Acetone

Other names:	(CH3)2CO
	2-PROPANONE
	Chevron acetone
	DIMETHYL KETONE
	Dimethylformaldehyde
	Dimethylketal
	KETONE PROPANE
	Ketone, dimethyl-
	Methyl ketone
	NSC 135802
	Propan-2-one
	Propanone
	Pyroacetic ether
	Rcra waste number U002
	Sasetone
	UN 1090
	«beta»-Ketopropane
	Â«betaÂ»-Ketopropane
Inchi:	InChI=1S/C3H6O/c1-3(2)4/h1-2H3
InchiKey:	CSCPPACGZOOCGX-UHFFFAOYSA-N
Formula:	C3H6O
SMILES:	CC(C)=O
Mol. weight [g/mol]:	58.08
CAS:	67-64-1

Physical Properties

Property code	Value	Unit	Source
af	0.3040		KDB
affp	814.30	kJ/mol	NIST Webbook
affp	811.50 ± 3.40	kJ/mol	NIST Webbook
affp	811.50 ± 3.40	kJ/mol	NIST Webbook
affp	815.20	kJ/mol	NIST Webbook
affp	812.00	kJ/mol	NIST Webbook
affp	812.60 ± 0.20	kJ/mol	NIST Webbook
aigt	738.15	K	KDB
basg	782.10	kJ/mol	NIST Webbook
basg	782.20	kJ/mol	NIST Webbook

basg	782.10 ± 1.50	kJ/mol	NIST Webbook
basg	784.70	kJ/mol	NIST Webbook
basg	782.10 ± 1.50	kJ/mol	NIST Webbook
basg	782.00 ± 0.20	kJ/mol	NIST Webbook
chg	-1821.40 ± 0.84	kJ/mol	NIST Webbook
chl	-1772.00	kJ/mol	NIST Webbook
chl	-1804.20	kJ/mol	NIST Webbook
dm	2.90	debye	KDB
dvisc	0.0003121	Pa×s	Densities and Viscosities of Binary Liquid Mixtures of Trichloroethylene and Tetrachloroethylene with Some Polar and Nonpolar Solvents
dvisc	0.0003000	Pa×s	Densities and Viscosities of 1-Butyl-3-methylimidazolium Tetrafluoroborate + Molecular Solvent Binary Mixtures
ea	0.00	eV	NIST Webbook
fll	2.60	% in Air	KDB
flu	12.80	% in Air	KDB
fpc	257.59	K	KDB
fpo	255.37	K	KDB
gf	-153.20	kJ/mol	KDB
gyrad	2.7400		KDB
hf	-217.10 ± 0.50	kJ/mol	NIST Webbook
hf	-217.70	kJ/mol	KDB
hf	-217.50 ± 0.67	kJ/mol	NIST Webbook
hf	-216.40	kJ/mol	NIST Webbook
hf	-218.50 ± 0.59	kJ/mol	NIST Webbook
hfl	-249.40 ± 0.63	kJ/mol	NIST Webbook
hfus	5.12	kJ/mol	Joback Method
hvap	31.27	kJ/mol	NIST Webbook
hvap	29.70 ± 0.00	kJ/mol	NIST Webbook
hvap	31.30	kJ/mol	NIST Webbook
ie	9.71	eV	NIST Webbook
ie	9.68	eV	NIST Webbook
ie	9.71	eV	NIST Webbook
ie	9.71 ± 0.01	eV	NIST Webbook
ie	9.71	eV	NIST Webbook
ie	9.72	eV	NIST Webbook
ie	9.74	eV	NIST Webbook
ie	9.71 ± 0.01	eV	NIST Webbook
ie	9.74 ± 0.03	eV	NIST Webbook
ie	9.68	eV	NIST Webbook

ie	9.71 ± 0.01	eV	NIST Webbook
ie	9.70 ± 0.10	eV	NIST Webbook
ie	9.68 ± 0.02	eV	NIST Webbook
ie	9.67	eV	NIST Webbook
ie	9.71 ± 0.03	eV	NIST Webbook
ie	9.71 ± 0.03	eV	NIST Webbook
ie	9.69 ± 0.01	eV	NIST Webbook
ie	9.71	eV	NIST Webbook
ie	9.80	eV	NIST Webbook
ie	9.72	eV	NIST Webbook
ie	9.71 ± 0.02	eV	NIST Webbook
ie	9.71 ± 0.01	eV	NIST Webbook
ie	9.71 ± 0.03	eV	NIST Webbook
ie	9.71 ± 0.01	eV	NIST Webbook
ie	9.71	eV	NIST Webbook
ie	9.75 ± 0.03	eV	NIST Webbook
ie	9.68	eV	NIST Webbook
ie	9.69 ± 0.01	eV	NIST Webbook
ie	9.70	eV	NIST Webbook
ie	9.70 ± 0.01	eV	NIST Webbook
ie	9.72	eV	NIST Webbook
ie	9.50	eV	NIST Webbook
ie	9.70	eV	NIST Webbook
ie	9.70 ± 0.00	eV	NIST Webbook
ie	9.71	eV	NIST Webbook
log10ws	-0.36		Crippen Method
logp	0.595		Crippen Method
mcvol	54.700	ml/mol	McGowan Method
nfpaf	%!d(float64=3)		KDB
nfpah	%!d(float64=1)		KDB
pc	4700.00	kPa	KDB
rhoc	234.06 ± 1.51	kg/m3	NIST Webbook
rhoc	272.97 ± 1.16	kg/m3	NIST Webbook
rhoc	268.91 ± 2.90	kg/m3	NIST Webbook
rhoc	252.06 ± 1.74	kg/m3	NIST Webbook
rhoc	278.20 ± 9.87	kg/m3	NIST Webbook
rinpol	472.00		NIST Webbook
rinpol	487.00		NIST Webbook
rinpol	474.00		NIST Webbook
rinpol	474.00		NIST Webbook
rinpol	500.00		NIST Webbook
rinpol	509.00		NIST Webbook
rinpol	512.00		NIST Webbook
rinpol	459.00		NIST Webbook

rinpol	471.00	NIST Webbook
rinpol	450.00	NIST Webbook
rinpol	476.00	NIST Webbook
rinpol	501.00	NIST Webbook
rinpol	470.00	NIST Webbook
rinpol	470.00	NIST Webbook
rinpol	450.00	NIST Webbook
rinpol	497.00	NIST Webbook
rinpol	500.00	NIST Webbook
rinpol	500.00	NIST Webbook
rinpol	460.00	NIST Webbook
rinpol	473.00	NIST Webbook
rinpol	473.00	NIST Webbook
rinpol	473.00	NIST Webbook
rinpol	465.00	NIST Webbook
rinpol	460.00	NIST Webbook
rinpol	469.00	NIST Webbook
rinpol	491.00	NIST Webbook
rinpol	474.00	NIST Webbook
rinpol	484.00	NIST Webbook
rinpol	468.00	NIST Webbook
rinpol	469.00	NIST Webbook
rinpol	478.00	NIST Webbook
rinpol	468.00	NIST Webbook
rinpol	503.00	NIST Webbook
rinpol	503.00	NIST Webbook
rinpol	471.00	NIST Webbook
rinpol	470.00	NIST Webbook
rinpol	479.00	NIST Webbook
rinpol	502.00	NIST Webbook
rinpol	477.55	NIST Webbook
rinpol	470.23	NIST Webbook
rinpol	469.50	NIST Webbook
rinpol	469.00	NIST Webbook
rinpol	478.00	NIST Webbook
rinpol	475.00	NIST Webbook
rinpol	465.00	NIST Webbook
rinpol	443.00	NIST Webbook
rinpol	447.00	NIST Webbook
rinpol	476.60	NIST Webbook
rinpol	487.00	NIST Webbook
rinpol	502.00	NIST Webbook
rinpol	502.00	NIST Webbook
rinpol	472.00	NIST Webbook

rinpol	496.00	NIST Webbook
rinpol	496.00	NIST Webbook
rinpol	500.00	NIST Webbook
rinpol	479.00	NIST Webbook
rinpol	466.00	NIST Webbook
rinpol	468.00	NIST Webbook
rinpol	466.00	NIST Webbook
rinpol	470.00	NIST Webbook
rinpol	472.00	NIST Webbook
rinpol	473.00	NIST Webbook
rinpol	472.00	NIST Webbook
rinpol	480.00	NIST Webbook
rinpol	471.00	NIST Webbook
rinpol	466.00	NIST Webbook
rinpol	468.00	NIST Webbook
rinpol	474.00	NIST Webbook
rinpol	471.00	NIST Webbook
rinpol	488.60	NIST Webbook
rinpol	500.00	NIST Webbook
rinpol	481.00	NIST Webbook
rinpol	475.30	NIST Webbook
rinpol	500.00	NIST Webbook
rinpol	503.00	NIST Webbook
rinpol	470.00	NIST Webbook
rinpol	510.00	NIST Webbook
rinpol	477.55	NIST Webbook
rinpol	503.00	NIST Webbook
rinpol	439.00	NIST Webbook
rinpol	447.00	NIST Webbook
rinpol	450.00	NIST Webbook
rinpol	475.00	NIST Webbook
rinpol	450.00	NIST Webbook
rinpol	472.00	NIST Webbook
rinpol	443.00	NIST Webbook
rinpol	459.00	NIST Webbook
rinpol	439.00	NIST Webbook
rinpol	444.00	NIST Webbook
rinpol	465.00	NIST Webbook
rinpol	443.50	NIST Webbook
rinpol	450.00	NIST Webbook
rinpol	437.00	NIST Webbook
rinpol	437.00	NIST Webbook
rinpol	475.00	NIST Webbook
rinpol	497.00	NIST Webbook

rinpol	485.00	NIST Webbook
rinpol	488.20	NIST Webbook
rinpol	488.70	NIST Webbook
rinpol	478.00	NIST Webbook
rinpol	477.00	NIST Webbook
rinpol	484.00	NIST Webbook
rinpol	481.00	NIST Webbook
rinpol	469.00	NIST Webbook
rinpol	469.00	NIST Webbook
rinpol	470.00	NIST Webbook
rinpol	470.00	NIST Webbook
rinpol	469.41	NIST Webbook
rinpol	469.28	NIST Webbook
rinpol	469.50	NIST Webbook
rinpol	469.67	NIST Webbook
rinpol	470.10	NIST Webbook
rinpol	470.70	NIST Webbook
rinpol	470.90	NIST Webbook
rinpol	470.23	NIST Webbook
rinpol	481.00	NIST Webbook
rinpol	500.00	NIST Webbook
rinpol	441.00	NIST Webbook
ripol	846.00	NIST Webbook
ripol	823.00	NIST Webbook
ripol	814.00	NIST Webbook
ripol	825.00	NIST Webbook
ripol	798.00	NIST Webbook
ripol	845.00	NIST Webbook
ripol	814.00	NIST Webbook
ripol	832.00	NIST Webbook
ripol	816.00	NIST Webbook
ripol	820.00	NIST Webbook
ripol	816.00	NIST Webbook
ripol	811.00	NIST Webbook
ripol	825.00	NIST Webbook
ripol	828.00	NIST Webbook
ripol	827.00	NIST Webbook
ripol	827.00	NIST Webbook
ripol	810.00	NIST Webbook
ripol	820.00	NIST Webbook
ripol	823.00	NIST Webbook
ripol	811.00	NIST Webbook
ripol	812.00	NIST Webbook
ripol	821.00	NIST Webbook

ripol	841.00	NIST Webbook
ripol	832.00	NIST Webbook
ripol	834.00	NIST Webbook
ripol	808.00	NIST Webbook
ripol	808.00	NIST Webbook
ripol	830.00	NIST Webbook
ripol	812.00	NIST Webbook
ripol	805.00	NIST Webbook
ripol	813.00	NIST Webbook
ripol	819.00	NIST Webbook
ripol	814.00	NIST Webbook
ripol	813.00	NIST Webbook
ripol	822.00	NIST Webbook
ripol	816.00	NIST Webbook
ripol	811.00	NIST Webbook
ripol	818.00	NIST Webbook
ripol	813.00	NIST Webbook
ripol	802.00	NIST Webbook
ripol	814.00	NIST Webbook
ripol	821.30	NIST Webbook
ripol	814.00	NIST Webbook
ripol	842.00	NIST Webbook
ripol	834.00	NIST Webbook
ripol	813.00	NIST Webbook
ripol	809.00	NIST Webbook
ripol	794.00	NIST Webbook
ripol	818.00	NIST Webbook
ripol	820.00	NIST Webbook
ripol	814.00	NIST Webbook
ripol	814.00	NIST Webbook
ripol	814.00	NIST Webbook
ripol	813.00	NIST Webbook
ripol	821.00	NIST Webbook
ripol	824.00	NIST Webbook
ripol	842.00	NIST Webbook
ripol	785.00	NIST Webbook
ripol	847.00	NIST Webbook
ripol	832.00	NIST Webbook
ripol	840.80	NIST Webbook
ripol	837.50	NIST Webbook
ripol	835.00	NIST Webbook
ripol	843.50	NIST Webbook
ripol	820.00	NIST Webbook
ripol	810.00	NIST Webbook

ripol	810.00		NIST Webbook
ripol	822.00		NIST Webbook
ripol	810.00		NIST Webbook
ripol	822.00		NIST Webbook
ripol	854.00		NIST Webbook
ripol	816.00		NIST Webbook
ripol	800.00		NIST Webbook
ripol	845.00		NIST Webbook
ripol	814.00		NIST Webbook
ripol	814.00		NIST Webbook
ripol	775.00		NIST Webbook
ripol	821.00		NIST Webbook
ripol	841.00		NIST Webbook
ripol	836.00		NIST Webbook
ripol	816.00		NIST Webbook
ripol	813.00		NIST Webbook
ripol	815.00		NIST Webbook
ripol	814.00		NIST Webbook
ripol	819.00		NIST Webbook
ripol	788.00		NIST Webbook
ripol	823.00		NIST Webbook
ripol	847.00		NIST Webbook
ripol	835.00		NIST Webbook
ripol	810.00		NIST Webbook
ripol	824.00		NIST Webbook
ripol	815.00		NIST Webbook
ripol	820.00		NIST Webbook
ripol	821.00		NIST Webbook
ripol	820.00		NIST Webbook
ripol	816.00		NIST Webbook
ripol	810.00		NIST Webbook
ripol	847.00		NIST Webbook
ripol	819.00		NIST Webbook
ripol	810.00		NIST Webbook
ripol	813.00		NIST Webbook
ripol	830.00		NIST Webbook
ripol	809.00		NIST Webbook
ripol	821.30		NIST Webbook
ripol	811.00		NIST Webbook
ripol	794.00		NIST Webbook
ripol	821.00		NIST Webbook
ripol	818.00		NIST Webbook
sl	200.00	J/mol×K	NIST Webbook
sl	200.40	J/mol×K	NIST Webbook

sl	217.60	J/mol×K	NIST Webbook
sl	220.50	J/mol×K	NIST Webbook
tb	329.25 ± 0.20	K	NIST Webbook
tb	329.20	K	KDB
tb	329.45 ± 1.00	K	NIST Webbook
tb	329.30	K	Vapor-liquid equilibrium in the production of the ionic liquid, 1-hexyl-3-methylimidazolium bromide ([HMIm][Br]), in acetone
tb	329.45 ± 1.00	K	NIST Webbook
tb	330.85 ± 1.00	K	NIST Webbook
tb	329.50 ± 0.50	K	NIST Webbook
tb	329.50 ± 1.00	K	NIST Webbook
tb	329.26	K	Isobaric vapor-liquid equilibrium for acetone + methanol system containing different ionic liquids at 101.3 kPa
tb	328.90 ± 1.00	K	NIST Webbook
tb	329.56	K	Isobaric vapour-liquid equilibrium measurements and extractive distillation process for the azeotrope of (N,N-dimethylisopropylamine + acetone)
tb	329.33	K	Isobaric Vapor-Liquid Equilibrium Data for the Acetone + Hexamethyl Disiloxane + Ethyl Acetate Ternary System at 101.3 kPa: Determination and Correlation
tb	329.35	K	Isobaric Vapor-Liquid Equilibrium of Acetone + Methanol System in the Presence of Calcium Bromide
tb	329.27	K	Measurement of Isobaric Vapor - Liquid Equilibria of Dimethyl Carbonate with Acetone, 2-Butanone and 2-Pentanone at 101.3 kPa and Density and Speed of Sound at 298.15 K
tb	329.55 ± 1.00	K	NIST Webbook
tb	330.15 ± 1.00	K	NIST Webbook
tb	329.55 ± 1.00	K	NIST Webbook
tb	329.40 ± 1.00	K	NIST Webbook
tb	329.15 ± 1.00	K	NIST Webbook
tb	329.34 ± 0.50	K	NIST Webbook
tb	330.20 ± 1.00	K	NIST Webbook
tb	329.25 ± 0.30	K	NIST Webbook

tb	329.30 ± 1.00	K	NIST Webbook
tb	329.30 ± 0.50	K	NIST Webbook
tb	329.25 ± 0.30	K	NIST Webbook
tb	329.40 ± 1.00	K	NIST Webbook
tb	329.30 ± 0.50	K	NIST Webbook
tb	329.70 ± 1.50	K	NIST Webbook
tb	329.65 ± 1.00	K	NIST Webbook
tb	329.15 ± 1.00	K	NIST Webbook
tb	329.30 ± 0.30	K	NIST Webbook
tb	329.65 ± 1.00	K	NIST Webbook
tb	329.45 ± 0.50	K	NIST Webbook
tb	329.35 ± 1.00	K	NIST Webbook
tb	329.65 ± 1.00	K	NIST Webbook
tb	329.25 ± 1.00	K	NIST Webbook
tb	329.48 ± 1.00	K	NIST Webbook
tb	329.20 ± 0.30	K	NIST Webbook
tb	329.30 ± 1.00	K	NIST Webbook
tb	329.23 ± 0.10	K	NIST Webbook
tb	329.22 ± 0.20	K	NIST Webbook
tb	329.35 ± 1.00	K	NIST Webbook
tb	329.15 ± 1.00	K	NIST Webbook
tb	329.65 ± 1.00	K	NIST Webbook
tb	329.32 ± 0.30	K	NIST Webbook
tb	329.50 ± 0.50	K	NIST Webbook
tb	329.17 ± 0.20	K	NIST Webbook
tb	329.35 ± 0.20	K	NIST Webbook
tb	329.30 ± 0.40	K	NIST Webbook
tb	329.35 ± 0.30	K	NIST Webbook
tb	329.45 ± 1.00	K	NIST Webbook
tb	329.35 ± 0.30	K	NIST Webbook
tb	329.35 ± 0.30	K	NIST Webbook
tb	330.00 ± 1.00	K	NIST Webbook
tb	329.39 ± 0.50	K	NIST Webbook
tb	329.40 ± 0.20	K	NIST Webbook
tb	328.95 ± 1.00	K	NIST Webbook
tb	328.75 ± 1.00	K	NIST Webbook
tb	329.65 ± 0.30	K	NIST Webbook
tb	329.55 ± 1.00	K	NIST Webbook
tb	329.30 ± 1.00	K	NIST Webbook
tb	328.57 ± 1.00	K	NIST Webbook
tb	329.15 ± 2.00	K	NIST Webbook
tb	328.85 ± 1.00	K	NIST Webbook
tb	331.15 ± 3.00	K	NIST Webbook
tb	329.26 ± 0.10	K	NIST Webbook

tb	329.15 ± 1.00	K	NIST Webbook
tb	329.20 ± 1.00	K	NIST Webbook
tb	329.65 ± 1.00	K	NIST Webbook
tb	329.25 ± 1.00	K	NIST Webbook
tb	329.75 ± 0.50	K	NIST Webbook
tb	329.20 ± 0.20	K	NIST Webbook
tb	329.25 ± 1.00	K	NIST Webbook
tb	329.40 ± 1.00	K	NIST Webbook
tb	329.30 ± 0.50	K	NIST Webbook
tb	329.27 ± 0.50	K	NIST Webbook
tb	323.40 ± 0.30	K	NIST Webbook
tb	329.15 ± 1.00	K	NIST Webbook
tb	329.30 ± 0.50	K	NIST Webbook
tb	329.65 ± 1.00	K	NIST Webbook
tb	329.35 ± 0.50	K	NIST Webbook
tb	328.65 ± 1.00	K	NIST Webbook
tb	329.65 ± 0.50	K	NIST Webbook
tb	329.30 ± 0.50	K	NIST Webbook
tb	329.50 ± 0.50	K	NIST Webbook
tb	329.28 ± 0.30	K	NIST Webbook
tb	329.33 ± 0.20	K	NIST Webbook
tb	329.35 ± 0.20	K	NIST Webbook
tb	329.35 ± 0.50	K	NIST Webbook
tb	329.35 ± 0.50	K	NIST Webbook
tb	329.35 ± 0.50	K	NIST Webbook
tb	329.25 ± 0.50	K	NIST Webbook
tb	329.15 ± 1.00	K	NIST Webbook
tb	329.28 ± 0.15	K	NIST Webbook
tb	329.35 ± 1.00	K	NIST Webbook
tb	329.25 ± 0.50	K	NIST Webbook
tb	328.90 ± 0.40	K	NIST Webbook
tb	329.75 ± 1.00	K	NIST Webbook
tb	329.25 ± 1.00	K	NIST Webbook
tb	329.33 ± 0.20	K	NIST Webbook
tb	326.15 ± 1.00	K	NIST Webbook
tb	329.15 ± 1.00	K	NIST Webbook
tb	329.26 ± 0.50	K	NIST Webbook
tb	329.25 ± 0.50	K	NIST Webbook
tb	330.95 ± 1.00	K	NIST Webbook
tb	329.26 ± 0.30	K	NIST Webbook
tb	329.25 ± 0.30	K	NIST Webbook
tb	329.21 ± 0.10	K	NIST Webbook
tb	329.25 ± 0.30	K	NIST Webbook
tb	329.25 ± 0.50	K	NIST Webbook

tb	329.40 ± 0.40	K	NIST Webbook
tb	329.30 ± 0.30	K	NIST Webbook
tb	329.30 ± 0.20	K	NIST Webbook
tb	329.00 ± 1.00	K	NIST Webbook
tb	329.15 ± 0.30	K	NIST Webbook
tb	329.40 ± 0.20	K	NIST Webbook
tb	329.38 ± 0.30	K	NIST Webbook
tb	329.25 ± 0.15	K	NIST Webbook
tb	329.25	K	NIST Webbook
tb	329.30 ± 0.06	K	NIST Webbook
tb	329.30 ± 0.20	K	NIST Webbook
tb	329.00 ± 0.20	K	NIST Webbook
tb	329.26 ± 0.30	K	NIST Webbook
tb	329.25 ± 0.05	K	NIST Webbook
tb	329.23 ± 0.05	K	NIST Webbook
tb	329.47 ± 0.50	K	NIST Webbook
tb	329.25 ± 0.30	K	NIST Webbook
tb	329.40 ± 0.40	K	NIST Webbook
tb	329.45 ± 0.30	K	NIST Webbook
tb	329.30	K	NIST Webbook
tb	329.21 ± 0.04	K	NIST Webbook
tb	329.00	K	NIST Webbook
tb	329.45 ± 0.30	K	NIST Webbook
tb	329.45 ± 0.15	K	NIST Webbook
tb	329.30 ± 0.20	K	NIST Webbook
tb	330.15 ± 0.50	K	NIST Webbook
tb	329.37	K	Vapor Liquid Equilibrium for the 1,1,1-Trifluorotrichloroethane + Sulfuryl Chloride System at 101.3 kPa
tb	329.65 ± 1.00	K	NIST Webbook
tc	508.10	K	KDB
tf	178.30	K	KDB
tt	177.60 ± 0.30	K	NIST Webbook
tt	176.60 ± 0.15	K	NIST Webbook
tt	178.50 ± 0.30	K	NIST Webbook
tt	177.60 ± 0.20	K	NIST Webbook
vc	0.209	m3/kmol	KDB
zc	0.2325190		KDB
zra	0.25		KDB

