Air	KDB
fpc	257.59	K	KDB
fpo	255.37	K	KDB
gf	-153.20	kJ/mol	KDB
gyrad	2.7400		KDB
hf	-217.70	kJ/mol	KDB
hf	-217.10 ± 0.50	kJ/mol	NIST Webbook
hf	-216.40	kJ/mol	NIST Webbook
hf	-217.50 ± 0.67	kJ/mol	NIST Webbook
hf	-218.50 ± 0.59	kJ/mol	NIST Webbook
hfl	-249.40 ± 0.63	kJ/mol	NIST Webbook
hfus	5.12	kJ/mol	Joback Method
hvap	31.30	kJ/mol	NIST Webbook
hvap	31.27	kJ/mol	NIST Webbook
hvap	29.70 ± 0.00	kJ/mol	NIST Webbook
ie	9.74	eV	NIST Webbook
ie	9.72	eV	NIST Webbook
ie	9.50	eV	NIST Webbook
ie	9.71	eV	NIST Webbook
ie	9.70	eV	NIST Webbook
ie	9.71 ± 0.03	eV	NIST Webbook
ie	9.71 ± 0.03	eV	NIST Webbook
ie	9.67	eV	NIST Webbook
ie	9.68 ± 0.02	eV	NIST Webbook
ie	9.70 ± 0.10	eV	NIST Webbook

ie	9.71 ± 0.01	eV	NIST Webbook
ie	9.71	eV	NIST Webbook
ie	9.74 ± 0.03	eV	NIST Webbook
ie	9.71 ± 0.01	eV	NIST Webbook
ie	9.71 ± 0.02	eV	NIST Webbook
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ie	9.68	eV	NIST Webbook
ie	9.68	eV	NIST Webbook
ie	9.71	eV	NIST Webbook
log10ws	-0.36		Crippen Method
logp	0.595		Crippen Method
mcvol	54.700	ml/mol	McGowan Method
nfpaf	%!d(float64=3)		KDB
nfpah	%!d(float64=1)		KDB
pc	4700.00	kPa	KDB
rhoc	252.06 ± 1.74	kg/m3	NIST Webbook
rhoc	268.91 ± 2.90	kg/m3	NIST Webbook
rhoc	278.20 ± 9.87	kg/m3	NIST Webbook
rhoc	272.97 ± 1.16	kg/m3	NIST Webbook
rhoc	234.06 ± 1.51	kg/m3	NIST Webbook
rinpol	500.00		NIST Webbook
rinpol	503.00		NIST Webbook
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rinpol	478.00		NIST Webbook
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tf	178.30	K	KDB
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tt	178.50 ± 0.30	K	NIST Webbook
tt	177.60 ± 0.30	K	NIST Webbook
tt	177.60 ± 0.20	K	NIST Webbook
vc	0.209	m3/kmol	KDB
zc	0.2325190		KDB
zra	0.25		KDB

Temperature Dependent Properties

Property code	Value	Unit	Temperature [K]	Source
cpg	98.66 ± 0.20	J/mol×K	439.20	NIST Webbook
cpg	100.50 ± 2.00	J/mol×K	438.00	NIST Webbook
cpg	99.40 ± 2.00	J/mol×K	428.00	NIST Webbook
cpg	96.80 ± 1.90	J/mol×K	422.60	NIST Webbook
cpg	93.30	J/mol×K	410.00	NIST Webbook
cpg	94.18 ± 0.94	J/mol×K	408.00	NIST Webbook
cpg	92.93 ± 0.19	J/mol×K	405.20	NIST Webbook
cpg	91.84 ± 0.92	J/mol×K	393.00	NIST Webbook
cpg	89.24 ± 0.89	J/mol×K	378.00	NIST Webbook
cpg	87.53 ± 0.88	J/mol×K	372.30	NIST Webbook
cpg	87.19 ± 0.17	J/mol×K	371.20	NIST Webbook
cpg	87.03 ± 0.87	J/mol×K	363.00	NIST Webbook
cpg	83.39 ± 0.83	J/mol×K	348.00	NIST Webbook
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cpg	80.58 ± 0.81	J/mol×K	332.60	NIST Webbook
cpl	121.30	J/mol×K	283.00	NIST Webbook
cpl	125.90	J/mol×K	293.20	NIST Webbook
cpl	133.90	J/mol×K	298.00	NIST Webbook
cpl	124.70	J/mol×K	289.40	NIST Webbook
cpl	123.80	J/mol×K	298.40	NIST Webbook
cpl	124.30	J/mol×K	260.00	NIST Webbook
cpl	124.68	J/mol×K	296.99	NIST Webbook
cpl	124.70	J/mol×K	298.00	NIST Webbook
cpl	124.70	J/mol×K	298.00	NIST Webbook
cpl	128.40	J/mol×K	302.40	NIST Webbook
cpl	128.24	J/mol×K	298.00	NIST Webbook
cpl	125.56	J/mol×K	298.20	NIST Webbook
cpl	126.30	J/mol×K	293.00	NIST Webbook
cpl	129.70	J/mol×K	298.00	NIST Webbook
cpl	125.90	J/mol×K	298.15	NIST Webbook
cpl	123.80	J/mol×K	298.15	NIST Webbook
cpl	123.80	J/mol×K	298.15	NIST Webbook
cpl	126.60	J/mol×K	298.15	NIST Webbook
cpl	126.60	J/mol×K	298.15	NIST Webbook
cpl	125.45	J/mol×K	298.15	NIST Webbook
cpl	123.80	J/mol×K	298.15	NIST Webbook
cps	96.00	J/mol×K	173.00	NIST Webbook

dvisc	0.0002839	Pa×s	313.15	Density and Viscosity of (2,2-Dichloro-N,N-di-2-propenylacetamide + Acetone) and (2,2-Dichloro-N,N-di-2-propenylacetamide + Ethanol) at T = (278.15 to 313.15) K
dvisc	0.0003060	Pa×s	298.15	Physical properties of the binary systems methylcyclopentane with ketones (acetone, butanone and 2-pentanone) at T = (293.15, 298.15, and 303.15) K. New UNIFAC-VISCO interaction parameters
dvisc	0.0002809	Pa×s	308.15	Excess parameter studies on the binary mixtures of toluene with ketones at different temperatures
dvisc	0.0003200	Pa×s	293.15	Dynamic Viscosities of the Binary Systems Cyclohexane and Cyclopentane with Acetone, Butanone, or 2-Pentanone at Three Temperatures T) (293.15, 298.15, and 303.15) K
dvisc	0.0003060	Pa×s	298.15	Dynamic Viscosities of the Binary Systems Cyclohexane and Cyclopentane with Acetone, Butanone, or 2-Pentanone at Three Temperatures T) (293.15, 298.15, and 303.15) K
dvisc	0.0003027	Pa×s	298.15	Excess parameter studies on the binary mixtures of toluene with ketones at different temperatures

