Acetic acid

Other names:	Acetasol
	Acetic acid, glacial
	Aci-jel
	Acide acetique
	Acido acetico
	Azijnzuur
	CH3COOH
	Essigsaeure
	Ethanoic acid
	Ethanoic acid monomer
	Ethylic acid
	Glacial acetic acid
	Kyselina octova
	Methanecarboxylic acid
	NSC 132953
	Octowy kwas
	Shotgun
	UN 2789
	Vinegar acid
Inchi:	InChI=1S/C2H4O2/c1-2(3)4/h1H3,(H,3,4)
InchiKey:	QTBSBXVTEAMEQO-UHFFFAOYSA-N
Formula:	C2H4O2
SMILES:	CC(=O)O
Mol. weight [g/mol]:	60.05
CAS:	64-19-7

Physical Properties

Property code	Value	Unit	Source
af	0.4470		KDB
affp	783.70	kJ/mol	NIST Webbook
aigt	699.82	K	KDB
basg	752.80	kJ/mol	NIST Webbook
chl	-874.50 ± 0.40	kJ/mol	NIST Webbook
chl	-875.16 ± 0.34	kJ/mol	NIST Webbook
chl	-872.40	kJ/mol	NIST Webbook
chl	-874.20 ± 0.20	kJ/mol	NIST Webbook
dm	1.30	debye	KDB

fll	5.40	% in Air	KDB
flu	16.00	% in Air	KDB
fpc	317.59	K	KDB
fpo	313.15	K	KDB
gf	-376.90	kJ/mol	KDB
gyrad	2.5950		KDB
hf	-435.40	kJ/mol	NIST Webbook
hf	-432.50 ± 1.60	kJ/mol	NIST Webbook
hf	-432.50 ± 1.60	kJ/mol	NIST Webbook
hf	-432.90 ± 1.50	kJ/mol	NIST Webbook
hf	-435.40 ± 4.30	kJ/mol	NIST Webbook
hf	-431.90 ± 1.50	kJ/mol	NIST Webbook
hf	-431.90 ± 1.50	kJ/mol	NIST Webbook
hf	-435.10	kJ/mol	KDB
hf	-432.90 ± 1.50	kJ/mol	NIST Webbook
hfl	-487.00	kJ/mol	NIST Webbook
hfl	-484.10 ± 0.40	kJ/mol	NIST Webbook
hfl	-484.50 ± 0.20	kJ/mol	NIST Webbook
hfl	-483.52 ± 0.36	kJ/mol	NIST Webbook
hfus	6.62	kJ/mol	Joback Method
hvap	51.60 ± 1.60	kJ/mol	NIST Webbook
hvap	50.30	kJ/mol	NIST Webbook
hvap	51.60	kJ/mol	NIST Webbook
hvap	51.60 ± 1.50	kJ/mol	NIST Webbook
ie	10.38 ± 0.03	eV	NIST Webbook
ie	10.87	eV	NIST Webbook
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ie	10.64 ± 0.00	eV	NIST Webbook
ie	10.66	eV	NIST Webbook
ie	10.65	eV	NIST Webbook

log10ws	1.22		Aqueous Solubility Prediction Method
logp	0.091		Crippen Method
mcvol	46.480	ml/mol	McGowan Method
nfpaf	%!d(float64=2)		KDB
nfpah	%!d(float64=3)		KDB
nfpas	%!d(float64=1)		KDB
pc	5787.00 ± 101.32	kPa	NIST Webbook
pc	5786.70 ± 26.66	kPa	NIST Webbook
pc	5829.01 ± 90.00	kPa	NIST Webbook
pc	5786.00	kPa	KDB
pc	5781.00 ± 20.00	kPa	NIST Webbook
pc	5786.00 ± 8.00	kPa	NIST Webbook
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ripol	1418.00	NIST Webbook
ripol	1444.00	NIST Webbook
ripol	1429.00	NIST Webbook
ripol	1460.00	NIST Webbook
ripol	1434.00	NIST Webbook
ripol	1448.00	NIST Webbook
ripol	1448.00	NIST Webbook
ripol	1464.00	NIST Webbook
ripol	1436.00	NIST Webbook
ripol	1443.00	NIST Webbook
ripol	1465.00	NIST Webbook
ripol	1480.00	NIST Webbook
ripol	1463.00	NIST Webbook
ripol	1486.00	NIST Webbook
ripol	1433.00	NIST Webbook
ripol	1467.00	NIST Webbook
ripol	1453.00	NIST Webbook
ripol	1424.00	NIST Webbook
ripol	1441.00	NIST Webbook
ripol	1434.00	NIST Webbook
ripol	1463.00	NIST Webbook
ripol	1479.00	NIST Webbook
ripol	1464.00	NIST Webbook
ripol	1460.00	NIST Webbook
ripol	1429.00	NIST Webbook
ripol	1449.00	NIST Webbook
ripol	1452.00	NIST Webbook
ripol	1461.00	NIST Webbook
ripol	1436.00	NIST Webbook
ripol	1460.00	NIST Webbook