Temperature Dependent Properties

Property code	Value	Unit	Temperature [K]	Source
cpg	80.58 ± 0.81	J/mol×K	332.60	NIST Webbook
cpg	92.93 ± 0.19	J/mol×K	405.20	NIST Webbook
cpg	80.96 ± 0.81	J/mol×K	334.00	NIST Webbook
cpg	81.50 ± 0.16	J/mol×K	338.20	NIST Webbook
cpg	83.35 ± 0.83	J/mol×K	347.80	NIST Webbook
cpg	83.39 ± 0.83	J/mol×K	348.00	NIST Webbook
cpg	87.03 ± 0.87	J/mol×K	363.00	NIST Webbook
cpg	87.19 ± 0.17	J/mol×K	371.20	NIST Webbook
cpg	87.53 ± 0.88	J/mol×K	372.30	NIST Webbook
cpg	89.24 ± 0.89	J/mol×K	378.00	NIST Webbook
cpg	91.84 ± 0.92	J/mol×K	393.00	NIST Webbook
cpg	94.18 ± 0.94	J/mol×K	408.00	NIST Webbook
cpg	93.30	J/mol×K	410.00	NIST Webbook
cpg	96.80 ± 1.90	J/mol×K	422.60	NIST Webbook
cpg	99.40 ± 2.00	J/mol×K	428.00	NIST Webbook
cpg	100.50 ± 2.00	J/mol×K	438.00	NIST Webbook
cpg	98.66 ± 0.20	J/mol×K	439.20	NIST Webbook
cpl	123.80	J/mol×K	298.15	NIST Webbook
cpl	133.90	J/mol×K	298.00	NIST Webbook
cpl	121.30	J/mol×K	283.00	NIST Webbook
cpl	125.90	J/mol×K	293.20	NIST Webbook
cpl	123.80	J/mol×K	298.40	NIST Webbook
cpl	124.30	J/mol×K	260.00	NIST Webbook
cpl	124.68	J/mol×K	296.99	NIST Webbook
cpl	124.70	J/mol×K	298.00	NIST Webbook
cpl	124.70	J/mol×K	298.00	NIST Webbook
cpl	128.40	J/mol×K	302.40	NIST Webbook
cpl	128.24	J/mol×K	298.00	NIST Webbook
cpl	125.56	J/mol×K	298.20	NIST Webbook
cpl	126.30	J/mol×K	293.00	NIST Webbook
cpl	129.70	J/mol×K	298.00	NIST Webbook
cpl	125.90	J/mol×K	298.15	NIST Webbook
cpl	123.80	J/mol×K	298.15	NIST Webbook
cpl	123.80	J/mol×K	298.15	NIST Webbook
cpl	126.60	J/mol×K	298.15	NIST Webbook
cpl	126.60	J/mol×K	298.15	NIST Webbook
cpl	124.70	J/mol×K	289.40	NIST Webbook
cpl	125.45	J/mol×K	298.15	NIST Webbook
cps	96.00	J/mol×K	173.00	NIST Webbook

dvisc	0.0003027	Pa×s	298.15	Excess parameter studies on the binary mixtures of toluene with ketones at different temperatures
dvisc	0.0002939	Pa×s	308.15	Density and Viscosity of (2,2-Dichloro-N,N-di-2-propenylacetamide + Acetone) and (2,2-Dichloro-N,N-di-2-propenylacetamide + Ethanol) at T = (278.15 to 313.15) K
dvisc	0.0002839	Pa×s	313.15	Density and Viscosity of (2,2-Dichloro-N,N-di-2-propenylacetamide + Acetone) and (2,2-Dichloro-N,N-di-2-propenylacetamide + Ethanol) at T = (278.15 to 313.15) K
dvisc	0.0002920	Pa×s	303.15	Physical properties of the binary systems methylcyclopentane with ketones (acetone, butanone and 2-pentanone) at T = (293.15, 298.15, and 303.15) K. New UNIFAC-VISCO interaction parameters
dvisc	0.0002809	Pa×s	308.15	Excess parameter studies on the binary mixtures of toluene with ketones at different temperatures
dvisc	0.0003200	Pa×s	293.15	Dynamic Viscosities of the Binary Systems Cyclohexane and Cyclopentane with Acetone, Butanone, or 2-Pentanone at Three Temperatures T) (293.15, 298.15, and 303.15) K

dvisc	0.0003060	Pa×s	298.15	Dynamic Viscosities of the Binary Systems Cyclohexane and Cyclopentane with Acetone, Butanone, or 2-Pentanone at Three Temperatures T) (293.15, 298.15, and 303.15) K
dvisc	0.0003200	Pa×s	293.15	Physical properties of the binary systems methylcyclopentane with ketones (acetone, butanone and 2-pentanone) at T = (293.15, 298.15, and 303.15) K. New UNIFAC-VISCO interaction parameters
dvisc	0.0003086	Pa×s	303.15	Density and Viscosity of (2,2-Dichloro-N,N-di-2-propenylacetamide + Acetone) and (2,2-Dichloro-N,N-di-2-propenylacetamide + Ethanol) at T = (278.15 to 313.15) K
dvisc	0.0003212	Pa×s	298.15	Density and Viscosity of (2,2-Dichloro-N,N-di-2-propenylacetamide + Acetone) and (2,2-Dichloro-N,N-di-2-propenylacetamide + Ethanol) at T = (278.15 to 313.15) K
dvisc	0.0003339	Pa×s	293.15	Density and Viscosity of (2,2-Dichloro-N,N-di-2-propenylacetamide + Acetone) and (2,2-Dichloro-N,N-di-2-propenylacetamide + Ethanol) at T = (278.15 to 313.15) K
dvisc	0.0003511	Pa×s	288.15	Density and Viscosity of (2,2-Dichloro-N,N-di-2-propenylacetamide + Acetone) and (2,2-Dichloro-N,N-di-2-propenylacetamide + Ethanol) at T = (278.15 to 313.15) K

dvisc	0.0003657	Pa×s	283.15	Density and Viscosity of (2,2-Dichloro-N,N-di-2-propenylacetamide + Acetone) and (2,2-Dichloro-N,N-di-2-propenylacetamide + Ethanol) at T = (278.15 to 313.15) K
dvisc	0.0002920	Pa×s	303.15	Dynamic Viscosities of the Binary Systems Cyclohexane and Cyclopentane with Acetone, Butanone, or 2-Pentanone at Three Temperatures T) (293.15, 298.15, and 303.15) K
dvisc	0.0003849	Pa×s	278.15	Density and Viscosity of (2,2-Dichloro-N,N-di-2-propenylacetamide + Acetone) and (2,2-Dichloro-N,N-di-2-propenylacetamide + Ethanol) at T = (278.15 to 313.15) K
dvisc	0.0002809	Pa×s	308.15	Density and Viscosity of Ketones with Toluene at Different Temperatures and at Atmospheric Pressure
dvisc	0.0003060	Pa×s	298.15	Physical properties of the binary systems methylcyclopentane with ketones (acetone, butanone and 2-pentanone) at T = (293.15, 298.15, and 303.15) K. New UNIFAC-VISCO interaction parameters
dvisc	0.0002918	Pa×s	303.15	Density and Viscosity of Ketones with Toluene at Different Temperatures and at Atmospheric Pressure

dvisc	0.0002918	Pa×s	303.15	Excess parameter studies on the binary mixtures of toluene with ketones at different temperatures
dvisc	0.0003027	Pa×s	298.15	Density and Viscosity of Ketones with Toluene at Different Temperatures and at Atmospheric Pressure
hfust	5.69	kJ/mol	177.60	NIST Webbook
hfust	4.77	kJ/mol	178.50	NIST Webbook
hfust	5.72	kJ/mol	176.60	NIST Webbook
hfust	5.72	kJ/mol	176.60	NIST Webbook
hfust	5.69	kJ/mol	177.60	NIST Webbook
hfust	5.71	kJ/mol	176.62	NIST Webbook
hvapt	30.70	kJ/mol	271.50	NIST Webbook
hvapt	32.10	kJ/mol	271.50	NIST Webbook
hvapt	35.00	kJ/mol	271.50	NIST Webbook
hvapt	29.09	kJ/mol	338.00	NIST Webbook
hvapt	31.10	kJ/mol	319.50	NIST Webbook
hvapt	32.60	kJ/mol	285.50	NIST Webbook
hvapt	29.10	kJ/mol	329.30	NIST Webbook
hvapt	31.80	kJ/mol	319.00	NIST Webbook
hvapt	9.20	kJ/mol	498.00	NIST Webbook
hvapt	15.30	kJ/mol	473.00	NIST Webbook
hvapt	21.70	kJ/mol	423.00	NIST Webbook
hvapt	26.10	kJ/mol	373.00	NIST Webbook
hvapt	31.90	kJ/mol	307.00	NIST Webbook
hvapt	32.70	kJ/mol	294.50	NIST Webbook
hvapt	32.80	kJ/mol	305.00	NIST Webbook
hvapt	29.70	kJ/mol	482.50	NIST Webbook
hvapt	30.60	kJ/mol	351.00	NIST Webbook
hvapt	33.80	kJ/mol	236.00	NIST Webbook
hvapt	32.90	kJ/mol	210.50	NIST Webbook
hvapt	29.90	kJ/mol	408.50	NIST Webbook
hvapt	32.10	kJ/mol	308.00	NIST Webbook
hvapt	29.50	kJ/mol	419.00	NIST Webbook
hvapt	29.12	kJ/mol	329.40	KDB

pvap	57.96	kPa	314.00	Isobaric vapor liquid equilibria for acetone + methanol + lithium nitrate at 100 kPa
pvap	60.29	kPa	314.76	Vapor-Liquid Equilibrium for Binary Systems of Diacetyl with Methanol and Acetone
pvap	64.95	kPa	317.00	Isobaric vapor liquid equilibria for acetone + methanol + lithium nitrate at 100 kPa
pvap	72.60	kPa	320.00	Isobaric vapor liquid equilibria for acetone + methanol + lithium nitrate at 100 kPa
pvap	80.96	kPa	323.00	Isobaric vapor liquid equilibria for acetone + methanol + lithium nitrate at 100 kPa
pvap	90.08	kPa	326.00	Isobaric vapor liquid equilibria for acetone + methanol + lithium nitrate at 100 kPa
pvap	100.00	kPa	329.00	Isobaric vapor liquid equilibria for acetone + methanol + lithium nitrate at 100 kPa
pvap	110.78	kPa	332.00	Isobaric vapor liquid equilibria for acetone + methanol + lithium nitrate at 100 kPa
pvap	122.47	kPa	335.00	Isobaric vapor liquid equilibria for acetone + methanol + lithium nitrate at 100 kPa
pvap	101.33	kPa	329.30	Vapor-liquid equilibrium in the production of the ionic liquid, 1-hexyl-3-methylimidazolium bromide ([HMIm][Br]), in acetone

pvap	101.30	kPa	329.26	Isobaric vapor-liquid equilibrium for acetone + methanol system containing different ionic liquids at 101.3 kPa
pvap	115.70	kPa	333.15	Thermodynamics of binary mixtures of N-methyl-2-pyrrolidinone and ketone. Experimental results and modelling of the solid-liquid equilibrium and vapou-liquid equilibrium. The Modified UNIFAC (Do) model characterization
pvap	101.30	kPa	329.56	Isobaric vapour-liquid equilibrium measurements and extractive distillation process for the azeotrope of (N,N-dimethylisopropylamine + acetone)
pvap	96.15	kPa	327.90	Vapor Liquid Equilibrium Data for Binary Mixtures of Acetic Acid + Anisole, Acetone + Anisole, and Isopropanol + Anisole at Pressure 96.15 kPa
pvap	68.20	kPa	318.15	Isothermal Vapor-Liquid Equilibria for Binary Mixtures of Methyl Nonafluorobutyl Ether + Acetone, Cyclopentyl Methyl Ether, Ethyl Acetate, n-Heptane, Methanol, and Toluene

pvap	101.30	kPa	329.33	Isobaric Vapor-Liquid Equilibrium Data for the Acetone + Hexamethyl Disiloxane + Ethyl Acetate Ternary System at 101.3 kPa: Determination and Correlation
pvap	101.00	kPa	329.35	Isobaric Vapor-Liquid Equilibrium of Acetone + Methanol System in the Presence of Calcium Bromide
pvap	297.10	kPa	364.51	Vapor Liquid Equilibrium for Six Binary Systems of C4-Hydrocarbons + 2-Propanone
pvap	70.07	kPa	318.77	Vapor Liquid Equilibrium for Binary Systems of 2,3-Pentanedione with Diacetyl and Acetone
pvap	65.08	kPa	316.80	Vapor Liquid Equilibrium for Binary Systems of 2,3-Pentanedione with Diacetyl and Acetone
pvap	60.29	kPa	314.76	Vapor Liquid Equilibrium for Binary Systems of 2,3-Pentanedione with Diacetyl and Acetone
pvap	55.29	kPa	312.52	Vapor Liquid Equilibrium for Binary Systems of 2,3-Pentanedione with Diacetyl and Acetone
pvap	50.40	kPa	310.08	Vapor Liquid Equilibrium for Binary Systems of 2,3-Pentanedione with Diacetyl and Acetone

pvap	45.21	kPa	307.33	Vapor Liquid Equilibrium for Binary Systems of 2,3-Pentanedione with Diacetyl and Acetone
pvap	35.22	kPa	301.25	Vapor Liquid Equilibrium for Binary Systems of 2,3-Pentanedione with Diacetyl and Acetone
pvap	30.23	kPa	297.67	Vapor Liquid Equilibrium for Binary Systems of 2,3-Pentanedione with Diacetyl and Acetone
pvap	101.30	kPa	329.37	Vapor Liquid Equilibrium for the 1,1,1-Trifluorotrichloroethane + Sulfuryl Chloride System at 101.3 kPa
pvap	100.00	kPa	329.00	1-Ethyl-3-methylimidazolium Dicyanamide as a Very Efficient Entrainer for the Extractive Distillation of the Acetone + Methanol System
pvap	100.00	kPa	329.00	Influence of Some Ionic Liquids Containing the Trifluoromethanesulfonate Anion on the Vapor Liquid Equilibria of the Acetone + Methanol System
pvap	81.65	kPa	323.15	Vapor-Liquid Equilibrium for Binary Systems of Diacetyl with Methanol and Acetone
pvap	70.07	kPa	318.77	Vapor-Liquid Equilibrium for Binary Systems of Diacetyl with Methanol and Acetone

pvap	65.08	kPa	316.80	Vapor-Liquid Equilibrium for Binary Systems of Diacetyl with Methanol and Acetone
pvap	294.30	kPa	364.51	Vapor Liquid Equilibrium for Six Binary Systems of C4-Hydrocarbons + 2-Propanone
pvap	56.64	kPa	313.15	Vapor-Liquid Equilibrium for Binary Systems of Diacetyl with Methanol and Acetone
pvap	55.29	kPa	312.52	Vapor-Liquid Equilibrium for Binary Systems of Diacetyl with Methanol and Acetone
pvap	50.40	kPa	310.08	Vapor-Liquid Equilibrium for Binary Systems of Diacetyl with Methanol and Acetone
pvap	45.21	kPa	307.33	Vapor-Liquid Equilibrium for Binary Systems of Diacetyl with Methanol and Acetone
pvap	40.00	kPa	304.37	Vapor-Liquid Equilibrium for Binary Systems of Diacetyl with Methanol and Acetone
pvap	38.04	kPa	303.15	Vapor-Liquid Equilibrium for Binary Systems of Diacetyl with Methanol and Acetone
pvap	35.22	kPa	301.25	Vapor-Liquid Equilibrium for Binary Systems of Diacetyl with Methanol and Acetone
pvap	30.23	kPa	297.66	Vapor-Liquid Equilibrium for Binary Systems of Diacetyl with Methanol and Acetone

pvap	302.56	kPa	365.43	Vapor Liquid Equilibrium for Six Binary Systems of C4-Hydrocarbons + 2-Propanone
pvap	296.80	kPa	364.52	Vapor Liquid Equilibrium for Six Binary Systems of C4-Hydrocarbons + 2-Propanone
pvap	296.40	kPa	364.51	Vapor Liquid Equilibrium for Six Binary Systems of C4-Hydrocarbons + 2-Propanone
pvap	135.11	kPa	338.00	Isobaric vapor liquid equilibria for acetone + methanol + lithium nitrate at 100 kPa
rfi	1.35860		293.15	Isobaric Vapor Liquid Equilibrium for Nine Binary Systems of Cracking C5 Fraction at 250 kPa
rfi	1.35890		293.15	Isothermal and Isobaric Vapor-Liquid Equilibria of the Ternary System of 2,2-Dimethoxypropane + Acetone + Methanol
rfi	1.35057		308.15	Densities, Refractive indices and Viscosities for Binary and Ternary Mixtures of Acetone, Ethanol and 2,2,4-Trimethylpentane at T = (288.15, 298.15, and 308.15) K

rfi	1.35605	298.15	Densities, Refractive indices and Viscosities for Binary and Ternary Mixtures of Acetone, Ethanol and 2,2,4-Trimethylpentane at T = (288.15, 298.15, and 308.15) K
rfi	1.36152	288.15	Densities, Refractive indices and Viscosities for Binary and Ternary Mixtures of Acetone, Ethanol and 2,2,4-Trimethylpentane at T = (288.15, 298.15, and 308.15) K
rfi	1.35553	298.15	Quaternary, Ternary and Binary LLE Measurements for 2-Methoxy-2-methylbutane + Furfural + Acetic Acid + Water at Temperatures between 298 and 341 K
rfi	1.35553	298.15	Quaternary, Ternary, and Binary LLE Measurements for 2-Methoxy-2-methylpropane + Furfural + Acetic Acid + Water at Temperatures between 298 and 307 K
rfi	1.35553	298.15	Ternary and Binary LLE Measurements for Solvent (4-Methyl-2-pentanone and 2-Methyl-2-butanol) + Furfural + Water between 298 and 401 K
rfi	1.35900	293.15	Solubilities of Some Phosphaspirocyclic Compounds in Selected Solvents

rfi	1.35620	298.15	Solubility of a-Carotene in Binary Solvents Formed by Some Hydrocarbons with 2,5,8-Trioxanonane, 2-Propanone, and Cyclohexanone
rfi	1.35550	298.00	Quaternary and ternary LLE measurements for solvent (2-methyltetrahydrofuran and cyclopentyl methyl ether) + furfural + acetic acid + water between 298 and 343 K
rfi	1.35900	293.15	Solubilities of (2,5-Dihydroxyphenyl)diphenyl Phosphine Oxide in Selected Solvents
rfi	1.35550	298.15	Ternary and binary LLE measurements for solvent (2-methyltetrahydrofuran and cyclopentyl methyl ether) + furfural + water between 298 and 343 K
rfi	1.34491	318.15	Density, Speed of Sound, and Refractive Index of 1-Ethyl-3-methylimidazolium Trifluoromethanesulfonate with Acetone, Methyl Acetate, and Ethyl Acetate at Temperatures from (278.15 to 328.15) K
rfi	1.35054	308.15	Density, Speed of Sound, and Refractive Index of 1-Ethyl-3-methylimidazolium Trifluoromethanesulfonate with Acetone, Methyl Acetate, and Ethyl Acetate at Temperatures from (278.15 to 328.15) K

rfi	1.35597	298.15	Density, Speed of Sound, and Refractive Index of 1-Ethyl-3-methylimidazolium Trifluoromethanesulfonate with Acetone, Methyl Acetate, and Ethyl Acetate at Temperatures from (278.15 to 328.15) K
rfi	1.36146	288.15	Density, Speed of Sound, and Refractive Index of 1-Ethyl-3-methylimidazolium Trifluoromethanesulfonate with Acetone, Methyl Acetate, and Ethyl Acetate at Temperatures from (278.15 to 328.15) K
rfi	1.35566	298.15	Properties of ionic liquid HMIMPF6 with carbonates, ketones and alkyl acetates
rfi	1.35880	293.15	Isothermal vapour-liquid equilibrium data for the binary systems 2-propanone + (2-butanol or propanoic acid)
rfi	1.35732	298.15	Phase equilibria and interfacial tensions in the systems methyl tert-butyl ether + acetone + cyclohexane, methyl tert-butyl ether + acetone and methyl tert-butyl ether + cyclohexane
rfi	1.35732	298.15	Vapor liquid equilibria and interfacial tensions for the ternary system acetone + 2,2 -oxybis[propane] + cyclohexane and its constituent binary systems

rfi	1.35597		298.15	Isobaric vapor liquid equilibria for mixtures of acetone, ethanol, and 2,2,4-trimethylpentane at 101.3 kPa
rfi	1.35900		293.15	Solubilities of Phosphorus-Containing Compounds in Selected Solvents
rfi	1.35820		298.15	Isobaric Vapor Liquid Equilibria for Binary Systems of Acetone + Isopropenyl Acetate, 2-Butanone + Isopropenyl Acetate, and Isopropenyl Acetate + Acetylacetone at 101.3 kPa
rfi	1.35900		293.15	Solubilities of 3,9-Dimethyl-3,9-dioxide-2,4,8,10-tetraoxa-3,9-diphosphaspiro[5.5]undec in Selected Solvents
rfi	1.35880		293.15	Solubilities of Methyldiphenylphosphine Oxide in Selected Solvents
rfi	1.35600		298.15	Physical properties and their corresponding changes of mixing for the ternary mixture acetone + n-hexane +water at 298.15K
rfi	1.35900		293.15	Solubilities of 2-(6-Oxido-6H-dibenz[c,e][1,2]oxaphosphorin-6-yl)-1,4-dihydroxy Phenylene in the Selected Solvents
rhol	785.00	kg/m3	298.15	Vapor Liquid Equilibrium Data for 2,3-Pentanedione + (Acetaldehyde or Acetone) at (100, 150, and 200) kPa

rhol	790.31	kg/m3	293.15	Excess Molar Enthalpies for Binary Mixtures of Ethanol + Acetone, + Octane, + Cyclohexane and 1-Propanol + Acetone, + Octane, + Heptane at 323.15
rhol	801.67	kg/m3	283.15	Experimental Densities and Excess Volumes for Binary Mixtures Containing Propionic Acid, Acetone and Water from 283.15 K to 323.15 K at Atmospheric Pressure
rhol	767.00	kg/m3	313.15	Thermophysical approach to understand the nature of molecular interactions and structural factor between methyl isobutyl ketone and organic solvents mixtures
rhol	779.00	kg/m3	303.15	Thermophysical approach to understand the nature of molecular interactions and structural factor between methyl isobutyl ketone and organic solvents mixtures
rhol	790.00	kg/m3	293.15	Thermophysical approach to understand the nature of molecular interactions and structural factor between methyl isobutyl ketone and organic solvents mixtures

rhol	778.63	kg/m3	303.15	Thermodynamics of amide + ketone mixtures. 1. Volumetric, speed of sound and refractive index data for N,N-dimethylformamide + 2-alkanone systems at several temperatures
rhol	784.43	kg/m3	298.15	Thermodynamics of amide + ketone mixtures. 1. Volumetric, speed of sound and refractive index data for N,N-dimethylformamide + 2-alkanone systems at several temperatures
rhol	790.19	kg/m3	293.15	Thermodynamics of amide + ketone mixtures. 1. Volumetric, speed of sound and refractive index data for N,N-dimethylformamide + 2-alkanone systems at several temperatures
rhol	785.10	kg/m3	298.15	Measurement and correlation of solubility and solution thermodynamics of 1,3-dimethylurea in different solvents from T = (288.15 to 328.15) K
rhol	778.70	kg/m3	303.20	Ionic liquid 1-hexyl-3-methylimidazolium hexafluorophosphate, an efficient solvent for extraction of acetone from aqueous solutions

rhol	796.04	kg/m3	288.15	Experimental Densities and Excess Volumes for Binary Mixtures Containing Propionic Acid, Acetone and Water from 283.15 K to 323.15 K at Atmospheric Pressure
rhol	790.36	kg/m3	293.15	Experimental Densities and Excess Volumes for Binary Mixtures Containing Propionic Acid, Acetone and Water from 283.15 K to 323.15 K at Atmospheric Pressure
rhol	790.66	kg/m3	293.20	Ionic liquid 1-hexyl-3-methylimidazolium hexafluorophosphate, an efficient solvent for extraction of acetone from aqueous solutions
rhol	790.21	kg/m3	293.15	(Liquid + liquid) equilibria for (water + 1-propanol or acetone + .beta.-citronellol) at different temperatures
rhol	783.80	kg/m3	298.15	Solubility and solution thermodynamics of sorbic acid in eight pure organic solvents
rhol	785.32	kg/m3	298.15	Extraction desulfurization process of fuels with ionic liquids
rhol	785.32	kg/m3	298.15	Effect of the alkyl side chain of the 1-alkylpiperidinium-based ionic liquids on desulfurization of fuels

rhol	785.23	kg/m3	298.15	Separation of sulfur compounds from alkanes with 1-alkylcyanopyridinium-based ionic liquids
rhol	773.07	kg/m3	308.15	Thermodynamic properties of binary mixtures of the ionic liquid [emim][BF4] with acetone and dimethylsulphoxide
rhol	784.40	kg/m3	298.20	Ionic liquid 1-hexyl-3-methylimidazolium hexafluorophosphate, an efficient solvent for extraction of acetone from aqueous solutions
rhol	784.40	kg/m3	298.15	Thermodynamic properties of binary mixtures of the ionic liquid [emim][BF4] with acetone and dimethylsulphoxide
rhol	789.99	kg/m3	293.15	Thermodynamic properties of binary mixtures of the ionic liquid [emim][BF4] with acetone and dimethylsulphoxide
rhol	784.24	kg/m3	298.15	Apparent molar volumes and compressibilities of tetrabutyl-ammonium bromide in organic solvents
rhol	784.23	kg/m3	298.15	Density and speed of sound of lithium bromide with organic solvents: Measurement and correlation
rhol	784.65	kg/m3	298.15	Excess molar enthalpies and volumes of binary mixtures of nonafluorobutylmethylether with ketones at T = 298.15 K
rhol	784.50	kg/m3	293.15	(Vapour + liquid) equilibria for (2-ethoxypropene + acetone) and (2-ethoxypropene + butanone)

rhol	761.26	kg/m3	318.15	Densities, viscosities, and refractive indices of binary and ternary mixtures of methanol, acetone, and chloroform at temperatures from (298.15-318.15) K and ambient pressure	
rhol	766.95	kg/m3	313.15	Densities, viscosities, and refractive indices of binary and ternary mixtures of methanol, acetone, and chloroform at temperatures from (298.15-318.15) K and ambient pressure	
rhol	784.64	kg/m3	298.15	Experimental Densities and Excess Volumes for Binary Mixtures Containing Propionic Acid, Acetone and Water from 283.15 K to 323.15 K at Atmospheric Pressure	
rhol	773.07	kg/m3	308.15	Experimental Densities and Excess Volumes for Binary Mixtures Containing Propionic Acid, Acetone and Water from 283.15 K to 323.15 K at Atmospheric Pressure	

rhol	773.08	kg/m3	308.15	Densities, viscosities, and refractive indices of binary and ternary mixtures of methanol, acetone, and chloroform at temperatures from (298.15-318.15) K and ambient pressure	
rhol	767.21	kg/m3	313.14	Experimental Densities and Excess Volumes for Binary Mixtures Containing Propionic Acid, Acetone and Water from 283.15 K to 323.15 K at Atmospheric Pressure	
rhol	761.29	kg/m3	318.14	Experimental Densities and Excess Volumes for Binary Mixtures Containing Propionic Acid, Acetone and Water from 283.15 K to 323.15 K at Atmospheric Pressure	
rhol	755.31	kg/m3	323.14	Experimental Densities and Excess Volumes for Binary Mixtures Containing Propionic Acid, Acetone and Water from 283.15 K to 323.15 K at Atmospheric Pressure	
rhol	784.39	kg/m3	298.15	Excess Molar Volumes and Surface Tensions of Xylene with Acetone or 2-Butanone at 298.15 K	

