

# tetra-ethoxylated trimethylol propane triacrylate (Acrylic acid

**2-[2-[2-(2-acryloyloxy-ethoxy)-ethoxy]-ethoxymethyl ester)**

**SMILES:** C=CC(=O)OCCOCCOCC(CC)(COCCOC(=O)C=C)COCCOC(=O)C=C

**Mol. weight [g/mol]:** 472.53

## Physical Properties

Property code	Value	Unit	Source
gf	-712.62	kJ/mol	Joback Method
hf	-1413.79	kJ/mol	Joback Method
hfus	57.18	kJ/mol	Joback Method
hvap	100.59	kJ/mol	Joback Method
log10ws	-1.71		Crippen Method
logp	1.637		Crippen Method
mcvol	367.830	ml/mol	McGowan Method
pc	969.88	kPa	Joback Method
rinpol	2852.00		NIST Webbook
rinpol	2852.00		NIST Webbook
tb	1031.00	K	Joback Method
tc	1269.70	K	Joback Method
tf	651.51	K	Joback Method
vc	1.399	m <sup>3</sup> /kmol	Joback Method

## Temperature Dependent Properties

Property code	Value	Unit	Temperature [K]	Source
cpg	1229.71	J/molxK	1031.00	Joback Method
cpg	1269.54	J/molxK	1229.92	Joback Method
cpg	1266.12	J/molxK	1190.13	Joback Method
cpg	1260.42	J/molxK	1150.35	Joback Method
cpg	1252.44	J/molxK	1110.57	Joback Method
cpg	1242.20	J/molxK	1070.78	Joback Method
cpg	1270.68	J/molxK	1269.70	Joback Method
dvisc	0.0000065	Paxs	1031.00	Joback Method

dvisc	0.0000084	Paxs	967.75	Joback Method
dvisc	0.0000113	Paxs	904.50	Joback Method
dvisc	0.0000159	Paxs	841.26	Joback Method
dvisc	0.0000237	Paxs	778.01	Joback Method
dvisc	0.0000378	Paxs	714.76	Joback Method
dvisc	0.0000661	Paxs	651.51	Joback Method

## Sources

<b>NIST Webbook:</b>	<a href="http://webbook.nist.gov/cgi/cbook.cgi?ID=R508560&amp;Units=SI">http://webbook.nist.gov/cgi/cbook.cgi?ID=R508560&amp;Units=SI</a>
<b>Crippen Method:</b>	<a href="http://pubs.acs.org/doi/abs/10.1021/ci9903071">http://pubs.acs.org/doi/abs/10.1021/ci9903071</a>
<b>Crippen Method:</b>	<a href="https://www.chemeo.com/doc/models/crippen_log10ws">https://www.chemeo.com/doc/models/crippen_log10ws</a>
<b>Joback Method:</b>	<a href="https://en.wikipedia.org/wiki/Joback_method">https://en.wikipedia.org/wiki/Joback_method</a>
<b>McGowan Method:</b>	<a href="http://link.springer.com/article/10.1007/BF02311772">http://link.springer.com/article/10.1007/BF02311772</a>

## Legend

<b>cpg:</b>	Ideal gas heat capacity
<b>dvisc:</b>	Dynamic viscosity
<b>gf:</b>	Standard Gibbs free energy of formation
<b>hf:</b>	Enthalpy of formation at standard conditions
<b>hfus:</b>	Enthalpy of fusion at standard conditions
<b>hvap:</b>	Enthalpy of vaporization at standard conditions
<b>log10ws:</b>	Log10 of Water solubility in mol/l
<b>logp:</b>	Octanol/Water partition coefficient
<b>mcvol:</b>	McGowan's characteristic volume
<b>pc:</b>	Critical Pressure
<b>rinpol:</b>	Non-polar retention indices
<b>tb:</b>	Normal Boiling Point Temperature
<b>tc:</b>	Critical Temperature
<b>tf:</b>	Normal melting (fusion) point
<b>vc:</b>	Critical Volume

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