dvisc	0.0002920	Pa×s	303.15	Dynamic Viscosities of the Binary Systems Cyclohexane and Cyclopentane with Acetone, Butanone, or 2-Pentanone at Three Temperatures T) (293.15, 298.15, and 303.15) K
dvisc	0.0003027	Pa×s	298.15	Density and Viscosity of Ketones with Toluene at Different Temperatures and at Atmospheric Pressure
dvisc	0.0002918	Pa×s	303.15	Density and Viscosity of Ketones with Toluene at Different Temperatures and at Atmospheric Pressure
dvisc	0.0002920	Pa×s	303.15	Physical properties of the binary systems methylcyclopentane with ketones (acetone, butanone and 2-pentanone) at T = (293.15, 298.15, and 303.15) K. New UNIFAC-VISCO interaction parameters
dvisc	0.0003849	Pa×s	278.15	Density and Viscosity of (2,2-Dichloro-N,N-di-2-propenylacetamide + Acetone) and (2,2-Dichloro-N,N-di-2-propenylacetamide + Ethanol) at T = (278.15 to 313.15) K
dvisc	0.0003657	Pa×s	283.15	Density and Viscosity of (2,2-Dichloro-N,N-di-2-propenylacetamide + Acetone) and (2,2-Dichloro-N,N-di-2-propenylacetamide + Ethanol) at T = (278.15 to 313.15) K

dvisc	0.0003511	Pa×s	288.15	Density and Viscosity of (2,2-Dichloro-N,N-di-2-propenylacetamide + Acetone) and (2,2-Dichloro-N,N-di-2-propenylacetamide + Ethanol) at T = (278.15 to 313.15) K
dvisc	0.0003339	Pa×s	293.15	Density and Viscosity of (2,2-Dichloro-N,N-di-2-propenylacetamide + Acetone) and (2,2-Dichloro-N,N-di-2-propenylacetamide + Ethanol) at T = (278.15 to 313.15) K
dvisc	0.0003212	Pa×s	298.15	Density and Viscosity of (2,2-Dichloro-N,N-di-2-propenylacetamide + Acetone) and (2,2-Dichloro-N,N-di-2-propenylacetamide + Ethanol) at T = (278.15 to 313.15) K
dvisc	0.0003200	Pa×s	293.15	Physical properties of the binary systems methylcyclopentane with ketones (acetone, butanone and 2-pentanone) at T = (293.15, 298.15, and 303.15) K. New UNIFAC-VISCO interaction parameters
dvisc	0.0002809	Pa×s	308.15	Density and Viscosity of Ketones with Toluene at Different Temperatures and at Atmospheric Pressure
dvisc	0.0003086	Pa×s	303.15	Density and Viscosity of (2,2-Dichloro-N,N-di-2-propenylacetamide + Acetone) and (2,2-Dichloro-N,N-di-2-propenylacetamide + Ethanol) at T = (278.15 to 313.15) K
dvisc	0.0002939	Pa×s	308.15	Density and Viscosity of (2,2-Dichloro-N,N-di-2-propenylacetamide + Acetone) and (2,2-Dichloro-N,N-di-2-propenylacetamide + Ethanol) at T = (278.15 to 313.15) K

dvisc	0.0002918	Pa×s	303.15	Excess parameter studies on the binary mixtures of toluene with ketones at different temperatures
hfust	5.72	kJ/mol	176.60	NIST Webbook
hfust	5.71	kJ/mol	176.62	NIST Webbook
hfust	5.69	kJ/mol	177.60	NIST Webbook
hfust	4.77	kJ/mol	178.50	NIST Webbook
hfust	5.69	kJ/mol	177.60	NIST Webbook
hfust	5.72	kJ/mol	176.60	NIST Webbook
hvapt	29.12	kJ/mol	329.40	KDB
hvapt	29.70	kJ/mol	482.50	NIST Webbook
hvapt	32.10	kJ/mol	271.50	NIST Webbook
hvapt	32.10	kJ/mol	308.00	NIST Webbook
hvapt	32.70	kJ/mol	294.50	NIST Webbook
hvapt	31.10	kJ/mol	319.50	NIST Webbook
hvapt	31.90	kJ/mol	307.00	NIST Webbook
hvapt	26.10	kJ/mol	373.00	NIST Webbook
hvapt	29.50	kJ/mol	419.00	NIST Webbook
hvapt	21.70	kJ/mol	423.00	NIST Webbook
hvapt	30.60	kJ/mol	351.00	NIST Webbook
hvapt	33.80	kJ/mol	236.00	NIST Webbook
hvapt	30.70	kJ/mol	271.50	NIST Webbook
hvapt	29.10	kJ/mol	329.30	NIST Webbook
hvapt	15.30	kJ/mol	473.00	NIST Webbook
hvapt	9.20	kJ/mol	498.00	NIST Webbook
hvapt	31.80	kJ/mol	319.00	NIST Webbook
hvapt	32.60	kJ/mol	285.50	NIST Webbook
hvapt	29.09	kJ/mol	338.00	NIST Webbook
hvapt	29.90	kJ/mol	408.50	NIST Webbook
hvapt	35.00	kJ/mol	271.50	NIST Webbook
hvapt	32.90	kJ/mol	210.50	NIST Webbook
hvapt	32.80	kJ/mol	305.00	NIST Webbook
pvap	70.07	kPa	318.77	Vapor Liquid Equilibrium for Binary Systems of 2,3-Pentanedione with Diacetyl and Acetone
pvap	57.96	kPa	314.00	Isobaric vapor liquid equilibria for acetone + methanol + lithium nitrate at 100 kPa

pvap	60.29	kPa	314.76	Vapor Liquid Equilibrium for Binary Systems of 2,3-Pentanedione with Diacetyl and Acetone
pvap	55.29	kPa	312.52	Vapor Liquid Equilibrium for Binary Systems of 2,3-Pentanedione with Diacetyl and Acetone
pvap	50.40	kPa	310.08	Vapor Liquid Equilibrium for Binary Systems of 2,3-Pentanedione with Diacetyl and Acetone
pvap	45.21	kPa	307.33	Vapor Liquid Equilibrium for Binary Systems of 2,3-Pentanedione with Diacetyl and Acetone
pvap	35.22	kPa	301.25	Vapor Liquid Equilibrium for Binary Systems of 2,3-Pentanedione with Diacetyl and Acetone
pvap	101.30	kPa	329.37	Vapor Liquid Equilibrium for the 1,1,1-Trifluorotrichloroethane + Sulfuryl Chloride System at 101.3 kPa
pvap	30.23	kPa	297.67	Vapor Liquid Equilibrium for Binary Systems of 2,3-Pentanedione with Diacetyl and Acetone
pvap	100.00	kPa	329.00	1-Ethyl-3-methylimidazolium Dicyanamide as a Very Efficient Entrainer for the Extractive Distillation of the Acetone + Methanol System