rhol	785.09	kg/m3	298.15	Thermodynamics of Ketone + Amine Mixtures. Part VIII. Molar Excess Enthalpies at 298.15 K for n-Alkanone + Aniline or + N-Methylaniline Systems
rhol	784.30	kg/m3	293.15	Liquid Liquid Equilibrium for the Ternary System Acetone Oxime Methyl Ether Acetone Water
rhol	784.30	kg/m3	298.15	Phase equilibrium study of binary and ternary mixtures of ionic liquids + acetone + methanol
rhol	778.66	kg/m3	303.15	Densities, viscosities, and refractive indices of binary and ternary mixtures of methanol, acetone, and chloroform at temperatures from (298.15-318.15) K and ambient pressure
rhol	785.23	kg/m3	298.15	Separation of pyridine from heptane with tricyanomethanide-based ionic liquids
rhol	785.09	kg/m3	298.15	Thermodynamics of ketone + amine mixtures. Part X. Excess molarenthalpies at 298.15 K for N,N,N-triethylamine + 2-alkanone systems.Characterization of tertiary amine + 2-alkanone, and ofamino-ketone + n-alkane mixtures in terms of DISQUAC

rhol	785.09	kg/m3	298.15	Thermodynamics of ketone + amine mixtures. Part IX. Excess molar enthalpies at 298.15K for dipropylamine, or dibutylamine + 2-alkanone systems and modeling of linear or aromatic amine + 2-alkanone mixtures in terms of DISQUAC and ERAS
rhol	790.00	kg/m3	293.00	KDB
rhol	786.77	kg/m3	298.15	Separation of ethylbenzene/styrene systems using ionic liquids in ternary LLE
rhol	783.81	kg/m3	298.15	Partial Molar Volumes of Butyltriethylammonium Iodide in Single Nonaqueous Solvents at 298.15 K
rhol	766.90	kg/m3	313.15	Density, Speed of Sound, and Derived Thermodynamic Properties of Ionic Liquids [EMIM]+[BETI]- or ([EMIM]+[CH3(OCH2CH2)2OSO3]- + Methanol or + Acetone) at T = (298.15 or 303.15 or 313.15) K
rhol	778.40	kg/m3	303.15	Density, Speed of Sound, and Derived Thermodynamic Properties of Ionic Liquids [EMIM]+[BETI]- or ([EMIM]+[CH3(OCH2CH2)2OSO3]- + Methanol or + Acetone) at T = (298.15 or 303.15 or 313.15) K

rhol	784.10	kg/m3	298.15	Density, Speed of Sound, and Derived Thermodynamic Properties of Ionic Liquids [EMIM]+[BETI]- or ([EMIM]+[CH3(OCH2CH2)2OSO3]- + Methanol or + Acetone) at T = (298.15 or 303.15 or 313.15) K
rhol	784.41	kg/m3	298.15	Densities, viscosities, and refractive indices of binary and ternary mixtures of methanol, acetone, and chloroform at temperatures from (298.15-318.15) K and ambient pressure
rhol	778.87	kg/m3	303.15	Thermodynamic properties of binary mixtures of the ionic liquid [emim][BF4] with acetone and dimethylsulphoxide
rhol	778.88	kg/m3	303.14	Experimental Densities and Excess Volumes for Binary Mixtures Containing Propionic Acid, Acetone and Water from 283.15 K to 323.15 K at Atmospheric Pressure
sfust	26.70	J/mol×K	178.50	NIST Webbook
sfust	32.00	J/mol×K	177.60	NIST Webbook
sfust	32.36	J/mol×K	176.62	NIST Webbook
sfust	32.03	J/mol×K	177.60	NIST Webbook
speedsl	1162.00	m/s	298.15	Vapor liquid equilibria for systems of diethyl carbonate and ketones and determination of group interaction parameters for the UNIFAC and ASOG methods

srf	0.02	N/m	328.15	Surface Tension of the Ternary System Water + Acetone + Toluene
srf	0.02	N/m	308.15	Surface Tension of the Ternary System Water + Acetone + Toluene
srf	0.02	N/m	298.15	Surface Tension of the Ternary System Water + Acetone + Toluene
srf	0.02	N/m	288.15	Surface Tension of the Ternary System Water + Acetone + Toluene
srf	0.02	N/m	327.88	Surface Tension of Pure Liquids and Binary Liquid Mixtures
srf	0.02	N/m	317.86	Surface Tension of Pure Liquids and Binary Liquid Mixtures
srf	0.02	N/m	307.86	Surface Tension of Pure Liquids and Binary Liquid Mixtures
srf	0.02	N/m	297.82	Surface Tension of Pure Liquids and Binary Liquid Mixtures
srf	0.02	N/m	287.81	Surface Tension of Pure Liquids and Binary Liquid Mixtures
srf	0.02	N/m	298.20	KDB
srf	0.02	N/m	318.15	Surface Tension of the Ternary System Water + Acetone + Toluene
srf	0.02	N/m	293.15	Investigation of Surface Properties and Solubility of 1-Vinyl-3-alkyl/Esterimidazolium Halide Ionic Liquids by Density Functional Methods

Pressure Dependent Properties

Property code	Value	Unit	Pressure [kPa]	Source
tbrp	329.30	K	2.70	NIST Webbook

Correlations

Information	Value
Property code	pvap
Equation	ln(Pvp) = A + B/(T + C)
Coeff. A	1.48340e+01
Coeff. B	-3.02945e+03
Coeff. C	-3.26710e+01
Temperature range (K), min.	240.93
Temperature range (K), max.	508.10

Information	Value
Property code	pvap
Equation	ln(Pvp) = A + B/T + C*ln(T) + D*T^2
Coeff. A	6.69693e+01
Coeff. B	-5.78459e+03
Coeff. C	-7.85881e+00
Coeff. D	7.10496e-06
Temperature range (K), min.	178.45
Temperature range (K), max.	508.20

Datasets

Viscosity, Pa*s

Temperature, K - Liquid	Pressure, kPa - Liquid	Viscosity, Pa*s - Liquid
303.15	101.30	0.0002974

Speed of sound, m/s

Temperature, K - Liquid	Pressure, kPa - Liquid	Speed of sound, m/s - Liquid
265.67	102.00	1298.93
265.67	504.00	1301.09
265.67	750.00	1302.5
265.67	1003.00	1303.56
265.67	1255.00	1304.92
265.67	1501.00	1306.29
265.67	2007.00	1308.73
265.67	2506.00	1311.28
265.67	3005.00	1313.79
265.67	4004.00	1318.59
265.67	5004.00	1323.31
265.67	6005.00	1328.01
265.67	8002.00	1338.69
265.67	10001.00	1349.81
265.67	12001.00	1358.7
265.67	14003.00	1366.48
265.67	16050.00	1375.48
265.67	18001.00	1384.77
265.67	20003.00	1394.13
265.67	23001.00	1407.08
265.67	27004.00	1425.32
265.67	31003.00	1439.78
265.67	34997.00	1455.17
265.67	39999.00	1476.05
265.67	45002.00	1494.61
265.67	49999.00	1511.6
265.67	55001.00	1528.77
265.67	60002.00	1548.05
265.67	65000.00	1564.17
265.67	18001.00	1384.54
273.16	102.00	1264.48
273.16	504.00	1266.57
273.16	750.00	1267.96
273.16	1003.00	1269.39
273.16	1255.00	1270.74

273.16	1501.00	1272.23
273.16	2007.00	1275.12
273.16	2506.00	1277.93
273.16	3005.00	1281.1
273.16	4004.00	1287.23
273.16	5004.00	1291.97
273.16	6005.00	1296.6
273.16	8002.00	1306.58
273.16	10001.00	1316.19
273.16	12001.00	1325.58
273.16	14003.00	1335.9
273.16	16005.00	1346.87
273.16	18001.00	1355.99
273.16	20003.00	1363.63
273.16	23001.00	1376.52
273.16	27004.00	1394.61
273.16	31003.00	1412.24
273.16	34997.00	1428.86
273.16	39999.00	1446.67
273.16	45002.00	1467.44
273.16	49999.00	1485.84
273.16	55001.00	1504.31
273.16	60002.00	1520.08
273.16	65000.00	1538.71
273.16	3005.00	1281.19
273.16	4004.00	1286.78
273.16	5004.00	1292.02
280.74	75.00	1233.63
280.74	504.00	1234.94
280.74	750.00	1236.39
280.74	1003.00	1237.85
280.74	1255.00	1239.22
280.74	1501.00	1240.59
280.74	2007.00	1243.75
280.74	2506.00	1247.03
280.74	3005.00	1249.76
280.74	4004.00	1255.13
280.74	5004.00	1259.87
280.74	6005.00	1264.8
280.74	8002.00	1275.79
280.74	10001.00	1286.82
280.74	12001.00	1296.79
280.74	14003.00	1306.71
280.74	16005.00	1316.23

280.74	18001.00	1325.46
280.74	20003.00	1335.09
280.74	23001.00	1350.8
280.74	27004.00	1366.99
280.74	31003.00	1384.26
280.74	34997.00	1401.32
280.74	39999.00	1423.56
280.74	45002.00	1441.37
280.74	49999.00	1461.28
280.74	55001.00	1480.44
280.74	60002.00	1498.97
280.74	65000.00	1514.56
280.74	75.00	1232.04
280.74	23001.00	1350.8
280.74	39999.00	1423.74
280.74	45002.00	1441.37
280.74	60002.00	1498.91
280.74	65000.00	1514.63
288.23	109.00	1199.01
288.23	504.00	1201.24
288.23	750.00	1202.57
288.23	1003.00	1204.2
288.23	1255.00	1205.54
288.23	1501.00	1206.97
288.23	2007.00	1209.92
288.23	2506.00	1212.8
288.23	3005.00	1215.61
288.23	4004.00	1221.35
288.23	5004.00	1227.14
288.23	6005.00	1232.71
288.23	8002.00	1244.39
288.23	10001.00	1255.4
288.23	12001.00	1265.22
288.23	14003.00	1275.63
288.23	16005.00	1286.61
288.23	18001.00	1296.18
288.23	20003.00	1305.89
288.23	23001.00	1319.61
288.23	27004.00	1339.05
288.23	31003.00	1357.37
288.23	34997.00	1373.83
288.23	39999.00	1395.08
288.23	45002.00	1416.1
288.23	49999.00	1435.37

288.23	55001.00	1453.97
288.23	60002.00	1473.72
288.23	65000.00	1492.08
298.23	109.00	1153.69
298.23	504.00	1156.23
298.23	750.00	1157.82
298.23	1003.00	1159.5
298.23	1255.00	1161.14
298.23	1501.00	1162.71
298.23	2007.00	1165.65
298.23	2506.00	1168.6
298.23	3005.00	1171.49
298.23	4004.00	1177.4
298.23	5004.00	1182.9
298.23	6005.00	1189.6
298.23	8002.00	1201.45
298.23	10001.00	1212.8
298.23	12001.00	1223.83
298.23	14003.00	1234.83
298.23	16005.00	1246.18
298.23	18001.00	1256.46
298.23	20003.00	1265.92
298.23	23001.00	1281.93
298.23	27004.00	1301.17
298.23	31003.00	1319.28
298.23	34997.00	1337.9
298.23	39999.00	1359.25
298.23	45002.00	1380.44
298.23	49999.00	1400.61
298.23	55001.00	1421.32
298.23	60002.00	1439.68
298.23	65000.00	1458.59
308.22	109.00	1107.74
308.22	504.00	1110.52
308.22	750.00	1112.35
308.22	1003.00	1114.01
308.22	1255.00	1115.75
308.22	1501.00	1117.34
308.22	2007.00	1120.69
308.22	2506.00	1123.91
308.22	3005.00	1127.11
308.22	4004.00	1133.79
308.22	5004.00	1140.56
308.22	6005.00	1146.98

308.22	8002.00	1159.51
308.22	10001.00	1171.13
308.22	12001.00	1182.37
308.22	14003.00	1194.75
308.22	16005.00	1205.76
308.22	18001.00	1216.57
308.22	20003.00	1227.49
308.22	23001.00	1244.05
308.22	27004.00	1263.44
308.22	31003.00	1283.97
308.22	34997.00	1302.46
308.22	39999.00	1324.45
308.22	45002.00	1347.09
308.22	49999.00	1367.79
308.22	55001.00	1388.31
308.22	60002.00	1408.07
308.22	65000.00	1427.73
318.22	109.00	1062.32
318.22	504.00	1065.38
318.22	750.00	1067.34
318.22	1003.00	1069.21
318.22	1255.00	1071.11
318.22	1501.00	1072.72
318.22	2007.00	1076.32
318.22	2506.00	1079.71
318.22	3005.00	1083.5
318.22	4004.00	1090.38
318.22	5004.00	1096.85
318.22	6005.00	1103.61
318.22	8002.00	1116.61
318.22	10001.00	1129.13
318.22	12001.00	1142.47
318.22	14003.00	1154.62
318.22	16005.00	1166.43
318.22	18001.00	1177.61
318.22	20003.00	1189.14
318.22	23001.00	1205.63
318.22	27004.00	1226.77
318.22	31003.00	1247.41
318.22	34997.00	1266.28
318.22	39999.00	1290.76
318.22	45002.00	1313.78
318.22	49999.00	1334.83
318.22	55001.00	1356.46

318.22	60002.00	1376.77
318.22	65000.00	1395.43
328.22	129.00	1018.3
328.22	504.00	1021.14
328.22	750.00	1023.1
328.22	1003.00	1024.91
328.22	1255.00	1026.82
328.22	1501.00	1028.6
328.22	2007.00	1032.33
328.22	2506.00	1036.21
328.22	3005.00	1040.05
328.22	4004.00	1047.07
328.22	5004.00	1053.76
328.22	6005.00	1060.7
328.22	8002.00	1074.73
328.22	10001.00	1088.42
328.22	12001.00	1101.14
328.22	14003.00	1113.97
328.22	16005.00	1126.49
328.22	18001.00	1138.92
328.22	20003.00	1151.08
328.22	23001.00	1168.1
328.22	27004.00	1190.45
328.22	31003.00	1211.66
328.22	34997.00	1231.66
328.22	39999.00	1256.67
328.22	45002.00	1280.63
328.22	49999.00	1303.34
328.22	55001.00	1324.67
328.22	60002.00	1345.81
328.22	65000.00	1365.57
338.22	170.00	972.8
338.22	504.00	976.32
338.22	750.00	978.88
338.22	1003.00	981.05
338.22	1255.00	982.97
338.22	1501.00	984.75
338.22	2007.00	988.42
338.22	2506.00	992.07
338.22	3005.00	996.0
338.22	4004.00	1004.26
338.22	5004.00	1011.63
338.22	6005.00	1019.16
338.22	8002.00	1033.34

338.22	10001.00	1047.72
338.22	12001.00	1061.19
338.22	14003.00	1074.71
338.22	16005.00	1088.09
338.22	18001.00	1100.33
338.22	20003.00	1112.58
338.22	23001.00	1131.02
338.22	27004.00	1154.76
338.22	31003.00	1176.19
338.22	34997.00	1197.57
338.22	39999.00	1223.29
338.22	45002.00	1248.14
338.22	49999.00	1271.19
338.22	55001.00	1293.31
338.22	60002.00	1315.39
338.22	65000.00	1336.38

Reference *https://www.doi.org/10.1016/j.jct.2003.12.001*

Temperature, K	Pressure, kPa	Speed of sound, m/s
248.15	100.00	1390.16
248.15	10000.00	1435.51
248.15	20000.00	1478.36
248.15	30000.00	1518.03
248.15	40000.00	1550.78
248.15	50000.00	1590.88
248.15	60000.00	1625.12
248.15	70000.00	1657.38
248.15	80000.00	1688.07
248.15	90000.00	1717.48
248.15	100000.00	1745.43
253.15	100.00	1367.16
253.15	10000.00	1413.84
253.15	20000.00	1457.41
253.15	30000.00	1497.98
253.15	40000.00	1536.44
253.15	50000.00	1572.26
253.15	60000.00	1606.9
253.15	70000.00	1639.52
253.15	80000.00	1670.88
253.15	90000.00	1700.93
253.15	100000.00	1729.65
258.15	100.00	1344.2

258.15	10000.00	1391.98
258.15	20000.00	1436.58
258.15	30000.00	1478.11
258.15	40000.00	1517.33
258.15	50000.00	1553.89
258.15	60000.00	1588.88
258.15	70000.00	1621.91
258.15	80000.00	1653.89
258.15	90000.00	1684.43
258.15	100000.00	1713.62
263.15	100.00	1320.92
263.15	10000.00	1370.35
263.15	20000.00	1415.99
263.15	30000.00	1458.25
263.15	40000.00	1498.44
263.15	50000.00	1535.81
263.15	60000.00	1570.98
263.15	70000.00	1604.54
263.15	80000.00	1637.11
263.15	90000.00	1668.1
263.15	100000.00	1697.74
268.15	100.00	1298.67
268.15	10000.00	1349.08
268.15	20000.00	1395.62
268.15	30000.00	1438.54
268.15	40000.00	1479.77
268.15	50000.00	1517.83
268.15	60000.00	1553.2
268.15	70000.00	1587.43
268.15	80000.00	1620.54
268.15	90000.00	1651.96
268.15	100000.00	1682.01
273.15	100.00	1275.4
273.15	10000.00	1327.81
273.15	20000.00	1375.45
273.15	30000.00	1418.96
273.15	40000.00	1461.31
273.15	50000.00	1499.98
273.15	60000.00	1535.77
273.15	70000.00	1570.58
273.15	80000.00	1604.15
273.15	90000.00	1636.02
273.15	100000.00	1666.45
278.15	100.00	1253.32

278.15	10000.00	1306.75
278.15	20000.00	1355.49
278.15	30000.00	1400.47
278.15	40000.00	1443.03
278.15	50000.00	1481.78
278.15	60000.00	1518.62
278.15	70000.00	1553.98
278.15	80000.00	1587.94
278.15	90000.00	1620.26
278.15	100000.00	1650.92
283.15	100.00	1231.06
283.15	10000.00	1285.85
283.15	20000.00	1335.74
283.15	30000.00	1381.4
283.15	40000.00	1424.94
283.15	50000.00	1464.43
283.15	60000.00	1502.23
283.15	70000.00	1537.63
283.15	80000.00	1571.96
283.15	90000.00	1604.68
283.15	100000.00	1635.63
288.15	100.00	1208.58
288.15	10000.00	1265.05
288.15	20000.00	1316.19
288.15	30000.00	1362.56
288.15	40000.00	1407.06
288.15	50000.00	1447.7
288.15	60000.00	1485.52
288.15	70000.00	1521.53
288.15	80000.00	1556.22
288.15	90000.00	1589.2
288.15	100000.00	1620.74
293.15	100.00	1186.09
293.15	10000.00	1244.14
293.15	20000.00	1296.86
293.15	30000.00	1344.06
293.15	40000.00	1389.16
293.15	50000.00	1430.7
293.15	60000.00	1469.0
293.15	70000.00	1505.7
293.15	80000.00	1540.77
293.15	90000.00	1574.27
293.15	100000.00	1605.01
298.15	100.00	1165.34

298.15	10000.00	1223.74
298.15	20000.00	1277.93
298.15	30000.00	1325.68
298.15	40000.00	1371.61
298.15	50000.00	1413.84
298.15	60000.00	1452.64
298.15	70000.00	1490.14
298.15	80000.00	1525.25
298.15	90000.00	1559.14
298.15	100000.00	1591.53

Reference *https://www.doi.org/10.1016/j.jct.2008.11.005*

Sources

Solid Liquid Phase Equilibrium and Phase Diagram for the Ternary Carbamazepine Succinic Acid Ethanol or Acetone System at (298.15 and 308.15) K:	https://www.doi.org/10.1021/je100918d
Determination and correlation of Avermectin B1a solubility in different binary solvent mixtures at temperatures from (283.15 to 313.15) K:	https://www.doi.org/10.1016/j.jct.2016.10.022
Solubilities of Phenylphosphinic Acid, Methylphenylphosphinic Acid, Hexachlorocyclotriphosphazene, and Hexaphenoxycyclotriphosphazene in Selected Solvents:	https://www.doi.org/10.1021/je1009812
Solubility of 1,1'-(Butane-1,4-diyl)-bis(3-methylpyridinium) Dihexylphosphate in Different Solvents from (288.15 to 333.15) K:	https://www.doi.org/10.1021/je9010954
Gas solubilities of carbon dioxide in methanol, acetone, mixtures of methanol and water, and mixtures of methanol and acetone:	https://www.doi.org/10.1016/j.fluid.2017.12.004
Determination and Correlation of Solubility of Tris(3-hydroxypropyl)phosphine Oxide in Selected Solvents:	https://www.doi.org/10.1021/je301029j
Crippen Method:	https://www.chemeo.com/doc/models/crippen_log10ws
Cosolvent effect on solubility of aripiprazole in mixed solvents and apparent thermodynamic analysis of dissolution process:	https://www.doi.org/10.1016/j.jct.2019.03.021
Solubility Measurement and Correlation of Two Erlotinib Hydrochloride Polymorphs in Alcohols and Ketones from 273.15 to 323.15 K:	https://www.doi.org/10.1021/acs.jced.6b00796
Activity coefficients at infinite dilution of organic solutes in the ionic liquid, methyl(trioctyl)ammonium thiosalicylate, [N1888][TS] by gas liquid chromatography at T = (303.15, 313.15, and 323.15) K:	https://www.doi.org/10.1016/j.jct.2010.12.020
Experimental measurement and modelling of solubility of inosine-5'-monophosphate disodium in pure and mixed solvents:	https://www.doi.org/10.1016/j.jct.2014.04.023
Solubility of 3-chloro-N-phenylphthalimide in ten organic solvents from T = (283.15 to 323.15) K: Determination and modelling:	https://www.doi.org/10.1016/j.jct.2016.01.003
Measurement and Correlation of the Solubility of L-Carnitine in Different Pure Solvents and Ethanol Acetone Solvent Mixture:	https://www.doi.org/10.1021/je500078n
Solubility of Fenofibrate in Different Binary Solvents: Experimental Data and Results of Thermodynamic Modeling:	https://www.doi.org/10.1021/acs.jced.6b00268
Solubility Determination and Thermodynamic Modeling of 2-Mercaptobenzimidazole in 12 Solvents from T = 278.15 K to T = 318.15 K:	https://www.doi.org/10.1021/acs.jced.9b00190
Solubility of Imidacloprid in Different Solvents:	https://www.doi.org/10.1021/je7004038
Activity coefficients at infinite dilution and physicochemical properties for organic solutes and water in the ionic liquid 4-(2-methoxyethyl)-4-methylmorpholinium bis(trifluoromethylsulfonyl)-amide:	https://www.doi.org/10.1016/j.jct.2011.11.021
Determination of Solubility and Thermodynamic Properties of [illegible]	https://www.doi.org/10.1021/acs.jced.8b00196
Solubility Determination and Modeling for 4,4'-Dihydroxydiphenyl Sulfone in Mixed Solvents of (Acetone, Ethyl Acetate, or Acetonitrile) + (Methanol and Acetone + Ethanol from (278.15 to 313.15) K:	https://www.doi.org/10.1021/acs.jced.6b00421
Solubilities of 2-(6-Oxido-6H-dibenz[c,e][1,2]oxaphosphorin-6-yl)-methanol in Selected Solvents:	https://www.doi.org/10.1021/je050488p
Solubilities of 3-Pentadecylphenol in Ethanol, 1-Butanol, Toluene, Acetone, Tetrachloromethane, and Ethyl Acetate:	https://www.doi.org/10.1021/je900346t
Solubility Measurement and Modeling for 2-Benzoyl-3-chlorobenzoic Acid and 1-Chloroanthraquinone in Organic Solvents:	https://www.doi.org/10.1021/je400246s
Solubility of Agomelatine Crystal Form I and Form II in Pure Solvents and (Isopropanol + Water) Mixtures:	https://www.doi.org/10.1021/acs.jced.5b00586
Activity coefficients at infinite dilution and physicochemical properties for organic solutes and water in the ionic liquid 1-(2-methoxyethyl)-1-methylpyrrolidinium tris(pentafluoroethyl)trifluorophosphate:	https://www.doi.org/10.1016/j.jct.2013.01.007
Measurement and Correlation for Solubility of L-valine in Water, Ethanol, N,N-Dimethylformamide, Acetone, and Isopropyl Alcohol between 293.15 and 343.15 K:	https://www.doi.org/10.1021/je500035r
Binary Solid-Liquid Solubility Determination and Model Correlation of Quizalofop-p-ethyl in Different Pure Solvents:	https://www.doi.org/10.1021/acs.jced.8b01181

