

# m-Fluorobenzamide

<b>Other names:</b>	3-fluorobenzamide Benzamide, 3-fluoro- Benzamide, m-fluoro-
<b>Inchi:</b>	InChI=1S/C7H6FNO/c8-6-3-1-2-5(4-6)7(9)10/h1-4H,(H2,9,10)
<b>InchiKey:</b>	YPIGHNIIXYSPKF-UHFFFAOYSA-N
<b>Formula:</b>	C7H6FNO
<b>SMILES:</b>	NC(=O)c1ccccc(F)c1
<b>Mol. weight [g/mol]:</b>	139.13
<b>CAS:</b>	455-37-8

## Physical Properties

Property code	Value	Unit	Source
affp	877.20	kJ/mol	NIST Webbook
basg	846.30	kJ/mol	NIST Webbook
gf	-146.44	kJ/mol	Joback Method
hf	-237.65	kJ/mol	Joback Method
hfus	21.50	kJ/mol	Thermodynamic Study of the Three Fluorobenzamides: Vapor Pressures, Phase Diagrams, and Hydrogen Bonds
hvap	50.68	kJ/mol	Joback Method
ie	9.60	eV	NIST Webbook
log10ws	-1.97		Crippen Method
logp	0.925		Crippen Method
mcvol	99.050	ml/mol	McGowan Method
pc	4450.38	kPa	Joback Method
tb	516.89	K	Joback Method
tc	742.75	K	Joback Method
tf	341.37	K	Joback Method
vc	0.372	m3/kmol	Joback Method

## Temperature Dependent Properties

Property code	Value	Unit	Temperature [K]	Source
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cpg	208.94	J/mol×K	516.89	Joback Method
cpg	218.62	J/mol×K	554.53	Joback Method
cpg	227.65	J/mol×K	592.18	Joback Method
cpg	236.05	J/mol×K	629.82	Joback Method
cpg	243.86	J/mol×K	667.47	Joback Method
cpg	251.10	J/mol×K	705.11	Joback Method
cpg	257.80	J/mol×K	742.75	Joback Method

## Sources

<b>NIST Webbook:</b>	<a href="http://webbook.nist.gov/cgi/cbook.cgi?ID=C455378&amp;Units=SI">http://webbook.nist.gov/cgi/cbook.cgi?ID=C455378&amp;Units=SI</a>
<b>Crippen Method:</b>	<a href="http://pubs.acs.org/doi/abs/10.1021/ci9903071">http://pubs.acs.org/doi/abs/10.1021/ci9903071</a>
<b>Crippen Method:</b>	<a href="https://www.chemeo.com/doc/models/crippen_log10ws">https://www.chemeo.com/doc/models/crippen_log10ws</a>
<b>Thermodynamic Study of the Three Fluorobenzamides: Vapor Pressures, Phase Diagrams, and Hydrogen Bonds:</b>	<a href="https://www.doi.org/10.1021/je100801s">https://www.doi.org/10.1021/je100801s</a>
<b>McGowan Method:</b>	<a href="https://en.wikipedia.org/wiki/Joback_method">https://en.wikipedia.org/wiki/Joback_method</a>
<b>McGowan Method:</b>	<a href="http://link.springer.com/article/10.1007/BF02311772">http://link.springer.com/article/10.1007/BF02311772</a>

## Legend

<b>affp:</b>	Proton affinity
<b>basg:</b>	Gas basicity
<b>cpg:</b>	Ideal gas heat capacity
<b>gf:</b>	Standard Gibbs free energy of formation
<b>hf:</b>	Enthalpy of formation at standard conditions
<b>hfus:</b>	Enthalpy of fusion at standard conditions
<b>hvap:</b>	Enthalpy of vaporization at standard conditions
<b>ie:</b>	Ionization energy
<b>log10ws:</b>	Log10 of Water solubility in mol/l
<b>logp:</b>	Octanol/Water partition coefficient
<b>mcvol:</b>	McGowan's characteristic volume
<b>pc:</b>	Critical Pressure
<b>tb:</b>	Normal Boiling Point Temperature
<b>tc:</b>	Critical Temperature
<b>tf:</b>	Normal melting (fusion) point
<b>vc:</b>	Critical Volume

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