pvap	100.00	kPa	329.00	Influence of Some Ionic Liquids Containing the Trifluoromethanesulfonate Anion on the Vapor Liquid Equilibria of the Acetone + Methanol System
pvap	81.65	kPa	323.15	Vapor-Liquid Equilibrium for Binary Systems of Diacetyl with Methanol and Acetone
pvap	70.07	kPa	318.77	Vapor-Liquid Equilibrium for Binary Systems of Diacetyl with Methanol and Acetone
pvap	65.08	kPa	316.80	Vapor-Liquid Equilibrium for Binary Systems of Diacetyl with Methanol and Acetone
pvap	60.29	kPa	314.76	Vapor-Liquid Equilibrium for Binary Systems of Diacetyl with Methanol and Acetone
pvap	56.64	kPa	313.15	Vapor-Liquid Equilibrium for Binary Systems of Diacetyl with Methanol and Acetone
pvap	55.29	kPa	312.52	Vapor-Liquid Equilibrium for Binary Systems of Diacetyl with Methanol and Acetone
pvap	50.40	kPa	310.08	Vapor-Liquid Equilibrium for Binary Systems of Diacetyl with Methanol and Acetone
pvap	45.21	kPa	307.33	Vapor-Liquid Equilibrium for Binary Systems of Diacetyl with Methanol and Acetone

pvap	40.00	kPa	304.37	Vapor-Liquid Equilibrium for Binary Systems of Diacetyl with Methanol and Acetone
pvap	38.04	kPa	303.15	Vapor-Liquid Equilibrium for Binary Systems of Diacetyl with Methanol and Acetone
pvap	35.22	kPa	301.25	Vapor-Liquid Equilibrium for Binary Systems of Diacetyl with Methanol and Acetone
pvap	30.23	kPa	297.66	Vapor-Liquid Equilibrium for Binary Systems of Diacetyl with Methanol and Acetone
pvap	302.56	kPa	365.43	Vapor Liquid Equilibrium for Six Binary Systems of C4-Hydrocarbons + 2-Propanone
pvap	296.80	kPa	364.52	Vapor Liquid Equilibrium for Six Binary Systems of C4-Hydrocarbons + 2-Propanone
pvap	296.40	kPa	364.51	Vapor Liquid Equilibrium for Six Binary Systems of C4-Hydrocarbons + 2-Propanone
pvap	294.30	kPa	364.51	Vapor Liquid Equilibrium for Six Binary Systems of C4-Hydrocarbons + 2-Propanone
pvap	297.10	kPa	364.51	Vapor Liquid Equilibrium for Six Binary Systems of C4-Hydrocarbons + 2-Propanone
pvap	101.00	kPa	329.35	Isobaric Vapor-Liquid Equilibrium of Acetone + Methanol System in the Presence of Calcium Bromide

pvap	101.30	kPa	329.33	Isobaric Vapor-Liquid Equilibrium Data for the Acetone + Hexamethyl Disiloxane + Ethyl Acetate Ternary System at 101.3 kPa: Determination and Correlation
pvap	68.20	kPa	318.15	Isothermal Vapor-Liquid Equilibria for Binary Mixtures of Methyl Nonafluorobutyl Ether + Acetone, Cyclopentyl Methyl Ether, Ethyl Acetate, n-Heptane, Methanol, and Toluene
pvap	96.15	kPa	327.90	Vapor Liquid Equilibrium Data for Binary Mixtures of Acetic Acid + Anisole, Acetone + Anisole, and Isopropanol + Anisole at Pressure 96.15 kPa
pvap	115.70	kPa	333.15	Thermodynamics of binary mixtures of N-methyl-2-pyrrolidinone and ketone. Experimental results and modelling of the solid-liquid equilibrium and vapou-liquid equilibrium. The Modified UNIFAC (Do) model characterization
pvap	101.30	kPa	329.26	Isobaric vapor-liquid equilibrium for acetone + methanol system containing different ionic liquids at 101.3 kPa

pvap	101.33	kPa	329.30	Vapor-liquid equilibrium in the production of the ionic liquid, 1-hexyl-3-methylimidazolium bromide ([HMIm][Br]), in acetone
pvap	135.11	kPa	338.00	Isobaric vapor liquid equilibria for acetone + methanol + lithium nitrate at 100 kPa
pvap	122.47	kPa	335.00	Isobaric vapor liquid equilibria for acetone + methanol + lithium nitrate at 100 kPa
pvap	110.78	kPa	332.00	Isobaric vapor liquid equilibria for acetone + methanol + lithium nitrate at 100 kPa
pvap	100.00	kPa	329.00	Isobaric vapor liquid equilibria for acetone + methanol + lithium nitrate at 100 kPa
pvap	90.08	kPa	326.00	Isobaric vapor liquid equilibria for acetone + methanol + lithium nitrate at 100 kPa
pvap	80.96	kPa	323.00	Isobaric vapor liquid equilibria for acetone + methanol + lithium nitrate at 100 kPa
pvap	65.08	kPa	316.80	Vapor Liquid Equilibrium for Binary Systems of 2,3-Pentanedione with Diacetyl and Acetone
pvap	72.60	kPa	320.00	Isobaric vapor liquid equilibria for acetone + methanol + lithium nitrate at 100 kPa

pvap	101.30	kPa	329.56	Isobaric vapour-liquid equilibrium measurements and extractive distillation process for the azeotrope of (N,N-dimethylisopropylamine + acetone)
pvap	64.95	kPa	317.00	Isobaric vapor liquid equilibria for acetone + methanol + lithium nitrate at 100 kPa
rfi	1.35620		298.15	Solubility of a-Carotene in Binary Solvents Formed by Some Hydrocarbons with 2,5,8-Trioxanonane, 2-Propanone, and Cyclohexanone
rfi	1.35900		293.15	Solubilities of (2,5-Dihydroxyphenyl)diphenyl Phosphine Oxide in Selected Solvents
rfi	1.34491		318.15	Density, Speed of Sound, and Refractive Index of 1-Ethyl-3-methylimidazolium Trifluoromethanesulfonate with Acetone, Methyl Acetate, and Ethyl Acetate at Temperatures from (278.15 to 328.15) K
rfi	1.35054		308.15	Density, Speed of Sound, and Refractive Index of 1-Ethyl-3-methylimidazolium Trifluoromethanesulfonate with Acetone, Methyl Acetate, and Ethyl Acetate at Temperatures from (278.15 to 328.15) K