Solubility of Desmosterol in Five Organic Solvents:	https://www.doi.org/10.1021/je8006088
Solubility of Tetrahydroxybenzophenone (THBP) in Acetone:	https://www.doi.org/10.1021/je700434b
(Solid + liquid) phase equilibria of tetraphenyl piperazine-1,4-diyldiphosphonate in pure solvents:	https://www.doi.org/10.1016/j.jct.2014.06.021
Solubility Behavior and Data Modeling of L-Proline in Different Neat and Binary Solvent Systems:	https://www.doi.org/10.1021/acs.jced.9b00802
Thermodynamics and activity coefficients at infinite dilution for organic compounds in the ionic liquid 1-hexyl-3-methylimidazolium chloride:	https://www.doi.org/10.1016/j.jct.2018.08.028
Surface Tension of Pure Liquids and Binary Liquid Mixtures:	https://www.doi.org/10.1021/je0201323
Solubility of Anthracene in C1-C3 Alcohols from (298.2 to 333.2) K and Their Mixtures with 2-Propanone at 298.2 K:	https://www.doi.org/10.1021/je1001945
Thermodynamic models for determination of 5-chloro-N-phenylphthalimide solubility in binary solvent mixtures of (acetone, ethyl acetate or 1,4-dioxane + methanol)	https://www.doi.org/10.1016/j.jct.2016.04.007
Experiment and Computation of Solubility and Dissolution Properties for Enalapril Maleate and Its Intermediate in Pure Solvents:	https://www.doi.org/10.1021/acs.jced.8b00333
Solubilities of Betulin in Fourteen Organic Solvents at Different Temperatures:	https://www.doi.org/10.1021/je700069g
Solubility of Cytidine 5'-Diphosphocholine Sodium in Water and Different Binary Mixtures from (278.15 to 298.15) K:	https://www.doi.org/10.1021/je900465r
Solubilities of succinimide in different pure solvents and binary methanol+ethyl acetate solvent mixtures:	https://www.doi.org/10.1016/j.tca.2012.03.007
Measurement of Isobaric Vapor - Liquid Equilibria of Dimethyl Carbonate with Acetone, 2-Butanone and 2-Pentanone at 101.3 kPa and Density and Speed of Sound at 298.15 K:	https://www.doi.org/10.1021/je0497000
Measurement and correlation of solubility and dissolution properties of [illegible] in pure and binary solvents:	https://www.doi.org/10.1016/j.fluid.2015.06.026
Solid Forms, Crystal Habits, and Solubility of Danthron:	https://www.doi.org/10.1021/acs.jced.5b00192
Separation of binary mixtures based on gamma infinity data using [OMMIM][NTf2] ionic liquid and modelling of thermodynamic functions:	https://www.doi.org/10.1016/j.jct.2017.12.012
Solubility of Acenaphthene in Different Solvents from (283.00 to 323.00) K:	https://www.doi.org/10.1021/je8006869
Solubilities of Palmitic Acid in Pure Solvents and Its Mixtures:	https://www.doi.org/10.1021/je8005979
Activity Coefficients at Infinite Dilution Measurements for Organic Solutes and Water in the Ionic Liquid 1-Butyl-3-methylpyridinium Trifluoromethanesulfonate:	https://www.doi.org/10.1021/je1000582
Determination of Activity Coefficients at Infinite Dilution of Organic Solutes in the Ionic Liquid 1-(2-Hydroxyethyl)-3-methylimidazolium Nonafluoro-1-butanesulfonate Using Gas-Liquid Chromatography:	https://www.doi.org/10.1021/acs.jced.8b00080
Solubility of Aspirin in Supercritical Carbon Dioxide with and without Acetone:	https://www.doi.org/10.1021/je0499465
Systems with ionic liquids: Measurement of VLE and c1 data and prediction of their thermodynamic behavior using original UNIFAC, mod. UNIFAC(Do) and COSMO-RS(Ol):	https://www.doi.org/10.1016/j.jct.2005.04.010
Determination and correlation of pyridoxine hydrochloride solubility in different binary mixtures at temperatures from (278.15 to 313.15) K:	https://www.doi.org/10.1016/j.jct.2015.09.026
Solubility of diosgenin in different solvents:	https://www.doi.org/10.1016/j.jct.2011.11.009
Solid-liquid equilibrium of sulbactam in pure solvents and binary solvent mixtures:	https://www.doi.org/10.1016/j.fluid.2014.09.008
Measurement and Correlation of the Solubility of 5-Fluorouracil in Pure and Binary Solvents:	https://www.doi.org/10.1021/acs.jced.8b00425
Solubilities of Some Phosphaspirocyclic Compounds in Selected Solvents:	https://www.doi.org/10.1021/je060138i
Correlation of solubility of pioglitazone hydrochloride in different binary solvents:	https://www.doi.org/10.1016/j.fluid.2013.05.008
Measurement and modelling of econazole nitrate in twelve pure organic solvents at temperatures from 278.15 K to 318.15 K:	https://www.doi.org/10.1016/j.jct.2016.07.043
Thermodynamics and activity coefficients at infinite dilution for organic solutes and water in the ionic liquid 1-butyl-1-methylmorpholinium tricyanomethanide:	https://www.doi.org/10.1016/j.jct.2013.08.030
Measurement and Correlation of Solubility of β-Hydroxy-2-naphthoic Acid in Pure and Binary Mixed Organic Solvents from T = (293.15 to 333.15 K):	https://www.doi.org/10.1021/acs.jced.8b00578
Solubility of Benzoin in Six Monosolvents and in Some Binary Solvent Mixtures at Various Temperatures:	https://www.doi.org/10.1021/acs.jced.7b00238
Solubilities of 3,9-Dimethyl-3,9-dioxide-2,4,8,10-tetraoxa-3,9-diphosphaspiro[5.5]undecane in Selected Solvents:	https://www.doi.org/10.1021/je9004189
Solubility of Decabromodiphenyl Ether in Different Solvents at (283.0 to 323.0) K:	https://www.doi.org/10.1021/je100316k
Solubilities of 1,1'-[1,2-Ethanediylbis(oxy-1,2-ethanediyl)]bis-[3-methyl-1H-imidazolium-1-yl] Dihexafluorophosphate in Water, Methanol, Ethanol, Dimethylformamide, Dimethyl Sulfoxide, and Acetone + Water Binary Solvent Mixtures:	https://www.doi.org/10.1021/je5004466
Quaternary, Ternary, and Binary LLE Measurements for 2-Methoxy-2-methylpropane + Furfural + Acetic Acid + Water at Temperatures between 298 and 307 K:	https://www.doi.org/10.1021/acs.jced.6b00149
Excess Molar Volumes and Surface Tensions of Xylene with Acetone or 2-Butanone at 298.15 K:	https://www.doi.org/10.1021/je0341763
Separation of sulfur compounds from alkanes with 1-alkylcyanopyridinium-based ionic liquids:	https://www.doi.org/10.1016/j.jct.2013.09.032
Measurements of activity coefficient at infinite dilution for organic solutes in tetramethylammonium chloride + ethylene glycol deep eutectic solvent using gas-liquid chromatography:	https://www.doi.org/10.1016/j.fluid.2018.01.019
Solution Thermodynamics of Benzotriazole in Different Pure Solvents:	https://www.doi.org/10.1021/acs.jced.7b01085
Thermodynamic models for determination of solid liquid equilibrium of the 6-benzyladenine in pure and binary organic solvents:	https://www.doi.org/10.1016/j.jct.2016.12.002
Solubility Modeling of 4-(Methylsulfonyl)benzaldehyde in Nine Organic Solvents at Elevated Temperatures:	https://www.doi.org/10.1021/acs.jced.5b01053
Measurement and Correlation of the Solubility of p-Coumaric Acid in Nine Pure and Water + Ethanol Mixed Solvents at Temperatures from 293.15 to 333.15 K:	https://www.doi.org/10.1021/acs.jced.6b00361

Solubility of Form III Piracetam in a Range of Solvents:	https://www.doi.org/10.1021/je1003934
Determination and correlation of solubility of 4-bromomethyl-2-cyanobiphenyl in acetone + (ethanol, n-propanol, n-butanol) mixtures:	https://www.doi.org/10.1016/j.jct.2016.06.032
Solubility of Artemisinin in Seven Different Pure Solvents from (283.15 to	https://www.doi.org/10.1021/je800515w
Activity coefficients at infinite dilution of organic solutes in the ionic liquid ethyl(2-hydroxyethyl)dimethyl-ammonium	https://www.doi.org/10.1016/j.fluid.2012.04.008
Activity coefficients at infinite dilution and physicochemical properties for	https://www.doi.org/10.1016/j.jct.2012.01.019
Activity coefficients at infinite dilution of organic solutes in 1-hexyl-3-	https://www.doi.org/10.1016/j.jct.2013.10.038
Solubility of Xylitol in Ethanol, Acetone, N,N-Dimethylformamide, 1-Butanol,	https://www.doi.org/10.1021/je060348v
Measurement and Correlation of Solubilities of	https://www.doi.org/10.1021/acs.jced.9b00275
Thermodynamics of solubility of thiomalic acid in different	https://www.doi.org/10.1016/j.fluid.2013.05.002
Activity coefficients at infinite dilution of organic solutes in the ionic liquid	https://www.doi.org/10.1016/j.jct.2013.05.030
Solubility Measurement and Modeling	https://www.doi.org/10.1021/acs.jced.5b00714
Solubilities of Diphenyl Sulfone in Selected Solvents	https://www.doi.org/10.1021/je800049b
Solubility of 2,5-Di-tert-butylhydroquinone and Process Design for Its Purification Using Crystallization:	https://www.doi.org/10.1021/je501049c
Solubility and Melting Properties of Salicylamide:	https://www.doi.org/10.1021/je060178m
Ternary Diffusion Coefficients of Glycerol + Acetone + Water by Taylor Dispersion Measurements at 298.15 K:	https://www.doi.org/10.1021/je050082c
Solubility Study of (2E)-1-(3-Pyridyl)-3-(dimethylamino)-2-propen-1-one in Different Pure Solvents and Binary	https://www.doi.org/10.1021/acs.jced.9b00047
Measurement and prediction of multicomponent diffusion coefficients in four ternary liquid systems:	https://www.doi.org/10.1016/j.fluid.2009.12.009
Determination of the solubility, dissolution enthalpy and entropy of suberic acid in different solvents:	https://www.doi.org/10.1016/j.fluid.2012.06.012
Solubility of 11r-Hydroxy-16r,17r-Epoxyprogesterone in Different Solvents between 283 K and 323 K:	https://www.doi.org/10.1021/je0495435
Solubility and Dissolution Thermodynamics for 1,3,5-Trichlorobenzene in Organic Solvents:	https://www.doi.org/10.1021/acs.jced.5b00607
Solubility of (+)-(S)-2-(6-Methoxynaphthalen-2-yl) Propanoic Acid in Acetone, Methanol, Ethanol, Propan-2-ol, and Ethyl Ethanoate at Temperatures between	https://www.doi.org/10.1021/je8008039
Vapor-Liquid Equilibrium for Binary Systems of Diacetyl with Methanol and	https://www.doi.org/10.1021/je060174h
Solubility Modeling and Mixing Thermodynamics of Thiamphenicol in Water and Twelve Neat Organic Solvents from T = (278.15 to 318.15) K:	https://www.doi.org/10.1021/acs.jced.7b00542
Speed of Sound and Refractive Index of	https://www.doi.org/10.1021/je9005618
1-Ethyl-3-methylimidazolium Trifluoromethanesulfonate with	https://www.doi.org/10.1016/j.jct.2016.07.017
Separation of aliphatic/aromatic hydrocarbons and sulphur compounds	https://www.doi.org/10.1016/j.jct.2014.04.025
Determination and Correlation of the Solubility of l-Fucose in Four Binary	https://www.doi.org/10.1021/acs.jced.8b00361
Ternary and binary LLE measurements	https://www.doi.org/10.1016/j.jct.2017.02.016
+ furfural + Water between 298 and 343 K:	https://www.doi.org/10.1021/je200305n
Measurement of Activity Coefficients at Infinite Dilution of Benzene, Toluene,	https://www.doi.org/10.1021/je049875+
Thermodynamics of binary mixtures of N-methyl-2-pyrrolidinone and ketone	https://www.doi.org/10.1016/j.jct.2004.11.007
Activity Coefficients at Infinite Dilution Measurements for Organic Solutes and	https://www.doi.org/10.1021/je900890u
Solubility of Physalin D in Ethanol, Methanol, Propanone,	https://www.doi.org/10.1021/je100227k
Trichloromethane, Ethyl Ethanoate, and Water at Temperatures from (283.2 to	https://www.doi.org/10.1016/j.jct.2015.02.004
Measurement and correlation of solubility and solution thermodynamics	https://www.doi.org/10.1016/j.jct.2015.12.032
Activity Coefficients at Infinite Dilution of Polar Solutes in	https://www.doi.org/10.1021/je700560s
Solid-liquid equilibria of the ternary system m-nitrobenzoic acid +	https://www.doi.org/10.1016/j.fluid.2009.07.011
Solubility of 16.alpha.,17.alpha.-Epoxyprogesterone in Six Different Solvents:	https://www.doi.org/10.1021/je050195w
Activity coefficients at infinite dilution measurements for organic solutes and water in the ionic liquid	https://www.doi.org/10.1016/j.jct.2011.02.012
Co-solvency phenomenon and thermodynamic properties of	https://www.doi.org/10.1016/j.jct.2019.06.018
Solubility of Caprolactam in 2-Propanol, Acetone, Acetonitrile, Methyl Acetate, Ethyl Acetate, and Butyl Acetate:	https://www.doi.org/10.1021/je900401z
Separation of binary mixtures hexane/hex-1-ene, cyclohexane/cyclohexene and ethylbenzene/styrene based on gamma infinity data measurements:	https://www.doi.org/10.1016/j.jct.2017.11.017
Measurement and Correlation of Solubility of Xylitol in Binary Ethanol + Acetone Solvent Mixtures with the Combined Nearly Ideal Binary Solvent/Redlich-Kister Equation:	https://www.doi.org/10.1021/je700131z

Thermodynamic properties of L-Theanine in different solvents:	https://www.doi.org/10.1016/j.jct.2016.10.017
Solubility of p-Aminobenzenesulfonamide in Different Solvents from (283.15 to 323.15) K:	https://www.doi.org/10.1021/je900038m
Solubility of 1-H-Tetrazole-1-acetic Acid in Different Solvents between 283 K and 323 K:	https://www.doi.org/10.1021/je700200b
Solubility determination and modelling for phthalimide in mixed solvents of (acetone, ethyl acetate or acetonitrile + methanol) from (278.15 to 313.15) K:	https://www.doi.org/10.1016/j.jct.2016.11.023
Solubility of Two Polymorphs of Erlotinib Hydrochloride in Isopropanol and Acetone from (273.15 to 303.15) K:	https://www.doi.org/10.1021/je500482k
The Yaws Handbook of Vapor Pressure:	https://www.sciencedirect.com/book/9780128029992/the-yaws-handbook-of-vapor-pressure
Solubilities of 4-(Hydroxymethyl)-1-oxido-2,6,7-trioxa-1-phosphabicyclo[2.2.2]octane in Organic Solvents:	https://www.doi.org/10.1021/je4009816
Measurement and Correlation of Solubility of Calcium Formate (Form alpha) in Different Binary Solvent Mixtures at Temperatures from 283.15 to 323.15 K:	https://www.doi.org/10.1021/acs.jced.9b00009
Solubilities of 2-(6-Oxido-6H-dibenz[c,e][1,2]oxaphosphorin-6-yl)-1,4-dihydroxy benzene in Selected Solvents:	https://www.doi.org/10.1021/je700404u
Measurement and Correlation of Solubility of Gatifloxacin in 12 Pure Solvents from 278.15 K to 318.15 K:	https://www.doi.org/10.1021/acs.jced.8b00902
Vapor-Liquid Equilibria of the Carbon Dioxide + Acetone System at Pressures from (2.36 to 11.77) MPa and Temperatures from (333.15 to 393.15) K:	https://www.doi.org/10.1021/je049887v
Solubility of Florfenicol in Different Solvents at Temperatures from (278 to 318) K:	https://www.doi.org/10.1021/je1008284
Surface Tension of the Ternary System Water + Acetone + Toluene:	https://www.doi.org/10.1021/je7000182
Solubility of Fumaric Acid in Propan-2-ol, Ethanol, Acetone, Propan-1-ol, and Water:	https://www.doi.org/10.1021/je9001637
Solubilities of Stearic Acid in Organic Solvents and in Azeotropic Solvent Mixtures:	https://www.doi.org/10.1021/je7006567
Solubilities of 1,3,2-Dioxaphosphorinane-2-methanol-a,a-5,5-tetramethyl 2-oxide in Selected Solvents at 273.15 K to 343.15 K:	https://www.doi.org/10.1021/je4003114
Determination of Infinite Dilution Activity Coefficients of Several Organic Solutes in N-Butylpyridinium Nitrate/N-Octylpyridinium Nitrate by Blend Inverse Gas Chromatography:	https://www.doi.org/10.1021/acs.jced.7b00244
Separation of ethylbenzene/styrene systems using ionic liquids in ternary LLE:	https://www.doi.org/10.1016/j.jct.2016.09.003
Solubility of Ammonium Thiocyanate in Different Solvents:	https://www.doi.org/10.1021/je800621p
Measurement and prediction of vapor-liquid equilibria of ternary systems containing ionic liquids:	https://www.doi.org/10.1016/j.fluid.2004.11.010
Coupled phase-reaction equilibrium for dihydromyrcene hydration system:	https://www.doi.org/10.1016/j.fluid.2016.11.007
Solubility of 2-Hydroxybenzoic Acid in Water, 1-Propanol, 2-Propanol, and 2-Propanone at (298.2 to 338.2) K and Their Aqueous Binary Mixtures at 298.2 K:	https://www.doi.org/10.1021/je300980k
Densities, viscosities and refractive indices of binary and ternary mixtures of methanol, acetone, and chloroform at temperatures from (298.15-318.15) K and ambient pressure:	https://www.doi.org/10.1016/j.fluid.2018.10.024
Solubility of D-phenylglycine methyl ester hydrochloride in water and in organic individual and mixed solvents: Experimental data and results of thermodynamic modeling:	https://www.doi.org/10.1016/j.fluid.2016.02.020
Solubility in Different Solvents, Crystal Polymorph and Morphology, and Optimization of Crystallization Process of AIBN:	https://www.doi.org/10.1021/acs.jced.7b00538
Solubilities of phosphoric acid, P,P0-1,4-phenylene P,P,P0,P0-tetraphenyl ester in the organic solvents:	https://www.doi.org/10.1016/j.fluid.2014.12.034
Critical Assessment of CO2 Solubility in Volatile Solvents at 298.15 K:	https://www.doi.org/10.1021/je101161d
Solubilities of Methyldiphenylphosphine Oxide in Selected Solvents:	https://www.doi.org/10.1021/je900277m
Solubility of Erythritol in Different Solvents:	https://www.doi.org/10.1021/je0501033
Solubility of d-Aspartic Acid in Several Neat Solvents: Determination, Modelling, and Solvent Effect Analysis:	https://www.doi.org/10.1021/acs.jced.9b00320
Solubilities of Phosphorus-Containing Compounds in Selected Solvents:	https://www.doi.org/10.1021/je100341q
Abraham model correlations for describing the thermodynamic properties of solute transfer into pentyl acetate based on headspace chromatographic and solubility measurements:	https://www.doi.org/10.1016/j.jct.2018.05.003
Solubility of Ionic Liquid [Bmim]Ac in Supercritical CO2 Containing Different Cosolvents:	https://www.doi.org/10.1021/acs.jced.7b01108
Measurement and calculation of solid-liquid phase equilibrium for ternary 4-chlorophthalic anhydride + 3-chlorophthalic anhydride + acetone system:	https://www.doi.org/10.1016/j.fluid.2013.08.032
Vapor-Liquid Equilibrium for Binary System of 2-Propanol + 2-Ethoxyethanol at 353 K and Ethyl ethanoate + 1,1-Diethoxyethane at 348 K and 2-Propanone + 1,1-Diethoxyethane at 328 K:	https://www.doi.org/10.1021/je049816w
Solubilities of Adefovir Dipivoxil in Different Binary Solvents at 298.15 K:	https://www.doi.org/10.1021/je700664g
Determination and thermodynamic modeling of solid-liquid phase equilibrium for 3,5-dichloroaniline in pure solvents and ternary 3,5-dichloroaniline + 1,3,5-trichlorobenzene + toluene system:	https://www.doi.org/10.1016/j.jct.2016.03.007
Solubility and mixing thermodynamic properties of (2,4,6-trimethylbenzoyl) diphenylphosphine oxide in pure and binary solvents:	https://www.doi.org/10.1016/j.fluid.2018.01.015
Solubility of CO2 in Alcohols, Glycols, Ethers, and Ketones at High Pressures from (288.15 to 318.15) K:	https://www.doi.org/10.1021/je101344v
Thermodynamic equilibrium of hydrocortisone form I in different solvent systems:	https://www.doi.org/10.1016/j.fluid.2017.09.007
Saturated Solubility and Thermodynamic Evaluation of l-Tryptophan in Eight Pure Solvents and Three Groups of Binary Mixed Solvents by the Gravimetric Method at T = 278.15–333.15 K:	https://www.doi.org/10.1021/acs.jced.9b00562
Solubility of Propyl p-hydroxybenzoate in Supercritical Carbon Dioxide with and without a Cosolvent:	https://www.doi.org/10.1021/je049752p
Solid-Liquid Phase Equilibrium and Thermodynamic Analysis of N,N'-Diethylthiourea in Different Solvent Systems:	https://www.doi.org/10.1021/acs.jced.9b00854

Solubilities of Betulinic Acid in Thirteen Organic Solvents at Different Temperatures:	https://www.doi.org/10.1021/je200531k
The dissolution behavior and apparent thermodynamic analysis of temozolomide in pure and mixed solvents:	https://www.doi.org/10.1016/j.jct.2018.11.026
Measurement and correlation solubility and mixing properties of L-malic acid in pure and mixture organic solvents:	https://www.doi.org/10.1016/j.fluid.2013.10.011
Activity coefficients at infinite dilution and physicochemical properties for organic solutes and water in the ionic liquid 1-(2-methoxyethyl)-1-methylpyrrolidinium bis(trifluoromethylsulfonyl)-amide:	https://www.doi.org/10.1016/j.jct.2012.03.015
Measurement and correlation of the solubility of maleic acid in acetone + ethyl acetate mixtures:	https://www.doi.org/10.1016/j.tca.2012.03.023
Determination and correlation of solubility of 11alpha,17alpha-dihydroxy-4-pregnene-3,20-dione in different solvents:	https://www.doi.org/10.1016/j.jct.2016.10.046
Solubility of alpha-form and beta-form of L-glutamic acid in different aqueous solvent mixtures:	https://www.doi.org/10.1016/j.fluid.2010.10.020
Solid Liquid Phase Equilibrium of Phosphoramidic acid, N,N'-1,2-Ethanediylbis-P,P,P',P'-tetraphenyl Ester in Selected Solvents:	https://www.doi.org/10.1021/je5010565
Thermodynamics of fenofibrate and solubility in pure organic solvents:	https://www.doi.org/10.1016/j.fluid.2014.01.029
Measurements of activity coefficients at infinite dilution for organic solutes and water in N-hexylisoquinolinium thiocyanate, [HiQuin][SCN] using GLC:	https://www.doi.org/10.1016/j.jct.2013.02.004
Activity coefficients at infinite dilution of organic solutes in the ionic liquid 1-butyl-3-methylimidazolium hexafluoroantimonate using gas liquid chromatography at T = (313.15, 323.15, and 333.15) K:	https://www.doi.org/10.1016/j.jct.2010.12.019
Solubility Determination and Thermodynamic Modeling of Buprofezin in Different Solvents and Mixing Properties of Solutions:	https://www.doi.org/10.1021/acs.jced.8b01099
Solubility measurement and modelling of ethyl 5-amino-4-cyano-3-(2-ethoxy-2-oxoethyl)-2-thiophenecarboxylate in four groups mixed solvents:	https://www.doi.org/10.1016/j.jct.2016.09.001
Solid-Liquid Equilibrium of Isomaltulose in Five Pure Solvents and Four Binary Solvents from (288.15 to 323.15) K:	https://www.doi.org/10.1021/acs.jced.8b00823
Thermodynamics of ketone + amine mixtures. Part X. Excess molar enthalpies at 298.15 K for N,N,N-triethylamine + 2-alkanone systems. Characterization of tertiary amine + 2-alkanone, and amino-ketone + n-alkane mixtures in terms of DISQUAC:	https://www.doi.org/10.1016/j.fluid.2013.07.037
Properties of ionic liquid HMIMPF6 with carbonates, ketones and alkyl acetates:	https://www.doi.org/10.1016/j.jct.2005.07.020
Activity Coefficients at Infinite Dilution of Organic Solutes in 1-Butyl-3-methylimidazolium Nitrate Using Gas-Liquid Chromatography:	https://www.doi.org/10.1021/je200050q
Solubility of β-Carotene in Binary Solvents Formed by Some Hydrocarbons with 2,5,8-Trioxanonane, 2-Propanone, and Cyclohexanone:	https://www.doi.org/10.1021/je060376d
Experimental and theoretically study of interaction between organic compounds and tricyanomethanide based ionic liquids:	https://www.doi.org/10.1016/j.jct.2014.12.027
Thermodynamic Functions for the Solubility of 3-Nitrobenzonitrile in 12 Organic Solvents from T/K = (278.15 to 318.15):	https://www.doi.org/10.1021/acs.jced.7b00615
Equilibrium study and diversified models of drug Norfloxacin in eight pure organic and binary solvents at T = (278.15 to 328.15) K:	https://www.doi.org/10.1016/j.fluid.2016.12.012
Lamotrigine Solubility in Some Nonaqueous Solvent Mixtures at 298.2 K:	https://www.doi.org/10.1021/acs.jced.5b00355
Excess Molar Enthalpies of CO2 + Acetone at Pressures from (9.00 to 18.00) MPa and Temperatures from (313.15 to 333.15) K:	https://www.doi.org/10.1021/je100220c
Isothermal (vapor + liquid) equilibria for the binary mixtures of (propylene oxide + ethanol) and (propylene oxide + 1-propanol) at several temperatures:	https://www.doi.org/10.1016/j.jct.2015.02.006
Effect of the alkyl side chain of the 1-alkylpiperidinium-based ionic liquids on desulfurization of fuels:	https://www.doi.org/10.1016/j.jct.2013.12.029
Measurements of the solubilities of m-phthalic acid in acetone, ethanol and acetic ether:	https://www.doi.org/10.1016/j.fluid.2008.01.014
Activity coefficients at infinite dilution of organic solutes in the ionic liquid trihexyltetradecylphosphonium hexafluorophosphate using gas liquid chromatography at T = (313.15, 333.15, 353.15 and 363.15) K:	https://www.doi.org/10.1016/j.jct.2012.01.004
Solubility determination and thermodynamic modeling of 5-nitro-8-hydroxyquinoline in ten organic solvents from T = (278.15 to 313.15) K and mixing properties of solutions:	https://www.doi.org/10.1016/j.jct.2016.04.018
Measurement and correlation of solubilities of apigenin and apigenin 7-O-rhamnosylglucoside in seven solvents at different temperatures:	https://www.doi.org/10.1016/j.jct.2010.09.013
Solubility of Hesperetin in Various Solvents from (288.2 to 323.2) K:	https://www.doi.org/10.1021/je800078j
KDB Vapor Pressure Data:	https://www.cheric.org/research/kdb/hcprop/showprop.php?cmpid=1191
Antisolvent Precipitation of Vitamin B6: A Thermodynamic Study:	https://www.doi.org/10.1021/je200853g
Solid-liquid phase equilibrium and mixing properties of 2-Cyano-4'-methylbiphenyl in pure solvents:	https://www.doi.org/10.1016/j.jct.2016.07.050
Influence of Different Solvent Properties and Composition for the Solubility of Iopromide:	https://www.doi.org/10.1021/acs.jced.8b00416
Thermodynamic study of molecular interaction-selectivity in separation processes based on limiting activity coefficients:	https://www.doi.org/10.1016/j.jct.2018.02.014
Measurement and Correlation for Solubilities of 16-Dehydropregnenolone Acetate in Different Solvents:	https://www.doi.org/10.1021/je1013262
Solubility and density of the disodium salt hemiheptahydrate of ceftriaxone in (acetone + water) at T=(298.15, 308.15, and 318.15) K:	https://www.doi.org/10.1016/j.jct.2003.10.006
Determination and Modeling of Solubility of 4-Aminobenzamide in Different Pure Solvents:	https://www.doi.org/10.1021/acs.jced.8b01144
Thermodynamics and selectivity of separation based on activity coefficients at infinite dilution of various solutes in ionic liquid [MMIM][BF4]:	https://www.doi.org/10.1016/j.jct.2018.05.017
Densities, Refractive indices and Viscosities for Binary and Ternary Mixtures of Acetone, Ethanol and 2,2,4-Trimethylpentane at T = (288.15, 298.15, and 308.15) K:	https://www.doi.org/10.1021/je050010l
Solubility and thermodynamic functions of tebuconazole in nine organic solvents from T = (278.15 to 313.15) K and mixing properties of solutions:	https://www.doi.org/10.1016/j.jct.2016.11.032
Solubilities of (2,5-Dihydroxyphenyl)diphenyl Phosphine Oxide in Selected Solvents:	https://www.doi.org/10.1021/je900730w
A relative headspace method for Henry s constants of volatile organic compounds:	https://www.doi.org/10.1016/j.fluid.2005.02.006