ripol	1427.00		NIST Webbook
ripol	1452.00		NIST Webbook
ripol	1402.00		NIST Webbook
ripol	1435.00		NIST Webbook
ripol	1445.00		NIST Webbook
ripol	1452.00		NIST Webbook
ripol	1449.00		NIST Webbook
ripol	1447.00		NIST Webbook
ripol	1400.00		NIST Webbook
ripol	1400.00		NIST Webbook
ripol	1408.00		NIST Webbook
ripol	1465.00		NIST Webbook
ripol	1465.00		NIST Webbook
ripol	1450.00		NIST Webbook
ripol	1428.00		NIST Webbook
ripol	1439.00		NIST Webbook
sg	282.84	J/mol×K	NIST Webbook
sl	158.00	J/mol×K	NIST Webbook
sl	193.70	J/mol×K	NIST Webbook
tb	391.00	K	Vapor liquid equilibria for water + acetic acid + (N,N-dimethylformamide or dimethyl sulfoxide) at 13.33 kPa
tb	391.20	K	Liquid-Liquid Equilibria of Water + Acetic Acid + Dimethyl Glutarate Ternary System
tb	390.93	K	Study of Vapor-Liquid Equilibria for Acetic Acid + n-Propyl Acetate + Isopropyl Acetate Systems
tb	391.20	K	(Liquid + liquid) equilibria of the (water + acetic acid + dibutyl phthalate) system
tb	391.03	K	Isobaric vapor-liquid equilibrium for water + acetic acid + 1-butyl-3-methylimidazolium dibutylphosphate at 101.32 kPa
tb	390.95	K	Ternary liquid-liquid phase equilibria of (water-carboxylic acid-1-undecanol) systems at 298.15 K
tb	391.05	K	Quaternary phase equilibrium of water-carboxylic acid mixture (formic-propionic acid or acetic-propionic acid)-solvent liquid systems at 298.15 K

tb	390.85	K	Measurements of Quaternary Liquid-Liquid Equilibrium for Water + Acetic Acid + Propionic Acid + Solvent (Butyronitrile, Benzyl Acetate, or Methyl Isobutyl Ketone) at 298.15 K
tb	391.05	K	Determination and correlation of liquid liquid equilibria for the (water + carboxylic acid + dimethyl maleate) ternary systems at T = 298.2K
tb	390.85	K	Isobaric vapor liquid equilibria for water + acetic acid + (N-methyl pyrrolidone or N-methyl acetamide)
tb	391.20	K	Liquid liquid equilibria of the ternary system water + acetic acid + dimethyl succinate
tb	390.95	K	Vapor liquid equilibria for the quaternary reactive system ethyl acetate + ethanol +water + acetic acid and some of the constituent binary systems at 101.3 kPa
tb	391.20	K	Liquid liquid equilibria of the ternary system water + acetic acid + dimethyl adipate
tb	391.00	K	KDB
tb	391.03	K	Isobaric Vapor-Liquid Equilibria for Water + Acetic Acid + 1-Ethyl-3-methylimidazolium Diethylphosphate at 101.32 kPa
tb	390.96	K	Investigation on Isobaric Vapor Liquid Equilibrium for Water + Acetic Acid + sec-Butyl Acetate
tb	391.15	K	Liquid-Liquid Equilibria of (Water + Acetic Acid + Diethyl Succinate or Diethyl Glutarate or Diethyl Adipate) Ternary Systems
tb	391.44	K	Isobaric Vapor-Liquid Equilibria for (Acetic Acid + Cyclohexane) and (Cyclohexane + Acetylacetone) at a Pressure of 101.3 kPa and for (Acetic Acid + Acetylacetone) at a Pressure of 60.0 kPa
tc	592.71	K	KDB
tf	289.70 ± 0.03	K	NIST Webbook

tf	289.93	K	Aqueous Solubility Prediction Method
tf	289.70	K	KDB
tf	289.69 ± 0.20	K	NIST Webbook
tf	289.95	K	Differential scanning calorimetry determination of phase diagrams and water activities of aqueous carboxylic acid solutions
tf	289.67 ± 0.05	K	NIST Webbook
tf	289.80 ± 0.30	K	NIST Webbook
tf	289.75 ± 0.05	K	NIST Webbook
tf	289.49 ± 0.05	K	NIST Webbook
tf	289.84	K	Solid liquid equilibrium in the ternary system acetic acid propanoic acid formamide
tf	289.62 ± 0.10	K	NIST Webbook
tf	289.85	K	NIST Webbook
tf	289.00 ± 1.50	K	NIST Webbook
tt	289.80 ± 0.15	K	NIST Webbook
tt	289.69	K	KDB
tt	289.80 ± 0.05	K	NIST Webbook
tt	289.69 ± 0.04	K	NIST Webbook
vc	0.171	m3/kmol	KDB
zc	0.2007690		KDB
zra	0.22		KDB