rfi	1.35597	298.15	Density, Speed of Sound, and Refractive Index of 1-Ethyl-3-methylimidazolium Trifluoromethanesulfonate with Acetone, Methyl Acetate, and Ethyl Acetate at Temperatures from (278.15 to 328.15) K
rfi	1.36146	288.15	Density, Speed of Sound, and Refractive Index of 1-Ethyl-3-methylimidazolium Trifluoromethanesulfonate with Acetone, Methyl Acetate, and Ethyl Acetate at Temperatures from (278.15 to 328.15) K
rfi	1.35900	293.15	Solubilities of 3,9-Dimethyl-3,9-dioxide-2,4,8,10-tetraoxa-3,9-diphosphaspiro[5.5]undec in Selected Solvents
rfi	1.35880	293.15	Solubilities of Methyldiphenylphosphine Oxide in Selected Solvents
rfi	1.35820	298.15	Isobaric Vapor Liquid Equilibria for Binary Systems of Acetone + Isopropenyl Acetate, 2-Butanone + Isopropenyl Acetate, and Isopropenyl Acetate + Acetylacetone at 101.3 kPa
rfi	1.35900	293.15	Solubilities of 2-(6-Oxido-6H-dibenz[c,e][1,2]oxaphosphorin-6-yl)-1,4-dihydroxy Phenylene in the Selected Solvents
rfi	1.35900	293.15	Solubilities of Phosphorus-Containing Compounds in Selected Solvents
rfi	1.35900	293.15	Solubilities of Some Phosphaspirocyclic Compounds in Selected Solvents

rfi	1.35890	293.15	Isothermal and Isobaric Vapor-Liquid Equilibria of the Ternary System of 2,2-Dimethoxypropane + Acetone + Methanol
rfi	1.35057	308.15	Densities, Refractive indices and Viscosities for Binary and Ternary Mixtures of Acetone, Ethanol and 2,2,4-Trimethylpentane at T = (288.15, 298.15, and 308.15) K
rfi	1.35605	298.15	Densities, Refractive indices and Viscosities for Binary and Ternary Mixtures of Acetone, Ethanol and 2,2,4-Trimethylpentane at T = (288.15, 298.15, and 308.15) K
rfi	1.36152	288.15	Densities, Refractive indices and Viscosities for Binary and Ternary Mixtures of Acetone, Ethanol and 2,2,4-Trimethylpentane at T = (288.15, 298.15, and 308.15) K
rfi	1.35860	293.15	Isobaric Vapor Liquid Equilibrium for Nine Binary Systems of Cracking C5 Fraction at 250 kPa

rfi	1.35553	298.15	Quaternary, Ternary and Binary LLE Measurements for 2-Methoxy-2-methylbutane + Furfural + Acetic Acid + Water at Temperatures between 298 and 341 K
rfi	1.35553	298.15	Quaternary, Ternary, and Binary LLE Measurements for 2-Methoxy-2-methylpropane + Furfural + Acetic Acid + Water at Temperatures between 298 and 307 K
rfi	1.35553	298.15	Ternary and Binary LLE Measurements for Solvent (4-Methyl-2-pentanone and 2-Methyl-2-butanol) + Furfural + Water between 298 and 401 K
rfi	1.35600	298.15	Physical properties and their corresponding changes of mixing for the ternary mixture acetone + n-hexane +water at 298.15K
rfi	1.35550	298.00	Quaternary and ternary LLE measurements for solvent (2-methyltetrahydrofuran and cyclopentyl methyl ether) + furfural + acetic acid + water between 298 and 343 K
rfi	1.35566	298.15	Properties of ionic liquid HMIMPF6 with carbonates, ketones and alkyl acetates

rfi	1.35880		293.15	Isothermal vapour-liquid equilibrium data for the binary systems 2-propanone + (2-butanol or propanoic acid)
rfi	1.35732		298.15	Vapor liquid equilibria and interfacial tensions for the ternary system acetone + 2,2 -oxybis[propane] + cyclohexane and its constituent binary systems
rfi	1.35597		298.15	Isobaric vapor liquid equilibria for mixtures of acetone, ethanol, and 2,2,4-trimethylpentane at 101.3 kPa
rfi	1.35732		298.15	Phase equilibria and interfacial tensions in the systems methyl tert-butyl ether + acetone + cyclohexane, methyl tert-butyl ether + acetone and methyl tert-butyl ether + cyclohexane
rfi	1.35550		298.15	Ternary and binary LLE measurements for solvent (2-methyltetrahydrofuran and cyclopentyl methyl ether) + furfural + water between 298 and 343 K
rhol	778.66	kg/m3	303.15	Densities, viscosities, and refractive indices of binary and ternary mixtures of methanol, acetone, and chloroform at temperatures from (298.15-318.15) K and ambient pressure

rhol	784.65	kg/m3	298.15	Excess molar enthalpies and volumes of binary mixtures of nonafluorobutylmethylether with ketones at T = 298.15 K
rhol	784.23	kg/m3	298.15	Density and speed of sound of lithium bromide with organic solvents: Measurement and correlation
rhol	784.24	kg/m3	298.15	Apparent molar volumes and compressibilities of tetrabutyl-ammonium bromide in organic solvents
rhol	789.99	kg/m3	293.15	Thermodynamic properties of binary mixtures of the ionic liquid [emim][BF4] with acetone and dimethylsulphoxide
rhol	784.40	kg/m3	298.15	Thermodynamic properties of binary mixtures of the ionic liquid [emim][BF4] with acetone and dimethylsulphoxide
rhol	778.87	kg/m3	303.15	Thermodynamic properties of binary mixtures of the ionic liquid [emim][BF4] with acetone and dimethylsulphoxide
rhol	773.07	kg/m3	308.15	Thermodynamic properties of binary mixtures of the ionic liquid [emim][BF4] with acetone and dimethylsulphoxide
rhol	785.23	kg/m3	298.15	Separation of sulfur compounds from alkanes with 1-alkylcyanopyridinium-based ionic liquids
rhol	785.32	kg/m3	298.15	Effect of the alkyl side chain of the 1-alkylpiperidinium-based ionic liquids on desulfurization of fuels

rhol	785.32	kg/m3	298.15	Extraction desulfurization process of fuels with ionic liquids
rhol	783.80	kg/m3	298.15	Solubility and solution thermodynamics of sorbic acid in eight pure organic solvents
rhol	790.21	kg/m3	293.15	(Liquid + liquid) equilibria for (water + 1-propanol or acetone + .beta.-citronellol) at different temperatures
rhol	790.66	kg/m3	293.20	Ionic liquid 1-hexyl-3-methylimidazolium hexafluorophosphate, an efficient solvent for extraction of acetone from aqueous solutions
rhol	784.40	kg/m3	298.20	Ionic liquid 1-hexyl-3-methylimidazolium hexafluorophosphate, an efficient solvent for extraction of acetone from aqueous solutions
rhol	778.70	kg/m3	303.20	Ionic liquid 1-hexyl-3-methylimidazolium hexafluorophosphate, an efficient solvent for extraction of acetone from aqueous solutions
rhol	785.10	kg/m3	298.15	Measurement and correlation of solubility and solution thermodynamics of 1,3-dimethylurea in different solvents from T = (288.15 to 328.15) K