Solubility of 3,7-Dinitro-1,3,5,7-tetraazabicyclo [3.3.1] Nonane in Ethanenitrile, Methanol, 1,1-Dichloroethane, Dimethyl Sulfoxide, Acetone, and Mixed Solvents	https://www.doi.org/10.1021/acs.jced.5b00033
Vapor Liquid Equilibrium for Binary System of 2-Propanol + 2-Methoxy-2-methylpropane and Ethyl Ethanoate + 2-Ethoxy-2-methylpropane at 333 K and 2-Propanol + 2-Ethoxy-2-methylpropane at 323 K:	https://www.doi.org/10.1021/je0342421
Isobaric vapor liquid equilibria for mixtures of acetone, ethanol and 2,2,4-trimethylpentane at 101.3 kPa:	https://www.doi.org/10.1016/j.fluid.2005.01.007
Solution thermodynamics of valnemulin hydrogen fumarate in different pure solvents:	https://www.doi.org/10.1016/j.jct.2015.07.023
The solubility of mercaptosuccinic acid in water + (methanol, ethanol, acetone) mixtures from (278.15 to 333.15 K):	https://www.doi.org/10.1016/j.fluid.2013.11.001
Measurement and Correlation of Solubilities of trans-Resveratrol in Ethanol + Water and Acetone + Water Mixed Solvents at Different Temperatures:	https://www.doi.org/10.1021/je800410b
Solubility of Acetylsalicylic Acid in Ethanol, Acetone, Propylene Glycol, and 2-Propanol:	https://www.doi.org/10.1021/je7005693
Solubility and Solvent Effect of 1-(2-Bromo-phenyl)-pyrrole-2,5-dione in 14 Pure Solvents from 278.15 to 323.15 K:	https://www.doi.org/10.1021/acs.jced.9b00560
Experimental isobaric vapour liquid liquid equilibrium data for the quaternary systems water (1) ethanol (2) acetone (3) n-butyl acetate (4) and water (1) ethanol (2) acetone (3) methyl ethyl ketone (4) and their partially miscible-constituent ternaries:	https://www.doi.org/10.1016/j.fluid.2006.11.011
Measurement of Activity Coefficients of Solutes at Infinite Dilution in (Dimethyl Sulfoxide + Acetamide) and (Formamide + Acetamide) at Different Temperature 298.15 K:	https://www.doi.org/10.1021/je030122h
Solubility and Dissolution Thermodynamic Data of Mefenamic Acid Crystals in Different Classes of Organic Solvents:	https://www.doi.org/10.1021/je400714f
Conductivities and Viscosities of the Ionic Liquid [bmim][PF6] + Water + Ethanol and [bmim][PF6] + Water + Acetone Ternary Mixtures:	https://www.doi.org/10.1021/je034078h
Solubilities of rutin in eight solvents at T = 283.15, 298.15, 313.15, 323.15, and 333.15 K:	https://www.doi.org/10.1016/j.fluid.2007.07.030
Solubility of Erythromycin A Dihydrate in Different Pure Solvents and Acetone + Water Binary Mixtures between 293 K and 323 K:	https://www.doi.org/10.1021/je0505265
Solubility and thermodynamic functions of cefazolin acid in pure and mixed solvents at (278.15 308.15) K:	https://www.doi.org/10.1016/j.fluid.2014.12.020
Aniracetam Solubility in Pure and Binary Solvents: Effect of Molecular Interaction and Analysis of Crystallized Products:	https://www.doi.org/10.1021/acs.jced.8b00023
Measurements of activity coefficients at infinite dilution for organic solutes in two quaternary ammonium-based ionic liquids [DDA][ClO4] and [DDA][BF4]:	https://www.doi.org/10.1016/j.fluid.2018.11.011
Determination and correlation thermodynamic models for solid liquid equilibrium of the Nifedipine in pure and mixture organic solvents:	https://www.doi.org/10.1016/j.jct.2016.07.035
Measurement of Infinite Dilution Activity Coefficients of Alcohols, Ketones, and Aromatic Hydrocarbons in 4-Methyl-N-butylpyridinium Tetrafluoroborate and 1-Butyl-3-methylimidazolium Hexafluorophosphate by Gas-Liquid Chromatography:	https://www.doi.org/10.1021/je800246v
Correlation and thermodynamic analysis of solubility of diphenhydramine hydrochloride in pure and binary solvents:	https://www.doi.org/10.1016/j.jct.2015.09.032
Determination and modeling of binary and ternary solid-liquid phase equilibrium for the systems formed by p-nitrobenzoic acid, m-nitrobenzoic acid and acetone:	https://www.doi.org/10.1016/j.jct.2016.10.004
Experimental Determination of Solubilities of Rutin in Acetone + Water and Ethanol + Water Mixed Solvents at T = (278.2, 288.2, 298.2, 308.2, and 318.2) K:	https://www.doi.org/10.1021/je7004177
Thermodynamic modelling for solubility of 4-nitrobenzaldehyde in different solvents at temperature range from (273.15 to 313.15) K and mixing properties of solutions:	https://www.doi.org/10.1016/j.jct.2016.09.012
Experimental and theoretical study on infinite dilution activity coefficients of various solutes in piperidinium ionic liquids:	https://www.doi.org/10.1016/j.jct.2013.01.005
VLE measurements using a static cell vapor phase manual sampling method accompanied with an empirical data consistency test:	https://www.doi.org/10.1016/j.jct.2014.08.024
Thermodynamic Models for Determination of the Solubility of Sulfanilic Acid in Different Solvents at Temperatures from (278.15 to 328.15) K:	https://www.doi.org/10.1021/je500867u
Activity coefficients at infinite dilution and physicochemical properties for organic solutes and water in the ionic liquid 4-(2-methoxyethyl)-4-methylmorpholinium trifluorotris(perfluoroethyl)phosphate:	https://www.doi.org/10.1016/j.jct.2012.05.017
Solubilities of 1,1?-(Propane-1,3-diyl)-bis(3-methyl-1H-imidazolium-1-yl) Dihexafluorophosphate in Different Solvents:	https://www.doi.org/10.1021/je9006585
Solubility measurement and modelling of 1,8-dinitronaphthalene in nine organic solvents from T = (273.15 to 308.15) K and mixing properties of solutions:	https://www.doi.org/10.1016/j.jct.2015.07.010
Solubility and Solution Thermodynamics of Diphenoxylate in Different Pure Solvents:	https://www.doi.org/10.1021/je5010033
Densities and Viscosities of Binary Liquid Mixtures of Trichloroethylene and Tetrachloroethylene with Some Polar and Nonpolar Solvents:	https://www.doi.org/10.1021/je034203p
Experimental Densities and Excess Volumes for Binary Mixtures Containing Propionic Acid, Acetone and Water from 283.15 K to 323.15 K at Atmospheric Pressure:	https://www.doi.org/10.1021/je030102f
Solubilities of Sulfadiazine in Methanol, Ethanol, 1-Propanol, 2-Propanol, Acetone, and Chloroform from (294.15 to 318.15) K:	https://www.doi.org/10.1021/je900742p
Determination and Modeling of Solid-Liquid Equilibrium for Ternary Systems of Terephthalaldehyde + 4-Hydroxybenzaldehyde + Ethyl Acetate/Acetone:	https://www.doi.org/10.1021/acs.jced.8b00801
Investigation of Surface and Solubility Properties of N-Vinylimidazolium Tetrahaloferrate(III) Magnetic Ionic Liquids Using Density Functional Theory:	https://www.doi.org/10.1021/acs.jced.5b00395
Determination and Correlation of Solubility of Quetiapine Fumarate in Nine Pure Solvents and Two Aqueous Binary Solvents:	https://www.doi.org/10.1021/acs.jced.7b00518
Solubilities of evodiamine in twelve organic solvents from T = (283.2 to 323.2) K:	https://www.doi.org/10.1016/j.jct.2012.11.021
Solubility and Thermodynamic Behavior of Veramoss in Different Pure Solvents and {Ethanol + Water} Mixtures from T = 283.15 to 338.15 K:	https://www.doi.org/10.1021/acs.jced.6b00646
Solubility and Solvent Effect of Acetamiprid in Thirteen Pure Solvents and Aqueous Solutions of Ethanol:	https://www.doi.org/10.1021/acs.jced.9b00294
Determination and correlation of solubility and solution thermodynamics of valnemulin hydrogen tartrate in different pure solvents:	https://www.doi.org/10.1016/j.fluid.2014.03.022
Solubility of Diflufenican in Pure and Binary Solvent Systems:	https://www.doi.org/10.1021/acs.jced.8b00632
Solubility of Chloroquine Diphosphate and 4,7-Dichloroquinoline in Water, Ethanol, Tetrahydrofuran, Acetonitrile, and Acetone from (298.2 to 333.2) K:	https://www.doi.org/10.1021/je8007099
Measurement of activity coefficients at infinite dilution of organic solutes in the ionic liquid 1-ethyl-3-methylimidazolium 2-(2-methoxyethoxy) ethylsulfate at T = (308.15, 313.15, 323.15 and 333.15) K using gas + liquid chromatography:	https://www.doi.org/10.1016/j.jct.2013.10.017

Solubility of 3-Aminopyridine in Acetone + n-Butyl Acetate from (288.15 to 323.15) K:	https://www.doi.org/10.1021/je800945e
Solubility of 3-Nitrophthalic Acid in Different Solvents between 278 K and 353 K:	https://www.doi.org/10.1021/je0604737
Solubility Measurement and Modeling of 1-(3-nitrophenyl)Ethanone and 1-(4-nitrophenyl)Ethenone in Nine Pure Organic Solvents from T = (278.15 to 318.15) K and Mixing Properties of Solutions:	https://www.doi.org/10.1021/acs.jced.8b00192
Isobaric Vapor-Liquid Equilibrium of Acetone + Methanol System in the Presence of Calcium Bromide:	https://www.doi.org/10.1021/je049688b
Determination and correlation of cefoperazone solubility in different pure solvents and binary mixture:	https://www.doi.org/10.1016/j.fluid.2013.11.005
Fluid Phase Behavior of Nitrogen + Acetone and Oxygen + Acetone by Molecular Simulation, Experiment and the Peng-Robinson Equation of State:	https://www.doi.org/10.1021/je400691f
Solubility Measurement and Thermodynamic Modeling of N-(4-Methylphenyl-Z-3-chloro-2-(phenylthio)propenamide in 12 Pure Solvents at Temperatures Ranging from 278.15 to 318.15 K:	https://www.doi.org/10.1021/acs.jced.7b01011
Thermodynamic functions for solubility of 3-nitro-o-toluic acid in nine organic solvents from T = (283.15 to 318.15) K and apparent thermodynamic properties of solutions:	https://www.doi.org/10.1016/j.jct.2017.02.017
Solubility Measurements at 296 and 310 K and Physicochemical Characterization of Abiraterone and Abiraterone Acetate:	https://www.doi.org/10.1021/acs.jced.8b00566
Solubilities of 1,1'-(Butane-1,4-diyl)-bis(pyridinium) Dihexafluorophosphate in Acetone + Water from (278.15 to 328.15) K:	https://www.doi.org/10.1021/je800893n
Solubility and polymorphic forms of antibiotic lasalocid sodium in different organic solvents:	https://www.doi.org/10.1016/j.fluid.2014.04.008
Interactions of volatile organic compounds with the ionic liquids 1-butyl-1-methylpyrrolidinium tetracyanoborate and 1-butyl-1-methylpyrrolidinium bis(oxalato)borate:	https://www.doi.org/10.1016/j.jct.2012.09.017
Separation of (water/butan-1-ol) binary systems based on activity coefficients at infinite dilution with phosphonium ionic liquids:	https://www.doi.org/10.1016/j.jct.2019.05.011
Solubility and Thermodynamic Modeling of Dimethyl Terephthalate in Pure Solvents and the Evaluation of the Mixing Properties of the Solutions:	https://www.doi.org/10.1021/acs.jced.9b00658
Solubility determination and thermodynamic modelling of terephthaldialdehyde in ten organic solvents from T = (273.15 to 318.15) K and mixing properties of solutions:	https://www.doi.org/10.1016/j.jct.2016.07.013
Activity coefficients at infinite dilution of organic solutes in 1-octyl-3-methylimidazolium nitrate using gas–liquid chromatography:	https://www.doi.org/10.1016/j.fluid.2012.05.006
Measurement and correlation of solubility of Tetracycline hydrochloride in six organic solvents:	https://www.doi.org/10.1016/j.jct.2012.08.007
Activity coefficients at infinite dilution for organic solutes and water in 1-ethyl-1-methylpyrrolidinium lactate:	https://www.doi.org/10.1016/j.jct.2015.05.014
Activity coefficient at infinite dilution measurements for organic solutes (polar and non-polar) in fatty compounds: Saturated fatty acids:	https://www.doi.org/10.1016/j.jct.2012.06.009
Excess molar enthalpies and volumes of binary mixtures of nonafluorobutylmethylether with ketones at T = 298.15 K:	https://www.doi.org/10.1016/j.jct.2006.03.003
Vapor Liquid Equilibrium Measurements of the Binary Mixtures Nitrogen + Acetone and Oxygen + Acetone:	https://www.doi.org/10.1021/je201058j
Isobaric vapor–liquid equilibria for acetone + methanol + lithium nitrate at 100 kPa:	https://www.doi.org/10.1016/j.fluid.2006.09.007
Solubility of tridecanedioic acid in pure solvent systems: An experimental and computational study:	https://www.doi.org/10.1016/j.jct.2015.05.026
An odd-even effect on solubility of dicarboxylic acids in organic solvents:	https://www.doi.org/10.1016/j.jct.2014.05.009
Research on the application of 7-chloro-quinaldine adducts in 7-chloro-quinaldine separation process:	https://www.doi.org/10.1016/j.jct.2019.02.002
Thermodynamic properties of binary mixtures of the ionic liquid [emim][BF4] with acetone and dimethylsulphoxide:	https://www.doi.org/10.1016/j.jct.2013.01.010
(Ternary liquid + liquid) equilibria for (water + acetone + alpha-pinene, or beta-pinene, or limonene) mixtures:	https://www.doi.org/10.1016/j.jct.2010.06.004
Vapor Liquid Equilibrium Data for 2,3-Pentanedione + (Acetaldehyde or Acetone) at (100, 150, and 200) kPa:	https://www.doi.org/10.1021/acs.jced.8b01196
Activity coefficients at infinite dilution and physicochemical properties for organic solutes and water in the ionic liquid 1-(2-methoxyethyl)-1-methylpiperidinium trifluorotris(perfluoroethyl)phosphate:	https://www.doi.org/10.1016/j.jct.2012.08.016
Solubility Determination and Thermodynamic Models for 2-Methylnaphthalene in Different Solvents from T = (278.15 to 303.15) K:	https://www.doi.org/10.1021/je5010627
Solubility of methylphosphonic acid in selected organic solvents:	https://www.doi.org/10.1016/j.fluid.2015.01.009
Henry's Constants of 1-Alkanols and 2-Ketones in Salt Solutions:	https://www.doi.org/10.1021/je0600956
Measurement and Correlation of Solubility of Resorcinol Bis(cyclic 2,2-dimethyl-1,3-propanediol phosphate) in Selected Solvents from T = 293.15 to 353.15 K:	https://www.doi.org/10.1021/acs.jced.7b00695
Solid-Liquid Phase Equilibrium of N,N'-Piperazinebis(neopentylglycol)phosphoramidate in selected solvents from T = (293.15 to 333.15) K:	https://www.doi.org/10.1021/acs.jced.7b00206
Vapor-Liquid Equilibrium for Six Binary Systems of C4-Hydrocarbons + 2-Propanone:	https://www.doi.org/10.1021/je050403k
Isobaric Vapor-Liquid Equilibrium Data for the Acetone + Hexamethyl Disiloxane + Ethyl Acetate Ternary System at 101.3 kPa: Determination and Correlation:	https://www.doi.org/10.1021/acs.jced.8b00488
Solubility Measurement and Thermodynamic Properties of Levonorgestrel in Pure and Mixed Solvents:	https://www.doi.org/10.1021/acs.jced.8b00720
Isobaric vapor liquid equilibrium for acetone + methanol system containing different ionic liquids at 101.3 kPa:	https://www.doi.org/10.1016/j.fluid.2017.11.035
Solid liquid phase equilibrium and phase diagram for ternary o-nitrobenzoic acid p-nitrobenzoic acid acetone system at 283.15 K and 313.15 K:	https://www.doi.org/10.1016/j.fluid.2008.02.007
The dissolution behaviour and apparent thermodynamic analysis of diprophylline in pure and mixed solvents:	https://www.doi.org/10.1016/j.jct.2019.06.007
Measurement and Correlation of Solubility of Marbofloxacin in 12 Pure Solvents from 283.15 to 328.15 K:	https://www.doi.org/10.1021/acs.jced.9b00490
Solubility of 2-(2,4,6-Trichlorophenoxy)ethyl Bromide in Methanol, Ethanol, Propanol, Isopropanol, Acetonitrile, n-Heptane, and Acetone:	https://www.doi.org/10.1021/je200074c
Thermodynamics and activity coefficients at infinite dilution for organic compounds and water in the ionic liquid 1-butyl-3-methylimidazolium perchlorate:	https://www.doi.org/10.1016/j.jct.2017.07.012

Thermodynamics and activity coefficients at infinite dilution for organic solutes in the ionic liquid 1-hexyl-2,3-dimethylimidazolium bis(trifluoromethylsulfonyl)imide.	https://www.doi.org/10.1016/j.jct.2016.07.021
The use of ionic liquids for separation of binary hydrocarbons mixtures based on gamma infinity data measurements:	https://www.doi.org/10.1016/j.jct.2018.07.024
Dynamic Viscosities of the Binary Systems Cyclohexane and Cyclopentane with Acetone, Butanone, or 2-Pentanone at Three Temperatures T) (293.15, 298.15, and 303.15) K:	https://www.doi.org/10.1021/je050110r
Thermodynamics of Ketone + Amine Mixtures. Part VIII. Molar Excess Enthalpies at 298.15 K for n-Alkanone + Aniline or + N-Methylaniline Systems:	https://www.doi.org/10.1021/je200333p
Solubility of 1-methyl-4-nitropyrazole in seventeen pure solvents at temperatures from 283.15 K to 323.15 K:	https://www.doi.org/10.1016/j.fluid.2018.05.023
KDB Pure (Korean Thermophysical Properties Databank):	https://www.cheric.org/research/kdb/hcprop/showprop.php?cmpid=1191
Solubility Measurement and Correlation for e-2,4,6,8,10,12-Hexanitro-2,4,6,8,10,12-hexaazaisowurtzitane in Five Organic Solvents at Temperatures between 283.15 and 318.15 K and in Different Ethyl Acetate Binary Solvents at Temperatures between 283.15 and 323.15 K:	https://www.doi.org/10.1021/acs.jced.6b00664
Activity Coefficients at Infinite Dilution of Various Solutes in [illegible]	https://www.doi.org/10.1021/acs.jced.8b00600
Activity coefficients at infinite dilution of organic solvents and water in [illegible]	https://www.doi.org/10.1016/j.fluid.2016.02.004
Thermodynamics and activity coefficients at infinite dilution for organic solutes, water and diols in the ionic liquid choline bis(trifluoromethylsulfonyl)imide.	https://www.doi.org/10.1016/j.jct.2014.04.024
Separation of binary mixtures based on gamma infinity data using [illegible]	https://www.doi.org/10.1016/j.jct.2018.09.003
Activity Coefficients at Infinite Dilution of Polar Solutes in 1-(2-Hydroxyethyl)-3-methylimidazolium Tetrafluoroborate Using Gas-Liquid Chromatography:	https://www.doi.org/10.1021/je900704b
Solubility of lansoprazole in different solvents:	https://www.doi.org/10.1016/j.fluid.2012.06.011
Solubility of KCl and MgCl2 in Binary Solvents Formed by Acetone and Water in the Temperature Range between (293.15 and 323.15) K:	https://www.doi.org/10.1021/je700017b
Determination of the solubility, dissolution enthalpy and entropy of pyrene in different solvents:	https://www.doi.org/10.1016/j.fluid.2012.01.019
Solubility and Data Correlation of b-Arbutin in Different Monosolvents from 283.15 to 323.15 K:	https://www.doi.org/10.1021/acs.jced.9b00703
Solution Thermodynamics of Imidazolium-Based Ionic Liquids and Volatile Organic Compounds: Benzene and Acetone:	https://www.doi.org/10.1021/je500986b
Diffusion Coefficients of Aromatic Compounds at Infinite Dilution in Binary Mixtures at 298.15 K:	https://www.doi.org/10.1021/je060289l
Solubility of p-Methylbenzene Sulfonic Acid in Pure and Modified Supercritical Carbon Dioxide:	https://www.doi.org/10.1021/je700639s
Determination and correlation of solubility and thermodynamic properties of L-methionine in binary solvents of water + (methanol, ethanol, acetone):	https://www.doi.org/10.1016/j.jct.2015.12.022
The study of activity coefficients at infinite dilution for organic solutes and water in 1-butyl-4-methylpyridinium dicyanamide, [B4MPy][DCA] using GLC:	https://www.doi.org/10.1016/j.jct.2013.09.007
Isobaric specific heat capacities of emulsified kerosene at high temperature and pressure:	https://www.doi.org/10.1016/j.tca.2018.04.018
Solubility of Butanedioic Acid in Different Solvents at Temperatures between 283 K and 333 K:	https://www.doi.org/10.1021/je900021g
Measurement and correlation of the solubility of maleic anhydride in different organic solvents:	https://www.doi.org/10.1016/j.fluid.2014.01.008
Solubility of Diamantane, Trimantane, Tetramantane, and Their Derivatives in Organic Solvents:	https://www.doi.org/10.1021/je800277a
Vapor-liquid equilibrium in the production of the ionic liquid, 1-hexyl-3-methylimidazolium bromide ([HMIM][Br]), in acetone:	https://www.doi.org/10.1016/j.fluid.2013.11.030
New ionic liquid [P4,4,4,4][NTf2] in bio-butanol extraction on investigation of limiting activity coefficients:	https://www.doi.org/10.1016/j.fluid.2018.07.028
Isothermal vapour-liquid equilibrium data for the binary systems 2-propanone + (2-butanol or propanoic acid):	https://www.doi.org/10.1016/j.fluid.2016.10.024
Solubility of Butyl Paraben in Methanol, Ethanol, Propanol, Ethyl Acetate, Acetone, and Acetonitrile:	https://www.doi.org/10.1021/je1006289
Solubility of Ibuprofen Sodium Dihydrate in Acetone + Water Mixtures: Experimental Measurement and Thermodynamic Modeling:	https://www.doi.org/10.1021/je5004093
Measurements of activity coefficients at infinite dilution of organic solutes and water in 1-propyl-1-methylpiperidinium bis{(trifluoromethyl)sulfonyl}imide ionic liquid using g.l.c.:	https://www.doi.org/10.1016/j.jct.2010.05.017
Isobaric vapor-liquid equilibria of the binary system acetone + acetone oxime at [illegible]	https://www.doi.org/10.1016/j.fluid.2012.12.028
Separation of pyridine from heptane with tricyanomethanide-based ionic liquids:	https://www.doi.org/10.1016/j.fluid.2015.03.027
Solubility and Thermodynamic Functions of Isatin in Pure Solvents:	https://www.doi.org/10.1021/je500396b
Solubility of Candesartan Cilexetil in Different Solvents at Various Temperatures:	https://www.doi.org/10.1021/je101020m
Solubility of Deflazacort in Binary Solvent Mixtures:	https://www.doi.org/10.1021/je800790g
Measurement and Correlation of Griseofulvin Solubility in Different Solvents at Temperatures from (281.95 to 357.60) K:	https://www.doi.org/10.1021/je100030j
Experimental solubility and thermodynamic aspects of methylene blue in different solvents:	https://www.doi.org/10.1016/j.tca.2019.03.024
Solubility Measurement and Thermodynamic Modeling of Benzoic Acid in Monosolvents and Binary Mixtures:	https://www.doi.org/10.1021/acs.jced.8b00025
Measurement and correlation of the solubility of genistin in eleven organic solvents from T = (283.2 to 323.2) K:	https://www.doi.org/10.1016/j.jct.2015.05.019
Solubility of Podophyllotoxin in Six Organic Solvents from (283.2 to 323.2) K:	https://www.doi.org/10.1021/je8007028
Determination and Correlation of Solubilities of 1,3,5-Trifluoro-2,4,6-triiodobenzene in Different Solvents from 289.25 to 332.15 K:	https://www.doi.org/10.1021/acs.jced.9b00381
Measurement and Correlation of Solubility of Loratadine in Different Pure Solvents and Binary Mixtures:	https://www.doi.org/10.1021/acs.jced.6b00721