Temperature Dependent Properties

Property code	Value	Unit	Temperature [K]	Source
cpg	98.37	J/mol×K	567.09	Joback Method
cpg	95.29	J/mol×K	537.78	Joback Method
cpg	92.08	J/mol×K	508.46	Joback Method
cpg	88.73	J/mol×K	479.15	Joback Method
cpg	85.25	J/mol×K	449.84	Joback Method
cpg	81.64	J/mol×K	420.52	Joback Method
cpg	77.89	J/mol×K	391.21	Joback Method
cpl	139.70	J/mol×K	332.00	NIST Webbook

cpl	135.50	J/mol×K	313.15	Heat Capacities of Binary Mixtures of Acetic Acid with Acetic Anhydride and Methenamine at Different Temperatures
cpl	123.10	J/mol×K	298.15	NIST Webbook
cpl	120.50	J/mol×K	298.00	NIST Webbook
cpl	121.30	J/mol×K	297.10	NIST Webbook
cpl	137.21	J/mol×K	333.15	Heat Capacities of Binary Mixtures of Acetic Acid with Acetic Anhydride and Methenamine at Different Temperatures
cpl	123.50	J/mol×K	298.00	NIST Webbook
cpl	159.80	J/mol×K	298.10	NIST Webbook
cpl	123.40	J/mol×K	294.70	NIST Webbook
cpl	137.00	J/mol×K	311.00	NIST Webbook
dvisc	0.0011140	Pa×s	298.15	Densities and Viscosities of Binary Mixtures of Acetic Acid with Acetic Anhydride and Methenamine at Different Temperatures
dvisc	0.0005700	Pa×s	353.15	Densities and Viscosities of N,N-Dimethylformamide + Formic Acid, and + Acetic Acid in the Temperature Range from (303.15 to 353.15) K
dvisc	0.0006320	Pa×s	343.15	Densities and Viscosities of N,N-Dimethylformamide + Formic Acid, and + Acetic Acid in the Temperature Range from (303.15 to 353.15) K

dvisc	0.0008040	Pa×s	323.15	Densities and Viscosities of N,N-Dimethylformamide + Formic Acid, and + Acetic Acid in the Temperature Range from (303.15 to 353.15) K
dvisc	0.0009200	Pa×s	313.15	Densities and Viscosities of N,N-Dimethylformamide + Formic Acid, and + Acetic Acid in the Temperature Range from (303.15 to 353.15) K
dvisc	0.0007870	Pa×s	323.15	Densities and Viscosities of Binary Mixtures of Acetic Acid with Acetic Anhydride and Methenamine at Different Temperatures
dvisc	0.0010590	Pa×s	303.15	Densities and Viscosities of N,N-Dimethylformamide + Formic Acid, and + Acetic Acid in the Temperature Range from (303.15 to 353.15) K
dvisc	0.0012110	Pa×s	293.15	Dynamic Viscosities, Densities, and Speed of Sound and Derived Properties of the Binary Systems Acetic Acid with Water, Methanol, Ethanol, Ethyl Acetate and Methyl Acetate at T = (293.15, 298.15, and 303.15) K at Atmospheric Pressure

dvisc	0.0010340	Pa×s	303.15	Dynamic Viscosities, Densities, and Speed of Sound and Derived Properties of the Binary Systems Acetic Acid with Water, Methanol, Ethanol, Ethyl Acetate and Methyl Acetate at T = (293.15, 298.15, and 303.15) K at Atmospheric Pressure	
dvisc	0.0012040	Pa×s	293.15	Densities and Viscosities of Binary Mixtures of Acetic Acid with Acetic Anhydride and Methenamine at Different Temperatures	
dvisc	0.0012040	Pa×s	293.15	Densities and Viscosities of Binary Mixtures of Acetic Acid with Acetic Anhydride and Methenamine at Different Temperatures	
dvisc	0.0006570	Pa×s	338.15	Densities and Viscosities of Binary Mixtures of Acetic Acid with Acetic Anhydride and Methenamine at Different Temperatures	
dvisc	0.0006970	Pa×s	333.15	Densities and Viscosities of Binary Mixtures of Acetic Acid with Acetic Anhydride and Methenamine at Different Temperatures	
dvisc	0.0006970	Pa×s	333.15	Densities and Viscosities of Binary Mixtures of Acetic Acid with Acetic Anhydride and Methenamine at Different Temperatures	