rhol	790.19	kg/m3	293.15	Thermodynamics of amide + ketone mixtures. 1. Volumetric, speed of sound and refractive index data for N,N-dimethylformamide + 2-alkanone systems at several temperatures
rhol	784.43	kg/m3	298.15	Thermodynamics of amide + ketone mixtures. 1. Volumetric, speed of sound and refractive index data for N,N-dimethylformamide + 2-alkanone systems at several temperatures
rhol	778.63	kg/m3	303.15	Thermodynamics of amide + ketone mixtures. 1. Volumetric, speed of sound and refractive index data for N,N-dimethylformamide + 2-alkanone systems at several temperatures
rhol	786.77	kg/m3	298.15	Separation of ethylbenzene/styrene systems using ionic liquids in ternary LLE
rhol	790.00	kg/m3	293.15	Thermophysical approach to understand the nature of molecular interactions and structural factor between methyl isobutyl ketone and organic solvents mixtures
rhol	779.00	kg/m3	303.15	Thermophysical approach to understand the nature of molecular interactions and structural factor between methyl isobutyl ketone and organic solvents mixtures

rhol	767.00	kg/m3	313.15	Thermophysical approach to understand the nature of molecular interactions and structural factor between methyl isobutyl ketone and organic solvents mixtures
rhol	785.00	kg/m3	298.15	Vapor Liquid Equilibrium Data for 2,3-Pentanedione + (Acetaldehyde or Acetone) at (100, 150, and 200) kPa
rhol	790.31	kg/m3	293.15	Excess Molar Enthalpies for Binary Mixtures of Ethanol + Acetone, + Octane, + Cyclohexane and 1-Propanol + Acetone, + Octane, + Heptane at 323.15
rhol	801.67	kg/m3	283.15	Experimental Densities and Excess Volumes for Binary Mixtures Containing Propionic Acid, Acetone and Water from 283.15 K to 323.15 K at Atmospheric Pressure
rhol	796.04	kg/m3	288.15	Experimental Densities and Excess Volumes for Binary Mixtures Containing Propionic Acid, Acetone and Water from 283.15 K to 323.15 K at Atmospheric Pressure

rhol	790.36	kg/m3	293.15	Experimental Densities and Excess Volumes for Binary Mixtures Containing Propionic Acid, Acetone and Water from 283.15 K to 323.15 K at Atmospheric Pressure
rhol	784.64	kg/m3	298.15	Experimental Densities and Excess Volumes for Binary Mixtures Containing Propionic Acid, Acetone and Water from 283.15 K to 323.15 K at Atmospheric Pressure
rhol	778.88	kg/m3	303.14	Experimental Densities and Excess Volumes for Binary Mixtures Containing Propionic Acid, Acetone and Water from 283.15 K to 323.15 K at Atmospheric Pressure
rhol	773.07	kg/m3	308.15	Experimental Densities and Excess Volumes for Binary Mixtures Containing Propionic Acid, Acetone and Water from 283.15 K to 323.15 K at Atmospheric Pressure
rhol	767.21	kg/m3	313.14	Experimental Densities and Excess Volumes for Binary Mixtures Containing Propionic Acid, Acetone and Water from 283.15 K to 323.15 K at Atmospheric Pressure

rhol	761.29	kg/m3	318.14	Experimental Densities and Excess Volumes for Binary Mixtures Containing Propionic Acid, Acetone and Water from 283.15 K to 323.15 K at Atmospheric Pressure	
rhol	755.31	kg/m3	323.14	Experimental Densities and Excess Volumes for Binary Mixtures Containing Propionic Acid, Acetone and Water from 283.15 K to 323.15 K at Atmospheric Pressure	
rhol	784.39	kg/m3	298.15	Excess Molar Volumes and Surface Tensions of Xylene with Acetone or 2-Butanone at 298.15 K	
rhol	785.09	kg/m3	298.15	Thermodynamics of Ketone + Amine Mixtures. Part VIII. Molar Excess Enthalpies at 298.15 K for n-Alkanone + Aniline or + N-Methylaniline Systems	
rhol	784.30	kg/m3	293.15	Liquid Liquid Equilibrium for the Ternary System Acetone Oxime Methyl Ether Acetone Water	
rhol	784.30	kg/m3	298.15	Phase equilibrium study of binary and ternary mixtures of ionic liquids + acetone + methanol	

rhol	784.10	kg/m3	298.15	Density, Speed of Sound, and Derived Thermodynamic Properties of Ionic Liquids [EMIM]+[BETI]- or ([EMIM]+[CH3(OCH2CH2)2OSO3]- + Methanol or + Acetone) at T = (298.15 or 303.15 or 313.15) K
rhol	778.40	kg/m3	303.15	Density, Speed of Sound, and Derived Thermodynamic Properties of Ionic Liquids [EMIM]+[BETI]- or ([EMIM]+[CH3(OCH2CH2)2OSO3]- + Methanol or + Acetone) at T = (298.15 or 303.15 or 313.15) K
rhol	766.90	kg/m3	313.15	Density, Speed of Sound, and Derived Thermodynamic Properties of Ionic Liquids [EMIM]+[BETI]- or ([EMIM]+[CH3(OCH2CH2)2OSO3]- + Methanol or + Acetone) at T = (298.15 or 303.15 or 313.15) K
rhol	783.81	kg/m3	298.15	Partial Molar Volumes of Butyltriethylammonium Iodide in Single Nonaqueous Solvents at 298.15 K
rhol	790.00	kg/m3	293.00	KDB

rhol	785.09	kg/m3	298.15	Thermodynamics of ketone + amine mixtures. Part IX. Excess molar enthalpies at 298.15K for dipropylamine, or dibutylamine + 2-alkanone systems and modeling of linear or aromatic amine + 2-alkanone mixtures in terms of DISQUAC and ERAS
rhol	785.09	kg/m3	298.15	Thermodynamics of ketone + amine mixtures. Part X. Excess molarenthalpies at 298.15 K for N,N,N-triethylamine + 2-alkanone systems.Characterization of tertiary amine + 2-alkanone, and ofamino-ketone + n-alkane mixtures in terms of DISQUAC
rhol	784.41	kg/m3	298.15	Densities, viscosities, and refractive indices of binary and ternary mixtures of methanol, acetone, and chloroform at temperatures from (298.15-318.15) K and ambient pressure
rhol	761.26	kg/m3	318.15	Densities, viscosities, and refractive indices of binary and ternary mixtures of methanol, acetone, and chloroform at temperatures from (298.15-318.15) K and ambient pressure