Solubility of 1-Fluoro-4-(methylsulfonyl)benzene in Five Pure Organic Solvents at Temperatures from (288.40 to 331.50) K	https://www.doi.org/10.1021/je400965y
Solubilities of Phosphonic Acid, P,P'-(1,4-Piperazinediyl)bis-,P,P,P',P'-tetraphenyl Ester in Selected Solvents	https://www.doi.org/10.1021/je400412z
Solubility of 2-isopropylimidazole in nine pure organic solvents and liquid mixture of (methanol + ethyl acetate) from T = (278.15 to 313.15) K	https://www.doi.org/10.1016/j.jct.2016.12.028
Ionic liquid 1-hexyl-3-methylimidazolium ... Experimental measurement and thermodynamic modelling ...	https://www.doi.org/10.1016/j.jct.2015.08.027
Solubility determination and correlation of cyromazine in sixteen pure solvents and mixture of acetone from ... and properties of solutions:	https://www.doi.org/10.1016/j.fluid.2018.07.024
Extractive Distillation: Isobaric Vapor-Liquid Equilibria for Acetone + Methanol + 1-Ethyl-3-methylimidazolium Trifluoromethanesulfonate	https://www.doi.org/10.1021/je0603170
Activity coefficients at infinite dilution of organic solutes in the ionic liquid ...	https://www.doi.org/10.1016/j.jct.2014.04.020
High-selective water/butan-1-ol separation on investigation of limiting activity coefficients with [P8,8,8,8][NTf2] ionic liquid	https://www.doi.org/10.1016/j.fluid.2017.06.001
Correlation of Warfarin Sodium ...	https://www.doi.org/10.1021/acs.jced.8b00977
Measurement and Correlation of the ... in Pure, Binary, and Ternary Solvent Mixtures:	https://www.doi.org/10.1021/je300647k
Solubility of Febuxostat in Four Organic Solvents at Various Temperatures	
Solubility evaluation and thermodynamic modeling of beta-lapachone in water and ten organic solvents at different temperatures	https://www.doi.org/10.1016/j.fluid.2018.05.005
Ternary phase diagram of phenanthrene and carbazole in ... and its application in the separation of them	https://www.doi.org/10.1016/j.jct.2018.11.025
Solubility of 4-Nitrobenzoic Acid in Supercritical Carbon Dioxide with and without Cosolvents:	https://www.doi.org/10.1021/je050354b
Solid–liquid equilibrium of dicyandiamide in different solvents:	https://www.doi.org/10.1016/j.fluid.2013.12.001
The determination and correlation of the solubility of naproxen in acetone and water mixtures:	https://www.doi.org/10.1016/j.fluid.2014.02.034
(Liquid + liquid) equilibria for (water + 1-propanol or acetone + .beta.-citronellol) at different temperatures:	https://www.doi.org/10.1016/j.jct.2015.02.013
Solubility of 1,1'-(Ethane-1,2-diyl)-bis(3-methylpyridinium) Dihexafluorophosphate in Different Solvents:	https://www.doi.org/10.1021/je900627q
Solubility of Adipic Acid in Acetone, Chloroform, and Toluene:	https://www.doi.org/10.1021/je8006887
Solubilities of Triadimefon in Acetone + Water from (278.15 to 333.15) K:	https://www.doi.org/10.1021/je900045z
Solubility and Melting Properties of Salicylic Acid:	https://www.doi.org/10.1021/je060134d
Measurement and Correlation of Solubility of Theobromine, Theophylline, and Caffeine in Water and Organic Solvents at Various Temperatures:	https://www.doi.org/10.1021/acs.jced.7b00065
Thermodynamic Properties of Mixtures Containing Ionic Liquids. 9. Activity Coefficients at Infinite Dilution of Hydrocarbons, Alcohols, Esters, and Aldehydes in ...	https://www.doi.org/10.1021/je050440b
Solubility of 11-Hydroxypregna-1,4,16-triene-3,20-dione in Different Solvents:	https://www.doi.org/10.1021/je800174u
Solid–Liquid Equilibrium Solubility, Thermodynamic Properties, and Solvent Effect ... Mixing ...	https://www.doi.org/10.1021/acs.jced.9b00362
... Bis(trifluoromethylsulfonyl)imide ... in ... Solvents at Different Temperatures: Temperature Range from 273.15 to 313.15 K:	https://www.doi.org/10.1021/acs.jced.7b00743
... Phosphoric Acid ... Density and ... Organic Solutions	https://www.doi.org/10.1021/je9009909
Isobaric Vapor Liquid Equilibria for Binary Systems of Acetone + Isopropenyl Acetate, 2-Butanone + Isopropenyl Acetate, and Isopropenyl Acetate + Acetylacetone at 101.3 kPa:	https://www.doi.org/10.1021/je700605t
Measurement and correlation of solubility of benzamide in supercritical carbon dioxide ...	https://www.doi.org/10.1016/j.fluid.2011.04.021
Thermodynamic study of solubility for pyrazinamide in ten solvents from T = (283.15 to 323.15) K:	https://www.doi.org/10.1016/j.jct.2017.04.014
Density and speed of sound of lithium bromide with organic solvents: Measurement and correlation:	https://www.doi.org/10.1016/j.jct.2007.04.006
Solubility of Lovastatin in Ethyl Acetate, Propyl Acetate, Isopropyl Acetate, Butyl Acetate, sec-Butyl Acetate, Isobutyl Acetate, tert-Butyl Acetate, and 2-Butanone, between (285 and 323) K:	https://www.doi.org/10.1021/je800120p
Thermodynamic Properties of Mixtures Containing Ionic Liquids. 8. Activity Coefficients ...	https://www.doi.org/10.1021/je0503554
Liquid densities of acetone and n-heptane and excess volumes of the binary system in a wide temperature ...	https://www.doi.org/10.1016/j.fluid.2010.10.028
... Speed of Sound, and Derived Thermodynamic Properties of Ionic ... 1-Butyl-3-methylimidazolium Bis(trifluoromethylsulfonyl)imide ...	https://www.doi.org/10.1021/je700648c
... Measurement about ... + Methanol or + Acetone at p = (298.15 ...) K	https://www.doi.org/10.1021/acs.jced.9b00460
... Thermodynamic ... for T = (278.15 to ...) K ... in 12 Neat ...	https://www.doi.org/10.1021/acs.jced.8b01250
Solid–Liquid Equilibria for Binary Mixtures of Methyl ... Solvents at Temperatures from 278.15 ...	https://www.doi.org/10.1021/acs.jced.7b00599
A Raman spectroscopic method for the determination of high pressure vapour ... Cyclopentyl Methyl Ether, Ethyl Acetate, n-Heptane, Methanol, and Toluene	https://www.doi.org/10.1016/j.fluid.2013.09.046
Solubility of Delphinidin in Water and Various Organic Solvents between (298.15 and 343.15) K:	https://www.doi.org/10.1021/je900851k
Experimental Measurements of the {CO2 (1) + Acetone (2) + Ivermectin (3)} System at High Pressure:	https://www.doi.org/10.1021/acs.jced.9b00072
Solubility in Different Solvents, Correlation, and Solvent Effect in the Different Crystallization Process of ...	https://www.doi.org/10.1021/acs.jced.8b01101
Determination and Correlation of Solubility and Thermodynamic Properties of trans-Cinnamyl Alcohol in Pure and Binary Solvents from 253.15 K to 293.15 K:	https://www.doi.org/10.1021/acs.jced.7b00665
Determination and Correlation of ... Solubility Data and Dissolution Thermodynamic Data of Cefixime Trihydrate in Seven Pure Solvents:	https://www.doi.org/10.1021/je5000605

Thermodynamic functions of 1-methyl-4-(methylsulfonyl)benzene solubility in nine organic solvents from T = (278.15 to 318.15) K:	https://www.doi.org/10.1016/j.jct.2016.08.021
Activity coefficients at infinite dilution of organic solutes in the ionic liquid trihexyl(tetradecyl)phosphonium tetrafluoroborate using gas liquid chromatography at T = (313.15, 333.15, 353.15, and 373.15) K:	https://www.doi.org/10.1016/j.jct.2010.12.005
Solubility of Camellianin A in Methanol, Acetonitrile, Acetone, and Water from (283.2 to 323.2) K:	https://www.doi.org/10.1021/je900554r
Experimental Measurement and Correlation of the Solubilities of 2,4-Dichloro-5-methoxypyrimidine in Ethyl Ethanoate, Methanol, Ethanol, Acetone, Tetrachloromethane, and Heptane at Temperatures between (295 and 320) K:	https://www.doi.org/10.1021/je9005689
Liquid–Liquid Equilibrium for the Ternary System of Water + Acetone + Phenol at 358 K to 368 K Under Pressure:	https://www.doi.org/10.1021/je501089z
Solubility of DL-p-hydroxyphenylglycine Sulfate in Binary Acetone + Water Solvent Mixtures:	https://www.doi.org/10.1021/je049908l
Activity Coefficients at Infinite Dilution in Methylimidazolium Nitrate Ionic Liquids:	https://www.doi.org/10.1021/je100998r
Solubility of Sulfachloropyridazine in Pure and Binary Solvent Mixtures and Investigation of Intermolecular Interactions:	https://www.doi.org/10.1021/acs.jced.7b01134
Solubility of Acephate in Different Solvents from (292.90 to 327.60) K:	https://www.doi.org/10.1021/je0603630
Crystal Structures, Reaction Rates, and Selected Physical Properties of Halo-Boronsubphthalocyanines (Halo = Fluoride, Chloride, and Bromide):	https://www.doi.org/10.1021/je3005112
Thermodynamic models for determination of the solubility of omeprazole in pure and mixture organic solvents from T = (278.15 to 333.15) K:	https://www.doi.org/10.1016/j.jct.2015.11.005
Solubilities of Ibuprofen in Different Pure Solvents:	https://www.doi.org/10.1021/je100255z
Solubility Measurement and Thermodynamic Model Correlation of 5-Nitrosalicylaldehyde in Different Solvents:	https://www.doi.org/10.1021/acs.jced.9b00232
Volumetric studies and thermodynamics of viscous flow of hydroxamic acids in acetone +water solvent at temperatures 303.15 and 318.15 K:	https://www.doi.org/10.1016/j.tca.2006.01.019
The behavior of mercury in water, alcohols, monoethylene glycol and triethylene glycol II. Elemental mercury solubility in alcohols, ethers and acetone; gas plant mercury distribution and speciation in monoethylene glycol solution:	https://www.doi.org/10.1016/j.fluid.2018.08.014
Solubility Measurement and Correlation of Fosfomycin Sodium in Six Organic Solvents and Different Binary Solvents at Temperatures between 283.15 and 323.15 K:	https://www.doi.org/10.1021/acs.jced.7b00638
Solubility of L-histidine in different aqueous binary solvent mixtures from 283.15 K to 318.15 K with experimental measurement and thermodynamic modelling:	https://www.doi.org/10.1016/j.jct.2016.09.039
Solubility, Model Correlation, and Solvent Effect of 2-Amino-3-methylbenzoic Acid in 12 Pure Solvents:	https://www.doi.org/10.1021/acs.jced.8b01226
Activity coefficients at infinite dilution for different alcohols and ketones in [Empy][ESO4]: Experimental data and modeling with PC-SAFT:	https://www.doi.org/10.1016/j.fluid.2015.09.003
Measurements of activity coefficients at infinite dilution for organic solutes and water in the ionic liquid 1-ethyl-3-methylimidazolium methanesulfonate:	https://www.doi.org/10.1016/j.jct.2012.03.005
High Pressure Vapor-Liquid Equilbria of Some Carbon Dioxide + Organic Binary Systems:	https://www.doi.org/10.1021/je0498560
Solubility of Ceftriaxone Disodium in Acetone, Methanol, Ethanol, N,N-Dimethylformamide, and Formamide between 278 and 318 K:	https://www.doi.org/10.1021/je0501989
Measurements of activity coefficients at infinite dilution of organic compounds and water in isoquinolinium-based ionic liquid [C8iQuin][NTf2] using GLC:	https://www.doi.org/10.1016/j.jct.2010.10.026
Solubility of Clopidogrel Hydrogen Sulfate (Form II) in Different Solvents:	https://www.doi.org/10.1021/je100022w
Activity coefficients at infinite dilution and physicochemical properties for organic solutes and water in the ionic liquid 4-(3-hydroxypropyl)-4-methylmorpholinium bis(trifluoromethylsulfonyl)-amide:	https://www.doi.org/10.1016/j.jct.2015.02.024
Thermodynamic study of solubility for 2-amino-4-chloro-6-methoxypyrimidine in twelve organic solvents at temperatures from 273.15 K to 323.15 K:	https://www.doi.org/10.1016/j.jct.2016.10.020
Solubility Determination, Correlation, and Solute–Solvent Molecular Interactions of 5-Aminotetrazole in Various Pure Solvents:	https://www.doi.org/10.1021/acs.jced.9b00385
Isothermal and Isobaric Vapor-Liquid Equilibria of the Ternary System of 2,2-Dimethoxypropane + Acetone + Methanol:	https://www.doi.org/10.1021/je050013y
Thermodynamic models for determination of the solid-liquid equilibrium of istradefylline in ethyl acetate plus (isopropanol, tetrahydrofuran, acetone) binary solvent mixtures:	https://www.doi.org/10.1016/j.jct.2017.03.015
A thermodynamic study of the ketoreductase-catalyzed reduction of 2-alkanones in non-aqueous solvents:	https://www.doi.org/10.1016/j.jct.2004.08.002
Measurement and Correlation of the Solubility of Dienogest in Twelve Pure and Water + Methanol Binary Solvents at Temperatures from 278.15 to 318.15 K:	https://www.doi.org/10.1021/acs.jced.9b00696
Solid-liquid equilibrium of diphenyl anilinophosphonate in the different organic solvents:	https://www.doi.org/10.1016/j.fluid.2015.03.036
Measurement and correlation of the solubility of N-(phosphonomethyl)iminodiacetic acid in different organic solvents:	https://www.doi.org/10.1016/j.fluid.2014.08.036
Solubility Measurement and Correlation of 2-Oxindole in 12 Pure Organic Solvents:	https://www.doi.org/10.1021/acs.jced.9b00308
Liquid-liquid equilibria for the system acetone + solketal + glycerol at 303.2, 313.2 and 323.2 K:	https://www.doi.org/10.1021/je500469a
Determination and Correlation of Solubilities of 2-Isopropylthioxanthone (ITX) in Seven Different Solvents from (299.15 to 329.65) K:	https://www.doi.org/10.1021/je501011t
Solubility of 1,6,6-Trimethyl-6,7,8,9-tetrahydrophenanthro[1,2-b]furan-10,11-dione in Four Organic Solvents from (283.2 to 323.9) K:	https://www.doi.org/10.1021/je800238y
Solubility measurement and correlation of the form A of ibrutinib in organic solvents from 278.15 to 323.15 K:	https://www.doi.org/10.1016/j.jct.2016.08.034
Solubility of 3,7,9,11-Tetraoxo-2,4,6,8,10-pentaaza[3.3.3]Propellane (TOPAP) in Different Pure Solvents at Temperatures between 273.15 and 318.15 K:	https://www.doi.org/10.1021/acs.jced.6b00349
Infinite dilution activity coefficients of volatile organic compounds in two ionic liquids composed of the tris(pentafluoroethyl)trifluorophosphate ([FAP]) anion and a functionalized cation:	https://www.doi.org/10.1016/j.jct.2013.05.035
Solubilities of 1,3,2-Dioxaphosphorinane-2-methanol-a,5,5-trimethyl-a-phenyl-2-oxide in Selected Solvents between 278.15 K and 333.15 K:	https://www.doi.org/10.1021/je500806u
Solubility of p-Aminobenzenesulfonamide in Supercritical Carbon Dioxide with Acetone Cosolvent:	https://www.doi.org/10.1021/je800700r
Solubility of Gastrodin in Pure and Mixed Solvents at 273.15-313.15 K and Its Correlation with Different Thermodynamic Models:	https://www.doi.org/10.1021/acs.jced.9b00286

Measurements of activity coefficients at infinite dilution for organic solutes and water in the ionic liquid 1-hexyl-3-methylimidazolium [illegible]	https://www.doi.org/10.1016/j.jct.2011.11.025
A novel non-intrusive microcell for sound-speed measurements in liquids. Speed of sound and thermodynamic properties of 2-propanone at pressures up to 160 MPa.	https://www.doi.org/10.1016/j.jct.2003.12.001
Solubility and thermodynamic properties of dirithromycin form A and B in pure solvents and binary solvent mixtures:	https://www.doi.org/10.1016/j.jct.2018.12.044
Solubility and Solution Thermodynamics of Gibberellin A4 in Different Organic Solvents from 278.15 K to 333.15 K:	https://www.doi.org/10.1021/acs.jced.5b00190
Measurement of the Phase Behavior of the Ternary System Carbon Dioxide + Acetone + Phenanthrene:	https://www.doi.org/10.1021/je0496239
Measurement of vapor-liquid equilibria (VLE) and excess enthalpies (HE) of binary systems with 1-alkyl-3-methylimidazolium bis(trifluoromethylsulfonyl)imide and prediction of these properties and γ∞ using modified UNIFAC (Dortmund):	https://www.doi.org/10.1016/j.fluid.2007.06.001
Activity coefficients at infinite dilution and physicochemical properties for organic solutes and water in the ionic liquid 1-(3-hydroxypropyl)pyridinium [illegible]	https://www.doi.org/10.1016/j.jct.2011.04.018
Thermodynamic analysis and correlation of solubility of Cefaclor in [illegible] solvents from 278.15 to 313.15 K:	https://www.doi.org/10.1016/j.fluid.2015.01.001
Measurement and correlation of solubility of trans-resveratrol in 11 solvents at T = (278.2, 288.2, 298.2, 308.2, and 318.2) K:	https://www.doi.org/10.1016/j.jct.2007.10.006
Diffusion Coefficients of Organic Compounds at Infinite Dilution in Mixtures Involving Associating Compounds. Experimental Determination and Modeling by Group Contribution Methods:	https://www.doi.org/10.1021/je700539w
Activity Coefficients at Infinite Dilution of Organic Solutes in [illegible] Tetrafluoroborate Using Gas-Liquid Chromatography:	https://www.doi.org/10.1021/je800218g
Phase equilibrium study of binary and ternary mixtures of ionic liquids + acetone + methanol:	https://www.doi.org/10.1021/je500418f
Phase equilibria of the ternary systems of (water + diethyl carbonate) with acetone, or 2-butanone at four temperatures:	https://www.doi.org/10.1016/j.jct.2016.06.033
Henry's Law Constants of Organic Compounds in Water and n-Octane at T = 293.2 K:	https://www.doi.org/10.1021/je900711h
Solubility of the Metastable Polymorph of Piracetam (Form II) in a Range of Solvents:	https://www.doi.org/10.1021/je300711r
Measurement and Correlation of Solubility of Two Isomers of Cyanopyridine in Eight Pure Solvents from 288.15 K to 318.15 K:	https://www.doi.org/10.1021/acs.jced.7b00301
Activity coefficients at infinite dilution for solutes in the trioctylmethylammonium bis(trifluoromethylsulfonyl)imide ionic liquid using gas-liquid chromatography:	https://www.doi.org/10.1016/j.jct.2009.08.012
Application of trihexyltetradecylphosphonium dicyanamide ionic liquid for various types of separations problems: Activity coefficients at infinite dilution measurements utilizing GLC method:	https://www.doi.org/10.1016/j.fluid.2019.03.023
Solubility Modeling and Solvent Effects of Allopurinol in 15 Neat Solvents:	https://www.doi.org/10.1021/acs.jced.8b00430
Phase Behavior, Densities and Isothermal Compressibility of CO2 + Pentane + Acetone Ternary System in Various Phase Regions:	https://www.doi.org/10.1021/je034160c
Thermodynamic modelling for solubility of 5-chloro-1-methyl-4-nitroimidazole in eleven organic solvents from T = (283.15 to 318.15) K:	https://www.doi.org/10.1016/j.jct.2016.10.006
Isobaric vapour-liquid equilibrium measurements and extractive distillation process for the azeotrope of (N,N-dimethylisopropylamine + acetone):	https://www.doi.org/10.1016/j.jct.2018.03.019
Separation of acetone and methanol azeotropic system using ionic liquid as entrainer:	https://www.doi.org/10.1016/j.fluid.2014.10.011
Solubility measurement and thermodynamic functions of 3-nitrobenzaldehyde in different solvents at elevated temperatures:	https://www.doi.org/10.1016/j.jct.2016.10.040
Solubility of Sulpiride in Pure Organic Solvents between (278 and 333) K:	https://www.doi.org/10.1021/je1002016
Activity Coefficients at Infinite Dilution in 1-Alkyl-3-methylimidazolium Tetrafluoroborate Ionic Liquids:	https://www.doi.org/10.1021/je050544m
p-Nitrophenylacetonitrile Solubility in Several Pure Solvents: Measurement, Correlation, and Solvent Effect Analysis:	https://www.doi.org/10.1021/acs.jced.9b00243
Measurement and Correlation of Solubility and Dissolution Thermodynamic Properties of Furan-2-carboxylic Acid in Pure and Binary Solvents:	https://www.doi.org/10.1021/je500012b
Thermochemistry of 1-bromoadamantane in binary mixtures [illegible] solvent:	https://www.doi.org/10.1016/j.tca.2005.11.035
Separation of binary mixtures hexane/hex-1-ene, cyclohexane/cyclohexene and ethylbenzene/styrene based on limiting activity coefficients:	https://www.doi.org/10.1016/j.jct.2017.03.004
Solubility of Trimethoprim (TMP) in Different Organic Solvents from (278 to 333) K:	https://www.doi.org/10.1021/je700497h
Solubilities of (a,Z)-2-Amino-a-(methoxyimino)-4-thiazoleethanethioic Acid S-2-Benzothiazolyl Ester in Different Pure Solvents and Binary Mixtures of Tetrahydrofuran + Dichloromethane or 1,2-Dichloroethane:	https://www.doi.org/10.1021/je100890f
Physicochemical properties and activity coefficients at infinite dilution for organic solutes and water in a novel bicyclic guanidinium superbase-derived protic ionic liquid:	https://www.doi.org/10.1016/j.jct.2012.09.033
Correlation of solubility and calculation of thermodynamic properties of [illegible] in different solvents:	https://www.doi.org/10.1016/j.fluid.2014.12.037
Solid-Liquid Phase Equilibria of N,N'-[1,3-Phenylenebis(methylene)]bis(phosphoramidic acid) P,P,P',P'-Tetraphenyl Ester in 10 Pure Solvents:	https://www.doi.org/10.1021/je4010917
Solubility of Andrographolide in Various Solvents from (288.2 to 323.2) K:	https://www.doi.org/10.1021/je100344z
Solubility of Enrofloxacin Sodium in Various Solvents at Various Temperatures:	https://www.doi.org/10.1021/je8005826
Solubilities of 3-Methoxy-N-phenylaniline and 3-(Methylthio)-N-phenylaniline in Five Organic Solvents (285 K to 333.75 K):	https://www.doi.org/10.1021/je3001816
Understanding the enhanced solubility of 1,3-benzenedicarboxylic acid in polar binary solvents of (acetone + water) at various temperatures:	https://www.doi.org/10.1016/j.jct.2016.10.011
Measurement and Correlation of Solubilities of Luteolin in Organic Solvents at Different Temperatures:	https://www.doi.org/10.1021/je060133l
Solubility Measurement and Modeling of 2-Amino-4,6-dichoropyrimidine in Ten Pure Solvents and (Ethyl Acetate + Ethanol) Solvent Mixtures:	https://www.doi.org/10.1021/acs.jced.8b00292
Thermodynamics and limiting activity coefficients measurements for organic solutes and water in the ionic liquid 1-dodecyl-3-methylimidzolium bis(trifluoromethylsulfonyl) imide:	https://www.doi.org/10.1016/j.jct.2016.08.008
Solubility and Molecular Interactions of Trimetazidine Hydrochloride in 12 Monosolvents and Solvent Mixtures of Methanol + (Ethanol, N,N-Dimethylformamide or Ethyl Acetate):	https://www.doi.org/10.1021/acs.jced.8b00235