dvisc	0.0007410	Pa×s	328.15	Densities and Viscosities of Binary Mixtures of Acetic Acid with Acetic Anhydride and Methenamine at Different Temperatures
dvisc	0.0007410	Pa×s	328.15	Densities and Viscosities of Binary Mixtures of Acetic Acid with Acetic Anhydride and Methenamine at Different Temperatures
dvisc	0.0011140	Pa×s	298.15	Densities and Viscosities of Binary Mixtures of Acetic Acid with Acetic Anhydride and Methenamine at Different Temperatures
dvisc	0.0007870	Pa×s	323.15	Densities and Viscosities of Binary Mixtures of Acetic Acid with Acetic Anhydride and Methenamine at Different Temperatures
dvisc	0.0008410	Pa×s	318.15	Densities and Viscosities of Binary Mixtures of Acetic Acid with Acetic Anhydride and Methenamine at Different Temperatures
dvisc	0.0008410	Pa×s	318.15	Densities and Viscosities of Binary Mixtures of Acetic Acid with Acetic Anhydride and Methenamine at Different Temperatures
dvisc	0.0009020	Pa×s	313.15	Densities and Viscosities of Binary Mixtures of Acetic Acid with Acetic Anhydride and Methenamine at Different Temperatures

dvisc	0.0011150	Pa×s	298.15	Dynamic Viscosities, Densities, and Speed of Sound and Derived Properties of the Binary Systems Acetic Acid with Water, Methanol, Ethanol, Ethyl Acetate and Methyl Acetate at T = (293.15, 298.15, and 303.15) K at Atmospheric Pressure
dvisc	0.0009640	Pa×s	308.15	Densities and Viscosities of Binary Mixtures of Acetic Acid with Acetic Anhydride and Methenamine at Different Temperatures
dvisc	0.0009640	Pa×s	308.15	Densities and Viscosities of Binary Mixtures of Acetic Acid with Acetic Anhydride and Methenamine at Different Temperatures
dvisc	0.0010370	Pa×s	303.15	Densities and Viscosities of Binary Mixtures of Acetic Acid with Acetic Anhydride and Methenamine at Different Temperatures
dvisc	0.0010370	Pa×s	303.15	Densities and Viscosities of Binary Mixtures of Acetic Acid with Acetic Anhydride and Methenamine at Different Temperatures
dvisc	0.0009020	Pa×s	313.15	Densities and Viscosities of Binary Mixtures of Acetic Acid with Acetic Anhydride and Methenamine at Different Temperatures

dvisc	0.0007090	Pa×s	333.15	Densities and Viscosities of N,N-Dimethylformamide + Formic Acid, and + Acetic Acid in the Temperature Range from (303.15 to 353.15) K
hfust	11.72	kJ/mol	298.70	NIST Webbook
hfust	11.72	kJ/mol	298.70	NIST Webbook
hfust	11.13	kJ/mol	290.06	NIST Webbook
hfust	11.52	kJ/mol	283.70	NIST Webbook
hfust	10.83	kJ/mol	289.80	NIST Webbook
hfust	11.73	kJ/mol	289.90	NIST Webbook
hsubt	67.00 ± 1.00	kJ/mol	221.50	NIST Webbook
hsubt	70.00 ± 1.00	kJ/mol	221.50	NIST Webbook
hvapt	38.80	kJ/mol	559.00	NIST Webbook
hvapt	38.10	kJ/mol	486.00	NIST Webbook
hvapt	38.70	kJ/mol	419.00	NIST Webbook
hvapt	41.60	kJ/mol	340.50	NIST Webbook
hvapt	43.00	kJ/mol	308.00	NIST Webbook
hvapt	40.30	kJ/mol	358.00	NIST Webbook
hvapt	41.60	kJ/mol	351.00	NIST Webbook
hvapt	42.00	kJ/mol	343.00	NIST Webbook
hvapt	40.90	kJ/mol	357.50	NIST Webbook
hvapt	37.90	kJ/mol	470.50	NIST Webbook
hvapt	23.70	kJ/mol	391.10	NIST Webbook
hvapt	23.68	kJ/mol	391.30	KDB
hvapt	39.10	kJ/mol	364.00	NIST Webbook
pvap	101.32	kPa	391.03	Isobaric Vapor-Liquid Equilibria for Water + Acetic Acid + 1-Ethyl-3-methylimidazolium Diethylphosphate at 101.32 kPa
pvap	49.34	kPa	368.10	Separation of Furfural from Ternary Mixtures
pvap	88.98	kPa	386.07	Separation of Furfural from Ternary Mixtures

pvap	96.15	kPa	388.91	Vapor Liquid Equilibrium Data for Binary Mixtures of Acetic Acid + Anisole, Acetone + Anisole, and Isopropanol + Anisole at Pressure 96.15 kPa
pvap	7.57	kPa	323.15	Vapor-Liquid Equilibria in the Propyl Acetate + Ethanoic Acid Binary System from (323.15 to 353.15) K: Measurement with a Static Method and Modeling with the NRTL, Wilson, UNIQUAC, and COSMO-SAC Approaches
pvap	18.66	kPa	343.15	Vapor-Liquid Equilibria in the Propyl Acetate + Ethanoic Acid Binary System from (323.15 to 353.15) K: Measurement with a Static Method and Modeling with the NRTL, Wilson, UNIQUAC, and COSMO-SAC Approaches
pvap	12.04	kPa	333.15	Vapor-Liquid Equilibria in the Propyl Acetate + Ethanoic Acid Binary System from (323.15 to 353.15) K: Measurement with a Static Method and Modeling with the NRTL, Wilson, UNIQUAC, and COSMO-SAC Approaches