rhol	766.95	kg/m3	313.15	Densities, viscosities, and refractive indices of binary and ternary mixtures of methanol, acetone, and chloroform at temperatures from (298.15-318.15) K and ambient pressure
rhol	773.08	kg/m3	308.15	Densities, viscosities, and refractive indices of binary and ternary mixtures of methanol, acetone, and chloroform at temperatures from (298.15-318.15) K and ambient pressure
rhol	784.50	kg/m3	293.15	(Vapour + liquid) equilibria for (2-ethoxypropene + acetone) and (2-ethoxypropene + butanone)
rhol	785.23	kg/m3	298.15	Separation of pyridine from heptane with tricyanomethanide-based ionic liquids
sfust	32.36	J/mol×K	176.62	NIST Webbook
sfust	26.70	J/mol×K	178.50	NIST Webbook
sfust	32.00	J/mol×K	177.60	NIST Webbook
sfust	32.03	J/mol×K	177.60	NIST Webbook
speedsl	1162.00	m/s	298.15	Vapor liquid equilibria for systems of diethyl carbonate and ketones and determination of group interaction parameters for the UNIFAC and ASOG methods
srf	0.02	N/m	327.88	Surface Tension of Pure Liquids and Binary Liquid Mixtures
srf	0.02	N/m	308.15	Surface Tension of the Ternary System Water + Acetone + Toluene

srf	0.02	N/m	298.15	Surface Tension of the Ternary System Water + Acetone + Toluene
srf	0.02	N/m	288.15	Surface Tension of the Ternary System Water + Acetone + Toluene
srf	0.02	N/m	293.15	Investigation of Surface Properties and Solubility of 1-Vinyl-3-alkyl/Esterimidazolium Halide Ionic Liquids by Density Functional Methods
srf	0.02	N/m	318.15	Surface Tension of the Ternary System Water + Acetone + Toluene
srf	0.02	N/m	317.86	Surface Tension of Pure Liquids and Binary Liquid Mixtures
srf	0.02	N/m	307.86	Surface Tension of Pure Liquids and Binary Liquid Mixtures
srf	0.02	N/m	297.82	Surface Tension of Pure Liquids and Binary Liquid Mixtures
srf	0.02	N/m	287.81	Surface Tension of Pure Liquids and Binary Liquid Mixtures
srf	0.02	N/m	298.20	KDB
srf	0.02	N/m	328.15	Surface Tension of the Ternary System Water + Acetone + Toluene

Pressure Dependent Properties

Property code	Value	Unit	Pressure [kPa]	Source
tbrp	329.30	K	2.70	NIST Webbook

Correlations

Information	Value
Property code	pvap
Equation	ln(Pvp) = A + B/(T + C)
Coeff. A	1.48340e+01
Coeff. B	-3.02945e+03
Coeff. C	-3.26710e+01
Temperature range (K), min.	240.93
Temperature range (K), max.	508.10

Information	Value
Property code	pvap
Equation	ln(Pvp) = A + B/T + C*ln(T) + D*T^2
Coeff. A	6.69693e+01
Coeff. B	-5.78459e+03
Coeff. C	-7.85881e+00
Coeff. D	7.10496e-06
Temperature range (K), min.	178.45
Temperature range (K), max.	508.20

Datasets

Viscosity, Pa*s

Temperature, K - Liquid	Pressure, kPa - Liquid	Viscosity, Pa*s - Liquid
303.15	101.30	0.0002974

Reference *https://www.doi.org/10.1021/je034204h*

Speed of sound, m/s

Temperature, K - Liquid	Pressure, kPa - Liquid	Speed of sound, m/s - Liquid

265.67	102.00	1298.93
265.67	504.00	1301.09
265.67	750.00	1302.5
265.67	1003.00	1303.56
265.67	1255.00	1304.92
265.67	1501.00	1306.29
265.67	2007.00	1308.73
265.67	2506.00	1311.28
265.67	3005.00	1313.79
265.67	4004.00	1318.59
265.67	5004.00	1323.31
265.67	6005.00	1328.01
265.67	8002.00	1338.69
265.67	10001.00	1349.81
265.67	12001.00	1358.7
265.67	14003.00	1366.48
265.67	16050.00	1375.48
265.67	18001.00	1384.77
265.67	20003.00	1394.13
265.67	23001.00	1407.08
265.67	27004.00	1425.32
265.67	31003.00	1439.78
265.67	34997.00	1455.17
265.67	39999.00	1476.05
265.67	45002.00	1494.61
265.67	49999.00	1511.6
265.67	55001.00	1528.77
265.67	60002.00	1548.05
265.67	65000.00	1564.17
265.67	18001.00	1384.54
273.16	102.00	1264.48
273.16	504.00	1266.57
273.16	750.00	1267.96
273.16	1003.00	1269.39
273.16	1255.00	1270.74
273.16	1501.00	1272.23
273.16	2007.00	1275.12
273.16	2506.00	1277.93
273.16	3005.00	1281.1
273.16	4004.00	1287.23
273.16	5004.00	1291.97
273.16	6005.00	1296.6
273.16	8002.00	1306.58
273.16	10001.00	1316.19

273.16	12001.00	1325.58
273.16	14003.00	1335.9
273.16	16005.00	1346.87
273.16	18001.00	1355.99
273.16	20003.00	1363.63
273.16	23001.00	1376.52
273.16	27004.00	1394.61
273.16	31003.00	1412.24
273.16	34997.00	1428.86
273.16	39999.00	1446.67
273.16	45002.00	1467.44
273.16	49999.00	1485.84
273.16	55001.00	1504.31
273.16	60002.00	1520.08
273.16	65000.00	1538.71
273.16	3005.00	1281.19
273.16	4004.00	1286.78
273.16	5004.00	1292.02
280.74	75.00	1233.63
280.74	504.00	1234.94
280.74	750.00	1236.39
280.74	1003.00	1237.85
280.74	1255.00	1239.22
280.74	1501.00	1240.59
280.74	2007.00	1243.75
280.74	2506.00	1247.03
280.74	3005.00	1249.76
280.74	4004.00	1255.13
280.74	5004.00	1259.87
280.74	6005.00	1264.8
280.74	8002.00	1275.79
280.74	10001.00	1286.82
280.74	12001.00	1296.79
280.74	14003.00	1306.71
280.74	16005.00	1316.23
280.74	18001.00	1325.46
280.74	20003.00	1335.09
280.74	23001.00	1350.8
280.74	27004.00	1366.99
280.74	31003.00	1384.26
280.74	34997.00	1401.32
280.74	39999.00	1423.56
280.74	45002.00	1441.37
280.74	49999.00	1461.28

