**Ethyl acetate**

**Other names:** 1-Acetoxyethane; Acetic acid, ethyl ester; Acetic ether; Acetidin; Acetoxyethane; Aethylacetat; CH3COOC2H5; Essigester; Ethyl acetic ester; Ethyl ester of acetic acid; Ethyl ethanoate; Ethylacetaat; Ethyle (acetate d'); Ethylester kyseliny octove; Etile (acetato di); NSC 70930; Rcra waste number U112; UN 1173; Vinegar naphtha; ac. acetic ethyl ester.

**InChI:** InChI=1S/C4H8O2/c1-3-6-4(2)5/h3H2,1-2H3

**InChI Key:** XEKOWRVHYACXOJ-UHFFFAOYSA-N

**Formula:** C4H8O2

**SMILES:** CCOC(C)=O

**Molecular Weight:** 88.11

**CAS:** 141-78-6

### Physical Properties

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<td>K</td>
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<tr>
<td>$T_{triple}$</td>
<td>189.30 ± 0.05</td>
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<td>$T_{triple}$</td>
<td>189.30 ± 0.20</td>
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<tr>
<td>$V_c$</td>
<td>0.28</td>
<td>m³/kg-mol</td>
<td>Joback Method</td>
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**Temperature Dependent Properties**

<table>
<thead>
<tr>
<th>Property</th>
<th>Value</th>
<th>Unit</th>
<th>Temperature (K)</th>
<th>Source</th>
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</thead>
<tbody>
<tr>
<td>$C_{p,gas}$</td>
<td>125.82</td>
<td>J/mol×K</td>
<td>360.0</td>
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<tr>
<td>$C_{p,gas}$</td>
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<tr>
<td>$C_{p,gas}$</td>
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<td>$C_{p,liquid}$</td>
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<td>J/mol×K</td>
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<td>$C_{p,liquid}$</td>
<td>168.94</td>
<td>J/mol×K</td>
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<td>NIST Webbook</td>
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<tr>
<td>$C_{p,liquid}$</td>
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<td>J/mol×K</td>
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<tr>
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<td>kJ/mol</td>
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<td>NIST Webbook</td>
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<tr>
<td>Property</td>
<td>Value</td>
<td>Unit</td>
<td>Temperature (K)</td>
<td>Source</td>
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<td>$\Delta_{\text{fus}}H$</td>
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<td>kJ/mol</td>
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<td>$\Delta_{\text{fus}}S$</td>
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<td>NIST Webbook</td>
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</tbody>
</table>

Sources


**NIST Webbook:** [http://webbook.nist.gov/cgi/inchi/InChI=1S/C4H8O2/c1-3-6-4(2)5/h3H2,1-2H3](http://webbook.nist.gov/cgi/inchi/InChI=1S/C4H8O2/c1-3-6-4(2)5/h3H2,1-2H3)

**Crippen Method:** [http://pubs.acs.org/doi/abs/10.1021/ci990307l](http://pubs.acs.org/doi/abs/10.1021/ci990307l)

Legend

**PAff:** Proton affinity (kJ/mol).

**BasG:** Gas basicity (kJ/mol).

$\Delta_{\text{c}}H^o_{\text{liquid}}$: Standard liquid enthalpy of combustion (kJ/mol).

$C_{p,\text{gas}}$: Ideal gas heat capacity (J/mol×K).

$C_{p,\text{liquid}}$: Liquid phase heat capacity (J/mol×K).

$\eta$: Dynamic viscosity (Pa×s).

$\Delta_iG^o$: Standard Gibbs free energy of formation (kJ/mol).
\( \Delta H^\circ_{\text{gas}} \): Enthalpy of formation at standard conditions (kJ/mol).

\( \Delta f^\circ_{\text{liquid}} \): Liquid phase enthalpy of formation at standard conditions (kJ/mol).

\( \Delta f_{\text{fus}}^\circ \): Enthalpy of fusion at standard conditions (kJ/mol).

\( \Delta f_{\text{fus}}^\circ \): Enthalpy of fusion at a given temperature (kJ/mol).

\( \Delta v_{\text{fus}}^\circ \): Enthalpy of vaporization at standard conditions (kJ/mol).

\( \Delta v_{\text{fus}}^\circ \): Enthalpy of vaporization at a given temperature (kJ/mol).

\( \text{IE} \): Ionization energy (eV).

\( \log P_{\text{oct/wat}} \): Octanol/Water partition coefficient.

\( P_c \): Critical Pressure (kPa).

\( \Delta f_{\text{fus}} S \): Entropy of fusion at a given temperature (J/mol×K).

\( S^\circ_{\text{gas}} \): Molar entropy at standard conditions (J/mol×K).

\( S^\circ_{\text{liquid}} \): Liquid phase molar entropy at standard conditions (J/mol×K).

\( T_{\text{boil}} \): Normal Boiling Point Temperature (K).

\( T_c \): Critical Temperature (K).

\( T_{\text{fus}} \): Normal melting (fusion) point (K).

\( T_{\text{fus}} \): Triple Point Temperature (K).

\( V_c \): Critical Volume (m\(^3\)/kg-mol).

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