

Calcium sulfate, anhydrous

Other names:	calcium sulfate calcium sulphate crysalba sulfuric acid calcium salt (1:1) sulfuric acid, calcium salt (1:1) thiolite
Inchi:	InChI=1S/Ca.H2O4S/c;1-5(2,3)4/h;(H2,1,2,3,4)/q+2;p-2
InchiKey:	OSGAYBCDTRGGQ-UHFFFAOYSA-L
Formula:	CaO4S
SMILES:	O=S1(=O)O[Ca]O1
Mol. weight [g/mol]:	136.14
CAS:	7778-18-9

Sources

Thermodynamic Modeling of Calcium Sulfate Hydrates in the CaSO₄-H₂O System
Contribution to the quaternary system H₂O-NaCl-Ca²⁺/Mg²⁺-K⁺-SO₄²⁻: Solid phases equilibria of Gypsum Hemihydrate and Anhydrite at the temperature range 0-120 °C and 0.001-0.01 MPa
Solubility Behavior and Morphology of Inorganic Salts: Difference between Gypsum in the Quaternary Binary System and in the Ternary System from Retry: Quaternary Region of the System Ca²⁺, Mg²⁺//NaCl-H₂O at 298.15 K:
Quaternary System Ca²⁺, Mg²⁺//SO₃²⁻, F⁻ at 298.15 K
Effect of Lipophilic Alcohols on Solubility, Physicochemical Properties, and Microstructure of Calcium Hydroxide Found in Natural Water A from (5 to 90) g/L
Solubility Behavior of Calcium Sulfate Dihydrate (Gypsum) in an Aqueous Sodium Chloride System in the Presence of Hydroxyalkyl Ammonium Acetate Ionic Liquids Additives: Morphology Changes and Physicochemical Solution Properties at 35 deg.C.

<https://www.doi.org/10.1021/acs.jced.9b00112>

<https://www.doi.org/10.1016/j.fluid.2015.10.038>

<https://www.doi.org/10.1021/acs.jced.5b00454>

<https://www.doi.org/10.1021/acs.jced.8b00093>

<https://www.doi.org/10.1016/j.tca.2013.10.024>

<https://www.doi.org/10.1021/acs.jced.9b00051>

<https://www.doi.org/10.1021/acs.jced.8b00385>

<https://www.doi.org/10.1021/acs.jced.8b00723>

<https://www.doi.org/10.1021/je101012n>

<https://www.doi.org/10.1021/acs.jced.9b00354>

<http://webbook.nist.gov/cgi/cbook.cgi?ID=C7778189&Units=SI>

Full-text version available from:

<https://www.chemeo.com/cid/53-766-3/Calcium-sulfate-anhydrous.pdf>

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