280.74	55001.00	1480.44
280.74	60002.00	1498.97
280.74	65000.00	1514.56
280.74	75.00	1232.04
280.74	23001.00	1350.8
280.74	39999.00	1423.74
280.74	45002.00	1441.37
280.74	60002.00	1498.91
280.74	65000.00	1514.63
288.23	109.00	1199.01
288.23	504.00	1201.24
288.23	750.00	1202.57
288.23	1003.00	1204.2
288.23	1255.00	1205.54
288.23	1501.00	1206.97
288.23	2007.00	1209.92
288.23	2506.00	1212.8
288.23	3005.00	1215.61
288.23	4004.00	1221.35
288.23	5004.00	1227.14
288.23	6005.00	1232.71
288.23	8002.00	1244.39
288.23	10001.00	1255.4
288.23	12001.00	1265.22
288.23	14003.00	1275.63
288.23	16005.00	1286.61
288.23	18001.00	1296.18
288.23	20003.00	1305.89
288.23	23001.00	1319.61
288.23	27004.00	1339.05
288.23	31003.00	1357.37
288.23	34997.00	1373.83
288.23	39999.00	1395.08
288.23	45002.00	1416.1
288.23	49999.00	1435.37
288.23	55001.00	1453.97
288.23	60002.00	1473.72
288.23	65000.00	1492.08
298.23	109.00	1153.69
298.23	504.00	1156.23
298.23	750.00	1157.82
298.23	1003.00	1159.5
298.23	1255.00	1161.14
298.23	1501.00	1162.71

298.23	2007.00	1165.65
298.23	2506.00	1168.6
298.23	3005.00	1171.49
298.23	4004.00	1177.4
298.23	5004.00	1182.9
298.23	6005.00	1189.6
298.23	8002.00	1201.45
298.23	10001.00	1212.8
298.23	12001.00	1223.83
298.23	14003.00	1234.83
298.23	16005.00	1246.18
298.23	18001.00	1256.46
298.23	20003.00	1265.92
298.23	23001.00	1281.93
298.23	27004.00	1301.17
298.23	31003.00	1319.28
298.23	34997.00	1337.9
298.23	39999.00	1359.25
298.23	45002.00	1380.44
298.23	49999.00	1400.61
298.23	55001.00	1421.32
298.23	60002.00	1439.68
298.23	65000.00	1458.59
308.22	109.00	1107.74
308.22	504.00	1110.52
308.22	750.00	1112.35
308.22	1003.00	1114.01
308.22	1255.00	1115.75
308.22	1501.00	1117.34
308.22	2007.00	1120.69
308.22	2506.00	1123.91
308.22	3005.00	1127.11
308.22	4004.00	1133.79
308.22	5004.00	1140.56
308.22	6005.00	1146.98
308.22	8002.00	1159.51
308.22	10001.00	1171.13
308.22	12001.00	1182.37
308.22	14003.00	1194.75
308.22	16005.00	1205.76
308.22	18001.00	1216.57
308.22	20003.00	1227.49
308.22	23001.00	1244.05
308.22	27004.00	1263.44

308.22	31003.00	1283.97
308.22	34997.00	1302.46
308.22	39999.00	1324.45
308.22	45002.00	1347.09
308.22	49999.00	1367.79
308.22	55001.00	1388.31
308.22	60002.00	1408.07
308.22	65000.00	1427.73
318.22	109.00	1062.32
318.22	504.00	1065.38
318.22	750.00	1067.34
318.22	1003.00	1069.21
318.22	1255.00	1071.11
318.22	1501.00	1072.72
318.22	2007.00	1076.32
318.22	2506.00	1079.71
318.22	3005.00	1083.5
318.22	4004.00	1090.38
318.22	5004.00	1096.85
318.22	6005.00	1103.61
318.22	8002.00	1116.61
318.22	10001.00	1129.13
318.22	12001.00	1142.47
318.22	14003.00	1154.62
318.22	16005.00	1166.43
318.22	18001.00	1177.61
318.22	20003.00	1189.14
318.22	23001.00	1205.63
318.22	27004.00	1226.77
318.22	31003.00	1247.41
318.22	34997.00	1266.28
318.22	39999.00	1290.76
318.22	45002.00	1313.78
318.22	49999.00	1334.83
318.22	55001.00	1356.46
318.22	60002.00	1376.77
318.22	65000.00	1395.43
328.22	129.00	1018.3
328.22	504.00	1021.14
328.22	750.00	1023.1
328.22	1003.00	1024.91
328.22	1255.00	1026.82
328.22	1501.00	1028.6
328.22	2007.00	1032.33

328.22	2506.00	1036.21
328.22	3005.00	1040.05
328.22	4004.00	1047.07
328.22	5004.00	1053.76
328.22	6005.00	1060.7
328.22	8002.00	1074.73
328.22	10001.00	1088.42
328.22	12001.00	1101.14
328.22	14003.00	1113.97
328.22	16005.00	1126.49
328.22	18001.00	1138.92
328.22	20003.00	1151.08
328.22	23001.00	1168.1
328.22	27004.00	1190.45
328.22	31003.00	1211.66
328.22	34997.00	1231.66
328.22	39999.00	1256.67
328.22	45002.00	1280.63
328.22	49999.00	1303.34
328.22	55001.00	1324.67
328.22	60002.00	1345.81
328.22	65000.00	1365.57
338.22	170.00	972.8
338.22	504.00	976.32
338.22	750.00	978.88
338.22	1003.00	981.05
338.22	1255.00	982.97
338.22	1501.00	984.75
338.22	2007.00	988.42
338.22	2506.00	992.07
338.22	3005.00	996.0
338.22	4004.00	1004.26
338.22	5004.00	1011.63
338.22	6005.00	1019.16
338.22	8002.00	1033.34
338.22	10001.00	1047.72
338.22	12001.00	1061.19
338.22	14003.00	1074.71
338.22	16005.00	1088.09
338.22	18001.00	1100.33
338.22	20003.00	1112.58
338.22	23001.00	1131.02
338.22	27004.00	1154.76
338.22	31003.00	1176.19

338.22	34997.00	1197.57
338.22	39999.00	1223.29
338.22	45002.00	1248.14
338.22	49999.00	1271.19
338.22	55001.00	1293.31
338.22	60002.00	1315.39
338.22	65000.00	1336.38

Reference *https://www.doi.org/10.1016/j.jct.2003.12.001*

Temperature, K	Pressure, kPa	Speed of sound, m/s
248.15	100.00	1390.16
248.15	10000.00	1435.51
248.15	20000.00	1478.36
248.15	30000.00	1518.03
248.15	40000.00	1550.78
248.15	50000.00	1590.88
248.15	60000.00	1625.12
248.15	70000.00	1657.38
248.15	80000.00	1688.07
248.15	90000.00	1717.48
248.15	100000.00	1745.43
253.15	100.00	1367.16
253.15	10000.00	1413.84
253.15	20000.00	1457.41
253.15	30000.00	1497.98
253.15	40000.00	1536.44
253.15	50000.00	1572.26
253.15	60000.00	1606.9
253.15	70000.00	1639.52
253.15	80000.00	1670.88
253.15	90000.00	1700.93
253.15	100000.00	1729.65
258.15	100.00	1344.2
258.15	10000.00	1391.98
258.15	20000.00	1436.58
258.15	30000.00	1478.11
258.15	40000.00	1517.33
258.15	50000.00	1553.89
258.15	60000.00	1588.88
258.15	70000.00	1621.91
258.15	80000.00	1653.89
258.15	90000.00	1684.43