pvap	28.20	kPa	353.15	Vapor-Liquid Equilibria in the Propyl Acetate + Ethanoic Acid Binary System from (323.15 to 353.15) K: Measurement with a Static Method and Modeling with the NRTL, Wilson, UNIQUAC, and COSMO-SAC Approaches
pvap	20.00	kPa	345.27	Isobaric Vapor Liquid Equilibrium for Binary Systems of Toluene + Acrylic Acid, Toluene + Acetic Acid, and Cyclohexane + Acrylic Acid at 20 kPa
pvap	7.72	kPa	323.15	Isothermal vapour liquid equilibrium with chemical reaction in the quaternary water + methanol + acetic acid + methyl acetate system, and in five binary subsystems
rfi	1.36870		293.15	Volumetric Properties of Highly Nonideal Binary Mixtures Containing Ethanoic Acid and Propanoic Acid with Butan-2-ol, Methyl-2-propanol, and 2-Methyl-2-butanol at Different Temperatures
rfi	1.36980		298.15	Densities, Excess Molar Volumes, Viscosity, and Refractive Indices of Binary Mixtures of Ethanoic Acid and Trichloroethylene with Dimethylbenzenes at Different Temperatures

rfi	1.37160	298.15	Geometric Structures of Associating Component Optimized toward Correlation and Prediction of Isobaric Vapor Liquid Equilibria for Binary and Ternary Mixtures of Ethanal, Ethanol, and Ethanoic Acid	
rfi	1.37000	298.15	Solubilities of Benzoic Acid and Phthalic Acid in Acetic Acid + Water Solvent Mixtures	
rfi	1.36960	298.15	Phase Equilibrium for the Esterification Reaction of Acetic Acid + Butan-1-ol at 101.3 kPa	
rfi	1.37200	293.15	Isobaric Vapor-Liquid Equilibria for the Binary Systems of Acetic Acid + Isopropenyl Acetate, Acetic Acid + Acetylacetone, and Water + Acetylacetone	
rfi	1.36960	298.15	Liquid-Liquid Equilibria for the System 1-Methyl Propyl Ethanoate (1) + Acetic Acid (2) + Water (3) at (283.15 and 323.15) K	
rfi	1.37160	293.15	Solubilities of Phosphorus-Containing Compounds in Selected Solvents	
rfi	1.37160	298.15	Isobaric Vapor-Liquid Equilibria for Binary and Ternary Mixtures of Methanol, Ethanoic Acid, and Propanoic Acid	

rfi	1.37180	293.15	Solubilities of Some Phosphaspirocyclic Compounds in Selected Solvents	
rfi	1.36969	298.15	Refractive Index, Surface Tension, and Density of Aqueous Mixtures of Carboxylic Acids at 298.15 K	
rfi	1.37190	293.15	Isobaric Vapor-Liquid Equilibria for Water + Acetic Acid + (n-Pentyl Acetate or Isopropyl Acetate)	
rfi	1.36809	303.15	Solubility, Thermodynamic Properties, and Derived Excess Properties of Benzoic Acid in (Acetic Acid + Water) and (Acetic Acid + Toluene) Binary Mixtures	
rfi	1.37012	298.15	Solubility, Thermodynamic Properties, and Derived Excess Properties of Benzoic Acid in (Acetic Acid + Water) and (Acetic Acid + Toluene) Binary Mixtures	
rfi	1.37224	293.15	Solubility, Thermodynamic Properties, and Derived Excess Properties of Benzoic Acid in (Acetic Acid + Water) and (Acetic Acid + Toluene) Binary Mixtures	

rfi	1.37203	298.00	Quaternary and ternary LLE measurements for solvent (2-methyltetrahydrofuran and cyclopentyl methyl ether) + furfural + acetic acid + water between 298 and 343 K
rfi	1.37130	293.15	Isobaric (vapour + liquid) equilibria data for the binary systems {1,2-dichloroethane (1) + toluene (2)} and {1,2-dichloroethane (1) + acetic acid (2)} at atmospheric pressure
rfi	1.37160	303.15	Phase equilibria measurements of ternary mixtures (sulfolane + a carboxylic acid + pentane) at 303.15 K
rfi	1.37160	303.15	Liquid liquid equilibria measurements of ternary systems (acetonitrile + a carboxylic acid + dodecane) at 303.15 K
rfi	1.36960	298.15	Study of liquid liquid equilibrium of the systems isobutyl acetate + acetic acid +water and isobutyl alcohol + acetic acid +water at different temperatures
rfi	1.37210	293.15	Liquid liquid equilibria of ternary systems (water + carboxylic acid + cumene) at 298.15K