Solubility of Lovastatin in Acetone, Methanol, Ethanol, Ethyl Acetate, and Butyl Acetate between 283 K and 323 K	https://www.doi.org/10.1021/je0500781
Solubility of Dioxopromethazine Hydrochloride in Different Solvents:	https://www.doi.org/10.1021/je700453v
Thermodynamic study of the solubility of 2,4'-dihydroxydiphenyl sulfone in nine organic solvents from T = (278.15 to 313.15) K and thermodynamic properties of dissolution:	https://www.doi.org/10.1016/j.jct.2016.07.001
Solubility of Rifapentine in Different Organic Solvents:	https://www.doi.org/10.1021/je7007457
Solubility Determination, Modeling, and Thermodynamic Dissolution Properties of Benzenesulfonamide in 16 Neat Solvents from 273.15 to 324.45 K:	https://www.doi.org/10.1021/acs.jced.9b00360
Densities and Viscosities of 1-Butyl-3-methylimidazolium Tetrafluoroborate + Molecular Solvent Binary Mixtures:	https://www.doi.org/10.1021/je8001128
Thermodynamics and activity coefficients at infinite dilution measurements for organic solutes and water in the ionic liquid 1-butyl-1-methylpyrrolidinium tetracyanoborate	https://www.doi.org/10.1016/j.jct.2011.06.007
Thermophysical approach to understand the nature of molecular interactions and structural factor between methylisobutylketone and organic solvents mixtures:	https://www.doi.org/10.1016/j.jct.2017.06.008
Measurement and Correlation for the Solubilities of Dihydrocapsaicin in n-Hexane, n-Pentane, Ethyl Acetate, Acetone, Ethanol, and Water:	https://www.doi.org/10.1021/je1011136
Density and Viscosity of (2,2-Dichloro-N,N-di-2-propenylacetamide + Acetone) and (2,2-Dichloro-N,N-di-2-propenylacetamide + Ethanol) at T = (278.15 to 313.15) K:	https://www.doi.org/10.1021/je900535d
Solubility of 3,3'-(Oxidi-2,1-ethanediyl)bis[1-methyl-imidazolium] Dihexafluorophosphate in Water, Methanol, Ethanol, Acetone, and Acetonitrile and Binary Mixtures (Water + DMF) from (283.00 to 328.00) K:	https://www.doi.org/10.1021/je501026m
Measurement of the Solubility of Diclazuril and Discussion of the Effect of Cosolvents:	https://www.doi.org/10.1021/acs.jced.9b00353
Solubility determination and thermodynamic modelling of allisartan isoproxil in different binary solvent mixtures from T = (278.15 to 313.15) K and mixing properties of solutions:	https://www.doi.org/10.1016/j.jct.2016.08.035
Solubilities of 1,1'-(Ethane-1,2-diyl)bis(pyridinium) Dihexafluorophosphate in Aqueous Acetone Mixtures:	https://www.doi.org/10.1021/je900524t
Solubility of dibenzothiophene in nine organic solvents: Experimental measurement and thermodynamic modelling	https://www.doi.org/10.1016/j.jct.2018.09.017
Solubility of Oleanolic Acid in Various Solvents from (288.3 to 328.3) K:	https://www.doi.org/10.1021/je700312r
Solubility Determination and Correlation of Glibenclamide in 11 Monosolvents and (Acetone + Acetonitrile) Binary Solvents from 283.15 K to 328.15 K:	https://www.doi.org/10.1021/acs.jced.8b00717
Apparent molar volumes and compressibilities of tetrabutylammonium bromide in organic solvents:	https://www.doi.org/10.1016/j.jct.2011.01.005
Measurements of Activity Coefficients at Infinite Dilution for Organic Solutes and Water in the Ionic Liquid 1-Butyl-1-methylpiperidinium Thiocyanate:	https://www.doi.org/10.1021/je101008y
Activity coefficients at infinite dilution and physicochemical properties for organic solutes and water in the ionic liquid 1-(2-hydroxyethyl)-3-methylimidazolium trifluorotris(perfluoroethyl)phosphate:	https://www.doi.org/10.1016/j.jct.2013.05.008
Influence of Typical Impurities on the Surface Tension Measurements of Binary Mixtures of Water and the Ionic Liquids 1-Butyl-3-methylimidazolium Tetrafluoroborate and Chloride:	https://www.doi.org/10.1021/je100949x
Interactions of Volatile Organic Compounds with the Ionic Liquid 1-Butyl-1-methylpyrrolidinium Dicyanamide:	https://www.doi.org/10.1021/je200822w
Solubility of Lovastatin in Acetone + Water Solvent Mixtures:	https://www.doi.org/10.1021/je800063d
Solubility of Organic Systems Containing 1,4-Dioxan-2-one:	https://www.doi.org/10.1021/je050406x
Solubility measurements and thermodynamic modeling of pyrazinamide in five different solvent-antisolvent mixtures:	https://www.doi.org/10.1016/j.fluid.2019.06.004
Thermodynamics of the solubility of lomefloxacin in methanol, ethanol, 1-propanol, 2-propanol, acetone, and chloroform from 293.15 to 323.15 K:	https://www.doi.org/10.1016/j.fluid.2012.02.011
Solubility determination and modelling of 4-nitro-1,2-phenylenediamine in eleven organic solvents from T = (283.15 to 318.15) K and thermodynamic properties of solutions:	https://www.doi.org/10.1016/j.jct.2016.11.014
Solubility of Hesperetin in Mixed Solvents:	https://www.doi.org/10.1021/je400513s
Solubility of 1,3,2-dioxaphosphorinane-5,5-dimethyl -2-(2,4,6-tribromophenoxy) 2-oxide in selected solvents:	https://www.doi.org/10.1016/j.fluid.2013.09.023
Effect of Solvent Properties and Composition on the Solubility of Ganciclovir Form I:	https://www.doi.org/10.1021/acs.jced.8b01080
Excess properties of binary mixtures containing 1-hexyl-3-methylimidazolium bis(trifluoromethylsulfonyl)imide ionic liquid and polar organic compounds:	https://www.doi.org/10.1016/j.jct.2011.11.007
Solubilities of 4,4'-Dichlorodiphenyl Disulfide in Six Organic Solvents between (303.15 and 333.15) K:	https://www.doi.org/10.1021/je301243f
Solubility determination and thermodynamic dissolution functions of 1,3-diphenylguanidine in nine organic solvents at evaluated temperatures:	https://www.doi.org/10.1016/j.jct.2016.03.011
Solubility and solution thermodynamics of ethyl 5-amino-4-cyano-3-(2-ethoxy-2-oxoethyl)-2-thiophenecarboxylate in nine organic solvents at evaluated temperatures:	https://www.doi.org/10.1016/j.jct.2015.11.025
Ammonium ionic liquids in extraction of bio-butan-1-ol from water phase using activity coefficients at infinite dilution:	https://www.doi.org/10.1016/j.fluid.2018.09.024
Solid Liquid Equilibria of Ternary 4-Nitro-2-benzofuran-1,3-dione + 5-Nitro-2-benzofuran-1,3-dione + 2-Propanone at 283.15 K and 323.15 K:	https://www.doi.org/10.1021/je301194e
Solubilities of 4-Bromo-1,8-naphthalic Anhydride in Different Pure Solvents and Binary Solvent Mixtures with the Temperature Range from (278.15 to 333.15) K:	https://www.doi.org/10.1021/je400259a
Pressurized Synthetic Methodology for Solubility Determination at Elevated Temperatures with Application to Paracetamol in Pure Solvents:	https://www.doi.org/10.1021/acs.jced.7b00118
Measurement, correlation and dissolution thermodynamics of biological active chalcone in organic solvents at different temperatures:	https://www.doi.org/10.1016/j.jct.2014.03.026
Vapor Liquid Equilibrium for the 1,1,1-Trifluorotrichloroethane + Sulfuryl Chloride System at 101.3 kPa:	https://www.doi.org/10.1021/je400544h
Vapor Liquid Equilibrium Data for Binary Mixtures of Acetic Acid + Anisole, Acetone + Anisole, and Isopropanol + Anisole at Pressure 90.15 kPa:	https://www.doi.org/10.1021/acs.jced.6b00700
On the solubility of nicotinic acid and isonicotinic acid in water and organic solvents:	https://www.doi.org/10.1016/j.jct.2013.01.024
Investigation of Surface Properties and Solubility of 1-Vinyl-3-alkyl/Esterimidazolium Halide Ionic Liquids by Density Functional Methods:	https://www.doi.org/10.1021/je500205z

Ternary and Binary LLE Measurements for Solvent (4- Methyl-2-pentanone and 2-Methyl-2-butanol) + Furfural + Water between 298 and 401 K:	https://www.doi.org/10.1021/acs.jced.5b00738
Solubility Measurement, Correlation and Molecular Interactions of 3-Methyl-6-nitroindazole in Different Neat Solvents and Mixed Solvents from T = 278.15 to 328.15 K:	https://www.doi.org/10.1021/acs.jced.8b01256
Partial Molar Volumes of Butyltriethylammonium Iodide in Single nonaqueous solvents at 298.15 K:	https://www.doi.org/10.1021/je900542y
Solubility of itaconic acid in different organic solvents: Experimental measurement and thermodynamic modeling:	https://www.doi.org/10.1016/j.fluid.2011.09.027
Thermodynamics of ketone + amine mixtures. Part IX. Excess molar enthalpies at 298.15K for dipropylamine, or dibutylamine + 2-alkanone systems and modeling of linear or aromatic amine + 2-alkanone mixtures in terms of DISQUAC and ERAS:	https://www.doi.org/10.1016/j.fluid.2013.01.011
Solubility of benzoic acid in acetone, 2-propanol, acetic acid and cyclohexane: Experimental measurement and thermodynamic modeling:	https://www.doi.org/10.1016/j.fluid.2010.06.021
Solubility Measurement and Phase Equilibrium Modeling of [illegible] Pure Solvents:	https://www.doi.org/10.1021/acs.jced.9b00350
Solubility Determination and Modelling of Sodium Mononitrobenzenesulphonate in Different Solvents at Temperatures Ranging from (283.15 to 323.15) K:	https://www.doi.org/10.1021/acs.jced.5b00783
Solubility of Amorphous Clopidogrel Hydrogen Sulfate in Different Pure Solvents:	https://www.doi.org/10.1021/acs.jced.5b00306
Assessment of Pyrrolidinium-Based Ionic Liquid for the Separation of Binary Mixtures Based on Activity Coefficients at Infinite Dilution:	https://www.doi.org/10.1021/acs.jced.9b00341
Solubility Measurement and Correlation of Probenecid in 12 Pure Organic Solvents and Thermodynamic Properties of Mixing of Solutions:	https://www.doi.org/10.1021/acs.jced.8b00863
Determination and Correlation of Solubilities for Succinic Acid in the Adipic Acid + Glutaric Acid + Acetone Mixture, and Adipic Acid in the Succinic Acid + Glutaric Acid + Acetone Mixture:	https://www.doi.org/10.1021/acs.jced.6b00145
Solubility of Naphthalene in Isobutyl Acetate, n-Butyric Acid, Ethyl Acetate, [illegible] N,N-Dimethylformamide, and Tetrahydrofuran,	https://www.doi.org/10.1021/je800428r
Solubility of CO2 in Propanone, 1-Ethyl-3-methylimidazolium Tetrafluoroborate, and Their Mixtures:	https://www.doi.org/10.1021/je100343v
Thermodynamics of 4'-bromomethyl-2-cyanobiphenyl in different solvents:	https://www.doi.org/10.1016/j.jct.2014.12.002
Solubilities of 4',5,7-Triacetoxyflavanone in Fourteen Organic Solvents at Different Temperatures:	https://www.doi.org/10.1021/acs.jced.6b00929
Solubilities of Adipic Acid and Succinic Acid in Glutaric Acid + Acetone or n-Butanol Mixture:	https://www.doi.org/10.1021/je500682v
Solid-liquid phase equilibrium of glyphosate in selected solvents:	https://www.doi.org/10.1016/j.fluid.2012.05.003
High-Pressure Multiphase Behavior of the Ternary Systems (Ethene + Water + Acetone) and (Ethane + Water + Acetone):	https://www.doi.org/10.1021/je100609r
Solubility Measurement and Thermodynamic Modeling for o-Toluenesulfonamide in 16 Solvents from T = 273.15 to 323.85 K:	https://www.doi.org/10.1021/acs.jced.9b00445
Thermodynamics of amide + ketone mixtures. 1. Volumetric, speed of sound and refractive index data for N,N-dimethylformamide + 2-alkanone systems at several temperatures:	https://www.doi.org/10.1016/j.jct.2016.02.016
Solubilities of 2,4,8,10-Tetraoxa-3,9-diphosphaspiro[5.5]undecane-3,9-dimethanol, 3,9-dioxide in Selected Solvents:	https://www.doi.org/10.1021/je700050r
Solubility measurement and thermodynamic functions of dehydroepiandrosterone acetate in different solvents at evaluated temperatures:	https://www.doi.org/10.1016/j.jct.2016.01.023
Study of Acoustic Parameters in Binary Mixture at Variable Frequencies:	https://www.doi.org/10.1007/s10765-013-1527-7
Solubility of 2-Aminopyridine in Acetone, Chloroform, and Ethyl Acetate:	https://www.doi.org/10.1021/je800768x
Measurement and correlation of the solubilities of tetra(5,5-dimethyl-1,3-dioxaphosphorinanyl-2-oxy) neopentane in different pure solvents:	https://www.doi.org/10.1016/j.fluid.2014.01.043
Solubility determination and thermodynamic modeling of 3-methyl-4-nitrobenzoic acid in twelve organic solvents from T = (283.15-318.15) K and mixing properties of solutions:	https://www.doi.org/10.1016/j.jct.2016.07.023
Activity coefficients at infinite dilution measurements for organic solutes and water in the 1-hexyloxymethyl-3-methyl-imidazolium and 1,3-dihexyloxymethyl-imidazolium bis(trifluoromethylsulfonyl)-imide ionic liquids. The cation influence:	https://www.doi.org/10.1016/j.fluid.2009.08.017
Measurement and correlation solubility and mixing properties of dimethyl [illegible] mixture at 278.15 K to [illegible] of solutions:	https://www.doi.org/10.1016/j.fluid.2013.06.037
Measurement and Modeling of Epigallocatechin Gallate Solubility in Supercritical Carbon Dioxide Fluid with Ethanol Cosolvent:	https://www.doi.org/10.1021/je901025f
Measurement and correlation of vapor liquid equilibria of binary systems containing the ionic liquids [EMIM][(CF3SO2)2N], [BMIM][(CF3SO2)2N], [MMIM][(CH3)2PO4] and oxygenated organic compounds respectively water:	https://www.doi.org/10.1016/j.fluid.2005.01.002
Solubility determination and correlation for 1,8-dinitronaphthalene in (acetone + [illegible] methanol) and [illegible]	https://www.doi.org/10.1016/j.jct.2015.10.015
Vapor-liquid equilibria for systems of diethyl carbonate and ketones and determination of group interaction parameters for the UNIFAC and ASOG methods:	https://www.doi.org/10.1016/j.fluid.2005.06.024
Separation of binary mixtures based on limiting activity coefficients data using [illegible] ammonium-based ionic liquid and modelling of thermodynamic functions:	https://www.doi.org/10.1016/j.fluid.2017.12.029
Determination of Henry's Law Constants Using Internal Standards with Benchmark Values:	https://www.doi.org/10.1021/je3010535
Measurement and correlation of phase diagram data for acetone andsulfate aqueous two-phase systems at different temperatures:	https://www.doi.org/10.1016/j.tca.2013.07.002
Activity Coefficients at Infinite Dilution for Hydrocarbons in Fatty Alcohols Determined by Gas-Liquid Chromatography:	https://www.doi.org/10.1021/je1005517
Solubilities of Some Chlorophenols in Supercritical CO2 in the Presence and Absence of Cosolvents:	https://www.doi.org/10.1021/je900328c
Activity Coefficients at Infinite Dilution of Polar Solutes in 1-Butyl-3-methylimidazolium Tetrafluoroborate Using Gas-Liquid Chromatography:	https://www.doi.org/10.1021/je060305e
Solubilities of Tris(o-phenylenedioxy)cyclotriphosphazene in Selected Solvents:	https://www.doi.org/10.1021/je200252c
Solubility and Thermodynamic Modeling of Sulfanilamide in 12 Mono Solvents and 4 Binary Solvent Mixtures from 278.15 to 318.15 K:	https://www.doi.org/10.1021/acs.jced.9b00411
Determination, Correlation, and Application of Sodium l-Ascorbate Solubility in Nine Pure Solvents and Two Binary Solvents at Temperatures from 278.15 to 323.15 K:	https://www.doi.org/10.1021/acs.jced.7b00846
Solubility of Tiamulin Hydrogen Fumarate in Acetone, Acetonitrile, Ethyl Acetate, Ethyl Formate, and Butyl Acetate from (288.2 to 318.2) K:	https://www.doi.org/10.1021/je800838w

Solubility of 4-Hydroxybenzaldehyde in Supercritical Carbon Dioxide with and without Cosolvents:	https://www.doi.org/10.1021/je401082x
Equilibrium Solubility, Model Correlation, and Solvent Effect of Indole-3-acetic Acid in Twelve Pure Solvents:	https://www.doi.org/10.1021/acs.jced.8b01265
Excess parameter studies on the binary mixtures of toluene with ketones at different temperatures:	https://www.doi.org/10.1016/j.jct.2009.12.008
Solid-liquid phase equilibrium and dissolution properties of ethyl vanillin in pure solvents:	https://www.doi.org/10.1016/j.jct.2016.10.029
Thermodynamics and activity coefficients at infinite dilution for organic solutes in the ionic liquid 1-butyl-1-methylpyrrolidinium dicyanamide:	https://www.doi.org/10.1016/j.fluid.2018.06.013
Solubilities of 3-Carboxy-3-hydroxypentanedioic Acid in Ethanol, Butan-1-ol, Water, Acetone, and Methylbenzene:	https://www.doi.org/10.1021/je101167z
Solubility of 2-(4-Ethylbenzoyl)benzoic Acid in Eleven Organic Solvents between 279.65 K and 343.15 K:	https://www.doi.org/10.1021/je700426k
Interactions of volatile organic compounds with the ionic liquid 1-ethyl-3-methylimidazolium tetracyanoborate:	https://www.doi.org/10.1016/j.jct.2011.09.028
Measurements of activity coefficients at infinite dilution for organic solutes and water in the ionic liquid 1-butyl-1-methylpyrrolidinium tricyanomethanide:	https://www.doi.org/10.1016/j.jct.2013.07.004
Solubility of Canthaxanthin in Pure Solvents from (293.15 to 343.15) K:	https://www.doi.org/10.1021/je800801x
Influence of sulfate-based anion ionic liquids on the separation factor of the binary azeotropic system acetone + methanol:	https://www.doi.org/10.1016/j.fluid.2012.12.006
Isobaric Vapor–Liquid Equilibrium for Nine Binary Systems of Cracking C5 Fraction at 250 kPa:	https://www.doi.org/10.1021/acs.jced.6b00180
Phase Behavior, Densities, and Isothermal Compressibility of CO2 + Pentane and CO2 + Acetone Systems in Various Phase Regions:	https://www.doi.org/10.1021/je034087q
Solubility of Cyclotrimethylenetrinitramine (RDX) in Binary Solvent Mixtures:	https://www.doi.org/10.1021/je7002463
Liquid–Liquid Equilibria of Water + Benzonitrile + Propanone, or + Dimethyl Sulfoxide:	https://www.doi.org/10.1021/je030115t
Isobaric vapor-liquid equilibrium for ternary mixtures of acetone + methanol + ionic liquids at 101.3kPa:	https://www.doi.org/10.1016/j.fluid.2017.03.014
Solubility of Artemisinin in Different Single and Binary Solvent Mixtures Between (284.15 and 323.15) K and NRTL Interaction Parameters:	https://www.doi.org/10.1021/je100125x
Solubility of Puerarin in Water, Ethanol, and Acetone from (288.2 to 328.2) K:	https://www.doi.org/10.1021/je050076g
Measurement and Correlation of Solubility of l-Valine in Water + (Ethanol, N,N-Dimethylformamide, Acetone, Isopropyl Alcohol) from 293.15 K to 343.15 K:	https://www.doi.org/10.1021/je500255d
Phase diagrams of acetonitrile or (acetone + water + EPPS) buffer phase separation systems at 298.15 K and quantum chemical modeling:	https://www.doi.org/10.1016/j.jct.2012.03.026
Solid–Liquid Equilibrium of Azacyclotridecan-2-one in 15 Pure Solvents from T = 273.15 to 323.15 K: Experimental Determination and Thermodynamic Modeling:	https://www.doi.org/10.1021/acs.jced.8b01193
Measurement and Correlation of Solubility of Azithromycin Monohydrate in Five Pure Solvents:	https://www.doi.org/10.1021/je400899e
Measurement and correlation of ternary solid-liquid equilibrium of 2-methyl-1,4-naphthoquinone + phthalic anhydride + acetone system:	https://www.doi.org/10.1016/j.fluid.2015.09.019
KDB:	https://www.cheric.org/research/kdb/hcprop/showprop.php?cmpid=1191
Solubility of 4-(3,4-Dichlorophenyl)-1-tetralone in Some Organic Solvents:	https://www.doi.org/10.1021/je060329l
Solubility of oxcarbazepine in eight solvents within the temperature range T = (288.15–308.15) K:	https://www.doi.org/10.1016/j.jct.2016.09.011
Solubility of 3-Chlorophthalic Anhydride and 4-Chlorophthalic Anhydride in Organic Solvents and Solubility of 3-Chlorophthalic Acid and 4-Chlorophthalic Acid in Water from (283.15 to 333.15) K:	https://www.doi.org/10.1021/je800869g
Thermodynamic equilibrium of hydroxyacetic acid in pure and binary solvent systems:	https://www.doi.org/10.1016/j.jct.2017.01.004
Vapour + liquid equilibria for (2-ethoxypropene + acetone) and (2-ethoxypropene + butanone):	https://www.doi.org/10.1016/j.jct.2005.10.001
Solubilities of 2,4-Dinitro-l-phenylalanine in Monosolvents at (273.15 to 368.15) K:	https://www.doi.org/10.1021/je100352r
Determination and modeling of the solubility of (limonin in methanol or acetone + water) binary solvent mixtures at T = 283.2 K to 318.2 K:	https://www.doi.org/10.1016/j.jct.2016.03.037
Phase Boundaries of CO2 + Toluene, CO2 + Acetone, and CO2 + Ethanol at High Temperatures and High Pressures:	https://www.doi.org/10.1021/je060099a
Solubility measurement, model evaluation and thermodynamic analysis of rivaroxaban polymorphs in organic solvents:	https://www.doi.org/10.1016/j.jct.2016.09.036
Vapor-Liquid Phase Boundaries of Binary Mixtures of Carbon Dioxide with Ethanol and Acetone:	https://www.doi.org/10.1021/je800371a
Vapor–Liquid Equilibrium for Binary Systems of 2,3-Pentanedione with Diacetyl and Acetone:	https://www.doi.org/10.1021/je7005924
Solubility of paroxetine hydrochloride hemi-hydrate in (water + acetone):	https://www.doi.org/10.1016/j.jct.2004.12.006
Measurement and correlation of solubility of dodecanedioic acid in different pure solvents from T = (288.15 to 323.15) K:	https://www.doi.org/10.1016/j.jct.2013.09.012
Determination and modelling of troxerutin solubility in eleven mono-solvents and (1,4-dioxane + 2-propanol) binary solvents at temperatures from 288.15 K to 323.15 K:	https://www.doi.org/10.1016/j.jct.2016.09.015
Solubility, dissolution enthalpy and entropy of L-Glutamine in mixed solvents of ethanol+water and acetone+water:	https://www.doi.org/10.1016/j.tca.2013.02.007
Solubility and solution thermodynamics of 2-methyl-4-nitroaniline in eleven organic solvents at elevated temperatures:	https://www.doi.org/10.1016/j.jct.2016.10.037
Solubility of β-Sitosteryl Maleate and Stigmasteryl Maleate in Acetone and Ethyl Acetate:	https://www.doi.org/10.1021/je7000396
Solubility of Flavonoids in Organic Solvents:	https://www.doi.org/10.1021/je7001094
Solubility of Difloxacin in Acetone, Methanol, and Ethanol from (293.15 to 313.15) K:	https://www.doi.org/10.1021/je800742d