258.15	100000.00	1713.62
263.15	100.00	1320.92
263.15	10000.00	1370.35
263.15	20000.00	1415.99
263.15	30000.00	1458.25
263.15	40000.00	1498.44
263.15	50000.00	1535.81
263.15	60000.00	1570.98
263.15	70000.00	1604.54
263.15	80000.00	1637.11
263.15	90000.00	1668.1
263.15	100000.00	1697.74
268.15	100.00	1298.67
268.15	10000.00	1349.08
268.15	20000.00	1395.62
268.15	30000.00	1438.54
268.15	40000.00	1479.77
268.15	50000.00	1517.83
268.15	60000.00	1553.2
268.15	70000.00	1587.43
268.15	80000.00	1620.54
268.15	90000.00	1651.96
268.15	100000.00	1682.01
273.15	100.00	1275.4
273.15	10000.00	1327.81
273.15	20000.00	1375.45
273.15	30000.00	1418.96
273.15	40000.00	1461.31
273.15	50000.00	1499.98
273.15	60000.00	1535.77
273.15	70000.00	1570.58
273.15	80000.00	1604.15
273.15	90000.00	1636.02
273.15	100000.00	1666.45
278.15	100.00	1253.32
278.15	10000.00	1306.75
278.15	20000.00	1355.49
278.15	30000.00	1400.47
278.15	40000.00	1443.03
278.15	50000.00	1481.78
278.15	60000.00	1518.62
278.15	70000.00	1553.98
278.15	80000.00	1587.94
278.15	90000.00	1620.26

278.15	100000.00	1650.92
283.15	100.00	1231.06
283.15	10000.00	1285.85
283.15	20000.00	1335.74
283.15	30000.00	1381.4
283.15	40000.00	1424.94
283.15	50000.00	1464.43
283.15	60000.00	1502.23
283.15	70000.00	1537.63
283.15	80000.00	1571.96
283.15	90000.00	1604.68
283.15	100000.00	1635.63
288.15	100.00	1208.58
288.15	10000.00	1265.05
288.15	20000.00	1316.19
288.15	30000.00	1362.56
288.15	40000.00	1407.06
288.15	50000.00	1447.7
288.15	60000.00	1485.52
288.15	70000.00	1521.53
288.15	80000.00	1556.22
288.15	90000.00	1589.2
288.15	100000.00	1620.74
293.15	100.00	1186.09
293.15	10000.00	1244.14
293.15	20000.00	1296.86
293.15	30000.00	1344.06
293.15	40000.00	1389.16
293.15	50000.00	1430.7
293.15	60000.00	1469.0
293.15	70000.00	1505.7
293.15	80000.00	1540.77
293.15	90000.00	1574.27
293.15	100000.00	1605.01
298.15	100.00	1165.34
298.15	10000.00	1223.74
298.15	20000.00	1277.93
298.15	30000.00	1325.68
298.15	40000.00	1371.61
298.15	50000.00	1413.84
298.15	60000.00	1452.64
298.15	70000.00	1490.14
298.15	80000.00	1525.25
298.15	90000.00	1559.14

298.15	100000.00	1591.53

Reference https://www.doi.org/10.1016/j.jct.2008.11.005

Sources

Thermodynamic Properties of Mixtures Containing Ionic Liquids. 8. Activity Coefficients at Infinite Dilution of Alkanals, Alcohols, Esters, and Alkylbenzenes and Water in the ionic liquid 1-Hexyl-3-methylimidazolium Bis(trifluoromethylsulfonyl)imide	https://www.doi.org/10.1021/je0503554
Activity coefficients at infinite dilution and physicochemical properties for 1-Butyl-3-methylimidazolium dicyanamide and 1-Butyl-1-methylpyrrolidinium dicyanamide	https://www.doi.org/10.1016/j.jct.2012.01.019
Determination and correlation of solubilities of pure and mixed Solvents:	https://www.doi.org/10.1016/j.jct.2015.09.026
Experimental Densities and Excess Volumes for Binary Mixtures Containing Propionic Acid, Acetone and Water from 283.15 K to 323.15 K at Atmospheric Pressure	https://www.doi.org/10.1016/j.jct.2019.06.007
High-pressure phase equilibrium data for carbon dioxide + dichloromethane + 1,3,2-Dioxaphosphorinane-2-methanol-a,5,5-trimethyl-a-phenyl-2-oxide	https://www.doi.org/10.1021/je400513s
Solubility Determination and Thermodynamic Models for 2-Methylnaphthalene in Different Organic Solvents at Different Temperatures:	https://www.doi.org/10.1021/je030102f
Surface Tension and Interfacial Tension of Binary Organic Liquid Mixtures:	https://www.doi.org/10.1016/j.fluid.2018.06.008
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Legend

af:	Acentric Factor
affp:	Proton affinity
aigt:	Autoignition Temperature
basg:	Gas basicity
chg:	Standard gas enthalpy of combustion
chl:	Standard liquid enthalpy of combustion
cpg:	Ideal gas heat capacity
cpl:	Liquid phase heat capacity
cps:	Solid phase heat capacity
dm:	Dipole Moment
dvisc:	Dynamic viscosity
ea:	Electron affinity
fll:	Lower Flammability Limit
flu:	Upper Flammability Limit
fpc:	Flash Point (Closed Cup Method)
fpo:	Flash Point (Open Cup Method)
gf:	Standard Gibbs free energy of formation
gyrad:	Radius of Gyration
hf:	Enthalpy of formation at standard conditions
hfl:	Liquid phase enthalpy of formation at standard conditions
hfus:	Enthalpy of fusion at standard conditions
hfust:	Enthalpy of fusion at a given temperature
hvap:	Enthalpy of vaporization at standard conditions
hvapt:	Enthalpy of vaporization at a given temperature
ie:	Ionization energy
log10ws:	Log10 of Water solubility in mol/l
logp:	Octanol/Water partition coefficient
mcvol:	McGowan's characteristic volume
nfpaf:	NFPA Fire Rating
nfpah:	NFPA Health Rating
pc:	Critical Pressure
pvap:	Vapor pressure
rfi:	Refractive Index
rhoc:	Critical density
rhol:	Liquid Density
rinpol:	Non-polar retention indices
ripol:	Polar retention indices
sfust:	Entropy of fusion at a given temperature
sl:	Liquid phase molar entropy at standard conditions

speedsl:	Speed of sound in fluid
srf:	Surface Tension
tb:	Normal Boiling Point Temperature
tbrp:	Boiling point at reduced pressure
tc:	Critical Temperature
tf:	Normal melting (fusion) point
tt:	Triple Point Temperature
vc:	Critical Volume
zc:	Critical Compressibility
zra:	Rackett Parameter

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