rfi	1.37200		293.00	Quaternary Liquid-Liquid Equilibrium of Water + Acetic Acid + Propionic Acid + Solvent (Amyl Alcohol, Cyclohexyl Acetate, or Toluene) Systems	
rfi	1.37170		293.15	Measurement and Modeling of Liquid Liquid Equilibrium for the Systems Vinyl Acetate + Acetic Acid/Ethanol + Water at 298.15 and 308.15 K	
rhol	1049.80	kg/m3	293.20	Liquid-liquid equilibrium data and thermophysical properties for ternary systems composed of water, acetic acid and different solvents	
rhol	1049.00	kg/m3	293.00	KDB	
rhol	1044.30	kg/m3	293.15	Phase equilibrium of (water + formic or acetic acid + ethyl heptanoate) ternary liquid systems at different temperatures	
rhol	1049.42	kg/m3	298.20	(Liquid + liquid) equilibria of the (water + carboxylic acid + dibasic esters mixture (DBE-2)) ternary systems	
rhol	1049.42	kg/m3	298.15	Liquid-liquid equilibria of water + acetic acid + 2-ethyl hexyl acetate ternary system	

rhol	1043.50	kg/m3	298.15	Isobaric vapor-liquid equilibrium of the binary system sec-butyl acetate + para-xylene and the quaternary system methyl acetate + para-xylene + sec-butyl acetate + acetic acid at 101.3 kPa
rhol	1044.60	kg/m3	298.15	Liquid - liquid equilibrium for the quaternary reaction system water p sec-butyl alcohol p sec-butyl acetate p acetic acid
rhol	1049.80	kg/m3	293.20	Liquid-liquid equilibrium data for ternary systems of water + acetic acid + acetate esters at 293.2 K and 303.2 K and ~ 95 kPa
rhol	1049.42	kg/m3	298.20	Liquid Phase Equilibria of the Water + Acetic Acid + Dimethyl Carbonate Ternary System at Several Temperatures
rhol	1030.00	kg/m3	313.15	Probe of Interactions of Acetic and Propionic Acids with 2',3'-N-Epoxypropyl-N-methyl-2-oxopyrrolidinium Salicylate Ionic Liquid
rhol	1035.80	kg/m3	308.15	Probe of Interactions of Acetic and Propionic Acids with 2',3'-N-Epoxypropyl-N-methyl-2-oxopyrrolidinium Salicylate Ionic Liquid
rhol	1041.40	kg/m3	303.15	Probe of Interactions of Acetic and Propionic Acids with 2',3'-N-Epoxypropyl-N-methyl-2-oxopyrrolidinium Salicylate Ionic Liquid

rhol	1046.90	kg/m3	298.15	Probe of Interactions of Acetic and Propionic Acids with 2',3'-N-Epoxypropyl-N-methyl-2-oxopyrrolidinium Salicylate Ionic Liquid
rhol	1052.50	kg/m3	293.15	Probe of Interactions of Acetic and Propionic Acids with 2',3'-N-Epoxypropyl-N-methyl-2-oxopyrrolidinium Salicylate Ionic Liquid
rhol	1048.93	kg/m3	293.15	Liquid-Liquid Phase Equilibria for Quinary, Quaternary, and Ternary Systems {Water + Furfural + Acetic Acid + Cyclopentyl Methyl Ether + CaCl2}: Measurement, Effect of Salt ,and Comparative Study
rhol	1044.12	kg/m3	298.15	Liquid Phase Equilibria of Aqueous Mixtures of Carboxylic Acids (C1-C4) with Ethylbenzene: Thermodynamic and Mathematical Modeling
rhol	993.00	kg/m3	343.15	Effect of temperature on intermolecular interactions between the organic solvents: Insights from density and excess volume
rhol	1004.00	kg/m3	333.15	Effect of temperature on intermolecular interactions between the organic solvents: Insights from density and excess volume

rhol	1015.00	kg/m3	323.15	Effect of temperature on intermolecular interactions between the organic solvents: Insights from density and excess volume	
rhol	1025.00	kg/m3	313.15	Effect of temperature on intermolecular interactions between the organic solvents: Insights from density and excess volume	
rhol	1028.13	kg/m3	313.20	Liquid-liquid equilibrium data for ternary systems of water + acetic acid + acetate esters at 293.2 K and 303.2 K and ~ 95 kPa	
rhol	1042.00	kg/m3	298.15	Effect of temperature on intermolecular interactions between the organic solvents: Insights from density and excess volume	
rhol	1052.00	kg/m3	293.15	Effect of temperature on intermolecular interactions between the organic solvents: Insights from density and excess volume	
rhol	1053.00	kg/m3	288.15	Effect of temperature on intermolecular interactions between the organic solvents: Insights from density and excess volume	
rhol	1026.72	kg/m3	313.20	The impact of uni-univalent electrolytes on (water + acetic acid + toluene) equilibria: Representation with electrolyte-NRTL model	