Measurement and Correlation of Solubility of Cefathiamidine in Water + (Acetone, Ethanol, or 2-Propanol) from (278.15 to 308.15) K:	https://www.doi.org/10.1021/acs.jced.5b00617
Synthesis and Solubility of 5,5-Dimethyl-2-(phenyl(phenylamino)methyl)-1,3,2-dioxaphosphinane 2-Oxide in Selected Solvents between 278.15 K and 347.15 K:	https://www.doi.org/10.1021/acs.jced.7b00585
Thermodynamics of risperidone and solubility in pure organic solvents:	https://www.doi.org/10.1016/j.fluid.2014.04.028
Solubility Determination of Nicotinamide and Its Application for the Cocrystallization with Benzoic Acid:	https://www.doi.org/10.1021/acs.jced.8b00560
Solubility determination and thermodynamic modelling for 2-amino-4-chlorobenzoic acid in eleven organic solvents from T = (278.15 to 313.15) K and mixing properties of solutions:	https://www.doi.org/10.1016/j.jct.2016.11.019
Solubility and Metastable Zone Width of 3-Amino-9-ethylcarbazole in Acetone:	https://www.doi.org/10.1021/je0600552
Activity coefficients at infinite dilution of organic solutes in the ionic liquid 1-ethyl-3-methylimidazolium methanesulfonate:	https://www.doi.org/10.1016/j.fluid.2010.10.008
1-Alkylcyanopyridinium-based ionic liquid in the separation processes:	https://www.doi.org/10.1016/j.jct.2016.01.017
Measurement and Correlation for the Solubility of Dimethyl 1,4-Cyclohexanedione-2,5-dicarboxylate in Different Solvents at Temperatures from (278.15 to 323.15) K:	https://www.doi.org/10.1021/je2000292
Activity coefficients at infinite dilution measurements for organic solutes and water in the ionic liquid 4-methyl-N-butyl-pyridinium bis(trifluoromethylsulfonyl)-imide:	https://www.doi.org/10.1016/j.jct.2009.06.011
Density and Viscosity of Ketones with Toluene at Different Temperatures and at Atmospheric Pressure:	https://www.doi.org/10.1021/je900523k
Activity Coefficients at Infinite Dilution of Polar Solutes in 1-Propyl-2,3-dimethylimidazolium Tetrafluoroborate Using Gas-Liquid Chromatography:	https://www.doi.org/10.1021/je700640r
Determination and Correlation of Solubility of Borneol, Camphor, and Isoborneol in Different Solvents:	https://www.doi.org/10.1021/acs.jced.9b00045
Crippen Method:	http://pubs.acs.org/doi/abs/10.1021/ci990307l
Measurement and Correlation of the Solubility of Tetramethylpyrazine in Nine Monosolvents and Two Binary Solvent Systems:	https://www.doi.org/10.1021/acs.jced.8b00888
Determination of Activity Coefficients at Infinite Dilution of Solutes in N,N-Di(2-ethylhexyl)isobutyramide Using Inverse Gas Liquid Chromatography:	https://www.doi.org/10.1021/acs.jced.8b00635
Measurement of Henry's Law Constants for Acetone, 2-Butanone, 2,3-Butanedione, and Isobutyraldehyde Using a Horizontal Flow Reactor:	https://www.doi.org/10.1021/je034137r
Solubilities of Nizatidine in Water, Methanol, Ethanol, Acetone, and Ethyl Acetate from (273.15 to 343.15) K:	https://www.doi.org/10.1021/je1000928
Solubility of propanoic acid 3-(hydroxyphenylphosphinyl)-sodium salts in different solvent:	https://www.doi.org/10.1016/j.fluid.2013.02.001
Thermostabilities and solubilities of flame retardant bisphenol S-bis (5,5-dimethyl-1,3-dioxaphosphorinanyl-2-oxy phosphate ester) as a function of temperature:	https://www.doi.org/10.1016/j.fluid.2014.04.006
Solubility Measurement and Thermodynamic Modeling of 4-Nitrophthalimide in Twelve Pure Solvents at Elevated Temperatures Ranging from (273.15 to 323.15) K:	https://www.doi.org/10.1021/acs.jced.6b00230
Thermodynamic Properties of Mixtures Containing Precursors of Vitamin B5:	https://www.doi.org/10.1021/je020060f
Phase equilibria and interfacial tensions in the systems methyl tert-butyl ether + acetone + cyclohexane, methyl tert-butyl ether + acetone and methyl tert-butyl ether + cyclohexane:	https://www.doi.org/10.1016/j.fluid.2008.08.012
Solubility determination and thermodynamic modelling of 3,5-dimethylpyrazole in nine organic solvents from T = (283.15 to 313.15) K and mixing properties of solutions:	https://www.doi.org/10.1016/j.jct.2017.02.011
Solubility of 6-Chloropyridazin-3-amine in Different Solvents:	https://www.doi.org/10.1021/je3001188
Solubility of Vitamin D3 in Six Organic Solvents at Temperatures from (248.2 to 273.2) K:	https://www.doi.org/10.1021/je300401c
Solubility Determination and Thermodynamic Mixing Properties of 5-Methyl-2-pyrazinecarboxylic Acid in Different Solvents:	https://www.doi.org/10.1021/acs.jced.9b00406
Measurement and Correlation of Solubility of Calcium-l-lactate Pentahydrate in Ethanol + Water and Acetone + Water Systems:	https://www.doi.org/10.1021/je500416k
Solubility of 2-Chlorophenylacetic Acid in 12 Pure Solvents from T = (273.15/283.15 to 318.15) K: Determination and Modeling:	https://www.doi.org/10.1021/acs.jced.8b00139
Determination and Correlation of Ethyl Vanillin Solubility in Different Binary Solvents at Temperatures from 273.15 to 313.15 K:	https://www.doi.org/10.1021/acs.jced.6b00972
4-Amino-3,6-dichloropyridazine Solubility Measurement and Correlation in Seven Pure Organic Solvents from (278.15 to 333.15) K:	https://www.doi.org/10.1021/je500286x
Solubility Measurement and Modeling of 3-Hydroxy-2-nitropyridine in Ten Pure Solvents and Two Binary Mixed Solvents for T = (278.15–318.15) K:	https://www.doi.org/10.1021/acs.jced.9b00564
Determination of the thermodynamic parameters of ionic liquid 1-propyl-3-methylimidazolium bromide by gas-liquid chromatography:	https://www.doi.org/10.1016/j.jct.2018.09.023
Determination of the solubility, dissolution enthalpy and entropy of 3-pentadecylphenyl acrylate in different organic solvents:	https://www.doi.org/10.1016/j.tca.2012.06.025
Solubilities of 1,1'-(Hexane-1,6-diyl)bis(pyridinium) Dihexafluorophosphate in (Water + Acetone):	https://www.doi.org/10.1021/je9004855
Experimental and Modeling Studies on the Solubility of 2-Chloro-N-(4-methylphenyl)propanamide (S1) in Binary Ethyl Acetate + Hexane, Toluene + Hexane, Acetone + Hexane, and Butanone + Hexane Solvent Mixtures Using Polythermal Method:	https://www.doi.org/10.1021/acs.jced.7b00288
Experimental Measurements and Modeling of the Solubility of Aceclofenac in Six Pure Solvents from (293.35 to 338.25) K:	https://www.doi.org/10.1021/je500038u
Solubility of L-Phenylalanine Anhydrous and Monohydrate Forms: Experimental Measurements and Predictions:	https://www.doi.org/10.1021/je201354k
Solubility of L-(+)-ascorbic acid in water, ethanol, methanol, propan-2-ol, acetone, acetonitrile, ethyl acetate, and tetrahydrofuran from (293 to 323) K:	https://www.doi.org/10.1021/je800056h
Solubility and solution thermodynamics of 2-methyl-6-nitroaniline in ten organic solvents at elevated temperatures:	https://www.doi.org/10.1016/j.jct.2016.05.027
Measurement and Correlation of the Solubility of Vanillic Acid in Eight Pure and Water + Ethanol Mixed Solvents at Temperatures from (293.15 to 323.15) K:	https://www.doi.org/10.1021/acs.jced.5b00619
Thermodynamic properties of acetone calculated from accurate experimental speed of sound measurements at low temperatures and high pressures:	https://www.doi.org/10.1016/j.jct.2008.11.005
Solubility Measurement and Thermodynamic Model Correlation and Evaluation of 2-Chloro-5-nitroaniline in 12 Pure Solvents:	https://www.doi.org/10.1021/acs.jced.8b00931

Experimental Investigation of the Solubility of CO_2 in (Acetone + Water):	https://www.doi.org/10.1021/je600571v
Solubility Determination of 2-Chloronicotinic Acid and Analysis of Solvent Effect:	https://www.doi.org/10.1021/acs.jced.9b00661
Abraham model linear free energy relationships for describing the partitioning and solubility behavior of nonelectrolyte organic solutes dissolved in pyridine at 298.15 K:	https://www.doi.org/10.1016/j.fluid.2016.10.009
Solubility of sodium diclofenac in different solvents:	https://www.doi.org/10.1016/j.fluid.2007.07.020
Activity coefficient at infinite dilution measurements for organic solutes (polar and non-polar) in fatty compounds - Part II: C18 fatty acids:	https://www.doi.org/10.1016/j.jct.2012.12.009
Solubility determination and thermodynamic modelling of 3-amino-1,2,4-triazole in ten organic solvents from T = 283.15 K to T = 318.15 K and mixing properties of solutions:	https://www.doi.org/10.1016/j.jct.2016.09.033
High-pressure solubility of naproxen, nicotinamide and their mixture in acetone with supercritical CO_2 as an antisolvent:	https://www.doi.org/10.1016/j.fluid.2014.03.029
Vapor-liquid-liquid equilibria of perfluorohexane +CO_2 + methanol, +toluene, and +acetone at 313K:	https://www.doi.org/10.1016/j.fluid.2005.12.041
Dissolution behavior and thermodynamic properties of lapatinib ditosylate in pure and mixed organic solvents from T = (283.15-323.15) K:	https://www.doi.org/10.1016/j.fluid.2019.01.005
Densities and Excess Molar Volumes for Binary Mixtures of Ionic Liquid 1-Ethyl-3-methylimidazolium Ethylsulfate with Solvents:	https://www.doi.org/10.1021/je1002237
A novel dynamic recirculating apparatus for vapour-liquid equilibrium measurements at moderate pressures and temperatures:	https://www.doi.org/10.1016/j.fluid.2013.07.044
Solubility modelling and thermodynamic dissolution functions of phthalimide in ten organic solvents:	https://www.doi.org/10.1016/j.jct.2015.10.024
Measurements of activity coefficients at infinite dilution in vegetable oils and capric acid using the dilutor technique:	https://www.doi.org/10.1016/j.fluid.2013.10.035
McGowan Method:	http://link.springer.com/article/10.1007/BF02311772
Solubilities of N-[(4-Bromo-3,5-difluorine)-phenyl]maleimide in Pyridine, Acetone, Tetrahydrofuran, Ethanol, Trichloromethane, and Acetonitrile at Temperature between (285.00 and 350.15) K:	https://www.doi.org/10.1021/je100819f
Solubility of CO_2 in Acetone 1-Butyl-3-methylimidazolium Tetrafluoroborate and Their Mixtures:	https://www.doi.org/10.1021/je300611q
Measurement and Correlation of the Solubility of 2-Cyanoacetamide in 14 Pure Solvents and the Mixing Properties of Solutions:	https://www.doi.org/10.1021/acs.jced.8b01205
Solubility determination and thermodynamic modeling of 2,4-dinitroaniline in nine organic solvents from T = (278.15 to 318.15) K and mixing properties of solutions:	https://www.doi.org/10.1016/j.jct.2016.07.009
Solubility of Fluorene in Benzene, Chloroform, Acetone, 1-Propanol, [illegible] (283 to 323) K:	https://www.doi.org/10.1021/je800676s
Activity coefficients at infinite dilution, physicochemical and thermodynamic properties for organic solutes and water in the ionic liquid ethyl-dimethyl-(2-methoxyethyl)ammonium tris(pentafluoroethyl)trifluorophosphate:	https://www.doi.org/10.1016/j.jct.2015.05.022
Excess Molar Enthalpies for Binary Mixtures of Ethanol + Acetone, + Octane, + Cyclohexane and 1-Propanol [illegible] at 323.15:	https://www.doi.org/10.1021/je020149l
In situ measurement of liquid-liquid equilibria by medium field nuclear magnetic resonance:	https://www.doi.org/10.1016/j.fluid.2017.01.027
Solid-liquid equilibrium of 1-hydroxyethane-1,1-diphosphonic acid in different solvents:	https://www.doi.org/10.1016/j.fluid.2012.11.001
Thermodynamic models for determination of the solid-liquid equilibrium of oxytocin in (acetonitrile + acetone) binary solvent mixtures:	https://www.doi.org/10.1016/j.fluid.2015.08.009
Dissociation Constants and Solubilities of Dalbergin and Nordalbergin in Different Solvents:	https://www.doi.org/10.1021/acs.jced.5b00098
Solid-liquid equilibrium and thermodynamic of 2,5-thiophenedicarboxylic acid in different organic solvents:	https://www.doi.org/10.1016/j.fluid.2014.05.003
Ionic Conductivity of Selected 2:1 Electrolytes in Dilute Solutions of Mixed Aqueous Organic Solvents at 298.15 K:	https://www.doi.org/10.1021/je700313j
Liquid-Liquid Equilibrium of (Water + Acetone) with Cumene or α-Methylstyrene or Phenol at Temperatures of (323.15 and 333.15) K:	https://www.doi.org/10.1021/je050500d
Liquid-liquid equilibria of 1,3-dimethylimidazolium methylsulfate with ketones, dialkyl carbonates and acetates:	https://www.doi.org/10.1016/j.fluid.2007.02.027
Solubilities of Phosphoramidic Acid, N-(phenylmethyl)-, Diphenyl Ester in Selected Solvents:	https://www.doi.org/10.1021/acs.jced.5b00007
Compact apparatus for rapid measurement of high-pressure phase equilibria of carbon dioxide expanded liquids:	https://www.doi.org/10.1016/j.fluid.2009.12.030
Measurement of Solubility of Thiamine Hydrochloride Hemihydrate in Three Binary Solvents and Mixing Properties of Solutions:	https://www.doi.org/10.1021/acs.jced.6b00613
Liquid Liquid Equilibrium for the Ternary System Acetone Oxime Methyl Ether Acetone Water:	https://www.doi.org/10.1021/je301319g
Experimental determination and calculation of the critical curves for the binary systems of CO_2 containing ketone, alkane, ester and alcohol, respectively:	https://www.doi.org/10.1016/j.fluid.2011.05.005
Measurement and Correlation of Solubility of Cefuroxime Acid in Pure and Binary Solvents at Various Temperatures:	https://www.doi.org/10.1021/je100397q
Activity coefficients at infinite dilution and physicochemical properties for organic solutes and water in the ionic liquid trihexyl-tetradecyl-phosphonium tricyanomethanide:	https://www.doi.org/10.1016/j.jct.2018.01.003
Determination of Henry's Law Constants and Activity Coefficients at Infinite Dilution of Flavor Compounds in Water at 298 K with a Gas-Chromatographic Method:	https://www.doi.org/10.1021/je0495942
Experimental Data and Thermodynamics Modeling (PC-SAFT EoS) for CO_2 + Acetone + Pluronic P 127 System at High Pressures:	https://www.doi.org/10.1021/acs.jced.8b01163
Dissociation Constants and Solubilities of Daidzein and Genistein in Different Solvents:	https://www.doi.org/10.1021/je4010905
Dortmund Data Bank Vapor-Liquid Equilibrium Data:	http://www.ddbst.com/en/EED/VLE/VLE%20Acetonitrile%3BAcetone.php
Solubility determination and thermodynamic modeling of paclobutrazol in nine organic solvents from T = (278.15 to 318.15) K and mixing properties of solutions:	https://www.doi.org/10.1016/j.jct.2016.09.038
Solubilities of Rutaecarpine in Twelve Organic Solvents from (283.2 to 323.2)	https://www.doi.org/10.1021/je4001334
Physical properties and their corresponding changes of mixing for the ternary mixture acetone + n-hexane +water at 298.15K:	https://www.doi.org/10.1016/j.tca.2006.01.008

Solubilities of Lauric Acid in n-Hexane, Acetone, Propanol, 2-Propanol, 1-Bromopropane, and Trichloroethylene from (279.0 to 318.3) K:	https://www.doi.org/10.1021/je800739y
Liquid-Liquid Equilibrium Data for the Separation of Acetone from n-Heptane Using Four Imidazolium-Based Ionic Liquids:	https://www.doi.org/10.1021/acs.jced.8b01105
Thermodynamic Functions for Solubility of 1-Hydroxybenzotriazole in Sixteen Solvents at Temperatures from (278.15 to 313.15) K and Mixing Property of Mixtures:	https://www.doi.org/10.1021/acs.jced.7b00316
Solubility of Thiotriazinone in Ethanol, 1-Propanol, 2-Propanol, 1-Butanol, Acetonitrile, Acetone, Ethyl Acetate, and Water from 293 K to 343 K:	https://www.doi.org/10.1021/je900177h
Determination and correlation of solubility and solution thermodynamics of saccharin in different pure solvents:	https://www.doi.org/10.1016/j.jct.2019.02.005
Equilibrium Solubility of Biapenem in Different Neat and Binary Solvents: Experimental Determination and Model Correlation:	https://www.doi.org/10.1021/acs.jced.8b01051
Solubility determination and thermodynamic functions of 2-chlorophenothiazine in nine organic solvents from T = 283.15 K to T = 318.15 K and mixing properties of solutions:	https://www.doi.org/10.1016/j.jct.2016.11.029
Pressure-Density-Temperature Behavior of CO2/Acetone, CO2/Toluene, and CO2/Monochlorobenzene Mixtures in the Near-Critical Region:	https://www.doi.org/10.1021/je0342771
Solubility of N-phenylanthranilic acid in nine organic solvents from T = (283.15 to 318.15) K: Determination and modelling:	https://www.doi.org/10.1016/j.jct.2016.08.017
Solubility measurement and correlation of 4-nitrophthalimide in (methanol, ethanol, or acetone) + N,N-dimethylformamide mixed solvents at temperatures from 273.15 K to 323.15 K:	https://www.doi.org/10.1016/j.jct.2016.08.013
Solubility of Isoniazid in Various Organic Solvents from (301 to 313) K:	https://www.doi.org/10.1021/je800156m
Quaternary, Ternary, and Binary LLE Measurements for 2-Methoxy-2-methylbutane + Furfural + Acetic Acid + Water at Temperatures between 298 and 341 K:	https://www.doi.org/10.1021/acs.jced.6b00150
Solubility Measurement and the Correlation of 1-Naphthaleneacetic Acid in Pure and Methanol + Water Binary Solvents from T = (278.25 to 323.55) K:	https://www.doi.org/10.1021/acs.jced.6b00816
Determination and Correlation of Pyridazin-3-amine Solubility in Eight Organic Solvents at Temperatures Ranging from (288.05 to 333.35) K:	https://www.doi.org/10.1021/je300517q
1-Ethyl-3-methylimidazolium Dicyanamide as a Very Efficient Entrainer for the Extractive Distillation of the Acetone + Methanol System:	https://www.doi.org/10.1021/je200972w
Solubility Determination and Correlation for Pyridoxal 5-Phosphate Monohydrate in Different Binary Solvents from 278.15 K to 318.15 K:	https://www.doi.org/10.1021/acs.jced.8b00309
Tetramethylammonium chloride + glycerol deep eutectic solvent as separation agent for organic liquid mixtures: Assessment from experimental limiting activity coefficients:	https://www.doi.org/10.1016/j.fluid.2018.06.003
High pressure phase equilibrium data for carbon dioxide p dichloromethane p acetone and carbon dioxide p dichloromethane p acetone p N-acetylcysteine (NAC):	https://www.doi.org/10.1016/j.fluid.2018.06.008
NIST Webbook:	http://webbook.nist.gov/cgi/cbook.cgi?ID=C67641&Units=SI
Joback Method:	https://en.wikipedia.org/wiki/Joback_method
Activity Coefficients at Infinite Dilution of Organic Solutes and Water in Tributylethylphosphonium Diethylphosphate Using Gas Liquid Chromatography: Thermodynamic Properties of Mixtures Containing Ionic Liquids:	https://www.doi.org/10.1021/acs.jced.5b00980
Apparent Molar Volume and Compressibility of [illegible] Bromide in [illegible] Solvents:	https://www.doi.org/10.1021/je900888e
Solubility and Crystal Size of Sirolimus in Different Organic Solvents:	https://www.doi.org/10.1021/je100626x
Excess Enthalpies of Dibromomethane with Acetone, 1,4-Dioxane, Pyridine, Diethyl Ether, Ethyl Methyl Ketone, and Tetrahydrofuran at 303.15 K:	https://www.doi.org/10.1021/je900547w
Solubilities of 1,4-Naphthoquinone in Acetone, Toluene, Xylene, Ethanol, and n-Butyl Alcohol:	https://www.doi.org/10.1021/je800405b
Separation of water/butan-1-ol based on activity coefficients at infinite dilution in 1,3-didecyl-2-methylimidazolium dicyanamide ionic liquid:	https://www.doi.org/10.1016/j.jct.2017.10.003
Quaternary and ternary LLE measurements for solvent (2-methyltetrahydrofuran and cyclopentyl methyl ether) + furfural + acetic acid + water between 298 and 308 K:	https://www.doi.org/10.1016/j.jct.2017.12.015
Solubility of KBr in Binary Solvents Formed by Acetone and Water in the Temperature Range between (288.15 and 323.15) K:	https://www.doi.org/10.1021/je8003639
Measurement and Correlation of the Solubility of Mifepristone in Eight Pure and Water + Methanol Mixed Solvents at Temperatures from 278.15 to 318.15 K:	https://www.doi.org/10.1021/acs.jced.8b01062
Solubilities of Organic Semiconductors and Nonsteroidal Anti-inflammatory Drugs in Pure and Mixed Organic Solvents: Measurement and Modeling with Hansen Solubility Parameter:	https://www.doi.org/10.1021/acs.jced.8b00536
Determination and Correlation of Solubility Data and Dissolution Thermodynamic Data of [illegible] in Different Pure Solvents:	https://www.doi.org/10.1021/je301014d
Solubility and Metastable Zone of Cefoperazone Sodium in Acetone + Water System:	https://www.doi.org/10.1021/je800837z
Vapor liquid equilibria and interfacial tensions for the ternary system acetone + 2,2'-oxybis[propane] + cyclohexane and its constituent binary systems:	https://www.doi.org/10.1016/j.fluid.2008.06.006
Screening of environmental friendly ionic liquid as a solvent for the different types of separations problem: insight from activity coefficients at infinite dilution measurement using (gas + liquid) chromatography technique:	https://www.doi.org/10.1016/j.jct.2015.08.017
Measurement and prediction of the solubilities of aromatic polyimide monomers in [illegible]:	https://www.doi.org/10.1016/j.jct.2012.04.013
Determination of Activity Coefficients at Infinite Dilution of 35 Solutes in the Ionic Liquid, 1-Butyl-3-methylimidazolium Tosylate, Using Gas-Liquid Chromatography:	https://www.doi.org/10.1021/je100410k
Physical properties of the binary systems methylcyclopentane with ketones (acetone, butanone and 2-pentanone) at T = (293.15, 298.15, and 303.15) K. New UNIFAC-VISCO interaction parameters:	https://www.doi.org/10.1016/j.jct.2005.08.002
Experimental Study of Thermodynamic Properties of Mixtures Containing Ionic Liquid 1-Ethyl-3-methylimidazolium Ethyl Sulfate Using Gas-Liquid Chromatography and Transpiration Method:	https://www.doi.org/10.1021/je0602723
Solubility of daidzin in different organic solvents and (ethyl alcohol + water) mixed solvents:	https://www.doi.org/10.1016/j.jct.2015.04.025
Densities and Excess Molar Volumes for Acetone + Propionic Acid + Water from (283.15 to 323.15) K at Atmospheric Pressure:	https://www.doi.org/10.1021/je034250h
Thermodynamic analysis and correlation of solubility of candesartan cilexetil in aqueous solvent mixtures:	https://www.doi.org/10.1016/j.fluid.2012.09.027
Solubility modelling and dissolution properties of 5-phenyltetrazole in thirteen mono-solvents and liquid mixtures of (methanol + ethyl acetate) at elevated temperatures:	https://www.doi.org/10.1016/j.jct.2017.04.019
Thermodynamics and selectivity of separation based on activity coefficients at infinite dilution of various solutes in 1-allyl-3-methylimidazolium bis{(trifluoromethyl)sulfonyl}imide ionic liquid:	https://www.doi.org/10.1016/j.jct.2016.06.028

Isobaric Vapor Liquid Equilibrium for Acetone + Methanol + Phosphate Ionic Liquids:	https://www.doi.org/10.1021/je5007373
Determination and Correlation of Solubility of Phenylbutazone in Monosolvents and Binary Solvent Mixtures:	https://www.doi.org/10.1021/acs.jced.6b00911
Measurement of solubility of erythromycin acetone solvate in aqueous acetone solution between 298K and 323K:	https://www.doi.org/10.1016/j.fluid.2008.10.016
Density, Viscosity, and Excess Properties of Binary Liquid Mixtures of Propylene Carbonate with Polar and Nonpolar Solvents:	https://www.doi.org/10.1021/je034204h
Solubilities of Triphenylphosphine in Ethanol, 2-Propanol, Acetone, Benzene, and Toluene:	https://www.doi.org/10.1021/je060350m
Surface Tension and Interfacial Tension of Binary Organic Liquid Mixtures:	https://www.doi.org/10.1021/je034062r
Binary Equilibrium Solubility of Amidinothiourea in Monosolvents: Experimental Determination, Model Correlation, and Solvent Effect Analysis:	https://www.doi.org/10.1021/acs.jced.9b00458
Solubility of Irbesartan (Form A) in Different Solvents between 278 K and 323 K:	https://www.doi.org/10.1021/je700296x

Legend

af:	Acentric Factor
affp:	Proton affinity
aigt:	Autoignition Temperature
basg:	Gas basicity
chg:	Standard gas enthalpy of combustion
chl:	Standard liquid enthalpy of combustion
cpg:	Ideal gas heat capacity
cpl:	Liquid phase heat capacity
cps:	Solid phase heat capacity
dm:	Dipole Moment
dvisc:	Dynamic viscosity
ea:	Electron affinity
fll:	Lower Flammability Limit
flu:	Upper Flammability Limit
fpc:	Flash Point (Closed Cup Method)
fpo:	Flash Point (Open Cup Method)
gf:	Standard Gibbs free energy of formation
gyrad:	Radius of Gyration
hf:	Enthalpy of formation at standard conditions
hfl:	Liquid phase enthalpy of formation at standard conditions
hfus:	Enthalpy of fusion at standard conditions
hfust:	Enthalpy of fusion at a given temperature
hvap:	Enthalpy of vaporization at standard conditions
hvapt:	Enthalpy of vaporization at a given temperature
ie:	Ionization energy
log10ws:	Log10 of Water solubility in mol/l
logp:	Octanol/Water partition coefficient
mcvol:	McGowan's characteristic volume
nfpaf:	NFPA Fire Rating
nfpah:	NFPA Health Rating

pc: Critical Pressure
pvap: Vapor pressure
rfi: Refractive Index
rhoc: Critical density
rhol: Liquid Density
rinpol: Non-polar retention indices
ripol: Polar retention indices
sfust: Entropy of fusion at a given temperature
sl: Liquid phase molar entropy at standard conditions
speedsl: Speed of sound in fluid
srf: Surface Tension
tb: Normal Boiling Point Temperature
tbrp: Boiling point at reduced pressure
tc: Critical Temperature
tf: Normal melting (fusion) point
tt: Triple Point Temperature
vc: Critical Volume
zc: Critical Compressibility
zra: Rackett Parameter