rhol	1043.67	kg/m3	298.20	The impact of uni-univalent electrolytes on (water + acetic acid + toluene) equilibria: Representation with electrolyte-NRTL model	
rhol	1055.01	kg/m3	288.20	The impact of uni-univalent electrolytes on (water + acetic acid + toluene) equilibria: Representation with electrolyte-NRTL model	
rhol	1028.13	kg/m3	313.20	Liquid-liquid equilibrium data and thermophysical properties for ternary systems composed of water, acetic acid and different solvents	
rhol	1037.00	kg/m3	303.15	Effect of temperature on intermolecular interactions between the organic solvents: Insights from density and excess volume	
sfust	38.36	J/mol×K	290.06	NIST Webbook	
sfust	40.47	J/mol×K	289.90	NIST Webbook	
srf	0.03	N/m	303.20	KDB	
srf	0.03	N/m	308.15	Surface Tension of o-Xylene + Acetic Acid and m-Xylene + Acetic Acid Binary Mixtures from 303.15 K to 343.15 K	
srf	0.03	N/m	313.15	Surface Tension of o-Xylene + Acetic Acid and m-Xylene + Acetic Acid Binary Mixtures from 303.15 K to 343.15 K	

srf	0.02	N/m	318.15	Surface Tension of o-Xylene + Acetic Acid and m-Xylene + Acetic Acid Binary Mixtures from 303.15 K to 343.15 K
srf	0.02	N/m	323.15	Surface Tension of o-Xylene + Acetic Acid and m-Xylene + Acetic Acid Binary Mixtures from 303.15 K to 343.15 K
srf	0.02	N/m	333.15	Surface Tension of o-Xylene + Acetic Acid and m-Xylene + Acetic Acid Binary Mixtures from 303.15 K to 343.15 K
srf	0.03	N/m	303.15	Surface Tension of o-Xylene + Acetic Acid and m-Xylene + Acetic Acid Binary Mixtures from 303.15 K to 343.15 K
srf	0.02	N/m	343.15	Surface Tension of o-Xylene + Acetic Acid and m-Xylene + Acetic Acid Binary Mixtures from 303.15 K to 343.15 K

Pressure Dependent Properties

Property code	Value	Unit	Pressure [kPa]	Source
tbp	341.30	K	17.00	Study of Vapor-Liquid Equilibria for Acetic Acid + n-Propyl Acetate + Isopropyl Acetate Systems

tbp	342.87	K	18.47	Study of Vapor-Liquid Equilibria for Acetic Acid + n-Propyl Acetate + Isopropyl Acetate Systems
tbp	345.03	K	20.01	Study of Vapor-Liquid Equilibria for Acetic Acid + n-Propyl Acetate + Isopropyl Acetate Systems
tbp	346.41	K	21.29	Study of Vapor-Liquid Equilibria for Acetic Acid + n-Propyl Acetate + Isopropyl Acetate Systems
tbp	348.49	K	23.18	Study of Vapor-Liquid Equilibria for Acetic Acid + n-Propyl Acetate + Isopropyl Acetate Systems
tbp	350.35	K	24.92	Study of Vapor-Liquid Equilibria for Acetic Acid + n-Propyl Acetate + Isopropyl Acetate Systems
tbp	352.88	K	27.44	Study of Vapor-Liquid Equilibria for Acetic Acid + n-Propyl Acetate + Isopropyl Acetate Systems
tbp	355.01	K	29.78	Study of Vapor-Liquid Equilibria for Acetic Acid + n-Propyl Acetate + Isopropyl Acetate Systems
tbp	357.54	K	32.78	Study of Vapor-Liquid Equilibria for Acetic Acid + n-Propyl Acetate + Isopropyl Acetate Systems
tbp	359.26	K	34.97	Study of Vapor-Liquid Equilibria for Acetic Acid + n-Propyl Acetate + Isopropyl Acetate Systems

tbp	363.04	K	39.99	Study of Vapor-Liquid Equilibria for Acetic Acid + n-Propyl Acetate + Isopropyl Acetate Systems	
tbp	368.56	K	48.99	Study of Vapor-Liquid Equilibria for Acetic Acid + n-Propyl Acetate + Isopropyl Acetate Systems	
tbp	372.27	K	55.76	Study of Vapor-Liquid Equilibria for Acetic Acid + n-Propyl Acetate + Isopropyl Acetate Systems	
tbp	374.54	K	60.04	Study of Vapor-Liquid Equilibria for Acetic Acid + n-Propyl Acetate + Isopropyl Acetate Systems	
tbp	376.66	K	64.44	Study of Vapor-Liquid Equilibria for Acetic Acid + n-Propyl Acetate + Isopropyl Acetate Systems	
tbp	379.05	K	69.39	Study of Vapor-Liquid Equilibria for Acetic Acid + n-Propyl Acetate + Isopropyl Acetate Systems	
tbp	381.60	K	75.23	Study of Vapor-Liquid Equilibria for Acetic Acid + n-Propyl Acetate + Isopropyl Acetate Systems	
tbp	383.32	K	79.98	Study of Vapor-Liquid Equilibria for Acetic Acid + n-Propyl Acetate + Isopropyl Acetate Systems	

tbp	384.82	K	83.66	Study of Vapor-Liquid Equilibria for Acetic Acid + n-Propyl Acetate + Isopropyl Acetate Systems
tbp	386.07	K	87.18	Study of Vapor-Liquid Equilibria for Acetic Acid + n-Propyl Acetate + Isopropyl Acetate Systems
tbp	389.58	K	97.08	Study of Vapor-Liquid Equilibria for Acetic Acid + n-Propyl Acetate + Isopropyl Acetate Systems

Correlations

Information	Value
Property code	pvap
Equation	ln(Pvp) = A + B/(T + C)
Coeff. A	1.47385e+01
Coeff. B	-3.05591e+03
Coeff. C	-8.92370e+01
Temperature range (K), min.	289.81
Temperature range (K), max.	591.95

Information	Value
Property code	pvap
Equation	ln(Pvp) = A + B/T + C*ln(T) + D*T^2
Coeff. A	6.82294e+01
Coeff. B	-7.02972e+03
Coeff. C	-7.79853e+00
Coeff. D	5.93100e-06
Temperature range (K), min.	289.81
Temperature range (K), max.	592.71

Datasets

Viscosity, Pa*s

Temperature, K - Liquid	Pressure, kPa - Liquid	Viscosity, Pa*s - Liquid
313.15	100.00	0.0009030
313.15	800.00	0.0009100
313.15	1600.00	0.0009170
313.15	2400.00	0.0009240
313.15	3200.00	0.0009320
333.15	100.00	0.0007130
333.15	800.00	0.0007200
333.15	1600.00	0.0007270
333.15	2400.00	0.0007350
333.15	3200.00	0.0007430
353.15	100.00	0.0005740
353.15	800.00	0.0005820
353.15	1600.00	0.0005900
353.15	2400.00	0.0005980
353.15	3200.00	0.0006050
373.15	100.00	0.0004710
373.15	800.00	0.0004780
373.15	1600.00	0.0004860
373.15	2400.00	0.0004940
373.15	3200.00	0.0005020
393.15	800.00	0.0003940
393.15	1600.00	0.0004020
393.15	2400.00	0.0004100
393.15	3200.00	0.0004180
413.15	800.00	0.0003330
413.15	1600.00	0.0003400
413.15	2400.00	0.0003470
413.15	3200.00	0.0003540
433.15	800.00	0.0002810
433.15	1600.00	0.0002870
433.15	2400.00	0.0002940
433.15	3200.00	0.0003020
453.15	800.00	0.0002440
453.15	1600.00	0.0002510
453.15	2400.00	0.0002580

453.15	3200.00	0.0002640
473.15	1600.00	0.0002130
473.15	2400.00	0.0002200
473.15	3200.00	0.0002260

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Legend

af:	Acentric Factor
affp:	Proton affinity
aigt:	Autoignition Temperature
basg:	Gas basicity
chl:	Standard liquid enthalpy of combustion
cpg:	Ideal gas heat capacity
cpl:	Liquid phase heat capacity

dm:	Dipole Moment
dvisc:	Dynamic viscosity
fll:	Lower Flammability Limit
flu:	Upper Flammability Limit
fpc:	Flash Point (Closed Cup Method)
fpo:	Flash Point (Open Cup Method)
gf:	Standard Gibbs free energy of formation
gyrad:	Radius of Gyration
hf:	Enthalpy of formation at standard conditions
hfl:	Liquid phase enthalpy of formation at standard conditions
hfus:	Enthalpy of fusion at standard conditions
hfust:	Enthalpy of fusion at a given temperature
hsubt:	Enthalpy of sublimation at a given temperature
hvap:	Enthalpy of vaporization at standard conditions
hvapt:	Enthalpy of vaporization at a given temperature
ie:	Ionization energy
log10ws:	Log10 of Water solubility in mol/l
logp:	Octanol/Water partition coefficient
mcvol:	McGowan's characteristic volume
nfpaf:	NFPA Fire Rating
nfpah:	NFPA Health Rating
nfpas:	NFPA Safety Rating
pc:	Critical Pressure
pt:	Triple Point Pressure
pvap:	Vapor pressure
rfi:	Refractive Index
rhoc:	Critical density
rhol:	Liquid Density
rinpol:	Non-polar retention indices
ripol:	Polar retention indices
sfust:	Entropy of fusion at a given temperature
sg:	Molar entropy at standard conditions
sl:	Liquid phase molar entropy at standard conditions
srf:	Surface Tension
tb:	Normal Boiling Point Temperature
tbp:	Boiling point at given pressure
tc:	Critical Temperature
tf:	Normal melting (fusion) point
tt:	Triple Point Temperature
vc:	Critical Volume
zc:	Critical Compressibility
zra:	Rackett Parameter